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U. S. Geological Survey.
Monograph.
No. 12. Pt. 2. 1886

GEOLOGY AND MINING INDUSTRY OF LEADVILLE.

PART II.

MINING INDUSTRY.

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PART II.

MINING INDUSTRY.

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CHAPTER I.

ORE DEPOSITS.

The preceding chapters have been devoted almost exclusively to the consideration of the geological structure of the district. This subject has been treated at considerable length, not only because it presents many facts which seemed of sufficient interest to geologists in general to justify such treatment, but also because a thorough knowledge of the geological structure of a region is an essential and indispensable basis for the study of its ore deposits; a fact which is too often lost sight of by those practically engaged in mining. For a time the miner may develop his mine successfully by simply following the ore lead, guided by the empirical rules which experience has taught him, and without regard to the geological phenomena presented by the country rocks, their structural conditions, or the probable origin and manner of formation of the deposits; but the time is sure to come when without this knowledge he will be liable to make mistakes which may cost him more than he has gained by all his previous labors.

N.B.

Before proceeding to a detailed description of the various ore deposits of the region studied in the course of this investigation, it may aid the reader to have a brief résumé of their principal characteristics and a concise statement of the conclusions which have been arrived at with regard to their origin and manner of formation.

CLASSIFICATION.

To a scientific description of natural objects the most valuable aid is a rational and universally accepted system of classification. The first obstacle one encounters in attempting the description of ore deposits is the absence of such a classification. The object of a system of classification is not only to afford a means of avoiding long and repeated circumlocutions in descriptions, but also to furnish a comprehensive view of the mutual relations of the classes of phenomena to which it is applied. Such systems must necessarily change from time to time as the scientific studies of the phenomena progress and knowledge with regard to them becomes more accurate and thorough. The unsatisfactory state of existing classifications of ore deposits is due in large degree to an imperfect knowledge of the subject on the part of those who have made them, but in part also to their being made from a false standpoint.

As the study of geology sprang originally from the empirical observations of those engaged in mining for the useful metals, so the first systems of classification of ore deposits were based on distinctions and characteristics established by the miners themselves in their daily work, and, as in carrying on this work the outward form of the deposit was the most essential characteristic, this naturally formed the basis of their classifications. But while general geology has made relatively more rapid progress than the study of ore deposits, which, being a matter of practical and economic importance, has seemed to many to belong to a lower sphere of scientific investigation than purely theoretical questions, the prevalent classifications still hold largely to the original basis of the practical miner. The form of a deposit might well constitute the basis of a classification, if it constituted an essential characteristic thereof, and if there were certain regular forms that belonged exclusively to particular classes of deposits, which had a necessary connection with the sum of their other characteristics. This is so far from being the case, however, that not only is no one form confined to any particular class of deposit, but the same class of deposit, that is, one which has undoubtedly the same origin and manner of formation, may have a great variety of different forms, as is the case with those about to be described.

That the scientific study of ore deposits has not kept pace with the advance in other branches of geology is due in great part no doubt to the inherent difficulty of the subject, but also in a measure to a want of scientific zeal or knowledge on the part of those who are practically engaged in mining. The phenomena to be investigated must be studied in the underground workings of mines, in which not only is a very small area open to observation as compared with the surface phenomena on which other geological reasonings are mainly based, but they are not in their nature as permanent as are the latter and soon become obscured by decay or entirely inaccessible. But, while the attainable facts are thus relatively meager, they have not all been made available to the student, for the reason that those practically engaged in mining are too often content with noting those alone which have an immediate practical bearing, and have neglected to put on record those of merely theoretical interest, which, nevertheless, if carefully observed, might afford a basis for scientific generalizations of great economic importance.

We can only hope to arrive at a satisfactory and rational classification, which shall be founded essentially on genetic principles, when our knowledge of ore deposits shall be vastly increased by the accumulation of a great number of scientific observations, based on correct geological studies, and towards this accumulation we must look to those practically conducting mines for a most essential contribution, since they alone have the opportunity of daily observation of the constantly changing phenomena which ore deposits present. Meanwhile it may be of use to review some of the more prominent systems of classification proposed by modern writers upon ore deposits, and to consider their relative applicability to the important class of deposits under consideration.

As the Germans were the first to write upon mines and ore deposits and the classifications adopted by other nations have been to a greater or less degree founded upon their work, the first place will be given to a mention of those most current in Germany at the present day. The original edition of B. von Cotta's treatise upon ore deposits appeared in 1853, and has not been essentially changed in the later edition here quoted. The next classification quoted is that of Dr. Joh. Grimm, professor of the School of Mines in Přibram, Bohemia. The third is that given in his course on mining at the School of Mines of Berlin, by Professor H. Lottner, and published by his successor, Professor A. Serlo. The last, that of Dr. A. von Groddeck, of the School of Mines at Clausthal, in the Hartz.

Von Cotta.[1]	Grimm.[2]	Serlo–Lottner.[3]	Von Groddeck.[4]
I. *Deposits of regular form.* 1. *Beds.* *a.* Beds of ore, coal, etc. *b.* Placer deposits. 2. *Veins.* *a.* Transverse or ordinary veins. *b.* Bedded veins. *c.* Contact veins. *d.* Lenticular veins. II. *Deposits of irregular form.* 1. *Stocks* (sharply defined bodies.) *a.* Stock works. *b.* Contact stocks. *c.* Cave fillings. *d.* and *e.* Pockets, kidney-shaped deposits (Butzen, Rachelen, Taschen, Nester, Rinner, Nieren).* 2. *Impregnations* (bodies not sharply defined). *a.* Independent impregnations. *b.* Dependent impregnations (connected with other deposits). * Untranslatable miner's terms.	I. *Disseminations or impregnations* (deposits forming an essential constituent of the country rock). 1. *Original impregnations.* 2. *Secondary impregnations.* II. *Distinct ore deposits* (forming an accessory constituent of the country rock). 1. *Sheet* (regular-shaped) *masses.* *a.* Bedded (sedimentary) deposits. *b.* Veins; crevice deposits; stringers (filling open fissures). *c.* Sheet-shaped segregations. 2. *Stocks and irregularly shaped deposits.* *a.* Bedded (sedimentary) deposits. *b.* Stocks (Butzen, Nester, etc.), filling pre-existing cavities. *c.* Stock works (reticulated veins).	I. *Inclosed or underground deposits.* 1. *Sheet* (regular-shaped) *deposits.* *a.* Veins. *b.* Beds. 2. *Mass* (irregular-shaped) *deposits.* *a.* Stocks. *b.* Stock works. 3. *Other irregularly-shaped deposits* (pockets, kidneys, &c.). II. *Superficial deposits.* 4. *Deposits of débris* (placers). 5. *Surface deposits in place* (bog-ore, &c.).	I. *Original or primary deposits.* A. *Contemporaneous with country rock.* 1. Deposits in stratified rocks. 2. Deposits in eruptive rocks. B. *Later than country rock.* 3. Deposits filling pre-existing cavities. *a.* Veins or lodes. *b.* Cave fillings. 4. Metamorphic (or metasomatic) deposits. II. *Secondary or detrital deposits.*

Von Cotta's classification is founded exclusively on the form of the deposit and recognizes no genetic principle as a basis of classification. Thus, such essentially opposed deposits as coal beds and placer deposits, on the one hand, and mineral veins and contact deposits, on the other, are put under one general heading; while cave-fillings, pockets, etc., which may be merely offshoots from a vein or contact deposit, come under a distinct main head.

Grimm's classification is also mainly founded on the outward form of the deposit, but he admits a few minor genetical distinctions, such as sep-

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[1] Die Lehre von den Erzlagerstätten. Freiberg, 1859.

[2] Die Lagerstätten der nutzbaren Mineralien. Prag, 1869.

[3] Bergbaukunde. Berlin, 1878.

[4] Die Lehre von den Lagerstätten der Erze. Leipzig, 1879.

arating bedded deposits of sedimentary origin from those which were formed later than the inclosing rocks. Lottner also bases his classification on outward form alone, but distinguishes secondary from original deposits. Von Groddeck lays much more stress on genetic distinctions, and not only brings in each of those recognized by the two previously named, but admits the existence of ore deposits of later formation than the country rock which do not necessarily fill pre-existing cavities or fissures.

F. Pošepný,[1] professor at Přibram, who has made an extensive study of ore deposits, including many of those of the United States, proposes an even more radically genetic subdivision of metalliferous deposits into (1) deposits in pre-existing cavities and (2) those formed by gradual replacement of the rock substances by the vein material or mineral, the first class being further subdivided into those filling cavities formed in a mechanical way, or *dislocation spaces*, and those formed by corrosive action in soluble rock, or *corrosive spaces*, which would correspond in general, though not necessarily in all cases, to the distinctions of Grimm and Groddeck of the fissure-fillings and cave-fillings.

In order that a classification should find general acceptation among mining men, it is essential, moreover, that it should be simple, concise, and of easy comprehension, qualifications which the first two of the above systems certainly do not possess. Thus, in this country, where mining geology has found its principal discussion in courts of law, in which Prime's translation of von Cotta has been generally accepted as authority, ore deposits of primary origin (leaving placers out of consideration) are practically divided into true fissure veins and deposits which are not true fissure veins, the latter class being somewhat loosely subdivided into contact deposits, blanket deposits, and rake, pipe, and gash veins.

The term "blanket deposit" is probably derived from the *manta* of the Spanish miners, a term which in Mexico and South America designates the richest and most productive ore bodies, but in the United States is apt to be applied in rather a derogatory sense to any horizontal sheet of ore. The last terms are derived from local usage in the lead regions of the north of

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[1] Archiv für praktische Geologie, p. 600. Wien, 1880.

England, and their general application is of very doubtful advisability, since authorities differ as to their exact definition. The term "gash vein" is the only one recognized in the classifications given below, and is there applied to a fissure which is confined to a particular rock or bed and which does not extend into the adjoining rocks.

The English literature of ore deposits is even more meager than the German. Of general treatises on this subject, the more prominent in this country are J. D. Whitney's Metallic Wealth of the United States, published in 1854; an article by R. W. Raymond, in his Mining Statistics for 1869; and an admirable but little known paper on ore deposits, in Johnson's Cyclopædia, by R. Pumpelly. J. S. Newberry has also published an article on the origin and classification of ore deposits in the School of Mines Quarterly for March, 1880. In England, J. Arthur Phillips published in 1884 an extended treatise on ore deposits. Of the classifications proposed by the above authors, those of Newberry and Phillips are nearly identical with that of Whitney and Raymond's is avowedly an adaptation of Lottner, the differences in either case being unessential for the purposes of the present discussion. Those of Whitney and Pumpelly alone are therefore given here, and to them is added that given by A. Geikie in his Text Book on Geology (London, 1882), mainly because of the different standpoint from which it is made.[1]

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[1] Prof. Joseph Le Conte has also published an article on the Genesis of Ore Deposits, in the American Journal of Science for July, 1883, in which a subdivision into (1) fissure veins, (2) incipient fissures, (3) brecciated veins, (4) substitution veins, (5) contact veins, (6) irregular ore deposits, is given.

J. D. Whitney.	R. Pumpelly.	A. Geikie.
I. *Superficial.* II. *Stratified.* *a.* Constituting the mass of a bed or stratified deposit. *b.* Disseminated through sedimentary beds. *c.* Originally deposited from aqueous solution, but since metamorphosed. III. *Unstratified.* Irregular: *a.* Masses of eruptive origin. *b.* Disseminated in eruptive rocks. *c.* Stock work deposits. *d.* Contact deposits. *e.* Fahlbands. Regular: *f.* Segregated veins. *g.* Gash veins. *h.* True or fissure veins.	I. *Surface deposits.* 1. Residuary deposits. 2. Stream deposits. 3. Lake and bog deposits. II. *Forms due to the texture of the inclosing rock or to its mineral constitution, or to both.* 1. *Disseminated concentrations.* *a.* Impregnations. *b.* Fahlbands. 2. *Aggregated concentrations.* *a.* Lenticular aggregations. *b.* Irregular masses (stocks) *c.* Reticulated veins (stock works). *d.* Contact deposits. III. *Forms due chiefly to pre-existing cavities or open fissures.* 1. Cave deposits. 2. Gash veins. 3. Fissure veins.	I. *Contemporaneous ores of stratified rocks.* II. *Contemporaneous ores of crystalline rocks.* III. *Subsequently introduced ores.* 1. Mineral veins or lodes. 2. Stocks and stock works (including gash veins).

All of the above are an advance upon von Cotta in that form is not in all cases the exclusive basis of classification. Whitney's first two subdivisions are distinctly genetic, but the third, which embraces the majority of metalliferous deposits, is an unsystematic grouping of a variety of forms having only one common quality, that of not being stratified. Whitney recognizes a genetic quality in his division *a*, that of being of eruptive origin, but few geologists of the present day agree with his wide application of this quality—for instance, to the great deposits of magnetic iron of Missouri and Lake Superior. In his "segregated veins" he recognizes the possibility of an unstratified deposit which is not the filling of a pre-existing cavity, while no such recognition is found in Geikie's classification. Geikie's term "subsequently introduced ores," on the other hand, is to be preferred to "unstratified deposits," as being based on a more essential characteristic of the deposit. This would involve, however, a definite statement as to the age of Whitney's Classes III, *a* and *b*, which his general term avoids.

Pumpelly's classification ignores the division of stratified or contemporaneous ore deposits, and in his text he states his belief that the greater number of ore deposits have been formed later than the inclosing rock; he also says that all metalliferous aggregations are the result of a process or series of processes of concentration.

Pošepný states his opinion on contemporaneous deposits even more strongly in the following words:[1]

> In the course of my nearly twenty-years studies of ore deposits I have yet met with no deposits (carrying sulphides) which answer to Werner's definition—that is, whose ores are contemporaneous with the country rock and which form a regular interstratified bed between other rock strata.

Like Pošepný, Pumpelly recognizes the importance of deposits which do not fill pre-existing cavities, devoting to these his subdivisions I and II. These he says fall under two heads, as regards the manner in which the space occupied by them was obtained: (1) by mechanical displacement of the inclosing material; (2) by a chemical replacement similar to that to which pseudomorphs owe their origin. His use of the form as a basis of subdivision for deposits filling pre-existing cavities seems more legitimate than in the case of those which his title seems to imply are merely concentrations of metallic minerals already existing in the rock, and the use of the word "concentration," as applied exclusively to the latter classes, seems unfortunate, as implying that the others are not concentrations also.

Geikie's classification has the merit of conciseness and his principal divisions are based on genetic principles, but his subdivisions, like those of von Cotta, recognize only differences of outward form.

In view of the difficulty, or even, in many cases, the apparent impossibility, of determining definitely the genesis of a given deposit, it may well be questioned how far it is advisable to adopt genetic relations as the basis of a classification, since it will frequently happen that an observer will be at a loss to determine under which subdivision the deposit he is studying should be placed. It seems to the writer, however, that in such a case, although his determination may not be final and may give rise to discussion and difference of opinion on the part of other observers in the same field,

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[1] Op. cit., p. 423.

he will be led by this very fact to make a more thorough and searching examination than if he were only required to define the deposit in question according to its outward form.

As regards the applicability of the foregoing classifications to the Leadville deposits, it will be seen from a perusal of the following pages that no one of the subdivisions proposed would adequately define them; either they would apply only to a limited portion of the deposits or else they would include them under the same head with deposits of an essentially different character.

Of von Cotta's, Lottner's, and Whitney's subdivisions, several would be applicable; thus, a large part of the deposits are contact deposits; other parts, however, not being at the contact of two different rocks, would be stocks when large and pockets, chambers, etc., when small. The same remark would apply to Pumpelly's subdivision of his Class II, 2. On the other hand his definition of gash veins, as filling open fissures, would not apply to those of this region. The deposits would come under only a single head of Grimm's, Geikie's, and von Groddeck's classifications. By the two former they would be classed under the general head of stocks, which really defines nothing except that they are of irregular shape and large. Finally, von Groddeck's term "metamorphic," or "metasomatic," applies to all the Leadville deposits and defines one most essential characteristic; without some modification, however, it would apply equally well to a large portion of the Rocky Mountain deposits in Archean rocks, which have been previously considered to be "true fissure veins."

LEADVILLE DEPOSITS.

Manner of occurrence.—By far the most important of the ores of Leadville and vicinity, both in quantity and in quality, occur in the blue-gray dolomitic limestone of the Lower Carboniferous formation, hence known as the Blue or ore-bearing Limestone, and at or near its contact with the overlying sheet of porphyry, which is generally the White or Leadville Porphyry. They thus constitute a sort of contact sheet, whose upper surface, being formed by the base of the porphyry sheet, is comparatively regular and well defined, while the lower surface is ill-defined and irregular, there being a gradual transition from ore into unaltered limestone, the former extending to varying depths from the surface, and even occupying at times the entire thickness of the Blue Limestone formation. This may be regarded as the typical form of the Leadville deposits; there are, however, variations from it, and also in the character of the inclosing rock, which do not necessarily involve any difference in origin or mode of formation. As variations in form, the ore sometimes occurs in irregularly-shaped bodies, or in transverse sheets not always directly connected with the upper or contact surface of the ore-bearing bed or rock; it also occurs at or near the contact of sheets of Gray or other porphyries with the Blue Limestone, and less frequently in sedimentary beds, both calcareous and silicious, and in porphyry bodies, sometimes on or near contact surfaces, sometimes along joint or fault planes.

Composition.—The prevailing and by far the most important ore, from an economical point of view, is argentiferous galena, with its secondary products, cerussite or carbonate of lead and cerargyrite or chloride of silver.

Lead is also found as anglesite or sulphate, as pyromorphite or chlorophosphate, and occasionally as oxide in the form of litharge or more rarely of minium.

Silver frequently occurs as chloro-bromide, less frequently as chloroiodide, and very rarely in the native state. Chemical investigation has failed to detect sufficient regularity in the proportions of chlorine, bromine, and iodine, combined with the silver, to justify the determination of distinct mineral species.

A frequent alteration product of mixed pyrite and galena, which occurs in considerable quantity, associated with the ore bodies, is generally called "basic ferric sulphate." It is an ocherous-looking substance of somewhat uniform outward appearance, but of varying composition, being mainly a mixture of jarosite, or yellow vitriol, and hydrated basic ferric sulphate, with more or less anglesite and pyromorphite.

Gold occurs in the native state, generally in extremely small flakes or leaflets. It is also said to have been found in the filiform state in galena.

As accessory minerals are:

Zinc blende and silicate of zinc or calamine.

Arsenic, probably as sulphide, and as arseniate of iron.

Antimony, probably as sulphide.

Molybdenum, in the form of molybdate of lead or wulfenite.

Copper, as carbonate or silicate.

Bismuth, as sulphide and its secondary product, a sulpho-carbonate.

Vanadium, as dechenite or the vanadate of lead and zinc.

Tin, indium, and cadmium have been detected in furnace products.

Iron occurs as an ore, though in the Leadville deposits in general it constitutes an essential part of the gangue or matrix in which the valuable ore is found. In the former case it occurs in considerable bodies as pyrite or sulphide and as anhydrous oxide or red hematite, with a little magnetite.

Gangue.—The other components of the ore deposits, which may be considered as gangue, although this term is perhaps more strictly applicable to non-metallic minerals, are:

Silica, either as chert or as a granular cavernous quartz, and chemically or mechanically combined with hydrous oxides of iron and manganese.

A great variety of clays or hydrous silicates of alumina, generally very impure and charged with oxide of iron and manganese, the extreme of purity being white normal kaolin, containing at times sulphuric acid in appreciable amount.

Sulphate of baryta or heavy spar.

Carbonate of iron, pyrite, and sulphate of lime are comparatively rare in the deposits of Leadville itself.

The miner's term, Chinese talc, has been retained for a substance which is found with singular persistence along the main ore channel, or at the dividing plane between White Porphyry and underlying limestone or vein material, and also at times within the body of the deposit. It is composed of silicate and a varying amount of sulphate of alumina, to which no definite composition can be assigned. It is compact, semi-translucent, generally white, and so soft as to be easily cut by the finger-nail. It is very hygroscopic; hardens and becomes opaque on exposure to the air.

Distribution.—With regard to the distribution of the above ores the principal generalizations to be made are:

I. *That the main mass of argentiferous lead ores is found in calcareo-magnesian beds.*

II. *That ores containing gold and copper are more frequently found in silicious beds, in porphyries, or in crystalline rocks.*

These associations have already been remarked in other mining districts.

Secondary alteration.—Here, as elsewhere, the ores found near the surface are mostly oxidized or chloridized ores, and those farther removed from it, or comparatively unexposed to the direct action of surface waters, are mostly sulphides. It may be observed, moreover, that the zone of secondary deposition, or that in which oxidized ores predominate over sulphides, varies in the depth to which it extends with the relative altitude of the deposit; or that in higher altitudes, where surface waters are imprisoned by frost during a larger portion of the year, the proportion of secondary products is less.

There is a contrast in this respect, however, between the deposits of Leadville and those of the more arid regions of the Great Basin. In the latter the surface zone, or zone of oxidation, is generally more sharply defined and extends down to what is known as the water level. This contrast is more apparent than real, for the zone of oxidation is there dry, because of the limited atmospheric precipitation, and in Leadville generally wet, partly because of the relatively great precipitation and partly because of the peculiar geological position of the deposits, which renders them more accessible to surface waters. The alteration of the ore deposits is produced, not by the water alone, but by the atmospheric agents which it brings from the surface with it; whereas in the case of deposits below the water level the water which reaches them, not coming directly from the surface, but through a relatively long underground passage, has during that passage been deprived of these active agents of oxidation or neutralized.

Mode of formation.—From the present investigation it has been assumed, with regard to the mode of formation of these deposits:

I. *That they were deposited from aqueous solutions.*

II. *That they were originally deposited mainly in the form of sulphides.*

III. *That the process of deposition was a metasomatic interchange with the material of the rock in which they were deposited.* That is, that the material of which they were composed was not a deposit in a pre-existing cavity in the rock, but that the solutions which carried them gradually dissolved out the original rock material and left the ore or vein material in its place.

IV. *That the mineral solutions or ore currents concentrated along natural water channels and followed by preference the bedding planes at a certain geological horizon, but that they also penetrated the adjoining rocks through cross joints and cleavage planes.*

Age of deposits.—As regards the time of deposition of the original ore deposits, it is proved:

That they were deposited not later than the Cretaceous period.

That they are later than the inclosing rock is proved by their mode of occurrence; and since they have partaken of the dynamic movements to which these rocks were subjected, and were folded and faulted with them, they must have been formed earlier than these dynamic movements, which, as the geological considerations already presented show, occurred not later than the close of the Cretaceous period.

Origin of the metallic contents.—With regard to the immediate source from which the minerals forming these deposits were derived, the following conclusions have been arrived at:

I. *That they came from above.*
II. *That they were derived mainly from the neighboring eruptive rocks.*

By these statements it is not intended to deny the possibility that the material may originally have come from great depths, nor to maintain that they were necessarily derived entirely from eruptive rocks at present immediately in contact with the deposits.

The facts and reasons on which these conclusions are based will be given in the following chapters.

CHAPTER II.

IRON HILL GROUP.

IRON HILL.

General description.—Of the three principal groups of mines, that of Iron Hill presents the simplest type, both in geological structure and in the character of its ore deposits. It is that of a block of easterly-dipping beds, with a fault on its western side, by whose displacement these beds have been lifted in places about one thousand feet above their western continuation, and in which the ore deposition has taken place at the upper surface of the limestone bed, along its contact with the overlying porphyry, and extending down at times into the mass of the limestone. This simple type obtains only on the south end of Iron Hill, and even then in a somewhat modified form, the north presenting, as will be seen later, the extreme of complication.

The area represented on the Iron Hill and North Iron Hill maps forms topographically one continuous ridge. The map has been printed on two sheets, partly because of its otherwise cumbersome size and partly because the geological character of the opposite ends of the hill is very different.

The Iron Hill map includes all of Iron Hill except its northern portion, together with a part of Dome or Rock Hill, the spur which lies between California and Iowa gulches. It thus takes in all the mines belonging to the Iron Silver Mining Company, to the La Plata Mining and Smelting Company, and to the Silver Cord Combination, which represent the principal developments outside the Adelaide-Argentine group in this portion of the Leadville region.

Iron Hill and its companion, Carbonate Hill, are flat-topped bosses or shoulders, on the main spur of the Mosquito Range between California and Evans gulches, whose form was evidently due originally to the displacement of Iron and Carbonate faults, though much modified by later erosion. The region, however, as distinguished from the other portions of Leadville, has been scarcely affected by glacial action, California gulch, in which erosion has been deepest, being, as has already been shown, essentially a valley of erosion. The slopes of the hills are steep, but extremely regular, and covered with an accumulation of "Slide," whose average depth may be considered to be from six to ten feet. This slide is distinguished from Wash by being not rounded, but angular and resulting from the disintegration of rock in place. It consists mainly of the débris of White Porphyry, which forms the top rock of either hill. The porphyry weathers into thin sherd-like fragments, which from their relative lightness are easily carried down by rain or snow, and therefore cover the greater part of the slopes of the hills, even where other rocks actually crop out. It is only along the steep slopes of the V-shaped valley of California gulch that actual outcrops of rock in place are found on either hill.

Geological structure.—The average strike of the formations on Iron Hill is a little west of north, and the beds dip east at an angle of about 12° to 25°, shallowing, however, to the eastward, and probably basining up toward the Mike fault. The south face of the hill has, by the erosion of the deep V-shaped valley of California gulch, been left so steep that its surface is but thinly covered by detrital material, and east of the Iron fault, whose line is marked by a slight depression down the slope, the outcrops of the succeeding sedimentary beds can be readily traced, in the numerous prospect holes, from the Lower Quartzite, immediately overlying the Archean, up to the main body of White Porphyry, which forms the summit of the hill.

The geological section represented on this slope is, then, in descending order:

1. White Porphyry capping, in which are included detached portions of the Weber Shales, represented in the Imes shaft by black shales and, along the outcrops on the Lime and Bull's Eye claims, by a greenish slate containing plentiful casts of *Lingula mytiloides*.

	Feet.
2. Blue Limestone	200
3. Parting Quartzite (outcrop obscure)	20
4. White or Silurian Limestone	140
5. Lower or Cambrian Quartzite	160
6. Archean gneiss (not exposed)	—

Later intrusive sheets.—Besides this normal series of beds, are two intrusive sheets of porphyry of later eruption than the White, and allied to, though not absolutely identical with, the Gray Porphyry. One of these is found at the top of the Blue Limestone, the other near its base. Their probable extent can be best seen by reference to the map and sections (Atlas Sheets XXIII, XXIV, XXV). The thicknesses there given are assumed from the position of outcrops, where they could be determined, and from other indirect evidence, and may differ considerably from the actual facts, as these porphyry sheets, especially the later ones, vary much in thickness in relatively short distances.

Upper sheet.—The rock of the former of these bodies is of a dark-gray color and consists of plates of altered mica and relatively large, opaque, white feldspars in a greenish-gray matrix. So far as seen it is in a too advanced state of decomposition to allow of a satisfactory determination of its original constituents. Externally, however, it resembles more closely the country rock of the Printer Boy mine than any other porphyry collected.

This sheet, while in general separating the White Porphyry from the Blue Limestone, does not always keep exactly the same horizon. In the bed of California Gulch, where the outcrops cross and where this porphyry seems to be thickest, it cuts into the Blue Limestone, leaving a portion of the latter above it, near the mouth of the La Plata tunnel. Farther west, on the hill slopes, it cuts up into the White Porphyry for a short distance, leaving a sheet of that rock between it and the Blue Limestone, and then again returns to the contact on the Lime claim, on Iron Hill, and west of the Dome fault, on Dome Hill. There is direct evidence that the sheet thins or wedges out from this crossing of California Gulch to the south, west, and north, but on the east no workings have yet reached a sufficient depth to cut it. It is not impossible that it may be an offshoot from some large body occupying a lower position in that direction—the Printer Boy body, for instance, which is at a lower geological horizon, though actually brought to a higher elevation by faulting.

Lower sheet.—The rock of the second body, as compared with that just described or with the normal Gray Porphyry, has in the hand specimen a much finer grain, and its minute feldspar crystals are generally of a flesh color. When thoroughly bleached by decomposition it can be distinguished from the White Porphyry by its speckled or mottled appearance, whence the name of "mottled porphyry" that is not infrequently applied to it. It is probably also a variety of Gray Porphyry, though, like the preceding, not found in sufficiently fresh condition for exact determination.

As nearly as can be determined from the various prospect holes on the slope of the hill, this body has its maximum thickness near the line of the Iron fault and thins out to the southeast. It is best seen in a tunnel driven in near the fault, on its contact with an underlying limestone, which is supposed to be the lower portion of the Blue Limestone, though, as the Parting Quartzite was not actually exposed below it, this cannot be regarded as beyond a doubt. A certain amount of iron-stained material is found at the contact, and it had been supposed by some that this repetition of a contact of porphyry and underlying limestone below the regular outcrop was evidence of another fault, the different character of the two porphyries having escaped observation.

This porphyry sheet is probably of much wider extent than the one previously described, although its actual outcrop is much more limited; as will be seen later, it probably extends under the greater part of Carbonate Hill, and inasmuch as sheets of Gray Porphyry are found in considerable development on the north end of Iron Hill, though at somewhat lower horizon, it is fair to assume, as has been done in the sections (Atlas Sheet XXIV), that it extends under Iron Hill also, gradually lowering in horizon toward the north. It is probable that the small bodies of Gray Porphyry found crossing the limestone in various points of the hill are offshoots from this body.

White Porphyry.—The White Porphyry, which forms the summit of the hill, is the normal rock already described. From the quarry in California gulch, just above Graham gulch, was taken the specimen chosen for complete analysis (see Appendix B, Table I). In this quarry, which is but a short distance west of the Iron fault, the jointing planes are strongly marked, those parallel with the plane of the fault being the most prominent.

Blue Limestone.—The Blue Limestone, as shown by the map, has an unusually broad outcrop in California gulch, owing to erosion and to the low angle at which it stands. From the bed of the gulch the outcrops extend up along the hill slopes on either side, only obscured by slide or surface débris, until cut off by the Iron and Dome faults, respectively. On the Montgomery claim, a cliff exposure of a very considerable thickness of the lower beds is afforded by an open cut, where the limestone was formerly quarried as a flux for the smelters. There is also a small outcrop west of the Emmet fault, near the bed of the gulch, below the Columbia tunnel. From the upper beds in the Silver Wave ground were taken the specimens illustrated in Plate VI (p. 64) and whose composition is shown in Appendix B, Table V. The characteristic ribbed structure is here very well developed. The thickness of the formation, as calculated from these outcrops, is two hundred feet or more, which is greater than that deduced from measurements on Carbonate Hill.

Silurian.—The White Limestone is disclosed in numerous prospect holes, and some shafts on the south side of the gulch have cut the characteristic Red-cast beds. The Parting Quartzite could not be unmistakably recognized, owing to its close resemblance underground to decomposed porphyry. There is, however, no reason to assume that it is wanting.

Cambrian.—The Lower Quartzite is best shown in the Globe and Garden City shafts, each of which has cut through it into the underlying Archean. The quartzite is of the usual normal type and the Archean is a coarse-grained granitoid gneiss.

Iron fault.—The average direction of the line of the Iron fault is a little east of north, but its course is very crooked, as shown on the map. Although this irregularity may be somewhat increased by erosion, i. e , be greater than if the line given on the map were its intersection with a horizontal plane, still it cannot be considered abnormal, since from the bed of California gulch northward to the Codfish Balls shaft it has been actually proved in so many cases as to render its delineation unusually exact.

It has been cut by the workings of the Garden City shaft; by the L. M. shaft, which was sunk perpendicularly to the depth of two to three hundred feet through White Porphyry, on the west side of the fault, into Lower Quartzite on the east side; by two shafts on the Lingula claim; and by numerous shafts and winzes in the claims of the Iron mine, some of the latter being sunk on the plane of the fault itself, and showing its average dip to be 60° to 65° to the westward, or nearly at right angles to the dip of the formation.

As the Blue Limestone has not yet been reached on the west side of the fault in the region represented on this map, its movement of displacement, or throw, cannot be accurately determined. Its maximum is probably not far from one thousand feet, since the City of Paris shaft, 1,200 feet north of the line of the map, was sunk to a depth of 800 feet without reaching the Blue Limestone. The dip of this bed carried back from the outcrop on Carbonate Hill, at the average angle, would reach at the line of the fault a much greater depth, probably not less than fifteen hundred feet; but there are good grounds for assuming that this dip shallows, and that the beds actually basin up, i. e., assume a westerly dip, before reaching the line of the fault. The movement of this fault may here be partly distributed among smaller parallel faults to the west, like the Carbonate fault, in which case the contact immediately adjoining the main fault may be found at a less depth than 1,000 feet. To the north, beyond the limits of this map, as has already been seen in the general description of the Leadville region, the movement of the Iron fault gradually decreases and it apparently passes into an anticlinal fold As regards the continuation of the fault south of California gulch, however, no definite data have been obtained, since the great accumulation of Wash and Lake beds there have been a barrier to underground explorations. It has been assumed that it gradually passes into a synclinal fold, as indicated on the map of Leadville. The movement of displacement south of California gulch is, however, distributed among two faults, the Dome and the Emmet, with which the Iron fault is connected by a cross-fault (the California fault), which follows approximately the bed of California gulch.

California fault.—The plane of this fault has not been actually cut, but its existence is proved by the discrepancy of the beds on either side of the gulch, the Blue Limestone outcropping near the Robert Emmet tunnel and opposite the Globe shaft, in which the Lower Quartzite is cut.[1]

Dome fault.—The Dome fault is in one sense the proper continuation of the Iron fault, since it forms the great break on Dome Hill, as Iron fault does on Iron Hill, and, like the latter, passes at its extremity into an anticlinal fold. Considered in this way, the Iron, California, and Dome faults would form a single fracture, somewhat irregular in direction, but having a general north-and-south trend, while the southern continuation of the Iron fault, as at present indicated, and the Emmet fault, would be simply branches, relieving the strain at the sudden bend of the fault in California gulch. To the east of this line of fracture are the principal outcrops of Blue Limestone and the main ore developments in this region, while to the west this horizon is more or less deeply buried beneath a covering of porphyry. The Dome fault proper has a general north-and-south direction. Its plane has been proved by underground workings only in the Vining tunnel, but the line as given on the map is tolerably closely determined by the developments of adjoining shafts and inclines, those on the west finding White Porphyry, underlaid by Gray Porphyry, on a level with Blue Limestone on the east, in the Rock and Dome workings.

Emmet fault.—This small fault, running in a southwest direction from the California fault, has a movement of displacement the reverse of the majority of the faults in this region — that is, the upthrow is to the west instead of to the east. Its plane has actually been proved by a drift running westward from a winze sunk in the Robert Emmet tunnel. It is further proved by the discrepancy in the position of the Blue Limestone and the overlying porphyries on either side of it, as shown in Section G, Atlas Sheet XXV. That it actually continues to its junction with the Iron fault to the south, as indicated on the Leadville map, is merely a matter of conjecture.

Dome Hill.—By reference to Atlas Sheet XXV, Sections E and F, it will be seen that the northern portion of the ridge of Dome Hill, adjoining Cal-

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[1] Since the close of field-work, developments in the Garden City mine have definitely located the position of the western end of this fault. The lower shaft on this claim was sunk perpendicularly 100 feet through limestone and vein material, and then passed into the Lower Quartzite, crossing the fault diagonally. At 120 feet a drift to the southwest cut the fault at 5 feet from the shaft, showing that its dip is to the south. At 75 feet from the shaft the same drift cut the plane of the Iron fault and passed into the White Porphyry on the west side of this fault.

ifornia gulch, was originally part of the Iron and Carbonate Hill ridge and that their present separation by the valley of California gulch is due tc erosion since the Glacial epoch. What is now the main crest of the ridge was once an arm or bay in the Arkansas lake, and the actual rock surface is buried to a great depth beneath the deposits formed in this lake and the later Wash Except, therefore, on the northern edge of the ridge adjoining California gulch, which is the portion shown on the Iron Hill map, data with regard to the actual rock surface are extremely meager. Its geological structure above and to the east is similar to, and practically a continuation of, that of Iron Hill, namely, a series of easterly-dipping beds, capped by porphyry, in which the ore bodies have been developed by following the contact of the Blue Limestone with the overlying porphyry. The main difference lies in the development of the intrusive sheet of Gray Porphyry below the White Porphyry, which is not, however, absolutely parallel with the bedding, inasmuch as on the summit of Dome Hill a small sheet of White Porphyry is left between the Gray Porphyry and the limestone and in the La Plata ground the Gray Porphyry cuts down through the upper part of the Blue Limestone.

West of the Dome fault the relative position of these two sheets of porphyry affords most valuable evidence as to the underground structure, and actually proves a basining-up of the beds towards the Dome fault, as has been assumed to be the case in regard to the beds west of the Iron fault. At the Bank of France shaft the Gray Porphyry actually comes to the rock surface. The City Bank and Oro City, on the other hand, pass through the White Porphyry into the Gray, as does the Vining shaft higher up on the hill. The Sullivan, Ben Burb, and Keno shafts have reached the contact and limestone after passing through the White and then a comparatively thin body of Gray Porphyry. The Blue Limestone is thus shown to be at no great depth below the surface near the Dome fault. On the other hand, at the Coon Valley shaft, near the head of Georgia gulch, the Blue Limestone is over six hundred feet deep, showing a comparatively steep dip from the fault westward.[1]

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[1] Since the completion of field-work the contact and even valuable bodies of ore have been proved in this region west of the Dome fault, notably in the Rosie, Sequin, and Vining claims. In the Sequin the contact was struck at 375 feet, in the Vining at 317 feet, in each case with a sharp dip to the westward.

The wedge-shaped block of ground between the Emmet and Iron faults may be considered a portion of the formation which, by compression between the adjoining blocks, has been lifted up relatively and compressed into an anticlinal fold. Actual outcrops of Blue Limestone are found near the bed of California gulch, opposite the Globe shaft. The Columbia tunnel was run in apparently on the very crest of the fold and developed considerable ore on the contact. From the line of the tunnel the formation dips gently to the eastward and very steeply to the westward, so that in the Crescentia shaft, a little west of it on the slopes of California gulch, at a depth of 335 feet the limestone had not yet been reached, but the shaft was in the Gray Porphyry beneath the White.[1] Section G, Atlas Sheet XXV, represents graphically the structure thus described.

Ore deposits.—The principal deposition of ore has taken place along the contact-plane between the Blue Limestone and overlying White Porphyry, and extended to greater or less depth into the mass of the limestone. In several instances large deposits have been formed within the body of the limestone, being probably on the line of some natural cleavage or joint plane which caused a deviation of the ore currents from their normal course.

The vein material or gangue consists of hydrated oxides of iron and manganese, silica, and clay. The iron varies from a hard, compact, more or less silicious brown hematite to a simple coloring matter of the clay. Manganese is found sometimes in fine, needle-like crystals of pyrolusite, but mainly occurs as a sort of wad, a black clayey mass known to the miners as "black iron." Silica occurs either as a blue-black chert or as a granular, somewhat porous mass, hardly distinguishable from quartzite. Clay is found in greatly varying degrees of impurity, from a white kaolin down, and is a product of the decomposition of porphyry. It occurs either in place or as an infiltrated mass. Besides this should be mentioned the Chinese talc of the miners, found mainly at the actual contact.

The ore is principally argentiferous galena and its secondary products are carbonate of lead, or cerussite, and chloride of silver. As accessory

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[1] Late developments in the lower Garden City shaft show that the Blue Limestone is considerably mineralized and that the formation dips very steeply to the southwest.

minerals, or those of less frequent occurrence, are sulphate of lead or anglesite, pyromorphite, minium, zinc blende, and calamine. Native sulphur is found in one instance as the result of the decomposition of galena, and native silver formed by the reduction of chloride.

MINE WORKINGS.

The principal mine workings in the area represented on the Iron Hill map may be divided into the following groups, commencing at the south:

1. The Rock and Dome.
2. The La Plata and Stone.
3. The Lime and Smuggler.
4. The Silver Wave and Silver Cord, including the South Bull's Eye.
5. The Iron mine proper, including the North Bull's Eye.

Rock and Dome.—These two claims are owned and worked by the Iron Silver Mining Company. The former is opened by a tunnel running southward on the strike, the latter by an incline running eastward on the dip. The ore bodies thus far developed in either mine are found near the surface of the hill and may belong to the same bonanza, if the same northeasterly direction of ore shoots prevails here as does on Iron Hill. On the hillside, at the present mouth of the Rock tunnel, was formerly an actual rock outcrop, consisting largely of hard carbonate, from which the mine derived its name and where the first ore in place was found in this region. From it were no doubt derived the heavy fragments which caused so much annoyance to the early gulch miners.

From this tunnel level the ore has been followed along the contact of limestone and porphyry a certain distance upward or toward the outcrop, but mainly eastward in the trough of a fold and then downward on the dip The workings have also been pushed southward with the intention of making a connection with the Dome workings. Beyond the crest of the fold to the eastward the contact has thus far proved comparatively barren, but at the lower extremity of the Rock incline ore has been found which may be the precursor of a second ore shoot.

In the Dome the rich ore has thus far been found near the mouth of the incline, in very considerable thickness and with a remarkable development of masses of Chinese talc in the ore body, at some distance from the contact. The incline has not yet reached a second ore shoot in depth, though there is every probability that one will eventually be found there.

The ore in both these mines is mainly a hard carbonate, very rich in lead, but of comparatively low grade in silver. It is very thoroughly oxidized, and in some cases a red oxide of lead has been found in it. It occurs in bodies sometimes of considerable thickness and always at or near the contact. At the contact the alteration of porphyry into the so-called Chinese talc is very persistent, and when found in the ore body, as in the Dome mine, shows that offshoots of the porphyry had probably penetrated the limestone previous to the replacement of the latter by vein material.

Sections E and G, Atlas Sheet XXV, which pass through the Rock workings, show the fold in the limestone, which affords a good illustration of the tendency of the ore currents to deposit their load immediately above any sharp bend in the stratification.

La Plata, Stone, and A. Y.—The La Plata claim is opened by a tunnel 800 feet long, running south from near the bed of the gulch. Its direction was intended no doubt to correspond with the strike of the formation, but in point of fact it diverges a little to the westward, so that while at the mouth it is at the actual contact of the White Porphyry and Blue Limestone, it departs from it more and more as it advances. At the extremity, however, the contact bends sharply down to the south, so that a winze has been sunk 70 feet to reach it. It is noticeable that this bend is on a line with the eastward continuation of the California fault.

Below the mouth of the tunnel and in the body of the limestone is found the Gray Porphyry sheet, which to the north and south is found above the Blue Limestone and separating it from the White Porphyry. The contact in this mine was not found very productive. A small body of ore was found east of the tunnel, near its mouth, and a prospecting drift running to the Gneisson shaft, and continued some distance beyond it, found the usual evidence of mineralizing action, but no pay ore; it showed, however, a steepening of the dip of the formation of 35°. This, with the sudden steepening at the end of the tunnel, shows how difficult it is to count on any regularity in the dip of the formation until it has been actually proved. The main ore developments have been in the body of the limestone, extending as much as one hundred feet below its surface, and are opened by the Rustin shaft. These and the similar ones in the Silver Wave ground are interesting as showing that the ore deposits are by no means confined to the surface of the limestone, as was originally supposed. The bodies are irregular in shape, but have their greatest extent in a nearly vertical direction. The slickensides found on their walls give evidence of some movement, and they were evidently formed by ore currents percolating along cross-joints or planes of fracture in the limestone, having a general north and south direction. The ore is oxidized and does not differ essentially in character from that in the adjoining mines.

The Stone claim was ingeniously outlined to take in the curving outcrop of the Blue Limestone as it crosses the gulch. The developments on it are mainly on the north side of the gulch, and have as yet opened no considerable ore bodies, though the evidences of replacement action are abundant. Probably a search below the contact for bodies similar to those of the La Plata might prove remunerative.

The shaft of the A. Y. mine, above the Stone claim, has developed an extremely interesting occurrence of unoxidized ore, a mass of galena, pyrite, and zinc blende, which was the only one reached in the Leadville region, though unfortunately not accessible, at time of visit. The ore is of low grade in silver, and hence of little value in competition with the more easily reducible oxidized ores.

Lime and Smuggler.—Directly opposite the Rock workings, and at a corresponding elevation on the north slope of California gulch, are the workings of the Lime, Smuggler, and adjoining claims, which, though not very extensive, are sufficient to give evidence of another zone where the limestone has been largely replaced by vein material. The minor folds in the limestone are here very sharp, the rock masses near the surface sometimes broken, and the replacement has been somewhat irregular, so that the continuity of the ore bodies is not always evident. Here, as in the Dome claim, a small thickness of White Porphyry separates the intrusive sheet of Gray Porphyry from the contact, as shown in Sections E and F. From the south incline of the Lime to the South Bull's Eye but little ore has yet been developed along the contact. On the extreme north end of the Lime claim, an incline, not indicated on the map, was driven in on the contact until cut off by a wall of Gray Porphyry, standing at an angle of 65° with a strike to the east and northeast. This would seem to be an offshoot from the intrusive sheet in the lower part of the Blue Limestone. The form of this offshoot, shown in Section E, must be understood to be, in the present state of developments, purely a matter of conjecture. That ore bodies have not been found at the contact here is, however, not necessarily a proof that they may not exist within the body of the limestone, as will be seen from the description of the next group. Thin beds of shales carrying *Lingula* are found at the contact in the Lime and Bull's Eye claims, near the outcrop.

South Bull's Eye, and Silver Cord Combination.—A considerable body of rich carbonate ore was found along the contact and near the outcrop at the south end of the Bull's Eye claim, which has been developed by the so-called South incline. It was extremely irregular in shape, extending in places fifteen or twenty feet below the contact; its probable continuation in the Silver Wave ground is apparently even thicker. As shown in Section D, this body, like that already described in the Rock mine, occurs just above and on the crest of a fold in the limestone, whose axis has a northeast direction parallel to that of the ore body. The ore was quite rich near the surface, but became poorer in depth. To the south it passes into black iron (wad), containing little or no silver. The incline, which runs diagonally across the body and follows approximately the contact plane, has at first an inclination of 12°, and after passing the crest of the fold steepens to an average angle of 25°, and for short distances reaches 45° or more; the contact is here barren, showing only iron-stained clay and a little Chinese talc. This body, like that of the Rock mine, was one of the earliest developments in the district.

In laying out the side lines of the Bull's Eye claim it was the intention of the original locators to include within them, as they did so successfully in the other claims of the Iron Silver Mining Company, the outcrop of the vein or of the upper surface of the Blue Limestone. As it happened, however, the limestone rises at this point over a secondary fold, and the line of contact bends backward, or up the hill, instead of following its normal grade along the slope, so that for a considerable distance the line of outcrop passes east of the Bull's Eye line and within the ground of the adjoining Silver Wave claim. A most valuable piece of ground was thus lost by an accident, which only an actual stripping of the limestone outcrop over its entire extent could have prevented. During the time this investigation was carried on, owing to pending litigation, the Silver Wave claim, which has since been consolidated with the claims adjoining it on the east, in what is known as the Silver Cord Combination, was not open to public inspection; nor could permission be granted to take copies of the maps of undergound workings, as was in general freely accorded by the Leadville mine owners. The workings and outlines of ore bodies in these claims, as given on the map, are hence necessarily incomplete, being made up from data obtained from outside surveyors and from notes gathered during a rather hasty personal inspection of the workings.

The mine is opened by two inclines from the surface in the northern part, and by shallow shafts from which inclined drifts follow the ore channels in a very irregular manner, in other portions of the claim.[1] The main or most northern incline runs at an angle of 15°, striking the contact at 10 feet from its mouth, and thereafter running in the body of the limestone at an ever-increasing depth below the contact. It thus passes beneath a drift run southward from the fifth level of the Iron mine, which follows a barren contact. On the contact little good ore has been developed, but very rich ore, and probably in very considerable quantity, said to have produced many hundreds of thousands of dollars, has been obtained from bodies in the mass of the limestone, and extending in some cases to a depth of one hundred feet below the contact. In those visited, the outlines of the body, although irregular as in most ore bodies in limestone, have in general a northeast direction and stand nearly vertical. At their upper limits can be generally distinguished a distinct crack or jointing plane in the limestone, extending up to the contact, as evidenced by the entrance of water through it. On the other hand, at the lower limits of these bodies no trace of any opening could be found through which the ore solutions might have come

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[1] On the map, by error in proof-reading, the parallel linings used to denote *inclines* have been omitted in this mine.

from below. In general outline the bodies seem pear-shaped, with the taper toward the top. Similar bodies are found around the Silver Cord shaft, also having a northeast trend, and, as the map shows, in a direct line with those in the Silver Wave and Grand View claims. East of the Silver Cord shaft a steeper dip in the formation comes in, which may be a continuation of the fold already noticed in the South Bull's Eye. It seems, then, that before the dynamic movement in this region there was a certain amount of fracturing of the beds, not, however, accompanied by any considerable displacement, and that along these planes of fracture the ore currents have penetrated into the body of the limestone, the ore deposition or replacement acting from their walls outward.

In the Silver Wave claim was also seen a freshly opened cave, one of the few that are found in the Leadville mines, and which is of interest as bearing on the generally-advanced theory that ore bodies in limestone are necessarily deposits in *pre-existing* cavities. It was somewhat funnel-shaped toward the top, about twenty-five feet in horizontal diameter, and contained no ore. Its walls, which had the wavy surface common to water-worn limestone, were covered with a thin coating of fine reddish ooze or slime. An examination of the walls showed that these were in part of unaltered limestone and in part of ore and vein material, which could not be distinguished from each other until the coating had been removed. It was thus evident that the cave was of comparatively recent formation, made by the percolation of surface waters and carved out of limestone and ore body indifferently; hence, that it is entirely posterior to the deposition of the ore, which was formed before surface waters, as the term is generally understood, could have reached to this depth.

Iron mine proper.—The underground workings of the group of claims which are exploited from the various shafts and inclines of the Iron mine cover an area of about twenty-five acres, being the most considerable of any single mine in the district. They have been driven a distance of over fifteen hundred feet along the contact eastward from the outcrop, or rather from the fault line, since at the peculiar eastward bend of the fault plane in the Iron and Iron Hat claims the limestone does not actually come to the surface.

The Iron Hill map shows the principal underground drifts in this area, taken from the actual working maps of the mine, the level of different points in these drifts being given by figures which denote their respective elevations above the 10,000-foot curve. The outlines of the ore bodies are given in generalized form, as deduced from the same maps and from personal observations. As in the case of all the mine maps, both drifts and ore bodies are indicated in projection, that is, as if the ground over them were transparent, and in this respect they differ from and are independent of the geological colors, which indicate the formations constituting the rock surface.

The mine is opened by three principal inclines, the North, Main, and South, the last of which is no longer in use. The bulk of the ore is extracted through the middle or Main incline, which is carried down at an angle of 12° to 13° in approximate conformity with the surface of the limestone; it is provided with powerful hoisting engines and has a double track. In Atlas Sheet XXIV a section is given through each of these inclines, designated A, B, C, respectively, the line of the last running partly through the North incline of the Bull's Eye claim, which adjoins the line of the Iron claim.

In this area the contact has been and is productive over an unusually large surface, the main ore body extending diagonally through the claims in a northeast direction from the croppings, with an average width of 200 feet. This productive zone is separated from the zone of the adjoining Silver Cord Combination by comparatively barren ground; that is, barren as far as present explorations have gone, although it is not absolutely certain that ore may not still exist in the body of the limestone. The irregularity with which the replacement action of the ore currents has acted upon the limestone is well shown in the Main incline. Here, after the ore had been extracted along the actual contact from the first to the fifth level and it was supposed that pay ore in this area was quite exhausted, it was found in one place to extend considerably below what was supposed to be the floor of the ore body, often simply a layer of black chert, and a lower drift was run back from the fourth station in the direction of the fault, disclosing a very large body of vein material and rich ore, extending nearly to the fault plane,[1]

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[1] Later developments have shown that this ore body extends actually to the fault plane. Several thousand tons of ore have been extracted from it through the new McKeon shaft, at about fifty feet below the contact.

and in places reaching a depth of 40 feet or more below the actual contact of limestone and porphyry.

In Section C, which passes through the North incline of the Bull's Eye, and a little south of the South incline of the Iron mine, the tendency of the rich ore to accumulate above a fold in the limestone, which has already been noticed in the Rock and South Bull's Eye, is quite apparent, the barren zone occurring on the steeper dip of the formation towards the Silver Cord claim.

In the section through the Main incline the folds are less prominent, but the same tendency always holds good, and it is one of the practical generalizations made by those working in the mine that rich ore bodies occur always in troughs of the limestone. The steeper dip of the formation beyond the accumulation of rich ore is quite evident. Just below the seventh level a small body of Gray Porphyry crosses the Main incline diagonally in a direction a little north of east. Here the incline is some distance below the contact, and it could not be definitely determined whether the porphyry extended up to the contact or not, though it has unintentionally been indicated as doing so in the section. To the westward, if it continues in that direction, it does not, as the contact has been explored on the line of its continuation without finding it. It is cut in the eighth level a short distance north of the Main incline, but in neither case is the limestone mineralized to any extent at its contact.

On the line of the North incline the general dip of the formation has become extremely shallow, as shown by the old drift, known as the Tucson incline, which followed the contact in all its curves and irregularities. This shallowing of the dip is probably due to a general basining-up of the formation to the northward, since on North Iron hill in the Adelaide and Argentine ground it curves in strike to the eastward and assumes a southerly dip. Thus at the Hynes shaft, which is on the same line of strike with the Tucson shaft, the contact stands about fifty feet higher than at the latter.

In this portion of the mine a second series of less important ore bodies occurs in a depression in the limestone to the west of the main bonanza and near the fault line. It has, like the latter, a general northeast trend.

Whether it will lead to more important developments in that direction explorations have not yet been sufficiently extensive to determine.

The main ore body on this line extends more or less continuously from a little above the fifth level eastward to the bottom of the Tucson shaft, being mainly concentrated between the fifth and eighth levels, where it extended at times to a depth of 30 feet or more below the contact. A very interesting feature of this remarkable ore body is the occurrence of a body of Gray Porphyry, cutting up into the limestone and at one point reaching the contact with the White Porphyry. It has no apparent connection with the body already mentioned in the Main incline. At the time of examination it was so little explored that but little could be ascertained as to its form or extent, and the representation given in Section A is almost entirely ideal. It is there drawn as extending across the contact into the White Porphyry, for the reason that Mr. Jacob found White Porphyry under it in the old Tucson drift, where it runs above the North incline. Both north and south of this line, however, it does not reach the contact, and ore and vein material are continuous over it. Later developments have shown that the ore extends to a considerable depth into the limestone along its contact and that its general direction is northwest and southeast.

It is probable that both these bodies of Gray Porphyry are irregular offshoots from the main intrusive sheet at the base of the Blue Limestone and differ from the ordinary dike. The fact that the one which crosses the general direction of the ore bodies is accompanied by a concentration of rich ore in its vicinity, while that which runs parallel with this direction is not, is in accordance with the conditions found in connection with such cross-cutting bodies of porphyry on Carbonate and Fryer Hills and with the theory that they are favorable to the concentration of ore when so situated, in that they would produce a retardation in the flow of the ore solutions and thus give them more time to deposit their load.

It was in one of the drifts running north from the North incline, at the sixth level, that a mass some two feet in diameter was found, composed mainly of native sulphur associated with a little carbonate of lead. As it was comparatively free from iron oxide, it seems evident that it must have resulted from the reduction of galena, the lead having been removed in the state of carbonate.

In the body of the limestone, on the eighth level not far from the North incline, a natural jointing plane, forming one wall of the drift, was observed to be coated with fine, silky, white crystals, which chemical examination proved to be calamine or silicate of zinc. If the sulphureted ores, which will undoubtedly be found when the mine workings shall have reached the limits of the zone of oxidation, are as rich in blende as those which have been found in the A. Y. mine, it seems singular that little or no zinc has hitherto been found associated with the oxidized ore. This occurrence would seem to show that, owing probably to greater solubility, the alteration products of blende have been removed during secondary deposition to a greater distance from their original location than those of the other sulphurets.

In the lower levels of the mine there has been a notable increase in the proportion of unaltered galena in the ore, but as yet no pyrites or other sulphurets have been found. While specimens of galena are still found which average as high as 1,200 ounces of silver to the ton, the general tenor of the ore is lower than near the outcrops, and the evidence afforded by the records of assays, which were very systematically kept in this mine, shows that there has been a gradual but comparatively steady decrease in the average tenor of the ore in silver with the progress in depth. These records further show, and their evidence was confirmed by numerous tests made in the laboratory of the Survey, that no reliance can be placed on a relation assumed by some to exist between the coarseness or fineness of grain of a galena and its contents in silver.

Explorations to the eastward beyond the Tucson shaft and in the lower part of the Main incline have been carried on along the contact line thus far without very remunerative results. It would seem probable that ore might be found in this direction in the body of the limestone, and possibly in more or less direct connection with the cross-cutting sheet of Gray Porphyry, from which those above mentioned are offshoots and which may be assumed to be at a considerable depth below the contact in this eastern region.

Relation of Iron fault to ore bodies.—In the area under consideration the plane of the Iron fault has been cut in so many places as to render its tracing practically continuous at its intersection with the contact. Its continuation has also been traced through the porphyry above the contact to the surface. Winzes have been sunk just north of the Main incline to a depth of 100 feet on the fault fissure, and from the McDonald shaft 65 feet, as shown in Sections A and B. Examinations of these workings, and descriptions of them where they were no longer accessible, render it very evident that the faulting has been posterior not only to the intrusion of the porphyry, but also to the deposition of the ore.

It has been soberly maintained by some experts when testifying in lawsuits that the faulting was previous to the eruption of the porphyry and that the latter flowed down over the successive benches formed by the faults, following their surfaces. A consideration of the general geological structure of the region, where instances abound showing that porphyry bodies and sedimentary beds were both folded and faulted together, should be sufficient to show how untenable is such a theory; but a sufficient refutation is found at this very point in the fact that the fault plane can be traced up to the surface through the overlying porphyry.

That the ore was originally deposited previous to the faulting is less self-evident, since in places there is a certain amount of alteration of the limestone adjoining the fault plane and since ore has been actually found in the fault fissure, which often has a width of three feet or more and is filled with a dark clayey mass, bearing a certain resemblance to vein material. The alteration is only such as might have been expected from the action of surface waters passing across the ends of the contact adjoining the fault, and consists merely in a slight impregnation or replacement of the limestone by oxides of iron and manganese. This action extends at most only a few feet into the limestone and is confined to a region comparatively near the surface. Had the original ore-bearing currents actually followed the plane of the fault, ore deposition would have extended to a much greater distance into the body of the limestone from the fault plane than from its upper surface, inasmuch as far easier access to percolating waters would have been afforded by the numerous bedding planes.

As regards the question of ore found along the fault plane, it may readily be conceived that in the dragging movement of the edges of two immense bodies of rock, the one against the other, during the fault displacement, a very considerable amount of the adjoining rock could be broken off and carried along for some distance from its original position. The greater part of this material would be clay from the porphyry, but with it would be mixed a certain amount of limestone and ore. The circulation of waters from the contact plane on either side, and therefore carrying more or less mineral matter in solution, might occasion a secondary replacement of this limestone by ore. It is even conceivable that in contact with the inorganic matter, which must have been present, sulphates might have been reduced to sulphides and galena have been deposited, but, unless the mineral were found in the limestone outside of the attrition material of the fault fissure, it would not be a proof that it was an original deposit before the fault movement.[1]

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[1] In the years that have elapsed since this was first prepared for the press, a new shaft has been sunk 110 feet south of the Main incline for the purpose of exploring the fault plane. The data obtained from this by personal observation and from information furnished by Mr. F. T. Freeland, engineer of the mine, and who was present during all the explorations, furnish a remarkable confirmation of the above views. The shaft was sunk to a vertical depth of 306 feet, but at an angle of 50°; drifts were run on the fault plane at four levels, that on the second level having a total length of 2,000 feet. In this level the sharp eastward bend of the fault plane has practically disappeared. On the first level, which corresponds to the third level of the Main incline, a considerable amount of ore was obtained from the lower ore body. Ore was also obtained at various depths on the fault fissure below this level. In regard to this ore, the following facts were observed: First, the ore was always found within the walls of the fault fissure; secondly, it occurred in masses rounded as if by attrition, and evidently foreign to the clayey filling of the fissure in which it was imbedded; thirdly, no ore was found outside of two vertical planes drawn through the intersection of the boundaries of the main Iron mine ore body with the fault plane. It is interesting to compare the actual section obtained in this shaft with that given in Section B, which was a theoretical deduction from data obtained at other points. The angle of the fault was found in depth to average 50°, instead of 65°, as had been deduced from observations near the surface.

Thickness of—	Section B.	McKeon shaft section.
	Feet.	*Feet.*
Slide and White Porphyry	40	45
Blue Limestone and vein material	200	192
Parting Quartzite	24	10
Gray Porphyry	48	72
Contact of White Limestone at	312	319

NORTH IRON HILL.

Atlas Sheet XXVI shows the topography, geology, and principal mine developments of the northern end of Iron Hill, overlooking Stray Horse gulch, and forms, as aforesaid, really a portion of the main map of Iron Hill. In this region ore was first discovered on the Camp Bird claim in the autumn of 1876. At present the principal mine workings belong to two companies, the Argentine and the Adelaide, the former of which owns the Camp Bird and Pine claims, and the latter the Terrible and Adelaide claims; the latter overlaps those of the former company, a fact which has given rise to much litigation.

General geological structure.—As compared with Iron Hill proper, its geological structure is one of extreme complexity, and also difficult of exact determination, for the reason that underground workings are few and accessible in but a comparatively small portion of the area.

In the region west of the Iron fault the structure indicated on the map is deduced from data obtained outside of its area, and to that extent is theoretical. That the Blue Limestone basins up to the eastward as it approaches the fault is proved in the Devlin shaft, which reached it at a depth of 200 feet, and in the Highland Mary and other shafts, in Stray Horse gulch just north of the limits of the map, which found it still nearer the surface. The outcrop indicated in the northwest corner of the map is a portion of the Blue Limestone, split off from the main body, corresponding to that cut in the Agassiz and adjoining shafts, and forming the south end of the Little Stray Horse Park synclinal basin, as explained in Part I, Chapter V.

East of the Iron fault the formations rise slightly to the northward, so that their strike assumes a more easterly and westerly direction and dips to the south and east. By the erosion of Stray Horse gulch, on the lower part of the steep northern slope of Iron Hill, a succession of Paleozoic formations down to the Lower Quartzite are exposed, while by the movement of the Adelaide fault, which crosses the northeast corner of the area mapped, a still lower series of beds is exposed beyond it.

The most striking peculiarity of the structure is the cutting across of the Blue Limestone formation by the White Porphyry, this region being on the line already mentioned as extending from Fryer Hill to West Sheridan, where this cutting down of the White Porphyry sheet occurs. Its effect is graphically shown on Atlas Sheet XXVII, Section B. It will be observed that whereas at the south end of the section the White Porphyry occurs, as it generally does, above and parallel with the Blue Limestone, at the northern end, where exposed by the workings of the Argentine mine, it crosses the basset edges of the Blue Limestone, and at the outcrop it probably comes in contact with the underlying Parting Quartzite. As the remainder of the Blue Limestone, above the cross-cutting of the White Porphyry, has been removed by erosion, it is not possible to determine whether it was mineral-bearing or not. As far as determined by the present workings the deposition of the richer ore has gone on, not as is ordinarily the case at the contact of the White Porphyry with the Blue Limestone, but at its contact with the Parting Quartzite.

Iron fault.—In this area the Iron fault is struck in the Iron Hat shaft, and on the Codfish Balls claim by a shaft and tunnel. Beyond this claim to the northward its location is only approximate, though beyond the limits of the map it is determined very closely by adjoining shafts on either side. Its movement is the same as it was at the south, namely, an upthrow on the east, but the amount of that throw is constantly decreasing as one goes north.

Adelaide fault.—The location of this fault is also approximate, owing to the infrequency of shafts in its neighborhood, and also to the fact that it often has porphyry on either side. Its movement is a slight upthrow on the northeast. Its location is determined by the discrepancy of the formations disclosed by the Laura Lynn, Park, and adjoining shafts in Adelaide Park, and by the Double Decker group of shafts opposite the Argentine tunnel, on the one side, and by the workings of the Adelaide and Argentine mines on the other.

Rock formations.—The sedimentary formations disclosed in this area are the same succession of Paleozoic beds, from the Lower Quartzite up to the Blue Limestone, that outcrop on the southern end of Iron Hill. The porphyry masses are, however, much more varied and numerous. It may be safely assumed that they are mostly intrusive sheets, but the underground workings are not yet sufficiently extensive to determine whether they may all be considered so or not. As has already been noticed in the general description, Part I, Chapter V, there is reason to suppose that one body of Gray Porphyry, extending from Adelaide Park up the south slope of Yankee Hill, has cut up across the formations from below.

The different bodies of porphyry that have been thus far disclosed in this portion of the hill may be enumerated as follows, commencing with those which stand the highest in geological horizon: (1) The main body of White Porphyry overlying the Blue Limestone; (2) a second sheet, cutting across the basset edges of the limestone and connected with No. 1; (3) a small body of Gray Porphyry between No. 2 and the Parting Quartzite; (4) a thin sheet of White Porphyry, splitting the Parting Quartzite into two parts; (5) a heavy body of Gray Porphyry, with two smaller sheets, probably offshoots, above and below it, respectively, all three in the White Limestone; (6) a lower sheet of White Porphyry, also in the White Limestone. The distribution of these bodies and their probable extent can be best seen by reference to Atlas sheet XXVII.

Section A, drawn at an oblique angle to the strike, passes first through the Argentine ground and then through the Adelaide, showing the distribution of the ore bodies in the latter. At its southeastern extremity only White Porphyry is given as cut by it, as it is supposed to be in the strike of the cross-cutting body of this rock. In the entire want of any actual data this theoretical representation may not be absolutely correct. The Blue Limestone is split into two wedge-shaped and probably overlapping bodies. The upper or northeast portion has been eroded off in the Adelaide and Argentine ground. Whether it has also been removed here, as represented in the section, or whether a portion should be shown in the White Porphyry, can only be determined by actual developments. The lower wedge-shaped portion of the Blue Limestone, extending to the south and west in normal contact with the Parting Quartzite, is supposed to come in a short distance southwest of this line, as shown in Section C, whose eastern end is nearly in the plane of Section A.

Section B, drawn approximately through the line of the Argentine tunnel and at right angle to the line of strike, gives the best representation of the geological structure, the lines having been determined by careful measurement. Section C, on the other hand, is rather a theoretical representation of what may probably be found on this line, reasoning from what is observed on either side of it, there being no underground explorations on its plane.

Ore deposits.—Deposition of ore in this region has been extremely irregular, as might have been expected from the complicated nature of the different intrusive bodies which have traversed the sedimentary formations. The main body of rich ore thus far discovered has been, as already mentioned, at the contact of White Porphyry and Parting Quartzite. This is found mainly in the Camp Bird and Pine claims, coming actually to the surface as an outcrop. It is not improbable that this and the small bodies found in the Adelaide mine are the replacement of isolated portions of the Blue Limestone, detached from the main body by the intruding porphyry. There is evidence also of considerable replacement action all along the contact of the White Porphyry with the Blue Limestone, both on the basset edges and on the upper surface of the latter.

In the Adelaide mine, as shown by the developments of the Ward and Adelaide shafts, the ore occurrence is extremely irregular. Lenticular bodies or pockets of sand carbonate are found between the White and Gray Porphyry and at the contact of the latter with the Parting Quartzite. Moreover, at the bottom of the Ward shaft a considerable body of vein material is said to have been opened in the lower Gray Porphyry, from which some silicates of copper were obtained. At the time of visit these workings were abandoned and could not be examined. The ore in general is carbonate of lead, with the usual gangue of iron oxide, but here rather silicious, as might be expected from the country rock. The masses of sand carbonate found in the Adelaide mine are remarkably pure, and have the appearance at a little distance of a white quartz sand. They contain, however, but little silver. A complete analysis of a specimen of one of these may be found in Appendix B, Table VIII. It contains about 95 per cent. of carbonate of lead, with a slight admixture of pyromorphite or chloro-phosphate of lead.

MINE WORKINGS.

The underground workings of this portion of Iron Hill are almost exclusively confined to the Argentine and Adelaide mines.

Argentine.—The Argentine mine is opened by the Camp Bird and Argentine tunnels and the Loker and Hynes shafts. The old workings on the Camp Bird claim are now mostly abandoned. The ore was found quite near the surface, resting on the Parting Quartzite, which is here 30 feet thick and contains no White Porphyry, as it does in the Argentine. This contact does not seem to have been followed in depth. Indeed, the geological relations of the ore bodies were so little understood in early times that no systematic exploration could be carried on.

In the Pine claim the main ore body was also found near the surface and above the level of the Argentine tunnel. It was afterwards traced down along the dip southward to a level 80 feet below the Argentine tunnel, then southeastward into the Adelaide claim, following nearly the line of the strike, but rising a little—that is, diverging to the eastward.

Argentine tunnel.—Ore is extracted through the Argentine tunnel, the Loker shaft being used simply for ventilation purposes This tunnel is over twelve hundred feet long, running first a little east of south and then bending to the west of south. It crosses the Adelaide claim, on agreement with that company, in order to explore the ground beyond. The geological structure, as exposed by this tunnel, was for a long time a complete puzzle to those who were working the mine, owing to the difficulty of distinguishing the different rocks from one other when bleached and altered. Even now a chemical test is often necessary. After passing through surface Wash the tunnel crosses the upper part of a body of White Porphyry into White Limestone. About seventy-five feet from the mouth a small gash vein in the porphyry, carrying galena, is said to have been found, upon which a winze was sunk. In the White Limestone a narrow sheet of bluish-gray porphyry is found before the tunnel enters the main body of Gray Porphyry, which at the contact is quite bleached by decomposition. Some iron-stained vein material is also found on the contact. Beyond, the tunnel again passes through White Limestone for 150 feet, another small sheet of porphyry being cut about midway in this distance. Parting Quartzite and

White Porphyry are then crossed, these being likewise very difficult to distinguish from each other underground. At the Blue Limestone contact, which occurs at the bottom of the Loker shaft, no ore is found on the tunnel level. A drift runs off to the eastward about one hundred and twenty-five feet from the Loker shaft, through which the ore stopes both above and below are reached. From the Loker shaft the tunnel runs for about five hundred feet in the Blue Limestone, which has an average dip of 15° to 20° to the southeast. Wherever raises have been made to the porphyry above, barren vein material has been found. This also reaches the tunnel level at times, following bedding or joint planes in the limestone. In one case a drift and winze have followed a considerable mass of vein material in the limestone, but without finding pay ore. Near the end of the tunnel the normal contact between limestone and porphyry is crossed. The limestone is here of lighter color, seamed with white calcite, and somewhat brecciated.

At the bottom of the Hynes shaft, with which the tunnel is intended to connect, drifts have been run upon the contact, disclosing some vein material. The dip of the formation is here shallower and to the southward. Above the contact are found quartzite and shales, belonging to the Weber Shale formation, between it and the White Porphyry, as in the Bull's Eye and Lime claims.

While the pay ore in this mine has been found at the contact, not of the Blue Limestone, but of the Parting Quartzite, it does not, so far as known, extend between these two formations. This would readily be accounted for on the theory that these ore bodies are the replacement of a portion of Blue Limestone left between the Parting Quartzite and the White Porphyry at the time of the intrusion of the latter.

Adelaide.—The ore bodies in the Adelaide mine are much more disconnected and irregular than in the Argentine. The mine has been mainly worked by a number of isolated shafts, and owing to complexity of the geological structure the ore bodies have not been systematically followed, so that, as many drifts were closed at the time of visit, the geological data are less complete than could be desired. The main difference in the formation between this and the Argentine is the occurrence of a later intrusive sheet of Gray Porphyry between the White Porphyry and the Parting

Quartzite, which seems to be somewhat irregular in form and of limited extent. The ore occurs both above this body, between it and the White Porphyry, and below it, or at its contact with the Parting Quartzite. A fragment of unreplaced Blue Limestone is also found resting on the Parting Quartzite and separating it from the overlying White Porphyry. This would seem to indicate the possibility that many, if not all, of the ore bodies are the replacement of similar fragments of limestone left by the irregular cutting of the porphyry.

The so-called Adelaide Discovery is a tunnel about fifty feet east of the smelter, adjoining Stray Horse gulch Here was an outcrop of three or four feet of hard carbonate ore, dipping 15° to 20° to the southeast and resting on Parting Quartzite. The tunnel, which starts in a little above the outcrop, ran into Blue Limestone, and a winze from its end is said to have struck the quartzite below.

The most important developments in the mine have been made in the Ward shaft This was sunk first through 170 feet of White Porphyry, which was much decomposed and for a considerable distance stained a brilliant red, apparently by anhydrous oxide of iron. At this depth was a layer of carbonate of lead, below which were 20 feet of decomposed Gray Porphyry and a second layer of ore resting on coarse-grained Parting Quartzite 15 feet in thickness. Below the quartzite was 20 feet of White Porphyry, and again five to six feet of quartzite, representing the balance of the Parting Quartzite formation. Below the quartzite the shaft passed through 75 feet of a very hard, jaspery material, consisting mostly of silica, with only about 5 per cent. of oxide of iron, which is a replacement of the White Limestone. When this material was freshly taken out it contained in seams and cavities a reddish gelatinous substance that resembled gelatinous silica in the process of deposition, which would indicate that the replacement of limestone and deposition of silicious matter are still going on. Below this the Gray Porphyry, very much decomposed, was penetrated to a depth of 57 feet. Almost all the decomposed iron-stained material taken from the shaft would assay one to five ounces of silver to the ton. At a depth of 330 feet from the surface the porphyry was impregnated for some eight feet with silicate and carbonate of copper; some red oxide and a little native copper were also found. The ore pockets found occurring above the Gray Porphyry consisted of remarkably pure white sand carbonate, free from admixture of clay and showing no galena. A complete analysis of a specimen taken from this horizon may be found in Appendix B. Those below the Gray Porphyry, whose connection with the ore body in the Argentine mine was afterwards traced, consisted of carbonate of lead, with iron-stained vein material, and in some cases, where the ore extended down into the quartzite, of unaltered galena.

The Adelaide No. 2 shaft found no ore at the contact of the White and Gray Porphyries. It was sunk through the lower ore body, the upper portion of the Parting Quartzite, the White Porphyry included in it, and into the lower body of Parting Quartzite.

The Terrible No. 2 shaft was sunk through Gray Porphyry, Parting Quartzite, White Porphyry, and Parting Quartzite again, into the White Limestone. No ore was found at the contact, and the White Limestone was not replaced, as in the Ward shaft, but was a crystalline rock with some decomposed iron-stained material at its upper surface, and with layers or lenticular bodies of white chalcedony throughout its mass, which are characteristic of this horizon.

The ore occurrence in these mines is distinguished from that of the majority of mines in this district by a total absence of manganese, a small amount of iron oxide, a relatively low tenor in silver, and a more frequent occurrence of gold, some of the fragments which occur in the quartzite being comparatively rich in this metal. The occurrence of copper ore in the Gray Porphyry is also exceptional, the nearest analogy being the body in the Little Johnnie and Uncle Sam, on Breece Hill, overlooking South Evans gulch.

Double Decker.—On the north side of Stray Horse gulch, opposite the Argentine, are the two shafts of the Double Decker mine, which have obtained from the Lower Quartzite a certain amount of gold ore. At this point all the overlying strata have been removed by erosion and the Lower Quartzite forms the rock surface. Both shafts have been sunk in this formation, and one of them has passed through it into the underlying crystalline rocks of the Archean. Neither was being worked at the time of examination, consequently no detailed information could be obtained, nor were any data as to amount or value of ore extracted available.

CHAPTER III.

CARBONATE HILL GROUP.

GENERAL STRUCTURE.

The geological structure of Carbonate Hill[1] is very similar to that of Iron Hill in that it is formed by a series of easterly-dipping beds broken on the west by a line of faulting or displacement. Outcrops are also exposed on its southern face by the erosion of California gulch, but in a less complete series, owing to its being shallower and proportionately wider, in consequence of which the bounding slopes are less steep and more thickly covered by surface débris. The fault is nearly parallel to that of Iron Hill, and, like it, merges into the axis of an anticlinal fold on the north. In the southern half of the hill, however, the movement of displacement is distributed in part to a second nearly parallel fault a short distance to the west. Of the southern continuation of these faults less satisfactory data are available, but they are supposed to merge together before crossing California gulch, and probably pass into an anticlinal fold under the Lake beds to the southwest, like the Dome fault, the normal continuation of the Iron fault. As on Iron Hill, there is also evidence of a basining-up of the beds of the relatively down-thrown mass on the west as they approach the fault; in other words, of a synclinal structure. Upon this evidence, which will be given later in full, depends the solution of the important question whether ore bodies exist under the present site of Leadville or not.

Rock formations.—The series of beds of which the hill is composed is essentially the same as that given in the Iron Hill section, but the distribution of the later intrusions of Gray or Mottled Porphyry differs somewhat in detail.

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[1] See Atlas Sheets XXVIII, XXIX, and XXX.

Where these cross the beds, either as dikes or sheets, there is a noticeable enrichment of the ore bodies. One main sheet of Gray Porphyry is found at or near the base of the Blue Limestone, which apparently cuts up to a higher horizon in different portions of the hill. A second sheet is found in White Limestone in California gulch, as shown on the map; but as none of the underground workings has penetrated as yet to this depth, there is no evidence to show whether this is a distinct sheet or merely an offshoot from the main body.

Carbonate fault.—The movement of displacement by faults on Carbonate Hill is considerably less than on Iron Hill, though its total amount cannot be definitely determined. As in the case of the former, the movement decreases to the north and the fault gradually passes into an anticlinal fold. In the southern portion of the area represented on the map this movement is distributed between two faults, the Carbonate and the Pendery. The Carbonate fault here runs nearly on the dividing line between the Carbonate and Ætna claims, cutting across the extreme southwestern corner of the former and the northeastern corner of the latter. It is proved in the No. 5 shaft of the Ætna claim, and in the Meyer shaft, which has been sunk following its plane till the contact on the west side was reached. As here shown, it stands with an inclination of about 60° west, shallowing somewhat in depth and having a movement of displacement of only about two hundred and fifty feet. The hanging wall has smooth and clearly defined slickensides surfaces, while the limestone in the foot wall is somewhat altered. The plane of the fault is occupied by selvage material, which is slightly impregnated with chloride of silver and contains occasional fragments of ore. The plane of the Carbonate fault has also been cut in the lower shaft of the Yankee Doodle claim. Beyond that point to the northward it has not actually been proved, and is located simply by discrepancies of level between adjoining underground workings. There is some reason to assume that to the northward, in the Waterloo claim, the movement of the fault has become nil, or is even reversed, that is, that there is a slight downthrow to the east, as shown in the Henriett-Waterloo section, Atlas Sheet XXIX.

Pendery fault.—A short distance west of the Glass shaft, a second fault, apparently nearly parallel and having the same angle of inclination with the Carbonate fault, cuts off the limestone, no explorations west of this line having reached below the White Porphyry. Its probable continuation has been traced southward to a connection with the Carbonate fault, and northward through the Washburne and St. Mary workings, where it appears to be accompanied by minor faults and folds, into a probable anticlinal fold running north through the Niles-Augusta, and then northwestward, opposite the Half Way House claim.

Morning Star fault.—In the workings of the Morning Star mine a small fault is found (shown in Section C, Atlas Sheet XXIX), in which the downthrow is to the east. It is probably only local in character and corresponds to the sharp bend in the beds observed in the Evening Star and Catalpa. It was only observed at one point, and is not therefore indicated on the surface maps, as its direction would be purely hypothetical.

Ore deposits.—The materials composing the ore deposits of Carbonate Hill are essentially the same as those of Iron Hill; they may perhaps be said to be poorer in bases of iron and manganese and proportionately richer in silica; therefore less favorable for the smelter; but this characteristic is rather one to be confined to individual mines or parts of a mine than applied in a general way. Silica occurs less frequently as chert and more commonly as a very finely granular and somewhat porous quartz rock than on either Iron or Fryer Hill. The ore is either galena or its secondary products, carbonate of lead and chloride of silver. In one instance native silver has been found. Dechenite, or the vanadate of lead, has been detected in ore from the Evening Star and Morning Star mines by Dr. M. W. Iles.[1]

Exceptionally good opportunites are offered for observing the action of replacement and the gradual passage from dolomite into the earthy oxides of iron and manganese. The workings not yet having reached the great distance from the surface that they have on Iron Hill, no such definite evidence is found of decrease in the action of surface waters producing oxidation and chlorination of the original deposits. The limit of the zone of oxidation would, moreover, be expected to be farther from the surface on account of its lower altitude.

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[1] American Journal of Science, May, 1882.

Southwest slope of Carbonate Hill.—The detailed map of Carbonate Hill (Atlas Sheet XXVIII), while including the principal mines, covers only the northern portion of the western slopes. For the geology of the southern portion, in which as yet no ore bodies of importance have been developed, reference must be had to the general map of Leadville (Atlas Sheet XIV), on which the limits of the detailed map are indicated. South of these limits the Prospect incline, the Rosebud and Deadbroke tunnels, on the north side, and the Jordan and Swamp Angel tunnels, on the south side of California gulch, follow the upper surface of the Blue Limestone. While this surface shows evidence of mineralization in the characteristic iron-stained material generally found at the contact and in the frequent occurrence of the so-called Chinese talc, supposed to be the product of the alteration of the porphyry by mineral solutions, it has thus far been found comparatively barren of rich ore bodies. In the Prospect incline, about three hundred and seventy-five feet from the mouth, the surface of the limestone, which had hitherto been wavy, as it is generally found, suddenly drops down almost perpendicularly for 125 feet; 60 feet farther on in the line of the incline, however, the limestone is found following its normal dip to the eastward. This is apparently a very sharp fold in the limestone, accompanied by a certain amount of faulting. It is approximately on the line with the sharp fold which will be described hereafter as running through the Carbonate and Yankee Doodle mines, but the accumulation of rich ore, which in these mines is found above the fold, is here wanting.

The only actual rock exposure in this region is that of the White Limestone in the quarry on the north side of California gulch. From a prospect hole sunk by placer miners in the bed of California gulch, above the flume, casts of a *Rhynconella*, in a sandy white limestone, are obtained.

Of the two porphyry sheets which are indicated here, the upper one, near the base of the Blue Limestone, has been traced from the Irish Giant shaft to the Silver Star, and from that, a little below the John Harlan, across the gulch to the Logan and Broadway shafts, on the south side. This is evidently the same sheet which to the north occurs near or at the base of the Blue Limestone and in the Henriett and Waterloo claims cuts up across it into the White Porphyry. The lower sheet of porphyry occurs in the

White Limestone and is best exposed in the bed of California gulch. It does not seem to extend far up on Carbonate Hill, as it was not struck in the O'Donovan Rossa shaft. Parting Quartzite, which forms the upper part of the Silurian limestone, is here very coarse-grained.

The evidence for determining the line of the Carbonate fault in this region is not very plentiful. A shaft and drill-hole have been sunk on the northern bank of California gulch, south of the Harrison smelter, to a depth of 200 feet in White Porphyry, and are thus evidently to the west of the fault. The Blind Tom shaft, on the road south of the California tunnel, which was sunk 130 feet in White Porphyry, is also west of the fault; while another shaft, 50 feet south of this, was sunk in Silurian limestone, and is hence east of the fault. The California tunnel (T–48) has been run into the hill about seven hundred feet in a direction E. 15° S., or magnetic east. The first 585 feet it is in White Porphyry, from which it passes suddenly into the Blue Limestone, across a clay selvage. This is supposed to be the line of the Pendery fault, although the average dip is only 30° to the westward, and it might possibly be supposed to be a folding of the limestone downward in that direction. Unfortunately, the bedding planes are not sufficiently distinct at this point to determine the question of a westerly dip. There seems to be little doubt, from the extensive slickenside surfaces, that, even if there be a westerly dip, there has also been considerable faulting movement. The tunnel runs for the rest of its extent in Blue Limestone, in which, toward the end, the bedding becomes quite distinct and the dip assumes the normal angle of 20° to the eastward. It is evident from its lower position relatively to the outcrops of the Blue Limestone on the hill above that this is a continuation to the south of the portion of that body in the Ætna claim which is west of the Carbonate fault and between it and the Pendery fault. The line of the Carbonate fault has therefore been drawn on the map according to this indication. It was observed that the timbers supporting the roof of the tunnel from its mouth to the limestone (which, owing to the soft and yielding character of the porphyry through which it runs, were placed exceptionally close together) had a slight and uniform inclination of 5° from the perpendicular to the west. It is not to be supposed that they were originally placed in this position, and the inference is therefore justifiable that the whole mass of rock above the tunnel has had a slight movement to the westward since the tunnel was run.

The underground workings of Carbonate Hill may, for convenience of description, be divided into two groups:

1. A southern, including the Carbonate, Little Giant, and Yankee Doodle claims, to the east of the main fault, and the Ætna, Glass-Pendery, and other claims, below or to the west of it.
2. A northern group, including the Crescent, Catalpa, Evening Star, Morning Star, Waterloo, Henriett, and adjoining claims.

SOUTHERN GROUP OF MINES.

The description of Carbonate, like that of Iron Hill, will commence with the southern end, reversing the order in which the sections are lettered, because the claims at this end were first opened and because the geological structure is more clearly and easily shown in their workings. In the Carbonate, Shamrock, Little Giant, and Yankee Doodle claims, east of the fault, the principal developments have been made on what is practically one ore body, running in a northeasterly direction from its outcrop on the Carbonate claim. A noticeable feature of the structure is that this ore body is bounded on the southeast by a prominent fold in the limestone, which bends down very sharply east and, rising again, forms a narrow trough. This is clearly shown in the Carbonate incline, Section I, Atlas Sheet XXX. The region to the southeast of this fold has thus far proved barren of rich ore, although explorations have hardly been carried out to a sufficient depth to warrant the conclusion that another bonanza may not exist in that direction. In the Yankee Doodle and Little Giant claims the ore body is narrow, but widens out as it approaches the surface in the Carbonate ground, with intermediate barren streaks which approximately correspond to minor folds more or less parallel with the main folds above mentioned. Practical evidence of the actual replacement of limestone by vein material is extremely common and well defined in the mines of Carbonate Hill. These, as will be shown in the detailed descriptions of mines which follow, are found in the sudden deepenings of the ore bodies on the limestone side of the contact plane, which by the miners are often confounded with actual waves in the limestone itself. Careful examination in such cases discloses the fact that the contact line is either not curved at all or only to a comparatively limited extent, while the ore impregnation extends downward abruptly to a very considerable depth, sometimes sixty or seventy feet, into the body of the limestone, from which it is separated by a transition zone, barren in general of pay mineral and consisting of limestone more or less impregnated with oxides of iron and manganese.

The first ore discoveries on Carbonate Hill were made on the ground of the present Carbonate claim in the vicinity of the Old incline. The claim forms part of the property belonging to the Leadville Consolidated Mining Company, which owns as well the adjoining claims of the Shamrock and West Shamrock, and also, by a recent consolidation, the Little Giant. For purposes of description the adjoining claim of the Yankee Doodle will be considered as forming part of this group, as the workings of this mine are connected with the others and the ore bodies in all these different claims are practically continuous. They are situated on the western slope of Carbonate Hill, midway between its steeper southwestern and more gentle northwestern inclinations. In the southern portion of the Carbonate claim, as will be seen by reference to the map, by the divergence of the line of fault from the line of contact of the upper surface of the limestone and porphyry, a zone of limestone, widening to the southward, is exposed beneath the slide. The actual position of the main fault line has not been traced beyond the No. 5 shaft of the Ætna mine. Its exact position to the south of this point is therefore somewhat hypothetical, the general direction being given by the developments of prospect shafts on the south slope of the hill, beyond the limits of the map. The actual outcrop of the ore body along the southern line of the Carbonate claim is also somewhat difficult to define, there being here one of the slight flexures in the limestone, of which mention has already been made, whose axis is at an angle with the fault line, crossing it somewhere in the neighborhood of the new (Meyer) shaft of the Ætna and from that point diverging to the southward. In the workings of the Ætna mine, which were parallel with and contiguous to the side line of the Carbonate, the ore body is said to have been practically horizontal on an east-and-west line for a short distance, and even to have had a slight inclination to the westward as it approached the fault. These workings, being now abandoned and partially filled up, could not be explored. Similar conditions are said to have been observed in the development of shaft No. 12, near the Carbonate line, while in the Carbonate ground itself, at the Combination incline, the outcrop of easterly-dipping beds comes practically to the surface at the mouth of the incline, and a drift, now closed up, running westward from the mouth of the incline, is said to have followed the ore body down on a westerly dip. The conditions of this outcrop have been thus fully described because it has been the cause of a long and expensive lawsuit between the Ætna and Carbonate mines. The owners of the Ætna claimed that they had the outcrop of the vein, or, in legal terms, "the apex," within their side lines. The owners of the Carbonate, on the other hand, maintained that theirs was the legal outcrop, inasmuch as the ore was found at a higher level within their side lines, and that they therefore possessed "the apex" of the vein, which is legally defined as that portion which is nearest the surface. The rulings of the judge were first favorable to the construction of the Carbonate, but were afterwards reversed when it was proved by actual measurement that ore was found within the Ætna claim one inch higher than at a corresponding point within the Carbonate.

Carbonate workings.—The southernmost workings on the Carbonate property consist, first, of the West Shamrock incline, which lies outside the limits of the map, commencing at a point near the southeast corner of the Carbonate claim, 240 feet south of the side line of the map, and running parallel with that side line 160 feet, inclining into the hill at an angle of 18°. This incline has developed no considerable amount of ore and is interesting only as showing the character of the formations along the contact line at this point. The mouth of the incline is opened in a pulverulent, blue limestone, with a floor of chert covered by a thin seam of iron-stained clay. The actual contact of the limestone with porphyry is probably above the line of the incline, which follows down on the chert floor for some distance, when it passes into solid blue limestone, more or less decomposed, or in which no bedding planes can be distinguished. Small seams, from one foot to two feet in thickness, of decomposed porphyry or clay material are found traversing the limestone, and half-inch seams of iron-stained clay on the cleavage planes. A shaft has been sunk on the hill above, which would connect it with the end of the incline, had both been continued a short distance farther. As an exploration for ore, therefore, the work done here is imperfect, inasmuch as, although no pay ore has been discovered, it does not definitely prove that it does not exist in the neighborhood. The occurrence of these bodies of black chert, which are rarely found on the northern portion of the hill, are not uncommon along the southern slopes. While by the miners in many parts of the district they are considered good indications of ore, this empirical test is by no means infallible, although an evidence of passage of silicious waters. One hundred and fifty feet below and west of the mouth of West Shamrock incline a perpendicular shaft (T–35) has been sunk to a very considerable depth on the Irish Giant claim, through Gray Porphyry into Blue and then into White Limestone, without developing any important ore bodies, though an impregnation of the limestone along the lower surface of the porphyry might not unreasonably be expected.

The next opening to the north is the Combination incline, just north of the mine offices. This incline has an average angle of 21°, flattening out a little in the upper ten feet. The limestone comes practically to the surface at the mouth of the incline, being covered with about nine feet of broken material or slide. The workings of the incline were abandoned at the time of visit, and apparently no considerable amount of ore had been extracted from them. Down to the first level, the section afforded by the incline itself, which cuts the contact between the limestone and the porphyry, shows a wavy outline of the latter. In the first drift north, the porphyry is considerably iron-stained and decomposed above the limestone, immediately adjoining which is the usual parting of Chinese talc. At 24 feet from the incline are old stopes, now filled up, in which the formation dips downward to the north. The first drift south cuts through the crest of short waves in the limestone and a body of iron from six to eight feet in thickness, while in the face the porphyry goes down with a steep dip to the east and south. Of the south drift, on the second level, the first 20 feet are in limestone, more or less replaced by iron oxide, which is succeeded by porphyry.

The north drift, on the second level, follows in its curves approximately the line of contact between limestone and porphyry, bending back to the westward nearly under the end of the drift on the first level, showing that a limestone ridge crosses the incline in a diagonal or southeast direction between the first and second levels. Below the second level the incline follows along the edge of the northern side of this limestone ridge for 25 feet, the north face of the incline being partially in limestone and clayey contact material, dipping sharply to the north, and the south face in solid limestone. Beyond, the limestone dips down to the east, and the rest of the incline is in porphyry, which is more or less iron-stained. At the end of the incline is a winze sunk 40 feet to the surface of the limestone, showing a rapid descent of the limestone to the northeast of the ridge, which has just been passed through, which is further evidenced by the appearance of bedding planes in the porphyry itself, which incline at an angle of 45° to the northeast. The drift running westward from the mouth of the Combination incline, which is said to have followed the contact on its slope to the westward, was not accessible at time of visit.

The Carbonate Old incline runs in 170 feet at an angle of 19° and presumably follows the contact, but it is now closed. Solid limestone is found eight feet from its mouth, covered by iron-stained vein material and broken porphyry. Between this and the mouth of the Main incline, under where the boarding-house now stands, was the ore body from which ore was first taken on this ground. The drifts are now filled up and abandoned, but it is evident that the body was of considerable size and very near the surface, lying in an approximately horizontal position, that is, near the crest of the fold already mentioned.

Carbonate incline.—The principal workings of the Carbonate mine are opened by the Main or Carbonate incline, which descends into the hill for a distance of 620 feet at an angle of 21° 30′ and in a direction E. 25° S. It is one of the comparatively few inclines in the district which have been driven straight, instead of following the irregularities of the limestone surface, the only true system for an incline from which it is expected to extract any considerable quantity of ore, and one which is probably more economical in the long run, since, in spite of the irregularity of the limestone surface in limited distances, the average dip is tolerably constant.

The section afforded by this incline is extremely interesting, as showing the irregularities of the limestone surface caused by undulations or slight flexures in the formation, and these are best seen in the graphic illustration afforded on Atlas Sheet XXX, Section I, which has been very carefully constructed from actual measurement. It will be seen by reference to this section that the original surface of the limestone presents a general wavy outline with one prominent fold, which, contrary to the general rule prevalent in the major flexures in the region, has its steeper side to the east. The deciphering and reconstruction of the original folds in the formation is a matter of some little delicacy, since the ore currents have eaten irregularly into the mass of the limestone, and the porphyry itself is also somewhat altered and mineralized. It is evident, however, that the dividing line between ore and limestone is one which must be entirely rejected for this purpose. The parting between ore and porphyry, on the other hand, in spite of occasional incursions of ore material into the mass of the porphyry, is practically much more definite, and is that which has been used in determining the points in the original surface of contact. One prominent fact to be observed in this section, and one which seems capable of a certain amount of generalization, is that the main rich ore body is found adjoining the crest of the prominent wave or fold in the limestone. It would seem that along the line of this sharp fold, which may very possibly have been accompanied by a slight displacement, there was an interruption in the ore currents, as is further evidenced by the fact that beyond the fold on the east side for a very considerable distance, indeed, to the extent of the present developments, there has been no considerable deposition of pay ore, the mineralized zone consisting of a most irregular replacement of the limestone by what is known to the miners as black iron, a mixture of oxide of manganese with clayey material, which passes by almost imperceptible transition into coarsely-crystalline black limestone. In this lower part of the incline is one of the most striking evidences of the fact that the ore deposit is an actual replacement of a limestone in place, the walls of the incline showing the clayey, ferro-manganiferous material penetrating irregularly into the limestone, now in thin, sheet-like bodies, following a cleavage or fracture plane and terminating in a point, and now replacing the whole mass for a distance of many feet. The line between the unaltered and replaced limestone has been accurately followed foot by foot, and a portion of the north wall of the incline is represented on a larger scale in Plate XXII, Fig. 3.

As well as can be seen through the timbers of the incline, its upper portion is practically in the porphyry and the limestone does not actually outcrop. Very probably this is the crest of a slight roll, and it is practically level at this point or may possibly have a very slight western inclination. The first limestone actually seen in place is about sixty feet from the surface, just above a drift which crosses under the floor of the incline. It is evident, however, that the limestone rises somewhat to the south, since it is said to have been found at six feet from the surface under the boarding-house. From this point downwards to the second level the incline is run practically in contact material or in decomposed iron-stained porphyry.

In the second level south, about fifty feet from the incline, a limestone floor is found rising rapidly to the eastward.

In the third level ore was found below the level of the incline, evidently in a depression of the contact line, but the workings are now closed.

The fourth level south, which is still open, runs at first in porphyry, a cross-cut to the left or eastward cutting the underlying limestone at a distance of 16 feet. This marks the point where the limestone rises towards the crest of the fold, shown on the incline section. While in this section the crest is comparatively narrow, it apparently widens out to the south in a broad, dome-shaped elevation. At about forty feet from the incline the fourth level cuts through a body of hard, compact iron, resting upon the limestone, and then bending to the eastward passes into the solid limestone. At a distance of one hundred and twenty-five or one hundred and thirty feet from the incline a streak of clayey matter was cut, carrying a little ore and running off to the southward, evidently a replacement body which followed a cleavage or fracture plane in the limestone. The drift then passes again into solid limestone, bending off to the south at its extremity, where it passes into porphyry, the limestone dipping sharply to the eastward at the bend. As will be seen by reference to the map, the drift at this point is nearly over the extremity of the south drift on the eighth level.

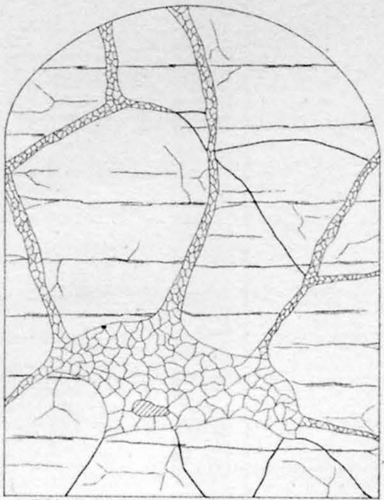

FIG. I. EVENING STAR INCLINE

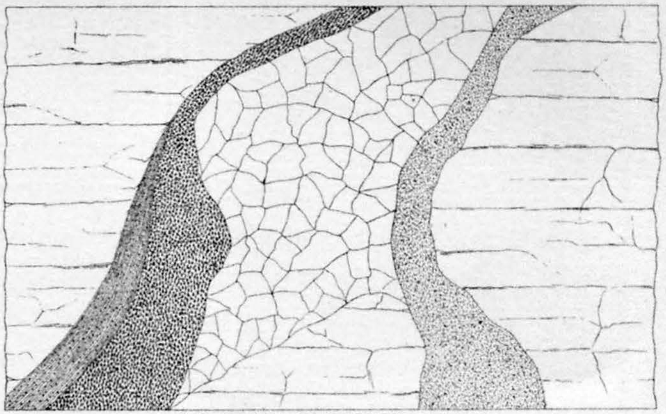

FIG. II. GLASS-PENDERY MINE

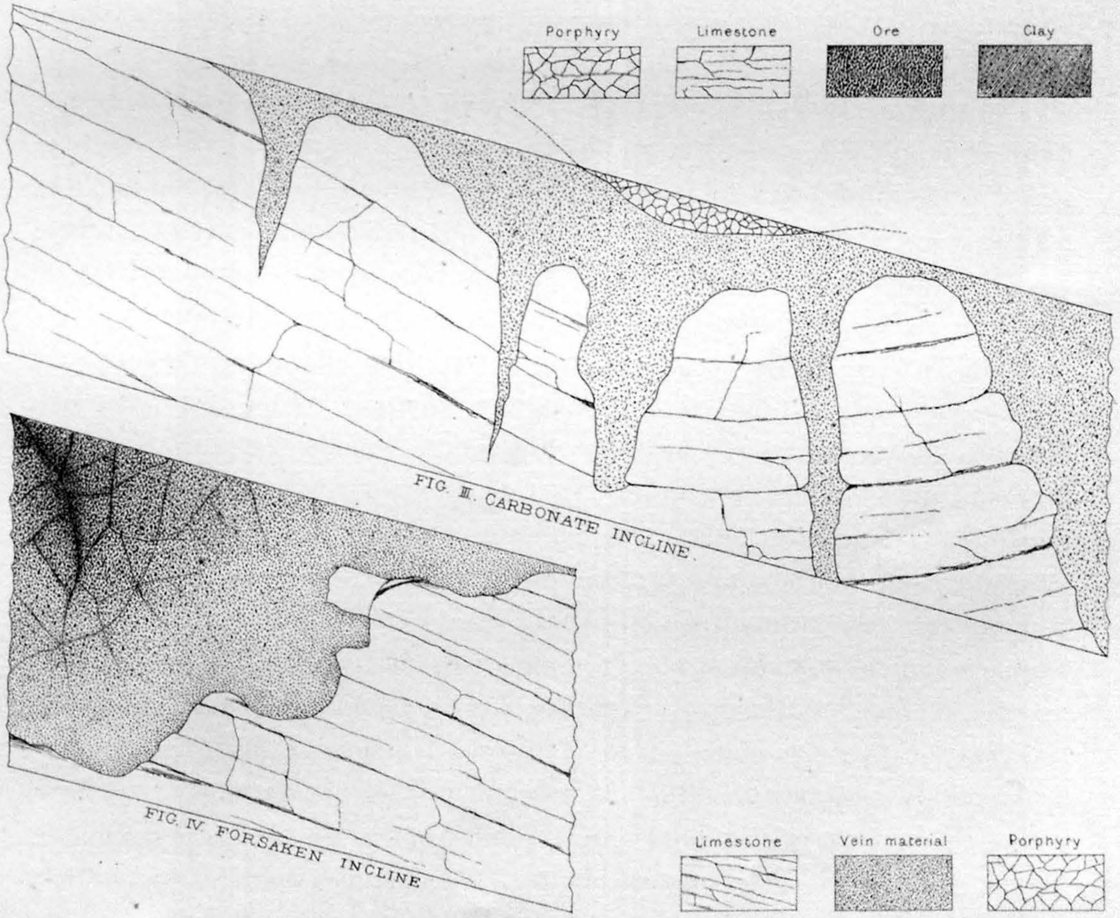

FIG. III. CARBONATE INCLINE.

FIG. IV. FORSAKEN INCLINE

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

VEIN PHENOMENA.

The fifth level south has only been driven about fifteen feet, being in the limestone, which extends at least that distance above its floor.

The sixth level south follows in its windings the contact of porphyry and limestone, which here dips steeply to the east and thus defines the eastern edge of the fold. The contact material consists of a thin streak of manganiferous clay and Chinese talc. At 30 feet from the end of this drift is an upraise on the right, 30 feet in height, following the almost perpendicular surface of the limestone, from the top of which a prospecting drift has been run out and is said to have been nearly connected with the drift on the fourth level; the drift beyond this point is in the body of the limestone.

On the seventh level no drift has been run to the southward, the incline being here entirely in porphyry.

The south drift on the eighth level has a general southwest course and is cut in a decomposed material which seems to result from the decomposition of the overlying porphyry and of the thin bed of quartzite which is frequently found between the porphyry and the limestone. It is a reddish clayey mass, having a gritty feel and containing fragments both of black chert and of quartzite. Within about twenty feet of the point where the drift forks the clay rises suddenly to the roof and the drift passes into a light-blue decomposed limestone of the pulverulent type, so common in the district, which crumbles between the fingers. The further continuations of these drifts are entirely in a body of fine-grained limestone of earthy texture, as distinguished from the crystalline, granular dolomite which is most frequently found. In the left-hand drift is a stope one set high and in the right-hand drift there is a large chamber from ten to fifteen feet in height and breadth and some twenty or thirty feet long, from which limestone has been quarried to be sold to the smelters as flux.

The south drift on the ninth level follows in its curves the contact between limestone and porphyry, having some black chert in the floor.

Of the workings on the north of the Carbonate incline those on the first and second levels are now mostly inaccessible. As well as can be ascertained, a thin sheet of ore extended from the incline, somewhat intermittently interrupted, as a rule, by limestone rising to the contact with the porphyry, whether caused by simple undulations in the limestone itself or by narrowings of the ore body cannot be determined. Pay ore was mostly in the form of sand carbonates, slightly impregnated with iron oxides, and in comparatively thin sheets, having a maximum thickness of about two or three feet. Connection was made from the second level to the end of the Ætna incline, to prevent the owners of the latter mine from following their ore shoot farther into the Carbonate ground, but is now closed up. Between the second level and the third and fourth is a large piece of comparatively barren ground, as indicated by the outlines of the ore body given on the map. These outlines, it must be borne in mind, are, from the necessity of the case, considerably generalized, it being impracticable to obtain accurate data with regard to their definite limits in the present abandoned condition of the workings.

One of the richest ore bodies in the mine was found immediately adjoining the incline on the north, between the third and fifth levels, occupying a slight depression immediately adjoining the crest of the main fold in the limestone. This ore body was in places two sets high (some fifteen feet thick), gradually thinning as it approached the crest of the fold, and to the northeastward spreading out in a thin and comparatively continuous sheet of sand carbonates. In the drift on the fourth level north, a body of unaltered limestone rises up in the floor, cutting off the ore about twenty feet from the incline; beyond this point ore is only found in detached bodies to the west of the drift, while to the east it extends in practically continuous sheets to the lower levels, gradually taking a steep easterly dip. The northern extension of the drift, after passing through a considerable stretch of barren contact, cuts several small, unimportant ore bodies, which form a continuation of those developed in the lower levels, and at its extremity was being run at the time of visit in solid limestone.

The fifth level is cut for a short distance from the incline in the limestone of the fold, then passes through the clay contact, showing considerable slickensides. The contact here is somewhat typical, showing first a blackening of the limestone by a certain amount of impregnation by manganese and iron oxides; above this, iron-stained clayey matter, containing the usual development of Chinese talc; and still above, a thin irregular body of quartzite (Weber Shales), followed by normal White Porphyry. The slickensides indicate a certain amount of movement or displacement along the steep eastern face of the ridge or fold. At about one hundred feet from the incline the drift bends under that of the next level above, and from here on follows in its sinuosities the line of contact, which has a steep dip to the eastward. From the point where the ore is struck in this drift almost to its northern extremity, the ore extends eastward to the two next lower levels and into the ground of the Little Giant claim, in a practically continuous body of iron-stained sand carbonates, generally one foot to two feet in thickness, only interrupted here and there by ridges of undecomposed limestone, reaching up to the porphyry contact. This ore sheet has a steep dip to the eastward, say on an average from 25° to 30°, and evidently stands on the eastern slope of the fold; it would seem, therefore, that the dome-shaped uplift in the limestone body between the fourth and sixth levels south of the incline has narrowed into a sharp ridge in the incline itself, and that it has then become a monoclinal fold to the north, the ore-body extending over the line of its crest and partly down its eastern slope into the trough.

The sixth level north is run for fifty feet on a barren contact, following the curve of the limestone ridge, then passes into an ore body extending right and left with a steep dip to the eastward, while the end of the drift bends sharply to the eastward and runs into the overlying body, in which the eastward dip is shown in the rude bedding planes.

The seventh level north for the first hundred feet is run in White Porphyry, more or less iron-stained or decomposed and probably some distance above the contact; limestone then comes up into the floor, covered by a streak of black iron and white Chinese talc a foot in thickness. Immediately beyond, a body of ore three sets high has been stoped out, which evidently represents the ridge of limestone, now entirely replaced by ore. This ore ridge at its maximum height is only about twenty feet in width, having a roof of breccia material consisting of fragments of black chert and quartzite in a matrix of clay and decomposed porphyry. The limit of the pay-ore body is found at a comparatively short distance east of this drift, the developments on the eighth level showing only a barren contact.

The north drift on the eighth level follows in its right fork the eastern boundary of the trough east of the ridge, and is cut in a yellow ocherous material, which occasionally passes into limestone, the roof being White Porphyry. The left-hand or west fork in the first half of its course is cut in White Porphyry and then passes into a decomposed, light-colored, and iron-stained limestone, which rises to the west. The drift, after a rise of 15 feet over a ridge of dark hard limestone, passes into soft clayey material, and then bending to the northward passes through two or three feet of quartzite into a body of Gray or Mottled Porphyry, in which it stops. This is the only point in which the Gray Porphyry has been discovered in this mine. No definite idea, therefore, can be formed as to the shape or origin of the body. It may be interesting to note in this connection, however, that farther in the hill, as shown by the developments of the Modoc shaft, there are several bodies of this Mottled Porphyry intercalated between the different beds of limestone. From this level, immediately adjoining the Main incline, an incline drift, following the contact, has been run westward to a point immediately under the fifth level. It is now mostly filled up, and is mainly interesting as proving the existence of the deep trough shown in the section.

From the eighth to the ninth level, a distance of over one hundred and twenty feet, no pay ore is found; but, as has already been mentioned, most interesting proof is afforded of the fact that the ore bodies are simply replacements of the limestone. The replaced material is largely a black, clayey matter, more or less iron-stained and in some cases passing into a red plastic clay, which would seem to have infiltrated into the mass from the porphyry above, while the limestone immediately adjoining this is itself more or less discolored by black oxide of manganese. A short distance below the intersection of the ninth level, which is in solid blue limestone, the limestone, with its contact seam of Chinese talc, dips steeply down to the eastward and the incline passes into porphyry. Just at the point where the limestone bends downward, a small ore body was discovered immediately over the roof of the incline.

Little Giant.—The workings of the Little Giant mine are practically an extension of those of the Carbonate along the continuation of the main ore sheet, where it has its steep dip to the eastward between the fourth and seventh levels. It is opened by a shaft 234 feet deep and is also connected by narrow sinuous drifts with the fifth level of the Carbonate. The ore is iron-stained sand carbonates, occurring, as a rule, in thicker bodies than in the Carbonate claim and being rather more irregular in shape. The workings consist of a main incline, now being driven a little south of east from the bottom of the shaft, and of an incline running nearly due east at an angle of 26°, from which levels have been run off at intervals of about fifty feet. The richest ore bodies have been found immediately adjoining the Carbonate claim on one side and that of the Yankee Doodle on the other. The larger portion of the ground included in the Little Giant claim, which lies to the southwest of the Shamrock, has not been prospected at all, because it has been considered beyond the southeastern limit of the pay-ore streak, as indicated on the map. The ground is probably barren.

Yankee Doodle.—The Yankee Doodle mine, east of the Carbonate fault, is opened by two independent shafts, not yet connected underground, and the present developments are confined to the workings from the upper shaft. The ore found here is a northeastern continuation of the Carbonate body. The upper shaft of the Yankee Doodle is 303 feet deep, the first station being at a depth of 296 feet from the surface. Limestone is said to have been struck in this shaft at 230 feet. The main drift, running eastward from the station, is cut in solid black crystalline limestone, showing some replacement action; but no pay ore is found until the first cross-drift is reached, at a distance of 170 feet from the shaft.[1] Twenty-five feet beyond this first drift is a winze 30 feet deep, which goes down at an angle of 70° to the eastward, following a sudden bend in the limestone, which is evidently the same that has been traced through the Carbonate and Little Giant claims, though having a steeper angle. Some ore has been found beyond the bend on either side of the main drift, but none as yet below the winze. The principal ore developments have occurred along the boundary line between this claim and the Little Giant, where an incline is being sunk at an angle of 30°. Considerable ore was found extending upwards along the surface of the limestone near the boundary line. These developments show a somewhat discontinuous body of pay ore, consisting of sand carbonates

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[1] By an oversight in the correction of proof the section (Section G G, Atlas Sheet XXX) shows White Porphyry, instead of vein material, below this drift from this point to the winze.

with a small proportion of unaltered galena. The ore sheet which extends from the Carbonate into this ground seems to be growing narrower and less continuous and may disappear entirely to the northeast. It is as yet, however, not safe to assume this as a fact, for the ground in the Yankee Doodle claim has not yet been thoroughly and systematically prospected; moreover, the shaft sunk on the Excelsior claim, which is nearly in the line of the probable continuation of this ore body, is said to have cut a large body of vein material, which, though not rich, is an evidence of mineralizing action.

From the middle shaft of the Yankee Doodle, which is about one hundred feet in depth, considerable ground has been explored by an incline, from which two sets of levels have been run off Both incline and levels are very irregular, following the varying inclination of the contact between limestone and porphyry.

From the end of the first level north a winze was sunk, apparently in a considerable body of vein material, in a sudden steepening of the limestone. As it will be seen by reference to the map and sections that this point is on a line with a similar sharp bend in the limestone, at the southern extremity of the drift on the sixth level of the Crescent mine, where indications of a body of rich ore are found, it would seem advisable to have pushed explorations farther at this point, with the prospect of developing one of the smaller bodies of ore, such as are found to the west of the main Carbonate body, between it and the Ætna line. The south drift on this level follows a barren contact of the usual character, viz, showing blackened decomposed limestone, with the usual parting of Chinese talc separating it from the iron-stained porphyry above, which contains fragments of black chert and quartzite.

On the second level some thin seams of sand carbonate are found to the north of the incline, and at the southern extremity a considerable body of this ore has been stoped out. This also is evidently worthy of further exploration. It thus appears that the apparently barren zone between the two rich ore bodies or bonanzas of Carbonate Hill is only a region where the ore is less continuous than in the bonanzas themselves, and that several small ore bodies have been found there, and probably by a systematic exploration of the ground many others might be found, especially if the search were not confined to the contact surface alone, but were also extended where indications of mineralized jointing planes seem to warrant it, into the body of the limestone below.

Area on top of hill.—In the area to the east of the claims of the southern group, on the top of Carbonate Hill, several prospecting shafts have been sunk in the White Porphyry, notably the Excelsior, William Wallace (300 feet), Tip Top (297 feet), Little Nell (440 feet), Thespian (400 feet), and the Modoc (600 feet). Of these, only the Excelsior and Modoc have reached the Blue Limestone. Neither was accessible at time of visit. The Excelsior is said to have found a heavy body of iron-stained vein material, but not sufficient pay ore to encourage further developments. The Modoc shaft had filled with water while awaiting better pumping machinery, but from data obtained from miners and from evidence afforded by the dump it appears that contact with a certain amount of vein material was struck at about five hundred feet. In the little drifting that was done no rich ore was found. The shaft was sunk 100 feet farther, passing through the Blue Limestone and two intrusive sheets of Gray Porphyry included within it, as ideally shown in Section I, Atlas Sheet XXX. It would thus appear that the sheet of Gray Porphyry, which on Iron Hill occurs near the base of the Blue Limestone, is here either cutting across this formation to a higher horizon or sending off offshoots, such as are found in some of the Iron Hill workings.

Area west of Carbonate fault.—In the Ætna and Glass-Pendery claims, which, with the exception of a small corner of the former, lie west of the Carbonate fault, little pay ore has been found on the contact; but the main ore bodies occur within the mass of the limestone, extending to a depth of fifty feet or sixty feet below its surface.

Ore was first discovered in a nearly vertical body, crossing the lower part of Glass No. 2 shaft in a southeast direction. The development of this body led to the discovery of other larger and more irregularly-shaped bodies extending beyond the Ætna line. These were worked by the Glass-Pendery owners, and much ore was taken from the Ætna ground before the owners of the latter were aware of its existence. It was on account of the litigation arising from the claim of damages by the Ætna mine that admission to the Glass-Pendery mine was for a long time utterly refused, only a single hasty visit being finally conceded for the purposes of this work. This refusal was particularly unfortunate, as the information to be obtained here has a most direct bearing on the question of the existence of ore further west under the city of Leadville. The existence of the Pendery fault, on which an incline is sunk 100 feet below the main level, was, however, ascertained beyond a doubt, with the strong probability of a slight western dip in the limestone adjoining the Pendery fault. (See Sections H and I, Atlas Sheet XXX.) Had the incline on the Pendery fault been continued, there seems to be little doubt that the limestone would have been struck in it at no very great depth, and definite data could thus have been obtained in regard to the ore horizon under Leadville.

The ore body which passes through the Glass shaft seems to have been a fracture or fissure in the limestone, partly filled with White Porphyry from the main sheet above. A section of it where it crosses a drift northwest of the shaft is given in Fig. 2, Plate XXII, in which it is seen that replacement action has followed the walls of the fissure on either side of the porphyry, the rich ore being confined, however, to the hanging wall.

In the Ætna ground a small body of Gray Porphyry is found in the drift just west of the main shaft. To the southeast of this shaft the limestone is singularly bleached and disintegrated, but no ore is found. The main large ore chambers occur north of this shaft, not far from the Pendery line. They extend up in places nearly, if not quite, to the contact plane, and are wedge-shaped or tapering toward the bottom. The ore in these claims is said to have contained little or no lead.

The Meyer shaft was sunk 50 feet perpendicularly to the fault, then followed the fault plane to the contact, from which point a drift was driven to meet these large chambers. The fault plane, like that of the Iron fault, was found to contain a certain amount of pay ore, mixed with attrition or selvage material, but none was found outside its walls in the limestone. The fissure is quite regular in its inclination, which shallows somewhat in depth, and is from one foot to three feet in width The evidence here, as in the Iron fault, shows that the original ore deposition was prior to the faulting, and that whatever ore is found on its plane was brought there mechanically or is the result of secondary deposition.

North of the Pendery and Ætna workings it was difficult to obtain accurate information in regard to the underground structure of the lower slopes of the hill, west of Carbonate fault. Many prospecting shafts have been sunk, a few of which penetrated the porphyry to the underlying limestone, but they had mostly been abandoned. The St Mary's was the only one which was accessible. From information obtained in this and by diligent questioning of persons who had visited the others, it appears that the limestone in this region probably falls off to the west in a series of irregular steps or benches, which may be actual faults or sharp flexures. The result is a probable dip to the westward, as indicated in a generalized form in Sections F and G, the only break or fold which could be actually located being that assumed as the northern continuation of the Pendery fault and which was actually seen in the St. Mary's workings. This fault is supposed to pass into an anticlinal fold in the northern half of the area mapped, as will be explained below.

In the lower part of the Yankee Doodle claim are several old prospecting shafts, now abandoned. That marked on the map as the lower shaft is said to have found limestone and ore at about fifteen feet, which was cut off to the westward by a sudden break. The break was followed by an incline seventy-five to one hundred feet farther, and work was then discontinued. This break is evidently the continuation of the Carbonate fault.

NORTHERN GROUP OF MINES.

In the northern half of the area shown on the Carbonate Hill map, another ore body parallel with that already described, but of much greater dimensions, has been developed east of the line of Carbonate fault, and a smaller, but very rich body, in somewhat peculiar relations, to the west of this line.

The first of these bodies extends northeastward from its outcrop in the Crescent claim through the Catalpa, Evening Star, Morning Star, and Waterloo claims, and probably beyond these into the ground of the Maid of Erin and Brookland claims. It obtains its maximum breadth of about two hundred feet and a thickness of pay ore of over fifty feet in the Evening Star ground, vein material having here replaced apparently the entire thickness of the Blue Limestone, which, as a result of this action, has shrunk to about one hundred feet. As in the Carbonate body, there is a noticeable steepening in the dip of the formation beyond the eastern limits of the ore current, but the amount of replacement by oxidized material has been so great that the minor waves in the limestone are difficult to trace.

It were too long to enter into a detailed description of the workings in each mine, as has been done in the case of the Carbonate; and, since the map and sections represent, so far as their scale permits, the results of thorough examination of every drift, only the salient points and general features will be mentioned in what follows.

Crescent mine.—The Crescent, like the Carbonate mine, is worked through a long incline, following the dip of the stratification. In this case, however, the angle of the incline varies from point to point in an attempt to keep on the contact; but, owing to the irregularities of the limestone surface, it runs, like the former, now into the limestone foot-wall and again into the porphyry above. Its average angle is at first 12° to 13°, but 50 feet beyond the No. 3 shaft it becomes 20° to 25°, continuing on this average slope to a distance of 800 feet from the mouth. For the first 80 feet it runs in Wash, composed of clayey gravel inclosing rounded bowlders of Sacramento Porphyry; above the contact are four feet of quartzite, supposed to belong to the Weber Shale horizon. This quartzite is very generally found below the porphyry in this and the adjoining mines to the north. It is ordinarily very thin and difficult to distinguish from the porous quartz which frequently constitutes the gangue or vein material. As in the Carbonate mine, the dip of the limestone is very shallow near the surface, and it is possible that it forms here as there the crest of a fold, but the explorations in this region were unfortunately too few to afford definite data on this point.

The main ore body is developed between the first and fourth levels and extends from the incline in a northeasterly direction to the Catalpa line. South of the incline, it was found only on the first level, extending beyond the No. 1 or Blacksmith shaft, where the limestone is rising rapidly to the surface. It was generally found as a thin and somewhat irregular sheet of sand carbonate, averaging perhaps a foot in thickness. Toward the Catalpa line on the fourth level it thickened to two and a half feet and carried 100 to 500 ounces of silver to the ton. Between the upper part or southwestern end of the body and the Catalpa line is an area which has proved barren of pay ore, so far as explored, though vein material is practically continuous through it, carrying always a certain amount of silver.

On the line of the incline, as is shown in the section, there is a ridge of limestone between the third and fourth levels and another just beyond the fifth. To the north, towards the Catalpa line, these two ridges have come together, and the steep dip, which in the incline is beyond the fifth level, is here between the fourth and fifth, as the converging of these two drifts shows. East of this line the contact has been found practically barren, though showing considerable replacement material, consisting of oxides of iron and manganese which all assay a few ounces in silver. Between the eighth and ninth levels is a deep trough, produced by a fold and possibly accompanied by some displacement, similar to that in the Carbonate mine. A winze was sunk here, said to have been 80 to 100 feet deep, in vein material, but it was no longer open, and the information is somewhat uncertain. The south drift on the eighth level runs along the edge of this trough on the contact, which dips 70° to the eastward. The extremity of the south drift on the sixth level, which follows the curves of the contact, rises 20 feet over the ridge of limestone which crosses the incline below No. 3 shaft, and finds a small body of ore, which deserves further prospecting.

Catalpa mine.—Although the first discovery of ore on this ground was made in the gossan, or iron outcrop at the surface, the mine was opened through a shaft sunk high up on the hill, which reached the contact at a depth of 170 feet and near the eastern limits of the main ore body. From this explorations were carried upward to the west, and a second shaft, the New Discovery, has lately been sunk to develop the ore shoot extending up toward the surface, which has not yet been thoroughly explored.

As the surveys in this mine had been carried on without any systematic determination of level in the drifts, as is often the case in Leadville mines, the outlines of ore bodies and contact may be less exact than in other mines. There seem to be two ridges in the formation, marked by sudden descent of the ore bodies to the east on the line of the section (E), but along the Evening Star line, above the general steepening of dip at the east limits of the bonanza, the inclination is more regular. The vein material in this ground extends to depths of thirty and forty feet below the contact, its total thickness not being in all cases ascertainable, as explorations are seldom extended in depth as far as the unaltered limestone. The rich ore bodies are generally found in its upper part, near the contact. The former is generally soft and clayey, sometimes, however, a hard silicious hematite, and in the vicinity of the ore bodies often a granular quartz, not unlike a quartzite in general appearance. In the Main shaft 40 feet of vein material was passed through before reaching unaltered limestone. This was barren, with the exception of a thin streak of galena, carrying 379 ounces of silver to the ton. The limestone below is of dark color, generally hard and crystalline, but sometimes soft and pulverulent, containing clay infiltrated through from above.

From the bottom of the New Discovery shaft the ore extends in a somewhat irregular body one to three feet in thickness to the northeastward, and along the Evening Star line is practically continuous eastward as far as the so-called "crib." In the direction of the Main shaft its continuity for a short distance is broken, but it comes in again in the northeast continuation of the Crescent body, increasing in thickness toward the "crib" on the Evening Star line, at the extremity of the drift running north from the Main shaft. At the "crib," so called from the structure, filled with waste, used to support the roof of the ore chamber, there is a sudden steepening of the ore body on both sides of the line between Catalpa and Evening Star. An almost solid mass of carbonate ore, 40 feet thick, was taken from this chamber. It was difficult to obtain definite data as to the bounding rocks, but it is evident that this deepening is due, not to a fold in the limestone, but to replacement action extending a little deeper, probably along some fissure or cleavage plane in the limestone. The general outline of the ore body here is shown in the longitudinal Section A, Atlas Sheet XXIX, whose line passes through this portion of the mines.

From the bottom of the Main shaft an incline is started in the limestone for the purpose of exploring the ground to the east; in it the bedding-planes of the limestone are very indistinct, but from data obtained at other points it is deduced that its dip must be nearly 45°.

In general it may be said of the ore bodies of the Catalpa mine that, while irregular and pockety, they have been much richer than the thicker bodies to the north.

Evening Star mine.—This claim, located on a narrow strip of ground little more than half the normal width of a claim, left between the Catalpa and Morning Star, included by good luck the thickest and widest portion of the bonanza, and has probably proved more profitable to its owners, as a legitimate mining enterprise, than any other in the region.

It is opened by two vertical shafts, known as the Main and Upper shafts, between which the ore body stretches in an almost continuous sheet and beyond which in either direction but little ore has been found.

Main shaft.—The Main shaft, as shown in Section D, was sunk through the White Porphyry, across a great thickness of iron vein material and through an underlying sheet of Gray Porphyry, into a second body of vein material, at the base of which was found a thin bed of quartzite. This is probably a portion of the Parting Quartzite, and the second iron body is therefore the replacement of a portion of the Blue Limestone split off from the main body by the intrusion of the Gray Porphyry. The fact that this underlying sheet has been actually cut here is extremely important, since its existence in the southern portion of the hill, between this point and the outcrops on the slopes toward California gulch, has been only inferentially proved by isolated masses supposed to be offshoots from it.

A dike-like body of Gray Porphyry is also cut in the upper workings adjoining the shaft. As shown here, it is six feet in width, runs in a northeast direction, and has a dip of 70° to the northwest. In places, especially toward the center of the mass, it is in exceptionally fresh condition, its matrix being a semi-translucent hornstone like mass, containing abundant crystals of limpid quartz with feldspar. By decomposition, which is often completed in a very short distance from the unaltered parts, the groundmass becomes perfectly opaque and white, and assumes a mottled appearance from the prominence given to small crystals of feldspar and the oxidation of the contained iron, so that it is not to be distinguished from the average Mottled Porphyry. This body can be traced but for a short distance in either direction. It is probably an offshoot from the main underlying sheet, and its position suggests a possible connection, or at least common origin, with that found in the Morning Star ground, although the shape of the latter, as shown in Section C, is more that of a sheet than of a dike. It must be borne in mind, however, that none of these later intrusive and cross-cutting sheets has the regularity of the normal dike as it is generally represented in geological text-books, and further that, as in the mines they are seldom exposed in more than a few isolated points, their graphic representation on the section is almost entirely ideal and subject to correction whenever further explorations furnish more facts in regard to them. The White Porphyry in this shaft was found to be highly decomposed throughout and so stained by iron oxides near the contact that the line of the latter could not be accurately determined. In it, about fifteen feet above the contact, was found a small body of ore, consisting of pyromorphite and cerussite, with a little sulphide, filling the interstices of small blocks of country rock. The rock contained over 80 per cent. of silica and may have been an included fragment of impure quartzite belonging to the Weber Shales. This was probably a secondary deposit.

The occurrence of a second body of vein material below the Gray Porphyry is extremely interesting, as showing that replacement of the limestone has taken place, at times, below this porphyry, as it has normally below the White Porphyry. In this case the body is exceptionally rich in manganese, being mainly the black iron of the miners. The jointing planes are covered with a coating of fine crystals of pyrolusite.

Upper shaft.—This shaft was sunk 290 feet through White Porphyry before reaching the contact. Here the porphyry was hard and exceptionally fresh, being what is locally known as "block porphyry"; moreover, it contained minute crystals of pyrite, disseminated through its mass. The occurrence of undecomposed pyrites in the porphyry is noteworthy in connection with the fact that for a considerable distance to the east of the shaft there is a close contact—that is, little or no replacement on the surface of the limestone. On the limestone surface only about eight to ten feet of vein material were found, which rapidly thinned out to the south and east.[1] Between these two shafts the rich ore body extends in a practically continuous sheet, reaching its greatest thickness of about forty feet toward the middle of the area. Its outlines are difficult to define, since the distinction between low-grade ore and high-grade vein material may vary at different times and since it could not be actually studied in every part; but where drifts were closed information had to be obtained from the miners. Its lower surface is very irregular, extending down into the vein material to depths which vary rapidly in a few feet; the upper limit, however, is more regular, being practically that of the contact, though every portion of this was not necessarily rich enough to be extracted. Toward the Morning Star mine it becomes thinner, but its lateral boundaries widen. East of the Main shaft, on the line of Section D, the contact is barren, but on the Morning Star line the ore extends to the line of steepening dip, as it does in the workings from the Morning Star Upper shaft. It is by no means certain that in this region the eastern limits of the ore body have been reached, and the outlines given on the map must be considered merely tentative. The diagram in Fig. 1, Plate XXII, is taken from the extremity of the incline running east from the Evening Star Upper shaft and shows a fragment of porphyry intruded into the body of the limestone, in this case unaltered; such an occurrence in a region of active replacement would account for the Chinese talc and clay which might be found entirely within an ore body.

No. 5 shaft.—Since the completion of field-work an exploring shaft has been sunk at the outcrop, which, after passing through a considerable thickness of vein material, is said to have found unaltered Blue Limestone. From this information the structure assumed in the section has been inferred, though, as will be shown in the discussion of the region west of the

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[1] Since the completion of field-work this shaft has been sunk 100 feet deeper, cutting alternately through solid limestone and replacement zones parallel with the bedding and containing iron vein material. The first of these zones was 7 feet thick, occurring at 15 feet below the contact; the second was 45 feet thick, containing in the middle from 5 per cent. to 20 per cent. of lead and occasional nodules of galena. The limestone on either side of this zone was decomposed and pulverulent, in the condition known to the miners as "lime-sand," but the bedding planes were often still distinct. From these developments it appears that the replacement zone on Section D should have been continued farther east. The shaft was not carried down to the Gray Porphyry, to determine whether pay ore exists at its contact, as in the Waterloo.

fault, there is no certainty that there may not be a decided shallowing in the dip of the formation as it approaches the assumed fault line.[1] No explorations have been made between this and the Main shaft except the Old Discovery shaft, which was sunk for a few feet in vein material and has long since been filled up and obliterated.

The ore of the Evening Star mine consists largely of sand carbonate and hard carbonate, but contains also a considerable amount of unaltered galena. Although less rich than that of the Catalpa, its working results give a high average, which may be estimated at 60 to 70 ounces of silver a ton. As compared with the Morning Star ore it runs lower in lead, but as a rule contains less silica and more iron and manganese, and is, therefore, more easily smelted. The hard carbonate, which is the characteristic ore of the western half of the mine, is a granular silicious material of peculiar steely or adamantine luster, either compact or porous and full of cavities. An examination of the cavities with a lens shows that they are more or less completely filled with transparent crystals of cerussite and little flakes of chloro-bromide of silver of pale-green color. The sand carbonates occur in streaks and lenticular bodies, generally not more than one or two feet in thickness. Chinese talc is occasionally found, but cannot be traced so regularly as when the ore bodies are thinner and the actual contact consequently more readily defined.

Morning Star mine.—To this mine belong both the Morning Star and Waterloo claims, the ground of the latter above the fault being as yet but little explored. In this mine, as in the Catalpa, no systematic levels were run, and, as many of the old workings were inaccessible at the time of visit, the data obtained as to details of form and occurrence of ore are less accurate than in the case of the Evening Star mine. The ore body continues its

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[1] According to Mr. Ricketts (The Ores of Leadville, Princeton, 1883) the data furnished by the sinking of this shaft through the Gray Porphyry and by a bore-hole drilled from its bottom as far as the Lower Quartzite show that the thickness given for the Gray Porphyry sheet in our ideal Section D is too great, its actual thickness being about fifty feet. He also states that he found no Parting Quartzite. As his information with regard to rocks passed through was obtained by examination of the dump, it might readily have escaped his observation. On the other hand, on the supposition that there was a nonconformity by erosion between Silurian and Carboniferous formations, it might have been eroded away at this point and be actually wanting; this would also account for a supposed less than normal thickness of the White Limestone.

normal northeast direction through this ground as far as explored, and apparently widens out to the northwest, the limits given to it in this direction on the map being merely those of explored areas, and not necessarily of the limits of ore deposition, as a great deal of ground is still unexamined. This body, which in the Evening Star ground had already commenced to shallow toward the boundary line, becomes very sensibly thinner throughout the Morning Star ground. From four to eight feet may be taken as the average thickness, though in places it deepens for a short distance to 20 or 30 feet. As a rule the ore runs much higher in lead than in the Evening Star, but is poorer in silver. It also contains more silica and less iron and manganese. Very white carbonate sands, consisting of almost pure cerussite, are found, especially along the contact. Here, as elsewhere, these seem to contain less silver than the more stained and impure carbonate ores. It may be that the latter have more silver in the form of sulphuret. As gangue the porous granular quartz is very prevalent, and often constitutes a good hard carbonate ore. Below the ore the ocherous yellow basic sulphate was frequently observed.[1]

The mine is principally worked through the Main shaft, from the bottom of which an incline follows the dip of the formation eastward, and levels are run southward to the Evening Star line and westward to connect with the old workings from the Lower shaft. From the incline levels are run at somewhat irregular distances, following the ore development, which has been mainly to the south in the upper part and in the lower to the north. A second shaft, known as the Upper shaft, has also been sunk to contact higher up on the hill, on the Waterloo ground, from which a level connects with the incline.

In this area the greatest east-and-west extent of the ore bodies has been along the Evening Star line, and its greatest thickness along the middle of the body, between the second and third levels. As far as could be ascertained, in one point only has unreplaced limestone been reached below the ore. This was at the southern extremity of the fourth level south, and it

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[1] According to Mr. L. D. Ricketts, who made a careful and detailed study of the Morning Star and Evening Star mines during the summer of 1882, this basic sulphate forms a distinct and practically continuous sheet under the ore body in the Morning Star ground.

was considered by the miners a limestone bowlder. Explorations generally stop in the iron body or barren vein material underlying the ore body. The steepening of dip to the east of the ore body is very marked at the present extremity of the incline, as well as near the bottom of the Upper shaft; its angle is here 25°. A fine body of carbonate ore is just being opened to the west of this point.

The old workings of the mine were reached from the Lower or Boarding-House shaft, and at time of examination were inaccessible. From information obtained it appears that a large mass of vein material was found here, and a layer of valuable ore at the contact, which reached the Evening Star line on the south, but was cut off at 125 feet east of the shaft by a break in the formation. This break was found, at the extremity of the west drift from the Main shaft, to consist of an actual displacement of 20 feet in the formation, bringing the basset edges of the limestone against a sheet of Gray Porphyry, which here overlies the contact To the west of this there is a sharp rise in the contact, which was explored with some difficulty through drifts rendered dangerous by the plastic condition of the decomposed Gray Porphyry. This body was found in no other workings, and what could be determined of its outlines is shown in Section C. Where it comes up through the limestone, as it undoubtedly must, is not therefore known, but a possible manner of offshoot from the main sheet below is shown in Section A.[1]

The normal continuation of this great ore body would be between the Maid of Erin and Brookland shafts. In the former contact was struck at 385 feet and 15 feet of iron were passed through. To the north of the shaft, in a drift rising between Gray and White Porphyry, was found a small body of galena. The developments are as yet too limited to furnish an accurate idea of the shape of this body of Gray Porphyry, which is therefore merely indicated in the section (B), as shown by present developments, with no suggestion as to its probable continuation. Ore is also said to have been struck in the Brookland shaft. The Big Chief and Clontarf shafts have also reached the contact and found vein material and ore, but as yet

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[1] Mr. Ricketts (op. cit., p. 41) states that a dike, eight to ten feet wide, crossing the limestone, has since been cut by one of the drifts of the mine, which he regards as the feeder of this sheet of Gray Porphyry.

developments are not carried on regularly in any of these shafts, owing to great influx of water and nothing definite can be said as to the extent or conditions of the ore body in that direction.[1]

Area west of Carbonate fault.—The original rock surface of Carbonate Hill west of the line of Carbonate fault slopes off very rapidly, as shown by the sudden deepening of the Wash in the various sections. The line where the slide of the steeper slopes gives way to actual Wash, or rounded bowlders and gravel of rearranged moraine material, marks a sort of beach-line in Glacial time, up to which the ice sheet must have extended in order to transport the bowlders, some of which (at the mouth of the Crescent incline, for instance) must have been brought from near the crest of the range. The depth of this mass of detrital material probably reaches 150 to 200 feet along the western edge of the map, and there is some evidence to show that the underlying Lake beds extend up to the base of the steeper rock-surface slope, as shown in the sections. Under such a mass of clayey gravel, which, like a sponge, permits the passage of water through it and yet keeps constantly saturated, the rock surface disintegrates and its mineral constituents are decomposed more readily than elsewhere, and the porphyries especially lose rapidly their distinctive characters. With these conditions of actual surface and rock surface the determination of geological structure is naturally difficult, and this difficulty is enhanced by the fact that, except at the northern and southern ends of the area mapped in the lower Henriett-Waterloo and Ætna-Pendery claims, respectively, the little underground exploring that has been done was simply for prospecting purposes, irregular, without system, and the workings are as a rule no longer accessible. The structure of this region, as represented on the sections, is the embodiment of information obtained at the expense of infinitely more time and labor than the examination of a large mine would have required, and yet is far from satisfactory in its character. For this reason, while the structure of the lower portion of the hill, as shown in Sections B, H, and I, may be taken as

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[1] Since the completion of field-work, at a depth of 633 feet contact has been reached in the Wolfe Tone shaft, which is a short distance east of the Brookland. Ore and vein material are said to have been about forty feet thick, the former occurring both as carbonate and as sulphuret. Below this a body of porphyry was found, which from description is apparently Gray Porphyry and may be the eastern continuation of the main sheet which has been developed in the Lower Waterloo workings.

essentially accurate, that given in the intermediate sections has a certain element of tentativeness and uncertainty. That in the Ætna-Pendery ground, represented by the two latter, has already been described.

Lower Henriett and Waterloo.—Section B shows the structure below the Carbonate fault on a line drawn between these two claims, whose lower workings are connected with each other. It is that of an anticlinal fold, whose axis corresponds very closely with the prolongation of the Carbonate fault line and whose crest has been planed off, with a sharp synclinal basin adjoining it on the west, along one side of which the lower sheet of Gray Porphyry cuts across the Blue Limestone up into the overlying White Porphyry, which has escaped erosion in the hollow of the basin. On the under side of this sheet of Gray Porphyry, at its contact with the limestone, the latter is mineralized, and a most valuable body of carbonate ore has been developed, extending into the hill at a rather steeper angle than the average dip of the formation. The existence of this lower ore sheet was supposed at first to indicate merely a faulted-down portion of the regular ore horizon, as it does in the Ætna-Pendery ground, the difference of level between the outcrops of vein material on the surface and that of the lower body, prolonged in dip into the hill, being quite what would be expected if the movement of the fault was normal, with a slight decrease in amount toward the north. It was observed, however, that the Half Way House and Henriett Lower shafts had passed through normal White Porphyry over limestone, whereas a short distance east of the latter shaft the White Porphyry gave way to Gray Porphyry, which thereafter continued to form the hanging wall of the ore body, limestone being in all cases its foot-wall. The White Porphyry contact, for reasons which the general geological description must have made apparent, is necessarily the top of the Blue Limestone; but the continuation of the Gray Porphyry contact soon came directly beneath the regular outcrop of iron vein material, which here has more than double the width that it has farther south. Therefore it was evident that the Gray Porphyry, although itself having the regular eastern dip, was in reality cutting across the Blue Limestone. No direct evidence of the fold in curving stratification lines has as yet been obtained, since where these would occur in mine workings the limestone has been entirely replaced and structure lines are obliterated. In the lower workings of the Henriett mine, moreover, near the prolongation of the fault line, Parting Quartzite was found in the floor of a drift, which proves that at this point the ore body is near the base, whereas in the shafts it was near the top, of the Blue Limestone. The line of Section B is apparently the line of greatest depression of the Gray Porphyry sheet, since to the north what is apparently a slight fault brings the limestone up, cutting off the ore body, and on the south, towards the New Waterloo shaft, the contact rises. This shaft was sunk entirely through vein material and Gray Porphyry, and apparently found no White Porphyry. The faulting movement has here become a slight down-throw to the east, comparable in direction and amount to the Morning Star fault, and which might readily be mistaken for a simple monoclinal fold. It is assumed to be the continuation of the Carbonate fault, though there is no direct proof, nor could a connection actually be traced, since the throw of the latter would become nil between here and where it is actually demonstrable. On the same line to the north, in Little Stray Horse Ridge, there is a displacement, which passes into an anticline on Fryer Hill, in the Dunkin ground.

West of the Halfway House shaft the contact between limestone and White Porphyry has been explored for ore, and is found to be cut off by a sudden deepening of the Wash, which evidently represents the shore-line of Lake Arkansas mentioned above. In the Jolly shaft the Wash is 140 feet deep and an east drift from it finds vein material and limestone.

The map shows an eroded anticline west of the Jolly shaft, which is the southern continuation of the quaquaversal, shown on the Leadville map, between the west ends of Fryer and Carbonate Hills.

Morning Star and Forsaken.— On the line of Section C the underground data are less complete, and the structure, which is even more complicated, is consequently determined with less certainty. The rocks passed through by the Waterloo Lower, Forsaken, and Portland shafts could not be determined by actual observation, and the information obtained may not be in every case geologically accurate. In the Forsaken incline the limestone dips regularly eastward, and the action of replacement, acting from the surface downwards, is very clearly shown. Figure 4, Plate XXII, represents a sketch of the north wall of the incline, near its head, in which, however, the transition between unreplaced limestone and vein material is less gradual than it is in nature. Above the contact in this mine, between vein material and White Porphyry, is a varying thickness of white quartzite, scarcely to be distinguished from porphyry. This is assumed to be a portion of Weber Shales left above the Blue Limestone, as is so frequently the case on Carbonate and Iron Hills. South of the Forsaken shaft the ore body, which is a layer of sand carbonate at the contact, extends apparently into the ground of the lower Evening Star, though these workings were not accessible during time of examination. West of the south shaft of the Forsaken, drifts run up along a barren contact into the Wash, which deepens rapidly along the lower end line of the Evening Star claim.

From the Forsaken incline south toward the Waterloo line, the formation rises apparently, though the connecting drift, which has a southeast course, descends near this line on a steepening dip eastward. From this drift, between the boundary line and the Main (or lower) shaft of the Waterloo, an incline follows for a short distance a rich body of sand ore at an angle of 25° to the east. At the head of this incline, directly over the ore, is a thin sheet of White Porphyry, which is overlaid by Gray Porphyry. This Gray Porphyry body dips steeply north and east and comes in actual contact with the ore at the end of the incline. It is also cut in the bottom of the Main shaft, where it is underlaid by vein material carrying a little galena. This shaft is said to have passed through vein material and then limestone before reaching the porphyry. South of the shaft a drift intended to connect with the New Waterloo shaft is in limestone, which has an apparent dip north. From the observations above noted, it would seem that the Gray Porphyry is here cutting across the limestone up into the White Porphyry in a southwesterly direction, as it is in a westerly direction on the line of Section B. Also that a slight anticlinal ridge runs along the Waterloo-Forsaken line, from which, however, the White Porphyry has not been entirely eroded off, as it has on the line of Section B. This structure may be seen graphically by supposing Section B to represent a north-and-south section across the Waterloo-Forsaken ore body on a line just west of the New Waterloo shaft. There would be the same bowl shaped syncline, though perhaps shallower, and a shorter anticline beyond it to the south, from which the White Porphyry had not been entirely eroded off. The conditions in the Waterloo Main shaft would be represented by those of the dotted lines, which on Section B denote the projection of the Harker shaft, and the Evening Star Lower shaft would occupy a corresponding position to the No. 3 Henriett. The ideal structure outlined here necessitates a very sudden rise in the original Blue Limestone surface to the northward, near the Waterloo-Forsaken line, as the movement of the Carbonate fault, which on the line of Section C is nearly 140 feet, would have become nothing and even be reversed before reaching the line of Section B.[1]

Niles-Augusta and Wild Cat.—South of the Forsaken the data to be obtained from workings west of the line of Carbonate fault were still more meager. The Evening Star Lower and Catalpa No. 2 shafts were both inaccessible; the former was said to have found a large body of vein material and ore, as shown in Section D. The latter was sunk 210 feet, and found the formation dipping nearly 45° east. From the dump it was evident that in its lower part it had passed through White Limestone and quartzite. It was assumed that it had passed across the fault line and reached the lower formations east of it.

From the Niles shaft three levels had been run. The upper drift ran east through White Porphyry and struck vein material at the Evening Star line. On the second level the drifts were mainly in limestone, with some vein material at the contact of overlying porphyry, necessitating a steep westward dip in the formation from the contact of the upper level. The lower level at 230 feet was entirely in limestone, whose stratification lines could not be distinguished. The limestone in this mine was of lighter color than is ordinary in the Blue Limestone.

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[1] Mr. Ricketts (loc. cit., p. 13) supposes a much simpler structure through the lower Waterloo and Forsaken mines, namely: that the formations continue westward on their normal dip until they reach the surface, there being no displacement anywhere along the assumed line of the Carbonate fault. His opinion is of weight, since he had the advantage of later and more extended underground workings and could give months of time where we could only give days. His explanation, however, takes no account of the White Porphyry west of the Carbonate fault line. Assuming that he is right and that our determinations of the existence of White Porphyry are at fault, some structural explanation similar to the above is required to the south of this line before the unmistakable conditions existing in the Ætna-Pendery grounds are reached.

In the Wild Cat the bottom of the shaft was in the same light-colored limestone. At a depth of about one hundred feet a level ran east along a waving contact, with considerable though irregular development of contact vein material and the same white quartzite above it that was found in the Forsaken mine. A north drift from this level followed a similar barren contact, the Wash at one point coming down into the roof of the drift. A cross-cut west from the end of this drift passed over a ridge in the formation and stopped just as it commenced to dip sharply westward.

Lower Crescent.—The workings of the Crescent Lower shaft were platted from a compass survey, and, while not as accurate as those from actual surveys, are sufficiently so for the purposes of this work. The shaft, which is 145 feet deep, passed through Wash 22 feet, iron-stained clay 10 feet, White Porphyry (soft above and hard block porphyry at bottom) 113 feet. The drifts followed clayey iron-stained material, with some development of specular iron, but no unaltered limestone was found. The east drift at the end had a dip of 45° east. The west drift passed over a slight ridge and stopped in a gentle westerly dip. Here the calculated down-throw of the Carbonate fault is about one hundred and forty feet.

The deductions made from the observations in these workings are, first, that there is a probable western dip in the formation on the lower slopes of the hill and, second, that an anticlinal structure is developed just west of the line of Carbonate fault, which to the north gradually merges into the axis of the fault, without, however, becoming strictly identical with it. The sections show a single anticlinal structure south of the Henriett-Waterloo line and a double anticline with a sharp included syncline on that line. It is possible that the structure south of that line will not prove exact in its details when more extended explorations shall be made; but errors of detail are in a measure unimportant, the great object being to determine whether there be a western dip in the formation along the lower slopes of the hill, which seems reasonably probable.

CHAPTER IV.

FRYER HILL GROUP.

GENERAL DESCRIPTION.

Fryer Hill, which has become so famous in the mining world by the richness of its ore deposits, forms a comparatively insignificant feature in Leadville topography, being simply the extreme shoulder of a minor spur extending to the westward from Breece and Yankee Hills, between Little Stray Horse and Big Evans gulches; its extreme elevation above these gulches is only 200 feet. At one time the Big Evans glacier probably covered its surface, and the moraine material which it left behind during its retreat, though partially removed by later erosion, still covers the surface of the hill to an average depth of about one hundred feet. For this reason the study of its geological structure has been a very laborious work, and one which could hardly have been accomplished were it not for the extensive underground developments made by the numerous mines among which its surface is divided up. As contrasted with the already described groups of Iron and Carbonate Hills, Fryer Hill exhibits an extreme of mineral replacement. The Blue Limestone is represented mainly by occasional detached patches of lime sand or disintegrated dolomite and irregular accumulations of iron vein material, more or less impregnated with rich carbonates of lead and chlorides of silver. The beds, which were horizontal previous to the action of mineralization, have been, during the dynamic movements which followed, compressed into gentle folds, and their crests have since been planed off by the great Evans glacier. These folds consist of a main anticlinal fold, whose axis has a northeasterly direction and forms a continuation of the Carbonate fault line, with a synclinal fold nearly parallel to it on the northwest. There is, moreover, a general dip of the folds and of the beds that compose them to the northeast. The structure is further complicated by minor crumplings, which are shown by the wavy outlines of the outcrops, but in which it is difficult to trace any regular law. In addition to all this the development of porphyry has been exceptionally great in this region. As already shown, the cross-cutting zone of White Porphyry passes through the southwestern edge of the hill, so that there is an underlying and an overlying body throughout the greater part of its area; and in the Amie ground the ore horizon is split up into three sheets, each bounded by White Porphyry. Gray Porphyry is found as an overlying sheet, as the continuation of the cross-cutting sheet of Carbonate Hill, in a dike-like body, and in several small and apparently detached sheets.

While on Iron and Carbonate Hills it is always possible to trace the limits of the original body of Blue Limestone, which shows but a moderate variation in thickness, on Fryer Hill it is difficult to explain the seemingly enormous contraction that this bed has suffered through the action of replacement. It is true that the intrusion of the porphyry mass has in many cases split the original body into several isolated sheets, whose originally irregular shape would account for a certain variation in the thickness of the present bodies of vein material. On the other hand, this explanation hardly seems adequate for certain extreme instances, such as that in the Little Chief mine, where within a distance of scarcely over one hundred feet the iron body varies from a thickness of six feet up to ninety feet. It would almost seem in such cases as if, in the plastic condition to which the presence of enormous quantities of surface waters had given rise, not only in the ore bodies but also in the surrounding porphyry, an alternating thickening and thinning of the body, perhaps already inaugurated by the shape of the original limestone, had been very much increased by subsequent compression within the mass. In other cases the finding of two bounding quartzites, that which represents the Weber Shales and that forming the Parting Quartzite, proves that there has been an absolute contraction due to replacement. The average thickness of the iron body on Fryer Hill will probably not exceed fifty feet, whereas, as has already been seen, the original Blue Limestone often reaches 200 feet in thickness.

Before proceeding to a detailed description of the individual mines, it may be well to mention briefly the locality and manner of occurrence of the main rock masses observed on Fryer Hill.

Gray Porphyry.—The bodies of this rock found on the hill have been indicated without any direct connection, simply from the fact that it has not been possible in the present state of development to trace each connection definitely, although it is very probable that many of the intrusive bodies may have a common origin. The principal body is that which is shown along the eastern and northern limits of the map, which is all that remains of the main sheet of Gray Porphyry, developed in such thickness in Little Stray Horse Park. It is the ordinary gray, somewhat decomposed rock, and has been proved in the Winnemuck shaft of the Little Pittsburgh, in several of the small shafts along the northern edge of the hill on the slopes of Big Evans gulch, and in all the shafts on the eastern edge of the map and immediately beyond it.

The second is a dike-like body, which extends probably from the Lee to the Chrysolite, although its continuity in a portion of this distance, between the Pittsburgh and the Amie claims, has not been definitely proved. This would seem to belong to the type of interrupted dikes, as it only reaches the surface in certain points, whereas in others the ore bodies extend continuously over it, but in depth it is doubtless continuous through its whole length. Where cut entirely through, it has an average thickness of forty-five to fifty feet. It is generally so decomposed that it is simply a soft, clayey mass, its only distinction from masses of White Porphyry in a similar condition being its mottled appearance, due to the forms of feldspar crystals and to iron stains resulting from the decomposition of hornblende and biotite. Occasionally, however, the characteristic large feldspars are distinctly visible, although the mass is so thoroughly altered that the pressure of the hand suffices to reduce it to a shapeless mass of plastic clay. In general this body seems to have a dip of about 45° to the northeast.

The third important mass of Gray Porphyry occurs within the lower White Porphyry, and has been cut in a drift connecting New Discovery No. 1 with New Discovery No. 5, and in the grounds of the Chrysolite between Vulture No. 1, Vulture No. 2, and Colorado Chief No. 2; in each instance characterized by its mottled appearance and by its more thorough decomposition than the inclosing White Porphyry. This is probably part of the same sheet which crosses the White Limestone under Little Stray Horse gulch, and is cut in a shaft south of New Discovery No. 5, in the New Gambetta shaft, and in the Eudora shaft. New Discovery No. 6 cuts a similar sheet of Gray Porphyry in the Blue Limestone, which evidently is part of the cross-cutting sheet of the Waterloo-Henriett claims. There is little doubt that all these bodies form part of the same intrusive sheet which is gradually rising in geological horizon to the westward, as shown graphically on the map.

Besides these three principal bodies, small irregular sheets are found overlying the iron in the southern portion of the Little Chief claim and in the Robert E. Lee mine, apparently conformable with the formation. In the eastern portion of the latter mine also is a dike-like sheet, five or six feet thick, cutting through the ore body and extending from the northern drift on first level to the eastern drift on second level, or in a northwesterly direction, with a dip, however, to the eastward.

White Porphyry.—The upper sheet of White Porphyry is generally very much decomposed, and within fifteen or twenty feet of the ore bodies it is reduced to a mixture of clay and quartz grains. Over the main summit of the hill there is little of it left, probably on an average not more than fifteen or twenty feet. Its decomposed state is evidently due to the abundant action of surface waters, which have free access through the superincumbent Wash, there being no solid rock above it.

The lower White Porphyry is relatively much less decomposed, although the microscope shows that decomposition has already progressed to a considerable extent within its mass. In some places, however, notably in the lower drifts of the Dunkin and Climax, where it approaches the lower limestone, it has been reduced to a clay and is so full of oxide of iron that it is difficult to distinguish it from the iron mass which has replaced the limestone. It is also characterized by a laminated appearance which makes it closely resemble decomposed shale. Its thickness has been proved in a number of shafts to vary from sixty to one hundred and sixty feet, as folfollows: Climax No. 1, 115 feet; Climax leased shaft No. 1, 60 feet; Climax No. 5, 115 feet; Dunkin No. 1, 110 feet; Montana shaft, 99 feet; Amie No. 2, 163 feet; Little Chief No. 1, 135 feet. It will be noticed on the map that this lower sheet of White Porphyry gradually passes up across the body of Blue Limestone to the southward towards Carbonate Hill, finally merging with the upper sheet.

Weber quartzite.—The overlying quartzite is coarse-grained and sometimes micaceous, as is common in the Weber formation. It occurs in detached patches at various points between the ore body and the overlying porphyry. Its most continuous body is in the Chrysolite ground, where it extends from the Roberts shaft towards Chrysolite No. 4, and varies in thickness from one foot to six feet. In the Roberts shaft, where its maximum thickness occurs, it is separated from the iron body by ten or fifteen feet of porphyry. In a drift from the Roberts shaft to Chrysolite No. 4 it is found sometimes resting directly on this iron, again separated by several feet of porphyry, and at other times split up into several bodies. Elsewhere on the hill it is generally found in more or less rounded fragments, included in the porphyry, directly above the iron.

Blue Limestone.—Small irregular bodies of pale-blue sand, generally near the surface of the ore, are frequently found in the vein material of Fryer Hill, especially in the Chrysolite ground. The occurrences of actual bodies of limestone in place are, however, extremely rare. Those observed are—

1. At the extreme western edge of Fryer Hill, as shown in the Colorado Chief and some of the adjoining prospect shafts, where it is struck directly beneath the Wash and apparently rests immediately upon the Parting Quartzite, with no intermediate body of White Porphyry. The outlines of this body could not, of necessity, be very definitely obtained, but it is probably of considerable extent and evidently represents the wedge-shaped portion separated by the cross-cutting sheet of White Porphyry. It is here a dark-blue, granular limestone, frequently somewhat impregnated with iron.

2. A large body of lime-sand occurs in the western portion of the Chrysolite, south of Vulture No. 2, adjoining the lower iron body.

3. A fragment of Blue Limestone is found below the main iron body, included in the porphyry, in the second western drift south of Vulture No. 3 shaft (see Section F, Atlas Sheet XXXIII).

4. A nearly continuous bed of lime-sand, in which are occasional portions of compact Blue Limestone, is found overlying the iron body, extending from near Vulture No. 3 over adjoining portions of the Vulture, Carboniferous, Chrysolite, New Discovery, and Little Chief claims.

5. In the Little Chief mine, about the middle of the claim and south of the Gray Porphyry dike, a body of limestone comes in suddenly, occupying the greater portion of the ore horizon for a considerable distance, an up-raise having been made through it; while below it has been proved to extend to the Parting Quartzite by a drift running along the contact of the two in an east-and-west direction. This limestone is partly disintegrated into sand and partly in a compact state. A little farther south the iron body is found resting directly on the Parting Quartzite, affording a direct proof that it is a replacement of the limestone. (See Section J, Atlas Sheet XXXIV.)

6. The most considerable body found is that cut at the northern end of the third level of the Dunkin mine. The drift runs through this body for a distance of about one hundred feet, the stratification lines showing at first a dip of 45° to the northward; the angle becomes shallower farther on, which may be due to a change in the strike. This body shows the characteristic ribbed structure of the Blue Limestone and contains imperfect casts of fossils. It has, moreover, every external appearance of a solid hard rock, but upon being broken down crumbles at once to fine sand.

The analyses of these lime-sands show no essential change in composition from the unaltered rock, as regards their contents in carbonate of lime and magnesia. The following are the proportions obtained:

	Carbonate of lime.	Carbonate of magnesia.
Dunkin lime-sand	55.14	44.29
Chrysolite lime-sand	54.09	43.79
Silver Wave limestone (type rock)	54.98	44.39

The disintegration is probably due to the dissolving out of the cementing material, which held the grains together, by percolating waters, and from the above analyses it would seem that in these dolomites, as in quartzites, the cementing material was essentially of the same composition as the rock itself.

Gangue.—The material which replaces the limestone on Fryer Hill does not differ essentially from that already described on Iron and Carbonate Hills. It is mainly an impure mass of oxides of iron and manganese, with a greater or less admixture of silica and clayey materials. It differs somewhat in different portions of the hill, being relatively richer in alumina and iron in the main mass of the hill, and very silicious in the eastern portion, or in the Lee group of mines. The black or manganiferous iron, which forms a large portion of it, though very irregular in its distribution, is, as elsewhere, generally barren. In the main mass of the hill silicious replacements consist of black chert, often forming large, almost solid bodies and at other times being thoroughly shattered into angular fragments, the seams more or less filled with clay. Considerable amounts of sulphate of baryta in crystalline form are scattered irregularly through the ore deposits, and are generally considered a good indication of ore, inasmuch as they usually accompany rich masses of chloride. The oxide of iron is generally hydrated, though sometimes mixed with a certain amount of anhydrous oxide. Frequently it forms a comparatively pure iron ore and is valuable as a flux, notably that occurring in the Amie mine. From this, with varying proportions of iron and silica, it passes through jaspery iron with conchoidal fracture into almost pure silica. It is frequently cavernous, the cavities being lined with crystals of quartz, cerussite, and sometimes of pyrolusite. Black iron contains from 10 per cent. of manganese upwards, though never approaching a pure manganese mineral in any large quantity, except in the Dunkin and in the adjoining workings of the Climax No. 3. No pyrites, so far as known, have ever been found in the mines and carbonate of iron is extremely rare.

Ore deposits.—Ore occurs either in the form of galena or its decomposition products, carbonate and a little sulphate of lead, with a small amount of chloro-phosphate or pyromorphite; and silver, either inclosed in the galena or impregnating the vein materials in the form of chloride, chloro-bromide, or iodide, or a mixture of the three. Galena occurs irregularly, generally in the center of a large mass of vein material, with its surface more or less oxidized and changed into carbonate. Besides this, considerable masses of sand carbonate with hard carbonates are found, which are always more or less stained with iron. Actual pseudomorphs of carbonate of lead after cubical crystals of galena have been found, as have also small crystals of pyromorphite and molybdenite. The galena, as a rule, contains a proportionately larger amount of silver than the carbonate of lead, as might naturally be expected, since, in the oxidization of galena by percolating waters, silver is removed in the form of chloride and has frequently been redeposited at some distance from its original position. In this way the Lee group of mines evidently owe their ore entirely to the later mineralizing action, since they are practically free from lead and consist of chert and highly silicious red and yellow ochers, impregnated with chloride and chloro-bromide of silver, without any lead The darker-colored sand carbonates are, as a rule, the richer; these found in the Amie mine, for instance have a dark-blue or greenish tinge and carry 300 ounces of silver to the ton, whereas the light-colored carbonates of the Morning Star mine only contain from forty to fifty ounces to the ton.

The extreme irregularity of the occurrence of the ore renders any generalization extremely difficult. It may be stated, however, that here, as in all the other hills, the rich ore is generally, though not invariably, found along the upper portion of the ore body. The main bonanzas or ore bodies, as will be seen by reference to the map, have a nearly east-and-west direction, parallel to the dike already noticed. It will also be noticed that the richest bodies have been found in comparative proximity to this dike and to the northeast of it, the main rich body to the southwest of this dike being that in the New Discovery, Little Chief, and Little Pittsburgh mines, oppoite what seems to be a partial break in the continuity of the dike. The influence of this dike on the deposition of ore has evidently been to cause an interruption or stagnation in the ore currents, by which their contents were precipitated more richly in a sort of eddy immediately adjoining it. On the northern portion of the hill no large ore bodies have been found as yet, although the existence of a large body of iron has been proved in which there are found small irregular pockets of ore. Exploration in this direction has been comparatively neglected on account of the great inrush of water wherever shafts have been sunk to the ore horizon. Erosion must have removed an enormous quantity of ore from the crest of the fold, and doubtless from the surface of many of the existing ore bodies. As a rule the lower iron bodies are comparatively barren. On the other hand, that very rich portion of the Lee body which extends into the Matchless and Hibernia grounds is immediately above the Parting Quartzite, or represents the very bottom of the Blue Limestone. A small body of ore was obtained from the lower bed of the Amie mine, which yielded seventy ounces to the ton, and from the workings of the Vulture No. 2 some pay ore was obtained in the same horizon.

The dip of the ore horizon is generally quite low, in the New Discovery not more than 5° and in the Little Pittsburgh and Little Chief the southern portion is quite horizontal. In the Chrysolite it dips from 10° to 11°, whereas in the Amie the dip is 20° to the northwest. On the east side of the anticlinal fold the dip is uniformly steeper, so that from the Lee to the Denver City the outcrop is relatively narrower.

Parting Quartzite.—The Parting Quartzite, where observed, is not over ten to fifteen feet in thickness and, as contrasted with the upper quartzite, very much decomposed, being generally disintegrated to a fine white sand and in every case very much iron-stained; it also contains considerable mechanical admixture of clay from the porphyry, so that it is not always possible to distinguish it with certainty from a highly decomposed White Porphyry from which the earthy bases have been largely removed.

White Limestone.—The White Limestone has been cut by several shafts on Fryer Hill. In its most characteristic form it is found in the Amie No. 2, where it was struck at a depth of 273 feet and has its peculiar light-drab color, compact texture, and the characteristic segregations of white chalcedony or chert. It is also found in the lower levels of the Dunkin and Climax mines, at their southern extremities, and in the Eudora and Pittsburgh shafts, toward Little Stray Horse gulch, where it comes up to the Wash; likewise in Chrysolite No. 6.

Lower Quartzite.—This quartzite has been cut by various prospect shafts along the western borders of the hill; among others, by the Little Eva No. 5, and by the Lida shaft, near Cumming & Finn's smelter, where a small intrusive body of White Porphyry was found in it. In other portions it has not yet been actually cut, but its existence is readily deduced from the position of the formations above it.

Fryer Hill map.—Before proceeding to a description of the underground workings of the various Fryer Hill mines, it may be well to explain what is intended to be represented on the map of that region (Atlas Sheet XXXI) in somewhat greater detail than is given in the legend. It has been attempted here to show on a single sheet the actual surface of the ground, the rock-surface beneath the Wash (distinguishing the different formations which make up that rock-surface), and the underground workings of the various mines, with the outlines of the bonanzas or ore bodies, as far as could be determined at the time of examination. The attempt to delineate so much on a single sheet has resulted in a complication, which it will require the reader's closest attention to unravel. He must first bear in mind that the black contours indicate the actual surface of the ground, and the figures attached to them, their relative elevation above the 10,000-foot curve. Second, that the geological colors indicate the outcrops of the various rocks and formations as they would appear if the superincumbent Wash material (which, as shown by the sections, has a thickness of from thirty to one hundred feet over the whole surface) were entirely removed. Third, that the drifts of the various mines and the outlines of the ore bodies, as determined by the explorations of those drifts and shown in black dots, are projected on a plane surface, or, in other words, are represented as if the rock material above them were entirely transparent and without regard to its thickness. These drifts, while in the main following an approximately horizontal plane, are in many mines on two or even three different levels. Owing to their approximate horizontality it was impossible, without too great complication, to indicate these different levels by any series of colors or conventional signs, but figures have been placed within the drifts to show their elevation in feet above the 10,000-foot curve. The map thus furnishes the data from which a section may be constructed along any given line, and from it the various sections represented on Atlas Sheets XXXII, XXXIII, and XXXIV have been so constructed. In unexplored portions these sections are more or less ideal, and actual exploration may prove them to be not absolutely correct. The intersections of the drifts with these sections show to what degree the plane of the section has been actually explored, but the outlines of the formations as indicated there are determined also from analogy and by deduction from observations made in the vicinity, but not actually on the section plane.

On the surface map the outlines of the various claims are indicated by broken lines. These are sometimes difficult to trace, owing to their coincidence with lines of drifts or with the blue lines representing the lines of the different sections. They are given as accurately as could be determined by the engineers who had been employed in surveying them; but, as invariably occurs in rich mining districts, there are many cases of contested boundaries between adjoining claims which have either been settled by compromise or are still in litigation, so that the lines here given cannot be assumed as officially and finally correct. The laying down upon an accurate topographical map of a mining district like this of finally correct side lines to the many claims that are there located, if not an absolute impossibility, would require an amount of time entirely incommensurate with the value of the result to be obtained, as can be readily understood by those familiar with the working of the land system of the United States as applied to mineral claims. In describing the various mine workings and ore bodies of this group, they will be taken up in geographical order, proceeding from west to east, without regard to priority of discovery.

MINE WORKINGS.

Chrysolite mine.—The property of the Chrysolite Mining Company consists of the following claims: Carboniferous, Chrysolite, Vulture, Little Eva, Colorado Chief, Pandora, Fair View, Kit Carson, and All Right. The greater part of the ore extracted from this property has been obtained from the first three claims. The area occupied by the others, as shown by the map, is mostly west and south of the outcrops of the main ore body, indicated in dark blue crossed in black; in other words, over this area the ore has been mainly removed by erosion. The Discovery shaft of the Chrysolite has, so far as known, discovered no ore. The first considerable ore body was opened by Vulture No. 1 shaft, near the northern end of this claim.

Soon after the discovery of this ore body, it was found by some clever persons that, as not unfrequently happens in the location of claims, a small triangular piece of ground in the immediate vicinity of this shaft, about thirty-six feet by sixty on the sides of the triangle, was left unclaimed. The shaft now called the Eaton shaft was at once sunk by these men and the continuation of the Vulture body there discovered. In self-defense the Chrysolite Company were obliged to buy out this little claim, generally known as the Triangle, at what at the time seemed a very high price, but which was repaid to them more than threefold by the ore which they extracted from the ground. A still smaller unoccupied piece of ground between the end lines of the Vulture and the Chrysolite, appropriately called the Sliver, was similarly taken up, and was bought out by the Chrysolite Company, but has thus far proved an unprofitable purchase. This can be distinguished upon the map by the two shafts, Sliver No. 1 and Sliver No. 2, which have been sunk upon it. The ore body at the Eaton and Vulture No. 1 shafts was near the outcrop of the upper iron body, and therefore its limits to the westward were soon reached. It was traced eastward, descending irregularly, but at a low angle, as far as the Carboniferous-Little Chief line. This ore body, though narrow, was extremely rich and yielded the greater part of the immense returns which were obtained from the mine in the earlier days of its working. Section B shows its outlines along an east-and-west line. From it were obtained large masses of chloride of silver, associated with cerussite, a single transparent mass of chloride which weighed several hundred pounds having been extracted.

A second ore body, parallel to this, was found about one hundred feet to the southward, which was traced in a southeasterly direction to the New Discovery-Vulture line, where it widened out and then disappeared. In the bottom of an east-and-west drift, a little south of this body, several masses of limestone were found below the main ore body, which in early days, before the character of the formation of the ore was understood, much puzzled those in charge of the mine. The extent of this unreplaced limestone was never determined, but it is now evident that it is simply a portion of the Blue Limestone, which existed wherever now the body of vein material is found, and which, for some reason or other, had not been replaced by vein material. Along the Vulture-New Discovery line considerable unreplaced lime-sand was also found on top of the iron body, extending into the New Discovery ground.

The main ore body, which was opened by the Vulture No. 1 and Eaton shafts, was traced a little south of east to the extreme southeast corner of the Carboniferous claim, with an average width of about fifty feet and a thickness of twelve to twenty feet. In the eastern portion it is opened by Chrysolite No. 3 and Carboniferous No. 1[1] shafts. It ends quite abruptly, both on the north and south, though barren vein material is found on either side of it. Cross-cutting drifts soon pass through this vein material into overlying White Porphyry, showing that the ore body was on the crest of a minor ridge or corrugation in the vein material, on either side of which is a shallow basin. That to the south has proved barren for a considerable distance into the New Discovery ground. Its form is shown in Section F, where it is seen that the upper portion consists largely of unreplaced limestone. An ore shoot was also followed, descending in a northeasterly direction from the Triangle workings, which later developed a large body of ore in the neighborhood of Chrysolite No. 4 shaft. In these older workings the rich ore consisted mainly of carbonate of lead and chloride of silver, with a comparatively small amount of galena. In the vein material a blue-black chert is prominently found, occurring in bodies up to ten feet in thickness. From this impure silica it passes into silicious iron and then into a clayey limonite, more or less impregnated with oxide of manganese, the extreme form of which, known to miners as "black iron," is a sort of impure wad. These were the early workings of the mine, made in the upper ore horizon.

Explorations were also conducted westwardly by a drift running a short distance south from Vulture No. 1 and then west to Vulture No. 2. The workings of Vulture No. 2 shaft disclosed a considerable body of vein material, about twenty-five feet in thickness, immediately underlying the Wash, containing a little ore, and passing to the south into lime-sand. To the west of this is a coarse decomposed quartzite, which is assumed to represent Parting Quartzite at the base of Blue Limestone. The connect-

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[1] Wrongly marked No. 7 on the map.

ing drift between Vulture No. 1 and Vulture No. 2 is mainly in very much decomposed Gray Porphyry, distinguishable by its mottled appearance and occasional large crystals of feldspar. White Porphyry separates this from the ore body of Vulture No. 2, and the contact between the two porphyries, which dips to the eastward, is slightly iron stained. The ore body of Vulture No. 2 evidently represents a small portion of the Blue Limestone, split off from the main body by the lower White Porphyry and by this intrusive sheet of Gray Porphyry, which is assumed to be the same found in New Discovery No. 5, which, extending across Little Stray Horse gulch, connects with the lower body on Carbonate Hill. The drift, in a southwesterly direction, from Vulture No. 1 to Colorado Chief No. 2, also crosses this body of Gray Porphyry. Midway between the two shafts the top of the drift cuts a body of fine conglomerate, resting immediately upon the porphyry and apparently belonging to the Lake bed formation. In the southeast corner of this drift a winze has been sunk to a depth of 86 feet, passing through the Gray Porphyry sheet into an underlying iron body, from which ore assaying 72 ounces to the ton was taken. The winze was abandoned on account of the difficulty of handling water; but its exploration proved sufficiently that the Gray Porphyry is a sheet dipping northeastward with the formation and that a second iron body occurs below it. A large outcrop of Blue Limestone, partially replaced on its upper surface and represented on the map in the Kit Carson, Fairview, Pandora, Colorado Chief, and New Discovery claims, belongs to this lower body, which is separated from the main upper ore body by the cross-cutting White and Gray Porphyries. The outlines given on the map are determined mainly by data derived from the dumps of a few abandoned shafts, and therefore are probably not absolutely correct. Its widest part is in the line of Section F, between Colorado Chief No. 1 and Pandora No. 3 shafts, the latter finding some vein material at its base.

The portions of the mine thus far described, and which are shown in the southern and western ends of Sections L, F, and B, respectively, were opened in the early days of ore development in the district, when it was supposed that the ore bodies would probably be found to take a downward direction towards the unknown sources below, from which they are generally supposed to come. Mining was, therefore, conducted without any definite system. Drifts were run here and there, up and down, wherever ore could be found, so that it was extremely difficult in traversing them to form a clear idea of the actual extent and form of the ore bodies or to know whether or not important ground still remained unprospected. When Mr. W. S. Keyes took charge of the mine a new and more rational system of development was adopted. A large three-compartment shaft, the Roberts shaft, was sunk in what it was supposed was the deepest part of the ore horizon, and from this shaft a system of horizontal drifts was run off at two or three different levels, with a regular system of rectangular cross-cuts at given distances. In this way it was possible to map out the shape of the ore bodies, and it soon became evident that the vein material occurred as an interstratified mass between two sheets of porphyry, somewhat irregular and corrugated, but basining up to the surface on the southwest and northwest. With the increased facilities for handling ore given by a large shaft and by level tramways leading from every part of the mine to it, the work of exploration could be pushed much more rapidly and the extraction of ore proportionately increased. The old workings of the mine were further explored and considerable ore was discovered where the bonanzas had been supposed to be exhausted. Entirely new bodies of ore were also found to the west and northwest of the shaft, continuing irregularly up to the outcrop as shown on Sections A, K, and L. None of these bodies was of so great continuous extent or so rich in silver as the main ore body extending eastward from Vulture No. 1. The most extensive and the richest was that developed near Chrysolite No. 4.

In the drifts running southward from the Roberts shaft to connect with the Chrysolite No. 1 workings a body of Gray Porphyry thirty to forty feet in thickness was cut, which has since been traced eastward as far as the Robert E. Lee mine. In the Chrysolite ground this body of porphyry has a dip of 45° to the northeast, but farther north it stands apparently nearly vertical, and is, therefore, assumed to be an interrupted dike. It extends much farther east and west than is shown on the map, being exposed by drifts from the lower levels where it is wanting in these directly above them, showing that it tapers upwards. In the northwest portions of the workings a few feet of Weber quartzite are found above the ore body, sometimes in direct contact with it and again separated by a slight thickness of White Porphyry; and in the extreme northwest workings the Parting Quartzite is found beneath the ore body, between it and the underlying White Porphyry, showing that here the entire thickness of the original Blue Limestone horizon is represented, compressed to a thickness of sixty to eighty feet. At the bottom of the Roberts shaft 10 feet of quicksand were passed through, which probably represents the disintegrated Parting Quartzite. The main west drift from the Roberts shaft on the 284-foot level, which runs a little south of west, passes for the first hundred feet through iron vein material containing some pay ore, then for 150 feet through block (White) Porphyry, then into a second body of iron vein material, at the extremity of which is some lime sand, succeeded by Parting Quartzite, all dipping gently to the eastward.[1] This is evidently the continuation of the ore horizon developed in Vulture No. 2, but it is noticeable that the Gray Porphyry found between that shaft and Vulture No. 1 is here wanting, showing that its northern limit has been reached. The lower iron body does not extend much north of this drift either, since, as shown in Section A, the drift westward from Chrysolite No. 4 finds the Parting Quartzite directly under the main or upper iron body.

In this northwestern quarter of the Chrysolite ground the body of vein material has averaged from sixty to eighty feet in thickness from its outcrops eastward and southward. In this vein material the bodies of rich ore are necessarily difficult to define, as they are simply concentrations of lead and silver minerals. The whole mass contains more or less of these metals, of which a certain arbitrary percentage is required to constitute pay ore. The ore consists, as in other parts of the mine, mainly of carbonate of lead and chloride of silver. The discarded iron vein material, which is extracted from the mine and accumulates on the dumps, constitutes a low-grade ore which it will doubtless some day be found profitable to work.

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[1] Since the completion of field-work this drift has been pushed farther westward than indicated on the map, and has passed through the White Limestone into the Lower Quartzite, showing that the outlines given on the map, though from somewhat meager data, are in the main correct and that the formations basin up to the westward. This drift has also discovered a westwardly dip in the Lower Quartzite, proving further the existence of the anticline which had been assumed to exist here.

To the northeast of the Roberts shaft pay ore is cut off by a body of black iron, into which it passes so abruptly that the latter often forms a wall 20 feet in height. Above the black iron is a body of blue lime-sand about one hundred feet in extent. Beyond the ore stopes in the vicinity of the shaft, exploring drifts on the lower (284-foot) level connect to the northward with Carboniferous No. 5 shaft, and from there to the westward with Chrysolite No. 5 shaft by an up-raise to the 316-foot level, all in barren vein material. From Carboniferous No. 5, the bottom of which is in disintegrated Parting Quartzite similar to that cut in the Roberts shaft, a drift runs due north through White (block) Porphyry and at 200 feet from the shaft cuts White Limestone, which is slightly iron-stained at the upper surface. Still farther north, beyond the limits of the Chrysolite claims, the Silver Wing shaft was sunk through White Porphyry into a body of iron vein material, which is evidently a replacement of the upper portion of the White Limestone. Explorations were conducted here under great difficulties, owing to the immense in-rush of water, and, so far as they went, did not disclose enough pay ore to justify the owners in pursuing them further.

The evidence of these northern workings is very conclusive as to the basining-up of the formation to the northwest, and this evidence is further confirmed by the several shafts to the east of the Silver Wing, the Buckeye, Hazzard, Hercules, Comique, and O. K., all of which have found a considerable body of iron vein material, either at the rock surface or under a thin covering of White Porphyry, which represents the outcrop in this direction of the Blue Limestone horizon. As in the Silver Wing, the great in-rush of water has proved a bar to extended explorations from these shafts.

The Gray Porphyry dike separates the two main ore shoots of the Chrysolite ground. Little can be determined about the form of this body in depth, as explorations have not proved it below the Blue Limestone horizon. It may be simply a transverse sheet, cutting diagonally across the formations and assuming a vertical position as it approaches the present rock surface. Still, its form, so far as traced, is sufficiently characteristic of the dike type to justify the assumption that it is rather a true dike than a transverse sheet, though the distinction, so far as the deposition of ore is concerned, is comparatively unimportant. It is distinctly later than the White

Porphyry, as are the transverse sheets of Gray Porphyry already noticed, and like them its influence has been favorable to the deposition of rich ore. It should not, however, be regarded as a dike cutting through the ore bodies, since it was evidently intruded before ore deposition commenced. Its exact relations to the original ore bodies are now difficult to define, for these were probably deposited in the form of sulphurets in a much larger proportion of unreplaced limestone than now exists, and the secondary action of oxidation, which has been going on ever since, has evidently increased the volume of vein material and reduced that of the unreplaced limestone. The probability is that, as in cutting across the formation this body probably interrupted some of the natural water channels along the contact planes of different rock formations, it caused a partial stagnation of the ore currents in its vicinity and thus favored precipitation and replacement action there.

The ore bodies are continuous around its western end from the Triangle workings to Chrysolite No. 4, and it is probable that its western limit is not far from that indicated by its outcrop on the map, as otherwise it would have been cut by some of the drifts in this portion of the mine, which, owing to the basining-up of the formation here, reach lower horizons than elsewhere. The ore bodies are also practically continuous across the line of the dike along the Carboniferous-Little Chief line, but here the dike is proved to exist under these ore bodies by drifts at lower levels, and the inference, therefore, is that, as the dike did not extend up to the upper surface of the Blue Limestone, ore deposition went on uninterruptedly across this break in its upper line. It was just to the north of the dike, in the Little Chief ground, that the thickest body of pay ore was found. The ore body in the extreme southeastern portion of the Carboniferous claim was also very thick; but, being among the earlier discoveries, the workings had caved at the time of visit and could not be examined; 12 feet of lime-sand and 24 feet of ore are said to have been cut by this shaft.

New Discovery.—The New Discovery claim adjoins the Carboniferous and Chrysolite on the south and the Vulture on the east, and geographically forms part of the ground just described, though it belongs to the Little Pittsburgh Mining Company, the claim of that name lying entirely to the east of the Little Chief, which separates it from the above claims. This and the Little Pittsburgh were the first mines worked in this region, and at the time of examination the larger ore bodies had been stoped out and the stopes were filled up, so that but imperfect data could be obtained with regard to them. The ore body was nearly continuous on a north-and-south line from the Carboniferous ground to New Discovery No. 2 shaft. It consisted mainly of sand carbonate, with chloride of silver, and had an unusual amount of barite in the gangue. This ore occurred mainly in the upper part of the ore horizon, resting in general on chert, with barren iron and clay below. This same upper ore body also covered a considerable area northwest of No. 2 shaft, and was expected to prove continuous over the greater part of the claim in that direction. As it approached No. 4 shaft, however, it gradually gave way to a mass of chert, which sometimes occupied the whole horizon, and which along the Vulture line was overlaid by a considerable body of lime-sand and unreplaced dolomite. On this northwest line a few small, scattered bodies of rich ore were found, but just to the northeast of it is the barren zone, already noticed in the Chrysolite ground, which seems to occupy a trough in the formation, the ore horizon, represented by comparatively barren vein material, descending towards its axis from either side These descents are sometimes so abrupt as to suggest a slight movement of displacement. To the southwest of this line the ore bodies, which are very irregularly distributed, extend up to the Wash. They follow two radiating lines from the main ore body, the one in the direction of No. 1 shaft, the other intermediate between that and the drifts running to No. 4. In either case the ore bodies descend to the southwest, which would at first seem a contradiction to the statement that the formation has a general dip northeast. The fact is, however, that the rock-surface, like the surface of the ground, descends here towards Little Stray Horse Creek, and these ore bodies, which are all that erosion has left, belong to the lower part of the ore horizon. It therefore suggests itself that, if this lower portion had been thoroughly prospected in other portions of the mine, other ore bodies might have been found. Owing to the imperfection or want of surveys, it is impossible to say whether this has been done or not.

New Discovery No. 1 shaft is that in which the original discovery of ore was made on the claim by George Fryer, at a depth of 60 feet. The iron body was only 20 feet thick, and this shaft then passed into the underlying White Porphyry. The small thickness of the iron body is here due to the fact that the upper portion of the ore horizon has been eroded off. In later times considerable exploration has been done from the shaft to determine whether the ground to the south is ore-bearing or not. Diamond-drill borings were made from an east drift at a depth of 165 feet below the top of the shaft, both eastward in a horizontal direction and vertically downwards. Neither found any ore bodies. The vertical drill penetrated to a depth of over one hundred and seventy-five feet, making a distance of 340 feet in all below the surface. It passed through the porphyry, finding a thin streak of iron vein material in its midst, into the Silurian formation, and apparently through that into the Lower Quartzite or Cambrian. Frequent assays of the cores were made by Mr. Rudolph Keck, and a slight trace of silver, amounting in some cases to ten ounces to the ton, was found in most of the material passed through, but no evidence of any ore bodies.

To the southward a drift was run, descending from 10,347 to 10,316 feet elevation, which passed through White and Gray Porphyries, finding a small streak of iron oxide at the contact of the two. In the Gray Porphyry body the drift turns abruptly east to connect with No. 5 shaft, which it does at 100 feet below the surface. This shaft was sunk to a depth of 185 feet, and, judging from the material on the dump, must have passed through the Gray and White Porphyry bodies and the Parting Quartzite into the White Limestone.

An exploring shaft (No. 6) was also sunk on the ridge south of Little Stray Horse gulch, at the southern extremity of the claim. It was driven somewhat intermittently, and could not therefore be closely followed. The rocks passed through were approximately as follows: Wash, 120 feet; Gray Porphyry, 40 feet; Blue Limestone, 60 feet; Parting Quartzite, 20 feet; White Limestone, 20 feet. This is on the south side of the shallow anticline assumed to exist under Little Stray Horse gulch. The structure, as well as could be deduced from the meager data obtainable in this part of the region, is shown on Sections C and K. The body of Gray Porphyry, which is here in the Blue Limestone, is assumed to be the same sheet which occurs in the lower White Porphyry at No. 5 shaft, and which is gradually cutting up to a higher horizon as one goes south, reaching the upper White Porphyry in the Lower Henriett ground.

From the relative elevation of the Blue Limestone in this shaft and in the adjoining shafts to the southeast, the Pearson (S–14) shaft of the Gambetta claim, the Joe Bates (S–26) shaft of the Stray Horse claim, and the Vanderbilt (S–25) shaft, there is evidently a break or a sharp fold in the formation to the east of this shaft. On the section both are assumed to exist, and the fault to be the northern continuation of the Carbonate fault. It must be stated, however, that it has not yet been cut on this ridge, and in so far its existence is a matter of pure hypothesis. There is unquestionably an anticlinal fold here, however, which can be traced northeastward into the Dunkin ground.

Little Chief.—This claim is analogous to that of the Evening Star, on Carbonate Hill, in that, being a narrow piece of ground left between two adjoining claims, it included within its area an unusually large proportion of ore-bearing ground. Its width is only 250 feet, instead of the normal 300, and the title to part of this was contested by the overlapping of the south end of the Little Pittsburgh claim. The outlines of the full claim are given on the map, as well as the broken line which was adopted as a compromise boundary between the contesting claims. As in the ground previously described, there are two main ore bodies, a southern and a northern, separated by the porphyry dike and an area of barren ground. The porphyry dike does not, however, reach the rock surface, as far as known, and in the western portion of the claim the ore body is continuous over it, and forms a connection between the northern and southern bodies along the Carboniferous and New Discovery lines. Here also, the southern body, at its outcrop immediately beneath the Wash, was the first opened. The original workings were reached through the small shafts Nos. 1, 2, 5,[1] and 7, and were driven irregularly, following the ore shoots. No. 1 found the ore directly beneath the Wash, at a depth of fifty to sixty feet below the surface, in a thickness of ten to twelve feet. The shaft was afterwards

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[1] No. 5, which is the southernmost shaft on the claim, is wrongly numbered on the map No. 3.

sunk through the underlying White Porphyry and reached the Silurian formation at a depth of 198 feet, finding 2 feet of iron[1] at the top and penetrating it 16 feet. No. 5 shaft, due south of this, after passing through 63 feet of Wash, struck the underlying White Porphyry, and reached the Silurian formation at a depth of 159 feet. The ore bodies reached from No. 1 shaft are at an elevation of about 10,400 feet, and lie directly beneath the Wash. Those opened by No. 7 shaft are about fifty feet lower, and are covered by White Porphyry and by a thin sheet of Gray Porphyry which is seen in a drift leading from No. 2 shaft. The ore body in this portion of the workings was nearly horizontal and from one to two and a half sets of timber in thickness (8 to 20 feet). North of No. 2 shaft, however, the formation dips rapidly to the northward, and on the line of Section J a considerable body of unreplaced Blue Limestone, occupying almost the whole thickness of the ore horizon and underlaid by Parting Quartzite, is developed by an up-raise from the 320-foot level; a little south of this up-raise iron is found to rest directly on the Parting Quartzite, thus affording a direct proof that it replaces the limestone. A drift runs east and west 150 feet, at the level of the bottom of the up-raise, in this body of unreplaced limestone. This body of limestone differs from the smaller masses of lime-sand hitherto observed, in that the ore deposition has gone on above rather than below it.

Gray Porphyry dike.—The dike lies immediately north of this body of unreplaced limestone. So far as observed it nowhere reaches the rock surface within this claim, but ends at the top in a rounded end, as shown in Section J. Shaft No. 3, near the Carboniferous line, is sunk through Wash into ore, and at its bottom is directly in the dike. By the outlines of the dike, shown on the sides of this shaft, it is seen that it here stands nearly vertical, dipping at a steep angle to the north. Drifts to the north and east from the bottom of the shaft pass out of the Gray Porphyry directly into the ore body, and cross-cuts south from the main eastern drift strike it again, in some cases stopping at the dike, in others passing through or over it to connect with the south workings. The ground along the Carboniferous-Little Chief line on the line of the dike was, at the time of visit, a mass of crushed

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[1] The term "iron," as used in these descriptions, is the miner's abbreviation for vein material carrying more or less iron oxide.

timbers, filling old stopes which were completely inaccessible, so that its connection between observed points in Little Chief and Carboniferous ground could not be examined. For this reason, in Section B, which passes through this portion, the dike has not been represented at all, though the plane of the section would cross it diagonally, and it undoubtedly is cut by that plane in some point. It is said that it is only 4 feet in thickness at this point, and that ore existed both above and below it; this statement must, however, be accepted cum grano salis.

Immediately north of No. 3 shaft an exceptionally thick body of ore was found, composed almost entirely of sand carbonates, mixed with a certain amount of clay and iron oxide. It extended at its maximum development eight sets of timber above and two below the level from the bottom of that shaft, or about ninety feet vertically, connecting to the westward with the Carboniferous ore body.

The newer workings of the mine are opened by two large three-compartment shafts, the Daly shaft and No. 4 shaft, from which regular rectangular systems of drifts are run. The greatest and most continuous development of ore is along the northern flanks of the dike, but the great thickness of pay ore found near No. 3 shaft seems to be local in character, as at 120 feet north it has decreased to 10 feet in thickness, and at the bottom of No. 4 shaft it is only five or six feet thick. No very large bodies of ore have been found north of No. 4 shaft, but a number of small chambers and pockets have been found south and west of the Daly shaft. This shaft passed through 103 feet of Wash, 20 feet of decomposed White Porphyry, and 50 feet of silicious iron. To the north of it several small bodies of dark, coarsely-crystalline blue limestone were found in the vein material, but no considerable ore bodies. The formation, as shown on Section J, is horizontal, or rising a little to the north, but to the northeast, beyond the Daly shaft, it soon commences to dip at a considerable angle, and yields considerable water, which forms a serious impediment in prospecting. Except in the northeastern portion, the Little Chief ground may be considered to have been very thoroughly prospected, and, as shown by the map, little or no useless expense has been incurred in prospecting at the southern end of the claim, where the ore horizon has been removed by erosion. The ore itself differs in no essential point from that taken from the adjoining mines. In the commencement it was smelted in a furnace, belonging to the company, situated near the shafts, and built on a very uncertain foundation, as with large chambers opened so near the surface the ground was bound to settle. It was soon found impracticable, moreover, to smelt with the ore of a single mine, and the advantage gained in transportation for its own ore was more than counterbalanced by the cost of that brought from other mines. This scheme was, therefore, soon abandoned, and the slags were afterwards used as a low-grade ore.

Little Pittsburgh.—Besides the New Discovery claim, already described, the Little Pittsburgh Company owns also the Little Pittsburgh and Dives claims, which occupy the area between the Little Chief and Amie claims, and overlap each, so that a compromise boundary line has been adopted in either case between them. As in the ground previously described, there are two distinct ore bodies; the one at the outcrop, the other immediately north of the dike. The dike itself is here more clearly defined than before, and stands with a dip of 70° to the north. In the body of vein material are found several thin sheets or stringers of porphyry, probably offshoots from the sheets of White Porphyry, which in the adjoining ground of the Amie mine have split the Blue Limestone, now represented by sheets of vein material, into three distinct portions, as shown in Section H.

The first prospecting shaft sunk on Fryer Hill was the Little Pittsburgh No. 1 shaft, and by a singular coincidence not only is this the point where the overlying Wash has the least thickness over the entire surface of the hill, but it is where the rock surface is highest west of the Amie claim, and is in the midst of one of the most important ore bodies of the region. The shaft is 36 feet deep, of which depth 20 feet is in Wash and 16 feet in ore. Near the bottom of the shaft is a large bowlder of Sacramento Porphyry which has fallen from the Wash, and whose under surface is polished and striated, showing that in its passage from the head of Evans gulch it probably was fastened in the bottom of the Evans glacier. The ore body opened by No. 1 shaft is only the relic of a much larger mass that has been partially removed by erosion, as is shown graphically in Section I. It will, therefore, be understood that the description applies to this relic, not to the original body. To the south it thins out rapidly, having diminished to 4 feet at 40 feet and to 18 inches at 70 feet from the shaft. In the end of the south drift, underlying White Porphyry and overlying Wash are both visible. An east-and-west drift explores the whole width of the ore body, reaching continuously into Little Chief ground. The body is nearly level, and in the southern portion has a slight inclination to the south. Its greatest thickness is from sixteen to twenty feet. Wherever its upper surface has been reached, the Wash is found resting directly on it. West of No. 1 shaft the underlying White Porphyry comes up to the level of the east-and-west drift just south of that drift and dips gently to the northward on the other side of the drift. East of the No. 1 shaft, near the boundary of the claim, the ore horizon consists principally of chert. Following this boundary northward the chert passes into black iron, and contains thin sheets of White Porphyry from one foot to two feet in thickness. In the abandoned workings just south of No. 2 shaft a winze was sunk 120 feet, entirely through White Porphyry, which was finally abandoned on account of great influx of water.

At No 2 shaft the ore body was 17 feet thick, and lay immediately beneath the Wash. About thirty feet north of this shaft the first White Porphyry in place was found overlying the ore. North of this line the ore horizon, which hitherto has been very flat, dips rapidly to the north, the incline which follows it descending 20 feet in a distance of 50 feet. From the foot of this incline run connecting drifts to the northern body, which develop, on the ore horizon and immediately above the underlying porphyry, masses of manganiferous iron and compact reddish chert, coated frequently with crystals of pyrolusite. The chert which is developed in the ore horizon, and which is one of the normal replacement products of the Blue Limestone, though very similar, yet differs somewhat from the concretions of chert found in the unaltered limestone, and which are very commonly included in the White Porphyry immediately above the contact. The latter is always compact and homogeneous, while the former readily splits into angular fragments, and its joints are frequently coated with delicate crystals.

The connecting drifts from the foot of the incline pass through this barren vein material, or through the White Porphyry under it, and crossing the Gray Porphyry dike, reach the northern ore body beyond it.

The foot-wall of the dike is here not very distinct, but the hanging or northern wall has a smooth, well-defined face, standing at an angle of 70°. The porphyry of the dike is often very much iron-stained in this ground, for which reason on the foot-wall it is sometimes almost impossible to distinguish between vein material and dike. On the hanging wall, however, there is generally a sort of clay selvage, with polished surfaces. The outline of the dike is very irregular, as shown by the fact that the hanging wall in this ground varies in angle from 75° to 45°, though the steeper dip is the prevailing one.

In the northern body the rich ore comes directly up to contact with the dike. It consists mainly of hard carbonate. Near the No. 4 shaft it is very thick, averaging about thirty feet, and in one part reaching 45 feet. It is practically continuous eastward to the No. 3[1] shaft, where it is again 30 feet thick, and beyond that into the Amie ground. To the northward, however, the rich ore bodies are very irregularly distributed in the ore horizon, and the ore horizon itself is apparently rather irregular. It has a general steep dip to the northward, and in the eastern part of the mine a tendency to dip also to the northwest. The boundaries between the rich ore bodies and the black iron or chert are very abrupt also, and often confounded with those of the formation. As the drifts were mostly run with no other system than to follow these rich ore bodies, it was very difficult, in the absence of any systematic mapping of the underground workings, to form a clear conception of the ore horizon and all its dips and strike.

As an instance of the confusion brought about in the minds of those working the mines by this want of system, it may be mentioned that a drift was run back southward from near No. 4 shaft into the porphyry dike for 30 feet, and then a raise was put up in search of the ore bed, which was abandoned, after being driven up 35 feet, on account of the danger of caving as they approached the Wash. The manner in which explorations were carried on from No. 5 shaft, to the north of No. 3, further illustrated this. The bottom of the shaft was in chert, which here forms the upper part of the ore horizon. A drift run north passed out of this chert in a

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[1] The number of this shaft has been omitted on the map. It can readily be distinguished, however, by its position near the eastern boundary line and a short distance north of the dike.

distance of 10 feet, and was then continued 70 feet in the overlying porphyry, at every foot increasing its distance from the ore horizon. The main level from this shaft running northeast also passed out of the chert into the overlying porphyry, and at about forty feet from the shaft a winze was started to search for the ore below; this was, however, abandoned after going 15 feet, and an up-raise was started which was persistently continued in the overlying White Porphyry to a height of 70 feet, when the Wash was reached.

Under these circumstances it is difficult to say how thoroughly the ground to the north has been prospected or whether the failure to find ore bodies there is to be taken as a conclusive proof that none exists. Owing to the steep dip of the formation a level was soon reached by exploring drifts, at which the influx of water was too great to be handled by the pumping appliances in use, and exploration became expensive and was easily discouraged when rich bodies were not readily found.

No. 6 shaft was sunk to a depth of over two hundred feet, passing through 93 feet of Wash, 75 feet of White Porphyry, and 42 feet of vein material with a porphyry streak in the middle, into Parting Quartzite, and then into the lower sheet of White Porphyry. Drifts to the northwest from this shaft find small masses of dark crystalline limestone in the vein material, similar to that found near the Daly shaft, in Little Chief ground. The two northern shafts, No. 8, on Dives ground, and Winnemuck shaft No. 7, had not reached the ore horizons at the time of examination, but had passed through a sheet of Gray Porphyry above the White Porphyry. This is probably a part of the main sheet of Gray Porphyry corresponding to that in Little Stray Horse Park, which once covered the whole of Fryer Hill, but has since been removed by erosion.

Beyond the limits of the Little Pittsburgh claim the Four Per Cent. shaft reached the ore horizon at a depth of about one hundred and sixty-five feet, finding vein material, but, so far as known, no considerable bodies of ore

Amie mine.—The Amie claim is very nearly parallel and next east to the Little Pittsburgh, and the rich northern ore body of the latter, as well as the porphyry dike, can be traced continuously from one into the other.

The porphyry dike is, as before, well defined on the hanging-wall side, having a clay selvage and some appearance of slickensides; its angle of dip is no longer as steep, averaging from 45° to 50°, and its thickness is also very variable, at one point being only 18 feet, at others thirty to forty feet, and in one case a drift was run in it 70 feet, and a raise was then made up to the Wash. It must be borne in mind, however, that the portion of the dike exposed by the few mine drifts which cut it is very small, relative to the whole mass, and that the variation in dip may, in many cases, only represent irregularities in the form of the body, and not variations in the dip of the mass as a whole.

The stringers of porphyry seen in the Little Pittsburgh ground have here enlarged into extensive sheets, which split up the ore horizon into three portions. The upper portion represents the greater part of the Blue Limestone body and furnishes the main supply of ore, the second and third ore bodies being simply irregularly-shaped portions, which were separated at the time of the injection of the porphyry, and have since been changed to vein material by the action of the ore currents. As these lower bodies have yielded but little pay ore, they have not been as thoroughly explored as the upper one, and their outlines, as given on Sections A and H, are more or less hypothetical.

The ore of the Amie mine is, as a rule, much richer than those already described. Even the iron vein material often averages ten to twelve ounces per ton in silver, in large masses, and, being comparatively free from silica, has been profitably employed as a flux by the smelters, in place of the Breece Iron ore which they had hitherto been using, and which was relatively much more expensive. The rich ore, mostly dark sand carbonates, generally occurs at the top of the ore horizon, immediately under the overlying porphyry, a clayey, iron oxide, with more or less manganiferous or black iron, forming the base of the horizon. Chert is much less widely developed than in the previously-described mines. A considerable amount of so-called "Chinese talc" is found throughout the rich ore bodies, doubtless the product of alteration of stringers of porphyry in the original limestone. South of the dike no considerable quantity of rich ore had been found at the time of examination, as the map shows; explorations had, however, by no means covered all the possible ground in which it might occur, so that the statement, that the southern ore body previously observed does not extend as far east as this, rests on rather negative evidence. In only one point in the southern workings had a raise been made which found the Wash resting directly on the ore horizon. Elsewhere the covering of White Porphyry still remained. The actual width of the outcrop of the ore horizon in this ground is deduced from observations in the adjoining mines.

The workings of the Amie mine have been intelligently and systematically conducted from the very commencement, so that it has had advantages in the cost of extraction of ore over other mines, and has been able to mine even the low-grade iron at a profit. Two compartment shafts, No. 1 and No. 2, each provided with cages, were sunk entirely through the first, and at that time the only known, ore horizon, near the east and west limits of the claim, respectively. These were connected by a main level, provided with a tramway, from which cross-drifts underrun the main ore body, so that in mining the ore requires but one handling, falling directly from the stopes, through ore shoots, into the mine cars in which it is taken to the surface. Shafts No. 3 and No. 4 were sunk later, to explore the ground to the north and south of the main ore body, respectively.

The thickness of different rock formations passed through by these shafts will serve to show their irregularities and part of the data on which the sections have been constructed. They are as follows:

	Wash.	White Porphyry.	Iron vein material.	White Porphyry.	Iron vein material.	White Porphyry.	Iron vein material.	White Porphyry.	Silurian formation.
	Feet.	*Feet.*	*Feet.*	*Feet.*	*Feet.*	*Feet.*	*Feet.*	*Feet.*	*Feet.*
No. 1 shaft	75	55	16	20	4	20	....	....	....
No. 2 shaft	75	25	10	70	15	20	4	46	77
No. 3 shaft	125	25	5	40	25	....	....	....	....
No. 4 shaft	50	25	25	12	3	12	3	60	....

The main level, which has an elevation of 10,365 feet, is 150 feet and 160 feet below the collars of No. 1 and No. 2 shafts, respectively, the collars of these shafts being placed about ten feet above the ground to allow space for the dump. No. 2 shaft, it will be observed, has been sunk to a considerable depth in the Silurian formation, in which it has developed the

White Limestone, with its characteristic white chert segregations, but no ore. It was intended at time of visit to continue it still farther, as soon as a Cornish pump could be put in to control the great influx of water, which is almost invariably found when a certain depth is reached. Although there is no geological impossibility of the occurrence of ore in depth, the facts of observation are so unanimously adverse to its probability that this may be said to be a misdirected expense of labor and money, and one which, if devoted to the exploration of the Blue Limestone horizon in any of its various subdivisions, would be far more likely to yield practical results.

The main ore horizon north of the dike has a general dip to the northeast, although, as defined by its general contact with the underlying porphyry, it inclines locally to the northwest near the Climax boundary. The main dip of the formation is, however, to the north, and, as in the previously described mines, this dip steepens rapidly in the northern part of the claim, though the gentle dip continues some distance north of the main body. The first lower sheet of porphyry, as developed by the drift driven from No. 1 shaft to connect with No. 3, is remarkably full of chert fragments, most of which appear to have been simply caught up in the porphyry flow; some, on the other hand, are apparently segregations in the mass of the porphyry since its consolidation.

The principal ore body occurs along the north flank of the dike, in some cases being seen to wedge out between this and the overlying White Porphyry. It is very variable in thickness; thus at No. 2 shaft the whole horizon is only 10 feet thick; at 70 feet to the southward it has thickened to nearly 50 feet, of which the upper 30 feet are in pay ore, mostly rich sand carbonates. A similar large body of sand carbonate, 20 feet in thickness, was found above the main level east of No. 1 shaft, which was 45 feet in length. As already mentioned, most of these bodies are in the upper part of the ore horizon; rich ore also occurs irregularly in different parts of the horizon and also in the lower ore sheets, though the latter contain as a rule a smaller proportion of pay ore. Explorations to the northward, as far as conducted, find a large proportion of barren ground in the ore horizon, and, as elsewhere, the influx of water as the formation descends renders exploration difficult and expensive. Several small ore bodies have, however, been opened by No. 3 shaft, which are sufficient to prove that ore does exist in this direction and to justify further exploration.[1]

Climax mine.—The Climax claim is parallel to and adjoins the Amie on the east. The structural conditions are, however, somewhat different in the two claims. The ore horizon is still split up into several parts, but, owing to the erosion of the crest of the anticlinal fold, which runs northeastward along the east boundary of the Climax claim, a much greater proportion of the ore horizon has been eroded off the area of the claim, and the outcrop of what remains runs northeastward nearly parallel to its side lines.

The mine has been worked only intermittently and without much system, and, as a considerable portion of the workings were inaccessible at the time of visit, information in regard to them could only be obtained by word of mouth, and leaves much to be desired in point of completeness and reliability. The general outlines of the structure were, however, sufficiently well determined by the examination of those workings which were accessible, and the uncertainty exists mainly with regard to details of ore distribution.

The mine workings consist of two disconnected groups, a southern and a northern, the former of which followed the eastern extension of the Amie body, the latter the western extension of the Dunkin body. Between these are the contract or leased workings, which, as their name implies, were worked by other parties under leases, of which no plats could be found, and about which little information could be obtained.

The southern workings are opened by shafts No. 3 and No. 5, shafts No. 4 and No. 6 having been sunk independently to explore the ground further south and not connected with these workings. Of No. 6 it is only known that it was sunk through 160 feet of Wash and reached a body of iron vein material in the top of the White Limestone. Shaft No. 4 cut two bodies of vein material, which are probably part of the lower ore horizon of the Amie mine, before reaching the White Limestone. It would seem probable that the drifts from this shaft might have cut the porphyry dike. Unfortunately at that time miners made no distinction between White and Gray Porphyry, and no definite information on that point could be obtained.

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[1] Since the close of field-work a considerable body of rich ore is said to have been opened by the Deer Lodge shaft, which is situated near the Climax line, not far from the Virginius shaft.

The first level from No. 3 shaft, at 100 feet below the surface, starts in the White Porphyry below the upper ore horizon, which pitches at 20° to 25° towards the Amie line. Drifts were run to the north and west and to the southwest on this level; the former passes at 20 feet from the shaft into the upper ore horizon, whose lower portion consists of soft red silicious iron, passing into soft black iron or into chert; turning westward it reaches the continuation of the sand carbonate ore body of the Amie, from which considerable rich ore was taken. The southwest drift runs mostly in the underlying porphyry, in which are several thin streaks of ore and of Chinese talc. To the west of this drift, near the Amie line, are old stopes, from which a peculiar white sand ore, lying at the top of the ore horizon, was taken. This ore is said to resemble a decomposed porphyry so much that at first it was supposed to be worthless, but on examination was proved to be extremely rich, assaying as high as 1,600 ounces of silver and giving mill runs of 300 ounces, but containing little or no lead. It is probably the result of a leaching of the original ore body, during or subsequent to the process of erosion.

No. 5 shaft was sunk later and passed through 125 feet of Wash, 5 feet of iron, and 110 feet of White Porphyry, stopping in the Parting Quartzite and White Limestone, which here dip gently northward. A drift to the northwest, on the 383-foot level, from this shaft, and a western branch from this drift in the direction of No. 3 shaft find a small body of iron in the porphyry, which may correspond to the second ore horizon of the Amie mine. A drift to the northward, on the other hand, finds quartzite in the midst of the porphyry, which is supposed to be a detached portion of the Parting Quartzite, as shown in Section G.

The northern workings are opened by No. 1 and No. 2[1] shafts, and also connect with the northern Contract shaft. The workings from the latter have developed little of importance; the shaft was sunk through 85 feet of Wash and 30 feet of iron vein material. Drifts to the south and west rise to the overlying Wash at their extremities, but develop no ore bodies. To the north the workings follow some thin streaks of pay ore standing nearly vertically in the iron vein material.

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[1] The number of this shaft has been omitted on the map; it lies near the Dunkin line, about one hundred feet southwest of No. 1.

Climax No. 2 shaft is 138 feet deep, of which the upper 90 feet are in Wash and the rest in the body of vein material, which is here separated into two parts by a sheet of very compact White Porphyry, as shown in Section E. This dividing porphyry, which is only 18 inches thick in the northwestern part of the workings, dips to northeast, increasing in thickness as it goes down, reaching four feet at the Dunkin line, and at Dunkin No. 2 shaft merging into the main lower body of White Porphyry. The principal ore body, which is found directly above this dividing porphyry, has a thickness of about 8 feet, increasing to 16 feet along the line of the Dunkin claim. The vein material here consists largely of highly manganiferous iron, with clay and chert generally at the base. The workings northeast of No. 2 have been carried up to the Wash, which here consists of sand and rounded pebbles, without meeting clay or the influx of water which are almost invariable accompaniments of the Wash. At Climax No. 1 shaft the ore horizon was found directly below 100 feet of Wash, in a single body 38 feet thick. This shaft was sunk to a depth of 220 feet, reaching some iron vein material in the bottom, which is supposed to be at the top of the Silurian formation, as in Climax No. 6. The lower north-and-south drift from No. 2 shaft, which is mainly run in the White Porphyry below the ore horizon, also cuts White Limestone at its southern end.

Virginius mine.—The extreme southern end of the Virginius claim overlaps the northern end of the Climax ground, and along the south line of the former two shafts have been sunk to the ore horizon and connected by drifts and winzes, the workings descending in steps towards the west. The main, or No. 2 shaft, was sunk through 136 feet of Wash, 40 feet of porphyry, and forty to fifty feet of vein material, passing at the bottom into dolomite and sand, which were dipping northwest. The vein material here was impure, containing much manganese, with clayey and sandy streaks. North of the shaft a cave was found near the top of the iron body, ten to fifteen feet in length and four to five feet in height.

Drifts and stopes connect with the No. 3 shaft, near which a small body of ore was taken out, which ran about forty ounces of silver to the ton. The main body of vein material in this part of the mine carries from two up to ten or fifteen ounces of silver to the ton. The east drift from this shaft was cut through the bottom of the ore horizon into underlying White Porphyry. Explorations have not been carried far to the northward on this ground, owing to the steep dip and great influx of water. Probably the yield of the ore did not seem to justify the expense that would be necessarily incurred in putting up a pumping plant capable of handling the water.

On Section G the continuation of the ore horizon to the north has been represented as unreplaced limestone, simply because it has not been proved to contain vein material, though it is impossible to say whether it does or not until it has been actually explored.

Dunkin mine.—The Dunkin claim lies next east to that of the Climax. From it a large amount of rich ore has been obtained, and exclusively from the continuation of the northern shoot, observed in the Climax ground. The eastern continuation of the main ore shoot of Fryer Hill, which lies on the north flank of the porphyry dike, has, as the map shows, been entirely eroded off the Dunkin ground and the claims to the south of it. Whether the outcrop of the porphyry dike is entirely wanting between the Amie and the northeast corner of the Big Pittsburgh claim, as represented on the map, is not known, since there are no underground explorations in this area from which data may be obtained. It is most probably continuous in depth, but has not been indicated as outcropping, on the principle of representing as far as possible only what is actually known. The great breadth of outcrop of the ore horizon on the Dunkin ground is due to the fact that it lies along the crest of an eroded anticline. There is some evidence to show that in some part of the area covered by this outcrop patches of White Porphyry still remain between the vein material and the Wash, but it is not sufficiently definite to locate or outline these patches, and their existence does not invalidate the general truth of the structure, as given by the outlines on the map.

The Dunkin mine is opened by three shafts, No. 1, at the south end; No. 2, in the middle; and No. 3, at the north, as shown in Section D. Besides these is an old No. 1 shaft, which, being in an area of barren vein material, is no longer used. The main working shaft is No. 2, near the center of the claim, as well as of ore developments. From this three sets of levels are run, at 10,425, 10,405, and 10,357 feet elevation, respectively. The Wash was here 90 feet and the ore horizon 40 feet thick, the first level starting about fifteen feet below the Wash and the second near the base of the ore horizon. The third level is in the Parting Quartzite at the shaft. No. 1 shaft was sunk through the underlying White Porphyry to the White Limestone, and is connected with No. 2 only on the second and third levels. Old No. 1 shaft found twenty to thirty feet of vein material above the porphyry. No. 3 shaft found the vein material directly beneath the Wash, and was sunk through it into Parting Quartzite and underlying porphyry. Only the third level connects directly with this shaft.

The most important ore body occurs between the first and second levels, extending southeastward from No. 2 shaft into the Matchless ground. It averaged from ten to sixteen feet in thickness and perhaps forty feet in width. Both ore body and ore horizon dip to the eastward on this side of the shaft, at an angle of about 15°. The drifts from the first level pass, to the southeast, rapidly into a body of black iron above the ore body and, to the southwest, into black iron and reddish silicious iron. At the end of a drift to the south a coarse sand is found at the top of the ore horizon, which in some cases is found to be impregnated with silver, and constitutes a rich ore. The west drifts in the second level, after passing through comparatively barren vein material, cut diagonally across the parting sheet of White Porphyry, which has already been noticed in the Climax ground, and reach the eastern end of the Climax ore body, immediately underlying this sheet of White Porphyry. The ore here consists of galena and sand carbonates.

In addition to the ore bodies above mentioned, later explorations have discovered numerous small bodies or patches of ore in the upper part of the ore horizon, immediately under the Wash. An interesting occurrence here was a mass of angular fragments of White Porphyry, cemented together by galena. The rich white sand noticed in the Climax ground was also found here in places. The galena in this mine is generally coarse grained, and sometimes exceptionally rich in silver; as elsewhere its tenor in silver is usually higher than that of the carbonates. A mill run of galena from the upper workings yielded 500 ounces of silver to the ton.

On the third level no pay ore has been found, but the developments are interesting from a structural point of view. It runs northeasterly through the middle of the claim from the bottom of No. 1 shaft to 160 feet beyond the bottom of No. 3 shaft. At No. 1 shaft, and for 40 feet north of it, it runs in White Limestone, dipping 20° northeast, which is more or less stained, and occasionally replaced by clayey iron oxide. It then runs into decomposed and iron-stained porphyry, which is in places so laminated that it might be mistaken for a shale. The Parting Quartzite, which is disintegrated and contains thin layers of bluish shale, comes in at 160 feet from No. 1 and continues for 100 feet, lying nearly horizontal, and probably represents a minor roll in the formation, as shown in Section D. Beyond No. 2 shaft the quartzite gives way to compact White Porphyry, in which a cross-cut to the east shows the iron body resting on it and dipping eastward. At the bottom of No. 3 shaft the Parting Quartzite is again cut, here being above the White Porphyry and immediately under the iron body or ore horizon. The drift then runs for 80 feet through the iron body and suddenly passes into decomposed Blue Limestone, which, on the sides of the drift, has all the appearance of the solid unaltered rock, showing the stratification planes dipping northeast at 40°, the characteristic ribbings of white spar, and an occasional fossil resembling a *Euomphalus*, but which when taken into the hand immediately crumbles into fine lime-sand. A partial analysis of this lime-sand is given in Appendix B, Table VI, which shows it to have the normal proportions of lime and magnesia contained in the unaltered rock. Toward the end of the drift the dip shallows, probably because it is becoming more nearly parallel with the strike of the beds. At the very end the roof of the drift has caved, showing decomposed White Porphyry immediately above the limestone and Wash a little distance above that This point, it will be observed, is almost opposite the workings of the Virginius claim. Beyond it the ore horizon has not been explored, nor is it likely to be until powerful pumping machinery is introduced capable of controlling the great influx of water.

Matchless.—The Matchless ground, which lies next east of the Dunkin claim, has been relatively little explored, probably because in early days it was considered unpromising ground, since the few prospecting shafts that were sunk did not strike rich ore. The indications afforded by the map, which show the condition of explorations at the time of this examination, show, however, that both the Dunkin or northern ore body and the Lee body, which is the main ore shoot of the hill, extend into it, and may be reasonably expected to join together on this ground. Moreover, it has a large extent of unexplored ground in the northeastern part of the claim, which, though less promising than the southern part, is certainly worth prospecting.

The Discovery shaft and the Main shaft were sunk, the one to the south, the other to the north of the continuation of the Dunkin ore shoot. The Main shaft was sunk through 110 feet of Wash, 20 feet of White Porphyry, 15 feet of chert, 30 feet of iron, and through the underlying White Porphyry to the Silurian formation, which it reached at a depth of about two hundred and fifty feet. (See Section E.) In a drift on the top of the iron body, several small layers of lime-sand were found immediately under the overlying White Porphyry, which was itself much decomposed and full of segregations of iron oxide The ore horizon, where cut by this shaft, contained little or no pay ore, but where the Dunkin ore body was found to extend to the Matchless line it was followed into the ground of the latter. Here it has a width of about forty feet and is from eight to sixteen feet in thickness. It extends in a northeasterly direction and descends rapidly to the eastward. The vein material is a cherty or silicious iron, and the pay ore a reddish clayey mass of sand carbonate,yielding much lead and silver. At the time of visit no connection had been made between this ore body and those to the east and south.

The Leonard or southern shaft was sunk to strike the continuation of the Lee body, which had been found to extend across the wedge-shaped portion of the Hibernia claim, between this and the Lee ground. It was sunk through 95 feet of Wash, 10 feet of ore, 15 feet of chert, and 12 feet of quartzite to the underlying White Porphyry. As far as at present explored, the rich ore body is confined to a narrow strip of ground along the Hibernia and Big Pittsburgh lines. It lies upon the Parting Quartzite, either directly or separated by a floor of chert, and therefore occupies the very lowest portion of the Blue Limestone horizon. In the northeast corner of the claim it abuts directly against the Gray Porphyry dike, which still dips to the northward at a steep angle. Although narrow, the ore body is very thick, reaching 30 feet in places, and is extremely rich. This ore differs from any hitherto observed on Fryer Hill in that it is almost entirely free from lead. Its silver exists as fine particles and films of chloride or chloro-bromide, disseminated through an ocherous sandy mass, and sometimes coating the cracks and cleavage faces of the chert. Another difference between this ore body and those of the portion of the hill already described is the small amount of manganese found, and the condition of the iron oxide, which is here more generally anhydrous, whereas in other parts of the region it is always hydrated or in the form of limonite.

Hibernia and Big Pittsburgh.—These two claims will be described together, as the only portion of them yet found productive is the extreme northern edge, where the western continuation of the Lee ore body extends a short distance across their lines. The Gray Porphyry dike is here about thirty feet wide and very well defined, cutting across the Blue Limestone horizon into the underlying White Porphyry; and pay ore has thus far been confined mainly to its northern flanks, though, as will be shown later, there is good evidence for assuming that the continuation of the southern ore shoot, as developed in the ground to the westward already described, once existed here also, and that it should be sought for further east, where the ore horizon has not been removed by erosion.

The Hibernia shaft, which was sunk just south of the dike, passed through 100 feet of Wash into soft black iron, with chert at the base and a little pay ore, having a total thickness of about twenty-five feet. Drifts run northward from the shaft across the dike to connect with the stopes in the little triangular or wedge-shaped point of the claim beyond. These stopes were five sets of timber high, and the little triangular area was an almost solid mass of rich ore; near the top was a layer of chert extending from the Lee ground, which was there supposed to be the top of the ore body; it was here broken through and the ore found to extend up to the Wash. The quartzite floor dips northward into the Matchless ground. Southward from the shaft a prospecting drift runs over two hundred feet in the underlying White Porphyry, striking the Parting Quartzite at the end. A cross-cut to the eastward from this drift finds barren iron resting on the White Porphyry, and a winze sunk in the floor of the drift is said to have found White Limestone.

Westward from the shaft, a drift runs through White Porphyry, which connects with the McCormick shaft of the Big Pittsburgh, on the dividing line between the two claims. This shaft was sunk by Mr. Tingley S. Wood, superintendent, after explorations in the southern portion of the claim had proved fruitless, with the idea that a portion of the Matchless body might be found within the Pittsburgh lines. His expectations were realized, and a narrow strip of very rich ore was found north of the dike, and directly under the Wash, being, as the map shows, the extreme southern or lower edge of the Blue Limestone outcrop. From this shaft two cross-cuts were run northward across the porphyry dike toward the north line of the claim, and an up-raise made along that line disclosed the ore body above the White Porphyry or Parting Quartzite, as the case might be. Up-raises were also made in the porphyry dike, which showed that it extended up to the Wash, or, in other words, outcropped. Owing to a surveyor's error the line drift was run on the Matchless side of the boundary line; but, the error being discovered, the Matchless was reimbursed for the ore taken from its ground. What is known about the balance of the ground owned by the Big Pittsburgh Company will be given below in the description of the southern portion of the map.

Robert E. Lee.— This claim, in spite of its small area, has been among the greatest silver producers in the district. It has been owned by different individuals, and for various reasons it has not been possible to obtain very trustworthy figures with regard to the actual value of its product. The ore has been remarkable for its high tenor in silver and its freedom from lead. It is also very silicious and contains a relatively small percentage of iron, for which reason it is by itself not so well adapted for smelting as the average ore of the district, and a great deal of the low-grade ore from the mine has been reduced by amalgamation. The ore horizon is here directly overlaid by a body of Gray Porphyry, whose thickness could not be ascertained. It is evidently of limited extent, as it was not cut by the shafts of the adjoining claims. It may be an offshoot from the dike, or, as indicated on the map, simply a small intrusive sheet. In the western portion of the claim this porphyry-covering, together with the normal sheet of White Porphyry and the main Gray Porphyry sheet above that, has been eroded off and the ore body extends up to the Wash. A thin quartzite, evidently belonging to the Weber formation, is often found directly above the ore horizon. The main ore body was almost perfectly continuous, varying in thickness from a few inches up to twenty-five feet, and generally overlaid as well as underlaid by dark-blue chert. At the time of visit a layer of ore was being followed which consisted of barite thoroughly impregnated with chloride of silver. The rich ore is sometimes a red sandy or clayey mass, and sometimes consists of chert or silicious iron, whose cracks and joints are lined with chloride of silver. The ore in general, as it comes from the mine, is characterized by its bright-red color, due to the presence of anhydrous iron and absence of manganese oxide.

The principal working shaft of the mine at time of visit was the No. 2, from which two levels were run; the old No. 1 or Ladder shaft was no longer used for the extraction of ore, and the new shaft to the northeast of these, designed to open the ore body on the dip, was not yet working.

The main thickness and the richest portion of the ore body lay to the south of shafts No. 1 and No. 2, between these and the dike. Directly south of No. 2 is a small irregular sheet of Gray Porphyry, cut in the lower level in a thickness of four to six feet, which seems to run partly with the stratification and partly across it. Too little of this body was exposed to afford sufficient data for determining its extent or origin, but it evidently acted favorably on the concentration of ore in its vicinity, probably by arresting the flow of the ore-bearing solutions and giving them time to precipitate the minerals they held in solution. The drifts in the western part of the mine had been extended south until they reached the porphyry dike, but, singularly enough, in the eastern part of the mine they stop before going so far south, it seeming to have been taken for granted that the dike would cut off the ore indefinitely in that direction; whereas there is every reason to believe that at no great distance to the eastward it will continue south over or across the line of the dike. It is hardly necessary to say that the outlines of the eastern end of the dike, as given on the map, are consequently founded only on general probability, there having been no exploration to determine its exact limits. The ore horizon in the Lee ground has a relatively steep dip to the northeast, which may be taken as averaging about 25°. Explorations on the dip to the northeast and northwest find the ore more irregularly distributed throughout the horizon and not so concentrated as in the older workings; nevertheless they indicate an extension of ore deposition in that direction sufficient to justify more extended explorations. In the early days no maps were made of the underground workings, the services of surveyors being only called upon from time to time to determine points for the connection of drifts and for the location of shafts on the surface. Those given on the map for this mine are the result of rough surveys made by us in the course of our examination, checked by measurements kindly given by the surveyors who had at various times been employed in the mine. They represent only the principal drifts which were in use at the time of examination, the intermediate ground being largely occupied by stopes and drifts no longer used.

The later workings are systematically conducted from two main levels, the 320 and the 350 foot, the station of the former being 192 feet below the collar of No. 2 shaft. The No. 4 shaft of the Lee claim, on the south side of Little Stray Horse gulch, finds the iron body directly under the Wash. As yet little attention has been given to this portion of the claim, although it certainly deserves it, as from analogy with other parts of the hill it would seem as likely that rich ore bodies should exist under the lee of the dike here as there, and they might extend still farther eastward.

Little Sliver.—On this claim, which lies next east of the Lee, a commencement of exploration of the ore body has been made, and very promising ore deposits are being found. The Sliver shaft was sunk through about one hundred and twenty-five feet of Wash to Gray Porphyry, and found the usual thin bed of shales and sandstones at the contact of this with the White Porphyry, which were here more or less replaced by iron vein material.

The Tip Top shaft, still further eastward, a little beyond the limits of the map, found these shales, with a certain amount of carbonaceous material, at a depth of 245 feet. In them were some small pockets of galena and carbonate ore.

Southeast corner of region mapped.—A considerable area still remains in the southeast part of the region represented on the map, from which the ore horizon has not yet been eroded. It has as yet been but little explored, partly because of its deep covering of Wash and of the great influx of water due to its position on the western rim of the Little Stray Horse Park basin, and partly because the possibilities of the existence here of valuable bodies of ore have not been generally understood The only actual developments thus far made have been by the Surprise shaft, on the May Queen claim, and by the Denver City shaft, on the claim of the same name. The former found vein material directly beneath the Wash, at a depth of 140 feet, consisting largely of chert and black iron at the base, with soft, clayey, low-grade ore above. An incline was run, following the pitch of the ore shoot to the southwest, although the dip of the formation is here to the eastward, as was soon shown by the western drifts, which cut the Parting Quartzite beneath the ore horizon. Some good chloride ore was afterwards found by up-raises which reached a higher portion of the horizon.[1]

The Denver City shaft, in the extreme southeastern corner of the map, is nearly on the crest of the moraine ridge which borders Stray Horse gulch on the north. The Wash was here 180 feet deep, beneath which the main sheet of Gray Porphyry was found in a thickness of about twenty feet. Under this was a thickness of some twelve feet of calcareous sandstone and shale, containing some low-grade ore, which was at first supposed to represent the ore horizon, though it is in reality only the irregular parting of Weber Shales left between the Gray and the White Porphyry. The true ore horizon was afterwards struck at a depth of 234 feet, and rich pockets of chloride ore were found in it. It was passed through by the shaft for about fifty feet, ending in a bed of chert, with White Porphyry, so full of chert fragments as to be called by the miners a conglomerate, below it.[2]

There is no question that a part of the Blue Limestone is already opened by the works of this mine, but the shaft is located so near the imaginary southeast-and-northwest line, where the lower White Porphyry cuts across the Blue Limestone, separating it into two wedge-shaped portions, that there is a possibility that a portion of this horizon may yet be left

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[1] Since the close of field-work, large and rich bodies of ore are said to have been opened in the Forest City ground, to the east of this claim.

[2] These data were obtained from Mr. Robert Bunsen, superintendent of the Denver City mine, since the completion of field-work, and are not the result of our own observations.

below the cross-cutting White Porphyry, since the Parting Quartzite, which defines the base of the Blue Limestone horizon, has not yet been reached. As it is a question of considerable economical importance for owners of property in this vicinity to know whether a second ore horizon is likely to be found beneath the second White Porphyry body, the evidence on which it has been indicated on Section C as probably not occurring beneath the Denver City shaft will be given in some detail.

The Denver City shaft is situated in strike between the Lee mine on the north and the Agassiz on the south, as may be seen by reference to the larger map of Leadville and vicinity. In the former and in the Surprise workings the whole Blue Limestone horizon is above the second White Porphyry, as evidenced by the occurrence of the Parting Quartzite at its base. The Agassiz mine, on the other hand, is near the south point of the wedge of Blue Limestone, while the greater part of this formation must be below the second White Porphyry, forming a continuous sheet, except when crossed by later intrusions of Gray Porphyry, from the outcrop on the west face of Carbonate Hill. This lower portion of Blue Limestone or ore horizon, on the other hand, must wedge out to the north and east, as the upper one does to the south and west, and the question to be decided is whether it has wedged out before the line of the Denver City shaft is reached or not. It is proved on Stray Horse Ridge, to the west of the Denver City and below the lower White Porphyry, by the Moyamensing, Joe Bates, Vanderbilt, Pierson, and other shafts, and in the valley of Little Stray Horse gulch by the Stonewall Jackson shaft. With regard to the latter, it is only known that a body of vein material has been found beneath the Wash. In the Pittsburgh shaft next north of this, however, White Limestone is found directly beneath the lower White Porphyry, showing that the wedging-out occurs between these two shafts and approximately as indicated on the map. The extension of the line of wedging-out to the southeast, which is the general direction of the cross-cutting porphyry, would pass to the west of the Denver City shaft, but in all probability not very far from it, so that there is a probability that no second ore horizon occurs there. On the other hand, as the porphyry sheets are necessarily somewhat irregular in shape, it cannot be said to be impossible that a thin sheet of vein material may be found, especially as the thickness already developed of about fifty feet is less than that found in many places, even where the entire thickness of the horizon has been replaced by vein material.

Little Stray Horse gulch.—There still remains to be described the region bordering this valley, from which the Blue Limestone or ore horizon has been removed by erosion. The data with regard to it were mainly derived from dumps of abandoned prospect shafts and from testimony of miners who had sunk the shafts, and are given by the outlines on the map, in addition to which there is not much to say. It will be understood that the relative accuracy of these outlines is dependent upon the proximity of these shafts, since there are absolutely no rock outcrops.

The Little Diamond shaft, on the Dolphin claim, just south of the end line of the Dunkin, found a considerable body of vein material beneath the Wash, which is the base of the ore horizon, where it has a local dip to the southward, as shown in Section D, the lower body of White Porphyry being exposed on the crest of this fold just east of it. The two May Queen shafts, near the base of the Denver City hill, find only White Porphyry. The Pittsburgh shaft, as already mentioned, passes through Wash and White Porphyry into White Limestone. The Little Daisy and Eudora shafts find White Limestone outcropping beneath the Wash. The new Gambetta shaft is in Gray Porphyry, supposed to be the cross-cutting sheet seen in New Discovery ground and on Carbonate Hill. This porphyry sheet was also cut at different horizons by the Eudora shaft, by the Vanderbilt on Stray Horse Ridge, by the Magnolia shaft in Stray Horse gulch, and, as already mentioned, by New Discovery shafts No. 5 and No. 6. The old Gambetta and Monarch shafts find the White Limestone directly beneath the White Porphyry, for which reason it is assumed that a portion of the Parting Quartzite has here been caught up by the porphyry and is somewhere in it at a higher horizon.

For the lower part of the valley direct data as to the outcrops are wanting. On the ridge south of it the Ida Nyce and Ypsilanti shafts have found the Blue Limestone beneath the Wash and a small amount of vein material in it. As well as could be determined, the formation has a slight easterly dip, as shown in Section C; but on the south side of the ridge in the main Stray Horse gulch there is some indication of a westerly dip beyond a slight anticlinal fold. It is unfortunate that no more exact information could be obtained with regard to it at this locality, as upon the verification of the westerly dip depends in large degree the probability of the occurrence of the ore horizon under the city of Leadville. Its existence at the west end of Fryer Hill is, however, definitely ascertained, and has been shown to be extremely probable along the west base of Carbonate Hill, which lends force to the supposition that it also occurs at this intermediate point.

RÉSUMÉ.

From the above descriptions it is apparent that, in spite of the greater complications of structure, the series of rock formations on this hill is essentially the same as that on Carbonate and Iron Hills, and that the processes of ore deposition have been essentially the same, though the secondary alteration of the deposits, which may be mainly ascribed to the action of surface waters, has been carried much farther. The Cambrian, Silurian, and Lower Carboniferous horizons are found in their normal positions, the Parting Quartzite being here, as elsewhere, of somewhat variable thickness, and the Blue Limestone horizon, which is often split up into several portions and entirely replaced by vein material, being then defined by this quartzite below and by the micaceous sandstone or quartzite of the Weber Shales above.

The intrusions of porphyry are more extensive and more varied and irregular in form. Above the normal sheet of White Porphyry, which here as there overlies the ore horizon, with detached portions of the Weber Shales left between, is the main sheet of Gray Porphyry, in great measure eroded off, which does not occur on the other hills. In addition to this, there is the second or lower sheet of White Porphyry, occurring generally at the base of the Blue Limestone horizon, but in some places cutting up across its lower portion and in others cutting down below the Parting Quartzite; further, there are ramifying offshoots from this lower White Porphyry, which have locally divided the Blue Limestone horizon into several different portions. Of later Gray Porphyry intrusions, there is the larger sheet, which, as on the other hills, is generally near the base of the Blue Limestone, but which does not conform strictly with the stratification planes, crossing them at low angles, and extending in geological horizon from the upper part of the White Limestone well up into the Blue Limestone, across the intermediate lower White Porphyry. Besides this are several smaller bodies of Gray Porphyry not found in the other hills, the most important of which seems to have the form of a transverse dike.

Of great faults like those on Carbonate and Iron Hills, there is no evidence, the force of compression having only produced gentle folds and some slight displacements of a few feet in extent, which are shown by sudden changes of level in the ore horizon; such a one has evidently occurred along the line of the north flank of the porphyry dike, which has slickensides surfaces, and shows in some cases a slight difference of level in the ore horizon on either side.

The process of ore deposition has been evidently the same metasomatic change or replacement of the limestone by ore and vein material, only it has been carried so much farther that, instead of a body of limestone with a little vein material extending irregularly from its surface downwards, there is found here only a mass of vein material with occasional irregularly-shaped residuary masses of unreplaced limestone or lime sand. Owing to the irregular distribution of the intrusive masses of porphyry, whose contact planes afforded channels by which the ore-bearing currents reached the limestone, the evidence is naturally less striking that these currents followed in general a downward course. Still it must be borne in mind that the greater mass of the present bodies of vein material are the result of secondary alteration by surface waters, and that this alteration having been much greater here, it is proportionately more difficult to trace the probable form or position of the original sulphuret deposit. In spite of this it may be observed that in the majority of cases the rich ore, which is presumably nearer its original position than the iron oxides, is found near the upper part of ore horizon. On the other hand, if the ore came directly from below, according to the idea which is generally advanced with regard to the source of ore deposits, the only channel which it could have followed would have been the walls of the porphyry dike. In this case we should expect to find evidences of the passage of the ore currents along these walls; but wherever they have been examined these evidences are conspicuously wanting. On the south flank the dike is generally separated from the ore body by a barren zone, containing often, it is true, iron vein material, but evidently of secondary origin. On the north flank the ore body extends up to the dike, but it is strictly confined to the ore horizon, and does not extend below that, the most that is found being a slight staining by iron oxides, readily accounted for by the percolation of surface waters descending through the ore horizon and carrying down some of its material with it. It is unfortunate that a more conclusive test could not be afforded by the cutting of the dike at a considerable depth below the ore horizon, but as this has not been done, we must reason from the evidence that is at hand.

The apparently abnormal variation in the thickness of the ore horizon is less readily accounted for, as has already been stated on page 446; but it must be borne in mind that the data from which the outlines of formations have been reconstructed are very limited and irregularly distributed, being derived from drifts run for the sole object of following known ore bodies and without any purpose of elucidating the structural conditions of the various strata.

The singular absence of lead in the Lee ore body is another exceptional feature of this region. It seems hardly probable that, in a district whose silver is so universally derived from argentiferous galena or its decomposition products, in this little spot alone silver should have been deposited by itself. The more natural explanation would seem to be, that the deposit is entirely secondary, and the result of the leaching of a larger body, now eroded off, by surface waters, which carried away the lead and left the silver. The geological position of the ore body favors this idea; it rests immediately on the Parting Quartzite, and therefore at the very base of the ore horizon; it is on the lower rim of a synclinal basin, which is known to carry an immense amount of water that would naturally drain out over its edges. It may be also that the absence of manganese would tend to the formation of the more soluble sulphate of lead, rather than the carbonate, which is generally found as the alteration product of galena in this district.

The influence of cross-cutting sheets of porphyry in producing a concentration of rich ore by causing a stagnation of the ore currents is shown by the distribution of the ore shoots. Thus the northeastern body in the Climax, Dunkin, and Matchless ground lies under the lee of the cross-cutting sheet of White Porphyry; the main ore shoot in the Amie, Pittsburgh, Little Chief, and Chrysolite lies in a similar position relatively to the Gray Porphyry dike; and the southern body in the last three claims lies just north of the lower cross-cutting sheet of Gray Porphyry.

The greater secondary alteration on this hill is readily accounted for by the fact that it is everywhere covered by a great thickness of Wash. This Wash, which is a loosely aggregated and permeable bowlder clay, acts like a wet sponge. It is constantly full of water at its contact with the rock surface on which it rests, which water is doubtless charged with air and decomposed vegetable matter, and thus acts more vigorously upon the rocks than would water flowing freely over the actual surface of the ground or that which percolates in minute channels through the solid rocks beneath the surface. This is shown by the fact that the upper sheet of White Porphyry, which lies immediately beneath the Wash, is generally reduced to a plastic mass, in which all trace of the original structure of the rocks is lost, while the lower sheet of the same rock is still a hard, compact rock, forming what the miners call block porphyry.

CHAPTER V.

OTHER GROUPS OF MINES.

MINES AND PROSPECTS IN THE LEADVILLE REGION.

It is from the mines included in the three groups already described that what may be considered the permanent ore supply of Leadville has been thus far derived, and it is in these mines alone that exploitation has been carried on so continuously and extensively as to afford an opportunity to study in detail the character and the form of the different ore bodies and their relations to the inclosing and neighboring rocks. For this reason they have been described with a detail that may, in the future, seem disproportionate to their relative importance, especially when, as is likely to be the case at no far distant day, the deposits of these limited areas shall have become nearly exhausted and the main supply is derived from what may now be considered outside areas. From the evidence obtained during this study it is fair to assume that a greater amount of as yet undiscovered ore exists outside these areas than has already been developed in the small groups already described, and that, while its exploitation will necessarily be more difficult, owing to greater depth and large influx of water, and its reduction will require more complicated processes, owing to a greater preponderance of sulphurets, these disadvantages will be offset by greater advantages of working, brought about by a more thorough knowledge of the geological relations of the ore deposits and by improvements introduced into the various processes of reduction.

With but few exceptions these outside mines have been hitherto but intermittently worked, and, owing to some minor differences in the character of their deposits or of their inclosing rocks, their geological structure has been more imperfectly understood by those in charge, and the work of exploration been carried on with less system and sometimes in an utterly aimless manner. Although it has been impracticable for these reasons to determine with the same accuracy and detail the relations of the ore bodies in these outside mines as has been done for those of Iron, Fryer, and Carbonate Hills, an explanation of their general geological structure will be of value as a guide for future exploitation, and a consideration of the relative amount of replacement action in different portions of the Leadville region, as shown by the developments thus far made in them, will afford a basis for determining the probable extent and direction of the original ore currents, and in consequence what part of the ore horizon, which the geological outlines have already located, is most likely to contain valuable ore bodies.

An examination of the relative distribution of vein material shown by the outcrops, as delineated in cross-lining on the Leadville map, shows two lines or zones along which the evidence of replacement action is most apparent, one running east from Fryer Hill to Little Ellen Hill, the second taking more of a southeasterly course from the southern end of Carbonate Hill to Long and Derry Hill. In the area between these two zones the surface is formed by porphyry bodies which overlie the ore horizon, so that no outcrops of vein material show on the map, except a few thin lines along the edges of fault planes. Under these porphyry bodies in Carbonate and Iron Hills a very large proportion of the area is already proved to be occupied by valuable ore bodies, and it may, therefore, be reasonably expected that similar bodies exist under the porphyry sheets between these zones farther east, although, owing to the greater depth of the ore horizon, they have not yet been reached by mine workings. There is evidence of still another zone of replacement extending north from Fryer Hill under Prospect Mountain, but, as the ore horizon has been reached in few points and the vein material has, at these few points, proved comparatively poor, the chances of finding any considerable development of rich ore in that direction have necessarily a smaller basis of probability than to the eastward, though the general geological conditions favor it.

In the description which follows, the mines will be grouped according to the main features of geological structure already outlined in Part I, Chapter V.

LITTLE STRAY HORSE SYNCLINE.

As shown in the previous chapter, the eastern portion of the area represented on the Fryer Hill map belongs structurally to the western rim of the Little Stray Horse basin, and the ore horizons of the Little Sliver, Forest City, and Denver City mines, if followed continuously eastward, would finally reach the bottom of the basin. The basin is bowl-shaped, its outlines being shown on the map by those of the Gray and White Porphyry bodies which fill its depression.

Southern rim.—Through its southern rim runs the zone of cross-cutting White Porphyry, in virtue of which the Blue Limestone along the southern and western rim is supposed to be split into two wedge-shaped sheets. Of these the lower one, which thickens to the south and constitutes the entire thickness of the horizon in the mines of Iron and Carbonate Hills, is buried under the whole overlying White Porphyry under the northern end of Graham Park, and its depth or condition of mineralization is not known. The points nearest to the axis of the basin at which it has been reached are in the Highland Mary (P–52), in Stray Horse gulch on the east, and in the Wolftone (T–5) on the west.

The upper portion, which wedges out to the south, outcrops under the Wash as indicated on the map, commencing to thin out near the southern edge of the area of the Fryer Hill map, and reaching its thinnest point at the Mahala (T–2). It is proved in the following shafts: The Moyamensing (S–12) strikes iron vein material, which probably forms the outcrop of the Denver City ore body, below the Wash at 115 feet. In the Robert Emmet mine, in Stray Horse gulch, this portion of the Blue Limestone is represented by 50 feet of manganiferous iron, with White Porphyry above and below, the overlying porphyry showing traces of original pyrites which have been dissolved out. The main shaft reached the contact at 110 feet, finding a dip of 30° N. E. and being sunk afterwards 150 feet in underlying porphyry. The ore thus far extracted has been taken between this shaft and the outcrop, which crosses Stray Horse gulch just below the tunnel (S–3). Farther south the Agassiz (T–3) finds the vein material and limestone 30 feet thick, at a depth of 45 feet, the Goneabroad (T–4) finds it at 80 feet, and the Cyclops (T–1), farther east, finds it at 148 feet below rock surface, with a thickness of 50 feet, and passes through it into underlying porphyry. In the Agassiz, as in the Robert Emmet, the vein material is manganiferous iron, with carbonate ore at its upper surface, sometimes five or six feet thick; about five feet of quartzite are found above the contact, as in Carbonate Hill. The dip is about 30° N. E. The Greenback shaft (O–53) found Lake beds, the northern continuation of the Graham Park area (see Atlas Sheet VI), beneath the Wash. The ore horizon consisted of 3 feet of iron and chert, 45 feet of limestone, and again 7 feet of iron. White Porphyry was penetrated 40 feet below this.

Southeastern rim.—On the eastern rim of the southern end of the basin the outcrops of the Blue Limestone are less continuously proved. The Indiana (P–53) finds limestone directly beneath the Wash, while the Highland Mary (P–52) reaches the lower body of Blue Limestone after passing through 122 feet of White Porphyry. The Rarus (P–61) passes through the edge of the Gray Porphyry sheet directly into limestone, showing that here the upper White Porphyry is wanting. It comes in again, however, in the Hunkidori (P–72) shaft, a little east, which penetrated it for 40 feet, after passing through 170 feet of Gray Porphyry and 5 feet of Weber Shales.

Western rim.—Along the western rim a number of shafts have been sunk in Gray Porphyry, in search of the continuation on the dip of the Lee and May Queen ore body, without having yet reached it, the influx of water making it difficult to sink their shafts. Some had penetrated the White Porphyry a short distance, and these had always found a portion of the Weber Shale, either as quartzite or as black carbonaceous shale impregnated with pyrites, at the contact of the two porphyries. All had found Wash from ninety to two hundred feet deep. In the Little Sliver (P–81) the White Porphyry was 41 feet thick; the Shamus O'Brien (P–73) had penetrated it 30 feet; the Tip Top (P–75), 38 feet; the Union Emma (P–79), 25 feet; and the Bangkok (P–77), 52 feet; while the Cora Bell (P–78), Forepaugh (P–76), Prince of Orleans (P–71), and Olive Branch (P–70) were still in the Gray

Porphyry and the Lickscumdidrix (P–68) bore-hole, in the middle of the basin, has gone through 400 feet without reaching its base.

Eastern rim.—East of the above shafts the El Paso (P–65) and Little Miami (P–58), at depths of 470 feet and 390 feet, respectively, were still in White Porphyry, after having passed through Gray Porphyry, and through varying thicknesses of Weber Shales both at the contact of the two porphyries and within the lower body. The Kennebec (P–55), Cullen (P–57), and Aztec (P–54) have reached the limestone after passing through the two porphyries, the former finding a second sheet of White Porphyry within the limestone. The same sheet is found in the Cordelia Edmondson (P–41), which is sunk in a large body of vein material, directly below the Wash. Several other shafts have struck the very considerable body of vein material which replaces the Blue Limestone on this rim of the basin, but as yet no important ore bodies have been found.

The most extensive workings are in the Chieftain and Scooper mines. The former is opened by a tunnel (P–43), which runs southeast through vein material, and then through limestone, compressed into gentle folds but apparently with a slight dip west, and, at a distance of 360 feet from the mouth, strikes granite which forms the foot-wall of the Iron fault. A decomposed porphyry, resembling White Porphyry, is found adjoining the fault. The limestone near the end of the tunnel is quite light colored; it may be the White Limestone, but the structure was not sufficiently shown to make this certain. The Scooper (P–44) shaft was sunk through 60 feet of Wash, 20 feet of Gray Porphyry, and 5 feet of White Porphyry, to iron and limestone. The contact with the porphyry is here very steep and runs in a direction a little east of south, being cut by several drifts; it was supposed to be the line of a fault. It is probably, however, simply an unusual steepening of the dip on this edge of the basin, as the Indiana (P–64) shaft, about 400 feet west of it, was still in Gray Porphyry at a depth of 330 feet. The limestone and vein material are crushed and folded even more than in the Chieftain, and probably from the same cause, viz, compression against the Iron fault. Considerable very silicious, hard carbonate ore, rich in chloride of silver, has been taken from this mine, but for some unexplained reason the developments have been very irregular and without system.

This basin has been described somewhat in detail to show how much fruitless labor has been expended in a region whose surroundings would lead one to suppose that it contains fine bodies of ore. Although so many shafts have been sunk to depths of 200 to 500 feet, the contact has been seldom reached, and even then but little explored. The main difficulty has been the great amount of water met in depth, which could not be controlled by the pumping apparatus in ordinary use. That such a basin should hold a large amount of water, especially when its outcrops are crossed by two such stream beds as those of Big Evans and Stray Horse gulches, is most natural, and it will probably be impracticable for any one mine to work in it alone. Work must be carried on by combination either of actual properties or else of working expenses, and powerful pumping apparatus must be established to drain the whole basin from its deepest point.

YANKEE HILL ANTICLINE.

On the western slope of Yankee Hill the J. B. Grant shaft found about eight feet of vein material between White Porphyry and White Limestone, which is only significant as showing that replacement action has gone on to a certain extent beneath the lower sheet of White Porphyry.

On the eastern side of Yankee Hill a large body of iron vein material has been found, extending from the Clara Dell and Little Champion northward through the Bevis and Boulder Nest mines, and thence eastward to the Andy Johnson, which reached it after passing through 200 feet of Gray and White Porphyries. This body consists of iron oxides and chert and is undoubtedly much more extensive than has been represented on the map; it contains some ore, but the data obtained with regard to it were too meager to do more than prove the same probability of the existence of valuable ore bodies in the synclinal basin to the eastward that exists in regard to Stray Horse Park. The Superior (K–61) and Mountain Boy (K–60) shafts struck a considerable body of limonite and chert on the southwestern edge of this basin, dipping at an angle of about 30° to the northeast. This body, which is some fifty feet thick, is supposed to be the replacement of a split-off portion of the Blue Limestone. This supposition accounts for the apparent want in the thickness of this horizon to the west on a line drawn through the Leavenworth shaft, as shown on the map and explained in Part I, Chapter V. On the other hand, it is the Weber Shales that are ordinarily found between the Gray and White Porphyry sheets, and in the Theresa (K–57) shaft, a short distance to the northeast, highly pyritiferous shales were found in this horizon at a depth of 325 feet.

Breece Iron mine.—In a similar position occurs the deposit of the Breece Iron mine, situated on a spur of Breece Hill, overlooking Adelaide Park. This remarkable deposit of iron ore is found at the surface in two distinct bodies, shown in open cuts, the one a short distance above the other. The lower body has a maximum thickness of 20 to 25 feet and rests on White Porphyry, with a mottled porphyry on the hanging wall. The upper body is not so thick and is overlaid by the main sheet of Gray Porphyry; both dip eastward, and the shaft (K–39) higher up the hill has been sunk through 30 feet of iron without reaching the bottom of the body, from which it may be supposed that the two bodies have here come together. The Gray Porphyry has either the characteristic large crystals of orthoclase or the cavities which they once filled. The intermediate porphyry is, however, of finer grain, of a pinkish color, and is full of minute cavities having the form of crystals of pyrite. This may possibly represent a tongue of Pyritiferous Porphyry extending between the two iron bodies. The lower (K–36) shaft has been sunk 350 feet in the underlying White Porphyry, which, near the iron body, is also impregnated with pyrite. The iron bodies are rather irregular in shape and send offshoots or stringers into the surrounding rocks. The ore is, however, massive and compact and remarkably free from earthy gangue. It has been largely used as a flux in the smelting works, being supposed to carry several ounces of silver to the ton, and has also been used by the Colorado Coal and Iron Company in the manufacture of Bessemer steel. It is a very pure hematite, with a certain admixture of magnetite which seems to occur mainly near the outcrop. It carries about 66 per cent. of metallic iron. The complete analysis of an average specimen, by Mr. Guyard, will be found in Appendix C. Undecomposed pyrites are found in the ore from the upper shaft. It seems probable that this body, like the iron bodies in the various silver mines, is the result of the oxidation of pyrites, which were concentrated at the junction of the three bodies of porphyry. It differs in being anhydrous, while all the others are hydrated; also, in the fact that it occurs high up on the hill and free from the ordinary covering of Wash, which, no doubt, promotes decomposition and the chemical combination of water in rocks that underlie it. From the staining of rocks eastward, along the outcrop of the contact, and from data obtained from shafts sunk higher up on the hill, it appears that an iron body extends, though not continuously, for some distance to the eastward

SYNCLINE EAST OF YANKEE HILL.

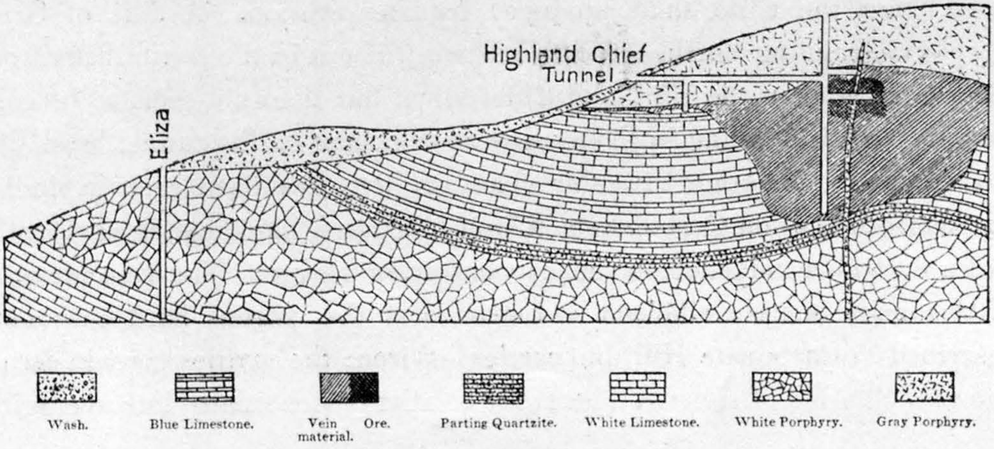
The geological structure of the area between Yankee Hill and Weston fault has been already explained in Part I, Chapter V, and is graphically shown in Sections C, D, and L. As yet no considerable ore bodies have been found at the ore horizon in this area; but it seems not to have received the attention it deserves, in view of the good indications afforded by the explorations already made. These may be briefly enumerated as follows:

The amount of vein material proved on its western rim has been mentioned above. On the south side, the Little Prince (K–32) shaft passed through 150 feet of Gray Porphyry and 80 feet of White Porphyry to the Blue Limestone horizon, which is here about one hundred and twenty feet thick and entirely replaced by a porous silicious material, not unlike the granular quartz gangue of the Morning Star mine. At first glance it somewhat resembles a decomposed porphyry, and in it small irregular bodies of sand carbonate, but no limestone, have been found. The Parting Quartzite was found below it. On the lower slopes of the hill the Nora (K–23), Bosco (K–28), Across the Ocean (K–31), and Great Hope (K–30) shafts were sunk through Gray Porphyry to the ore horizon, the first two without finding any intervening White Porphyry. In all more or less vein material was found replacing the limestone. The Onota, from which the typical Gray Porphyry was taken for analysis, also reached limestone near the middle of the basin. In the Great Hope quartzite was reached after passing through 60 feet of vein material. In the iron vein material some large layers of lime-sand were found, and at 105 feet from the surface a streak of galena five to six feet thick is said to have been passed through, but the main ore of the mine was taken from the quartzite, which, as well as the lower portion of the iron body, was impregnated with gold. The gold was very coarse. Some four hundred to five hundred tons of quartzite gold ore, averaging one and a half ounces of gold and four ounces of silver to the ton, are said to have been taken from the mine, and one lot of 3,000 pounds is said to have yielded thirty-one ounces of gold to the ton. A dike of White Porphyry, cut in the eastern portion of the workings, may have influenced the concentration of ore at this point. The quartzite is probably Parting Quartzite, though the thickness of 60 feet given for the Blue Limestone horizon seems small, especially as it is said to have been over ninety-six feet thick in the adjoining Across the Ocean shaft. The vein material on the dump of the latter contained a great many quartz-lined cavities, and it is possible that what was taken for quartzite in the Great Hope was simply granular silicious gangue like that in the Little Prince.

SOUTH EVANS ANTICLINE.

Highland Chief mine.—The vicinity of the South Evans anticline has evidently been a favorable locality for ore deposition, but the development of the ore bodies has been retarded by the difficulty of understanding the geological structure and the relative positions at which they occur. The Highland Chief mine has been the most important ore producer of this portion of the district. It is opened by a shaft on the brow of the hill overlooking South Evans gulch and by a tunnel run in to meet it part way down the slope. Fig. 2 represents an ideal section on a broken line drawn through the Highland Chief shaft (L–1), the tunnel (G–54), and the Eliza shaft (G–58).

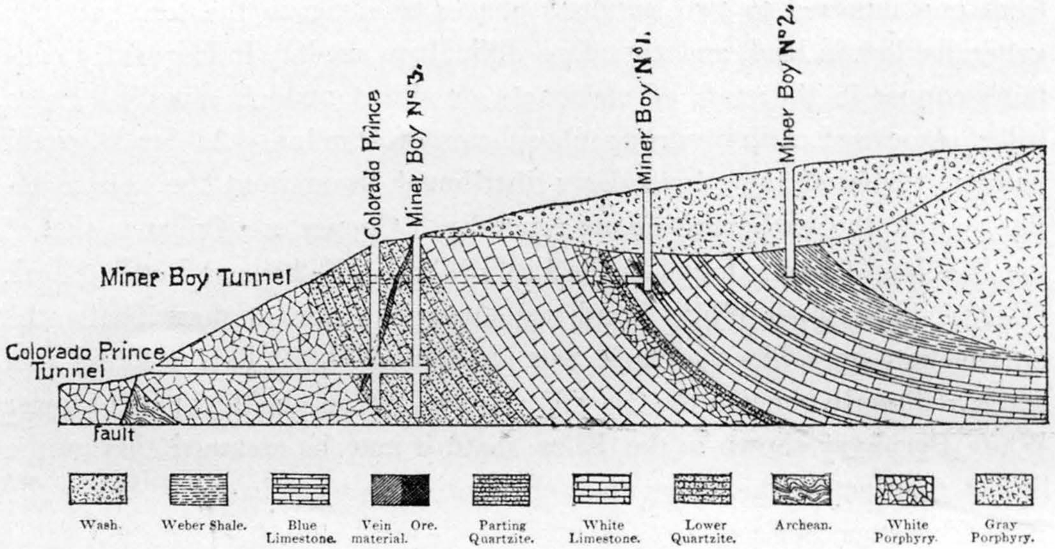
FIG. 2.



Both tunnel and shaft are driven through Gray Porphyry to the contact, the upper White Porphyry being here wanting. No limestone is struck in either, but a winze sunk a short distance from the mouth of the shaft found it, as did the Highland Mary (G–55) and Curran (G–56) shafts close by, which found some vein material, but no ore. Vein material rises rapidly in the floor of the tunnel as it approaches the shaft, and the latter is sunk in it 170 feet below the tunnel level. The lower surface of the porphyry sheet has a dip to the northeast, which is also seen in the eastern portion of the workings beyond the line of the section, while in the southwest workings there is a dip westward. In the Chemung (K–5) tunnel, which is driven 400 feet, along the contact of Gray and White Porphyries and then in Blue Limestone, in a southeasterly direction, from a point just below the road about seven hundred feet southwest of the mouth of the Highland Chief tunnel, the limestone at the end is found to be dipping westward. It thus appears that there is a slight ridge or lateral fold in the formation, which is shown on the map by the curve in the strike of the formations. The section, whose line is taken along the crest of this fold, shows at right angles to it a more pronounced folding, forming an anticlinal and synclinal structure parallel to the South Evans anticline, which is necessitated by the intersections of formation lines obtained in the Highland Chief and Eliza (G–58) shafts. The intermediate outcrops are obscured by the moraine material or Wash, left on the shoulder of the hill by the South Evans glacier. The section shows the rounded outlines of rock surface left under the Wash by this glacier and the abrupt slope below the Eliza shaft produced by later erosion. A dike of Gray Porphyry, running northeast and southwest, is cut in the south drifts from the second level of the Highland Chief mine, but does not seem to be continuous, as it is wanting in the southwest workings on the tunnel level. It has evidently the same irregular character that has been seen in similar transverse bodies on Carbonate and Iron Hills, and, like them, has evidently had a favorable influence on the concentration of ore. The vein material of this mine is very silicious and resembles the porous hard carbonate described in Carbonate Hill, but carries less iron; the cavities contain cerussite and chloride of silver. The ore is relatively rich, many lots averaging from one hundred to two hundred ounces of silver to the ton, but it is generally low in lead, and therefore difficult to smelt. It frequently contains copper in the state of carbonate or silicate, which gives it a green color, associated with hydrous phosphates of alumina. As far as could be seen it is much more irregularly distributed throughout the vein material than in the carbonate mines generally. Copper ore similar to that of the Highland Chief is found in the Little Johnny (F–29) and Rattling Jack (F–28). The upper White Porphyry comes in again at these shafts, the intermediate ones, like the Uncle Sam (F–32 and F–33), having found vein material directly beneath the Gray Porphyry. From the thickness of lower White Porphyry shown in the Eliza shaft, it may be assumed that on the line of this section the entire mass of White Porphyry has gone below the Blue Limestone horizon.

Colorado Prince group.—The lower formations, which form a cliff face overlooking South Evans gulch between the Highland Chief mine and the Colorado Prince fault, have been exposed by numerous mine workings. These have been exploited in such an irregular and intermittent manner that it has been impossible to obtain satisfactory data as to the amount or quality of ore extracted from them. The geological structure shown by the various shafts and tunnels that were examined is, however, interesting, showing the rising of the beds over the South Evans anticline and the slight displacement caused by the Colorado Prince fault. Just above the Colorado Prince tunnel there is an actual rock outcrop of White Porphyry and overlying Lower Quartzite, forming a steep cliff; but, on the shoulder above, the rock surface is deeply buried under the Wash, a relic of the lateral moraine of the South Evans glacier. The following diagram shows an ideal section, drawn through the Colorado Prince tunnel (G–43) and shaft (G–47), and the shafts of the Miner Boy, No. 3 (G–48), No. 1 (G–50), and No. 2 (G–51). The Miner Boy or Kentucky tunnel (G–42), which is a little beyond the line of the section, is indicated in dotted lines.

FIG. 3.



The most important mine of this group is the Colorado Prince, which has its own stamp mill for crushing and amalgamating its ore. According to the reports of experts who have examined it, it has a very large body of rich ore, but the practical results of work do not thus far seem to have justified their prognostications. The management of the mill has been so frequently changed that it was impossible to learn the actual working results of treatment over a long enough period to determine whether the apparent want of success was due to the quality of the ore itself or to faulty methods of reduction.

The deposit of the Colorado Prince and Miner Boy mines is somewhat in the nature of a gash vein in the Lower Quartzite. It stands at an average pitch of 75° to the east, varying between 60° and the vertical. Its strike is about N. 15° W. It varies in width from a few inches to five or six feet, and in the upper workings is said to have been stoped out on a width of 20 feet. It has no distinct walls or clay selvages, and the matrix of the ore is principally decomposed quartzite, more or less stained by iron oxides. Near the Miner Boy No. 3 shaft it splits into two branches, one following its general direction, the other being more nearly north and south. Following the vein is a thickness of one to three feet of light-colored decomposed rock, called by the miners trachyte and considered by one expert to be a propylite dike. It is, however, only a fine grained conglomerate of rounded pebbles of quartz in a clay matrix. So far as explored, the vein is confined to the horizon of the Lower Quartzite, the upper 20 feet of which are calcareous. It may have extended up into the White Limestone, but it has not yet been followed into the White Porphyry below and shows a tendency to pinch in that direction. The ore is essentially a free-gold ore and is generally stained by oxide of iron and carbonate of copper. Galena occurs sparingly, having been only observed in a few spots. Carbonate of lead is said to have been found. The first and second class ores are generally sent to the smelting works at Argo, only the third class going to the stamp mill.

The vein seems to be the result of the action of percolating waters along a fracture plane in the formation, which was very probably formed at the time of the displacement of the Colorado Prince fault, with which it has a general parallelism. The ore is probably the result of the oxidation of pyrites, and whether it was originally concentrated in this form in its present position or brought in as a secondary deposition from the surrounding rocks it is difficult to say definitely.

The general dip of the formations in these workings is steeply to the south, sometimes varying a little to the west and again to the eastward. Its observed angle also varies from 35° to 60° in the Colorado Prince workings, the steeper dip occurring near the Miner Boy shaft. The intersections obtained in the Colorado Prince tunnel are shown on the section. The Kentucky or Miner Boy tunnel was run for the following distances through the successive formations: Lower Quartzite, 200 feet; White Limestone, 200 feet; White Porphyry, 40 feet; Parting Quartzite, 35 feet; Blue Limestone, 50 feet to the bottom of shaft No. 1. These figures, it must be remembered, are not actual thicknesses of the formations. The lower part of the Blue Limestone as exposed in Miner Boy No. 1 shaft is thoroughly impregnated with oxide of iron and is said to have yielded some very good assays. The No. 2 shaft of the Miner Boy found black shales of the Weber Shale formation directly beneath the Wash.

LITTLE ELLEN HILL.

On Little Ellen Hill the Blue Limestone has been found to be replaced to a considerable extent and some argentiferous lead ore has been obtained from it; but present explorations cover only a small proportion of its area. The principal mines are the Virginius and Little Ellen. The Virginius (G–24) is opened by a tunnel run southwards along the strike of the formation from the north side of the hill, facing Big Evans gulch. The limestone is largely disintegrated and in the condition of lime-sand. The vein material, as usual an impure iron oxide, is very silicious. Galena is thickly scattered through it, but the ore is of rather low grade in silver. The tunnel is 250 feet long, and drifts from it have been driven eastward 150 feet on the dip. On the hill above, in the Cleveland (G–27), 15 feet of vein material, carrying galena, are found above the limestone, and at a depth of 30 feet in the limestone a cross-cutting body of Gray Porphyry, which is also cut in the Last Chance (G–31) shaft, immediately below the Wash. The other shaft (G–30) of the Last Chance finds Parting Quartzite below the Wash. The Australian (G–28) and Tenderfoot (G–26) also find vein material at the contact.

The Little Ellen mine, higher up and on the slope of the hill facing South Evans gulch, finds a very large body of low-grade lead ore at the same horizon, where the strike has changed to east and west. This contact is traced eastward through the Lulu, Gnome, and Alps workings, showing considerable replacement action, but as yet no large ore bodies.

BREECE HILL.

In the large area lying between the regions above described and Iowa gulch there is an immense development of igneous rocks, and, on the theories deduced from the studies made in this region that there is a direct connection between igneous action and ore deposition, or rather that the latter is more abundant where the former has been most active, this area should contain large deposits of ore. Unfortunately there is little else than theory upon which to base this assumption. The ore horizon throughout the greater part of the area is so deeply buried beneath the surface that, in the uncertainty that exists as to how deep it may be necessary to sink a shaft in order to reach it, no mine owner has yet had the enterprise to attempt its exploration. It is difficult to predicate the probable thickness of a sheet of porphyry from the width of its outcrops or even from an observed thickness at some point but little removed, since it is liable to change rapidly, both in thickness and in horizon, from causes which are not apparent at the surface. The depths given for the Blue Limestone horizon under the Pyritiferous Porphyry by the various sections which pass through this area cannot be expected to be so close an approximation to the actual facts as when the proportion of sedimentary beds is greater. They have, however, by no means been given at hap-hazard, but are the result of most careful weighing of probabilities, based on a study of the whole region.

The nearest approach to an actual development of ore in the Blue Limestone under Pyritiferous Porphyry is in the Mike mine, and the indications afforded by this favor the supposition that it will be found to be ore-bearing when reached farther east. Another favorable sign is found in the numerous evidences of mineral concentration in the porphyry itself. Scattered bodies of ore have also been found in the beds of the Weber horizon, which are sufficient to show considerable mineral action.

GREEN MOUNTAIN.

In the shales and sandstones of the Weber formation and in the adjoining Pyritiferous Porphyry bodies numerous small deposits of ore have been struck, but of somewhat different character from those developed at the Blue Limestone horizon. The most important of these have been found on Green Mountain, in regard to which the following information has been obtained.

The Ontario mine is opened by a tunnel 380 feet long and a shaft 70 feet deep, run through Pyritiferous Porphyry into micaceous sandstone and black shale. The deposit is found in a vein running north and south, and traceable from about one hundred feet from the mouth of the tunnel through the porphyry into the overlying sandstones. A body of coarse-grained galena 20 feet long by 3 feet in width was found in the porphyry, but in the sandstone the vein proved barren. A similar deposit in the porphyry is found in the adjoining Tiger (E–22) mine. It is a gash vein six to eight inches wide, running northeast, with a dip of 60° to the southeast. The ore is a mixture of galena and pyrites, one lot of which yielded 22 ounces of silver and 22 per cent. of lead to the ton.

In the Green Mountain (E–12) mine the ore occurs in the sandstone, and is a free-gold ore, sometimes very coarse. It was not accessible at time of visit, and nothing certain could be learned of the form of the deposit, which is very probably also a gash vein.

LONG AND DERRY HILL.

Ready Cash mine.—On the steep south wall of Iowa gulch formed by Upper Long and Derry Hill is the Ready Cash mine, which is interesting on account of its occurrence in Archean rocks rather than from its importance as an ore-producer. Its country rock is a coarse-grained, reddish granite similar to that found at the heads of Iowa and Empire gulches. The ore occurs in two small veins from one foot to four feet in width, which come together at the surface. The one strikes N. 25° E. and dips 50° to the east; the other strikes N. 45° E. and has a still steeper dip. The mine is opened by tunnels which run into the hill in a southerly direction and cut the veins at about three hundred and twenty feet from the surface; in one of the upper tunnels a dike of quartz-porphyry is cut between the two veins, in which are some gold-bearing seams. The vein filling consists principally of altered granite, carrying free gold and some chloride of silver, with relatively well defined walls. The deposit is evidently formed along fracture planes in the granite, filling any cavity that may have existed and replacing the constituents of the adjoining country rock.

Long and Derry mines.—The Long and Derry mines were among the first discovered in the Leadville region, and are still held by their original owners, a very exceptional circumstance. They have been worked very slowly and irregularly, and at the time of examination no maps had been made of the underground workings. Many of the drifts had become inaccessible, so that only a very general idea of the form of the deposits could be obtained. They occur as a replacement of the Blue Limestone, but, as contrasted with the deposits of Iron and Carbonate Hills, rather in the form of large chambers than along the contact with the overlying White Porphyry.

These chambers are irregularly distributed through the upper part of the Blue Limestone, and, as the latter has been more or less eroded, some of them formed actual outcrops of ore on the surface, and were thus easily discovered in the commencement. The mine is opened by both tunnels and shafts, but the ore is mainly extracted through the former. The Faint Hope (M–4) tunnel runs 180 feet through limestone and vein material and 360 feet in White Porphyry, to the bottom of the Porphyry (E–37) shaft. The latter passed through 98 feet of porphyry and was sunk 46 feet in vein material containing nodules of chert and low-grade ore. The Long and Derry (E–32) tunnel runs in a southerly direction from the hillside overlooking Iowa gulch, starting in White Porphyry and being expected to reach the contact at 500 feet. In the Dana (M–3) shaft were found many chert nodules in the porphyry, which carried casts of fossils, principally *Pleurophorus* and *Spirifera.* Such included chert nodules are frequently found not far from the contact plane and were evidently caught up in the mass of the porphyry as it forced its way along between the strata. The whole body of the limestone on Long and Derry Hill seems to be more or less impregnated with oxides of iron and manganese and with silicious material, as shown by the black outcrop at the Belcher (M–5) mine, near the base of the formation. In this outcrop the replacement evidently proceeded from above, as the ore wedges out at the bottom. The apparently broken character of the formation in the vicinity of the ore deposits may be due in a measure to its proximity to the surface and want of a protecting covering of Wash. In one portion of the mine was a large accumulation of angular blocks of porphyry in the limestone, which had probably fallen from above into a cave that had been dissolved out by surface waters.

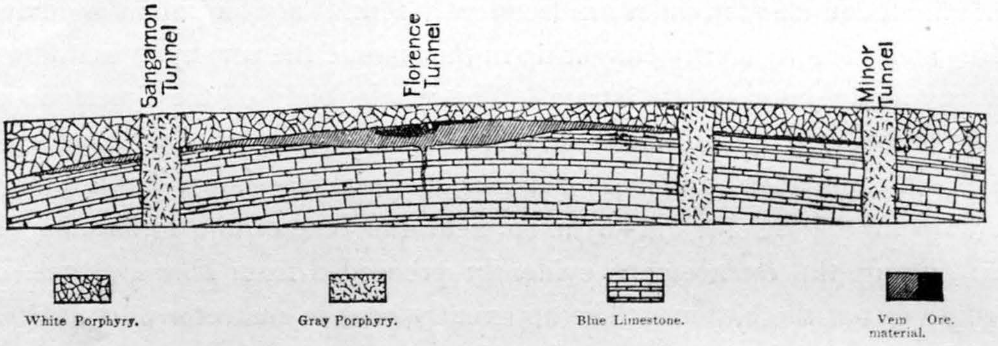
The ore is carbonate of lead and chloride of silver near the surface and galena in depth, the latter, as usual, being generally richer in silver than the carbonate. Only small specks of pyrite have been found in it. The dike of Gray Porphyry, which crosses the hill in front of the Faint Hope tunnel, may be supposed to have been an important factor in the concentration of ore here by causing a stagnation in the mineral-bearing currents.

PRINTER BOY HILL.

The existence of the three transverse dikes crossing the slopes of Printer Boy Hill and of the cross-cutting bodies of Pyritiferous Porphyry on its north side, in the upper part of California gulch, is an indication of probable mineral concentration in this region, and developments, so far as they have gone, tend to confirm this idea.

On the south slope the curving outcrop of the contact of Blue Limestone and White Porphyry has been very extensively prospected near the surface, and evidence of replacement action has been found at almost every point. Figure 4 shows an ideal section of the middle portion of the south slope and the relative amount of vein material, as far as can be determined by present workings.

FIG. 4.



The principal development of ore has been thus far found in the Florence mine, which is at the crest of the fold. The vein material, which is the usual clayey matter, carrying oxides of iron and manganese, has here a maximum thickness of thirty to thirty-five feet, and seems to thin out to the east and west, respectively, but to the north or into the hill it has as yet been explored but a short distance. The ore occurs at the surface and rather irregularly through this vein material; besides the usual carbonate of lead and chloride of silver, it contains several minerals not common in the district, among which may be mentioned native gold, visible to the naked eye, and a sulfo-carbonate of bismuth. (See Appendix C.) In a tunnel directly below the main workings of the Florence a gash vein of galena in limestone several feet in thickness was found. Above, it connects with the ore body of the Florence, but, as followed in depth by a winze, was found to pinch out entirely, furnishing one of many proofs that the ore entered the limestone from above and not from below.

The dikes of Gray Porphyry on either side of the Florence are thirty to fifty feet wide and can be traced along the hill for considerable distances. They have the characteristic large crystals of orthoclase porphyritically distributed through the mass.

East of the Florence the Minor tunnel has been run several hundred feet on the contact without finding ore, but good ore is said to have been obtained from the First National mine, still farther east. In the latter the limestone seems to be split into two parts by the cross-cutting zone of White Porphyry.

To the westward the contact has been developed by the Sangamon tunnel, the Wilson, Brian Boru, G. M. Favorite, and other mines, and some ore has been shipped, but no certain information was obtained in regard to its quantity or its quality.

Although some prospects showing ore may have escaped observation, the above descriptions suffice to show that on either side of Iowa gulch in the vicinity of the three dikes of Gray Porphyry the Blue Limestone contact is ore-bearing over a comparatively large proportion of its extent. It can, however, hardly be said to have been thoroughly explored as yet, nor are the ore developments as rich or extensive as the geological conditions would lead one to expect. It may, therefore, reasonably be assumed that future explorations in this area, if systematically conducted, will prove remunerative. The most promising direction for exploration would seem to be to the northward under Printer Boy Hill. Next to that, the easterly continuation of the contact in depth beyond the First National affords a promising field. It must be borne in mind, however, that in this direction the zone of cross-cutting White Porphyry will soon be reached, where the Blue Limestone will be found split into two wedge-shaped masses, and thus there will be two contacts between limestone and porphyry.

IOWA GULCH.

In Iowa gulch, to the westward, the contact is cut off by the Mike fault, after crossing the gulch just beyond the G. M. Favorite. A little farther down the gulch, however, opposite the Rock and Dome workings, it comes to the surface again and is explored in the Nisi Prius and adjoining claims. In the Nisi Prius very large bodies of highly manganiferous vein material are found, which, though not rich themselves, are generally considered to indicate the proximity of a concentration of chloride of silver The tunnel on this claim cut through a series of limestone ridges alternating with depressions filled by vein material. That the ridges in this case do not represent actual folds in the formation is proved by the fact that the stratification lines can be seen to run horizontally across them.

This is the westernmost point at which the contact has been reached in Iowa gulch, and, while it may be found at no very great depth immediately west of Dome fault, the probabilities are that beyond that, as shown by the cross-sections N, O, and P, it is too deeply buried beneath Wash and Lake beds to render its exploration advisable unless it were followed down continuously on its western dip. The thickness of superincumbent detrital material given in these sections is deduced from that found in the Coon Valley and Black Cat and is probably a pretty close approximation to the actual facts. The limits of the ore horizon to the westward, as indicated by the western outcrops of the Blue Limestone, are theoretical deductions, and may vary somewhat from the facts; but the limit of error here is not a very wide one, as actual outcrops of Archean on this strike-line are found at a very short distance south of the boundary of the map.

HEAD OF CALIFORNIA GULCH.

On the north side of Printer Boy Hill and in the upper part of California gulch the geological conditions are extremely complicated, owing to the number and variety of porphyry bodies and their rather unusual structural relations. The cross-cutting zone of White Porphyry passes through here, as shown by the fact that on the north side of the gulch it comes in contact with the Parting Quartzite and on the south with the Blue Limestone. A body of Pyritiferous Porphyry is also found cutting up across the formation, being in the White Limestone at the bottom of the gulch, and to the northward apparently joining the main body, which overlies both White Porphyry and Blue Limestone. In addition to these are several bodies of Gray Porphyry, whose position and structural relations are very imperfectly known. For these reasons a considerable concentration of ore might naturally be looked for in this region, but it would probably be found more irregularly distributed and less easy to follow than in the normal groups of Carbonate and Iron Hill. The ore currents would no longer have been confined to one or two principal channels, but would have had a variety of contact planes which they might have followed, and whose position cannot be predicated beforehand. Unfortunately the small amount of exploration already made not only gives but an imperfect idea of the actual geological relations of the various bodies, whose representation on the map and sections may in consequence be somewhat imperfect, but the actual development of ore has been so slight as to afford no indication as to what particular contact is the more likely to contain ore. The normal contact of Blue Limestone and White Porphyry extends eastward from the Pilot fault along the north slope of Printer Boy Hill, until in the Eclipse the limestone commences to wedge out Ore is found in considerable quantity in the Pilot tunnel, apparently along the very plane of the fault, and a certain amount of vein material occurs at the contact in the Lovejoy, Eclipse, and other shafts. On the other side of the gulch the same contact that carries ore in the Adelaide-Argentine, viz, Parting Quartzite and White Porphyry, is found in the Iron Duke and others and is apparently mineralized to a certain extent.

GOLD DEPOSITS.

Between Pilot and Mike faults, on the northwest slope of Printer Boy Hill, is a body of porphyry not identical with any other found in the region and which is notable for containing the Printer Boy and Five-Twenty gold-bearing lodes.

The Printer Boy lode was discovered before the existence of carbonate ores in this region was known, and produced a large amount of gold between 1866 and 1870, of which no record can be obtained. It is a practically vertical deposit along a jointing or fracture plane in the porphyry, having a direction a little east of north. It now consists of two claims, the Upper Printer Boy and Lower Printer Boy, each opened by vertical shafts.

Only the latter was accessible at time of visit, and here the vein is double, the two branches being separated by ten or twelve feet of decomposed porphyry. The gangue is nothing more than thoroughly decomposed porphyry, being a white, clayey material in which only the quartz grains of the original porphyry remain unaltered; scarcely any metallic contents are visible. In the old workings, where the ore was rich, free gold could be seen, and in the deeper workings considerable iron and copper pyrites and some galena and tennantite were found. Gold occurred in both pyrite and galena, and a piece of ore containing galena crystals, connected by a filament of wire gold, was one of the show specimens of the mine. Selected specimens are said to have contained 122 ounces of gold to the ton, and the average assay is given at three to four ounces. The vein has varied in thickness from one inch to four feet, with an assumed average of seven inches. On the west wall, from the surface down to a depth of 200 feet, were branches or stringers extending out into the porphyry, sometimes as much as three feet thick and containing ore of similar quality to the vein, though of different color and hardness. South drifts from the shaft of the Upper Printer Boy were said to be cut off, successively, at a distance of a few hundred feet, by a cement deposit, which, from the description given by those working the mine at that time, would seem to be a portion of Lake beds deposited in the bay that, as already shown, existed where now is the spur separating California from Iowa gulch.

In the Upper Printer Boy the normal Printer Boy Porphyry, which is a coarse-grained, greenish-gray rock with large feldspar crystals, was overlaid by a white, fine-grained porphyry, resembling decomposed Pyritiferous Porphyry from which the pyrite has been dissolved out. It is said that the former was only found at one hundred to two hundred feet below the surface, but this does not agree with observations which show the Printer Boy Porphyry in all the prospect holes around the Upper Printer Boy shaft. Both rocks occur in the Gray Eagle tunnel (M–53), west of Eureka gulch, and it is probable that part of the main body of Pyritiferous Porphyry originally covered this portion of the hill and that from the pyrite in it the waters derived the metallic contents which now are deposited in the veins. A number of other small gold-bearing veins are also found in the Printer Boy porphyry, the most important of which is the Five-Twenty, from which a considerable amount of gold ore is said to have been taken. The tunnel by which it was opened also cut a body of carbonate ore, which at that time was not considered worth extracting.

These deposits, together with that of the Colorado Prince mine, would seem at first glance to be of different character, and perhaps of different origin and manner of formation, from the normal carbonate deposit of the Leadville region, and more in the nature of the ordinary fissure vein. The somewhat limited study which it has been possible to make of them tends to show, however, that the process of ore deposition was essentially the same in both cases. There is no evidence of any pre-existing cavity which was filled by foreign gangue material; but the ore currents, following in this case a fracture or jointing plane, instead of a stratification plane, deposited gradually their load of metallic sulphides by a chemical interchange with the material of the rock through which they passed. Gold seems to have been principally deposited in the more acidic or silicious rocks and silver in the calcareous or basic ones. This preference is not confined to the Leadville region, but is seen in other mining districts. Its reason, however, is not yet satisfactorily explained from a chemical standpoint.

As regards their relative age, the little evidence that can be obtained seems to indicate that the carbonate deposits are older than the gold veins under consideration, since the former must have been originally deposited before the folding and faulting took place; whereas, if the assumption that the latter occupy fault planes be correct, it is probable that they were formed subsequently to this period.

PLACER DEPOSITS.

It is worthy of note that California gulch has furnished almost all the placer gold which made this region known long before its wealth in silver was even suspected, while Iowa and Evans gulches, adjoining it on either side, which are carved out of the same series of rocks and are both larger and contain more detrital material, have thus far yielded little or no return to the placer miner. The question naturally arises, therefore, why should the smaller gulch contain, as it did, exceptionally rich gravels and its neighbors be practically barren? Although placer mining had been virtually abandoned before this investigation was undertaken and these deposits no longer constituted an essential part of the mining industry of Leadville, nor could detailed studies be made of them, yet some suggestions are afforded by the consideration of the general geological conditions of the region that are not without value.

The richest portions of the California gulch diggings are said to have been, first, in the bend below Oro, at the mouth of Nugget gulch; next, in the bend at the La Plata mine; and then in that below Graham gulch. The first was exceptionally rich, and in the narrow bed of the gulch at Oro a gold-bearing cement, containing hydrated oxide of iron, was found below the gravel, which yielded one ounce of gold to the ton. The gulch gold was worth $17 to $19 per ounce, while that from the mines was worth only $15. From its propinquity the Printer Boy Porphyry, known to contain actual gold veins, suggests itself as the source of these rich gravels, with the oxide of iron resulting from the decomposition of pyrite in the Pyritiferous Porphyry as a cementing material. Moreover, the Weber sandstones, at the head of the gulch, have been found to carry gold veins, and from their abrasion also gold-bearing gravels would have been carried down the gulch. These probable sources of gold seem, however, inadequate to account for the greater relative richness of California gulch over its neighbors, since both Evans and Iowa gulches contain gold-bearing veins and the amount of sandstone and porphyry débris that has been carried away through either of their beds must have been far greater than that swept through California gulch.

It seems very doubtful whether in general all, or even the greater part, of the gold contained in placer gravels is derived from the abrasion of actual gold veins. Traces of gold may be found in a very large proportion of the massive rocks which form the earth's crust. Gold veins are concentrations of this mineral in sufficient quantity to attract attention and yield a profit to the labor of man; but doubtless there are a vast amount of smaller concentrations which may escape his notice. As the rock disintegrates and is worn away by atmospheric agencies the gold from these smaller deposits, as well as from the larger, is set free from its inclosing rock and subjected to the concentrating action of mountain streams. Placer deposits are the results of nature's vast sluicing processes. To bring them into the condition in which they may be made available by man requires not only the gold-bearing rock, which her agencies may grind up into sand and gravel, but the sifting power of rapid streams, which may carry down the lighter and coarser material, and a suitable channel, in which the heavier particles may lodge, as in the riffles of a sluice-box. All mountain gravels, all sands of rivers coming from the mountains, contain a certain amount of gold, but it is only under peculiarly favorable conditions that the gold is so concentrated as to render the gravel or sand remunerative to the labor of man. Among the most favorable of these conditions is a comparatively narrow channel, having a hard and compact bed-rock and ridges or bends in its course, which, by causing a partial arrest in the rapidity of the current, shall allow the heavier particles of gold to settle to the bottom and hold them there when once they have settled.

From this point of view there is a very evident reason why California gulch should have furnished rich placers and why the gold which may exist in Iowa and Evans gulches should not yet have been extracted, even though the detrital material which has been carried down either gulch should originally have been equally rich in gold. California gulch, as has already been explained, is a valley of erosion, formed entirely by the action of running water and since the Glacial period. It has, therefore, a bottom or bed of hard rock; its transverse section is V-shaped and therefore, like the Spitzkasten, favorable for the concentration of heavy particles at its bottom. When comparatively full of water, its numerous bends formed eddies in the down-flowing currents and allowed a longer time at these points for the settling of the suspended particles; and, as it cuts across many different formations in its course, its bed must have transverse ridges, which have caught some of the gold and prevented it from being carried farther down the stream. Evans and Iowa gulches, on the other hand are glacier-carved valleys; their courses are straight, their bottoms broad and comparatively smooth. The moraine material with which they are largely filled has not been subjected to the sifting or jigging process, to which gravel is subjected in the bed of a stream. Moreover, as shown in Sections M, N, and O, the lower part of their present bed is cut, not out of rock, but out of the loose gravelly formation of the Lake beds. It is not probable that this later bed, along which the material brought down by post-Glacial erosion has been carried, has a sufficiently hard and permanent bed-rock to allow of the concentration of gold on its surface; at any rate, such bed-rock has not yet been found. It is probable that the actual rock surface, which formed the bed of the Glacial valley, may be found gold-bearing if it is ever explored; but it could hardly be expected that in it gold will be found so concentrated as it was in California gulch, and its exploration will necessarily be attended with many mechanical difficulties, owing to its probable depth below the present surface.

The Lake bed deposits themselves undoubtedly contain a large amount of free gold, and it is probable that this may be sufficiently concentrated at some points on the bed-rock of the original lake to be worked with profit; but the form of this bed is as yet too little known to give any definite idea of where such points may be. Probably the valley of the east fork of the Arkansas would be the most favorable point for such an exploration, since its drainage area is the most extensive of any of the valleys tributary to this ancient lake; but it would be impossible to determine beforehand how far up this valley the ancient shore-line of the lake extended.

ORE PROSPECTS IN UNEXPLORED AREAS.

Besides the regions mentioned in the above description there are a few other unexplored areas in which the contact may be reasonably expected to be found productive which will be enumerated below in the order of theirrelative availability, or of the proximity to the surface or to known deposits at which the contact may be found.

On Dome Hill, as shown in Section H, Atlas Sheet XIX, the contact is near the surface on either side of Dome fault, and is practically unexplored south of the Dome workings on the one side and of the Ben Burb on the other. From the Dome mine southward to Iowa gulch, east of the fault and along the west side of the fault for a corresponding distance, there is good reason to assume that the contact will be found productive, and explorations should be conducted on the dip in either direction, though to the westward the assumed limit of probable mineralization will soon be reached, and on the lower part of Dome Ridge, beyond the Coon Valley, present indications do not promise any large ore bodies.

On Iron Hill the region west of Iron fault probably contains the continuation of the Iron mine bonanzas, but the depth at which it will be found may be such as to render its exploitation expensive; a shaft sunk to reach the contact opposite the Iron mine might have to go nearly a thousand feet. As there is some uncertainty as to the actual position of the bonanzas in this region, it might be more prudent to follow the contact downward from some point where its depth is known, say in the neighborhood of the Devlin shaft, where it is only 200 feet deep.

On Carbonate Hill the contact should be followed westward from the line of the Pendery fault. It has already been shown in Part I, Chapter V, that there is a good probability that on the northwestern slopes of the hill the formations dip westward, and that farther south in the Pendery ground, though broken by a fault, the displacement of the latter is probably not great. Were the contact followed westward from a shaft reaching the limestone anywhere along the lower slopes of the hill, it would not take long to determine whether it still exists under Leadville or had been removed by erosion previous to the deposition of the Lake beds. Should the latter prove to be the case it would avoid the great and, in this contingency, useless expense of sinking a deep shaft under Leadville itself, which otherwise will doubtless some day be undertaken.

MINES AND PROSPECTS OUTSIDE THE LEADVILLE DISTRICT.

The region outside the area covered by the map of Leadville and vicinity was examined primarily with the object of determining its general geological structure and studying the series of formations where they were less metamorphosed and better exposed than in that district, in order that by the experience thus gained it might be easier to unravel the complicated geological problem there presented. In the few short months that could be devoted to this work it was impossible to make a very complete or systematic study of the ore deposits opened by the various mines in this outside region, and only those that came within the line of work and were accessible at time of visit were examined. The following notes, the result of that examination, are offered, notwithstanding their incompleteness, mainly because of the information they afford with regard to the geological distribution of ore deposits in the area as a whole, which is sufficiently accurate, although the details of ore occurrence, especially in some of the larger mines, could not be ascertained, owing to their inaccessibility. Those mines whose names are followed by an (L) were not seen by the writer, but the information given has been obtained from the notes of Prof. Arthur Lakes, who assisted in the examination. The same general order of topographical description that was followed in Part I, Chapter IV, is preserved here.

NORTHEASTERN REGION.

Monte Cristo mine.—This mine is situated just beyond the northern limits of the map, on the steeper slope of the spur running eastward from Quandary Peak, just below timber line. Its ore is a low-grade argentiferous galena, occurring in a bed of quartzite of Cambrian age, with little or no accompaniment of vein material or of other minerals. The formation here dips eastward at an angle of 35°, which is approximately the slope of the steeper eastern flank of the spur. The quartzite stratum which carries the ore is the outcropping rock on this steeper part, so that it has only been necessary to strip off the thin surface accumulations of soil and débris to reach the mineral, which is thus exposed over an area of several acres. The galena is rather coarse-grained and crystalline, and is irregularly intergrown in the quartzite, occurring on its upper surface from a few inches to a foot in thickness, but in one case extending eight feet below the surface of the bed. It is evident at a glance that it could not have been deposited contemporaneously with the quartzite, and no evidence was seen of any pre-existing cavities; whence it is assumed that its deposition was a metasomatic change by percolating water, like the limestone deposits of the region. Although the ore, from its manner of occurrence, can be mined very cheaply, its tenor in silver is so low and it is so intimately mixed with the quartzite that it probably cannot be smelted profitably without previous concentration in ore-dressing works.

On North Peak ridge the North Star mine (L) is situated at the eastern end of its higher part, a short distance above and west of the saddle marked by the Blue Limestone outcrop. Evidence afforded by the dump shows that its shaft went through the Lower Quartzite cap of the ridge into the Archean gneiss below. At the foot of the steep slope below this mine and just west of the saddle are several prospect holes in a bed of dark-green hornblendic rock, near the upper part of the formation, which is highly impregnated with pyrites. In the amphitheater north of the saddle copper pyrites are found in the Archean gneiss, in a gangue of quartz. On the saddle itself a shallow prospect hole called Sammy's Barrel shows a gash vein in the White Limestone a few inches in thickness, which carries galena in a gangue of calc spar. At the eastern end of Hoosier Ridge, at the extreme head of Beaver Creek, is the outcrop of a large deposit of iron ore in the upper Carboniferous or Triassic beds, which apparently follows the bedding, and from which specimens of chrome iron are said to have been obtained.

MOUNT LINCOLN.

The silver deposits on Mount Lincoln were first discovered in the summer of 1871. In spite of their great altitude, being nearly fourteen thousand feet above sea level, they were rapidly opened; a number of mining towns sprang up at the foot of the mountain; quartz-mills were built and smelting works erected. For a time they enjoyed great prosperity, but of late years have been in great measure abandoned, and the mining towns of Quartzville and Montgomery are now practically deserted. The deposits are principally in limestone. The cause of their abandonment may be found in part in the inherent difficulty of regular development of limestone deposits, owing to their frequent want of continuity and to the misconception on the part of the miners of the character of the deposits. In great part it is probably also due to the excitement attendant on the discovery of the rich deposits of Leadville, which drew away the fickle miners to new fields. Everything tends to show, however, that the region is one exceptionally rich in metallic deposits, and that under systematic development its prosperity may be revived at no very distant future.

The Russia mine (L) is situated about five hundred feet below the summit of the peak, in a direction south-southeast. The deposits of this mine are found in the Blue Limestone, which here forms the surface of the spur, the overlying sheet of Lincoln Porphyry having been eroded off. In places, yellowish quartzite and a bed of yellow, compact, argillaceous rock, called by the miners porphyry, is still found above the limestone. The ore is principally galena, with some sulphates and carbonates of copper, associated with a gangue material of sulphate of baryta. The deposit, where found at the contact, averages about three feet in thickness, and frequently opens out into large irregular chambers extending down into the mass of the limestone. Some of these chambers are 25 feet high, the general average being about ten to fifteen feet. In working the mine a thin streak of mineral, carrying gypsum associated with galena or barite, was followed until it opened out into a chamber or pocket. The ore appears to be richer as it approaches the summit of the peak. About seventy-five feet from the entrance to the tunnel, a slide or fault was found, pitching 45°, with a downthrow to the east. Very little mineral was found on this faulted surface, but following it up from where it was struck in the tunnel it led to a great ore body 45 feet above.

On the extreme of the northeast spur of Lincoln, overlooking the town of Montgomery, are a number of abandoned openings in the Blue Limestone, showing a similar character of irregular deposits of galena, associated with barite, which have been worked principally by open cuts.

MOUNT BROSS.

Moose mine.—The principal mine on Mount Bross is the Moose mine, which was discovered in July, 1871, though the Dwight, which is supposed to be an extension of the same deposit, was first found in 1869. As the mine was closed to visitors at the time of examination, the data with regard to its workings have been obtained from former superintendents. It is situated on the northeast slope of Mount Bross, near the summit of the wall overlooking the Cameron amphitheater. Its ore is galena and copper pyrites, with their various oxidation products, carbonates and sulphates, and the gangue largely barite or heavy spar. It is found in the Blue Limestone, at or near the contact with the overlying Lincoln Porphyry. The ore has been very rich, yielding from two hundred to three hundred ounces of silver per ton. The formation here dips to the southeast. Besides the overlying Lincoln Porphyry, a dike of White Porphyry was observed, crossing the formation a little west of the mine, and doubtless an examination of the underground workings might have disclosed other bodies of porphyry. The tongue of Blue Limestone, which forms the surface of the spur between the Cameron and Bross amphitheaters east of the Moose mine, is, like that of the corresponding spur of Mount Lincoln, honeycombed with open cuts, from which apparently considerable quantities of ore similar to that of the Moose mine have been taken. So far as observed, the silver deposits in this region have been principally confined to the Blue Limestone horizon.

Dolly Varden mine, on the eastern slope of Mount Bross, at a little lower level than the Moose mine, is also in the Blue Limestone. Its ore is generally more oxidized than that of the Moose, and occurs in the body of the limestone, near its contact with a nearly vertical dike of White Porphyry about forty feet thick. The main dike has a strike of between N. 15° E. and N. 30° E., and dips 60° to the northwest. The ore has been found along the southeast face of this dike in a vertical extent of 150 feet, and extending southeastward into the limestone with the dip of the formation to a distance of 100 feet. On the east of the main dike is a second porphyry dike, with a strike nearly east and west and a dip of 45° to the south, which may be simply a branch of the main dike. On the west side of the main dike as yet no considerable amount of ore has been discovered.

In the little valley on the southeast slope of Mount Bross, known as Mineral Park, is the tunnel of a company which was organized during the prosperous times of this region with the avowed intention of piercing the mountain to strike the source of the rich deposits of the Moose and other mines. It was not visited by the writer, hence an exact description of the rocks through which it has passed cannot be given, but it deserves mention, as illustrating the pernicious habit among western miners of spending large sums in running tunnels through the solid rock with no sounder basis of hope than the chance of striking some unknown rich deposit. This tunnel is said to start in quartzite and to have been driven about seven hundred feet. Whether it commences at the Parting Quartzite or at once in the Lower Quartzite is a matter of comparatively little importance. It starts, at all events, below the horizon of the Blue Limestone, and therefore can by no possibility strike any of the ore bodies contained in that stratum. As will readily be seen by reference to Section L of Atlas Sheet X, the tunnel will soon, if it has not already done so, reach the crystalline rocks of the

Archean formation. Its horizontal distance from the Moose mine is about two miles. The expense per foot of driving such a tunnel after the first few hundred feet, as any person familiar with mining can readily see, must be very great. The only ore bodies which it can strike will be those occurring at considerable depths in the Archean rocks. In this extent of two miles a valuable ore body might be struck, especially near some dike of eruptive rock, but the probabilities are against it. As far as present developments show, the veins discovered in the Archean in this region are generally small and of no great value. Even in the case of a well-defined vertical fissure, the advisability of running so long a tunnel to strike it in depth, unless it had been proved by actual exploration to be rich at the depth at which it would be reached by the tunnel, would be extremely doubtful.

BUCKSKIN CAÑON.

The neighborhood of Buckskin Cañon is a region of important developments of ore bodies, which are not confined to the horizon of the Blue Limestone, but occur also at lower horizons, some small bodies having even been discovered in the Archean rocks. The most important of these is the Phillips mine, which was discovered in 1860 by Joseph Higginbotham, commonly known as "Buckskin Joe," from whom the neighboring town received its name. The ore occurs in the Lower Quartzite, near the bottom of the valley, on the south side of the stream. Here an immense lenticular body of iron pyrites, with a little copper pyrites and a gangue of sulphate of baryta, follows the strike of the rocks up the gentle slope of the valley, rising from the stream bed to the foot of the cañon walls. Erosion having laid bare this body in a length of about two thousand feet, it was mined by the early settlers in an open trench from 20 to 30 feet wide and about fifteen to twenty feet deep. Being exceptionally exposed to the action of water and of the atmosphere, the upper part of the deposit was entirely oxidized. By a natural process of concentration this oxidized portion became very rich, so that in the early days, even with the rude processes then in use, consisting mainly in sluicing and grinding in arastras, it was worked at a profit, and from a quarter to a half million dollars are said to have been taken from it. When the unoxidized portion was reached, the ore became poorer, yielding only about six dollars per ton in gold, and the difficulty of treating it in the ordinary stamp-mill was such that the mine was after a few years practically abandoned. The ore body, which at the surface had practically vertical walls, was found to split up in depth into smaller bodies, dipping eastward and following approximately the stratification planes of the quartzite. On the east side of the open cut a body of porphyry forms the wall of the ore body for a short distance, and is probably an offshoot from a larger body of Lincoln Porphyry which crops out in the stream bed a few hundred feet below the mine. Explorations on this ore body have been carried to a comparatively small distance below the surface, and it is not at all improbable that, if systematically exposed in depth, other large bodies might be developed, which, under present conditions, with smelting works within easy reach, might be mined at a profit.

Criterion mine.—On the north wall of Buckskin Cañon, a little above the town of Buckskin Joe and about half way up the cliff, is the now abandoned Criterion mine. Here, just above the junction of the Lower Quartzite with the underlying Archean, in a shallow ravine, is exposed a body of green porphyrite, impregnated with pyrites. A little above this, on the east side of the ravine, is an immense open chamber, about sixty feet in height by one hundred feet or more in length and ten to twenty feet wide, striking N. 35° E. It crosses the beds almost perpendicularly, and seems to have once contained a body of ore, which has been removed either by atmospheric agencies or possibly by some earlier miners of whom no trace is left. The white quartzite adjoining the body is stained with iron oxide and somewhat disintegrated. Nearly adjoining the upper part of this body on the northwest, an irregularly lenticular-shaped ore body has been developed by the tunnel of the Criterion mine. Its form is rather horizontal than vertical, but it connects with the upper part of the large cavity. The vein material is crumbly iron-stained quartz, evidently resulting from the decomposition of the quartzite. The irregular-shaped chamber from which the ore has been taken is in places twenty to twenty-five feet in height, and has been opened for a length of over one hundred feet. No chemical examination was made of the ore to ascertain its actual value, but it contains carbonate of lead and probably silver. The evidence of extended mineralization at this point is so strong that it is well worthy of more thorough exploration. In all probability the ore was originally a sulphuret, though containing a smaller portion of pyrites than the Phillips deposit, but so far as examined it has been entirely oxidized. In the White Limestone above, apparently on the same vertical plane with the large cavity, an opening has been made on two small bodies of galena, following vertical joining planes in the limestone. It seems probable that the ore which once filled the large cavities was originally deposited along a jointing plane or the plane of a small fault, and that the cavity has been enlarged by secondary alteration. A small fault with a movement of about twenty feet can be traced along the bed of the shallow ravine which indents the face of the cliff. It is worthy of note that all theses planes have a common strike, northeast and southwest, which is also that of the fault on the north face of Loveland Hill; this may be observed on the opposite side of the cañon, about a mile above this point and on a line due southwest from here.

Excelsior mine (L).—A little west of the Criterion and higher up on the face of the cliff is the Excelsior mine, likewise abandoned, whose position is marked by the skeleton of an old building standing perched on the edge of the cliff, overlooking the precipice below. Here an irregular body of porphyry or porphyrite traverses the White Limestone, which is somewhat contorted at its contact. The mine is opened by a tunnel, on either side of which the eruptive mass crops. But little ore has been developed, but the rocks are deeply stained with iron, and at the contact of the limestone and porphyry is found a little free gold and silver, with copper and iron pyrites.

The spur of Mount Bross, above these mines, is covered with prospect-holes, the more recent of which are generally in the Blue Limestone. The older prospects, which were made at the time when only gold was sought for, are generally confined to the silicious beds below this horizon.

Colorado Springs mine.—In the Red amphitheater, higher up the cañon, on the south face of Mount Bross, is the Colorado Springs mine, which is still being worked, and obtains rich galena ore from the lower part of the Blue Limestone, at or near its contact with the Parting Quartzite. The ore body averages two or three feet in thickness. The lower portion of the Blue

Limestone is here, contrary to what its name would indicate, quite light-colored. The ore is extremely rich, but very irregularly distributed; it is evidently a replacement of the limestone, there being no evidence of pre-existing cavities.

Dominion mine (L).—In a prospect hole to the north of this, called the Dominion, ore is also found at the contact of the limestone and Parting Quartzite. The ore body averages two feet in thickness and is said to assay from $68 to $500 in silver per ton. It carries brittle silver, silver glance, sulphate and carbonate of copper, and some antimony.

Sweet Home mine.—On the face of the cliff, a little southeast of the Red amphitheater, in the gneiss below the Paleozoic rocks, is the Sweet Home mine. Its ore is found near a body of very much decomposed Lincoln Porphyry, whose surfaces are covered with a yellow coating of oxide of iron, containing arsenic and antimony. This mine is interesting from the varieties of mineral species thus far obtained from it. Among these are cuprite, fluorite (pink and blue), jamesonite, melanterite, rhodocrosite, and zinkenite.

From the Tanner Boy mine, on the southwest side of the gulch, said to be a fissure deposit in gneiss, are obtained remarkably fine crystals of deep red rhodocrosite, or carbonate of manganese, containing a little magnesia. It occurs in very well-defined rhombic crystals, isomorphous with calcite. As might be expected from the prevalence of eruptive bodies, ore is found in small veins in almost every part of the Archean rocks exposed in Buckskin Cañon. As yet, however, no large and important discoveries have been made, and a priori the conditions for concentration of large ore bodies would be less favorable here than in the overlying Paleozoic rocks.

On the south side of Buckskin gulch, at the eastern end of the cliff wall, shown in Plate XIII (p. 128), and formed by the lower Paleozoic beds sloping up toward the crest of Loveland Hill, small ore bodies have been opened up, following jointing planes in the White Limestone and the Lower Quartzite. These deposits are generally extremely thin and of limited extent. They frequently send out branches for a short distance into the adjoining rock, following the stratification planes. Their ore is galena, black sulphurets, and sometimes native silver. The direction of these joints is, as at the Criterion mine, generally northeast and southwest. Their plane is practically at right angles to the stratification. A slight discordance is frequently observed in the beds on either side of the plane, showing a displacement of a few feet. Among those noticed is the Ernest, at the foot of the cliffs just back of the Phillips mine. Here may be seen a phenomenon not uncommon in unstratified rocks, the crumpling of a certain bed, which does not extend to the adjoining beds. This may be supposed to be the result of unequal plasticity, by which, as an effect of lateral compression, one bed has been crumpled or folded, when the others, being more plastic, have expanded sufficiently to allow for the longitudinal contraction. The Ernest is in the upper part of the White Limestone, as is the Rock Island, a little higher up the cañon. The Northern Light, which is shown in Plate XIII, is in the Lower Quartzite, and has a gangue of calc spar. Its vein material has a maximum thickness of 6 feet. Still to the west of this is the Rock Island, which is in the upper part of the White Limestone.

LOVELAND HILL.

In the Blue Limestone, which covers the surface of Loveland Hill, are a great number of prospect holes and small mines, which have opened irregular bodies of ore, principally argentiferous galena and its decomposition products.

Fanny Barrett mine (L).—The most important of these is the Fanny Barrett, which is situated not far from the edge of the cliff overlooking Buckskin gulch, and apparently near the line of fault noticed on the south wall of Buckskin gulch. It is about on the line of strike of the Criterion deposit. The ore fills a so-called fissure some four feet in width, containing galena and what is called by the miners "hard carbonate," i. e., a sulphate of lead, forming greenish-white concentric layers like agate, with some carbonate of copper. The gangue is a soft hydrated oxide of iron, with considerable oxide of manganese, which gives it a black color. In places the ore forms branching deposits and lenticular bodies of considerable size along the stratification planes. The fissure is supposed to have been traced quite across the crest of Loveland Hill to the cliffs facing Mosquito and Buckskin gulches on either side. On the Buckskin side, a little above the contact of the Archean with the Lower Quartzite, a body of ore has been opened in the gneiss, at the junction of the porphyry and the white quartzite. A little below this a vertical fissure has been opened in the gneiss, about two feet in width, containing galena and spathic iron. From the description, this deposit is evidently of the same nature as the other limestone deposits in the region, in this case, however, following mainly a vertical jointing plane. It is hardly probable that such a plane would be found to be actually continuous for any great length, but very likely, from the persistence of its direction, these planes have been developed more frequently in the neighborhood of the line of the fault, and may have been accompanied by a slight displacement. This movement would naturally extend for some distance into the underlying Archean, and the ore currents would follow similar joints formed in these rocks. From analogy with other deposits in this district it is hardly probable that the so-called fissure veins in the gneiss have been filled from below.

Higher up on Loveland Hill, on the ridge which connects it with Buckskin Peak, ore is found in the White Limestone and in the Lower Quartzite. Among the prospects the La Salle ore body occurs between the limestone and the quartzite; Little Nell and Julia are on narrow vertical veins, also striking northeast, in the Lower Quartzite. On the very top of the hill, in a small patch of Lower Quartzite, are the Mountain Lion and Silver Exchange claims, whose ore is galena, with green carbonate of copper, carrying gold and silver. On the Mosquito side of the hill is the Kansas mine, at the junction of the White Limestone and Lower Quartzite, in proximity to the porphyrite body, below which is the Christian Aid.

On the southeastern end of Loveland Hill, towards Mosquito gulch, numerous bodies of ore have been opened in the Paleozoic rocks. Of these the most important is the Orphan Boy, which is abandoned and was not visited. Its ore is found in quartzite, which probably belongs to the Cambrian formation. The Jersey and other claims near Park City, in the bottom of the gulch, have struck galena in the White Limestone.

BETWEEN MOSQUITO AND HORSESHOE GULCHES.

South of Mosquito gulch, on Pennsylvania Hill, are many prospects in the Blue Limestone, and on Ball Mountain in the Weber Grits. In none of these, so far as observed, have any ore bodies of value been found. The prospects in the Weber Grits are especially barren, and it is difficult to conceive what inducement has been offered to the miners to expend the amount of labor they have in this region. In general they seem to have been attracted by the darker beds, containing carbonaceous matter, whether sandstone or shale, which, in general, have a little pyrites in them and sometimes a trace of gold. Along the Sacramento gulches also, and as far as the ridge which divides them from Horseshoe, the labor of the miner seems to have been barren of practical results. This is the intermediate region between the Lincoln Porphyry intrusions around the Mount Lincoln massive and those of White Porphyry along the line of Horseshoe gulch.

Sacramento mine.—In the Blue Limestone of the ridge south of Little Sacramento gulch, and east of the London fault, is the Sacramento mine, which has been worked for a number of years and produced a good deal of rich ore. At the mine itself the Blue Limestone forms the surface, the overlying porphyry having been removed by erosion, so that it is impossible to say what proportion of the original ore body may have existed in the upper part of the limestone. The ore, which consists of galena and sand carbonate of high grade in silver, is found generally in irregular masses throughout the limestone beds. Some clayey iron oxide occurs as gangue, and considerable sulphate of baryta is found associated with the mineral. Several hollow caverns have been found by the explorations of the mine, in one of which remarkably beautiful stalactites of white fibrous arragonite occur. These caves have evidently been formed since the deposition of the ore and are connected with natural jointing planes in the limestone. Some of the caverns are partially filled with a clayey material, which runs four to five ounces per ton in silver, and is evidently an infiltration from the ore bodies, which continue on either side of the cave without bearing any relation in form to it. The limestone above the horizon of the ore is generally darker colored than that below, which is of a light-gray color and more crystalline structure; but the distribution of the ore bodies is too irregular to consider them a definite deposit along the dividing plane between these two varieties of limestone. The lighter limestone has the ribbed structure, characteristic of the Blue Limestone horizon, very beautifully developed.

On the ridge between Spring Valley and Horseshoe gulch the Mudsill claim has found galena at the same horizon, forming a contact deposit between the darker and lighter limestones. To the east of this the Little Christine has developed small veins of galena in the porphyry overlying the limestone. On the eastern slopes of the range south of Horseshoe gulch, as well as on Sheep Mountain ridge, the contact of Blue Limestone and White Porphyry has been but little prospected, and, although no considerable ore bodies have yet been discovered in this portion, there is a fair probability that more systematic search will yet develop them, Sheep Mountain being a most especially promising locality.

CREST OF THE MOSQUITO RANGE.

A great deal of prospecting has been done along the very crest of the Mosquito range, the bare, precipitous, and often almost inaccessible slopes of high mountains seeming to offer especial attractions to the hardy prospector. This may be, in part at least, explained by the fact that in these elevated regions the rock surfaces are generally exposed and swept clean of obscuring soil and other surface accumulations, and therefore the "indications" which they follow, such as a staining of iron oxide or of carbonate or of silicate of copper, are more readily seen than in the lower country, where, when found, ore is likely to be more remunerative. It seems hardly necessary to say that the idea that prevails among some prospectors, and to which credence is given even by some educated persons, that rich ore bodies are more frequent near mountain-tops, has no basis of geological evidence to support it. But few mines have been actually worked at the crest of the range, and these only intermittently, which is hardly surprising when one considers that it is frozen up nine months in the year.

At the head of Mosquito gulch, and just west of the London fault, the New York mine is said to have developed considerable rich ore at the contact of the Blue Limestone with a small overlying sheet of White Porphyry of local occurrence.

London mine.—This is the most promising of these elevated deposits, and has been known for a long time; but, besides the natural difficulties of its position, apparent bad management and contested ownership have combined to retard its development. Fuel being expensive at such an altitude, the novel idea was adopted of using wind as a motive power, and a powerful and strongly constructed windmill was built near the mine; but it proved incapable of resisting the force of the fierce winter storms which sweep over Mosquito Pass, and was blown down the first winter after its erection.

The geological horizon of the deposits, as explained in Chapter IV (p. 142), is apparently near the junction of the White and Blue Limestones, at the contact with a sheet of White Porphyry lying parallel with the stratification. It is there shown that the strata are here turned up at a steep angle against the London fault on its west face, and that as their strike makes an acute angle with the plane of the fault, they are gradually cut off by it, successively higher horizons coming in contact with the fault line as one goes across London Mountain from north to south. This is shown in a general way on the map of the Mosquito Range; but it must be borne in mind that the line of fault could not be traced with absolute accuracy, owing to its covering of débris, and that, even if it could have been, the scale of the map is too small to represent the necessary details of structure. This is a point of great interest as regards both its geological structure and the light its exploration might throw upon the genesis of ore deposits, and it is to be regretted that the underground workings of the mine had not been sufficiently extended to afford more definite data on these points. In the single, somewhat hurried visit that could be made only a cursory view could be obtained of the ore deposits and their surroundings. They consist of two so-called veins, parallel and about forty feet apart, standing nearly vertical and striking about northwest and southeast Of these, the southwestern has an average thickness of about four feet and is extremely well defined. It carries sulphurets of lead, copper, and iron, with some gold. The northeastern body is a free-gold ore, impregnating the country rock for about two feet in thickness. Both have been proved in a length and in a height above the tunnel level of several hundred feet, and the ore was said to average $40 per ton and upwards, consisting in considerable proportion of free-milling ore.

The main working tunnel was run into the base of the mountain on a level with the mine buildings, a little west of south in direction, and cut the ore bodies at 510 and 590 feet from its mouth, respectively. From this drifts run along the veins in southeasterly direction, from which it was intended to carry stopes upwards toward the crest of the mountain, which, at its summit, is seven hundred to eight hundred feet above the tunnel level. The direction of these ore bodies converges towards the fault plane, and it was estimated that the latter would be cut by these drifts under the crest or a little south of that point, the distance at which it would be reached depending on the angle of dip of the fault plane, of which nothing definite could be ascertained. If this dip were very steep or near the vertical, the distance of intersection would be correspondingly greater. It was evident from what was seen during the examination that the ore bodies, in spite of their appearing at first glance to be vertical veins, are following, at least approximately, the stratification of the sedimentary beds, and in this respect resemble the majority of the deposits of the region. They were, therefore, probably formed before the folding and faulting took place; in which case they will end at the fault plane, their continuation beyond that plane having been carried up by the movement of the fault, and since eroded away. It is, of course, possible that the ore currents by which they were deposited came up along the fault plane itself, in which case the ore bodies would naturally be found to continue downwards on that plane. Should this be found to be the case it would be the first instance observed in all this region of original ore deposition on one of these great fault planes.[1]

Peerless mine.—At the head of Horseshoe gulch are a number of claims, two of which have developed a considerable amount of ore. Of these, the Badger Boy is in the Blue Limestone, which is exposed in the bed of the north fork of Four-Mile Creek. The Peerless mine is at the very crest of the ridge, just south of Peerless Mountain. Its ore is found about twenty feet below the top of the limestone in irregular lenticular bodies. The limestone is here very silicious, the upper part resembling quartzite, but it

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[1] Since this visit a wealthy company has developed the mine, building a railway eight miles long for transporting ore and supplies and a mill for the reduction of the ore. The fault plane is said to have been cut by the drifts along the vein, but trustworthy accounts are wanting as to the results of developments bearing upon the above points.

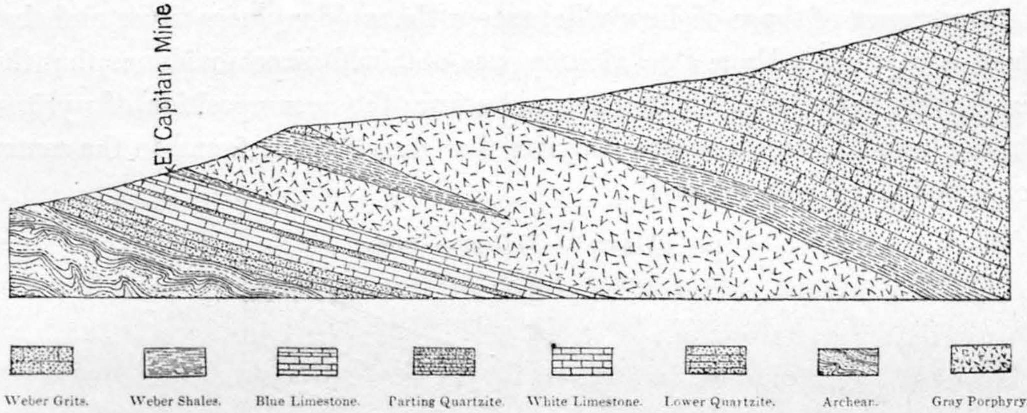
is also very much iron stained, and it seems probable that this is merely a secondary deposition, resembling the gangue matter of some of the Leadville mines. The ore is galena and hard carbonates, with a little copper, generally in the form of carbonate.

WESTERN SLOPES OF MOSQUITO RANGE.

On the west side of the range, outside the limits of the Leadville map, are many prospect holes, both at the contact of the Blue Limestone and White Porphyry and in the underlying and overlying rocks indiscriminately. The former have generally developed a little ore, even as far south as Weston Pass, but the developments are extremely superficial and unsystematic. Here, as elsewhere, miners seem particularly attracted by the black layers which are common in the Weber Grits, and considerable labor has been fruitlessly expended in exploring the bed of black shales on Empire Hill, which contains fossils whose casts have been replaced by iron pyrites. In the region north of Leadville, on the west side of the range, owing to the fact that the surface is principally occupied by the Weber formation, the by no means inconsiderable labor expended by prospectors has been comparatively fruitless. Slight traces of ore are frequently found in contact with the numerous bodies of porphyry which traverse this formation, but as yet none worthy of any extended development. The teachings of this examination are that remunerative ore bodies are far more likely to be found in the calcareous than in the silicious beds, and for this reason prospectors would be wise to direct their labors principally to prospecting the outcrops of limestone strata. The actual ore contact between Blue Limestone and overlying porphyry occurs on the wooded slopes bordering Tennessee Park, and is so much obscured by surface accumulations that it is practically unprospected.

El Capitan mine.—This mine is situated just west of the extreme northwest corner of the Mosquito map, the location of its shaft being given in the margin. Greater importance attaches to it than the economical value of the ore bodies would otherwise warrant, because it is the first discovery of rich ore in Blue Limestone north of the Leadville district. No indications of ore were found upon the surface in this region, and it was simply because it was recognized as the "Leadville contact" that the limestone was prospected. The general geological structure is shown in Figure 5, which represents an east-and-west section approximately along the northern boundary of the area mapped.

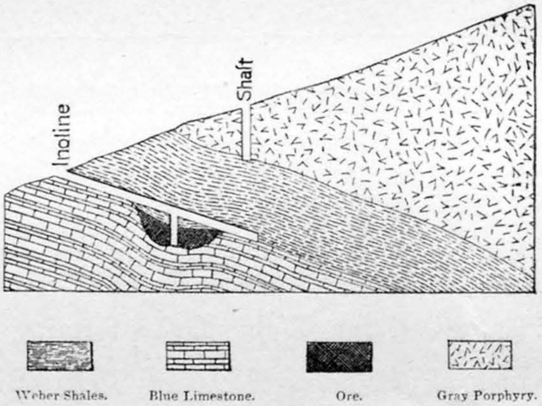
Figure 5.—Taylor Hill.



The Cambrian, Silurian, and Blue Limestone beds here outcrop along the base of the wooded hill slopes, dipping at about 20° to the eastward. Above these is the main sheet of Gray or Eagle River Porphyry in great thickness, the sandstones of the Weber Grits forming the upper part of Chicago ridge. The lower part of the Weber Grits, or the Weber Shales, which here consist mainly of quartzite, with only a few thin layers of shale, are split into two portions by the porphyry body and separated from the main body of Weber Grits above. A fragment of these, about forty feet thick at the outcrop, is left resting directly on the Blue Limestone below the porphyry body. The Blue Limestone has the ribbings of white calc spar, which characterize it at Leadville and at the Sacramento mine. No ore has been found at the actual porphyry contact, but it occurs at the upper surface of the Blue Limestone, extending downward to a depth of 15 to 20 feet, as shown in Figure 6, which is a section on a larger scale through the mine itself.

As at Leadville, the ore occurs as a replacement of the limestone, but yields free gold instead of lead and silver, the latter occurring only in small quantity, and lead as yet not having been found. The vein material consists of iron-stained clay, with layers of light-colored banded chert, from a few inches to two or three feet in thickness, which pass into coarse silicious sandy material, resembling quartzite; this latter constitutes the richest gold ore. Impure kaolin and Chinese talc are also found in the vein material, and are generally gold-bearing. The deposit is thus in almost every respect a counterpart of those of Leadville, except that gold replaces silver and that lead is wanting. From this vicinity was obtained direct evidence that the iron oxide in these vein materials results from the decomposition of pyrites, in the kernel of undecomposed pyrite that sometimes is found in the center of a nodule of vein material.

Figure 6.—El Capitan Mine.



The following table shows the result of a chemical examination of one of these nodules from No Name gulch, to the south of the El Capitan mine. Analysis No. 1 is the pyrite kernel, No. 2 the dark zone next to it, and No. 3 the lighter-colored outer zone. From these it would seem that the process of change is not only an oxidation of the sulphide, but a gradual removal, first of the sulphuric acid and later of part of the hydrated oxide of iron, and a replacement of this by insoluble material, mainly silica. From the two analyses of rich ore (Nos. 4 and 5) in the same table, the richer of the two appears to be that in which this change has proceeded furthest, and which is called "silicious ore," although its contents in alumina would seem to indicate a mechanical as well as a chemical interchange.

	Substance.	Insoluble.	FeS_2	Fe_2O_3	Al_2O_3	CaO MgO	Ignition.	Ag	Au	Total.
1	Pyrite kernel, No Name gulch	7.36	56.31	34.44	....	Trace.	....	....	....	98.11
2	Darker intermediate zone	16.36	....	73.42	....	0.09	10.13	....	....	100.00
3	Lighter outer zone..	54.28	....	39.84	....	0.06	5.82	....	....	100.00
4	"Silicious ore," El Capitan mine.....	*Silica* 82.40	*PbO* Trace	10.38	4.16	1.217	*Water* 1.83	0.002	0.011	100.00
5	"Gold ore," El Capitan mine........	11.56	None	77.60	....	0.609	10.23	0.0002	0.0004	99.9996

TEN-MILE DISTRICT.

The ore deposits of this district, which lies north of the western half of the Mosquito map, will be made the subject of a separate monograph, and only a brief mention of their prominent characteristics need be made.

The larger and more prominent ore bodies of this district resemble the Leadville deposits in that they occur as a replacement of limestone strata, and generally along the upper surface of these strata. In this respect they are contact deposits, but the immediately overlying rock is generally a sandstone instead of a porphyry sheet. The development of eruptive rocks is even more remarkable than at Leadville, and they occur generally as intrusive sheets nearly parallel with the stratification, sometimes also in the form of transverse sheets and dikes. Thus, although the ore body may not be found in actual contact with an eruptive rock, it is never very far removed from one.

The ore bodies themselves are less completely oxidized than those of Leadville, probably because of the greater elevation above sea-level, in consequence of which surface waters are imprisoned by frost during a greater proportion of the year. The unoxidized portion of these larger ore bodies consists entirely of sulphides, mainly pyrite, zinc blende, and galena, while the oxidized portion, near the surface, consists of iron-stained, clayey vein material, carrying carbonate of lead and chloride of silver, scarcely to be distinguished from the average Leadville ore, though it is generally less rich in silver. The manner of occurrence of ore bodies along the upper surface of the limestone, and extending irregularly from that surface downwards, gives evidence also that the ore currents acted from above downwards. Moreover, in the most important mine in the district, the Robinson mine, the actual crack, either fault or jointing plane, through which these currents may have reached the limestone can be traced in the roof of the main ore body and parallel with its longest dimension. Along this a certain amount of pyrite has been deposited in the micaceous sandstone which overlies the limestone.

The beds in which these deposits are found belong to the Upper Coal Measure formation, whose base has been arbitrarily taken at the Robinson Limestone, so called from the mine of that name in which it was first examined.

DEPOSITS IN THE ARCHEAN.

On the opposite side of Tennessee Park from the El Capitan mine, one of the prominent peaks of the Sawatch range is known as Homestake Peak, from the mine of that name, which occurs in the Archean rocks on its slopes. It was discovered before the carbonate deposits of Leadville, and is said to have produced considerable rich galena, carrying from 30 to 60 per cent. of lead and from 200 to 250 ounces of silver to the ton. In it has been found a small amount of an arsenical nickel mineral, supposed to be gersdorffite. As the mine was not visited, nothing further can be said as to the character of the deposit.

In the Buckskin and Arkansas amphitheaters are many prospects and a few small mines, on what are assumed by their owners to be true fissure veins. None of the mines were being worked at time of visit, and they were therefore mostly inaccessible. They were examined, however, whenever they came within the line of work, and it was found, so far as such examination extended, that they were invariably a simple mineralization of the country rock along some plane which admitted the percolation of water, whether a jointing plane or the plane of a fault, and that the mineral occurred as an impregnation or replacement of the country rock, the vein materials being simply an alteration product of this rock, and not foreign material filling pre-existing fissures, as is supposed to be the case in a true fissure vein.

CHAPTER VI.

GENESIS OF LEADVILLE DEPOSITS.

The ore deposits of the region having been described in the foregoing chapters, with as much detail as space allows, it remains to gather together briefly the evidence upon which the general conclusions with regard to their origin and manner of formation, given in Chapter I, have been arrived at.

Before entering upon the discussion of this evidence, however, I would state that I do not claim that these conclusions are absolute or final, nor do I wish to be considered as offering them as general theories applicable to all ore deposits. They are those that seem to best accord with the facts now at my command, and these facts have been presented carefully, conscientiously, and without bias in favor of any preconceived theory. Other facts may come to light which may lead to a modification of these conclusions, and I reserve the right to adopt such modifications wherever the evidence afforded by further developments in this or any other district which may come under my notice shall seem such as to require it. The study of the underground structure of a mining district necessarily presents many problems which cannot be definitively solved until the entire area has been explored; nevertheless, deductions may be made from observed facts, and from analogy with facts gathered in other carefully studied regions, which will justify a hypothetical forecast that may be of great practical service to those engaged in mining, even should that forecast be proved later to be incorrect in some of its details.

In studying the genesis of the ore deposits of Leadville, one main difficulty at the time that this work was carried on was their universally oxidized condition, resulting from secondary alteration by surface waters, which rendered it difficult to arrive definitely at the original form in which they were laid down. When the mine workings shall have been extended to such depths that the mineral is found practically sheltered from the secondary action of surface waters the study will be rendered more easy and the results arrived at will be more certain.

In this discussion it will not always be possible to follow the exact order given in Chapter I, which was adopted rather for the purpose of presenting a concise and readily comprehensible statement than as the logical sequence of the processes actually involved in the ore deposition.

MANNER OF OCCURRENCE.

Perhaps the most suggestive fact in the manner of occurrence of the Leadville ores is their predominance in the beds of the Lower Carboniferous or Blue Limestone, and one is naturally led to seek a cause for this preference. It being admitted that the ores were deposited from aqueous solutions, it is readily apparent that the more soluble limestone beds would be more easily acted upon by these solutions than the other sedimentary beds of the region, which consist mainly of sandstones and argillaceous shales, and are much less susceptible to the action of percolating waters. This, however, does not explain why the Blue Limestone should have been chosen rather than any of the other calcareous beds of the region.

There was a theory at one time prevalent among mining men in the West that the great silver deposits in limestone were peculiar to a definite geological horizon, a conception that perhaps had its origin in the tendency of some geologists to generalize from the coincidence that certain classes of deposits have been found in different parts of the world at the same geological horizon. Indeed, some have gone so far as to base their determination of the horizon of certain beds, when other evidence was wanting, on the occurrence in them of this class of deposits. The great silver deposits in limestone of the western United States are, in point of fact, found throughout the whole range of the Paleozoic system, and the horizons of no two districts have as yet been proved to be absolutely identical. As a case in point, the vast deposits in the Ten-Mile district, only 16 miles from Leadville, occur in the Upper Carboniferous Limestones, and *not* in the Blue or Lower Carboniferous Limestone. It is true that these deposits in the western United States are mainly confined to Paleozoic beds, but this is hardly to be wondered at when one considers that the other horizons are almost entirely silicious or argillaceous and contain as a rule a very limited development of limestone. It is evident that broad generalizations on the basis of geological horizon alone are not only unfounded but are misleading, and that it is more logical to seek for an explanation in the local conditions of each individual district.

The causes that may have influenced the concentration of ore in any particular bed of limestone may have been physical or chemical; that is, the structural or physical conditions of the region may have been such that the solutions were naturally directed to that particular bed, or the composition of that bed may have been such as to render it peculiarly susceptible to the action of waters reaching it from adjoining rocks or to cause the precipitation of the minerals held in solution by those waters.

The *physical* or structural conditions of this region have been shown by geological descriptions to be peculiarly favorable to the concentration of percolating waters in the Blue Limestone. The great intrusive sheets of porphyry are found to follow it most persistently, mainly along the upper surface, less frequently along its under surface, and also cutting transversely across it. These intrusive bodies are also found at other horizons, it is true, but at none so persistently and so uniformly as at this. Thus both ascending and descending currents would readily reach these beds, the latter trickling through the uniformly permeable eruptive rock, the former following up the walls of the channel through which it was erupted Such waters, after passing through a medium of different composition, would be ready for a chemical interchange with the limestone, but in the case of ascending waters it does not appear evident why this interchange should have taken place along one wall of the channel rather than the other, while with descending waters this action would naturally commence on the upper surface of the limestone bed. Thus the physical conditions afford a reason for the predominant choice of this horizon.

As regards the *chemical* composition of the bed, the evidence is less conclusive. Some authors have been inclined to regard dolomitic limestone as the peculiar habitat of lead and silver deposits, from the fact that many silver-bearing limestones have been found to be dolomitic, making thus a generalization out of a coincidence. But the evidence here fails to confirm any such coincidence. Not only are the other limestone beds of the region, which do not carry metals in quantity, equally dolomitic, but in the adjoining Ten-Mile district the ore occurs in non-dolomitic limestones, and in the Robinson mine is confined to the upper part of a limestone which is almost chemically pure carbonate of lime, while it does not extend into the lower part of the horizon, which carries nearly 7 per cent. of carbonate of magnesia.

A reason which might have been offered for the greater susceptibility of dolomite to the decomposing action of percolating waters, viz, the supposition that the carbonate of lime in a dolomite is first attacked, and that thus a disintegration of the rock is readily commenced, has been disproved in the case of the rocks of this region by the chemical experiments made, which show that the waters act simultaneously upon both carbonates, or rather upon the double carbonate and not upon either of its component parts.

As regards the minor differences in composition between the different dolomites of the region, the chemical investigations which it has been possible to make furnish little more than suggestions. The beds whose physical conditions are most similar to those of the Blue Limestone, and which are also most frequently found mineral-bearing after it, are those of the Silurian or White Limestone.

The first striking difference between the two is the darker color of the former, which is presumably due to organic matter and possibly in part to sulphide of iron, as suggested by Mr. Guyard. Chemical analysis confirms the indications given by outward appearances, showing an appreciable amount of organic matter in the Blue Limestone and none in the White. If the metals were brought in in the state of sulphates, the organic matter would promote their reduction to sulphide. On the other hand, the Robinson limestone, already cited, affords an instance opposed to this view, for there it is the light-colored limestone which carries the ore, while the darker limestone, which it may be assumed has more organic matter, is quite barren. A second difference between the Blue and the White Limestone is that the former contains no silica and the latter over 10 per cent. It may be that owing to this difference in composition the former is more soluble, but this could only be satisfactorily determined by a practical experiment which should be carried on for a sufficiently long time to imitate in some degree the processes of nature. Before undertaking such an investigation, which would require several years for its proper conduct, it would be advisable to gather data from various districts to determine whether in point of fact the silicious limestones or dolomites are in general less frequently ore-bearing than those of normal composition.

From the above considerations it seems that in this district the main cause of concentration of ore in the Blue Limestone has been its physical or structural conditions, and that the influence of its peculiar composition has been at best of minor importance

COMPOSITION OF ORES.

As at the time the materials for these investigations were gathered underground explorations had not yet penetrated to depths which were beyond the oxidizing influence of surface waters, a great part of the ores and vein materials collected were necessarily of secondary origin; their composition therefore affords only indirect evidence in regard to the composition and genesis of the original deposits by indicating the agencies and processes by which these alterations were effected.

Independently of this evidence, however, there exist good a priori grounds for the assumption that the original deposits were in the form of sulphides: first, in the fact that, in ore deposits in general, oxidized ores almost universally give place to sulphides in depth and beyond the reach of surface waters; and, second, in that in the analogous deposits of the adjoining Ten-Mile district oxidized ores similar to those of Leadville are seen to result directly from the alteration of a mixture of galena, pyrite, and zinc blende.

Carbonate ores.—In the following table are given the analyses of three specimens of carbonate ore, selected as being especially free from impurities, together with that of an average taken from a thousand specimens of carbonate ore as it is delivered by the mines to the smelters, and containing therefore a considerable admixture of what might be considered gangue.

No. 1, from the Adelaide mine, is an extremely pure white sand, made up of small crystals of cerussite, and in the mass looking not unlike a coarse white sandstone. These white sands are, as a rule, less rich in silver than the discolored and relatively impure ones, but this has exceptionally little silver. It does not occur in the Blue Limestone or in the immediate vicinity of it, but between the White and Gray Porphyries.

No. 2, from the Little Chief mine, is a rich sand-carbonate, discolored and appearing to contain more impurities than it really does.

No. 3, from the Waterloo mine, is also a discolored carbonate, less rich than the former, in which only the constituents of cerussite and pyromorphite were determined qualitatively.

No. 4 is the analysis of a mixture of ores made by Mr. Th. Fluegger, chemist of the Harrison Reduction Works. It was made for the purpose of showing the average proportions of the most common elements in the ore, and it may be assumed that neither did the laboratory facilities at his command admit of, nor his purpose demand, an equal standard of accuracy with the others in regard to the rarer elements and their combinations.

Carbonate ores.

	1. Adelaide.	2. Little Chief.	3. Waterloo.	4. Average ore.
PbO	80.352	75.408	77.980	24.77
Al_2O_3	0.444	1.415		3.99
Fe_2O_3	0.467	1.940		24.86
FeO	0.299			0.89
MnO	0.137	0.074		4.03
MnO_2		1.386		$\left\{\begin{matrix}K_2O\\Na_2O\end{matrix}\right\}$ 0.98
CoO	Trace	Trace		
ZnO		0.095		Trace
CaO	0.303	0.335		2.36
MgO	0.068	0.056		3.04
SiO_2	0.651	1.972		22.59
Sb_2O_5		0.121		Sb= 0.02
As_2O_5	Trace	Trace		As= 0.01
P_2O_5	1.532	Trace	6.480	
SO_3	Trace	0.486		S= 0.90
CO_2	14.700	14.251	10.180	5.58
Cl	0.255	0.288	0.840	0.09
H_2O	0.395	1.140		5.53
Ag	0.009	0.777	0.047	0.31
Au	Trace	Trace		Trace
Total	99.612	99.744	95.527	99.95

In discussing the above analyses the combinations of different elements in the ores and the indications afforded as to the state in which they existed prior to alteration will first be considered, next the possible processes, and finally the agencies, by which this alteration may have been effected.

Gold exists in these ores only in traces; it has never, so far as known, formed any considerable value in the limestone ores, except in the Florence mine, where it has been observed in the native state. It is in this form without doubt wherever it occurs. It is generally supposed to be most commonly associated with pyrite in sulphuret deposits, and the assays of the Pyritiferous Porphyry given below (Table IV, Appendix B) confirm this hypothesis. They show, however, that it may be associated with galena, and it is said that from the Printer Boy mine a specimen of galena was obtained in which two crystals of this mineral were connected by a thread of gold. Most of the small amount of gold that is produced by Leadville mines comes from deposits in porphyry or sandstone, which sometimes carry a little copper also. The greater part of the gold in Leadville bullion comes from ore shipped to the smelting works from deposits in the Archean rocks of neighboring districts.

Silver exists in the oxidized ores invariably in the form of chloride, as far as can be judged by actual observation. With a strong lens, minute crystals and flakes of chloride can be detected on the crystals of cerussite, or coating cleavage surfaces and cracks in the various vein materials, even in comparatively poor ores. The above analyses confirm this observation. The first three show more than enough chlorine for combination with the silver they contain. In the mixture the amount of chlorine is 0.01 less than is required by the amount of silver; if it could be assumed that the determination of chlorine was absolutely accurate, a small portion of the silver might be assumed to be combined with the sulphur, antimony, and arsenic. Since there is a reasonable doubt on this point, the weight of evidence is still in favor of its probable combination with chlorine.

The association of minerals in the Leadville ore deposits generally furnishes a priori grounds for the assumption that in the original deposit silver was in the form either of simple sulphide or of some of the antimonial or arsenical sulphides in which it commonly occurs in nature. As regards the probability of its having existed in the latter combinations, it is significant that in the Adelaide ore, which contains scarcely any silver, only a trace of arsenic and no antimony are found; whereas both are present and in appreciable quantities in the richer ores, 2 and 4. It will be seen later that these substances are generally detected in the vein materials, and are found in considerable quantities in the smelting products, antimony being in relatively larger proportions than arsenic. It will also be seen that all the vein materials and the country rocks adjoining the ore bodies contain a small but persistent percentage of silver.

Lead occurs mainly as carbonate, sometimes as sulphate, and quite often in the form of chloro-phosphate. In the Little Chief ore a little sulphate still exists, but no appreciable amount of pyromorphite. In the Adelaide ore, on the other hand, only a trace of sulphate is found, but a notable proportion of pyromorphite, while in the Waterloo ore this mineral amounts to 32.07 per cent., as against 61.78 per cent. of cerussite. Analysis 4 shows no pyromorphite, but the evidence of many tests of ores and vein materials besides those given above and the composition of smelting products, all show that this mineral is very common and widespread throughout the region. The amount of carbonic acid given in Analysis 4 is notably insufficient for the lime, magnesia, and oxide of lead. If it is assumed that the sulphur is all combined with the latter, the mixture contains only one part of anglesite to three parts of cerussite. Even this is, however, opposed to the evidence of observation, for anglesite can but rarely be seen in the ores, whereas cerussite is most common.

That the lead originally occurred in the form of galena is shown by the frequent occurrence of the unaltered mineral in the center of masses of carbonate. A chemical test of one of these galena nodules showed that between it and the carbonate was an extremely thin crust, made up mostly of sulphate. It is, therefore, probable that in the alteration it passed first into sulphate and then into carbonate, although the intermediate product is not always distinguishable.

In the Iron mine a considerable amount of native sulphur, associated with cerussite, was found as an alteration product of galena.

Iron and *manganese* might be more properly considered gangue materials. They are mainly in the form of hydrated sesquioxide and protoxide, respectively. A little protoxide of the former and peroxide of the latter ore was found. The former may be combined as basic sulphate, which, as will be seen later, sometimes forms considerable bodies. The latter is probably anhydrous, as pyrolusite is frequently distinguishable in actual crystals and sometimes forms considerable ore masses. Although no actual pyrite was observed in the Leadville deposits, there is little doubt that iron existed in this form in the original deposits. With regard to the original form of manganese there is more uncertainty, as the sulphides of this metal are relatively rare. It sometimes occurs as carbonate, in association with sulphides of other metals, losing its carbonic acid when they are oxidized. It is so common an associate of iron in oxidized ores and so seldom noticed in unaltered sulphides that it might be thought to have been in part brought in as oxide during secondary alteration. It is possible that some of the iron in the ores may be combined with silica as silicate and with arsenic as arseniate. These were not absolutely proved to exist, as it was not considered of sufficient importance to give the time necessary for the tests.

Zinc occurs in the above ores in very small proportion, and probably in the form of silicate (calamine), since this is the only mineral of zinc that has been observed in the Leadville deposits. It is rarely visible, and generally forms fine, needle-like, silky white crystals, lining drusy cavities and cracks or joints in the vein material and limestone. There is little doubt that it originally occurred as zinc blende, and, from analogy with the Ten-Mile deposits, it may be presumed that it formed a much larger proportion of the deposit in its original form than it does now. The much greater solubility of its sulphate than that of the other metals would account for its more thorough removal by surface waters.

Of other metals cobalt is the only one not already mentioned that is detected by the above analysis. In addition to this, copper, bismuth, molybdenum, and vanadium have been locally observed in mineral combination, and Mr. Guyard claims to have detected nickel, tin, indium, selenium, tellurium, and cadmium in the furnace products.

The *earthy bases*, alumina, lime, and magnesia, preserve a certain proportion between each other in the three analyses. This accords with the assumption that the deposits are a replacement of limestone, and the fact that the apparently pure cerussite from the porphyry in the Adelaide ground contains as much magnesia and lime as the Little Chief ore is in so far a confirmation of the assumption on structural grounds (see p. 407) that the ore bodies in the former mine are the replacement of fragments of Blue Limestone caught up by the porphyry intrusion.

Silica is in relatively higher proportion in the mixture (4) as compared with the earthy bases, but preserves a comparatively even relation to iron oxide, which would be expected from their association in the ores in general. It is a question whether the silica is in part in actual combination with iron or whether it all exists as free silica. Mr. Guyard suggests that some of the lead may exist as silicate, but this also is not definitely proved.

Chloride ores.—In the following table is given the proportional amounts of chlorine, bromine, and iodine, respectively, in three typical chloride ores relatively free from other minerals, tested for the purpose of determining whether sufficiently definite quantitative relations exist between these substances to justify the recognition of more than one mineral species. I and II were obtained from the pale-green mineral, and III from a colorless specimen.

Chlorine, bromine, and iodine.

	Mine.	Cl.	Br.	I.	Total.
I.	Robert E. Lee	13.78	85.63	0.59	100.000
II.	Amie	9.80	89.99	0.21	100.000
III.	Big Pittsburgh	99.925	....	0.075	100.000

The results are somewhat negative, and, so far as they go, lead to the conclusion that the chloride ores are merely mixtures, in varying proportions, of chloride, bromide, and iodide of silver. The green chlorides, which are of very common occurrence, are generally called embolite, and this application of the term is justifiable if this mineral is considered simply an indefinite mixture. In the analyses of embolite given in Dana's Mineralogy the relations of bromine to chlorine vary from 1:0.33 to 1:5.67, but no iodine is given. In I and II, given above, the relations are 1:0.16 and 1:0.11, and all three contain a small amount of iodine.

According to Moesta,[1] in the mines of Chañarcillo, Chili, there is a regular gradation in these compounds according to depth, the pure chlorides being found in the upper levels down to a depth of 20 meters; below these come mixtures, containing proportions of bromide increasing with the depth; still lower iodide of silver is added to the mixture, and pure iodide of silver occurs at sixty to seventy meters in depth, directly over the deposits of galena and pyrite, in which the first sulphides of silver are found. It has not been possible to detect any regularity whatever in the distribution of these compounds in the Leadville deposits. The colorless and green chlorides are the prevailing varieties, but minute yellow crystals of iodide of silver have been observed in the deposits of the Chrysolite mine.

In Chañarcillo among secondary deposits about half the silver occurs in the native state, and of the other half the far greater proportion occurs as chloride or chloro-bromide, iodides being rare and the proportion of iodine in the mixture being very small. Thus, according to Moesta, from whose work all these statements are taken, the relative proportions of chlorine, bromine, and iodine are essentially the same in which they exist in sea-water. It may be added that the same relations exist in surface waters, and in rocks whenever these substances have been detected in them. Moesta's theory with regard to the chlorides, &c., of Chañarcillo, that they were formed by the action of sea-water (since there is evidence that the region has been covered by the ocean within comparatively recent times), would not apply to the Leadville deposits, for the reason that they have not been submerged since the upheaval and erosion that brought them within the reach of surface waters.

Basic ferric sulphates.—In the following table are given the analyses of three specimens, contributed by Mr. L. D. Ricketts, from material observed by him to frequently constitute a persistent bed under the rich ore bodies, especially on Carbonate Hill. This material is of ocherous-yellow color,

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[1] Chlor-, Brom- und Jodverbindungen des Silbers in der Natur. Dr. Fr. A. Moesta. Marburg, 1870.

somewhat like a dry clay, and easily recognized by its external appearance, though, as will be seen, of very variable composition:

Basic sulphates.

	Mine.	SiO_2	Fe_2O_3	Al_2O_3	CaO	MgO	K_2O	Na_2O	H_2O	PbO	Bi_2O_3	As_2O_5	P_2O_5	SO_3	Cl	Ag	Au	Totals.
I.	Maid of Erin.	None	46. 70	None	0. 06	0. 06	5. 33	1. 68	10. 54	4. 27	0. 08	0. 46	0. 08	30. 53	0. 02	0. 0048	Trace	99. 8148
II.	Morning Star.	0. 30	42. 98	0. 20	0. 64	None	6. 31	0. 83	10. 12	8. 27	None	0. 42	1. 58	27. 81	0. 26	0. 0036	None	99. 7236
III.	Lower Waterloo.	0. 36	44. 40	0. 23	None	None	0. 15	0. 37	8. 99	19. 50	None	0. 39	0. 11	25. 07	0. 04	0. 075	None	99. 685

These substances are somewhat complex basic sulphates, and might be considered to be a mixture of jarosite with varying proportions of basic ferric sulphate. They are evidently an alteration product of pyrite and galena, although, while nodules of galena rich in silver are occasionally found in them, pyrite has not yet been detected. The absence of zinc in the specimens analyzed is noteworthy, and is in accordance with the observation already made, that it has been further removed from the original ore bodies than the other metals, presumably on account of the ready solubility of its sulphate. The persistent percentage of the alkalies, which were found in sensibly the same proportions in three other specimens tested, would suggest that the waters which produced this alteration reached the ore bodies after passing through decomposed porphyry. Their chief interest lies in the definite evidence they afford that they result from the oxidation of sulphides. Similar products have frequently been observed in old mine openings where large bodies of pyrite have been long leached by surface waters. Copperas first formed gradually loses a portion of its water on exposure to the air, and the protoxide of iron becomes sesquioxide. Further exposure leads to the formation of limonite.

Processes of alteration.—Assuming that the metals in the original deposits existed in the form of sulphides, it comes next in order to consider the possible processes by which the combinations noted above may have been derived from them.

By oxidation all the metallic sulphides may be transformed into sulphates,[1] and the relative solubility of the latter is in inverse proportion to the

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[1] Percy's Metallurgy of Silver and Gold, Part I. London, 1880.

distance to which, at Leadville, the metals appear to have been removed during secondary alteration from the original locus of deposition.

According to J. Roth,[1] of such sulphates 100 parts of water dissolve, respectively:

At 11° C., 0.004383 parts } sulphate of lead.
At 13° C., 0.003155 parts } sulphate of lead.
At 12° C., 21.300 parts sulphate of iron.
At 41° C., 41.300 parts sulphate of zinc.

The sulphate of silver is less soluble than either that of iron or of zinc, but probably more so than the sulphate of lead; and at 100° C. it is said to be soluble in 88 parts of water.

Sulphide of silver may be reduced to native silver by the action of water at 100° C., during which, according to Moesta, the water itself is decomposed and SO_3 and HS are formed. Native silver is slowly converted into chloride in waters containing alkaline chlorides.[2]

Sulphide of silver is converted directly into chloride of silver at ordinary temperatures when exposed to the action of sulphate of sesquioxide of iron, chloride of sodium, and water.[2] The presence of air is not necessary for this reaction, but if the sulphate of protoxide of iron is substituted for the basic sulphate, chloride of silver is not produced without the presence of air. This indicates that a salt of sesquioxide of iron must be formed before the sulphide of silver is decomposed.

Moesta's[3] experiments show that this reaction may take place with a solution of NaCl alone at 100° C., and even at 20° C., but that it is quickened by the presence of chloride of magnesium, and still more by powdered pyrite; also that the combination with iodine is more rapid than with chlorine.

Sulphate of lead is transformed at the ordinary temperature into carbonate by solutions of fixed alkaline carbonates, and also by those containing bicarbonate of lime and atmospheric air. Carbonate of lead (cerussite) is soluble in 7,144 parts of water saturated with carbonic acid.[4] In conversion 100 parts by weight of sulphide of lead become 126.78 of sulphate, and this in turn 111.71 parts of carbonate of lead.[5] The increase

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[1] Allgemeine Geologie, p. 59. Berlin, 1879.

[2] Percy's Metallurgy of Silver and Gold, Part I. London, 1880.

[3] Chlor-, Brom- und Jodberbindungen des Silbers in der Natur, p. 40. Dr. Fr. A. Moesta. Marburg, 1870.

[4] Percy's Metallurgy of Lead, p. 40. London, 1870.

[5] J. Roth, op. cit., p. 243.

in volume from sulphide to carbonate is 28.13 per cent. Such changes of weight and volume might account for the prevailing sandy condition of the carbonate ores. It may be assumed that they were once comparatively solid masses of galena, and during these transformations occupied, first a larger, then a smaller space, thus leaving interstices between the minute crystals of cerussite of which the sand carbonate consists. In the presence of phosphoric acid and alkaline chlorides, sulphate of lead may be transformed into pyromorphite. Sulphate of zinc is readily transformed into carbonate, and, as a further stage of alteration, finally becomes a silicate.

Sulphate of protoxide of iron may become carbonate in the presence of earthy carbonates. Carbonate of protoxide by oxidation and hydration becomes hydrated sesquioxide or limonite, which on loss of its water becomes hematite; the latter in contact with organic matter may be reduced to the protoxide or magnetite.

The sulphate of the protoxide may by oxidation change directly to sulphate of sesquioxide or basic ferric sulphate, which in turn becomes limonite and hematite.

The fact that carbonate of iron is so rarely found in the Leadville deposits would suggest that their limonite might have been formed in the latter way, and that the basic ferric sulphates, of which analyses are given above, may have been formed directly from the sulphide.

Agents of alteration.—The alteration having taken place through the agency of waters coming from the surface, it remains to consider whence they may have derived the substances which would have facilitated the alteration of the sulphide. Surface waters in general are said to carry a certain amount of atmospheric oxygen, of organic matter, of chloride of sodium, and of phosphoric acid. An analysis of surface water at Leadville, taken from the reservoir in Big Evans gulch, was found by Mr. Hillebrand to contain, in a million parts, $K_2O = 1.12$, $Na_2O = 1.92$, $SO_3 = 7.20$, and $Cl = 1.14$.[1]

Of the rocks through which they may have passed, out of eight eruptive rocks analyzed, six were found to contain phosphoric acid and five

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[1] The Delaware River at Trenton contains 1.20 and two glacier streams in the high Alps an average of 0.4 Cl in 1,000,000 parts.

chlorine. Traces of phosphoric acid were found in both of the limestones analyzed. In the Lingula shales, which directly overlie the Blue Limestone, 5.14 per cent. phosphoric acid was found in the portion containing casts of Lingula and 0.35 per cent. in that free from these casts. Chlorine was found in each of the sixteen specimens of dolomitic limestones, from different horizons and localities, that were tested for that substance. Bromine and traces of iodine have also been detected, but they exist here, as in sea-water, in far smaller proportions than chlorine Organic matter is found abundantly in most of the rocks of the Carboniferous formations. The specimen of Blue Limestone analyzed gave only 0.03 per cent., but in the overlying Weber Grits not only are there frequent beds of carbonaceous shales, passing at times into actual coal, but the sandstones sometimes contain as high as 4 per cent. of carbonaceous matter.

Relative richness of galena and cerussite.—The greater richness in silver of galena over cerussite in this region is very noteworthy. Mr. L. D. Ricketts,[1] who made a detailed study of the ores of Carbonate Hill, states that the average tenor of cerussite in that locality is less than 40 ounces of silver to the ton, while galena averages 145 ounces of silver to the ton. He also states that assays of five galena nodules, and of the carbonate crusts on each, showed that in proportion to the amount of lead present there was six times as much silver in the galena as in the cerussite. Table XIV, Appendix B, which gives the assays of various specimens of ores, vein materials, and adjoining country rocks collected during the investigation, shows a similar relation in the silver contents of galena (No. 15) and of its cerussite crust (No. 16), 420 ounces and 28.6 ounces, which are in even greater contrast than in the cases cited above.

The fact that silver is found disseminated throughout the vein materials and adjoining country rocks, even where little or no lead is found, shows that during secondary alteration silver has been further removed from its original locus and more widely disseminated than lead. In fact, it may be assumed that the outlines of the present bodies of lead ore vary but little from those of the original deposits, but it would hardly be safe to make such an assumption in regard to silver ores. It is apparent that this

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[1] Ores of Leadville, p. 87. Princeton, 1883.

relative distribution of the two metals was brought about by surface waters, and is therefore dependent on the relative solubility of the combinations of the respective metals formed during alteration.

Silver sulphate is probably more soluble than either the sulphate or the carbonate of lead, but, as it is not known to occur as a mineral, it cannot be assumed that silver necessarily passed through sulphate during its change from sulphide to chloride.

The chloride of silver is said to be insoluble in pure water at ordinary temperatures, but Vogel[1] and Hahn[2] have shown it to be soluble to a certain extent in the alkaline chlorides. According to Stas,[3] it is, in a measure, soluble in pure cold water, its solubility varying according to its physical condition and the temperature. Its solubility is greatest when in the flaky state, as precipitated in the cold from a sufficiently dilute solution of silver, and diminishes as the flakes shrink. It is precipitated from the solution by the addition of an alkaline chloride. Still further evidence of its solubility is afforded by the occurrence of the mineral Huantajayit, discovered by Raimondi in Peru,[4] which consists of 11 per cent. Ag Cl and 89 per cent. NaCl. According to Sandberger, this is readily soluble in a little water, but an excess of water produces a precipitation of the chloride of silver. It may therefore be assumed that chloride of silver is soluble in surface waters under certain conditions, but is very readily precipitated from its solution.

Although the statements given above as to the relative solubilities of lead and silver salts are not so definite as might be wished, the fact of the relatively greater richness of galena over cerussite in these deposits seems so well established as to justify inverse reasoning, namely, the deduction of an argument in favor of the greater insolubility of the lead salts. It might be assumed from Stas's experiments that freshly-formed chloride of silver (flaky) would be more soluble than carbonate of lead, but that after the lapse of sufficient time the latter might become more soluble than the chloride of silver, especially in water charged with carbonic acid.

Outcrop deposits richer than those in depth.—There is a fair foundation for the generalization that in the deposits, as developed at the time of this investigation, the ores were growing poorer in silver as exploration extended far-

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[1] Wagner's Jahres-Ber. 1874, 22, p. 481.

[2] Trans. A. I. M. E., 1873-'74, p. 99.

[3] Compt.-rend., 1870-'73, p. 998.

[4] Neues Jahrb. f. Mineralogie, 1874, p. 174.

ther from the surface. In the case of the Iron mine the statement to this effect by those in charge of the mine was supported by actual figures; in other cases it was an opinion founded on general observation, which was still worthy of credence. Whether it will continue when the unoxidized deposits are reached remains to be proved.

A ready explanation for this condition of things may also be found in the relative solubility of the products of alteration and in the fact that the deposits near the present surface may be considered to be simply the relics of larger deposits, gradually removed by erosion as the alteration by surface waters went on. The original deposits may be assumed to have been a mixture of galena, pyrite, and blende in proportions which, while subject to a wide local variation, would bear a certain average relation throughout the region. Given a ton of this mixture, by the alterations which have been above noted it would probably decrease in weight, though its volume might remain sensibly the same; zinc and iron would be removed as soluble sulphates, the former in greater part, the latter in very appreciable amount, and their volume would be replaced in part by the increase in volume of the lead during its transformation into carbonate, in part also by silica and earthy bases brought in mechanically by the waters. Although silver salts have been assumed above to be, in this case, more soluble than those of lead, they are much less so than those of iron and zinc, and a relatively small proportion of the silver would have been removed from the given ton of ore. The space formerly occupied by this ton of ore would, therefore, be occupied by oxidized material weighing much less than the original ton, but which would contain, in proportion to that weight, more lead and silver and less zinc and iron. In accordance with this explanation, the proportion of lead to the ton of ore should also decrease in depth as a general average of the district.

Again, as by gradual erosion the deposits—say, for instance, those in the easterly-dipping beds of Carbonate and Iron Hills—approached the surface, actual surface waters running down along the dip would meet bodies composed mainly of carbonate of lead and chloride of silver. If carbonate of lead is soluble in water containing free carbonic acid, the portion thus dissolved would be carried away entirely, for in their continued passage through limestone it may be assumed that the waters would preserve their excess of carbonic acid. The chloride of silver, on the other hand, might be assumed to have become insoluble in the long time that had elapsed since it was formed; or, if any were taken up, it would probably be thrown down again in a short distance by a slight change in the character of the descending solutions. Thus a gradual increase in the proportions of silver over lead would be taking place in the zone which was being brought by the erosion of overlying rocks nearer and nearer to the surface.[1]

COMPOSITION OF VEIN MATERIALS.

In the remaining vein materials, which constitute the relatively valueless portion of the deposits, iron and manganese, generally in the form of hydrated oxides, are the most prevalent metals. Carbonate of iron is very rarely found, and pyrite never among the oxidized ores. The comparative absence of zinc has already been remarked, yet this metal is quite uniformly detected in the products of smelting. Arsenic and antimony are

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[1] In a paper read at the Chattanooga meeting of the American Institute of Mining Engineers in May, 1885, by Mr. F. T. Freeland, superintendent of the Iron Silver Mining Company, the following analyses are given of sulphide ore from the Minnie mine, situated on the southeast slope of Iron Hill, adjoining the Colonel Sellers, made by William R. Boggs, jr. They are interesting as showing the relative distribution of silver in the different components of the unoxidized ore:

	I.	II.	III.	IV.
	Galena.	Blende.	Pyrite.	Mixture.
Pb	72.65	6.71	2.21	50.86
Zn	5.66	55.08	14.24	12.86
Fe	1.60	4.00	35.40	9.30
S	15.66	32.44	44.76	24.50
Ag (ounces)	(41.5)	(94.5)	(4.5)	(11.5)
Ag (per cent)	0.14	0.324	0.014	0.039
Au	Trace	Trace	....	Trace
Residue	4.12	0.92	2.70	1.88
Total	99.83	99.474	99.324	99.439

His mixture, calculated from the relative amounts of Pb, Zn, and Fe, in column IV, would consist of a little less than three parts of galena to one each of pyrite and blende. The tenor of silver given for this mixture is, however, only a little over a quarter of what would result from a mixture in these proportions of these three minerals, each having the tenor in silver given in its separate column. It seems doubtful, therefore, if these separate values can be taken as a fair representation of the average tenor of the different sulphides in the ore. That given for zinc blende is evidently abnormally high, as may be seen on comparison with the pyrite column. This pyrite contains an admixture of about one-fifth of its weight in zinc blende, and if this zinc blende ran as high in silver as that given in column II, this fifth alone would yield nearly 20 ounces of silver, whereas the whole mass is said to contain only 4½ ounces. The silver tenor of the galena, on the other hand, represents a fair average of the galenas found in the older mines, though these, as is shown by the assays in Appendix B, vary very much and may contain over ten times as much silver as this.

only accessory constituents in either ore or vein material. The composition of some of the typical forms of vein material, taken from Table X, Appendix B, is given below.

Vein materials.

No.	Name.	Locality.	Insoluble.	FeS_2	Fe_2O_3	MnO_2	ZnO	CaO	MgO	CO_2	H_2O and ignition.	Ag	Au	Total.
✓ 1	"Hard Carbonate."	Scooper mine.	80.00	......	13.91	0.10	......	......	......	....	5.978	0.012	Trace	100.00
✓ 2	Silicious hematite.	Chrysolite mine.	8.20	......	54.14	22.36	2.56	......	......	....	12.709	0.031	None.	100.00
✓ 3	"Iron ore" ...	Kenosha mine.	21.10	......	7.70	65.98	1.00	......	......	....	4.22	0.012	None.	100.00
4	Pyrite nucleus.	No Name gulch.	7.36	53.61	34.44	......	......	......	Trace	....	....	......	......	......
5	Dark intermediate zone.		16.36	......	73.42	......	......	......	0.09	[a]Trace	10.13	......	......	......
6	Light outer zone.		54.28	......	39.84	......	......	......	0.06	[a]Trace	5.82	......	......	......
7	Altered limestone light	Garden City mine.	[b]10.08	......	0.81	0.39	......	27.17	19.49	42.78	....	......	......	100.72
8	Altered limestone dark		[b]20.02	......	0.97	1.56	......	24.15	16.71	37.37	....	......	......	100.78
9	Chert nodule .	El Paso mine.	[b]96.3	3.06 per cent. soluble in solution of potassium hydrate, remainder chiefly Fe_2O_3 and Al_2O_3.										
10	Chert under ore body.	Little Pittsburgh mine.	[b]97.9											
11	Granular quartz under ore body.	Waterloo mine.	[b]81.29	3.93 per cent. soluble in solution of potassium hydrate, remainder chiefly Pb, CO_3, and Fe_2O_3.										
12	Porphyry breccia, with ore cement.	Evening Star mine.	[b]84.2	Cement chiefly pyromorphite, with some cerussite and galena and a little calcite.										

[a] SO_3 [b] SiO_2

No. 1 is a cavernous red rock, with slightly greasy luster, which, though containing only a trace of lead carbonate and too little silver to pay for working, is yet classed as "hard carbonate" from its outward appearance, this miner's term being quite elastic in its application. It is mainly silica, combined perhaps in part with oxide of iron. Rock from this same mine, of quite similar appearance, but with a few delicate crystals of cerussite and horn silver lining the cavities, constitutes a very rich ore.

No. 2 is a dark-colored "iron ore," intermediate between the black iron and the jaspery iron of the Fryer Hill mines. It contains a rather unusual percentage of zinc, and its silica contents are much lower than its outward appearance would indicate. Both contain a little antimony, and the former a little PbO, which were not determined quantitatively. Both contain also a small percentage of silver, but not enough to constitute pay ore.

No. 3, though classed as an iron ore by the miner, is more properly an ore of manganese.

Nos. 4, 5, and 6 are given to illustrate the different stages in the alteration of pyrite, which would lead to the above vein materials. The specimen was obtained at some distance from Leadville, it not being possible to find any pyrite in the mines themselves; it occurs, however, at the Blue Limestone horizon and in analogous conditions to the Leadville ores, except that no rich silver-lead ores had yet been found at the locality. No. 4 is the comparatively unaltered pyrite nucleus; No. 5, the inner zone of alteration; and No. 6, the lighter-colored outer zone. Only traces of sulphuric acid are found in either. The percentage of metallic iron has slightly increased from No. 4 to No. 5, as has that of insoluble matter; sulphur alone has, therefore, been removed, probably as sulphate of lime. From No. 5 to No. 6 iron has disappeared rapidly and been replaced by insoluble material; it has presumably been carried away largely as hydrated oxide, which is probably commencing to replace the carbonates of lime and magnesia in the adjoining rock. The rapid increase in insoluble matter could not, it would seem, be entirely original matter, but must in part have been brought in by the decomposing waters. The conditions here may be contrasted with those which probably existed where the basic ferric sulphates above described were formed. In the latter case the decomposition probably took place within the mass of a large body of metallic sulphides; here it was a small amount of sulphide in direct contact with the country rock, and free oxygen would have been present in relatively greater proportion, so that the sulphide would have been more readily decomposed by the oxidation of the iron.

Nos. 7 and 8 represent the early action of waters coming from such a decomposition upon the country rock, in this case a fragment of Blue Limestone found on the south slope of Iron Hill, near the fault. No. 7 is the lighter portion, No. 8 the darker, in which the replacement has apparently proceeded much farther. In the latter the oxides of iron and manganese have increased only slightly in amount, and the decrease in carbonates is mainly supplied by the increase in silica, an increase which, as in the former case, must be mainly accounted for as coming from an extraneous source; it was probably taken up by the waters in their passage through the porphyry.

Nos. 9, 10, and 11 represent the most silicious forms of vein material, comparatively free from bases. Nos. 9 and 10 are the black cherts so common in the ore bodies, and 11 the granular quartz which frequently replaces it, especially on Carbonate Hill. In each the silica is still partly soluble in a moderately strong solution of potash. Besides silica the two former contain iron and alumina and probably a little organic coloring matter. These cherts are thoroughly compact and generally form barren streaks or floors in the ore bodies; sometimes, however, chloride of silver is found coating their cleavage surfaces. The granular quartz, on the other hand, which is very porous, frequently contains crystalline cerussite partly filling the minute cavities and then constitutes an ore, though, as in the present case, it is liable to be mistaken in the hand specimen for a white quartzite.

No. 12 is the breccia of White Porphyry in the small ore-body which was found above the Blue Limestone horizon, just west of the small dike in the Evening Star ground. The percentage of silica is above the normal, showing that the ore-bearing waters have removed a portion of the bases. In the cementing material galena is altered partly to pyromorphite, partly to cerussite, and the calcite may indicate that the ore-bearing solutions reached here after passing along the limestone contact or may simply result from the decomposition of the porphyry.

The assays of vein material and limestones in Table XIV, Appendix B, show that all the specimens tested carry silver in appreciable amounts, though in the case of limestones which were gathered in the various mines, and in that of the Breece Iron ore, the tenor in silver is much less than is generally credited to those used as flux by the smelters. They are such amounts as might be expected to have been carried into them by the surface waters after leaving the ore bodies.

Kaolin and Chinese talc.—These names are given in the mines to certain substances, evidently alteration products of porphyry, occurring with great persistency along the contact of limestone and porphyry, where they sometimes form the only vein material, and also within the ore bodies, sometimes at quite a distance from the contact. In the latter case they probably result in most cases from small offshoots of the porphyry, such as have been mentioned as occurring in the Little Pittsburgh mine, and, as are shown in Fig. 1, Plate XXII, penetrating the as yet unaltered limestone. The characteristic Chinese talc is compact, with conchoidal fracture, somewhat translucent, with a sort of opalescent luster, and is easily cut by the fingernail when fresh, but becomes opaque and hardens on exposure to the air. White when pure, it is generally more or less discolored and veined by oxides of iron and manganese. The miners often carve it into pipes and figures. The so-called kaolin is white, opaque, and generally plastic, but also hardens on exposure. No true kaolin was found among the specimens collected. In the following table, I and II would be considered kaolins, and III, IV, and V, Chinese talcs. VI, VII, and VIII are specimens from the Lower Waterloo mine, contributed by Mr. L. D. Ricketts; in spite of their different composition, they are not to be distinguished in the hand specimen from the Chinese talcs.

Kaolin and Chinese talc.

	Mine.	SiO_2	Al_2O_3	Fe_2O_3	FeO	ZnO	CaO	MgO	K_2O	Na_2O	H_2O	SO_3	P_2O_5	Totals.
I.	Amie	48.72	34.01	0.56	0.66	......	......	1.11	9.88	0.67	4.42	......	......	100.03
II.	New Discovery	43.66	37.78	......	......	......	0.22	0.30	Trace	Trace	17.95	Trace	......	99.91
III.	Big Pittsburgh	4.55	35.60	2.26	......	......	Trace	Trace	2.73	5.28	15.05	34.55	......	100.02
IV.	Morning Star	24.47	38.05	0.93	0.77	......	0.23	0.30	2.72	1.30	16.67	15.48	0.23	101.15
V.	Swamp Angel	27.89	33.79	......	......	......	0.53	1.14	2.83	1.56	16.51	15.75	......	100.00
VI.	Lower Waterloo	35.33	10.38	....	......	33.05	1.62	0.71	....	......	19.06	......	....	100.15
VII.	....do	35.97	8.81	......	......	35.40	1.87	0.80	......	......	17.46	......	......	100.31
VIII.	....do	37.54	24.76	0.64	......	18.43	0.63	0.71	0.66	0.36	16.37	......	......	100.10

351
+4276
422
-3070
475
452 - 483

24-9-80 +81
" " "
" " "
" " "
" " "
25 " "
24 " "

The compositions given above show that these substances are mixtures of hydrated silicates of alumina, with more or less sulphate of alumina, which, in the case of the last three, is replaced by silicate of zinc. As is generally the case with such alteration products, it is difficult to consider them distinct minerals. The occurrence of zinc in the last three is somewhat unexpected. Of the others, III is perhaps the most abnormal. In the mine it was a pure white, extremely plastic mass, which could be molded like plaster or potter's clay, and became quite hard on exposure to the air. Sulphuric acid is a very common constituent of these substances, having been found qualitatively in each one of five other specimens taken from widely separated parts of the district. Where the substance occurs in the immediate vicinity of an ore body this acid may be readily conceived to have come from the oxidation of the metallic sulphides; but in the case of those occurring on barren contacts, far away from any known body of metallic minerals, as is the case with V, it would seem that their formation might date back to the passage of the sulphurous waters which brought in the original ore deposits.

Lime and magnesia salts.—Although lime and magnesia are found in small quantities in both ores and gangue materials, it is rather remarkable, when one reflects that the country rock is a dolomitic limestone, that their minerals are so uncommon. Calcite occurs as incrustation on crevices and lining cavities or druses in the iron, but never in any large amount. Gypsum is rarely found, although it seems evident that it must have been one of the most important products of alteration. It must therefore be assumed that, owing to its ready solubility, it has been entirely carried away.

Barite.—Barite, or heavy spar, is a not uncommon constituent of the gangue, but it is very irregularly distributed. It generally occurs in aggregations of tabular crystals, frequently concentrated in considerable masses, and more or less stained by iron oxide. Chloride of silver is generally found associated with it. This association of barite and chloride of silver is noteworthy. Among the miners the presence of the former mineral is considered a sure indication of rich chloride ore. In this connection it is interesting to recall Miller's experiments, mentioned by Sandberger,[1] showing that sulphide of barium dissolves pyrargyrite, or ruby silver, without decomposition. The frequent presence of antimony in the ores and vein materials renders it probable that a part of the silver may have originally existed as ruby silver. Moreover, although there are no experimental proofs, it is probable that waters containing sulphide of barium would dissolve the

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[1] Neues Jahrbuch für Mineralogie, 1869, p. 309.

sulphide, or sulpho-salts, of silver under the conditions of time and supply which probably prevailed during the process of decomposition of the original ore deposits. From such a solution the chlorine in the limestone might have precipitated the silver at the same time that the sulphide of barium was transformed into sulphate.

Manganese.—Another empirical generalization of the miners in this region is that, where a large amount of manganese is found in the iron vein material, rich chloride deposits are likely to be found in the immediate vicinity. It is worthy of note in this connection that barium is a most frequent constituent of manganese ores. The fact that the oxides of manganese when treated by hydrochloric acid evolve chlorine is also suggestive. If it is assumed that silver in its passage from sulphide to chloride passes through sulphate, and that hydrochloric acid was formed in the surface waters, say, by the action of sulphuric acid formed in some of the reactions that may have taken place, the presence of manganese oxide would favor the liberation of chlorine, which, in turn, would form chloride of silver from the sulphate when the latter came in contact with carbonate of lime, and the lime be carried away as sulphate.

ORES DEPOSITED AS SULPHIDES.

That the ores were originally deposited as sulphides would legitimately be assumed, from the almost universal observation in nature that such oxidized ores pass into sulphides in depth. So generally is this accepted as a rule in ore deposits that it would require special demonstration to prove beyond a doubt that the native metals or their oxides and chlorides (except perhaps gold, tin, and the platinum group of metals) are in any particular case original, and not the result of secondary alteration from sulphides. Analogy with the deposits in the neighboring Ten-Mile district affords more direct evidence in favor of this assumption. Moreover, since the completion of the field work of this investigation, explorations on the dip in Carbonate and Iron Hill have proved that the oxidized ores actually do pass into sulphides. As yet no systematic description of these sulphide deposits has been published from which it may be learned whether the metals exist exclusively as sulphides, but in the absence of any statement to the contrary it seems fair to assume that such is the case.

It remains, then, to consider the possible reactions which may have brought about the deposition of ores under the circumstances and in the manner assumed above; and in considering these it must be borne in mind that it is not possible to reproduce in the laboratory all the conditions that may have prevailed at the depths within the earth's crust at which these deposits were formed, and that therefore reactions may have taken place and combinations may have been formed under these conditions which, from laboratory experience alone, might not be deemed possible.

The sulphides of the heavy metals may be precipitated, according to Roth,[1] from various solutions: first, where they exist as sulphides, by sulphides of the alkalies and alkaline earths; second, where they exist as carbonates and sulphates, when they come in contact with solutions containing the alkalies and alkaline earths or sulphureted hydrogen; third, where they exist as sulphates, which in contact with organic matter are reduced to sulphides. The metallic sulphides are soluble in waters containing alkaline sulphides or sulphureted hydrogen, and silica and the earthy bases in water containing alkaline carbonates. Solfataric waters (that is, hot waters charged with mineral matter arising from some unknown source below) are known to contain sulphureted hydrogen and the alkaline sulphides and carbonates. On the supposition that the metals of these deposits came up from the unknown source below or were derived from pyrite and galena in neighboring rocks, it might be assumed that the iron and lead at least were actually brought in as sulphides; in this case, however, it is somewhat difficult to conceive the reaction by which the sulphides should replace the cabonates of lime and magnesia, and, so far as laboratory experience teaches, it would seem necessary that the carbonates should have already been dissolved out and carried away before the sulphides were deposited. This apparently involves the pre-existing cavity theory. It is, however, conceivable that the dissolving out of the former so immediately preceded the deposition of the latter that the process was practically an interchange of substance for substance, or the commencement of a change from sulphide to sulphate may have taken place in presence of the carbonate, and

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[1] Allgemeine Geologie, p. 563.

the sulphate have been immediately reduced to sulphide again by organic matter or other reducing agency.[1]

That the direct replacement of the dolomite by sulphides is not impossible, however, seems proved by the fact that galena, zinc blende, and pyrite are found in nature as pseudomorphs after calc spar,[2] and the last two, also, as pseudomorphs after dolomite.[3] The sulphates of the metals are more or less soluble in water, especially when it contains some free sulphuric acid. Their reduction to sulphide through the agency of organic matter is a matter of common observation.

The reactions by which the Leadville deposits might have been made from solutions carrying the metals as sulphates are more readily conceivable. In contact with dolomite containing organic matter the sulphates would be reduced to sulphides with the formation of carbonic acid. The waters thus charged with an excess of carbonic acid would dissolve and remove the carbonates of lime and magnesia, which would be replaced by the metallic sulphides. Any excess of sulphuric acid would form soluble sulphates of lime and magnesia, which would also be carried away. If these sulphates were reduced to sulphides they would render the waters more capable of dissolving out the dolomite.

The metals might have been taken up in the form of sulphates by waters percolating through rocks, where they might have been brought into this combination by the oxidation of sulphides or by the decomposition of silicates. It might also be conceived that during their passage these sulphates would be reduced to sulphides by contact with organic matter before they reached the locus of the deposit.

The most important objection to the hypothesis that the metals were brought in as sulphates is that the lead sulphate is so insoluble compared with that of iron or zinc; and yet the amount of galena in the ores was probably greater than that of zinc blende.

Sulphide of barium would be precipitated as sulphate of baryta in contact with the limestones, owing to its relatively greater insolubility than the sulphates of lime and magnesia.

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[1] J. Roth, Allg. Geologie, p. 235.

[2] J. Roth, op. cit., p. 171.

[3] J. Roth, op. cit., p. 184.

Silica, when brought in by waters containing alkaline carbonates, in which it is notably soluble, might form silicates of the alkalies, the carbonic acid of the latter serving, as suggested in the case of sulphides, to render the waters capable of carrying away the earthy carbonates. Later the combined alkalies might be in part replaced by other bases, such as oxide of iron, and in part actually dissolved out, leaving free silica.

MODE OF FORMATION.

It has been assumed that the ores of this region were originally deposited, first, from aqueous solutions; secondly, by a metasomatic interchange[1] with the country rock; and thirdly, in the form of sulphides. Direct evidences of processes which went on in former geological epochs at great depths below the surface are necessarily difficult to obtain, especially where, as in the present case, the field of observation was confined to material which has been more or less altered since those processes had ceased. It is therefore necessary in the commencement to assume the more probable among possible processes, and then to see to what extent the assumed process may be reconciled with observed facts.

The agencies by which mineral matter may be carried from one place to another within the earth's crust are heat and water, or a combination of the two. It was only in the very infancy of geology that heat alone was seriously admitted to be a possible agent for the formation of mineral deposits in depth. The nature of such deposits was soon found to be such as to preclude the possibility that they might have resulted from the consolidation of a fused mass. Sublimation, on the other hand, as a means of forming such mineral masses, involves a combination of heat, pressure, and water, and may therefore in one sense be considered to be a form of aqueous solution. Its practical demonstration, however, is confined to laboratory experiments, which can at best be but an imperfect imitation of the process of nature. The removal of the materials of which ore deposits are formed by the agency of water alone may be observed to be going on in nature at the present day. Hence this agency, to which, under the comparatively un-

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[1] By metasomatic interchange is meant an interchange of substance, without necessarily involving, as does pseudomorphism, the preservation of the original form of the substance replaced or even of its original volume.

known conditions which prevailed where deeper-seated deposits were formed, a certain amount of heat may have been added, is the one adopted by the majority of students of vein phenomena to account for the removal of the vein materials from place to place within that portion of the earth's crust that comes under our observation.

That it was from aqueous solutions that the Leadville vein materials were deposited is a necessary corollary of the assumption that the deposition took place as a metasomatic interchange between them and the country rocks, since the various materials of which they consist could have been brought in and the dolomite and other rock substances have been removed only by the agency of water. A further necessary corollary of the metasomatic interchange is that the ores were not deposited in pre-existing cavities, as is generally assumed to be the case in ore deposits, particularly those in limestone. The three assumptions being thus interdependent, evidence in favor of either may be considered, in so far, a proof of the others, and it will not be necessary to consider them separately. Direct evidence that the original sulphides in the region were deposited in this manner is necessarily difficult to obtain, where secondary alteration has gone so far as it has in Leadville; but indirect and negative evidence is abundant. When the unoxidized deposits have been thoroughly opened by future explorations, so that it will be possible to study them in their unaltered and original condition, an opportunity will be offered for testing the correctness of the deductions here made.

Indirect evidence.—In their present condition there can be no doubt that the ore bodies are a replacement of the country rock. In the case of the limestone deposits they grade off gradually into the country rocks, the only regular outlines of the bodies being those which are formed by the contact of the limestone with the adjoining porphyry; the other outlines are irregular and ill-defined. Not only are fragments of unaltered limestone found entirely inclosed within the ore bodies, but the latter sometimes occupy the entire space between surrounding sheets of porphyry, which the geological structure shows must have been formerly occupied by the original limestone bed. The chemical analyses of the ores and vein materials given above show lime and magnesia to be constant constituents, decreasing proportionally from the outer limits of the body toward its interior. When these substances are in sufficiently large proportion to be visible to the eye, they are seen to be, not in the crystalline condition in which they would be expected to be if they were brought into a pre-existing cavity and then deposited, but in the same granular condition in which they exist in the country rock. Although it may be said that the present outlines of the oxidized ore bodies are not necessarily the same as those of the original sulphide deposits, it is probable, from the study that has been made of the processes of alteration, that they preserve a general proportion and relation to those outlines, and do not vary from them sufficiently to invalidate the deduction that the original deposits could not have been made in open caves. The deposits in rocks other than limestone consist of metallic minerals and of altered portions of the country rock, in which the structure of the latter can sometimes be still traced, and are not the regular layers of matter foreign to the country rock, which results from the filling of a pre-existing fissure or cavity by materials brought in from a distance and deposited along the walls.

In the case of the still unaltered sulphide deposits of Ten-Mile district, which may reasonably be assumed to have been formed in an analogous way, the arrangement of the particles of the original rock can frequently be seen to be preserved in the metallic minerals, which maintain a certain parallelism with the original bedding planes in the lines defined by minute changes in these minerals.

Negative evidence is afforded by the absence of that condition of things which would naturally be expected to exist if the ore bodies had been deposited in pre-existing cavities, as has been assumed to be the case by those who have contented themselves with this a priori assumption founded on the theory generally given in text-books, without taking time to study the phenomena as they actually exist. The common character of caves which have been dissolved out of limestone is that their walls are coated with a layer of silt or clay, which has been left undissolved by the percolating waters, and that these walls, where undisturbed, have a peculiar surface of little cup-shaped irregularities. There is also almost invariably an accumulation at the bottom of the cave of irregular fragments of limestone, which have broken off from its sides or roof. Observation shows us, moreover, that deposits of mineral matter made in pre-existing cavities are in more or less regular layers, parallel with the walls of the cavity, and that where this approaches a spherical shape, even in a slight degree, these layers are concentric. The most perfect type of this arrangement is seen in the agates which fill geodes.

Were the ore bodies of Leadville the filling of pre-existing cavities, not only would it be expected that a certain parallelism within the walls or an arrangement in layers of their various mineral constituents should have existed, but these walls would have been defined by a distinct clayey selvage, all of which could hardly have been entirely obliterated by the secondary alteration which has taken place. Further, an examination of the outlines of these ore bodies afforded by the maps and sections shows the physical impossibility of their having once been open cavities. What would have supported the roofs of such broad continuous openings as they would represent, or, in cases where they occupy sensibly the entire space between two sheets of porphyry, why did these sheets not close together? Again, how could such cavities have been formed at the depth at which these deposits were originally formed, which it has been shown must have been about 10,000 feet below the rock surface?

The caves which now exist in the limestones of this region are of extremely recent origin, and, as has been shown, cut through limestone and ore bodies indiscriminately. The action of the surface waters which formed them is therefore not only recent, but more recent than that which produced the greater part of the secondary alteration of the ore bodies. Those who maintain that the deposits in limestone have necessarily been deposited in pre-existing cavities do not in all cases,[1] it is true, distinctly state that these cavities must have been formed by surface waters; but it yet remains to be proved that any of the caves which are so commonly found in limestones have been formed at any great distance from the surface. The majority certainly have not, and, since it is generally admitted that the power of easily dissolving limestones is acquired by the waters

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[1] Prof. J. S. Newberry (School of Mines Quarterly, March 1880) distinctly states that the western deposits in limestone have been deposited in caves which, like the Mammoth Cave, were formed by surface waters, and that probably these deposits will prove of limited extent in depth, since the excavation of limestones "must be confined to the zone traversed by surface drainage."

from the free carbonic acid and organic acids which they take up at the surface, it seems probable that these would have been neutralized and the solvent power largely lost long before they could have reached depths comparable with those at which these deposits were formed.

ORIGIN OR SOURCE OF THE METALLIC MINERALS.

Ascension or lateral secretion.—The origin of the metallic contents of ore deposits has been, from the very earliest days of geology, a most fruitful theme of speculation and theorizing, probably for the very reason that so little has been done toward obtaining data, founded upon actual observations or experiments, to support one theory or exclude another. In the days of the bitter contests between Neptunists and Plutonists the supporters of either school allowed only the extreme alternatives, that the vein materials were washed into the veins from the surface (descension theory) or that they were forced into them in a molten condition from below (ascension theory). Probably in either case, in the heat of the contest, they went beyond the real opinion of the originators of the school, for it does not appear from the writings of Werner,[1] the father of the Neptunist school, that he himself went further than to maintain that veins were filled by deposit from solutions reaching them from above, without attempting to indicate the source from which these solutions derived their metallic contents. The idea of the original ascensionists, or more properly, injectionists, that the mineral contents of ore deposits could have been injected into their present position in a fused state, is so opposed to all observed facts that it has long since been abandoned; and probably no one would maintain that original ore deposits are derived from waters at present flowing on the surface.

There still remains a tendency among writers to separate themselves into upholders of modifications of one or the other of these original theories, but an impartial examination of their views shows that, so far as their foundation in well-ascertained facts or on legitimate deductions from these facts goes, there is really no great essential difference between them. Thus the French geologists, who, by the prominence given to the synthetic experiments of Sénarmont, Daubrée, and others, may be considered to be the

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[1] A. G. Werner, Neue Theorie der Entstehung der Gänge. 1791.

legitimate inheritors of the ascension or injection theory, now modified into the sublimation theory, themselves admit that it is practically an aqueous solution, even if in a gaseous form, from which they derive the metallic contents of their deposits.[1] And among those who professedly maintain that most ores are probably deposited by percolating waters, but who would distinguish lateral secretion from ascension or descension, there is fundamentally so much held in common that differences seem slight.[2]

That *some* ore deposits have necessarily been deposited from solution is admitted by all geologists who have made special studies of the subject, and that the *greater part* of them have been so formed is maintained by a large and ever-increasing class. Geological investigations have also shown that within the rocks forming the crust of the earth, so far as observation has yet reached, there is a constant circulation of waters carrying more or less mineral matter in solution, and that no rock is absolutely impermeable. There are therefore both upward and downward currents, it being generally assumed that the latter are surface waters sinking under the influence of gravity and the former the same waters rising under that of the internal heat of the earth. It will be readily apparent, however, that such movement is not necessarily vertical in either direction, but will take its immediate direction from the character of the rock mass through which it is passing; that there will be a tendency of waters, filling capillary passages and minute fissures, to seek larger channels on joint, fault, and stratification planes, along which their movement will be more free; further, that, in case of waters passing along such channels and carrying mineral matter in solution, this mineral matter will be deposited where the conditions of the inclosing rock are such as to favor a chemical precipitation or interchange, and that such precipitation will be most abundant where for any cause there is some

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[1] A. de Lapparent, Traité de Géologie, p. 1170. Paris, 1883.

[2] Thus Joseph Le Conte, in his article on the Genesis of Ore Deposits (Amer. Jour. of Sci., May, 1883), while devoting much space to disproving the theory of Sandberger, as exposed in his recent Researches on Ore Deposits, is really of the same opinion as the latter on most essential points. The main point of difference between the two appears to be that, while Le Conte maintains that the phenomena of Steamboat Springs and Sulphur Bank, where deposits are actually going on at the present day, should be taken as a type of all deposits and may serve as a basis for a general theory, Sandberger considers them exceptional cases and their conditions not necessarily the same that prevailed with deposits formed at a great depth below the surface.

interruption in the regular flow of the current, as rapidly moving waters deposit much less readily than those whose movement is very slow.

Admitting the above conditions, it would seem, a priori, impossible to assign any general direction of movement to currents from which ores are deposited, and that each individual deposit must be studied by itself in order to determine, by its geological relations, from which direction the depositing solutions probably came. The admission that ore bodies have been deposited by currents of circulating waters logically involves the admission that they may have been upward, downward, or lateral currents, according as the conditions at time of deposition favored either direction of approach to the locus of deposit. While the determination of this direction in a special district is of the utmost importance from an economical point of view, since by it the explorations for the continuation of ore bodies must be largely guided, its theoretical importance as bearing upon the general question of the origin of ore deposits seems to have been hitherto much exaggerated.

Source of metals.—In speculations as to the source from which the metallic contents of ore deposits are derived, a distinction should also be made between the immediate and the ultimate source.

The ultimate source is as much a purely speculative matter as the nebular hypothesis. Since according to this hypothesis the earth in its present condition is the result of gradual cooling from an incandescent mass, and since moreover the specific gravity of the rocky crust which is exposed to observation is very much less than that of the whole mass of the earth, it is a legitimate conclusion that the heavy metals must be in much larger proportions in the interior of the earth than in the rocky crust. Although this view is generally admitted by geologists at the present day, it is evident that its basis is somewhat negative, since, like the nebular hypothesis upon which it is founded, it cannot be proved in the present state of science by actual experiment or observation. Volcanic emanations and thermal springs have been found to contain metallic minerals, as have also the waters of the ocean, but it cannot be stated definitely from what depth they have come in the former case, nor whether, in the latter case, they may not have been ultimately derived from the same indefinite, deep-seated source. It would seem proper, therefore, in a practical treatise like the present, to leave out of consideration altogether the ultimate and purely speculative source and to confine the investigation to the more immediate source, about which it is possible to obtain some actual and demonstrable evidence.

As circulating waters must take up as well as throw down their metallic contents, it is evident that under varying conditions the same material may have been deposited more than once and in more than one form since it reached that part of the rocky crust of the earth which is open to actual observation. There may be, therefore, intermediate sources between the ultimate and the immediate, but which, like the ultimate, are removed from actual demonstration.

It is common practice to say of any ore deposit, not distinctly sedimentary, that it has *come from below*, and to rest content with this statement, which, even if not susceptible of direct proof, has the merit that in one sense it cannot be disproved. This practice evidently had its origin in the fact that early writers upon ore deposits used as the type deposit, upon which to found their theories, the nearly vertical fissure vein. This they assumed to be the filling of a pre-existing open crack, extending indefinitely toward the center of the earth, by heated solutions arising from great depths; as these solutions approached the surface of the earth and were consequently relieved of the great pressure to which they were subject in the depths, by reason of that relief they gradually deposited their contents on the walls of the fissure until it was completely filled. While under the theoretical conditions assumed this hypothesis might afford an adequate explanation of the manner of deposition in such a vein, it is by no means proved that such conditions exist in nature, and therefore the explanation, so readily given in most cases, is generally inadequate and not founded upon a sufficient study of the geological conditions.

In the case of the Leadville deposits the inadequacy and even falsity of this explanation, except as applied to the ultimate source from which the metals may have been derived, is readily apparent.

In the first place, the geological study of the district has shown that they must have been formed beneath a thickness of at least ten thousand feet of superincumbent rocks and an unknown amount of sea water. If they had been deposited from hot ascending solutions, as the result of a relief of pressure, it would naturally be expected that the bulk of the deposit would have been found in the upper part of this mass of rocks, where the pressure was least, instead of at its base.

Secondly, as at the time of deposit the sedimentary beds in which they occur were horizontal and relatively undisturbed, if the deposit had been made from ascending currents it would naturally be expected that the process of deposition should have acted from the lower surface of the beds upwards, instead of from the upper surface downwards, as is shown to have been the case in the Blue Limestone, which carries the bulk of the ores.

Thirdly, as far as present investigations have extended, there is a noticeable absence, in the region of greatest ore development, of channels extending downwards, through which the ascending solutions might have come. The vast majority of eruptive bodies are in the form of nearly horizontal sheets, parallel with the stratification. The few approximately vertical bodies that have come under observation afford no evidence that their walls form part of a channel through which the ore currents came up from below.

The above considerations seem sufficiently conclusive evidence against adopting upward currents as the direct source of the ore deposits of Leadville. The principal water channel at the time of deposition was evidently the upper contact of the Blue Limestone with an overlying porphyry, and from this surface they penetrated downwards into the mass of the limestone. It may be assumed, therefore, that the currents were descending under the influence of gravity, rather than ascending under the influence of heat.

It is well known that percolating waters circulate freely in every direction through massive or eruptive rocks, owing to the effect which cooling and weathering have of splitting them into irregular blocks, while in sedimentary rocks, however permeable, the bedding-planes are naturally the easiest for them to follow. If, then, at the time of deposition the prevailing direction of the ore currents had been downwards, it is easy to conceive that they would have descended freely through the overlying porphyry masses and would have been diverted temporarily from a vertical to a horizontal course along the stratification plane of the first sedimentary bed they reached, and that, when this was a comparatively soluble rock like the dolomitic limestone of Leadville, they would eat their way gradually into it, either from this surface or from cracks through which they were here and there able to penetrate its mass. A downward current seems, therefore, to best suit the facts thus far observed with regard to the Leadville deposits. It might be objected that a downward current would not necessarily be hot, but it has been found by experiment and observation that metallic minerals may be taken up by cold water, though not so rapidly as by hot. Moreover, it is probable that the intrusive bodies retained for a long time sufficient heat to sensibly raise the temperature of waters coming in contact with them.

In looking, then, for the immediate source whence the waters by which these ores were deposited derived their metallic contents, which source should be within a limited distance of the locus of the deposits, since it cannot be supposed that the waters would travel for a great distance through rocks of varying composition without suffering considerable change in the material they held in solution, it would seem natural to consider the rocks in the vicinity of these deposits, and especially those overlying them.

Metallic contents of country rocks.—It was resolved in the early stages of the work to make a careful chemical examination of all the rock varieties of the region, as far as circumstances would admit, selecting comparatively unaltered rocks, which might be supposed to retain most of their metallic contents, and those sufficiently removed from any known ore channels to be free from the suspicion of having received these contents from the waters exuding from such channels.

This investigation was undertaken solely for the purpose of obtaining facts which might explain the condition of things existing in the region, and was conducted without the bias of any preconceived general theory. Indeed, in the opinion of the writer, our knowledge of the ore deposits of the world is still too limited and superficial to admit of the formulating of any generally and universally applicable theory. On the other hand, the weight of what may be considered actual evidence, as distinguished from pure speculation, seems to be in favor of the lateral-secretion theory in its broader acceptation. Geologists whose acknowledged ability and wide experience give weight to their opinion have already made the generalization that the majority of the ore deposits, whether in crystalline or in sedimentary rocks, are found, if not in actual contact with, at least in the immediate vicinity of, eruptive rocks. The experience of the writer would lead him to qualify this generalization by adding that it is with the older and generally intrusive rocks of eruptive origin that valuable ore bodies are most frequently associated, while they are rare in regions where these rocks only form surface flows or are outpourings of actual volcanic vents. As there is no ground for assuming that the latter rocks would be freer from heavy metals than the former, the reason for these associations would have to be found in the fact that the older rocks have been exposed longer to the action of percolating waters and the deeper rocks have been more accessible to the waters containing materials that would readily dissolve the metals.

As regards the derivation of ore materials from neighboring rocks, G. Bischof, who has rendered most important services to geology in removing it from a speculative to an inductive basis, first gave an authoritative and decided opinion in these words: [1]"As a general consequence of the relations between the matrices of lodes, the rocks adjoining them, and their condition, as well as those between different lodes, it may be inferred that all the substances contained in the lodes have been derived from the adjoining rocks."

Both Breithaupt[2] and von Cotta[3] admit the probability of this derivation, provided the existence of the vein materials in the country rock can be proved. Bischof had already proved this for the gangue materials, but his investigations had not been carried further, and the existence of the heavy metals in the country rocks still remained to be demonstrated. To this task Dr. F. Sandberger has devoted himself, as he tells us,[4] since 1873. Up to that time he had been an advocate of the ascension theory, but after having by careful analysis detected all the vein materials of a certain dis-

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[1] G. Bischof, Chemical and Physical Geology, III, p. 548.

[2] Paragenesis der Mineralien, p. 119. 1849.

[3] Lagerstätten der Erze. I, p. 177; II, pp. 297 et seq. 1859

[4] Untersuchungen über Erzgänge. Wiesbaden, 1882.

trict in the adjoining country rocks he was led to make extended investigations of the country rocks of ore deposits in general, not contenting himself with a simple lump analysis of the rock, but separating out its individual constituents and analyzing them separately. He now claims that he has thus been able to discover all the metals occurring in veins, and that they are mainly contained in the basic silicates of crystalline rocks, such as mica, hornblende, and augite. He has also analyzed the waters of many thermal springs, and concludes that the metals contained in these are not deposited in their channels, but only at the mouths, where they are practically in contact with the atmosphere, and, while he considers further similar investigations desirable, he holds that what he has already determined proves the general inapplicability of the thermal-spring origin for fissure veins, which have mostly been formed at depths where the influence of the atmosphere would not be felt. Whether Dr. Sandberger's conclusions be accepted in their entirety or not, the results of his investigations are certainly very suggestive. Both he and Bischof consider the silicates more probable sources of the metals than the disseminated pyrites so abundantly found in eruptive rocks, which they hold not to be original constituents thereof.

In the present investigation it was not feasible to follow Sandberger's method of analyzing the separate constituents of all the different rocks, which involves a great expenditure of time and the use of elaborate chemical apparatus. Moreover, the porphyries in the vicinity of Leadville contain no basic silicates in a sufficiently undecomposed state to be separated out. Lump analyses alone were then practicable, but by the employment of dry methods it was possible to make a greater number of tests and detect extremely minute traces of silver and gold, which by Sandberger's method could hardly have been found. Lead and barium were also sought for in the wet way. The other principal constituents of the ores, silica, iron, and manganese, are so universally disseminated that a special search was considered unnecessary. The methods pursued and the details of the results obtained are set forth in Appendix B. More than twice the number of assays there given for gold and silver were originally made, but after they had all been completed a possible source of error was discovered, which led to a repetition of the test in the case of all for which material remained.

The discrepancies in these were not such as to have affected the final conclusion, but it was considered best to publish only those whose positive results were beyond a doubt.[1] Comparatively few sedimentary rocks were tested, for the reason that in all cases where not evidently exposed to the influence of the ore currents they were found to be barren of all the essential vein materials.

Baryta determinations.—Baryta forms an essential constituent of one variety of feldspar (hyalophane), which contains from 9 to 20 per cent. of it, and it has also been found in small amounts in andesine, oligoclase, and orthoclase feldspars, where it is generally associated with strontia; probably both might have been found much more frequently if analysts had made a special search for them. As these substances are generally recognized as possible constituents of eruptive rocks, it was not considered necessary to make a great many tests in order to establish the possibility that the barite in the Leadville deposits might have been derived from the neighboring rocks. As strontium replaces barium to a certain extent in the Leadville mineral, its occurrence is also of significance. Neither of these substances was found in the sedimentary rocks analyzed, but from eruptive rocks were obtained as follows:

Barium and strontium in eruptive rocks.

Rock.	Locality.	BaO	SrO
White Porphyry	California gulch quarry	0.03	Trace
Pyritiferous Porphyry	White's Hill	0.098	
Do	do	Trace	
Do	White's Hill, Malvina tunnel	Trace	
Gray Porphyry	Near Onota, Johnson gulch		0.08
Lincoln Porphyry	Summit of Mount Lincoln		0.07
Porphyrite	North Mosquito amphitheater		Trace
Hypersthene-andesite	Buffalo Peaks	Trace	Trace

In one specimen each of Mount Zion Porphyry, porphyrite, and rhyolite, neither baryta nor strontia was found. Thus out of eleven specimens tested five contained baryta, five strontia, and in three neither was detected.

Lead determinations.—Lead has also been detected in feldspars, though less frequently than baryta.

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[1] For this reason the averages vary slightly from those assumed at the time the abstract of this report was written.

Among the rocks tested during this investigation it was found in a single specimen only of sedimentary rock, a sandstone impregnated with pyrite, probably derived from the adjoining Pyritiferous Porphyry. As the quantitative results are given in full in Table III, Appendix B, they need only be presented in the following abbreviated form here:

Lead in eruptive rocks.

Rock.	Number of specimens tested for PbO.	Number containing PbO.	Number in which no PbO was found.
White Porphyry	2	2	0
Gray Porphyry	1	1	0
Lincoln Porphyry	3	2	1
Pyritiferous Porphyry	8	8	0
Porphyrite	2	1	1
Granite	2	1	1
Total	18	15	3

Of the above specimens all except granite and porphyrite belong to a higher geological horizon than the Blue Limestone. The results given in Table III are those which were obtained from that portion of the rock soluble in strong acids. In only three cases was the insoluble part fused and subjected to further treatment. In these cases more lead was obtained from the insoluble part than from the soluble. It may therefore be assumed that had the same treatment been pursued in each case the results would have been even more conclusive as to the prevalence of lead in appreciable quantities in the igneous rocks of the region. The greater portion obtained in the second treatment existed undoubtedly in the form of silicate, since most of the sulphide contained in the rock, whether original or secondary, would have yielded to the acid treatment. But it does not necessarily follow that all the lead obtained by the first treatment was associated with the pyrite, for the analysis of pyrite from the Pyritiferous Porphyry treated separately gives only 0.0019 per cent. PbO, while the average of eight specimens of the rock gives 0.0020 in the soluble portion alone.

As a further illustration, Table III gives the result of an examination for zinc of two rocks from the Ten-Mile district, where the ores are highly zinciferous. They are 0.008 and 0.0043 per cent. ZnO, respectively, with an appreciable amount of CoO as well. Of the granites, the one containing lead is of eruptive type, the other rather a granite gneiss.

Had time and laboratory facilities permitted, these results might have been greatly multiplied and the analyses made much more exhaustive; but, although the existence of lead in such connection inferentially implies that of gold or silver also, it was thought wiser to obtain direct proof of the occurrence of these metals as well.

Silver and gold determinations.—The quantitative results of these determinations, together with a statement of the methods employed, will be found in Appendix B. The following table presents the results given in Table IV in a condensed and more comprehensive form:

Silver and gold in eruptive rocks.

Rock.	Number of specimens tested.	Number containing silver.	Number containing gold.	Number in which neither was found.
White Porphyry	11	3	(?)	8
Gray Porphyry	3	3	1	0
Lincoln Porphyry	6	5	0	1
Pyritiferous Porphyry	10	9	3	1
Sacramento Porphyry	1	1	0	0
Green Porphyry	1	1	0	0
Diorite (augite)	1	1	0	0
Porphyrite	6	6	0	0
Andesite (hornblendic)	1	1	0	0
Rhyolite	1	1	0	0
Trachyte	1	1	0	0
Granite	2	0	0	2
Total	44	32	4	12

In discussing the above results it is important to consider the present condition, position, and composition of the different rock masses. It must be borne in mind also that the negative results are not absolute, but mean merely that the specimen tested does not contain more than 0.0000068 per cent. of silver, or 0.002 ounce to the ton. Such quantities are, it is true, almost infinitely small; so also is the amount of time and water allowed for their leaching almost infinitely great.

The White Porphyry is the most universally decomposed rock in the region. Nowhere was it possible to obtain it in an absolutely fresh state, but the rock of dikes was found to most nearly approach this condition. The three specimens which were found to contain silver came from dikes, while all the others came from the main sheet overlying the Blue Limestone. Considering this fact alone, it might be assumed that the metallic contents had already been leached out of this sheet. On the other hand, the rock in its unaltered, normal condition apparently contained a very small proportion of basic silicate; and of pyrite, if it existed as an original constituent, but little trace is left.

On the other hand, the Pyritiferous Porphyry, which stands at the opposite end of the scale as regards its contents in gold and silver, although generally decomposed at the surface, is less so than the White Porphyry, and its interior is less deeply exposed, either by erosion or by underground workings. It also contains a larger proportion of basic silicates. Its most striking feature is the enormous amount of pyrite that it contains, amounting, on an average, to about 4 per cent. of its mass. Part, at least, of this pyrite is original, as it is found included within the crystals of quartz. Both pyrite and galena are occasionally found, however, coating the jointing planes of the rock, in which case they are undoubtedly secondary.

From a consideration of the quantitative results given for this rock in Table IV, it is evident that, while the traces of gold in the rock might have been contained in the pyrite, all the silver could not thus be accounted for. The average assay of the ten specimens of Pyritiferous Porphyry is 0.2773 ounces Ag per ton; but one of these might be considered abnormal, since it alone contains more than the sum of the other nine. Rejecting this, the average of the other nine specimens is 0.0265 ounce Ag. The pyrite, separated and assayed alone, gave 0.390 ounce Ag, or 0.00134 per cent.; but in a rock containing 4 per cent. of such pyrite, which was the estimate obtained by a careful mechanical separation of this material from a Pyritiferous Porphyry of average composition, there would only be 0.0156 ounce Ag per ton, or less than three-fifths of the above average of nine specimens.

A mixture of galena and pyrite, also separated from the rock and assayed by itself, gave 2.4 ounces Ag to the ton, or 0.00823 per cent. From Table III it is found that the eight specimens of Pyritiferous Porphyry tested have an average of 0.002025 per cent PbO in the soluble portion, or what may be assumed to have been in the form of sulphide. This would correspond to 0.002277 per cent. of impure galena, assuming impure galena to bear the relation to PbO of eight to nine. If this galena carries 0.00823 per cent. Ag, as above, $0.002277 \times 0.00823 = 0.0000177$ per cent. Ag, or 0.0051 ounce to the ton, would be the amount it contributed to the total silver contents of the rock. If this be added to the amount to be derived from pyrite, $0.0156 + 0.0051 = 0.0207$ ounce, it is still less than the average, 0.0265 ounce, given for the above average rock.

But the tests for lead show that a considerable portion is contained in the silicates (in the three specimens in which this test was made, about three-fifths of the whole amount); and, if the silver is assumed to be necessarily associated with the lead in the rock, this would amply account for the remaining 0.0058 ounce.

Of the other porphyries the most significant are the Lincoln and the Gray, which, as has already been shown, are practically the same type of rock. Both have a much larger proportion of basic silicates than the White Porphyry. Of the two the Gray is to outward appearance the more decomposed, but in the Lincoln Porphyry microscopical examination shows that alteration of the basic silicates has already set in, and it is probable that the more decomposed appearance of the Gray Porphyry is due to the action of surface waters on the other constituents, mainly the feldspars. It is noticeable that the Lincoln Porphyry from Clinton gulch, which contains the most silver, is the only one which contains pyrite; also, that the others are from a region where there has been a considerable concentration of metals in ore deposits, which is not the case, so far as known, in Clinton gulch.

With the exception of the Sacramento Porphyry the other rocks have no apparent association with important ore deposits. It is significent that the diorite given in Table IV contains augite, hornblende, and mica, whereas two other diorites assayed, and in which no silver was found, contained a very small proportion of basic silicates. In the recent eruptive rocks there is also an apparent relation between the amount of basic minerals and the contents in silver. In Nevadite, in which they are almost entirely wanting, no silver was found, while the trachyte and andesite, which contain more of these minerals than the Black Hill rhyolite, also contain more silver.

Although the above facts are not sufficiently conclusive to afford absolute proof that the metallic contents of the deposits were entirely derived from the eruptive rocks, they certainly show the possibility and even probability that this source furnished a part at least of the vein materials.

The actual percentage of metals found may seem very small; on the other hand, it should be remembered that the amount of time and of water allowable for the leaching process may have been almost indefinitely large. The present porphyry bodies, moreover, are of enormous extent as compared with the actual size of the deposits, while the amount of porphyry that has been removed by erosion since the deposits were first made, though it cannot be accurately estimated, must have been even larger.

Possible contents of porphyry bodies.—In order to show that even with the small percentages given in the above table the possible contents of the porphyry bodies are amply adequate to account for the amount of ore thus far developed in the district, a hypothetical calculation will be made based on these percentages and on the probable bulk of one of the porphyry bodies, taking first the amount assumed to exist now and second a conservative estimate of the amount which existed at the time of original ore deposition, and before any of it had been removed by erosion. For this purpose the Pyritiferous Porphyry will be chosen, since a greater number of tests of this rock have been made than of any other.

The present area of outcrop of the Pyritiferous Porphyry may be taken, in round numbers, as 5,000 × 10,000 feet = 50,000,000 square feet. If it is assumed that it originally extended westward to the foot of Carbonate Hill, north to the line of Yankee Hill, south to that of Printer Boy Hill, and but little beyond its present boundaries to the eastward, it would have covered a square area of 10,000 × 20,000 = 200,000,000 square feet, or four times the assumed area of its present outcrop. The specific gravity of Pyritiferous Porphyry, obtained as an average of four specimens, is 2.608. Therefore one ton (2,000 pounds) of this rock would occupy 12.27 cubic feet; say $12\frac{1}{2}$ for convenience of calculation.

This porphyry is assumed to contain 4 per cent. of pyrite ($FeS_2=4.00$). From Table III, Appendix B, its contents of protoxide of lead, as an average of the eight specimens tested, is 0.002025 per cent. in the soluble portion; or, assuming, from the proportion found in the insoluble portion of the three specimens in which it was tested, that this represents only two-fifths of the entire lead contents, the average contents of the whole rock would be $PbO=0.0050625$ per cent. From Table IV the average of ten specimens assayed for silver is found to be 0.2773 ounce per ton, or, rejecting the richest of these ten specimens as above the normal, the average of the remaining nine specimens is 0.0265 ounce silver per ton of Pyritiferous Porphyry.

The probable thickness of the porphyry sheets it is rather difficult to estimate. The sections as drawn give a maximum thickness of about fifteen hundred feet, and an unknown thickness has been eroded away. It may not be unreasonable to assume 1,000 feet as the average thickness of the original body. From the above-assumed data would be obtained, as the contents of the present and original areas of Pyritiferous Porphyry, respectively, and on the basis of the two different values for lead and silver given above, the following:

Contents of Pyritiferous Porphyry.		Designation.	In area of present outcrop.	In area of assumed original body.
Amount of porphyry		Tons	4,000,000,000	16,000,000,000
Amount of pyrite, at 4 per cent		Tons	160,000,000	640,000,000
Amount of galena	at 0.002025 per cent. PbO	Tons	8,675,000	34,700,000
	at 0.0050625 per cent. PbO	Tons	21,687,750	86,751,000
Amount of silver	at 0.0265 ounce per ton	Ounces	106,000,000	424,000,000
	at 0.2775 ounce per ton	Ounces	1,109,000,000	4,436,000,000

To obtain an actual average of the metallic contents of this or any other body of porphyry would have required a systematic sampling of the rock and the taking of specimens at given and equal distances, not only on its surface but through its mass in depth, for the tests already made show that the metals are not evenly distributed, but vary in an apparently arbitrary manner. Such a sampling is manifestly not practicable, nor would the expenditure of labor and time required by it be advisable if it were, since in the present state of explorations in this region it is impossible to trace with any degree of certainty the processes of original ore deposition. The most that could be hoped for was to indicate the possible methods by which the deposition might have taken place and to weigh the probabilities afforded by ascertained facts in favor of one or the other of these methods. The foregoing reasons seem to favor the probability that the ores may have been derived, in part at least, from one or more of the bodies of porphyry which occur in the region, and the above figures show that the small percentages of the metals still existing in these rocks might furnish an adequate amount of material to form the known ore bodies.

The most uncertain element in all these calculations and hypotheses is the form and extent of the porphyry bodies in depth beyond the limit of present explorations. The form given to these bodies in the sections is, as has already been stated, only hypothetical, though founded on deductions from many actual observations and in all probability correct in its main outlines. Still it is probable that there are more vents or channels from below through which the bodies have reached their present position than are shown there, but their number and position can only be determined by actual exploration. It is possible that in future years, when mine workings shall have been extended over areas where the ore horizon exists at considerable depths below the surface and other eruptive channels have been found and critically examined, evidence may be obtained that ore solutions have ascended along these channels from below. Such evidence will not, however, necessarily preclude the derivation of part of the metals from the country rocks, and at present that derivation is the only one which has the support of actual though somewhat indirect proof.

Another element of uncertainty, and one which renders it difficult to decide from what particular variety of porphyry the metals of the deposits were derived, is the impossibility of determining the form and character of the porphyry bodies, which have been removed by erosion, as they existed at the time of original ore deposition, and upon this point future exploration will throw little or no light.

APPENDIX B

CHEMISTRY

BY

W. F. HILLEBRAND

CONTENTS.

ERUPTIVE ROCKS.

LIMESTONES.

ORES AND VEIN MATERIALS.

TABLES OF ANALYSES AND NOTES ON METHODS EMPLOYED.

ERUPTIVE ROCKS.

TABLE I.—*Complete analyses of eruptive rocks and constituent minerals.*

[The Coll. No. is the number given to the specimen in the collection of Leadville rocks deposited in the Nationa Museum.]

	1	2	3	4	5	6	7	8	9	10	11	12	13	14
Sp. grav	2.680	......	2.629	2.670	2.636	......	2.768	2.740	......	......	2.742	......	......	3.307
At temp. of.	16° C.	......	26° C.	16° C.	16° C.	......	16° C.	16° C.	......	......	16° C.	......	......	23° C.
SiO_2	70.74	45.03	73.50	66.45	68.10	62.22	56.62	64.81	74.45	65.04	56.19	51.70	51.16	50.04
TiO_2	......	......	......	0.10	0.07	......	......	0.08	......	......	......	......	......	......
Al_2O_3	14.68	38.14	14.87	15.84	14.97	*a*20.33	16.74	15.73	14.72	*b*20.40	16.12	1.72	2.15	2.91
Fe_2O_3	0.69		0.95	2.59	2.78		4.94	1.67	......	......	4.92	0.30	......	......
FeO	0.58	......	0.42	1.43	1.10	......	3.27	2.91	0.56	......	4.43	18.00	*c*18.36	*d*17.81
MnO	0.06	......	0.03	0.09	0.09	......	0.15	0.08	*e*0.28	......	Trace	0.36	0.36	0.12
CaO	4.12	Trace	2.14	2.90	3.04	2.95	7.39	4.22	0.83	0.79	6.99	2.87	3.81	6.70
BaO	0.03	......	......	......	......	......	......	......	......	......	Trace	......	......	......
SrO	Trace	......	......	0.07	0.08	......	......	Trace	......	......	Trace	......	......	......
MgO	0.28	(*f*)	0.29	1.21	1.10	......	4.08	2.82	0.37	......	4.60	25.09	24.25	21.74
K_2O	2.59	9.44	3.56	2.89	2.93	8.31	1.97	1.43	4.53	9.74	2.37	(*g*)	(*g*)	......
Na_2O	2.29	0.71	3.46	3.92	3.46	3.45	3.50	3.98	3.97	4.11	2.96	(*g*)	(*g*)	0.27
Li_2O	(?)	......	......	Trace	......	Trace	(?)	FeS_2 0.90	Trace	Trace	......	(*g*)	(*g*)	......
H_2O	2.09	4.06	0.90	0.84	1.28	*h*1.90	0.92	0.62	0.66	0.29	1.03	(*g*)	(*g*)	......
CO_2	2.14	......	......	1.35	0.92	......	1.15	1.08	......	......	......	......	......	......
P_2O_5	......	......	......	0.36	0.16	......	Trace	0.23	0.01	......	0.27	......	......	......
Cl	Trace	......	......	0.05	0.03	......	......	0.03	......	......	0.02	......	......	......
Totals ...	100.29	......	100.12	100.09	100.11	99.16	100.73	100.59	100.38	100.37	99.90	100.64	100.09	99.59

a Very little Fe_2O_3 present; FeO not determined.
b Includes traces of Fe_2O_3.
c Includes any Fe_2O_3 present.
d Calculated from the Fe_2O_3 found.
e MnO_2.
f A little MgO present.
g Water and alkalies not tested for.
h By ignition; want of material prevented a second determination.

1. White Porphyry. Coll. No. 27p. Quarry in California Gulch, southwest slope of Iron Hill, near Leadville. Composed chiefly of quartz orthoclase, and plagioclase; no biotite or hornblende; little magnetite; no apatite; feldspars attacked; muscovite and calcite secondary.

2. Muscovite crystals from White Porphyry. Sheet on south slope of Little Zion. Analysis only approximate, being made to prove the mica to be muscovite.

3. Mount Zion Porphyry. From Little Harry shaft, Prospect Mountain.

4. Lincoln Porphyry. Coll. No. 75. Summit of Mount Lincoln, Park County. Rock nearly fresh; contains quartz, orthoclase (large crystals), plagioclase, biotite; biotite partly changed to chlorite, containing rutile (?) needles; some magnetite and apatite; calcite.

5. Gray Porphyry. Coll. No. 59a. Near Onota claim, Johnson Gulch, Leadville. Rock nearly fresh; contains quartz, orthoclase (large crystals), plagioclase, and biotite. Details same as for 4.

6. Large pink crystals of feldspar from Gray Porphyry (No. 5). Coll. No. 59a.

7. Porphyrite. Coll. No. 126 (Type V of table, see Appendix A). Near the Northern Light mine, Lower Buckskin gulch, Park County; intrusive sheet. Contains abundant plagioclase and hornblende, and little biotite; quartz and orthoclase (?) in groundmass. Magnetite and little apatite.

8. Porphyrite. Coll. No. 260. Type VI of table (see Appendix A). Head of North Fork of Mosquito gulch, Park County. Dike in Archean. Fresh. Much biotite and plagioclase; some quartz in groundmass. Very little hornblende, magnetite, pyrite, apatite.

9. Nevadite. Coll. No. 397. Northeast point of Chalk Mountain, Ten-Mile district. Contains sanidine, plagioclase, and quartz in abundance; little biotite and magnetite.

10. Sanidine from Nevadite. Coll. No. 136. Southern edge of Chalk Mountain.

11. Hypersthene-andesite. Coll. No. 144. Buffalo Peaks, Park County. Very fresh. Plagioclase and hypersthene abundant; considerable augite, magnetite, apatite; glass base.

12. Hypersthene from hypersthene-andesite. Coll. No. 144 (see 11). The hypersthene for this and the two following analyses was separated from the other constituents of the rock by hydrofluoric acid, as recommended by Fouqué in Santorin et ses éruptions.

13. Same as 12, but of a different sample.

14. Hypersthene from hypersthene-andesite. Coll. No. 150. Buffalo Peaks, Park County. Particles of undecomposed feldspar were visible attached to the crystals and fragments of hypersthene analyzed.

TABLE II.—*Silica and alkali determinations.*

No.	Coll. No.	Rock.	Variety.	Locality.	SiO_2	K_2O	Na_2O	Li_2O
1	......	Porphyry	Mount Zion	Little Harry shaft, Prospect Mountain	73.50	3.56	3.46	None
*a*2	27p	do	White	Quarry, California gulch	70.74	2.59	2.29	......
3	230	do	do	Above Blue Limestone, Dyer Mountain.	66.51	......	......	......
*a*4	75	do	Lincoln	Summit of Mount Lincoln	66.45	2.89	3.91	Trace
5	78	do	do	Dike in eastern spur of Mount Lincoln	64.16	2.42	3.15	......
6	74	do	do	El Paso mine, Clinton gulch	69.00	......	......	......
*a*7	59a	do	Gray	Near Onota Claim, Johnson gulch	68.10	2.93	3.45	None
*b*8	......	do	Pyritiferous	White's gulch	......	4.62	2.91	......
9	103	do	Green	South head of Mosquito gulch	65.05	......	......	......
10	98	do	do	Between Buckskin and Mosquito gulches.	63.85	......	......	......
11	95a	do	Mosquito	North Mosquito amphitheater	68.01	......	......	......
*b*12	......	do	do	North face of Mount Lincoln	......	4.36	4.26	......
13	85a	do	Sacramento	North of East Leadville	65.08	2.57	3.55	......
14	109a	do	Silverheels	Summit of eastern spur of Mount Silverheels.	60.42	......	......	......
*b*15	......	do	do	Mount Silverheels	......	2.70	4.08	......
*a*16	126	Porphyrite	Hornblende	South wall of Buckskin gulch	56.62	1.97	3.50	......
17	131	do	do	Arkansas amphitheater	57.33	......	......	......
18	......	do	Hornblende and biotite.	Ten-Mile amphitheater	57.76	......	......	......
19	132	do	Hornblende	North Mosquito amphitheater	54.54	......	......	......
20	121	do	do	Buckskin amphitheater	65.73	......	......	......
*a*21	260	do	Biotite	North Mosquito amphitheater	64.80	1.43	3.98	None
22	130	do	do	Arkansas amphitheater	66.29	......	......	......
23	129	do	do	do	59.26	......	......	......
24	137	Rhyolite	Nevadite	Chalk Mountain, eastern edge	71.44	......	......	......
*a*25	397	do	do	Northeast point of Chalk Mountain	74.45	4.53	3.97	Trace
26	140	do	Liparite	Southern end of Black Hill	69.54	......	......	......
27	268	do	do	South bank of Empire gulch	68.05	3.50	2.17	......
28	......	do	do	Near Granite, Chaffee County	76.84	......	......	......
29	269a	do	Showing quartz	Little Ellen shaft, McNulty gulch	65.75	......	......	......
30	269a	do	Compact	do	65.21	......	......	......
31	......	do	Tufa	Five miles below Salt Works, South Park.	70.30	......	......	......
32	142	Trachyte	Quartziferous	Head of Little Union gulch	61.22	......	......	......
33	143	Andesite	Hornblende	Buffalo Peaks, northwest Peak	57.60	......	......	......
34	149	do	Dacite	Buffalo Peaks, central amphitheater	66.50	2.57	3.87	None
*a*35	144	do	Hypersthene	Buffalo Peaks, northeast spur	56.19	2.37	2.96	None
36	150	do	do	Buffalo Peaks, base of middle peak	60.36	......	......	......

a Complete analysis in Table I. *b* Analyzed by L. G. Eakins.

NOTE.—Blanks in the above table denote simply that no tests were made for the substance indicated.

TABLE III.—*Determinations of lead, zinc, cobalt, and barium, chiefly in eruptive rocks.*

No.	Coll. No.	Rock.	Variety.	Locality.	PbO in soluble portion.	PbO in insoluble portion.	ZnO	CoO	BaO
*a*1	27p	Porphyry	White	Quarry, California gulch	0.0030	......	......	......	0.08
2	54b	do	do	Dike in Dolly Varden mine, Mount Bross.	0.0028	......	......	......	......
*a*3	75	do	Lincoln	Summit of Mount Lincoln	None.	......	......	......	......
4	59a	do	Gray	Onota mine, Johnson gulch	0.0024	0.0034	......	......	......
5	......	do	Eagle River.	From surface at El Capitan mine, Tennessee Pass.	Trace.	......	......	......	......
6	56	do	do	From shaft of El Capitan mine, Tennessee Pass.	Trace.	......	......	......	......
7	......	Limestone	Blue	Twelve feet below ore body, El Capitan mine.	None.	......	......	......	......
8	130	Porphyrite	Biotite	Arkansas amphitheater	None.	......	......	......	......
9	124	do	Hornblende and biotite.	Bartlett Mountain	0.0006	......	......	......	......
10	326	Porphyry	Eagle River.	Main fork of Eagle River	......	......	0.0080	0.0008	......
11	269a	Rhyolite	......	Little Ellen shaft, McNulty gulch	......	......	0.0043	0.0010	......
12	90a	Porphyry	Pyritiferous	Hartford mine, Breece Hill	0.0053	......	......	......	......
13	90c	do	do	White's Hill, west of Pilot fault	0.0013	......	......	......	......
14	87	do	do	White's Hill, between Printer Girl and Golden Edge.	0.0013	0.0029	......	......	......
15	......	do	do	White's Hill	......	......	......	......	0.098
16	93a	do	do	do	......	......	......	......	Trace
17	......	do	do	White's Hill, Melvina tunnel	Trace.	......	......	......	Trace
18	......	do	do	Head of White's gulch	0.0030	0.0034	......	......	......
19	......	do	do	Printer Boy Hill	0.0006	......	......	......	......
20	94a	do	do	Rebel Warrior mine, Ball Mountain	0.0030	......	......	......	......
21	94	do	do	Wednesday tunnel, Ball Mountain	0.0017	......	......	......	......
22	288	Sandstone	do	Snow Bird claim, head California gulch	0.0026	......	......	......	......
23	120	Pyrite	From Pyritiferous Porphyry.	Lalla Rookh mine, Breece Hill	0.0019	......	......	......	......
24	17	Granite	Archean	Garden City shaft, California gulch	None.	......	......	......	......
25	217	do	do	Northwest slope of Mosquito Peak	0.0008	......	......	......	......

a Complete analysis in Table I.

N. B.—The blank spaces under the headings PbO, etc., do not indicate absence of the respective oxides. Where no results are given, no tests were made.

REMARKS ON TABLES I, II, AND III.

Insoluble silicates were decomposed by fusing with alkaline carbonates for the determination of silica, titanic acid, and all bases except the alkalies.

Ferric oxide and alumina were separated either by pure potassium hydrate, or more generally by ammonium sulphide, after addition of tartaric acid and ammonia.

For the determination of ferrous oxide, treatment with sulphuric acid in sealed tubes at about 200° C. was employed in cases where complete decomposition of the silicates could thus be effected, and the solution titrated with potassium permanganate. For the decomposition of refractory silicates, pure hydrofluoric acid, distilled from a platinum retort, was employed, the solution being effected in platinum vessels with careful exclusion of air. The iron was then determined as above.

Barium and strontium were looked for in the precipitated calcium oxalate after ignition, and estimated by the method given by Bunsen in his treatise on Mineral Water Analyses; the purity of the resulting compounds of these elements was ascertained by means of the spectroscope.

When a rock was examined merely to ascertain the presence or absence of barium, a considerable portion (10 grams) was decomposed with hydrofluoric and sulphuric acids, the soluble salts were extracted with water after expulsion of the hydrofluoric and excess of sulphuric acids, the residue was fused with sodium carbonate, extracted with water, and the insoluble part collected on a filter. After solution in hydrochloric acid, the barium, if present, was thrown down with sulphuric acid, the precipitate ignited and weighed, and after decomposition with sodium carbonate, tested spectroscopically.[1]

For the estimation of the alkalies, decomposition was effected in the earlier analyses by hydrofluoric and sulphuric acids; in the later by heating in a platinum crucible with calcium carbonate and ammonium chloride. The potassium was thrown down, after weighing the mixed chlorides, as potassium-platinic chloride, and calculated from the weight of the latter, the sodium being found by difference. Lithium could never be detected spectroscopically in the potassium-platinic chloride, but occasionally in the sodium salt.

Chlorine was determined by fusing with alkaline carbonate, extracting with water, acidifying the filtrate with nitric acid, and precipitating with silver nitrate.

Phosphorus pentoxide was always determined in a separate portion of the powder, and water by ignition in a hard glass tube and absorption in a weighed calcium chloride tube. The loss in weight by treatment with acid in a suitable apparatus gave the carbon dioxide.

For the detection and estimation of lead, large quantities (30–50 grams) were employed. Pyrite and other soluble salts were first extracted with nitric or nitrohydrochloric acid; the filtrate, together with copious washings, evaporated nearly to dryness several times with nitric acid; the residue digested with dilute nitric acid, and the solution and undissolved matter separated by filtration. As the insoluble part might contain a trace of lead sulphate, a warm ammoniacal solution of ammonium tartrate was passed repeatedly through the filter, and to the filtrate ammonium sulphide added. Through the previous nitric acid solution a strong current of hydrogen sulphide gas was passed for a considerable length of time, the precipitate, mainly sulphur from reduction of iron salts, collected on a filter, well washed, dried, and ignited gently with the filter paper to volatilize the sulphur. A few drops of nitric acid were then added, and heat was applied to dissolve the lead, mostly reduced to the metallic state by the carbon of the filter paper. The solution was filtered onto a watch-glass, and to this was added the nitric-acid solution of any lead sulphide that might

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[1] Later investigation by the writer seems to indicate that baryta may be a far more frequent constituent of eruptive rocks than has hitherto been supposed. The failure to detect it in the ignited calcium oxalate, where it is usually looked for, cannot be regarded as a proof of its absence from the eruptive rock examined. Experience in a number of cases has shown that where baryta and lime are in solution in as high a proportion as 1 of the former to 4 or 5 of the latter, in presence of considerable ammonium chloride, an almost complete separation of the two is effected by double precipitation of the lime by ammonium oxalate. The solubility of barium oxalate appears to be increased by the presence of magnesium salts. An ordinary spectroscope repeatedly failed to show the faintest evidence of barium in the ignited calcium oxalate. This subject will be more fully investigated. (W. F. H.)

have appeared in the ammonium tartrate solution above mentioned. As a trace of lead sulphate might have been formed by the ignition of the precipitate by hydrogen sulphide and have escaped solution in the nitric acid added, the residue on the filter was exhausted with ammonium tartrate and tested with ammonium sulphide. The contents of the watch-glass were then evaporated with two or three drops of sulphuric acid, and finally gently heated to expel nitric acid. If lead was present it could now invariably be seen at the center as white powder. This was collected on the smallest possible filter, washed with alcohol, dried, ignited, and weighed as sulphate. The latter was then scraped as far as possible on charcoal, and carefully reduced with a very little soda. The yellow coating of lead oxide was invariably formed, and in the soda appeared minute metallic buttons, malleable and soluble in nitric acid. The solution, concentrated to a drop or two, showed a bluish-black precipitate with hydrogen sulphide.

The portion of the rock insoluble in nitric or nitrohydrochloric acids, composed of quartz and silicates, was decomposed with hydrofluoric and sulphuric acids purified by distillation from a platinum retort, and dissolved in slightly acidified water, after expulsion of the hydrofluoric and excess of sulphuric acids. Solution and possible residue were then treated as in the foregoing for the separation and estimation of lead.

For the estimation of zinc and cobalt, large quantities (30 grains) were taken and decomposition was effected by hydrofluoric acid. After evaporating with sulphuric acid and igniting to expel the excess of the latter, solution was effected in hot water slightly acidified; the solution saturated with hydrogen-sulphide; from the filtrate, after oxidation, alumina and iron thrown down by ammonia; the precipitate redissolved in hydrochloric acid after filtration, and reprecipitated. This being repeated once more, the combined filtrates were evaporated to a moderate volume; the alumina still in solution was thrown down while boiling by ammonia, and this precipitate redissolved and reprecipitated. To the again combined filtrates ammonium sulphide was added to throw down zinc, manganese, cobalt, and nickel, if present; the precipitate was treated on the filter with a mixture of one part hydrochloric acid of 1.12 sp. gr. and six parts solution of hydrogen sulphide. The zinc and manganese in solution were thrown down again by ammonium sulphide, the manganese (being present in very small quantity) extracted by dilute acetic acid, while the zinc sulphide on the filter was then brought into a weighed platinum crucible by means of hydrochloric acid, evaporated to dryness, and ignited with mercuric oxide in the manner recommended by Volhard. The oxide, after weighing, gave the characteristic green coloration on igniting with cobalt nitrate. The cobalt sulphide left on the filter after extraction of zinc and manganese was ignited with the filter, digested with nitrohydrochloric acid; the solution rendered alkaline with ammonia; ammonium carbonate added; the slight precipitate separated by filtration and the cobalt thrown down by potassium hydrate. The ignited oxide tested by the method of Jorissen showed no trace of nickel.

TABLE IV.—*Gold and silver determinations.*

No.	Coll. No.	Formation.	Locality.	Silver, percentage.	Silver, ounces per ton.	Gold, ounces per ton.
1	53	White Porphyry	Buckskin amphitheater, débris slope	0. 000, 024, 0	0. 007	None
2	44a	do	"Horseshoe," south wall		None	Do.
3	45a	do	South base of White Ridge		do	Do.
4	45b	do	Summit of White Ridge		do	Do.
5		do	Four-Mile gulch		do	Do.
6	44	do	Head Four-Mile gulch		do	Do.
7	46	do	West end of Lamb Mountain		do	Do.
8		do	Northwest slope of Sheep Mountain		do	Do.
9	47	do	do		do	Do.
10	54a	do	Dolly Varden mine, west dike	0. 001, 269	0. 37	(?)
11	54b	do	Dolly Varden mine, east dike	0. 000, 481	0. 14	(?)
12	72	Lincoln Porphyry	Arkansas Valley, east side, above Howland		None	None
13	74	do	Clinton gulch	0. 000, 155	0. 045	Do.
14	75a	do	Mount Lincoln summit	0. 000, 034	0. 01	Do.
15	75a	do	Mount Lincoln summit, different specimen from above.	0. 000, 024	0. 007	Do.
16	78	do	Mount Lincoln, dike on south face	0. 000, 017, 1	0. 005	Do.
17	251a	do	East spur Mount Lincoln		Trace	Do.
18	59a	Gray Porphyry	Onota shaft, Johnson gulch	0. 000, 034, 3	0. 01	Do.
[a]19	59a	do	North of Onota shaft	[a]0. 000, 137	0. 040	Trace
20	60c	do	Lickscumdidrix bore-hole, Little Stray Horse Park.	0. 000, 137	0. 040	None
21	84	Sacramento Porphyry.	Between Sacramento and Pennsylvania gulches.	0. 000, 034, 3	0. 01	Do.
22	98	Green Porphyry	Mosquito gulch, north wall	0. 000, 017, 1	0. 005	Do.
23	90h	Pyritiferous Porphyry	Shaft west of Tribune, Breece Hill	0. 008, 68	2. 533	Do.
24	90e	do	Hartford shaft, Breece Hill	0. 000, 275	0. 08	Trace
25	88	do	Montana shaft, Breece Hill	0. 000, 045	0. 013	Do.
26		Galena and pyrite	South of Ace of Hearts, Breece Hill	0. 008, 23	2. 400	0. 033
27	249	Pyrite	Shaft west of Ohio Bonanza, Breece Hill	0. 001, 34	0. 390	0. 130
28		Pyritiferous Porphyry	South fork of White's gulch	0. 000, 113	0. 033	(?)
29		do	Comstock Tunnel, White's Hill	0. 000, 113	0. 033	(?)
30		do	West of Pilot fault, White's Hill	0. 000, 103	0. 030	Trace
31	90c	do	Head of White's gulch	None	None	None
32		do	Above Pilot mine, Printer Boy Hill	0. 000, 034, 3	0. 01	Do.
33	93b	do	Above Oro City, south side of California gulch	0. 000, 054, 8	0. 016	(?)
34	94a	do	Rebel Warrior shaft, Ball Mountain	0. 000, 079, 1	0. 023	None
35	118	Diorite	Red amphitheater, Buckskin gulch	0. 000, 013, 7	0. 004	Do.
36	124	Porphyrite	Bartlett Mountain, near summit	0. 000, 192, 5	0. 056	Do.
37	119b	do	Buckskin amphitheater, débris slope	0. 000, 048	0. 014	Do.
38	121	do	do		Trace (?)	Do.
39	266	do	"Horseshoe" dike in limestone		do	Do.
40	129	do	Arkansas amphitheater	0. 000, 017, 1	0. 005	Do.
41	131	do	do	0. 000, 006, 8	0. 002	Do.
42	140	Rhyolite	Black Hill, South Park	0. 000, 068, 7	0. 020	Do.
43	142	Trachyte	Head of Union gulch	0. 000, 092, 8	0. 027	Do.
44	143	Hornblende-andesite	West slope of Buffalo Peaks	0. 000, 103	0. 030	Do.
45		Granite	Big Evans gulch		None	Do.
46		do	Yankee Hill, south of Logan shaft		do	Do.

[a] Complete analysis of this specimen in Table I.

REMARKS ON TABLE IV.

For the estimation of such extremely small quantities of silver and gold as it was supposed some of the eruptive rocks from the Leadville region might contain, and even for their detection alone, a most extreme degree of care and precaution was imperative. It being necessary to operate upon large quantities of material, it was decided to make the determinations by crucible assay, this process combining the greatest accuracy with the least expenditure of time. It was found, however, after a number of tests, that none of the lead or litharge obtainable was sufficiently free from silver for the present purpose. The silver contained in the lead or litharge used for an assay was generally so largely in excess of that in the powdered rock mixed with it that the prills of silver obtained from the regular assay upon rock known to contain silver and from a check assay upon the lead or litharge alone frequently differed in weight only within the allowable limits of error. Recourse was then had to lead acetate, of which several lots were examined. These were all found much freer from silver than either of the substances previously tested, and one lot of commercial acetate from Mallinckrodt & Co., of Saint Louis, Mo., was used for all the assays tabulated above.

Preparatory to using, it was dehydrated by fusing in a large iron vessel till sudden swelling up and solidification of the whole mass took place, and then finely pulverized. This material, containing about 73 per cent. of lead by assay, was found by repeated tests, conducted, as given below, upon the same amounts as used for the rock assays, to carry 0.004 ounce silver per ton of 2,000 pounds, or 0.0000137 per cent., including a trace of gold far too small for estimation. The latter was left on solution of the silver in nitric acid as a minute black speck, indistinguishable without the aid of a lens. By collecting into one button the silver from 500 to 600 grams of dehydrated lead acetate, parting with great care, bringing the gold upon a sheet of white writing paper and flattening it out with a knife blade, the yellow reflection of gold could readily be observed by examination with a lens, and sometimes with the naked eye.

The process of assay was as follows: Four Hessian crucibles, of suitable size, were each charged with one assay ton (29,166 milligrams) of the sample to be assayed. two and one-half assay tons of the dehydrated lead acetate, and a proportionate amount of a flux consisting of soda, borax, and a little argol. After mixing well, a layer of salt was placed on top, and, if much pyrite was present, an iron nail inserted, The four charged crucibles were then placed covered in a wind furnace fired by coke, and left, with proper regulation and final strong increase of temperature, till fusion was complete. The contents were then poured into molds, the lead reguli, weighing each about 55 grams, reduced by scorification in a muffle to a smaller size, the reduced reguli united two and two and rescorified, and the two resulting therefrom again united and reduced by scorification to a single button of suitable size for cupellation. Toward the end of cupellation, which was always conducted with the greatest care and, as nearly as possible, under the same conditions of temperature for each assay, the button was poured from its cupel into another one immediately behind the first, in order that the cupellation might be finished upon a smooth bottom. If this precaution was neglected, the silver button was occasionally not to be found in the roughened surface of the cupel. After a little experience, no loss need be apprehended in pouring from one cupel to the other. The silver was then weighed upon an Oertling assay balance, indicating a difference in weight of 0.02 milligram with great exactness and of 0.01 milligram with tolerable accuracy. After deducting from the weight of silver found that due to the lead acetate, which, where ten assay tons had been used, would be 0.04 milligram, division of the remainder, if any, by the number of assay tons of rock taken gave directly the contents in ounces and decimal fractions of an ounce troy per ton of 2,000 pounds avoirdupois, since 29,166.6 ounces troy make one ton of 2,000 pounds avoirdupois and an assay ton contains 29,166.6 milligrams. The silver was then dissolved in nitric acid, but the presence of a trace of gold, derived from the lead acetate, rendered the detection of gold from the rock impossible, unless its amount considerably exceeded that of the lead salt. An example will best show the degree of accuracy attainable. Suppose rock and lead acetate to have been taken in the usual amounts: Four assay tons (116.66 grams) of the former to ten assay tons (291.66 grams) of the latter, and the final silver button to weigh 0.06 milligram. From this is to be deducted 0.04 milligram, and the remainder divided by 4, the number of assay tons of rock tested, gives 0.005 ounce per ton as the accurate result. Had the weight been 0.05 milligram, the correctness of the result, 0.0025 ounce, might be more open to doubt, as the balance cannot be counted upon to indicate differences of only 0.01 milligram with certainty. Hence, for the above quantities of sample and lead acetate, 0.005 ounce per ton is about the limit of accuracy.

There will be noticed in the table occasional instances, notably in No. 41, where lower figures are given. In these cases the amount of rock assayed had been increased without at the same time increasing the lead acetate. In the case of No. 41 it was impossible to decide from 4 assay tons whether silver was present or not, though the weight seemed to slightly exceed 0.04 milligram. By doubling the amount of sample and using still only ten assay tons of lead acetate, the weight of the silver sensibly increased, thus showing beyond reasonable doubt that the rock was argentiferous. It did not appear advisable, however, as a rule, to reduce the proportion between the weights of sample and lead acetate much below 4:10 for fear the reduced lead might not be sufficient to extract and collect the silver entirely.

LIMESTONES.

TABLE V.—*Complete analyses. Dolomitic limestones.*

	CaO	MgO	MnO	FeO	CO_2	SiO_2	Al_2O_3	Fe_2O_3	K_2O	Na_2O	H_2O	SO_3	P_2O_5	Cl	I	Org.	FeS_2	Totals.
I ...	26.60	17.41	......	0.83	40.01	11.84	1.66	1.51	0.017	0.029	0.48	......	Trace	0.05	Trace *a*	.....	......	100.436
II ...	30.79	21.14	Trace	0.24	46.84	0.21	0.27	0.21	0.030	0.062	0.22	Trace	Trace	0.10	....	0.03	Trace	100.142

a 0.000,022, 5 per cent.

I. Type of the Silurian or White Limestone. Coll. No. 164. From quarry in California gulch.
II. Type of the Lower Carboniferous or Blue Limestone. Coll. No. 170. Silver Wave claim, Iron Hill.

REMARKS ON TABLE V.

The carbon dioxide and the water of the above analyses were estimated as in the case of the eruptive rocks; the one by loss in weight upon treatment with hydrochloric acid in a suitable apparatus, the other by absorption in a calcium chloride tube.

The organic matter of Analysis II was determined by an ordinary combustion analysis, after dissolving a considerable quantity of the rock in dilute hydrochloric acid and collecting and drying the insoluble matter upon an asbestus filter. The carbon dioxide formed was caught in potash bulbs and weighed. For 58 parts of carbon found, 100 parts of organic matter were assumed, as recommended by Fresenius.

The trace of iodine shown in Analysis I was detected and estimated by dissolving one pound of the dolomite in nitric acid, precipitating the chlorine and iodine as silver salts, reducing the latter by zinc and sulphuric acid, separating the iodine by addition of potassium nitrite, collecting it in carbon disulphide, and titrating with a dilute solution of sodium hyposulphite. Bromine could not be detected. The chlorine was determined on from five to ten grams of rock by precipitation with silver nitrate from a nitric acid solution.

The alkalies were estimated by igniting twenty grams of the finely powdered rock in small portions in a platinum crucible to expel carbon dioxide, extracting with water and proceeding as in ordinary alkali determinations. As the amounts of alkalies found did not exceed those required by the chlorine to form chlorides, but rather fell slightly below, due perhaps to partial volatilization during the preliminary calcination, it seems probable that the chlorine is combined with sodium and potassium, and possibly small quantities of calcium and magnesium.

It was found that by boiling the powder with water without previous calcination a portion of the chlorine and alkali could be extracted, and that the amount increased as the pulverization was more perfect. The total amount of chlorine thus capable of extraction never equaled that actually present in the rock, however. Microscopical examination showed the dolomites to be full of extremely minute fluid inclusions. If the chlorine was derived from these inclusions, where it might be held as sodium and potassium chlorides, a ready explanation is afforded for the incomplete extractibility of the chlorine by boiling water. By no mechanical pulverization could such a perfect subdivision of the particles be effected as to expose all the inclusions; a considerable proportion would still remain intact and retain a corresponding amount of chlorine.

TABLE VI.—*Lime, magnesia, and chlorine determinations.*

No.	Coll. No.	Horizon.	Locality.	$CaCO_3$	$^aMgCO_3$	Cl
1	153	Cambrian	Below Excelsior mine, Buckskin gulch	25.43	4.03	Trace
2	153c	do	Monte Cristo Ridge, Quandary Peak	46.05	36.71	Do.
3	153a	Silurian	Dyer Mountain, marbleized bed	53.16	43.43	Do.
b4	164	do	Quarry in California gulch	47.50	36.56	0.05
5	......	do	Below Dyer Mine, Dyer Mountain (light blue, compact).	36.27	21.73	Trace
6	......	do	Below Dyer mine, Dyer Mountain (pinkish, decomposed).	34.30	20.12	Do.
7	162	do	Red amphitheater, Buckskin gulch	49.30	32.49	Do.
b8	170	Blue Limestone	Silver Wave claim, Iron Hill	54.98	44.39	0.10
9	286	do	Dunkin mine, Fryer Hill (lime-sand)	55.14	44.29	Trace
10	285	do	Chrysolite mine, Fryer Hill (lime-sand)	54.09	43.79	Do.
11	170b	do	Empire Hill	......	......	Do.
✓ 12	170a	do	Ridge south of Sacramento mine	56.80	41.89	✓ Do.
13	171b	do	Sacramento mine, Spring Valley	54.30	44.33	✓ Do.
14	174	do	South slope of White Ridge, Horseshoe gulch	52.77	36.01	Do.
15	......	do	London mine, London Mountain	52.86	40.49	......
16	291	Weber Grits	Ridge west of Mount Silverheels, Park County	54.34	43.82	✓ Trace
17	197	do	East bank of Beaver Creek	54.32	43.24	✓ Do.
18	292	do	Beaver Creek, west base of Mount Silverheels	53.91	43.73	✓ ......
19	......	Upper Coal Measure.	Robinson limestone, lower bed	87.87	6.93	✓ ......
20	......	do	Robinson limestone, upper bed	97.11	Trace	✓ ......
21	198	Trias	(Con. fracture) Silverheels, between Fairplay and Como	95.78	1.53	✓ ......
22	198a	do	(Con. fracture) first ridge west of Crooked Creek	99.11	0.36	✓ ......
23	416	do	(Con. fracture) Jacque Mountain, near summit	97.54	0.76	✓ ......
24	......	Trias ?	Calcareous shale, near flume northeast of Fairplay	16.18	12.01	......
25	......	do	do	34.84	26.35	......
26	......	Lake beds	Marl from Greenback shaft, Graham Park	48.63	1.42	......

a Includes $(FeMn)CO_3$. *b* Complete analysis in Table V.

N. B.—Blank spaces in above table denote that no tests were made.

REMARKS ON TABLE VI.

The above figures for calcium and magnesium carbonates have been calculated from the lime and magnesia actually found. With exception of two numbers indicated by *b*, the chlorine was determined only qualitatively and noted as "trace," although generally present in quantity sufficient for estimation.

TABLE VII.—*Serpentine and amphibole from dolomitic limestones.*

No.	Coll. No.	SiO_2	Al_2O_3	Fe_2O_3	FeO	MnO	CaO	MgO	K_2O	Na_2O	H_2O	CO_2	P_2O_5	Cl	Total.
I	......	40.15	0.93	1.28	1.05	......	2.08	40.03	......	......	*a*12.88	1.60	Trace	......	100.00
II	161	17.64	0.99	0.62	0.18	Trace	32.24	19.01	Trace	0.07	3.72	25.33	0.05	0.08	99.93
III	161a	55.21	2.02	0.38	3.19	0.20	13.50	23.20	......	Trace	1.07	......	......	......	98.97

a By difference.

I. A layer of pure serpentine on limestone from east side of Red Amphitheater, Buckskin gulch, Park County.

II. Limestone thoroughly impregnated with serpentine, producing a light-yellow rock. No pyroxene or amphibole visible under the microscope. Same locality.

III. Amphibole with some pyroxene from limestone containing serpentine. Same locality.

REMARKS ON TABLE VII.

In connection with the serpentine of Analysis I, it was found by treatment of the powdered rock with very dilute hydrochloric acid that the carbon dioxide was combined entirely with lime. That the rock is a true serpentine appears from a calculation of the oxygen ratios.

Oxygen percentages.	Oxygen ratios.
SiO_2 = 21.413	RO(FeOMgO) : SiO_2 : H_2O
FeO = 0.233 } 16.245	1 : 1.318 : 0.705
MgO = 16.012 }	3 : 3.95 : 2.11
H_2O = 11.449	Required by theory 3 : 4 : 2

Analysis II, after deducting calcium carbonate, shows the residue to have a composition approximating to that of serpentine.

The mineral of which III is the analysis was obtained in an apparently pure state by treatment with dilute hydrochloric acid, whereby 16½ per cent. of the whole went into solution and was found to consist chiefly of calcium carbonate, with a little magnesia, ferrous oxide, and phosphorus pentoxide.

ORES AND VEIN MATERIALS.

TABLE VIII.—*Sand carbonates.*

	I.	II.	III.
	Adelaide.	Little Chief.	Waterloo.
PbO	80.352	75.408	77.980
Al_2O_3	0.444	1.415	
Fe_2O_3	0.467	1.940	
FeO	0.299		
MnO	0.137	0.074	
MnO_2		1.386	
CoO	Trace	Trace	
ZnO		0.095	
CaO	0.303	0.335	
MgO	0.068	0.056	
SiO_2	0.651	1.972	
Sb_2O_5		0.121	
As_2O_5	Trace	Trace	
P_2O_5	1.532	Trace	6.480
SO_3	Trace	0.486	
CO_2	14.700	14.251	10.180
Cl	0.255	0.288	0.840
H_2O	0.395	1.140	
Ag	0.009	0.777	0.047
Au	Trace	Trace	
Total	99.612	99.744	
Less O for Cl	0.057		
	99.555		

I. Sand Carbonate, Ore Coll. No. 36, Adelaide mine, North Iron Hill.
II. Sand Carbonate, Ore Coll. No. 33, Little Chief mine, Fryer Hill.
III. Sand Carbonate, Waterloo mine, Carbonate Hill.

REMARKS ON TABLE VIII.

Consideration of Analysis I shows that the carbon dioxide is insufficient for combination with all the lead oxide, and that the chlorine (entirely soluble in nitric acid) and the phosphorus pentoxide bear to one another the exact ratio of chlorine and phosphorus pentoxide in the mineral pyromorphite; $3(3PbO, P_2O_5)+PbCl_2$.) Calculating from the chlorine there is found to be 9.75 per cent. of this mineral. The carbon dioxide is, then, somewhat more than sufficient for the remaining lead oxide, forming 86.60 per cent. of cerussite, $PbCO_3$.

The silver exists in the state of chloride, as shown by extracting it from a large amount of ore with ammonia.

A portion of the ferrous oxide is present as magnetite. The slight excess of carbon dioxide above that required for the lead is probably combined with ferrous, manganous, and calcium oxides.

Analysis II shows plainly that pyromorphite is practically absent from the specimen of ore from the Little Chief mine. The lead exists mainly as carbonate, with a little sulphate, and probably a small amount of antimoniate. A yellow substance left unattacked with the silica and silver chloride ore, on treating with nitric acid, gave reactions for lead and antimony. The silver exists altogether in the state of chloride and could be completely extracted by ammonia. The total chlorine, found by fusion with alkaline carbonates, extraction with water, and subsequent precipitation with silver nitrate, is slightly in excess of that required by the silver, but it was found that a few hundredths of 1 per cent. was present in a combination soluble in water.

Only the chief constituents from the ore in the Waterloo mine (Analysis III) were estimated. Starting from the chlorine, which represents that soluble in nitric acid alone, the ore is found by calculation to contain 32.07 per cent. of pyromophite and 61.78 per cent. of cerussite, the carbon dioxide exactly sufficing for the lead oxide left after combining the elements of pyromorphite. An excess of 1.44 per cent. phosphorus pentoxide is probably combined with alumina, of which a considerable amount was found to be present.

TABLE IX.—*Chloro-bromo-iodides of silver.*

	I.	II.	III.	
Cl	13.78	9.80	99.925	I. Ore Coll. No. 39. Robert E Lee mine, Fryer Hill.
Br	85.63	89.99	....	
I	0.59	0.21	0.075	
	100.00	100.00	100.000	II. Ore Coll. No. 30c. Amie mine, Fryer Hill.
AgCl	21.59	15.75	99.966	
AgBr	77.99	84.09	....	III. Ore Coll. No. 37. Big Pittsburgh mine, Fryer Hill.
AgI	0.42	0.16	0.034	
	100.00	100.00	100.000	

REMARKS ON TABLE IX.

The figures in the upper series represent the relative proportions of Cl, Br, and I; those in the second series, the percentages of the corresponding silver salts.

The ore specimens, having first been treated with nitric acid to extract any soluble chlorine salts, were then subjected to the reducing action of zinc and sulphuric acid, whereby the silver salts were entirely reduced. To the filtered solution, containing all the chlorine, bromine, and iodine, potassium nitrite was added, the liberated iodine collected in carbon disulphide, separated with the latter by filtration, and estimated by titration with dilute sodium hyposulphite solution. The chlorine and bromine were then thrown down by silver nitrate, the precipitate was washed thoroughly by decantation, brought entirely into a tared vessel, fused, and weighed. As sufficient material had been taken to insure several grams weight of mixed chloride and bromide, the estimation of the halogens by entire conversion into silver chloride in a current of chlorine gas was repeated on different portions with very closely agreeing results, of which the above are the mean. In Analysis III a qualitative test failed to indicate the presence of a trace of bromine, and the fused silver chloride, when heated in chlorine gas, showed no change whatever in weight. The silver in the ore, reduced by the action of zinc and sulphuric acid, was not estimated. The figures in the second horizontal series above are therefore obtained by calculation from the chlorine, bromine, and iodine found. In Analysis I the proportion of AgCl:AgBr is 4:11, while in Analysis II it is 1:4.

TABLE X.—*Various ores and vein materials.*

N. B.—With the exceptions noted, blanks in the table denote "no tests."

1. "Hard carbonate," Scooper mine, Yankee Hill.
2. "Silicious ore," El Capitan mine, Taylor Hill.
3. "Gold ore," Ore Coll. No. 60, El Capitan mine, Taylor Hill.
4. Silicious hematite, Ore Coll. No. 84, Chrysolite mine.
5. "Iron ore," Ore Coll. No. 90, Kenosha mine, Long and Derry Hill.
6. Altered limestone (light material), Garden City mine.
7. Altered limestone (dark material), Garden City mine.

8, 9, and 10. Specimen showing pyrite altering to a light ocherous mass. Ore Coll. No. 44, No Name gulch, Lake Co.
8. Nucleus of pyrite. 9. Dark zone. 10. Light outer zone.

11. White filling in chert nodule from porphyry, Ore Coll. No. 300b, Ben Burb shaft.
12. Chert nodule, Ore Coll. No. 299, El Paso shaft.
13. Breccia with ore cement, Ore Coll. No. 53, Evening Star mine.
14. Chert under ore body, Little Pittsburgh mine.
15. Granular quartz under ore, Ore Coll. No. 3a, Waterloo mine.

	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15
Insoluble.	80.00			8.20				7.36	16.36	54.28					
SiO_2		82.400	11.560			10.08	20.02				*a*98.03	*b*96.3	*c*84.2	97.9	*d*81.29
PbO		Trace	None												
FeS_2								53.61							
Fe_2O_3	13.91	10.380	77.600	54.14	7.70	0.81	0.97	34.44	73.42	39.84					
Al_2O_3		4.160			None										
MnO_2	0.10			22.36	65.98	0.39	1.56								
ZnO				2.56	1.00										
CaO		*e*1.217	*e*0.6094			27.17	24.15								
MgO						19.49	16.71	Trace	*e*0.09	*e*0.06					
CO_2						*f*42.78	*f*37.37								
SO_3									Trace	Trace					
H_2O	*g*5.978	1.830	10.230	*h*12.709	4.22										
Ignition									10.13	5.82					
Ag	0.012	0.002	0.0002	0.031											
Au	Trace	0.011	0.0004	None											
Totals	100.000	100.000	100.000	100.000	*i*78.90	100.72	100.78	*j*95.41	100.00	100.00					

a Remainder Fe_2O_3 and Al_2O_3, no water.

b Remainder Fe_2O_3 and Al_2O_3. Of the SiO_2 3.06 per cent. was soluble in a moderately strong solution of potassium hydrate.

c Cementing material chiefly pyromorphite, with some galena and cerussite, also a little calcite.

d Remainder chiefly $PbCO_3$ and Fe_2O_3; of the silica 3.93 per cent. was soluble in a moderately strong solution of potassium hydrate.

e By difference.

f Calculated.

g By difference; includes some PbO and Sb_2O_5.

h By difference; includes a little Sb_2O_5.

i Remainder chiefly SiO_2.

j Remainder is SO_3 and H_2O.

REMARKS ON TABLE X.

The analyses of the above table were made without view to completeness, the object being in the majority of cases to ascertain merely the general nature of the ore or material under hand. As this appears at a glance from the tabulated results, further remarks are unnecessary except in the case of 8, 9, and 10. The specimen showed a nucleus of granular pyrite in process of decomposition, ferric oxide being observable throughout the mass. This very irregular nucleus was inclosed in an envelope of dark-brown hydrated ferric oxide, the boundaries being in places rather sharply defined, in others indistinct. The dark zone was in turn surrounded by a zone of light-brown oxide, the line of demarcation being very regular and sharply defined. The dark oxide was compact and flinty; the light oxide also compact, but less hard.

TABLE XI.—*Alteration products of porphyry.*

No.	Ore Coll. No.	Local name.	SiO_2	Al_2O_3	Fe_2O_3	FeO	ZnO	CaO	MgO	K_2O	Na_2O	H_2O	SO_3	Totals.
1	55b	Kaolin......	48.72	34.01	0.56	0.66	....	....	1.11	9.88	0.67	4.42	....	100.03
2	55a	Chinese talc	43.36	37.78	....	......	....	0.22	0.30	Trace		17.95	Trace	99.91
3	56c	Kaolin......	4.55	35.60	2.26	......	....	Trace	Trace	2.73	5.26	15.05	34.55	100.00
4	56	Chinese talc	24.47	38.05	0.93	0.77	....	0.23	0.30	2.72	1.30	16.67	15.48	101.15*a*
5	56b	....do	27.89	33.79	....	......	....	0.53	1.14	2.83	1.56	*b*16.51	15.75	100.00
6	105a	... do	35.33	10.38	....	......	33.05	1.62	0.71	......	.....	19.06	......	100.15
7	105a	....do	35.97	8.81	....	......	35.40	1.87	0.80	......	......	17.46	....	100.31
8	105b	... do	37.54	24.76	*c*0.64	......	18.43	0.63	0.71	0.66	0.36	16.37	....	100.10

a Includes 0.23 P_2O_5. *b* By difference. *c* Present as a visible impurity.

1. Amie mine, in ore body. Ore Coll. No. 55b.
2. New Discovery mine. Ore Coll. No. 55a.
3. Big Pittsburgh, contact of Gray Porphyry. Ore Coll. No. 56b.
4. Morning Star mine. Ore Coll. No. 56.
5. Swamp Angel tunnel, contact of White Porphyry. Ore Coll. No. 56b.
6. Lower Waterloo mine. Ore Coll. No. 105a.
7. Lower Waterloo mine. Ore Coll. No. 105a.
8. Lower Waterloo mine. Ore Coll. No. 105b.

REMARKS ON TABLE XI.

Owing to the indefinite nature of the greater part of the peculiar products of alteration represented by analysis in the above table, it is impossible to ascribe to them distinctive names. Notwithstanding the great external similarity of all but the first of the specimens examined, they have been found to differ most widely in composition, though, aside from the above exception, three distinct groups may be recognized, namely: First, simple hydrated aluminium silicates allied to kaolinite; second, mixed aluminium silicates and aluminium and alkali sulphates, likewise hydrated; and, third, certain hydrated aluminium and zinc silicates, also mixtures.

In the following are given the distinctive physical and chemical characteristics, accompanied by brief discussions of the analytical results:

No. 1, grayish white; compact, but of hardness considerably less than 1, rubbing off on the fingers; luster, pearly; insoluble in hydrochloric acid. Evidently derived directly from porphyry, since honeycombed remnants of feldspar crystals, and even large crystals, an inch in length, showing rough faces, occur imbedded in the mass. Under the microscope it appears to consist of crystalline scales without definite form.

In order to obtain material for analysis free from undecomposed feldspar, it was slightly crushed and stirred with water in a beaker, whereby it became thoroughly disintegrated, the fine matter floating and imparting to the water a beautiful satiny appearance similar to that frequently observable in streams receiving the tailings from stamp mills, while the gritty particles fell to the bottom. By pouring off the suspended matter, allowing to settle, decanting the supernatant liquid, and drying the slimy deposit, an apparently pure matter was obtained, showing the pearly luster of the original mass, and containing, like that, when air-dried, about one-quarter of 1 per cent. of hygroscopic moisture. This is not included in the above analysis. No further loss occurred on prolonged heating until a temperature considerably above 100° C. was reached, while a strong red heat was requisite for complete expulsion of the water.

No altogether satisfactory formula can be deduced from the figures in the table. On dividing the molecular value by that for water, as being most accurately determined, the ratio is found to be

$$\begin{array}{cccc} SiO_2 : & Al_2O_3 : & R_2(R)O : & H_2O \\ 9.93 & 4.09 & 1.87 & 3.00 \end{array}$$

or, approximately,

$$10 : 4 : 2 : 3.$$

As no other specimens of a similar nature, from the Amie or other mines, have been observed, by analysis of which it could be ascertained whether the above ratio remains constant or not, it would be rash to affirm that the material analyzed represents a distinct mineral species, the final product of the alteration of the porphyry from which it is derived.

No. 2, pure white, veined with manganese dioxide; compact, hardness about 2, rubbing off on the fingers when dry. When fresh and moist, frequently greenish in color, opaline in appearance, and semi-transparent, especially on the thin edges, becoming opaque on exposure. Insoluble in hydrochloric acid. Portions free from MnO_2 taken for analysis.

It was found that after two or three years' exposure to the air a large amount of water, 3.36 per cent. of that given in the analysis, was still retained in a very weak state of combination, apparently as hygroscopic moisture, since it escaped over sulphuric acid. No further loss occurred on heating at 100° C., nor below 160° C. to 170° C., although blackening took place, due to carbonization of organic matter. Dried over sulphuric acid or at 100° C., the powder was so extremely hygroscopic that it was deemed advisable to make the analysis upon air-dried material. The percentage of loosely combined or hygroscopic water was found to decrease slowly on long exposure of lumps to the air, so slowly as to be perceptible only at intervals of a month or more. Deducting all water driven off at 100° C., the molecular ratio $SiO_2 : Al_2O_3 : H_2O$ is 1.98 : 1.00 : 2.20, thus showing the substance to be closely allied to kaolinite.[1]

Nos. 3, 4, and 5. In general appearance 4 and 5 differ little from the substance last described. Color, white, streaked frequently with iron and manganese oxides; hardness, after long exposure, in case of 5, about 2½. Practically insoluble in hydrochloric acid. No. 3 is pure white, and resembles 1; it contained no hygroscopic water. No. 4 contained but 1.23 per cent.; while No. 5 retained 4.58 per cent. of the same (included in the analysis), after long exposure to the air in the form of lumps.

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[1] The same is found in the Morning Star consolidated group of mines according to L. D. Ricketts, one of whose published analyses (The Ores of Leadville, Princeton, 1883) shows a ratio $SiO_2 : Al_2O_3 : H_2O = 2 : 1 : 3$, probably including hygroscopic or weakly combined water.

The air-dried material of 4 and 5 was analyzed, since, when dried over sulphuric acid or at 100° C., the hygroscopicity was such as to render accurate weighing out of the question. In very few hours No. 5 reabsorbed, when exposed in the air, over half of the 4.58 per cent. of moisture lost at 100° C.

Consideration of the analyses, coupled with the observed insolubility in hydrochloric acid, shows beyond reasonable doubt that these bodies are mixtures of alunite, $K_2SO_4+(Al_2)S_3O_{12}+2H_6(Al_2)O_6$, corresponding in formula to the jarosite of the following table, or of an allied mineral, with different indefinite hydrated aluminium-calcium-magnesium silicates. If the supposed alunite is calculated on the basis of the sulphuric acid, the residual amounts of silica, alumina, lime, magnesia, alkalies, and water are found to have widely different and not very definite molecular ratios in each analysis.

From the fact of No. 3, which is mainly an aluminium-alkali sulphate, containing no weakly combined or hygroscopic water, the hygroscopicity appears to be a property of the hydrated aluminium silicates.

Nos. 6, 7, and 8. Similar in appearance to the simple hydrated aluminium silicates represented by analysis 2. Nos. 6 and 7 were taken by Mr. L. D. Ricketts from one locality in the mine, No. 8 from another. The first and second were not to be distinguished from each other by the eye, being brilliantly white (greenish under certain conditions of light), opaline and semi-transparent, while the third was veined with iron and manganese oxides and possessed in a less degree the pronounced conchoidal fracture of the others. On exposure they became opaque, and after some months possessed a hardness of about 3. Nos. 6 and 7 were entirely and readily decomposed by strong hydrochloric acid when finely pulverized, while upon 8 the action of the acid was not so marked, though still energetic. The hygroscopicity of these substances, especially of the first and second, is extraordinary. Over sulphuric acid No. 6 lost 11.64 per cent. of water, while No. 7 lost 10.26 and No. 8 but 5.30 per cent., these amounts being included in the tabulated results of analysis. Exposure to a temperature of 100° C., and even 150° C., occasioned no further loss in weight, but the presence of organic matter made itself manifest by the blackening of the powder. The dried material reabsorbed moisture with great rapidity. It was at first supposed that these were mixtures of calamine with some hydrated aluminium silicate. But if from the molecular values those for zinc oxide are eliminated and proportionate amounts for silica and water subtracted, on the supposition that calamine is present, the ratios between the remaining molecular values are not the same as should be the case if the mixture consisted in all these cases of calamine and one other definitely constituted mineral. Moreover, on decomposing with hydrochloric acid no gelatinization takes place, and *not even a trace* of silica goes into solution, an argument against the possibility of the presence of either calamine or willemite. The molecular values, considered altogether for each analysis, do not present relations sufficiently definite to allow of supposing any one of the specimens to represent a single mineral species.

TABLE XII.—*Alteration products of galena and pyrite.*

No.	Ore Coll. No.	SiO_2	Fe_2O_3	Al_2O_3	CaO	MgO	K_2O	Na_2O	H_2O	PbO	Bi_2O_3	As_2O_5	P_2O_5	SO_3	Cl	Totals.
1	106a	None	46.70	None	0.06	0.06	5.33	1.68	10.54	4.27	0.08	0.46	0.08	30.53	0.02	99.81
2	106b	0.30	42.98	0.20	0.64	None	6.31	0.83	10.12	8.27	None	0.42	1.58	27.81	0.26	99.72
3	106c	0.36	44.40	0.23	None	None	0.15	0.37	8.99	19.50	None	0.39	0.11	25.07	0.04	99.1
4	106d	......	......	......	......	......	2.12	0.57	......	......	......	......	......	......	......	......
5	106e	......	......	......	......	......	0.36	0.77	......	......	......	......	......	......	......	......
6	106f	......	......	......	......	......	1.96	0.18	......	......	......	......	......	......	......	......
7	106g	......	......	......	......	......	4.04	0.57	......	......	......	......	......	......	......	......

1. From Maid of Erin mine, under White Porphyry. Contains 0.0048 Ag and trace of Au.
2. From Morning Star (Forsaken) mine, under Gray Porphyry. Contains 0.0036 Ag.
3. From Lower Waterloo mine, under Gray Porphyry. Contains 0.075 Ag.
4. From Morning Star (Forsaken), under Gray Porphyry.
5. From Morning Star (Forsaken), under Gray Porphyry.
6. From Morning Star (Forsaken), under Gray Porphyry.
7. From Silver Cord mine, under White Porphyry.

REMARKS ON TABLE XII.

Notwithstanding the great similarity in appearance of all the specimens of which the above are analyses they are rather complex mixtures in very varying proportions of several mineral substances. They all show a similar chemical behavior. Heated in a closed tube, water is first evolved, the substance then changes from an ocherous, or sometimes brownish yellow, to dark brown, and later sulphuric and sulphurous acids escape. The same changes occur in an open tube. On charcoal with soda there appears sometimes a slight coating of arsenic trioxide, accompanied by a smell of arsenic and the reaction for lead. Entirely insoluble in boiling water. Nitric acid in the cold extracts part of the lead oxide; also, arsenic and phosphorus pentoxides and usually chlorine. Continued boiling with nitric acid seems to decompose the iron minerals completely. Warm hydrochloric acid effects complete decomposition, and no trace of a ferrous salt can be detected, even if solution has been effected in an atmosphere of carbon dioxide. Caustic alkalies also decompose them completely, all the sulphur trioxide and arsenic and phosphorus pentoxides going into solution, whereby the analysis is materially simplified. Various tests, combined with a consideration of the three complete analyses, show that the lead is present as anglesite and pyromorphite[1]: in No. 2 as the latter mineral alone, the lead oxide, phosphorus pentoxide, and chlorine being in the exact proportions required for the formula $(3PbO, P_2O_5)+PbCl_2$. The As_2O_5 is not present in the corresponding chloro-arseniate of lead, as shown by the fact that the proportion of P_2O_5 to that part of the Cl not combined with silver is always the same as in pyromorphite, and that there is insufficient lead for both phosphorus and arsenic pentoxides together, as in No. 2, where it exactly suffices for the phosphorus pentoxide. The arsenic pentoxide is therefore undoubtedly present as a hydrated ferric arseniate. By combining in the first place chlorine and phosphorus pentoxide with lead oxide and the remainder of the latter with sulphur trioxide, definite conclusions may be reached as to the composition of the remainder of the mixture. Analyses

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[1] The pyromorphite is sometimes visible in bunches of small crystals.

1 and 2 (especially the latter) show that the chief constituent in these cases is probably *jarosite*, $K_2SO_4+(Fe_2)S_3O_{12}+2H_6(Fe_2)O_6$. As there remains a slight excess of SO_3 and Fe_2O_3 after combining the constituents of this mineral on the basis of the alkalies present, a basic ferric sulphate is to be assumed as a further constituent of the mixture.

Analysis 3 shows little pyromorphite, much anglesite, little jarosite, and much hydrated basic ferric sulphate, of which latter it is impossible to determine the formula definitely, since it is not known how much ferric oxide and water may be combined with the arsenic pentoxide.

The remaining partial analyses were made to ascertain whether or not the alkalies were constant constituents of this class of products of alteration of the original vein material. Qualitative tests showed that pyromorphite and anglesite were occasionally present in greater amount than shown in analyses 1, 2, and 3. Arsenic pentoxide was found in very considerable quantity in the material from the Silver Cord mine (7).

TABLE XIII.—*Miscellaneous alteration products.*

No.	Rock.	Locality.	SiO_2	$\left.\begin{matrix}Fe_2O_3\\Al_2O_3\end{matrix}\right\}$	K_2O	Na_2O	Soluble in HCl.	Loss by ignition.	Totals.
1	Altered White Porphyry	London mine	71.2	14.5	3.3	0.7	[a]10.0		99.7
2	do	do	72.8						
3	do	New York mine	75.8						
4	Altered White Limestone	Buffalo Peaks	97.1						
5	Altered Blue Limestone	do	97.7						
6	do	do	78.9						
7	Flint Kernel[b]	do	[c]98.16					1.84	100.00
8	Altered opaline coating[d]	do	[e]97.42					2.58	100.00

[a] Consists of calcium carbonate, with a little magnesium and less manganese carbonate.

[b] Specific gravity at 18½° C. = 2.570. Hardness, 6.

[c] Soluble in strong solution of potassium hydrate after four to five hours' digestion = 65.73 per cent.

[d] Specific gravity at 16½° C. = 2.023. Hardness 5.5.

[e] Soluble in strong solution of potassium hydrate after four to five hours' digestion = 97.42.

TABLE XIV.—*Assays of ores, vein materials, and country rocks.*

No.	Ore Coll. No.	Name.	Mine.	Remarks.	Ag.	Au.
		Ores.			*Oz. per ton.*	*Oz. per ton.*
1		Sand carbonate	Matchless, near Hibernia	Crystals of cerussite	682. 90	None
2	35b	do	Morning Star, fourth level north	Compact	10. 00	Trace
3	35d	do	Morning Star, first level south	White sand	41. 00	None
4	35c	do	Waterloo, on Forsaken line	Dark colored	13. 80	Do.
5	35c	do	do	do	25. 40	Trace
6	91	do	Long and Derry		22. 10	None
7	40	Chloride	Evening Star, on Catalpa line, second level.	With cerussite	82. 20	Trace
8	45	Hard carbonate	Dunkin, between first and second levels.		17. 90	Do.
9	42a	do	Evening Star, on Morning Star line.		33. 80	None
10		do	Niles-Augusta		93. 40	Do.
11		do	do	In limestone	81. 00	Trace
12	41	do	Scooper, east drift	Very silicious	3. 50	Do.
13	26a	Galena and cerussite	Dunkin, upper evel		353. 20	None
14	27	do	Henriett lower		516. 40	Do.
15	27e	Galena	Waterloo, near Forsaken line		420. 00	Trace
16	27e	Cerussite crust on No. 15.	do		28. 60	None
17		Galena	Dunkin, above ore body	Cementing porphyry breccia.	443. 80	Do.
18	27b	Galena, blende, and pyrite.	A. Y. mine, 150 feet deep		9. 80	Do.
19	28d	Galena and pyrite	Ontario, Breece Hill	Gash vein in porphyry.	20. 00	Trace
20	51	Copper ore	Florence, winze on north drift		2. 90	Strong tr
21	49	Gold ore	do		2. 00	0. 10
22	48	do	Lower Printer Boy, middle shaft, second station.		0. 50	Strong tr
		Vein materials.				
23	13	Silicious hematite	Vanderbilt, at 120 feet		16. 00	Trace
24	84	do	Chrysolite		9. 10	None
25		Silicious iron	Amie, lower body	Hydrated	2. 90	Do.
26		do	New Discovery	Jaspery	0. 50	Do.
27		do	Across the Ocean	Cavernous	1. 10	Do.
28		Black iron	Chrysolite-Vulture No. 1	Manganiferous	7. 80	0. 20
29		do	Climax, upper working	do	3. 90	None
30		Hematite	Breece Iron, upper shaft	do	0. 30	Do.
31		do	do	With pyrite	Trace	Do.
32		Pyrolusite	Crescent, lower shaft	Barren contact	4. 00	Do.
		Country rocks.				
33		Blue Limestone	Catalpa, east incline, near bottom	Black and altered	0. 50	Do.
34		do	do	Another specimen	0. 30	Do.
35		White Limestone	Climax, lower level south	Altered contact	0. 40	Do.
36		do	Amie No. 2, 273 feet deep	Altered	0. 50	Trace
37		Lime-sand	Dunkin, north end third level	Altered Blue Limestone	0. 70	None
38		do	Chrysolite, west of Vulture No. 1	Near rich ore body	85. 70	Trace
39		do	Chrysolite, near Little Chief line	In ore body	0. 70	Do.
40		do	Chrysolite, second level, northwest end.	In porphyry	0. 40	None

APPENDIX C

METALLURGY

BY

ANTONY GUYARD

CONTENTS.

SECTION IV. PRODUCTS OF SMELTING.

SECTION V. THEORETICAL DISCUSSION.

ARGENTIFEROUS LEAD SMELTING AT LEADVILLE.[1]

INTRODUCTION.

When Mr. S. F. Emmons intrusted me with the duty of reporting on the smelters of Leadville, he directed me to insist on the mechanical appliances offering a special interest, to lay special stress on the chemical phenomena of the blast furnaces, and to examine carefully the different furnace products of the smelting works. The fact that metals and other substances sparingly distributed throughout the mineral deposits of the Leadville mining camp are concentrated in these products gave a special interest in this study, which had a direct bearing upon the geology and mineralogy of the district under survey.

The fine analytical laboratory of the Survey in Denver made this study possible; and I trust that the numerous new facts and discoveries resulting from it will prove interesting and useful, not only to the metallurgist and miner, but also to the chemist and geologist.

In order to render as intelligible as possible the description of the plant, apparatus, and implements used in smelting in Leadville, I took the measurement of the most interesting portions, and made from them rough sketches, which were afterwards completed and corrected by Mr. W. H. Leffingwell, assistant topographer of the Survey. The disposition of the inside of inaccessible parts, such as the dust-chambers, was explained to me by the superintendents of the smelters. Some tracings of furnaces, blast apparatus, and dust-chambers were kindly given by Messrs. Billing and Eilers, owners of the Utah smelter; James Brierton, of the Harrison Reduction Works; August Werner, of the Elgin Smelter; and by Messrs. Fraser and Chalmers, of Chicago, manufacturers of a great number of the furnaces and smelting implements used in Leadville. The full description of the crushers and blowers was also kindly communicated to me by the respective manufacturers. The sketches accompanying this report, most of which are drawn to scale, have been prepared from these data by Mr. Morris Bien, assistant topographer, and engraved by Mr. Julius Bien, the well-known engraver.

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[1] In consequence of the sudden death of Mr. Antony Guyard at Paris, France, on the 29th of March, 1884, it was impossible to have his aid in the final revision of this report for the press. In making this revision I have received most important assistance from Mr. W. F. Hillebrand, who was present in the laboratory at the time Mr. Guyard was making his analyses. We have confined our changes to obvious clerical errors in calculations, and to minor alterations of language which would render his meaning clearer, changes which in no way affect his conclusions. It is possible that could Mr. Guyard have been personally consulted he might have made other modifications in his report, but, in view of his great experience and reputation as a metallurgical chemist, I have not felt authorized to do more than offer a few suggestions in foot-notes, where his meaning seems obscure or liable to be misunderstood. (S. F. E.)

During my visits to the smelters, I was received everywhere with the greatest courtesy, both by miners and superintendents, supplied with all information, and allowed to inspect thoroughly every part of the works. I was at the same time requested not to publish the names of the smelters in connection with information which might betray their private interests. In compliance with this request, in the following report the principal smelters have been designated by letters in all cases where it seemed possible that the publishing of their names might be detrimental to their interests. All analyses, where not otherwise specifically designated, have been made by me in the laboratory of the Survey.

SECTION I.

PRELIMINARY CONDITIONS OF SMELTING.

LEADVILLE.

Situation.—The city of Leadville is situated 10,150 feet above the level of the sea, in the Rocky Mountains, valley of the Arkansas, California mining district, Lake County, State of Colorado. It is placed in direct communication with Denver, the capital of Colorado, and thence with the east, by means of the Denver and South Park Railroad and the Denver and Rio Grande Railroad, both lines running on the same track between Buena Vista and Leadville.

The young city of Leadville, which did not exist three years ago, is full of bustle, life, and excitement, and had a population at the last census (1880) of about 15,000 inhabitants, but which fluctuates a good deal. It is built on a mesa or terrace formed of rearranged moraine material brought down by large glaciers which once existed there, and is surrounded on all sides by hills and mountains rising from three to four thousand feet above its level. The most conspicuous points are: on the west side, Mount Elbert and Mount Massive; on the east side, Ball Mountain, the Mosquito Pass, and Mount Sheridan; and on the north side, Mount Zion, above the Arkansas River.

MINES.

Most of the lead and silver mines are situated to the eastward of Leadville (northeast, due east, and southeast), in the localities known as Fryer Hill, Carbonate Hill, Iron Hill, Printer Boy Hill, Long and Derry Hill, Little Ellen Hill, Stray Horse gulch, and Iowa gulch. As the names of the mines and of their ores recur frequently in this report, some information concerning them has been tabulated below.

The geological information was kindly communicated by Mr. E. Jacob, geological assistant to Mr. S. F. Emmons, and the output data were taken from the Leadville Weekly and Monthly Circular, which receives its information direct from the mine superintendents.

In Table I is given the daily output of the working mines, whose names are arranged alphabetically and grouped according to locality.

TABLE I.—*Leadville mines.*

Name of mine.	Chief rocks passed through by the shafts.	Daily output in tons.	Date of output.
Fryer Hill.			
Amie	Wash and White Porphyry	30	Aug. 21, 1880
Chrysolite	do	75	Aug. 21, 1880
Climax	do	20	Aug. 21, 1880
Dunkin	do	8	Aug. 21, 1880
Hibernia	do	12	July 15, 1880
Little Chief	do	110	Aug. 21, 1880
Little Pittsburgh	do	25	Aug. 21, 1880
Matchless	do	25	Dec. 4, 1880
Robert E. Lee	do	40	Aug. 21, 1880
Virginius	do	10	Aug. 24, 1880
Carbonate Hill.			
Agassiz	White Porphyry	5	Dec. 20, 1879
Carbonate	do	35	Dec. 20, 1879
Catalpa	do	10	Aug. 21, 1880
Crescent and Ætna	do	10	Aug. 21, 1880
Evening Star	do	60	Dec. 4, 1880
Forsaken	do	20	Aug. 21, 1880
Half Way House	do	10	Dec. 20, 1879
Little Giant	do	5	Aug. 21, 1880
Morning Star	do	65	Aug. 21, 1880
Pendery and Glass	do	20	Aug. 21, 1880
Yankee Doodle	do	5	Dec. 20, 1879
Stray Horse gulch.			
Double Decker	Quartzite	2½	Dec. 20, 1879
North Iron Hill.			
Adelaide	White and Gray Porphyry	5	Dec. 20, 1879
Argentine and Camp Bird	White Porphyry	10	Dec. 20, 1879
Iron Hill.			
Iron mine	White Porphyry	150	Aug. 21, 1880
La Plata	do	20	Aug. 21, 1880
Silver Wave	do	18	Aug. 21, 1880
Smuggler	do	5	Dec. 20, 1879
Tucson	do	15	Dec. 20, 1879
Dome Hill.			
Rock and Dome	White Porphyry	25	Aug. 21, 1880
Yankee Hill.			
Chieftain	White Porphyry and limestone	8	Aug. 21, 1880
Scooper	Gray Porphyry and limestone.	10	Dec. 20, 1879
Breece Hill.			
Breece Iron	Quarry in Gray Porphyry		
Colorado Prince	Quartzite	25	Aug. 21, 1880
Highland Chief	Gray Porphyry	100	Aug. 21, 1880
Little Prince	do	5	Dec. 20, 1879
Miner Boy	Limestone	5	Dec. 20, 1879
Little Ellen Hill.			
Little Ellen	White Porphyry	10	Aug. 7, 1880
Virginius	Blue Limestone	5	Dec. 20, 1879
Long and Derry Hill.			
Belcher	Limestone	5	Dec. 20, 1879
Long and Derry	White Porphyry	5	Dec. 20, 1879

The aggregate daily output of the mines whose output has not been indicated in Table I reaches 30 tons, and the average daily output for all the mines may be said to reach from 700 to 800 tons during the year. It will be seen that the smelting capacity of the camp of Leadville is about 700 tons per 24 hours, so that the margin left for the shipment of ore is rather small.

ORES.

In Leadville the ore deposits are almost invariably found in limestone, which they apparently have replaced. Table I shows that the ore deposits are only reached through masses of various porphyries, and occasionally, at the outcrops, through limestone.

Description.—The ores of Leadville, composed chiefly of carbonate of lead or cerussite and of galena or sulphuret of lead, are divided into two great classes, the hard carbonates or lumps and the sand or soft carbonates, and each class is subdivided and designated by letters or numbers, according to the assay contents and value. In these ores silver exists chiefly in the state of chloride and of chloro-bromo-iodide. Some of the constituents of the ores have been found in an isolated state; pyromorphite, or chloro-phosphate of lead, and wulfenite, or molybdate of lead, in the Little Chief mine; anglesite, or sulphate of lead, in most mines; silicate of lead in small reddish crystals; an as yet unknown mineral in the Evening Star mine (this mineral was found by Mr. Emmons and examined by myself, but the quantity was not sufficient to make a complete examination);[1] cerargyrite, or chloride of silver, in the Chrysolite mine, and embolite, or chloro-bromide of silver, in most mines; shapbachite, or sulphuret of bismuth, lead, and silver, in the Florence mine; and bismuthiferous lanarkite, or sulfato-carbonate of lead and bismuth, in the same mine.

Chief ores.—The following description will show what are the chief ores in the principal mines:

Adelaide.—Large crystals of cerussite, cemented by coarse clay.

Agassiz.—Sand. Light-yellow ocher.

Belcher.—Hard. Compact masses of mixed oxides of iron and manganese, impregnated with small and indistinct crystals of galena and cerussite.

Catalpa.—Hard. Flinty-looking masses of even grain, impregnated with indistinct cerussite crystals. This is the typical "hard carbonate" of the camp.

Chrysolite.—Hard. Masses of indistinct crystals of cerussite, cemented by oxides of iron and manganese, both anhydrous and hydrated; color, brown, reddish, and yellow.

Crescent.—Sand. Pale-yellow ocher and pale-yellow and whitish masses.

Dunkin.—Hard. Fine crystalline galena, imbedded in a hard silicious cement; also distinct and indistinct crystals of galena in a kind of chert.

Dunkin.—Sand. Light-yellow ocher.

Dyer.—Hard. Flint, impregnated with galena.

Evening Star.—Hard. Hard carbonate.

Florence.—Sand. Masses with a dull-blackish tinge (bismuthiferous lanarkite), and also shapbachite, with a metallic luster similar to bismuthinite and stibnite.

Great Hope.—Hard. Hard carbonate.

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[1] It is probably dechenite, which has since been found in determinable quantity by Dr. Iles.

Highland Chief.—Sand. Masses of indistinct crystals of cerussite, not cemented.

Homestake.—Hard. Masses of galena and pyrites in small crystals, cemented with dolomite and siderite.

Hibernia.—Sand. Soft clay, with indistinct cerargyrite, but no cerussite.

Iron.—Hard. Crystalline masses of distinct and indistinct crystals of cerussite, cemented with oxide of iron; part of the carbonate of lead colorless, part with a blackish tinge (in examining the lanarkite of the Florence mine it was found that this tinge is due to sulphuret of silver); also hard masses of pure galena in large crystals, cemented by small ones; also masses of large and small distinct crystals of cerussite, cemented by manganiferous oxide of iron; also masses of hematite and occasionally red and yellow ocher; also nodules of cerussite, cemented with ocher.

La Plata.—Sand. Bluish-black and yellow masses of small crystals of cerussite.

Matchless.—Sand. Soft silicious masses impregnated with cerussite.

Morning Star.—Hard. Hard carbonate.

Rock.—Hard. Masses of crystalline cerussite, with yellow spots of ocher and greenish spots of embolite; also most varieties previously described.

Robert E. Lee.—Sand. Chloride ore, ocherous yellow (a qualitative examination of this ore showed that the ocher contains a considerable quantity of antimoniate of iron); the silver exists in the state of embolite, containing a very small quantity of iodide of silver.

Virginius.—Hard. Uniform masses impregnated with cerussite.

It must be noticed that the principal varieties described here are to be found in most mines, although the bismuth ores are confined to the Florence mine, and those free from lead, or dry ores, are found mainly in the Lee and Hibernia mines. A glance at the assays of ores made at different smelters will give a correct idea of the relative contents of the ores in lead and silver, and also in gangue and iron, so that these assays have a mineralogical and geological signification, as well as a commercial one

COMPOSITION OF ORES.

A complete examination of the ores from every mine would have formed a most interesting chapter, but the field of investigation opened in Leadville was so vast that its accomplishment would have involved a far greater time than could be given to this study, which has therefore been perforce restricted to the most important points.

Carbonate ores.—The interesting observation was made that when cerussite is tinged with black this color is due to sulphuret of silver, so that to some extent the richness of the ore in silver can be ascertained by the eye. The following exhaustive analyses of the carbonate ore of two of the principal mines of Leadville were made by Dr. W. F. Hillebrand in the laboratory of the Survey at Denver:

Analysis I.—Adelaide ore. Sand in lumps, formed of masses of small crystals of colorless cerussite, cemented by cerussite, clay, and oxide of iron.

Analysis II.—Little Chief ore (sand in lumps). Formed of masses of distinct and indistinct small crystals of cerussite with a bluish-black tinge, full of whitish and yellowish spots and cavities.

ANALYSIS I.—*Adelaide ore.*

Oxide of lead	79.550
Chloride of lead	0.990
Peroxide of iron	0.467
Alumina	0.444
Protoxide of iron	0.299
Protoxide of manganese	0.137
Oxide of cobalt	Trace
Lime	0.303
Magnesia	0.068
Gold	Trace
Silver	0.009
Arsenic	Trace
Water	0.395
Phosphoric acid	1.532
Sulphuric acid	Trace
Silica	0.651
Carbonic acid	14.700
Total	99.545

(Hillebrand.)

Gold, $\frac{1}{20}$ of an ounce to the ton.

Silver, $2\frac{1}{2}$ ounces to the ton.

Discussion.—This analysis shows that the Adelaide ore is composed of:

Carbonate of lead (PbO, CO_2)	85.605
Pyromorphite 3 ($3PbO, P_2O_5$) + $PbCl_2$	9.748
Carbonates, silicates, and oxides	4.647
	100.000

All the constituents of these minerals, which are widely distributed in Leadville, will be traced later in the analytical study of the smelting process.

ANALYSIS II.—*Little Chief ore.*

Oxide of lead	75.408
Peroxide of iron	1.940
Alumina	1.415
Oxide of cobalt	Trace
Protoxide of manganese	0.074
Peroxide of manganese	1.386
Oxide of zinc	0.095
Lime	0.335
Magnesia	0.056
Silica	1.970
Arsenic acid	Trace
Antimonic acid	0.121
Water	1.140
Carbonic acid	14.251
Phosphoric acid	Trace
Sulphuric acid	0.486
Chlorine	0.288
Silver	0.777
Gold	Trace
Loss	0.258
Total	100.000

(Hillebrand.)

Silver, 226.62 ounces to the ton.

Discussion.—In this ore pyromorphite is replaced by anglesite, the sulphuric acid of which plays an important part in smelting. Some of the constituents of this ore—arseniates, sulphates, and antimoniates—will be found in an almost unaltered condition in some of the furnace products (accretions). The presence of zinc gives a special interest to this analysis, for the reason that this metal is found in every furnace product—lead fumes, accretions, bullion, and slags. But, although the two preceding analyses are typical of the carbonate ores and contain their most important elements, they do not show the metals molybdenum, titanium, bismuth, nickel, cadmium, copper, and tin, which have also been found in the furnace products, and appear to be pretty widely distributed in the camp, though generally in very small quantities. It is, however, evident that certain substances are entirely wanting in some ores. For instance, a special examination of the Little Chief Smelter slag failed to detect more than traces of titanic acid, showing that this substance does not exist in the Little Chief ore, although it is to be found in most Leadville slags.

Chloride ores.—In Leadville the chlorides of silver have long been known as cerargyrite and embolite, of which they have the characteristic appearance. The writer made a special qualitative examination of the embolite found in the Robert E. Lee mine (old workings). This embolite is disseminated in the peculiar antimoniferous ocher already spoken of. One specimen gave chlorine and bromine, with a little iodine; another gave less iodine than the preceding, but more bromine. Dr. W. F. Hillebrand has made the following analyses of the chief chlorides of Leadville:

ANALYSIS III.—*Robert E. Lee (chloride ore).*

Chloride of silver	21.580
Bromide of silver	77.986
Iodide of silver	0.425
	100.000

ANALYSIS IV.—*Amie (chloride of silver contained in hard carbonate).*

Chloride of silver	15.755
Bromide of silver	84.091
Iodide of silver	0.154
	100.000

The formula of this chloride is $4AgBr + AgCl(I)$, a very small quantity of iodine replacing chlorine.

ANALYSIS V.—*Big Pittsburgh ore.*

Chloride of silver	99.965
Bromine of silver	None
Iodide of silver	0.035
	100.000

At the Chrysolite mine a magnificent block of very pure cerargyrite was found, weighing several hundred pounds.

In order to determine the relative proportions of chlorine, bromine, and iodine throughout the district, a mixture was made of lead fumes collected in the dust chambers of eight smelters. The following figures, obtained by analysis of this mixture, may be taken to represent the relative proportions of chlorine, bromine, and iodine in the silver ores:

Chloride of silver	89.10,	equivalent to chlorine	82.45
Bromide of silver	10.45,	equivalent to bromine	16.83
Iodide of silver	0.45,	equivalent to iodine	0.72
	100.000		100.000

Special stress has been laid upon the composition of the chloride ores, for the reason that they play an important part in lead smelting in Leadville. To chlorine, bromine, and iodine is due a great part of the loss in lead, not only because chloro-bromo-iodide of lead is a very volatile compound, but also because chloro-bromo-iodo-phosphates and sulphurets of lead are found which are also remarkable for their great volatility.

Average of ores.—Mr. Th. Fluegger, assayer of the Harrison Reduction Works, in Leadville, has published in the Engineering and Mining Journal of March, 1880, an analysis of a sample from 1,000 tons, representing specimens from every producing mine in Leadville. This analysis, made with probably insufficient means in the laboratory of one of the smelters, has evidently no pretension to scientific accuracy, since some of the elements—sulphur, arsenic, antimony—are left uncombined and since all the rare elements are not indicated. It is given here, however, because upon it have been based the main features of the chemical discussion of the blast furnace.

If it is assumed that the quantity of silver reported in this analysis is correct, it represents an average quantity of silver of nearly 90.5 ounces to the ton. This figure appears exaggerated, for the reason that the proportion of silver to lead is one ounce to five pounds, while in practice mixtures aimed at contain one ounce of silver to six pounds of lead. But the average percentages of lead (23), iron (18), and silica (22.5) agree precisely with the general composition of the smelting charges in Leadville.

ANALYSIS VI.—*Average ore.*

Carbonic acid	5.58
Oxide of lead	24.77
Silica	22.59
Sulphur	0.90
Protoxide of iron	0.89
Peroxide of iron	24.86
Protoxide of manganese	4.03
Silver	0.31
Lime	2.36
Magnesia	3.04
Arsenic	0.01
Antimony	0.02
Potash and soda	0.98
Chlorine	0.09
Water	5.53
Alumina	3.99
Gold, copper, zinc	Trace
	99.95

Silver, 90.5 ounces to the ton; lead, 23 per cent.; iron, 18 per cent.; silica, 22.59 per cent.

Assays of various ores.—The assays of the first division of the following table were made by the writer in June, 1880, for one of the smelters. The silver assays present a certain interest, since they were made in crucibles, whereas the ordinary method of silver assay at Leadville is that by scorification; the lead assays, however, are the ordinary fire assays usually made in the region.

The assays given in the rest of the table were made by the different assayers attached to the respective smelting works.

TABLE II.

I.—ORE ASSAYS.

Name of mine.	Lead.	Silver to the ton.	Gold to the ton.
	Per cent.	*Ounces.*	*Ounces.*
Adelaide (sand)	15. 00	47. 70	None
Do	15. 00	30. 00	0. 85
Do	26. 45	8. 80	0. 05
Do	54. 90	12. 60	0. 10
Do	21. 45	12. 00	0. 10
Do	9. 45	34. 15	0. 05
Amie	45. 80	23. 30	....
Do	0. 50	36. 00	....
Do	11. 30	78. 00	....
Amie (special lot)	31. 70	1, 189. 70	....
Amie (black lumps)	1. 70	31. 40	....
Amie (lumps)	1. 50	31. 20	....
Amie	42. 00	217. 30	....
Do	7. 00	194. 85	....
Do	2. 00	37. 35	....
Amie (lumps)	5. 80	213. 80	....
Amie	6. 70	119. 95	....
Do	0. 50	76. 80	....
Do	17. 00	230. 90	....
Do	1. 90	36. 25	....
Belcher	5. 65	33. 10	....
Do	6. 50	39. 50	....
Chrysolite	19. 30	235. 30	....
Do	27. 50	55. 90	....
Evening Star	44. 35	99. 05	....
Hibernia (clayish)	1. 80	33. 05	....
Little Giant	28. 00	41. 10	....
Do	40. 00	62. 90	....
Do	54. 00	87. 70	....
Virginius	25. 20	8. 10	....
Do	34. 50	22. 50	....

II.—ASSAYS MADE AT ELGIN SMELTER, AUGUST, 1880.

Amie	14. 00	180. 00	....
Camp Bird	46. 00	12. 00	....
Carbonate	30. 00	197. 00	....
Catalpa	33. 00	80. 00	....
Evening Star	37. 50	55. 00	....
Iron	33. 70	45. 00	....
Little Chief	35. 00	65. 00	....
Loveland	46. 50	85. 00	....
Morning Star	48. 00	56. 00	....
Pine	28. 00	12. 00	....
R. E. Lee	None	185. 00	....

TABLE II—Continued.

III.—ASSAYS MADE AT MESSRS. CUMMING & FINN'S SMELTER, JULY, 1880.

Name of mine.	Lead.	Silver to the ton.	Gold to the ton.
	Per cent.	*Ounces.*	*Ounces.*
Adelaide	22 to 44	12 to 20	0. 5 to 0. 75
Amie	2 to 10	20 to 1,100	
Chrysolite	27. 00	40 to 80	
Evening Star	22. 00	51. 00	
Hibernia	.5 to 1	60 to 180	
Little Giant	12 to 40	14 to 80	
Morning Star	40 to 55	35 to 40	
Virginius	25 to 35	8 to 32	

IV.—ASSAYS MADE AT CUMMING & FINN'S SMELTER, AUGUST, 1880.

Amie (lumps)	3. 00	40. 00	
Amie (screenings)	None	100. 00	
Hibernia (clayish)	None	33. 00	
Do	None	60. 00	
Homestake	None	70. 00	
Do	8. 00	60. 00	
Morning Star (sand)	55. 00	38. 00	
Morning Star (hard)	40. 00	32. 00	
Do	40. 00	36. 00	
Morning Star (sand)	47. 00	30. 00	
Do	55. 00	38. 00	

V.—ASSAYS MADE AT THE HARRISON REDUCTION WORKS, JULY, 1880.

Agassiz	14. 00	49. 60	
Alpine	45. 00	141. 00	0. 50
Amie	23. 50	165. 00	
Camp Bird	48. 70	12. 00	
Carbonate	27. 00	218. 00	
Chieftain	7. 00	76. 00	
Chrysolite	29. 00	97. 50	
Colorado Prince	None	15. 00	17. 70
Double Decker	None	35. 50	2. 70
Dunkin	8. 00	119. 00	
Forsaken	13. 50	62. 00	
General Shields (Sawatch Range)	None	53. 50	0. 10
Gold Cup (Sawatch Range)	None	110. 00	
Gold Ore (Sawatch Range)	None	6. 70	4. 10
Independence (Sawatch Range)	None	8. 60	5. 76
Iron	40. 50	79. 00	
Long and Derry	16. 00	73. 00	
Little Chief	16. 00	56. 50	
Little Pittsburg	36. 40	266. 00	
Morning Star	65. 00	61. 00	
Nevada	23. 70	16. 00	2. 20
Pine	31. 00	12. 00	
Ready Cash	3. 70	125. 00	9. 20
Robert E. Lee	None	146. 50	

At the smelters, silica, or rather that mixture of silica and refractory silicates insoluble in acids, and known as gangue, is determined, as well as the per cent. of iron.

The following will give an idea of their relative proportions:

TABLE II—Continued.

VI.—ASSAYS MADE AT THE CALIFORNIA SMELTER, JULY, 1880.

Name of mine.	Lead.	Silver to ton.	Iron.	Gangue.
	Per cent.	*Ounces.*	*Per cent.*	*Per cent.*
Amie	5.00	40.00	40.00	20.00
Brian Boru	40.00	40.00	6.00	12.00
Iron	24.00	30.00	24.00	46.00
Morning Star	55.00	40.00	5.00	16.80
Robert Emmet	10.00	12.00	49.00	12.00
Tucson	50.00	30.00	10.00	10.00

VII.—ASSAYS MADE AT THE HARRISON REDUCTION WORKS, JULY, 1880.

Chrysolite	29.00	97.50	8.35	11.36
Climax	Not determined		37.10	30.80
Iron	40.00	79.00		19.40
Little Chief	16.00	56.50	21.00	17.30
Robert E. Lee	None	146.50	17.30	49.40
Rock			22.10	47.00

VIII.—ASSAYS MADE AT THE GRANT SMELTING WORKS, JULY, 1880.

Catalpa (hard)	33.60	79.00	4.25	41.30
Do	45.60	64.60	5.40	18.50
Catalpa (sand)	43.70	83.00	39.50	27.90
Chrysolite (B, hard)	21.90	47.75	25.30	16.40
Chrysolite (A, sand)	42.50	75.00	13.70	14.60
Dyer	1.00	80.00	3.70	58.30
Evening Star			7.00	49.00
Do	23.90	44.50	6.80	14.20
Henrietta	38.80	45.60	45.20	8.60
Hibernia	None	52.22	25.70	22.10
Highland Chief	15.40	144.00	2.50	15.10
Iron	33.70	20.90	11.70	12.60
Little Chief	20.00	55.00	22.40	17.00
Do	33.30	116.80	13.00	14.30
Little Chief (hard)	10.00	25.00	29.75	12.50
Little Chief	18.80	99.00	14.40	33.60
Little Chief (sand)	37.50	80.00	14.70	17.50
Little Chief	9.35	35.72	19.90	28.30
Little Chief (sand)	27.50	100.00	15.75	27.50
Little Chief (galena)	55.00	5.50	3.85	15.00
Little Pittsburgh			18.20	26.50
Morning Star	42.30	18.60	11.00	14.20
New Discovery			17.15	33.70
Silver Wave	7.15	35.35	49.30	17.70
Do	5.95	36.52	32.90	12.80

TABLE II—Continued.

IX.—ASSAYS MADE AT MESSRS. BILLING & EILERS'S SMELTER, JULY, 1880.

Name of mine.	Lead.	Silver to ton.	Iron.	Gangue.
	Per cent.	*Ounces.*	*Per cent.*	*Per cent.*
Amie			25.00 to 45.00	12.00 to 25.00
Chrysolite (hard)	20.00 to 25.00	65.00	23.50	15.00
Dome			23.00	16.00
Dunkin (sand)	1.00 to 13.00	50.00 to 100.00	34.00	32.50
Iron (hard)	50.00	80.00		
Rock (hard)	36.00	18.00	18.10	6.50
Rock (sand)			18.15	12.50
Average assays of all the ores smelted from June, 1879, to June, 1880	30.27	62.58		

X.—ASSAYS OF VARIOUS ORES MADE AT MESSRS. CUMMING & FINN'S SMELTER, JULY, 1880.

Adelaide (sand)	44.00	20.00	8.00	15.00
Amie (lumps)	3.00	40.00	35.00	13.00
Amie (first class)	8.50	300.00	33.50	24.00
Amie (second class)	4.50	30.00	35.00	18.00
Chrysolite (sand)	34.00	72.00	23.60	16.00
Chrysolite (hard)	16.00	45.00	26.00	18.00
Evening Star	22.00	51.00	5.00	52.00
Hibernia	None	33.00	28.00	29.00
Do	None	60.00	20.00	48.00
Homestake	5.00	70.00	9.00	57.00
Do	8.00	60.00	14.00	36.00
Little Giant (first class)	37.00	84.00	4.00	32.00
Little Giant (second class)	19.00	38.00	18.00	14.00
Little Giant (third class)	12.00	17.00	34.00	19.00
Morning Star (hard)	43.00	39.00	4.00	35.00
Morning Star (sand)	55.00	38.00	5.00	20.00
Morning Star	53.00	53.00	3.50	20.50
Morning Star (sand)	55.00	38.00	5.00	20.00
Morning Star	53.50	27.00	6.70	14.00
Morning Star (hard)	40.00	36.00	6.00	24.00
Morning Star	7.00	30.00	5.00	20.00
Virginius	34.00	22.00	15.00	22.00

XI.—ASSAYS FOR IRON AND GANGUE MADE AT MESSRS. CUMMING & FINN'S SMELTER, JULY, 1880.

Name of mine.	Iron.	Gangue.	Name of mine.	Iron.	Gangue.
	Per cent.	*Per cent.*		*Per cent.*	*Per cent.*
Adelaide	17.00	14.00	Amie	38.10	25.10
Do	16.50	15.00	Do	37.50	24.00
Chrysolite	21.00	20.00	Evening Star	4.50	60.30
Do	19.80	11.00	Forsaken	25.00	32.00
Do	13.50	20.00	Hibernia	38.00	29.50
Do	26.30	23.50	Little Giant	14.20	34.60
Evening Star	13.70	45.00	Do	5.40	31.80
Do	12.70	32.00	Do	12.50	45.50
Do	5.45	42.50	Do	6.00	29.00

TABLE II—Continued.

XII.—ASSAYS OF LOTS AND MIXTURES OF ALL SORTS MADE AT THE LA PLATA SMELTER IN THE YEAR 1880.

Name of mine.	Iron.	Gangue.	Name of mine.	Iron.	Gangue.
	Per cent.	*Per cent.*		*Per cent.*	*Per cent.*
11 lots	23.90	20.30	7 lots	20.70	18.60
14 lots	22.30	19.40	11 lots	17.90	18.80
11 lots	23.70	32.00	Average of 64 lots	21.50	19.70
10 lots	20.40	19.60			

Discussion.—The preceding tables are valuable as furnishing, not only data for reference, but also proofs of the activity of mining and smelting in Leadville. It is also evident from their examination that there is no relation whatever between the lead and silver contents of the ores. This could scarcely be otherwise, if it is considered that lead exists in the state of carbonate or sulphide, and silver in the state of sulphide or chloro-bromo-iodide, compounds which have no common properties. A carbonate ore rich in lead may contain a large quantity of residual or untouched sulphide of silver, and be rich in silver, or its silver may have been carried away in the state of chloride and the ore be poor in silver. This chloride of silver carried away may be redeposited in any kind of mineral, in porous quartz or in clay, and the ore may be very rich in silver and contain no lead. In other cases both carbonate of lead and chloride of silver are carried away and deposited in the same gangue, giving ore rich in both lead and silver.

SMELTING WORKS.

Location.—Since Leadville became an important mining camp sixteen distinct smelting works have been erected. Two smelters only are situated in Leadville proper, the Harrison Reduction Works and the Grant Smelting Works, which both stand on the northern bank of California gulch. In the outskirts of the city, and at the junction of the upper and the lower roads of this bank, stood the Leadville smelter, now pulled down. Then come in succession, but still on the northern bank of California gulch, the La Plata, the American, Billing & Eilers's, and the California Smelting Works. At the lower end of California gulch is situated the small town of Malta, near which were erected the Malta and Lizzie smelters. In Adelaide, on Iron Hill, stood the Adelaide smelter, which belonged to the Adelaide mine, but has long since ceased running. On Fryer Hill and immediately above the Little Chief mine stood the Little Chief smelter. This smelter has since been pulled down on account of the sinking of the ground upon which it was erected, and its furnace is now running at Messrs. Cumming & Finn's smelter. On the southwestern bank of Big Evans gulch are found in succession, going westward, the Ohio and Missouri, Cumming & Finn's, Gage, Hagaman & Co.'s, Raymond, Sherman & McKay's, and the Elgin Smelting Works.

At the time this report was made (August, 1880) several smelters had entirely ceased running, viz, the Adelaide, Little Chief, American, Malta, Lizzie, Leadville, Gage, Hagaman & Co.'s, and Raymond, Sherman & McKay's. Since that time the American and Malta have resumed work, the former successfully, its plant being in a perfect state of preservation; the latter rather unsuccessfully, mainly by reason of its imperfect plant and machinery.

The following are the names of the different smelting works, of the superintendents, and the dates at which they commenced smelting:

Smelting works.	Superintendent.	Date.
Harrison Reduction Works	James Brierton	Oct., 1878
Grant	J. B. Grant	Sept., 1878
Leadville		—, 1877
La Plata	M. E. Smith	June, 1878
American	Carl Heinrich	May, 1879
Billing & Eilers's	Fritz Wolf	May, 1879
California	M. E. Smith	June, 1879
Malta	F. Fohr	—, 1875
Lizzie		—, 1876
Little Chief	S. T. Tyson	Aug., 1879
Ohio and Missouri	N. R. Wilson	June, 1879
Cumming & Finn's	Thomas MacFarlane	July, 1879
Gage, Hagaman & Co's		June, 1879
Raymond, Sherman & McKay's		June, 1879
Elgin	Aug. Werner	June, 1879
Adelaide		—, 1879

Cost of plant.—From data obtained by the census investigation of the precious metals for 1880, carried on under the supervision of Mr. S. F. Emmons, it is found that the smelters of Leadville were erected at an aggregate cost of about $800,000.

In the following table will be found the cost of plant of twelve of the principal smelters in arbitrary numerical order:

TABLE III.

	Cost of plant.
Smelter I	$43,000
II	160,000
III	80,000
IV	25,000
V	37,000
VI	65,000
VII	95,000
VIII	60,000
IX	48,000
X	37,000
XI	30,000
XII	20,000

General disposition of smelting works.—The slope of the banks of gulches is particularly favorable to the construction of smelters, most of which are divided into several levels, which allow of a rational division of labor and economize constructions, hoisting machinery, and manual labor.

On California gulch most smelters have the following levels:

1. The furnace, slag-heap, and bullion level, connected by means of inclined ways with the main lower road of the gulch running at the foot of the slag-heap.
2. The feeding-floor level, which is also that for crushing, sampling, ore-beds, and ore-bins. This level is always provided with a wagon road, branch of the upper road, and communicates with the upper and lower levels by means of inclined ways.

3. First ore-bin level, with a wagon road between the rows of bins, allowing the discharge of ore-wagons into the lower row of ore-bins, and the wheeling away of the ore extracted in barrows from the upper row of ore-bins to the feeding and crushing level. This level communicates, like the preceding, with the upper and lower levels by means of inclined ways.

4. Second ore-bin level, with disposition similar to preceding.

5. Third ore-bin level, with disposition similar to preceding.

6. Charcoal and coke-bin level, with dispositions similar to preceding.

7. Ore-dumps, fluxes, charcoal, coke and wood reserves, or upper level. A glance at Figure 2, Plate XXXI, representing smelter C, will give an idea of the disposition of levels.

The general arrangement just described is that adopted at the most favorable points, but sometimes levels 3, 4, 5, and 6 are reduced to three, two, and even one level. This is particularly the case on Big Evans gulch, whose banks are far from being as high as those of California gulch. There the levels are reduced to two:

1. The furnace, slag-heap, and bullion level.

2. The feeding-floor level, used also for crushing, sampling, ore-beds, and ore-bins. The fuel-bins, ore-dumps, fluxes, and wood reserves are generally placed at the back of the ore-bins.

The works are always inclosed, from the furnace level to the back of the feeding floor, in a light wooden structure. Where there are several levels of ore-bins they are independent of the main building; but where there is only one upper level the ore-bins are placed in the building. The offices and laboratory always occupy a detached building. The office is always provided with large wagon scales, varying in capacity from 10 to 20 tons, and used for weighing the wagons loaded with ore or bullion and taring them after unloading.

The boilers, engines, and blowers are always placed on the furnace level, on one side of the furnaces, as are the smith's and mechanic's shops, which, however, often occupy a small detached building.[1]

ORE BUYING.

Method.—The manner in which ore is purchased by the smelters of Leadville is somewhat different from the method usually pursued in other camps. The ore is purchased outright for cash from the mines, a certain deduction being made for the loss of silver in smelting, and a certain amount being charged for what is called the cost of treatment.

In the following table is shown, as a sample, for a few of the principal mines—

1. The deduction for the loss of silver in smelting.
2. The cost of treatment.
3. The price given for the lead contained in the ores.

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[1] At the time this report was made all the smelters in California gulch were connected by side tracks with the railroad, and preparations were made to connect the railroad with the smelters on Big Evans gulch.

Unless the ore contains more than a certain percentage (5 per cent. to 30 per cent.), the lead is not paid for at all.

Mine.	Deduction for loss of silver in smelting.	Cost of treatment per ton of ore.	Price paid for lead per unit of 20 pounds.
	Per cent.		
Amie	10	$25	$0.25
Carbonate	7½	20	25
Chrysolite	5	20	25
Dunkin	5	22	25
Evening Star	7½	28	25
Iron Silver	5	18	30
Morning Star	5	15	30
Tucson	5	21	25

These rates are subject to constant fluctuation, according to—

1. The price of fluxes.
2. The amount of fluxes required in smelting.
3. The price of charcoal and coke.
4. The character of the ore: whether large lumps or sand; whether highly sulphureted or highly silicious; whether rich or poor in lead; whether rich in oxide of iron or without it.

The cost of treatment has varied during the year ending June 1, 1880, from $15 to $30 per ton of ore. The price paid for silver and lead in the ore varies naturally with the New York market. During the year ending June 1, 1880, the variation for silver has been from full New York quotations and no discount to a discount of 10 per cent., the average discount having been about 5 per cent. off silver quotations.

Lead is bought by the unit, i. e., 1 per cent., or 20 pounds in the ton; and its price has varied from 15 cents to 45 cents per unit during the year 1879–'80. The price per unit of lead depends on individual agreement, and also on the contents of the ore in lead. At some smelting works the cost of treatment will be $16 to $25, with a deduction of 5 per cent. off silver, and the price of lead 20 cents to 25 cents per unit when the ore contains above 30 per cent. At others, the cost of treatment will equal $20, the deduction off silver 5 per cent., and the price for lead 15 cents per unit when the ore contains above 5 per cent.

Gold is paid for at the rate of $18 per ounce, but only when its amount exceeds one tenth of an ounce per ton of ore.

Cost of transportation.—When the ore is bought direct from the mine, its transportation is paid for by the mine owners, and the cost of handling varies from $1 to $1.85 per ton of ore, according to distance; but when the ore is purchased at the sampling works, the smelters have to pay for its transportation to their bins at the above rates.

SAMPLING.

Method.—The general method of sampling carried on in the camp is the following: In shoveling the ore from the ore-wagon to the ore-bin, every tenth shovelful is thrown aside into a wheelbarrow. Thence the sample thus obtained is wheeled to the sampling-floor and passed through the crusher. It is then well mixed with the shovel, laid in a thin layer on the floor, and quartered down very carefully until small enough to be dried easily. The amount of moisture is determined by desiccation of this sample, previously weighed. When dry it is passed through Cornish rolls set to one-eighth of an inch or through small mills. It is once more well mixed and quartered down until small enough to be ground on the buck-plate or in the mortar, and passed through fine sieves, about 70 meshes to the linear inch. This done, the sample is once more well mixed and divided into three parts, one of which is assayed by the smelter, the other at the mine, and the third by an independent assayer, or more generally kept in reserve for reference in case of dispute. Sometimes the bulk of the sample obtained from every tenth shovelful from the wagon is reduced by setting apart every fifth shovelful. This reduced sample is afterwards subjected to the treatment which has just been described in detail.

Sampling works.—Every smelter in Leadville possesses a sampling floor, with ore-beds, crushers, and ore-bins; but there are besides three large sampling works, which are independent of the smelters and where the buying, assaying, crushing, drying, sampling, and selling of ore only are carried on. These works belong to Messrs. A. R. Meyer & Co., Eddy & James, and Gillespie & Ballou.

The sampling works are provided with a large number of bins for the preparation and classification of ores of every grade and from every mine; and, as at the smelters, the machinery, crushers, Cornish rolls, and mills are driven by steam-power. Large open spaces are kept for the accumulation of ore-dumps and the preparation of ore-beds of a given composition. These are made by spreading layer upon layer of ores of known weight and contents in silver and lead. Drying is carried on on a large scale, the driers consisting of large parallelopipedic cuts in the ground, about six feet wide and twenty feet long, provided with a coal fire-place at one end, connected with a sheet-iron stack at the other, and covered over on a level with the ore floor with sheet-iron, upon which the ore to be dried is spread in layers.

The advantages offered by these works are twofold. The prospectors and small miners can always dispose of their small lots of ore, and the large ones of those ores which are in any way exceptional or out of the usual run. On the other hand, the smelters can always find their supplies of ores of a given composition ready for the furnace, or special ores to modify or complete the composition of their own ore-beds or mixtures.

CRUSHING.

Sand ores do not require crushing; in fact, they are already in dust or pieces too small for the furnace, and require mixing in convenient proportion with crushed ore in order to be fit for use. But hard ore and sand ore in lumps require crushing, as well as the limestone, iron-stone, and old slags which are used as fluxes. This is effected, both at smelters and at sampling works, by means of compact but powerful stonebreakers or crushers, always driven by steam-power.

Machines used.—The crushers mostly used in Leadville are Blake crushers manufactured by the Blake Crusher Company, New Haven, Conn., and by the Farrel Foundry and Machine Company, Ansonia, Conn. At the sampling works one or two Alden crushers manufactured by E. T. Copeland, New York, are also in use.

The table below gives the principal types used, their numbers, nominal horse-power required, and capacity.

Capacity numbers of the crushers.	Opening between jaws.	Revolutions of fly-wheel per minute.	Horse-power required.	Crushing capacity per hour.
	Inches.			*Tons.*
No. 1	10 by 4	300	4	4
No. 2	10 by 7	275	6	6½
No. 4	10 by 4	300	4	4
No. 5	15 by 9	275	9	9⅓
No. 0	10 by 4	300	4	4
No. A	10 by 4	300	4	4

Blake crushers.—The crushers manufactured by the Blake Crusher Company belong to two styles: (1) The older style or eccentric pattern and (2) the Challenge Rock-Breaker, or Sectional Cushioned Crusher. Of the eccentric pattern, a horizontal and vertical section will be found in Plate XLI, Figures 1 and 2; the drawing given is a copy of that furnished by the company for a No. 2. The circle *D* is a section of the fly-wheel shaft, which should make from 225 to 250 revolutions per minute. The dotted circle *E* is a section of the eccentric. *F* is a pitman or connecting-rod, which connects the eccentric with the toggles *G G*, whose bearings form an elbow or toggle-joint. *H* is the fixed jaw; this rests against the end of the frame *A*. *P P* are chilled iron plates, between which the rock is crushed. When worn at the lower end they can be inverted and thus present a new wearing surface. The cheeks *I I* fit in recesses on each side and hold the chilled plates *P P* in place. By changing the position of the cheeks from right to left when worn, both will have a new surface. *J* is the movable jaw. It is supported by the round bar of iron *K*, which passes freely through it and forms the pivot upon which it revolves. *L* is a spring of india-rubber, which is compressed by the forward movement of the jaw and assists its return. *M M* are bolt-holes. *B* is the fly-wheel. *C* is the driving-pulley. *Q Q Q Q* are oiling tubes; *R R R R*, steel bearings; *O*, the toggle-block; *N*, the wedge; *Y*, the wedge-nut; *S*, set-screws for tightening toggle-block; *T*, bush and key. The frame *A A* and supports *Z Z* are made of cast iron. This crusher is being gradually superseded by the Challenge Rock-Breaker, manufactured by the same firm, which has many points of superiority over the preceding, and is not quite so delicate in construction or so apt to get out of order.

The Challenge or sectional cushioned crusher is represented in perspective and vertical section in Plate XLI, Figures 3 and 4, which are copied from the company's drawing of a No. 5 crusher. Its crushing capacity per hour is 9 tons when the jaws are set 1½ inches apart and when its speed is 275 revolutions per minute. Flint, hard ores that break with a snap, dolomite, hematite, and old slags go through the crusher at that rate in the same conditions, but with sand or soft ore the capacity is sensibly diminished. The 9 horse-power indicated as being necessary to drive this crusher is purely nominal, and represents, so to speak, an average; in practice the driving engine should have greater power, in order to overcome irregular or unexpected resistance.

The Challenge crusher consists of a three-sided frame-work *F*, of cast iron, with a broad flanged base, holding the movable jaw in suspension, which forms the front part of the machine, between the upright convergent jaws of which the stone is crushed.

The jaw-shaft K is held in place by wrought-iron or steel clamps C, which serve to take part of the strain due to crushing in the upper part of the jaw space, and also serve as walls thereof. In the lower part of the three-sided frame, or front part of the crusher, and on each side of it are holes in the casting to receive the main tension rods R, which connect the front and rear part of the machine. The rear part B is called the main toggle block. It is also provided with holes to receive the main tension-rods R R, corresponding to those in the front casting. The tension-rods R R are provided with screw-threads and nuts N N, by means of which their length, and in consequence the opening between the jaws, are readily adjusted to crush coarse or fine.

The front and rear castings are supported on parallel timbers G G, to the under side of which are bolted the boxes carrying the main eccentric shaft, provided with fly-wheels and pulley. These timbers take the transverse strain, which comes upon the pitman connecting the main shaft and the toggle-joint, situated in the rear of the movable jaw, and between it and the main toggle-block. Between the broad flanged bases of the front and rear castings and the timbers on which they rest are placed flat rubber cushions C' C', one-fourth to three-eighths of an inch thick. Every revolution of the shaft brings the toggles more nearly into line and throws the movable jaw forward. It is withdrawn by the rod provided with rubber spring L. In this way a short vibratory movement is communicated to the movable jaw. The pitman R' H is constructed so that it can be lengthened or shortened, and thus change the inclination of the toggles O O, and consequently the length of the movable jaw J.

The great advantage of this machine over the old style is that of possessing elastic parts, rigid enough to allow the performance of the work desired, but giving way under accidental strains, such as the introduction of a steel hammer between the jaws. The frame A is made of timber. The best method of setting up this stone-breaker is to place its frame on four timbers 15 by 15 inches, disposed as is shown at X X' and Y. These timbers are pinned or bolted together.

The following are the main parts of the machine and the letters used to indicate them in the drawing:

A, timber frame.
B, main toggle-block.
C′ C′, rubber cushions.
D, fly-wheel.
E, main pulley.
F, main cast-iron frame.
G, timber supports.
H, pitman half-box.
I, cheeks.
J, movable jaw.
K, jaw shaft.
L, rubber spring.
L′, spring rods.
M, pitman-rod nuts.
N N′, main tension-rod nuts.
O, toggles.
P, jaw (chilled plates).
R, main tension rods.
R′ H, pitman.
R′, pitman-rods.
S, main eccentric shaft.
T, toggle bearings.

The Farrel Foundry and Machine Company's Blake crusher is used a good deal in Leadville. It is constructed on very nearly the same principles as the Blake Crusher Company's eccentric pattern. It presents the same appearance, it requires the same amount of power to produce the same quantity of work in the same time, and a complete description of it would be superfluous, since it answers exactly to the description of the eccentric pattern. It differs from it, however, in one respect, namely the substitution of a crank shaft for the eccentric shaft.

Alden crusher.—The Alden crusher and pulverizer, is not in use at smelters, which, when they have any pulverizing to do, use Cornish rolls; but it is used at the sampling works, where a considerable amount of pulverizing is done. The jaws of this crusher differ essentially from those of the others in this respect—that their grooves are perpendicular to the length of the jaws, while in the others these grooves are parallel to the length.

Fig. 1, Plate XLV, gives a perspective view of the Alden crusher, in which portions of the jaws and jaw-faces are shown in section. The jaws are hung upon wrought-iron trunnions, the ends of which project through and are supported by the sides of the frame. Motion is imparted by links connected with the trunnion ends, and driven by studs projecting from a sliding yoke beneath. This yoke is connected with a crank-shaft by a pitman. The rotation of the crank moves the yoke to and fro on a nearly horizontal plane, alternately moving and pushing the movable ends of the two jaws, and imparting a rubbing motion, which is the main feature of the machine. The jaws may be adjusted at varying distances, so as to obtain a product of varying degrees of fineness.

The Cornish rolls, used by both the smelters and samplers for grinding their samples, consist of two steel cylinders, 12 inches long and 6 inches in diameter, connected by cog-wheels, driven by pulley and transmission belt, and fed by means of a thin sheet-iron funnel, having the shape of an inverted truncated pyramid. These rolls are usually set one-eighth of an inch apart.

ASSAYING.

In Leadville assaying is quite an important branch of the mining and smelting industries. In addition to the assayers attached to all the smelting and sampling works and to the principal mines, there are no less than twenty independent assayers residing in the city and having their own assay offices. Besides being employed as referees and experts in cases of dispute between mines and smelting works, the latter are patronized by the prospectors and small miners.

The chief assays made in the camp are silver, gold, lead, iron, and gangue assays, and at some smelters specific-gravity determinations of slags.

Furnaces.—The laboratories are generally provided with permanent crucible and muffle furnaces, made of common brick, lined with fire brick, and placed side by side, as is shown in Plate XXXIX; but very often the two furnaces are separate.

By means of the dampers D′ and D′ in the chimney, the assayer can regulate the draft and the intensity of heat in the furnaces. The apertures A B C D are closed by means of sheet-iron plates, easily removed by tongs. Occasionally, portable clay furnaces, of American and English manufacture, are used for cupellation.

Pulverization.—The ores and slags are, first of all, coarsely pounded in a cast-iron mortar (Fig. 12, Plate XLIII), a form of mortar that is not well adapted for this use, since it is too thin and very often breaks before the stone does. The coarsely pounded material is then ground on the buck-plate. This consists of a cast-iron plate (Figs. 9 and 10, Plate XLIII), about an inch thick, faced on one side, and provided or not with flanges on each side. It rests on a firm table or timber support. The ore is laid on the plate and ground with the bucker. The bucker (Fig. 11, Plate XLIII) is a mass of cast iron, with a cylindrical lower surface, faced on the plate side, and fixed to a wooden handle. Grinding is performed by placing the left hand on the bucker, holding the handle in the right hand, and moving the bucker forwards and backwards, at the same time lifting and lowering the handle, and exerting a slight pressure with the left hand. While all this is going on the bucker is also moved from the left to the right side, and inversely, so as to increase the grinding surface. All this is much more easily performed than described.

The pulverized ore is then passed through sieves of 70 to 80 meshes to the linear inch, represented in Fig. 8, Plate XLIII, in elevation. The metallic cloth of the sieve is made of brass. It is adjusted to a tinned-iron circular frame, *b*, fitting in a circular tinned-iron box, or dust-receiver, *a*. This is a very convenient arrangement, the loss in dust is very small, and the mixing of the dust takes place at the same time as the sifting.

Crucibles and scorifiers.—Figs. 3, 4, and 5, Plate XLIII, represent the crucibles, scorifiers, and gold-annealing cups, which are manufactured by the Denver Fire Clay Company. The gold-annealing cups and scorifiers are similar to the European ones in appearance, but greatly inferior to them in quality. The assay crucibles, three-sixteenths of an inch thick, are probably the thinnest clay pots used in assaying in any country. They are very convenient for the reason that, with a low temperature in the furnace, the assay fluxes become easily fluid, but they never stand more than two runs in the crucible furnace.

Cupels.—Cupels are always made in the assay laboratories in brass molds, the process being too well known to demand description. Their form and size are shown in Fig. 6, Plate XLIII.

Muffles.—The muffles made by the Denver Fire Clay Company are good. They are generally large enough to hold from 12 to 16 scorifiers, enabling the assayer to assay three or four samples of ore at the same time.

Tools.—The scorifier tongs, cupel tongs, crucible tongs, raking rods, anvils, hammers, chisels, etc., are similar in every respect to those universally used in assaying.

Slag molds.—The molds into which are poured the crucible and scorifier slags are peculiar, and are represented in Figs. 1 and 2, Plate XLIII. They consist of a sheet of cast iron, divided into 12 conical molds. They are very convenient, the lead buttons and slags cooling rapidly on account of the thinness and large surface of the mold.

Fuel.—Coke is used in the crucible furnaces and charcoal in the muffle furnaces, but sometimes coke and charcoal are mixed in the muffle furnaces.

Balances.—Balances capable of weighing from four pounds to one-sixteenth of an ounce are used for the estimation of moisture in the ore; balances weighing from 100 grams to 1 milligram, for the weighing of scorifying and crucible assays; and those sensitive to the tenth of a milligram, for the weighing of silver prills and gold partings. These balances are generally manufactured by Becker & Sons, of New York. They offer no peculiarity in construction.

The *weights* used in assaying are gramme weights for lead, iron, and gangue assays, and silver prills, or gold partings; but the ore, slags, and bullion are weighed in assay tons, whose symbol is A. T., or its subdivisions. The weight boxes contain one-tenth of an assay ton, or $\frac{1}{10}$ A. T., $\frac{2}{10}$ A. T., $\frac{5}{10}$ A. T., 1 A. T., 2 A. T. Some boxes contain besides $\frac{1}{20}$ A. T. and 5 A. T. The system of assay-ton weights introduced by Prof. C. F. Chandler, of the School of Mines, Columbia College, New York, is as simple as it is ingenious. The ton of 2,000 pounds avoirdupois is equal to 32,000 ounces avoirdupois, or to 29,166 ounces troy, or to 907,180,000 milligrams. The weight of the assay ton is 29,166 milligrams, consequently each milligram represents one ounce troy, and 29,166 milligrams represent one ton. When the material to be assayed for precious metals is weighed by the assay ton or its multiples, the weight of the precious metals in milligrams, or multiples of the milligram, corresponding to those of the assay ton, expresses in troy ounces the weight of gold or silver contained in one ton of ore or bullion. A few examples will illustrate this:

1. Twenty-nine thousand one hundred and sixty-six milligrams of bullion, or one assay ton, give after cupellation a button of silver weighing 205.5 milligrams. This shows that one ton of this bullion contains 205.5 ounces troy of silver.

2. One-half an assay ton of slags gives, after assaying, a button weighing 1½ milligrams; this shows that one ton of slag contains 3 ounces troy of silver.

3. One-tenth of an assay ton of ore contains 3 milligrams of silver; this shows that one ton of ore assays 30 ounces troy of silver.

The laboratories are provided also with sand-baths, flasks, beakers, dishes, burettes, and a few of the principal reagents used in assaying by the wet way. Iron and gangue assays are regularly made in the wet way, and occasionally the ore is assayed for sulphur and arsenic, the slags for lead, the ores and fluxes for lime and magnesia.

Silver assays.—The general process used by common consent in Leadville for ore assays is the scorification process, a rapid and accurate method. Some mines, however, require crucible assays. The scorification process is so well known and so fully described in text books that it will not be insisted upon. The assays of each sample are made in three or four scorifiers. One tenth of an assay ton is weighed for each scorifier, and then mixed with ten times its weight, or one assay ton, of pure granulated lead, or rather with a granulated lead whose contents in silver are known and subsequently subtracted from the silver buttons obtained. The silver-prills are weighed to the tenth of a milligram, and each of these divisions corresponds to an ounce to the ton. A little borax is always used to scorify the oxide of iron and other bases. Slag, like ores, is assayed by scorification; but this process ought to be abandoned and the crucible process substituted for it, chiefly for the reason that in the crucible the assay may be made with one assay ton if necessary, this quantity not being excessive for the estimation of 1 or 1½ ounces of silver to the ton. The crucibles used in crucible assays are those drawn to scale in Figs. 3 and 5, Plate XLIII. A mixture of

Powdered ore	½ assay ton.
Litharge	1 assay ton.
Bicarbonate of soda	½ assay ton.
Borax	¼ assay ton.
Argol	$\frac{1}{10}$ assay ton,

or some similar mixture, for each assayer has his favorite flux, is fused in them, in the presence of an iron nail or rod, which, however, some assayers dispense with altogether. The mixture is generally covered with a layer of borax or common salt.

Bullion assays.—The assays are generally made on a car-load sample, representing 10 tons. Two pieces of lead are detached from the top and bottom part of each bar of bullion forming the car-load (in general 400 bars); all these are melted together in a plumbago crucible, under a cover of live charcoal; the charcoal and scum are then removed; the sample, well mixed by stirring, is poured into an ingot mold (a bullion one); the bar obtained is about one inch thick (Fig. 7, Plate XLV). Four pieces are detached from it with chisel and hammer, as shown in *a*, Fig. 7. One-half an assay ton is weighed from each piece, and cupelled, and the assay carried on as usual.

Gold assays.—Gold assays are made by dissolving the silver buttons in weak nitric acid, as usual.

Lead assays.—Ores and slags are assayed for lead in the crucible. Five grams of the pulverized ore or slag are mixed with 15 grams of a flux composed of

Borax	1 part.
Bicarbonate of soda	4 parts.
Argol	1 part.
Flour	½ part.

or some analogous flux. The mixture is fused, with or without the addition of an iron nail or rod, either in the crucible or the muffle-furnace. When the muffle is used, the crucibles, represented in Fig. 5, Plate XLIII, are placed in it, together with large pieces of charcoal, to produce a reducing atmosphere, and the front of the muffle is kept closed. In both crucible and scorification assays the lead buttons and slags, when taken out of the furnace, are rapidly poured into the molds, shown in Figs. 1 and 2, Plate XLIII. In lead assaying the button of lead, detached from the slag after cooling, is weighed in grams and its fractions, and the result, multiplied by 20, gives the percentage.

Iron assays.—The ores are assayed for iron by Marguerite's well-known burette process, with a standard solution of permanganate of potash.

Estimation of gangue.—Gangue is determined by dissolving the ore in strong hydrochloric acid, or aqua-regia, collecting the insoluble residue on a filter, washing well, calcining, and weighing. Some assayers evaporate the solution to dryness at 100° C. before filtering, in order to estimate both gangue and soluble silica.

Estimation of moisture.—Moisture is determined in the ores by desiccation of one pound of ore placed in a copper pan over the muffle-furnace, or over a sand-bath heated by a kerosene lamp.

Specific gravity determinations.—This operation is performed every day at a few smelters on the slags of each furnace. It seems an unnecessary operation, first, because superintendents ought to rely solely upon careful assays for lead and silver; second, because, with a little practical experience, the mere appearance of the slag is more reliable than its specific gravity; third, because those who determine daily the specific gravity of slags and their contents in lead and silver have never been able to find a relation between the three data. In the analytical study on the slags made specially for this report it will be seen that there is no relation whatever between the contents of lead and silver; and at the smelters it is admitted that the specific gravity of slag may be raised by other substances than lead—by iron, for instance.

The specific gravity determinations are carefully made by means of the Jolly specific gravity spring-balance, represented in Fig. 2, Plate XXXVIII. This instrument consists of a wooden gallows-frame, at the end of whose horizontal beam is suspended a delicate wire spring, provided with a small ivory index, *J*, and a small brass pan, *P*, suspended from the spring by three wires. On the face of the vertical beam, looking towards the spring, is a mirror, carefully graduated in millimeters. A beaker, three-fourths filled with distilled water, is placed on a stand, *S*, which is provided with a set-screw, and moves up and down the vertical beam.

To make a specific gravity determination the eye is placed in front of the mirror in such a position that the pupil of the eye, the upper part of the ivory index, the graduation on the mirror, and the image of the pupil in the mirror are brought into line. The number of divisions at this point is x. A small piece of slag is then placed in the pan P; the division to which the ivory index is lowered is then carefully noted; let this be called x'; then $x'-x$ represents the weight of the slag in the air, expressed in divisions. The stand S is then raised until the slag dips into the water and the index rises. The number of divisions is once more carefully noted; let it be expressed by x''; $x'-x''$ represents the weight of the volume of the water displaced by the slag, consequently the specific gravity will be given by the formula $\frac{x'-x}{x'-x''}$. A little correction is necessary with this instrument; x'' should in reality be $x''+x'''$; x''' being the number of divisions lost by the pan when immersed in water.

The writer has devised a little instrument, easy to carry, easy to construct, and self-correcting, for the determination of specific gravity. It consists of a test-tube ballasted with distilled water and floating in a proof-glass filled with distilled water (see Fig. 7, Plate XLII). The test-tube is carefully graduated; the level of the water x, outside of the tube, is noted, as well as the level of the water y, inside of the tube. A small piece of slag or mineral is introduced into the tube, which sinks a certain number of divisions x'; x' represents its weight. The water is raised inside of the tube a certain number of divisions y'; y' represents its volume; $\frac{x'}{y'}$ gives its specific gravity corrected for temperature. One of the great advantages of this instrument is that specific gravity determinations can be made with almost as much accuracy with common water as with distilled, the weight and volume of water being self-correcting.

SECTION II.

MATERIALS USED IN SMELTING.

GENERAL CONSIDERATIONS.

Smelting is conducted on exactly the same principle by all the smelters throughout the camp. Ab uno disce omnes. The ore is invariably smelted in blast furnaces lined with fire-brick, and provided with water-jackets at the zones of agglomeration and fusion; dolomite, hematite, and old slag being used as fluxes, and a mixture of charcoal and coke as fuel. In one smelter only a little metallic iron (old horseshoes) is used for the reduction of galena when present in certain proportions in the ore, but even at this smelter it is an accidental rather than a normal operation. The facilities afforded to the smelters by nature in the Leadville region are really very great; there smelting is practically reduced to its elementary principles. The ore is, so to speak, "roasted by nature," since cerussite is evidently in all cases the result of the oxidation of galena; it requires no preliminary preparation save crushing, and for about one-fifth of the ore, which comes out of the mine in the state of sand, this is, of course, dispensed with; the quantity of matte and speiss formed is small; a good quality of hematite is found on Breece Hill, though it is used but in small quantity, owing to the fact that the ores themselves often contain the requisite quantity of iron to form slag, and to reduce arsenical, antimonial, and sulphuret compounds of lead. Dolomite, as will be seen later, forms as good a flux as carbonate of lime; its chief defect is that the slag formed is less fusible than pure lime-and-iron slag.

Before the railroads reached Leadville the smelters were compelled to use dolomite. Since that time it is said that a smelting firm has adopted the use of limestone with good results and that its use is likely to become general in the camp.

Smelting in Leadville at the present day is never badly performed, chiefly for the reason that all the furnaces are constructed on the same principles and are provided with the latest improvements. The imperfections in smelting are generally intentional, and are based on economical grounds which are in themselves unattackable and render criticism useless. Still, it must be stated that a few smelting firms have brought smelting in Leadville to actual practical perfection, and in their economic results these are the most successful.

STATISTICS OF LEADVILLE SMELTERS.

In Table IV will be found the following information, compiled from data gathered by special experts for Mining Statistics of the Tenth Census and by the writer, for the year ending June 1, 1880, each smelter being designated by a letter:

I. Annual consumption of ore.
II. Annual consumption of fluxes; their nature and cost.
III. Annual consumption of fuels; their nature and cost.
IV. Annual production of bullion; its contents, and cost of transportation.
V. Relations between ore, fuel, fluxes, bullion, and silver.
VI. Plant of each smelter.
VII. Labor; amount, time employed, and cost.

TABLE IV.

I. ORE.

	A.	B.	C.	D.	E.	F.	G.	H.	I.
Tons...	10, 236	38, 000	18, 590	4, 200	8, 411	5, 793	25, 464	12, 000	(a)

a No data.

II. FLUXES.

1. Dolomite.
2. Hematite.
3. Average price of dolomite per ton.
4. Average price of hematite per ton.

		A.	B.	C.	D.	E.	F.	G.	H.	I.
1.	Tons ..	232	5, 312	4, 170	250	440	964	2, 467	(a)	(a)
2.	..do	280	143	1, 774	292	587	1, 162	2, 968	(a)	(a)
3.	Dollars	2. 80	4. 00	3. 50	4. 00	3. 50	3. 50	b 1. 25	3. 50	(a)
4.	..do ...	8. 00	10. 00	9. 50	11. 50	6 to 7	9. 00	8. 50	10. 00	(a)

a No data. b Cost of hauling.

III. FUELS.

1. Charcoal, in bushels.
2. Charcoal, in tons.
3. Coke, in tons.
4. Proportion of charcoal to coke at each smelter.
5. Pine wood, for boilers, in cords.
6. Average weight of cord of pine wood used.
7. Cost of charcoal per bushel.
8. Average price of charcoal per ton.
9. Cost of coke per ton.
10. Cost of pine wood per cord.

		A.	B.	C.	D.	E.	F.	G.	H.	I.	Average.
1.	Bushels	188, 760	1, 094, 870	506, 558	76, 791	200, 000	279, 498	563, 087	(a)	(a)	....
2.	Tons ...	1, 342½	7, 664	3, 546	537½	1, 400	1, 956½	3, 941½	(a)	(a)	....
3.	...do ...	3, 300	4, 890	2, 810	263	700	810	2, 550	(a)	(a)	....
4.	...do....	0.4 : 1	1. 5 : 1	1. 2 : 1	2 : 1	2 : 1	2. 4 : 1	1. 5 : 1	(a)	(a)	1.33 : 1[b]
5.	Cords ..	1, 040	3, 600	1, 200	400	760	750	(c)	1, 200	800	....
6.	Pounds.	3, 000	2, 800	3, 000	3, 000	3,000 to 3,500	2,000 to 3,200	2,000 to 2,800	2,000 to 2, 800	(a)	....
7.	Cents ..	10 to 15	10 to 17	10 to 15	12 to 18	10 to 18	10 to 18	13	10 to 12	(a)	....
8.	Dollars.	18. 57	18. 57	18. 57	18. 57	18. 57	18. 57	18. 57	18. 57	(a)	....
9.	...do....	28. 60	25. 58	30. 45	30. 56	28. 60	25. 60	25. 56	27. 50	(a)	....
10.	...do....	4. 50	4. 75	5. 00	4. 75	4. 50	4. 00	4. 50	4. 50	(a)	....

[a] No data. [b] Proportion for whole camp obtained from 2 and 3. [c] Charcoal screenings, but little wood.

IV. BULLION.

1. Tons of bullion produced.
2. Average tenor of bullion in silver (ounces per ton).
3. Average tenor of bullion in gold (ounces per ton).
4. Total amount of silver in ounces.
5. Freight to the East per ton of bullion.

		A.	B.	C.	D.	E.	F.	G.	H.	I.
1.	Tons	1, 752	6, 200	4, 436	503	1, 240	1, 321	4, 012	5, 000	(a)
2.	Ounces..	404. 5	328. 53	250	250	300	300	450	300	(a)
3.	...do.....	2. 08	None	None	None	. 15	None	None	. 15	(a)
4.	...do.....	708, 684	2, 036, 886	1, 109, 000	125, 750	372, 000	396, 300	1, 805, 400	1, 500, 000	(a)
5.	Dollars .	40 to 45	(b)	(b)	(b)	(b)	27 to 35	35. 00	35. 50	(b)

[a] No data. [b] Paid by refiner.

V. PROPORTIONAL RELATIONS.

1. Parts of dolomite to 100 parts of ore.
2. Parts of hematite to 100 parts of ore.
3. Parts of fuel to 100 parts of ore.
4. Parts of fuel to 100 parts of smelting charges.
5. Bullion extracted to 100 parts of ore.
6. Percentage of lead extracted in smelting.
7. Percentage of silver extracted in smelting.
8. Charges for smelting per ton of ore, in dollars.
9. Cost of smelting per ton of ore, in dollars.
10. Average assay of slag, in ounces of silver per ton.
11. Average assay of flue-dust, in ounces of silver per ton.

	A.	B.	C.	D.	E.	F.	G.	H.	I.	Average.
1	2.27	13.98	22.43	5.95	5.23	16.64	9.69	No data	No data	[a]10.88
2	2.73	.37	9.54	6.95	6.98	20.06	11.65	No data	No data	[a]8.3
3	45.35	33.04	34.19	19.06	24.96	47.76	25.49	No data	No data	[a]32.83
4	36.25	23.33	22.60	15.33	19.33	32.50	19.00	No data	No data	[a]24.03
5	17.11	16.31	23.86	11.98	14.74	22.8	15.75	41.66	No data	[a]20.53
6	85 to 88	86 to 91	88	85 to 95	85 to 90	90 to 93	87	90	85 to 90	88
7	100	95 to 97	97	88 to 95	95	97	98.5	97.5	96	96.5
8	15 to 30	15 to 30	15 to 30	12 to 25	16 to 30	15 to 30	15 to 30	15 to 30	15 to 30	22.00
9	12 to 18	18 to 23	10 to 15	13 to 16	15 to 18	13.00	13.68	15.00	16 to 18	15.25
10	2	4	0.5	1.5	1.5	1.5	1.5	4	1.5	2
11	36	37	36	35	35	36	36	37	37	36

[a] These five averages were obtained by dividing by seven the sum of the respective proportions given for each smelter from which data were obtained. This gives a true average of the proportions for each smelter, but it might be considered that a truer average for the camp would be obtained directly from the totals of ore, fluxes, and fuel consumed during the year by these seven smelters. Calculated in this way, the average proportions are, respectively, dolomite to ore, 12.50; hematite to ore, 6.51; fuel to ore, 31.99; fuel to charge, 23.31; bullion to ore, 19.94.

VI. PLANT OF SMELTERS.

Smelter.	A.	B.	C.	D.	E.	F.	G.	H.	I.
1. Furnaces:									
(a) Number in use	2	6–3	2	2	2	2	4	3–1	2
(b) Shape	Round.	Round-square.	Square	Square.	Round.	Square.	Square.	Round-square.	Round.
(c) Working capacity: tons per 24 hours.	35 to 40	180	70	40	50	60	120	100	50
2. Steam-engines:									
(a) Number in use	1	2	1	1	1	1	2	1	1
(b) Horse power	40	160	50	40	40	50	70 and 50	100	60
(c) Average steam pressure, pounds.	60	70	60	65	70	65	60	80	70
3. Stone-breakers:									
(a) Number in use	2	3	3	1	2	2	2	1	1
(b) Capacity numbers	No. 5.	Nos. A, 2, and 5.	Nos. A, 2, and 5.	No. A.	Nos. 1 and 5.	Nos. 0 and 4.	Nos. 2 and 5.	No. 5.	No. A.
4. Cornish rolls:									
Number in use	None.	3	1	1	1	None.	2	1	None.
5. Other crushers	3 stamp-battery.	Pulverizer.	None.	None.	None.	Small mill.	None.	None.	None.
6. Blowers:									
(a) Number in use	2	9	2	2	2	2	4	4	2
(b) Capacity numbers	No. 5.	Nos. 4, 4½, and 5.	No. 5½.	No. 5.	Nos. 4½ and 5½.	No. 5.	Nos. 4½, 5, 5½, and 6.	Nos. 5 and 5½.	No. 5½.
7. Dust chambers:									
(a) Number in use	1	2	2	1	1	2	4	1	2
(b) Construction material.	Bricks.	Sheet-iron.	Limestone.	Sheet-iron.	Bricks.	Sheet-iron.	Bricks.	Sheet-iron.	Sheet-iron.

VII. LABOR.

1. *Number of each class of employees per 24 hours, when works are in full blast.*

	A.	B.	C.	D.	E.	F.	G.	H.	I.
Staff	4	9	7	4	6	5	5	5	5
General foremen	....	....	....	....	....	....	2	....	....
Foremen	....	6	2	1	....	....	3	2	1
Head smelters	2	27	8	4	4	4	13	8	4
Slag wheelers	....	32	....	....	2	4	12	9	3
Feeders	....	27	....	....	4	4	13	8	4
Helpers	32	48	26	10	....	8	8	10	5
Engineers	....	7	....	....	....	....	....	....	....
Fuel men	....	....	....	....	2	....	....	....	....
Day laborers	....	81	60 to 70	10 to 15	12	20	20 to 25	10	20

2. *Length of shift for employees (in hours).*

General foreman	....	....	....	....	....	....	13	....	....
Foreman	12	8	12	12	12	12	8	12	12
Head smelter	12	8	12	12	12	12	8	12	12
Slag wheeler	12	12	12	12	12	12	8	12	12
Feeders	12	8	12	12	12	12	8	12	12
Helpers	12	12	12	12	12	12	12	12	12
Day laborers	12	10	10	10	10	10	10	10	10
Engineers	12	8	12	12	12	12	8	12	12
Fuel men	12	8	12	12	12	12	8	12	12

3. *Wages per shift of employees.*

General foreman	....	....	....	....	....	....	$5 00	....	....
Head smelter	$4 25	$3 00	$4 00	$4 00	$4 00	$3 50	3 00	$4 00	$4 00
Foreman	....	4 00	4 to 6 00	5 00	5 00	....	3 00	4 00	4 50
Slag wheelers	....	3 00	....	....	4 00	3 00	2 50	3 00	3 00
Feeders	....	3 00	....	4 00	3 50	3 50	3 50	3 50	4 00
Helpers	2 50	3 00	3 00	3 00	3 00	3 00	3 00	3 00	3 00
Day laborers	2 50	2 50	2 50	2 50	2 50	2 50	2 50	2 50	2 50
Engineers	....	3 50	....	....	....	....	....	....	....
Fuel men	....	....	....	....	3 00	....	....	....	....

4. *Aggregate salary of staff per month.*

Aggregate	$870	$1, 350	$1, 400	$700	$1, 100	$800	$900	$800	No data

5. *Total salaries and wages per twenty-four hours.*

Total	$197 50	$697 50	$328 52	$122 51	$110 16	$140 30	$232 58	$176 30	No data

CONSTRUCTION MATERIALS.

Common brick.—The bricks used in the construction of outer walls and dust-chambers are made from clays found in California and Big Evans gulches. They are made in a very simple way: Into a wooden mold (Figs. 9 and 10, Plate XLIV), divided into three compartments having the shape and dimensions of bricks, a lump of the clay, brought to the proper degree of consistency, is jammed at one blow without subsequent effort or pressure. The excess of clay, represented by *c* in Fig. 9, is cut off by means of an iron wire, both ends of which are fixed to a wooden handle (Fig. 11, Plate XLIV). The mold is then reversed and gently shaken. The detached bricks are dried as usual in long rows in the air. They are then piled up in large stacks and burned.

Fire-brick.—The fire-bricks used for the lining of furnaces are sent to Leadville chiefly from the manufactory of Messrs. Evans & Howard, St. Louis, Missouri, and also from the Cambria Fire-brick Company, Golden, Colo., and from the Denver Fire-clay Company, Denver, Colo.

Tapping clay.—Good plastic and refractory clay is needed for tymp-stones, tap-holes, tamping, and steep (brasque) used in the lining of furnace crucibles. The following analysis of tapping clay found in Big Evans gulch and used at the Grant smelter was made by Dr. M. W. Iles, of the Grant Smelting Works:

ANALYSIS VII. TAPPING CLAY.

Silicate of alumina	74.5
Water	14.0
Oxide of iron	3.0
Magnesia	Trace
Carbonate of lime	6.8
Free silica	1.0
	99.3

(M. W. Iles.)

Other materials.—When the smelting works are erected on the plans of superintendents, the smelting implements are derived from various sources. The castings, however, such as water-jackets, iron pillars, plates for supports and frames of crucibles, ingot-molds, slag-pots, etc., are generally made by Messrs. Hendey & Meyer, of Denver; while the boilers and engines are made by different foundries. In many cases smelters have found it more convenient and advantageous to obtain the whole of their smelting plant from Messrs. Fraser & Chalmers, of Chicago, Ill., who are prepared to furnish a complete smelting outfit, from the crushers and furnace down to ingot-molds and tamping-rods.

FUELS AND FLUXES.

Coke.—Coke is made in El Moro, on the Rio Grande Railroad, from Cretaceous coals found there; it is known in Leadville as El Moro coke. It is also made in Como, on the South Park Railroad, from Como Cretaceous coals, and is then known in Leadville as South Park coke. From Messrs. Billing & Eilers, prominent smelters of Leadville, the following information respecting coke has been obtained. The composition of the ash was determined in their laboratory.

ANALYSES VIII AND IX. COKES.

	VIII, El Moro coke.	IX, South Park coke.
Percentage of ash	22.0	9.5
Composition of ash:		
Silica	84.5	29.1
Peroxide of iron	7.1	47.8
Alumina, lime, etc	8.4	23.1
	100.0	100.0

The weight of coke per bushel is about 40 pounds, so that one ton of coke contains about 50 bushels. Detailed information respecting the consumption of coke, its price, and relation to charcoal has already been given in Table IV.

Charcoal.—This fuel is made from the spruce tree, which abounds in the vicinity of Leadville. The pine wood, cut in lengths of four feet, is converted into charcoal by the usual process of slow burning in pits or kilns. The pits consist of stacks of wood 40 feet long, 12 feet high, and 15 feet wide, entirely covered with earth. Apertures provided at the base of this rough kiln allow the slow combustion of wood to take place. When the operation is completed the apertures are stopped with earth and the whole mass is allowed to cool thoroughly. The charcoal made in this way is not of very good quality; that made in kilns is much better.

Charcoal kilns.—In the valley of the Arkansas, south of Malta, there are several establishments each provided with nine or eleven beehive-shaped kilns, erected especially for the purpose of supplying the smelters with charcoal. The oldest establishment of this kind is to be found in California gulch, in close proximity to the south bank and opposite to Messrs. Billing & Eilers's smelter; these kilns were erected by Mr. McAllister, who was the first to introduce them in the vicinity of Leadville, and from him was obtained the following information: His establishment consists of six kilns, similar in every respect, one of which, drawn to scale, is represented on Plate XXVIII, Figs. 3 and 4. The kilns are beehive-shaped; they are made of fire-brick cemented with lime-and-sand mortar, each kiln being made of 18,000 bricks. The greatest diameter is 22 feet, the height 21 feet. In front of the kiln is a charging and discharging opening, *A*, 5 feet 5 inches high and 5 feet wide, closed by a sheet-iron door, and at the back and upper part of the kiln is a feed-hole or door, *B*, similarly closed, 4½ feet high and the same in width. This feed-hole is placed at a height of 16 feet from the ground. It is connected by a tramway, running over a bridge, with the wood-stacks on the upper part of the bank of the gulch. This wood is already cut in lengths of four feet. At the base of the kiln are three rows of apertures 3 inches by 4 inches and two feet apart. The rows are 1 foot apart and contain from 22 to 25 apertures. These holes may be closed at will with bricks and clay.

The pine wood, cut in lengths of four feet, as has been previously stated, is first piled through the lower opening, *A* (large stacks of wood stand on this level), and afterwards through the upper door, *B*, and in this way the kiln is completely filled.

Both doors being left open to create a draft, a charcoal and dry-wood fire is kindled at the door *A*. Both doors are then closed and hermetically sealed with clay, and the combustion is regulated by means of the apertures *O*, which are left open or are closed, according to the intensity or direction of the wind. The air enters at the lower row and the smoke escapes at the upper.

For the complete transformation of wood into charcoal in these kilns it requires from four to eight days, according to whether the wood is dry or green. Dry wood produces a greater percentage of charcoal and of better quality than green wood. When the combustion is completed, all the apertures are hermetically sealed by means of bricks and clay and the kiln is allowed to cool thoroughly. The cooling requires about four days.

Each kiln holds from 25 to 27 cords of wood, or about 3,350 cubic feet; one cord of wood produces about 50 bushels of charcoal. In consequence, each kiln yields on an average 1,300 bushels of charcoal in 10 days. During each operation about two gallons of creosote tar runs out at the lower part of the ground door, but no use is made of it. The charcoal made in this way is of excellent quality and gives great satisfaction.

The weight of one bushel of charcoal is about 14 pounds; consequently there are about 142½ bushels of charcoal to the ton.

Composition of ash.—At the Leadville smelters charcoal is said to contain about 2.5 per cent. of ash. This figure is probably quite correct for pit-charcoal. A rough examination was made of a fine jet-black piece of charcoal from McAllister's kilns, picked from the heap at Messrs. Billing & Eilers's smelter. This gave only 1.62 per cent. of ash, containing 0.42 per cent. of soluble salts (carbonate of potash and soda, with some chlorides) and 1.20 per cent. of alumina, silica, lime, phosphates, etc. A rough examination was also made of some charcoal ash found in the laboratory of the Cumming & Finn smelter; the proportion of soluble alkaline salts was about the same as in the preceding, but the insoluble residue was chiefly composed of alumina. In all probability the composition of charcoal ash varies according to the nature of the soil upon which the trees grew. In the discussion on smelting, 2.5 per cent. has been adopted as the average percentage of ash in charcoal.

Dolomites.—The dolomites are extracted chiefly from the Dugan and Montgomery quarries and from the Glass-Pendery and Carbonate mines. The consumption, price, and proportion of dolomite used, etc., will be found in Table IV. The samples which were analyzed in the laboratory of the Survey were prepared by mixing equal weights of typical specimens picked up on the heap at various smelters.

Analysis X is that of the Dugan dolomite from the Dugan quarry, near Mount Zion, Arkansas Valley. This dolomite is in rather large, indistinct crystals, with a bluish-black tinge and with white, creamy-yellow, and red spots. The specimens were found at the Cumming & Finn and Elgin smelters.

Analysis XI is that of the Montgomery dolomite, from the Montgomery quarry, on Iron Hill, California gulch. This dolomite is in compact, homogeneous masses, with a fine crystalline structure and a bluish-black tinge. The specimens were collected at the American, California, Grant, Harrison, La Plata, and Billing & Eilers's smelters.

Analysis XII is that of dolomite from the Glass-Pendery mine, on Carbonate Hill. This dolomite has a very peculiar appearance. It is formed of homogeneous and very friable masses, composed of indistinct and exceedingly small crystals, with a uniform grayish tinge. The specimen analyzed was found at Grant smelter, where a large quantity of it is used.

ANALYSES X, XI, AND XII. DOLOMITES.

Elementary.

	Analysis X (Dugan).	Analysis XI (Montgomery).	Analysis XII (Glass-Pendery).
Carbonic acid	46.9262	43.7947	47.3922
Chlorine	0.1429	0.0618	0.0408
Sulphur	Trace	Trace	None
Sulphuric acid	None	Trace	None
Phosphoric acid	0.1224	0.0676	0.0328
Sodium	0.0697	0.0273	0.0119
Potassium	0.0386	0.0151	0.0110
Magnesium	Trace	Trace	Trace
Calcium	Trace	Trace	Trace
Lime	30.4297	27.2586	29.9671
Magnesia	20.7843	20.0455	21.5230
Protoxide of manganese	0.0533	0.0620	0.1998
Protoxide of iron	0.3827	0.5742	0.1275
Protoxide of lead	Trace	Faint trace	None
Peroxide of iron	0.1062	0.0974	0.2232
Silica	0.7064	7.7652	0.2748
Alumina	0.1673	0.1068	0.0400
Organic matter	0.0250	0.0676	0.0152
Water	0.0440	0.0550	0.0707
Loss	0.0013	0.0012	0.0700
Total	100.0000	100.0000	100.0000

Rational.

	Analysis X (Dugan).	Analysis XI (Montgomery).	Analysis XII (Glass-Pendery).
Carbonate of lime	54.0837	48.5353	53.4445
Carbonate of magnesia	43.6470	42.0955	45.1983
Carbonate of manganese	0.0862	0.1003	0.3232
Carbonate of iron	0.6165	0.9251	0.2054
Carbonate of lead	Trace	Faint trace	None
Sulphate of lime	None	Trace	None
Phosphate of lime	0.2652	0.1464	0.0710
Chloride of sodium	0.1774	0.0694	0.0456
Chloride of potassium	0.0738	0.0288	0.0181
Chlorides of magnesium and calcium	Traces	Traces	Traces
Sulphide of iron or calcium	Trace	Trace	None
Silica	0.7064	7.7652	0.2748
Alumina	0.1673	0.1068	0.0400
Peroxide of iron	0.1062	0.0974	0.2232
Organic matter	0.0250	0.0676	0.0152
Water	0.0440	0.0550	0.0707
Loss	0.0013	0.0072	0.0700
Total	100 0000	100.0000	100.0000

Discussion.—The dolomites used in smelting are true dolomites, in which small quantities of carbonate of lime and magnesia are replaced by carbonates of manganese and iron. The presence of chlorides was at first puzzling, for the reason that about half the total quantity of chloride was soluble in boiling water when the dolomites were finely powdered, the other half insoluble. The insoluble chloride might have been contained in apatite crystals combined with calcium. A microscopical examination by Mr. Whitman Cross disclosed, however, no apatite, but a very great number of minute fluid inclusions. Both he and Dr. W. F. Hillebrand were led to believe that the chlorides were contained in the inclusions. Dr. Hillebrand partly proved it by levigating the dolomites and thereby extracting nearly three-fourths of the chlorides by treatment with water. Mr. Emmons directed me to make experiments on dolomites broken into small pieces, but not powdered. These were digested with water over a water-bath for 48 hours. The first solution contained only traces of chlorides, and the experiment being repeated a second time in the same conditions the second solution did not contain any chlorine, thus proving that the chlorides are not impregnating the mass of the rock, but are contained within the crystals of dolomite. The fact that the Glass-Pendery dolomite, which is half disintegrated, contains much less chloride than the dolomites in compact masses corroborates these views.

The Dugan and Montgomery dolomites have another point of interest, which should not be overlooked. These dolomites contain traces of sulphides, whether of iron or of calcium there was no time to determine. The fact is, however, that no sulphide of iron is visible in the microscopical section and that the dolomites treated by weak acids evolve unmistakable sulphureted hydrogen. Should the presence of sulphide of calcium eventually be proved beyond a doubt it would give a great practical value to the observation made by the writer, that carbonate of lime is extremely soluble in sulphide of calcium. This reaction is so striking that it seems probable that it plays a great part in nature and that carbonates of lime may be carried away in alkaline solutions as well as in acid ones and deposited from these. On the other hand, there seems to be a relation between the quantity of organic matter and the quantity of sulphides contained in the dolomites. The Glass-Pendery dolomite, which contains only traces of organic matter, has no sulphides, while the Montgomery dolomite, which contains the largest proportion of organic matter, contains also the largest amount of sulphides. These relations may, however, be purely accidental. The dolomites were examined for the precious metals, but no silver could be detected in either of them, although it is said in Leadville that the Glass-Pendery dolomite contains from one to two ounces and the Carbonate Mine dolomite from two to six ounces of silver to the ton.

To complete the discussion of dolomites a few analyses made at various Leadville smelters are given below. The average composition of dolomites, which has been adopted in the discussion on smelting, was derived from them.

Analysis of a Glass-pendery dolomite once used at the California smelter, made at the time by the superintendent, Mr J. E. Hardman.

ANALYSIS XIII. DOLOMITE.

Carbonate of lime	50.03
Carbonate of magnesia	35.16
Silica	1.14
Protoxide of iron	0.41
Alumina	2.62
Moisture	10.64
	100.00

Analyses of dolomites made in the Grant Smelting Works by Dr. M. W. Iles.

	XIV.	XV.	XVI.	XVII.	XVIII.	XIX.
Carbonate of lime	66. 50	54. 94	49. 57	57. 95	51. 60	55. 35
Carbonate of magnesia	25. 10	41. 93	37. 08	39. 65	39. 77	39. 35
Carbonate of iron	None	None	6. 23	None	None	None
Silica	2. 70	0. 93	4. 22	0. 76	2. 50	2. 80
Alumina and peroxide of iron	6. 40	1. 31	3. 53	1. 65	6. 13	2. 50
Organic matter	None	0. 21	None	None	None	None
Total	100. 70	99. 32	100. 63	100. 01	100. 00	100. 00

NOTE.—Analyses XIV and XV, locality not given; XVI, Glass mine dolomite; XVII, Carbonate mine dolomite, said to contain from two to six ounces of silver to the ton; XVIII and XIX, Glass-Pendery dolomite, said to contain from one to two ounces of silver to the ton.

The superintendents in Leadville do not like dolomite as a flux. It is probable that before long limestone will be substituted for it. Already Messrs. Billing & Eilers have experimented at their smelter with perfectly pure arragonite from the Duncan quarry, Arkansas Valley, close to Leadville, and the results have been most satisfactory.

Limestone.—Should limestone be used instead of dolomite, it might be brought from Robinson, in the Ten-Mile District, 16 miles distant, or from Cañon City, about one hundred and thirty miles south of Leadville, on the Rio Grande Railroad. These limestones are similar in appearance to lithographic limestone. That from Robinson (Upper Carboniferous) contains 97.11 per cent. of carbonate of lime, as determined by Dr. W. F. Hillebrand. The following analysis of the Cañon City limestone (Cretaceous) was made by Dr. M. W. Iles.

ANALYSIS XX. CAÑON CITY LIMESTONE.

Carbonate of lime	88. 90
Carbonate of magnesia	6. 30
Silica	3. 10
Alumina and oxide of iron	1. 50
	99. 80

Hematite.—The hematite used as a flux at the smelters is chiefly extracted from the Breece Iron mine, on Breece Hill, but at one smelter some Silver Wave mine iron ore is also much used as a flux. This ore was not, however, examined. The sample of Breece Hill hematite, which was examined in the laboratory of the Survey, was made from specimens collected on the hematite heaps of the following smelters: American, California, Elgin, Harrison, and Billing & Eilers.

The following is the description of the specimens and the color of their streaks:

1. Black, submetallic luster; red spots; reddish violet streak.
2. Red and yellow; silicious appearance; deep brick-colored streak.
3. Very compact; submetallic luster; magnetic; black streak.
4. Compact; dull luster; light brick-colored streak.
5. Black; submetallic luster; brownish streak.

In Analysis XXI the decimals are carried to six figures, in order to introduce both gold and silver.

ANALYSIS XXI. BREECE IRON ORE.

Elementary analysis.

Iron	66.443392
Manganese	0.007280
Nickel and cobalt	Trace
Zinc	0.025201
Copper	0.022597
Gold	0.000102
Silver	0.000404
Arsenic	0.007174
Antimony	Trace
Oxygen	27.430173
Chlorine (traces calculated)	0.000132
Water	0.290000
Carbonic acid	2.444655
Phosphoric acid	0.100740
Titanic acid	0.052250
Silica	2.388500
Lime	0.121800
Magnesia	0.619900
Alumina	0.045000
Loss	0.000700
	100.000000

Silver, 0.13 ounce to the ton.
Gold, 0.06 ounce to the ton.

Rational analysis.

Peroxide of iron	71.843540
Magnetic oxide of iron (Fe_3O_4)	18.009740
Carbonate of iron	6.445000
Chloride of silver	0.000536
Gold	0.000102
Arsenic acid (combined with Fe_2O_3)	0.011000
Oxide of copper	0.028300
Oxide of zinc	0.031400
Peroxide of manganese	0.011500
Oxides of cobalt, nickel, and antimony	Trace
Phosphate of lime	0.218140
Titanic acid (in the state of titanate of iron)	0.052250
Silica	2.388500
Lime	0.004400
Magnesia	0.619900
Alumina	0.045000
Water	0.290000
Loss	0.000692
	100.000000

Discussion.—The hematite was not examined either for bromine or iodine, with which silver is generally combined in Leadville. Chromium, tungsten, molybdenum, and vanadium were carefully sought for, but no traces of these metals could be detected.

Titanium could only be found by a method which was specially devised for its detection, and which is the following: The hydrochloric solution of hematite is reduced to the minimum of oxidation by sulphureted hydrogen and then boiled to expel the excess of this gas. The solution is then as nearly as possible neutralized with an alkali and boiled with an excess of hyposulphite of soda, which precipitates titanic acid, alumina, and a little soluble silica. The precipitate collected on a filter, washed thoroughly and calcined, is treated in a platinum vessel with a mixture of sulphuric, hydrochloric, and hydrofluoric acids, and the whole is evaporated to dryness. The residue is fused with bisulphate of potash, and titanic acid is extracted, as usual, by boiling the dilute solution.

Although magnetic oxide of iron is reported in the analysis with the formula Fe_3O_4, this is not exact. The writer succeeded in isolating this oxide in a state of great purity by alternately extracting it with the magnet and rubbing it with the finger on filter paper until it no longer soiled the paper, to which the non-magnetic oxides remained attached. It was then analyzed, and its composition is represented by the formula $Fe_{20}O_{27}=6(FeO)+7(Fe_2O_3)$, instead of $7(FeO)+7(Fe_2O_3)$, which would be equivalent to the formula Fe_3O_4. It is only quite natural that magnetic oxide formed in the midst of peroxide of iron should contain an excess of this oxide. The writer assumes that the force of adhesion was used for the first time in this instance for the mechanical separation of substances. It has been employed since in connection with the use of the magnet in investigations on the nature of different metallurgical products, and in each case it has led to interesting results.

ORE-BEDS.

Smelting charges consist of mixtures of ore with fluxes and fuel in definite but somewhat varying proportions, previously determined, so as to produce a desired chemical combination.

The ore entering into the smelting charge may be an unmixed ore of known composition, or a previously-prepared mixture of ores, called an ore-bed, or a combination of the two.

Ore-beds are prepared by superposing layers of different ores of known weight and composition in such proportion as to produce mixtures of known contents in lead, silver, iron, and silica.

Composition of ore-beds.—Ore-beds are generally made to contain equal parts of metallic iron, metallic lead, and silica or gangue, or from 20 per cent. to 25 per cent. of each. The relation between lead and silver is about six pounds of lead to one ounce of silver; but this relation often varies, as well as the percentage of lead, while, on the contrary, the percentage of iron and gangue remains pretty constant.

The great advantage derived from the preparation of ore-beds, besides giving mixtures of known composition, is that of drying the ore, an operation which if carried on in the furnace would absorb an enormous amount of heat.

In Table V will be found the following particulars in regard to seven different ore-beds:

1. Humid weight of each ore-bed in pounds.
2. Average percentage of moisture for each ore-bed.
3. Dry weight of each ore-bed in pounds.
4. Percentage of silica or gangue for each ore-bed.
5. Total weight of silica in pounds for each ore-bed.
6. Percentage of iron for each ore-bed.
7. Total weight of iron in pounds for each ore-bed.
8. Average tenor of silver in ounces to ton for each ore-bed.
9. Total weight of silver in ounces for each ore bed.
10. Percentage of lead for each ore-bed.
11. Total weight of lead in pounds for each ore-bed.

TABLE V.—*Composition of ore-beds.*

Number of ore-bed.	Ore.			Silica.		Iron.		Silver.		Lead.	
	Humid weight.	Moisture.	Dry weight.	Per cent.	Total weight.	Per cent.	Total weight.	Ounces to the ton.	Total weight.	Per cent.	Total weight.
	Lbs.	*P. ct.*	*Lbs.*		*Lbs.*		*Lbs.*		*Ounces.*		*Lbs.*
1	410,355	10. 2	368, 430	21. 54	79, 375	21. 48	79, 160	42. 94	7, 912	19. 50	71, 868
2	340,915	10. 8	304, 099	26. 50	80, 840	23. 20	70, 779	39. 12	5, 946	19. 60	59, 875
3	302,690	10. 7	270, 257	22. 30	60, 354	22. 10	59, 975	35. 03	4, 734	20. 00	54, 105
4	283,000	9. 4	256, 358	22. 00	56, 475	26. 40	67, 876	61. 62	7, 898	21. 00	54, 135
5	279,475	12. 0	245, 902	20. 00	49, 321	16. 60	40, 841	66. 02	8, 146	28. 20	69, 437
6	330,805	10. 2	296, 920	25. 40	75, 556	21. 86	64, 696	65. 03	9, 690	19. 00	56, 267
7	......	....	235, 340	17. 35	40, 831	24. 78	58, 321	56. 53	6, 651. 9	23. 45	55, 239
Totals and averages..	......	......	1, 977, 306	22. 40	442, 752	22. 30	441, 648	51. 56	50, 977. 9	21. 30	420, 926

No. 1 is made at Smelter H of ore from the Amie, Hibernia, Homestake, and Morning Star mines.

No. 2 is made at Smelter H of ore from the Amie, Chrysolite, Evening Star, Morning Star, and Virginius mines.

No. 3 is made at Smelter H of ore from the Amie, Evening Star, Hibernia, Homestake, Little Giant, Morning Star, etc.

No. 4 is made at Smelter H of ore from the Amie, and Evening Star.

No. 5 is made at Smelter H of ore from the Amie, Adelaide, and of flue-dust.

No. 6 is made at Smelter H of ore from the Amie, Morning Star, etc.

No. 7 is made at Smelter B of ore from the Catalpa, Evening Star, Henriett, Hibernia, Highland Chief, Morning Star, and Silver Wave mines.

A consideration of Table V shows—

1. That the ore beds vary a good deal in weight; in the examples given, from 117 to 189 tons.
2. That the mixtures contain on an average about the same quantity of silica, iron, and lead.
3. That on an average the relation of silver to lead by weight is as 1 to 120.4, or one ounce of silver to 8¼ pounds of lead.
4. That the amount of moisture is pretty constant.

SMELTING CHARGES.

By smelting charges will be designated the combined weights of ore, fluxes, and fuel thrown at the same time into the furnaces, and by charges, the weights of ore and fluxes entering into the composition of the smelting charges. The word ore embraces ore beds and unmixed ores, and the word fluxes, dolomite, hematite, and old slags. The weights of smelting charges differ a good deal, according to the capacity of the furnaces. The term fuel will always be used for the mixtures of coke and charcoal used in Leadville. Although the amount of fuel used in smelting will always be given in weight, it must be remembered that coke and charcoal are not weighed at all smelters, but are as often measured by the shovel or the barrow; the volume has been converted into weight for comparison.

SMELTER A.

The information obtained at this smelter is not very satisfactory. The smelting charges are made up of—

Ore, 150 pounds.		Flux, 50 pounds.		Fuel, 35 pounds.	
Ore-bed	100	Dolomite	10	Charcoal	15
Unmixed ore	50	Hematite	10	Coke	20
		Old slags	30		

Charge (ore and flux), 200 pounds. Smelting charge (ore, flux, and fuel), 235 pounds.

When coke is scarce the above fuel is used, but when coke is plentiful the fuel preferred is 35 pounds of a mixture of 60 per cent. coke and 40 per cent. charcoal.

The proportions are as follows:

Flux to ore	33⅓
Fuel to ore	23⅓
Fuel to charge	17½

This would form a very fair smelting charge; but, if we reconstruct an average charge from the consumption of ore, flux, and fuel, given for this smelter in Table IV, we find the following result:

Ore, 150 pounds.		Flux, 27.5 pounds.		Fuel, 68 pounds.	
Ore	150	Dolomite	3.4	Charcoal	19.7
		Hematite	4.1	Coke	48.3
		Old slags	30.0		

Charge (ore and flux), 187.5 pounds. Smelting charge (ore, flux, and fuel), 255.5 pounds.

The discussion of this average charge leads to the following results:

Proportion of flux to ore	25
Proportion of fuel to ore	45⅓
Proportion of fuel to charge	36¼

After inspecting these figures no one will be surprised to hear that the superintendent of this smelter complains bitterly of his furnace. The furnaces are undoubtedly very clumsy, but they are constructed on the same plan as all the other furnaces in the camp, and the fault lies chiefly in the fact that less hematite and dolomite is used at this smelter than at any other, that the slags are less fluid than any others in the camp, and that the enormous percentage of fuel exhausts itself uselessly on refractory charges. The number of smelting charges run through each furnace in twenty-four hours is equal to 300.

SMELTER B.

Smelting charges made in August, 1880.

No. 1.

Ore, 510 pounds.		Flux, 200 pounds.		Fuel, 140 pounds.	
Ore-bed	200	Dolomite	50	Charcoal	80
Low-grade ore	100	Old slags	150	Coke	60
Various rich ores	200				
Lead scraps	10				

Charge (ore and flux), 710 pounds. Smelting charge (ore, flux, and fuel), 850 pounds.

No. 2.

Ore, 510 pounds.		Flux, 190 pounds.		Fuel, 140 pounds.	
Ore-bed	100	Dolomite	40	Charcoal	80
Low-grade ore	100	Old slags	150	Coke	60
Various rich ores	300				
Lead scraps	10				

Charge, 700 pounds. Smelting charge, 840 pounds.

SMELTER B.

Smelting charges made in August, 1880—Continued.

No. 3.

Ore, 510 pounds.		Flux, 190 pounds.		Fuel, 140 pounds.	
Ore-bed	150	Dolomite	40	Charcoal	80
Low-grade ores	100	Old slags	150	Coke	60
Various rich ores	250				
Lead scraps	10				

Charge, 700 pounds. Smelting charge, 840 pounds.

No. 4.

Ore, 500 pounds.		Flux, 150 pounds.		Fuel, 140 pounds.	
Ore-bed	300	Dolomite	50	Charcoal	70
Various rich ores	200	Old slags	100	Coke	70

Charge, 650 pounds. Smelting charge, 790 pounds

At Smelter B fuel is measured by the wheelbarrow. Eighty pounds of charcoal represent one charcoal barrow made of thin sheet-iron and holding about 5½ bushels. Sixty pounds of coke represent an ore barrow used also for coke. In Fig. 2, Plate XXXI (elevation of Smelter C), both kinds of barrows are indicated.

In the smelting charges Nos. 1, 2, 3, and 4, the average proportions are:

Flux to ore	36
Fuel to ore	27½
Fuel to charge	20¼

If we reconstruct an average smelting charge from data given in Table IV we find—

Ore, 500 pounds.		Flux, 206.7 pounds.		Fuel, 165.1 pounds.	
Various ores	500	Dolomite	69.85	Charcoal	100.8
		Hematite	1.85	Coke	64.3
		Old slags	135		

Charge (ore and flux), 706.7 pounds. Smelting charge (ore, flux, and fuel), 871.8 pounds.

The figures represented here are normal; the great amount of old slags used at this smelter accounts for the relatively small proportion of hematite. The percentage of fuel is in excess of that given in the preceding examples, for the reason that part of the fuel at the smelters is used for assaying, heating, and various other purposes besides smelting.

The proportions in the average charge are as follows:

Flux to ore	41⅓
Fuel to ore	33
Fuel to charge	23⅓

SMELTER C.

Smelting charges made in August, 1880.

No. 1.

Ore, 306 pounds.		Flux, 157 pounds.		Fuel, 100 pounds.	
Ore-bed	123	Dolomite	90	Charcoal	50
Rock Mine ore	123	Hematite	7	Coke	50
Evening Star ore	19	Old slags	60		
Dunkin Mine ore	41				

Charge (ore and flux), 463 pounds. Smelting charge (ore, flux, and fuel), 563 pounds.

No. 2.

Ore, 315 pounds.		Flux, 148 pounds.		Fuel, 100 pounds.	
Ore-bed	100	Dolomite	84	Charcoal	50
Rock Mine ore	100	Hematite	4	Coke and screenings	50
Evening Star ore	44	Old slags	60		
Rock Mine ore	71				

Charge, 463 pounds. Smelting charge, 563 pounds.

No. 3.

Ore, 332 pounds.		Flux, 127 pounds.		Fuel, 95 pounds.	
Ore-bed No. 1	106	Dolomite	64	Charcoal	50
Ore-bed No. 2	53	Hematite	3	Coke and screenings	45
Dunkin Mine ore	40	Old slags	60		
Rock Mine ore	133				

Charge, 459 pounds. Smelting charge, 554 pounds.

At Smelter C the above charges are smelting charges according to our definition, but they are called semi-charges. The slags are not weighed, but measured by the ore-shovel; they are not mixed with the ore and flux, but with the fuel. Fuel is measured by the fuel-shovel in the proportion of two shovels of charcoal for one of coke. One shovel of charcoal (fuel-shovel) is equal to seven pounds, and one shovel of coke (fuel-shovel) to 14 pounds. One shovel of slags (ore-shovel) weighs about 15 pounds.

In smelting charges Nos. 1, 2, and 3, the average proportions are—

Flux to ore	45⅛
Fuel to ore	31
Fuel to charge	21⅛

At Smelter C the smelting charges are model ones, like everything else connected with this smelter. The slags obtained from the above smelting charges have the composition of singulo-silicates. They are very fluid at a relatively low temperature, and carry less lead and silver than any others in the camp; and the average charge representing the work done during a whole year will show with what regularity work is carried on at this smelter.

The average charge, deduced from data given in Table IV, is as follows:

Ore, 310 pounds.		Flux, 157.6 pounds.		Fuel, 105.9 pounds.	
Ores	310	Dolomite	69.53	Charcoal	59.10
		Hematite	29.57	Coke	46.80
		Old slags	58.50		

Charge (ore and flux), 467.6 pounds. Smelting charge (ore, flux, and fuel), 573.5 pounds.

In the average charge the proportions are—

Flux to ore	50¾
Fuel to ore	34⅕
Fuel to charge	22⅗

Being in possession of data obtained at Smelter C for the month of July, 1880, these data will be discussed, for in the opinion of the writer everything connected with Smelter C is worth recording.

	Tons.
Ore smelted in July, 1880	1,500
Dolomite smelted in July, 1880	525
Hematite smelted in July, 1880	85
Bullion produced in July, 1880	435½

The bullion produced in twenty-four hours is equal to 14 tons, assaying 136 ounces of silver to the ton.

The average charge for the month of July, 1880, is as follows:

Ore, 310 pounds.		Flux, 144 pounds.		Fuel, 105 pounds.	
Ores	310	Dolomite	67	Fuel	105
		Hematite	17		
		Old slags	60		

Charge (ore and flux), 454 pounds. Smelting charge (ore, flux, and fuel), 559 pounds.

In the preceding charge the average proportions are—

Flux to ore	46½
Fuel to ore	34
Fuel to charge	23

Production of bullion per charge, 90 pounds.

As has previously been stated, the mixture of ore, dolomite, and hematite, weighing about eight hundred pounds, is called the charge. It is made to contain about 20 per cent. of lead, of which about 88 per cent. is extracted in the state of bullion. Consequently each charge will contain 160 to 161 pounds of lead, of which 141 to 141.5 pounds are extracted in the state of bullion. This quantity of bullion requires 200 pounds of fuel for its extraction, showing that one part of bullion requires about one and one-half parts of fuel for its reduction.

As the quantity of material to be smelted in each charge weighs about 920 pounds, from which 141 pounds are extracted in the state of bullion, the remaining 779 pounds constitute the slag, containing about 2 per cent. of lead, or 15 pounds, and the loss in fumes is equal to about four pounds of lead per charge. Each charge gives about 5.67 parts of slag for one part of bullion. The furnaces run about two hundred charges in 24 hours, yielding: Bullion, 16 tons; slag, 80 tons; and consuming: Rich ores, 63½ tons; fuel, 20 tons; charges, 143 tons.

SMELTER D.

Smelting charges made in August, 1880.

Ore, 700 pounds.		Flux, 330 pounds.		Fuel, 160 pounds.	
Ore-beds	500	Dolomite	80	Charcoal	95
Various ores	200	Hematite	170	Coke	65
		Old slags	80		

Charge (ore and flux), 1,030 pounds. Smelting charge (ore, flux, and fuel), 1,190 pounds.

In the preceding smelting charges the proportions are—

Flux to ore	47
Fuel to ore	22¾
Fuel to charge	15½

The composition of this smelting charge is a normal one. The composition of the average smelting charge, calculated from the data given in Table IV, is the following:

Ore, 700 pounds.		Flux, 170.3 pounds.		Fuel, 133.4 pounds.	
Various ores	700	Dolomite	41.65	Charcoal	89.6
		Hematite	48.65	Coke	43.8
		Old slags	80		

Charge (ore and flux), 870.3 pounds. Smelting charge (ore, flux, and fuel), 1,003.7 pounds.

In the average smelting charge the proportions are—

Flux to ore	24⅓
Fuel to ore	19
Fuel to charge	15⅓

The percentage of fuel is the smallest which has been yet observed, and this last smelting charge would prove the most perfect if it were not for the important element, *time*, which has been purposely neglected. The question will be discussed after exhausting the composition of smelting charges of the various smelters, and it will be seen whether it is advisable to aim at the lowest percentage of fuel in smelting charges.

SMELTER E.

Smelting charges made in August, 1880.

Ore, 300 pounds.		Flux, 80 pounds.		Fuel. 73 pounds.	
Ore-bed	200	Dolomite	15	Charcoal	36.5
Various ores	100	Hematite	15	Coke	36.5
		Old slags	50		

Charge (ore and flux), 380 pounds. Smelting charge (ore, flux and fuel), 453 pounds.

In the preceding smelting charge the proportions are—

Flux to ore	26⅔
Fuel to ore	24⅓
Fuel to charge	19⅕

Average charge deduced from data given in Table IV.

Ore, 300 pounds.		Flux, 86.6 pounds.		Fuel, 74.9 pounds.	
Various ores	300	Dolomite	15.69	Charcoal	49.9
		Hematite	20.91	Coke	25
		Old slags	50		

Charge (ore and flux), 386.6 pounds. Smelting charge (ore, flux, and fuel), 461.5 pounds.

In the average smelting charge the proportions are—

Flux to ore	29
Fuel to ore	25
Fuel to charge	19⅓

The general rule observed at this smelter is the following: The ore is made to contain 30 per cent. of gangue, 30 per cent. of iron, and 20 per cent. of lead, and the charges are smelted with 20 per cent. of fuel.

SMELTER F.

Smelting charges made in August, 1880.

Ore, 500 pounds.		Flux, 280 pounds.		Fuel, 150 pounds.	
Ore-bed (very siliceous)	500	Dolomite	110	Charcoal	130
		Hematite	120	Coke	20
		Old slags	50		

Charge (ore and flux), 780 pounds. Smelting charge (ore, flux, and fuel), 930 pounds.

In the above smelting charge the proportions are—

Flux to ore	56
Fuel to ore	30
Fuel to charge	19¼

Average charge calculated from data given in Table IV.

Ore, 500 pounds.		Flux, 233.4 pounds.		Fuel, 238.8 pounds.	
Various ores	500	Dolomite	83.2	Fuel	238.8
		Hematite	100.2		
		Old slags	50		

Charge (ore and flux), 733.4 pounds. Smelting charge (ore, flux, and fuel), 972.2 pounds.

In the average charge the proportions are—

Flux to ore	46⅗
Fuel to ore	47¾
Fuel to charge	32½

SMELTER G.

Smelting charges made in August, 1880.

Ore, 645 pounds.		Flux, 142 pounds.		Fuel, 140 pounds.	
Iron Mine ore	85	Dolomite	62	Charcoal	80
Robert E. Lee ore	123	Old slags	80	Coke	60
Morning Star ore	185				
Robert E. Lee ore	77				
Evening Star ore	175				

Charge (ore and flux), 787 pounds. Smelting charge (ore, flux, and fuel), 927 pounds.

In the preceding charge the proportions are—

Flux to ore	22
Fuel to ore	21¾
Fuel to charge	17⅘

Average charge deduced from data given in Table IV.

Ore, 600 pounds.		Flux, 203 pounds.		Fuel, 153 pounds.	
Various ores	600	Dolomite	58.1	Charcoal	92
		Hematite	69.9	Coke	61
		Old slags	75		

Charge (ore and flux), 803 pounds. Smelting charge (ore, flux, and fuel), 956 pounds.

In the average charge the proportions are—

Flux to ore	33⅘
Fuel to ore	25½
Fuel charge	19

At this smelter the fuel is measured by the shovel.

SMELTER H.

Smelting charges made in August, 1880.

No. 1.

Ore, 530 pounds.		Flux, 175 pounds.		Fuel, 130 pounds.	
Ore-bed	500	Dolomite	55	Charcoal	110
Adelaide ore	30	Old slags	120	Coke	20

Charge (ore and flux), 705 pounds. Smelting charge (ore, flux, and fuel), 835 pounds.

No. 2.

Ore, 550 pounds.		Flux, 155 pounds.		Fuel, 130 pounds.	
Ore-bed	500	Dolomite	35	Charcoal	110
Adelaide ore	50	Old slags	120	Coke	20

Charge, 705 pounds. Smelting charge, 835 pounds.

No. 3.

Ore, 480 pounds.	Flux, 160 pounds.	Fuel, 120 pounds.
Ore-bed 480	Dolomite .... 60 Old slags 100	Charcoal............ 100 Coke............... 20

Charge, 640 pounds. Smelting charge, 760 pounds.

No. 4.

Ore, 500 pounds.	Flux, 135 pounds.	Fuel, 120 pounds.
Ore-bed............ 450 Adelaide ore....... 50	Dolomite .... 35 Old slags 100	Charcoal........... 100 Coke................ 20

Charge, 635 pounds. Smelting charge, 755 pounds.

In the preceding smelting charges the proportions are—

	No. 1.	No. 2.	No. 3.	No. 4.
Flux to ore..........	33	28⅕	33⅓	27
Fuel to ore	24½	23⅗	25	24
Fuel to charge......	18⅖	18⅖	18¾	19

Data relative to the consumption of ore, fluxes, and fuel not being obtainable at this smelter, one of the most important in the camp, the construction of an average smelting charge is impossible; but the general rule observed at the works in the composition of the smelting charges is the following: The ore-beds are made to contain equal parts of gangue and metallic iron, 20 to 25 per cent. of each, and from 16 to 25 per cent. of lead, about six pounds of lead for one ounce of silver. When the proportion of gangue and iron is equal in the ore-bed, the ore is mixed with 10 per cent. of dolomite; but when gangue is in excess hematite is added in sufficient quantity to make the balance. At this smelter the slags obtained are called acid slags. The fuel-shovels used at this and other smelters are drawn to scale in Figs. 1, 2, and 3, and the ore and slag shovels are shown in Figs. 4 and 5, Plate XLIV.

SMELTER I.

Smelting charges made in August, 1880.

[At this smelter the ore-bed was made with Morning Star, Dunkin, Iron mine, and Agassiz ore.]

Ore, 526 pounds.	Flux, 273 pounds.	Fuel, 147.5 or 137 pounds.
Ore-bed........... 263 Virginius......... 156 Chrysolite (sand). 107	Dolomite ... 66 Hematite ... 67 Old slags.... 140	Charcoal.... 60.5 } or { 57 Coke........ 87 } or { 80

Charge (ore and flux), 799 pounds. Smelting charge (ore, flux, and fuel), 946.5 or 936 pounds.

In the preceding charges the proportions are—

Flux to ore	52
Fuel to ore	26 and 28
Fuel to charge	17⅛ and 18⅛

Charges made at California works in February, 1880.—Mr. J. E. Hardman had the kindness to communicate the following experimental charges, which he made at the California smelting works while superintendent, chiefly with a view to avoid the use of old slags.

	No. 1.	No. 2.	No. 3.	No. 4.	No. 5.	No. 6.
Ore-beds	500	400	250	425	450	500
Various ores	....	....	150	....	....	....
Dolomite	45	40	50	50	65	55
Hematite	78	100	85	90	20	50
Old slags	18	18	....	....	....	....
Charcoal	100	100	100	100	90	110
Coke	40	24	24	24	48	32
Total weights:						
Ore	500	400	400	425	450	500
Flux	141	158	135	140	85	105
Charges	641	558	535	565	535	605
Fuel	140	124	124	124	138	142
Smelting charges	781	682	659	689	673	747
Number of charges in 24 hours	80	......	106	90	66	54

In the preceding charges the proportions are—

	No. 1.	No. 2.	No. 3.	No. 4.	No. 5.	No. 6.
Flux to ore	28⅕	39½	33¾	33	19	21
Fuel to ore	28	31	31	29⅕	30⅔	28⅖
Fuel to charges	21¾	22⅕	23⅕	22	25¾	23½

The following table will aid in the correct interpretation of these figures:

Tons per 24 hours of—	No. 1.	No. 3.	No. 4.	No. 5.	No. 6.
Ore consumed	20.0	21.2	19.1	14.8	13.5
Charge consumed	25.6	28.35	25.4	17.65	16.31
Fuel consumed	5.6	6.57	5.58	4.55	3.83
Slag formed	17.0	16.75	18.5	16.5	12.0
Bullion formed	2.0	2.5	3.0	5.25	2.25

It results from this table that for each ton of fuel burned the quantity of ore smelted is equal to: In No. 1, 3.5 tons; in No. 3, 3.2; in No. 4, 3.4; in No. 5, 3.2; and in No. 6, 3.5—showing that there is no advantage in using flux instead of old slag, and that there is a disadvantage in doing so, since fluxes are costly and are apt to carry away no inconsiderable quantities of lead and silver, while old slag costs nothing and is already saturated with lead and silver.

RÉSUMÉ.

General discussion.—By comparing, in a tabulated form, the average data obtained in the preceding discussions relative to each smelter with the relation between the actual and the nominal capacity, it will be possible to some extent to realize the relation between the composition of smelting charge and the time.

	A.	B.	C.	D.	E.	F.	G.	H.
Proportion of flux to ore	25.0	41.3	50.8	24.3	29.0	46.6	33.8	30.4
Proportion of fuel to ore	45.3	33.0	34.2	19.0	25.0	47.7	25.5	24.3
Proportion of fuel to charge	36.2	23.3	22.6	15.3	19.3	32.5	19.0	18.6
Tons of ore per 24 hours (actual capacity)	28.0	104.0	51.0	11.5	23.0	15.9	69.7	32.9
Nominal capacity	35.0	180.0	70.0	40.0	50.0	60.0	120.0	100.0
Relation of actual to nominal capacity	0.80	0.57	0.72	0.29	0.46	0.26	0.57	0.32

NOTE.—The relations of actual and nominal capacity, as here given by Mr. Guyard, cannot be relied upon, as he has assumed that each smelter was running 365 days during the year, whereas in point of fact the running time must have been much less and must have varied widely in the cases of different smelters, none of which were probably running at their full capacity for any great length of time. (S. F. E.)

SECTION III.

PLANT AND SMELTING OPERATIONS

SMELTING PLANT IN GENERAL.

Furnaces.—All the furnaces of Leadville are built on the same general principles and contain the same essential parts, but they belong to two distinct styles: the rectangular or square and the circular or round. The following description is made from furnaces of both styles used at Smelter B, but a glance at all the other furnaces sketched for this report, in which the same parts are designated by the same letters, will show that the differences are only in details.

Square furnaces (see Plate XXVI).—The general appearance of these furnaces is that represented in elevation, Fig. 1. The furnace is formed of two independent parts: (1) The masonry *C*, supported on a main cast-iron plate support, *O*, resting on cast-iron pillars, *P*. (2) The crucible *A* upon which rest the water-jackets *B*. The space between the water-jackets and the masonry is filled up with fire-brick, *b*. This arrangement, as it is easy to perceive, is most convenient for repairs of parts exposed to injury or destruction, and cannot be too highly commended. It is universally adopted in the camp. The masonry is firmly bound by braces *Q*, the system adopted for bracing varying with almost every furnace. Immediately above the feeding-floor, *P'*, are to be seen the feed-holes, *H*, provided with sliding doors, *S'*. The smelting charges are thrown into the furnace through these holes.

The different parts of the masonry are the following (see vertical section, Fig. 3): *C* is the shaft of the furnace. The portion of the shaft immediately below the feed-holes is called the throat. It is seen also in horizontal section in Fig. 4. *D* is the chimney. *E* is the stack. The stack can be closed or opened by means of the damper *G*. The stack is also connected with the dust-condensing chambers by means of the sheet-iron flue *F'*. *C'* represents the walls of the furnace. The wall placed above the slag-gutter *U*, which is always considered as the front part of the furnace, is called the front wall. The opposite wall, at the rear, is the back wall; on each side are the side walls, in which apertures are provided for the feed-holes. A wooden hood, *W*, and chimney, *W'*, are placed in front of the furnace and above the slag-gutter to carry off the fumes from slags. The crucible *A* is formed of strong cast-iron plates, *a*, firmly screwed and bolted together, and it is covered with a cast-iron plate, *d* (see Figs. 3 and 5, Plate XXIX). The crucible is lined with fire-brick or steep (brasque). In front of the crucible projects the fore-hearth *X*, to which is adapted the slag-gutter *U*. On one side of the crucible is placed the lead-pot *L*, communicating with the hearth or crucible *A'* (Figs. 1, 2, and 3) by means of the siphon *L.'* This arrangement is called the siphon-tap or automatic tap. It constitutes one of the greatest improvements ever introduced in the construction of blast furnaces, for by its means lead keeps always at the same level and thus escapes as much as possible the oxidizing action of the blast The lead pot is always inclosed in a cast or wrought iron box or frame, *a'*, projecting outside of the crucible. The portion of the hearth designated by *A''* (Fig. 3) is the dam. *X'* is the steep of which the hearth, fore-hearth, and lead-pot are made.

The water-jackets *B* constitute also one of the greatest improvements ever introduced in the construction of blast-furnaces. When properly cared for they never get injured; occasionally they may get shifted or spring a leak between the joints, but this rarely affects the jackets themselves. It is sufficient to state that smelting campaigns of thirteen months are known in the camp, to give an idea of the importance of this arrangement. The water-jackets *B* are hollow boxes, indicated in elevation, Fig. 1, and in section, Fig. 3. They are made of cast-iron, wrought-iron, or steel boiler-plates. In the furnace now under description they are made of cast iron. In the water-jackets water can circulate freely, so that the temperature of this portion of the furnace wall, where the most intense heat reigns in the interior, never exceeds 60° to 70° C. The water-jacket arrangement is always sectional, so as to afford every facility for the removal of the jackets when the furnaces need important repairs. The sectional disposition admits of the expansion and contraction of this portion of the furnace without altering the relative positions of the parts, and on this account must be highly commended.

In the furnace under description there are twelve jackets: two in front, called the front or breast jackets; two at the back; and four on each side. In horizontal section, Fig. 2, the manner in which they are formed is shown very clearly. The jackets are firmly screwed, bolted, and braced together. Each jacket is provided with one or more circular apertures for the introduction of the nozzles of the tuyeres. In Fig. 2 the arrangement and disposition of the tuyeres *N* is plainly seen. Each jacket is provided with a cast-iron feeder, *R*, forming an integral portion of the jacket and cast with it, for the introduction of water. The level of this feeder is higher than the upper part of the jacket, so as to fill it completely with water. Small pipes, *S*, screwed to the feeders, act as outlets for the hot water, which is carried away by the water-gutter *T*. Cold water is introduced in the feeders by means of the taps *Y* supplied from the main water-pipe, *M*. In Leadville the tuyeres are never provided with any special arrangement for cooling them by water, for the reason that the water-jackets act as perfect coolers of the tuyeres. The tuyeres are generally made of thin galvanized sheet-iron, provided with sliding valves, *l*, used to observe the interior of the furnace, and also as safety-valves, for they are left partially opened. The front jackets are always provided with an open space, *V*, varying in shape and dimensions, and closed with a plug of tapping-clay. This plug is called the tymp-stone. The tap-hole *Z* is perforated through the tymp-stone for the exit of molten slag.

Circular furnaces (Plate XXVII).—The circular furnaces are constructed on the same principles as the square ones, differing only in that their masonry is always hidden from view by a wrought-iron casing or jacket, *J′*, painted black. This jacket is made of riveted wrought-iron plates about one-fourth inch thick. The round furnaces, like the square ones, are made of two independent parts; the masonry supported on a cast-iron plate, *O*, resting on cast-iron pillars, *P;* and the crucible or hearth *A*, upon which rest the circular water-jackets *B*, always made of wrought-iron plates riveted. The interval between the water-jackets and the masonry is also filled in with fire-brick, *b*. The main cast-iron plate support *O* is provided with a circular vertical flange, *O′*, and with four projecting horizontal flanges, *O″*, corresponding to the pillars. These horizontal flanges are supported by brackets *r;* they rest on the flanges (designated also by *O″*) of the capitals of the pillars, supported by brackets *t*. The masonry jacket *J′* (Fig. 3) is incased by the flange *O′* of the main cast-iron plate support *O*, and rests on this plate, as does the masonry *C′*. The wall *C′* is made of fire-bricks. The stack *E*, a continuation of the jacket *J′*, is not lined with fire-bricks. A wooden hood, *W*, and chimney, *W′*, are placed in front of the furnace, above the slag-gutters. The hexagonal induction blast-pipe *I* supplies the branch pipes *J* and the tuyeres *N* with the blast. *K* represents the canvas hose or wind-bags connecting the branch pipes *J* with the tuyeres *N*.

The hearth or crucible *A* and fore-hearth *X* are made of strong cast-iron plates, firmly bound together and covered with a cast-iron plate, *d*. The lead-pot, which projects from the crucible, is framed in a wrought-iron or cast-iron box. The fore-hearth is provided with two slag-gutters, *U*. The hearth *A′*, fore-hearth *X′*, and lead-pot *L* are made of steep. The lead-pot and crucible communicate, as in the case of square furnaces, by means of the siphon *L′*. The circular water-jackets are made of two sections, firmly bound together. Each section is provided with a water-supply pipe, *M*, at base, and an outlet, *M′*, at top, to carry off hot water. In these jackets no other feeders are used.

Holes in the jackets allow the introduction of the tuyeres. The horizontal section (Fig. 2) through the tuyeres shows the disposition of the blast apparatus in the furnace. The deflecting elbow or sheet-iron flue *F′* forms the connection between the chimney *D* and the dust-condensing chamber *D′*. The stack *E* is provided with a damper, *G*. The furnace has but one feed-hole, *H*, with its sliding door *S*.

Blast apparatus.—The blast apparatus in general use in Leadville consists of rotary positive blasts, driven by steam power; of galvanized-iron pipes, for the distribution of the blast; and of thin galvanized-iron tuyeres, connected with the branches of the blast-pipes by means of canvas hose or wind-bags. The blower mostly adopted is Baker's rotary forced-blast blower, but at one smelter Root's forced-blast blower is also in use. The average pressure of the blast introduced in the furnaces is one inch of mercury. The following table gives the character, capacity, etc., of the blowers used at each smelter:

TABLE VI.—*Blowers in use.*

At smelter.	Number in use.	Kind of blower.	Volume of blast per revolution.	Number of revolutions per minute.	Volume of blast per minute.	Extreme limits of pressure in inches of mercury.
			Cubic feet.		*Cubic feet.*	
A	2	Baker's, No. 5	25	80	2,000	1 to 1½
B	6	......do	25	90	2,250	1 to 1½
B	2	Baker's, No. 4½	16½	90	1,485	1 to 1½
B	1	Baker's, No. 4	12	90	1,080	1 to 1½
C	2	Baker's, No. 5½	30	85	2,250	¾ to 1¼
D	2	Baker's, No. 5	25	88	2,200	½ to 1½
E	1	Baker's, No. 4½	16½	80	1,320	1
E	1	Baker's, No. 5½	30	80	2,400	1
F	2	Baker's, No. 5	25	80	2,000	¾ to 1⅛
G	1	Baker's, No. 4½	16½	108	1,782	¾ to 1¼
G	1	Baker's, No. 5½	30	100	3,000	¾ to 1¼
G	1	Baker's, No. 6	45	85	3,825	¾ to 1¼
G	1	Root's, No. 5	23¼	129	3,000	¾ to 1¼
H	3	Baker's, No. 5	25	85	2,125	1
H	1	Baker's, No. 5½	30	100	3,000	1
I	2	......do	30	80	2,400	1
J	2	Baker's, No. 5	25	80	2,000	1
K	1	Baker's, No. 4½	16½	80	1,320	1
L	1	Baker's, No. 5	25	85	2,125	1
M	1	Baker's, No. 4½	16½	80	1,320	1
N	1	Baker's, No. 4	12	90	1,080	1
O	2	Baker's, Nc. 5	25	80	2,000	1

The horse-power required to drive the blowers at a given rate is obtained by the following empirical formula: V being the volume of blast in cubic feet to be delivered in one minute; P, the pressure shown by the manometer in the blast-pipes, expressed in ounces per square inch; H, the horse power; and h, the power required to overcome friction, varying with the size of the blowers:

$$H = \frac{V \times P \times 0.003}{11} + h$$

The power required to run one of the blowers is proportionate to the pressure of the blast, volume delivered, and friction; blowers of different sizes require the same power when they deliver the same volume of blast.

Baker's rotary forced-blast blower.—These blowers are manufactured by Messrs. Wilbraham Brothers, of Philadelphia. They are compact and constructed on very simple principles. They deliver a positive blast, the volume of which is proportionate, for each size, to the number of revolutions. They hardly ever get out of order and give universal satisfaction in Leadville. (The sketches corresponding to the following description are those of a blower No. 5, taken from the maker's catalogue.) They consist of a cast iron case, *A* (Figs. 1 and 2, Plate XLII), strongly ribbed and bolted, rectangular in plan and section, and of an arched top, *B*. Inside of this works a drum, *D*, carrying two tapering arms, *C C′*, which sweep round so close to the interior periphery that no air escapes. There are, besides, in the cast-iron case *A* two other drums, *E* and *F*, acting as valves; each is provided with a crescent-shaped abutment and recess, which allow the wings of the fan *C C′* to pass it. The three drums are connected by suitable gearing on the outside of the case, as is shown in Fig. 1, in such a manner that the revolutions of the drum-valves draw air isochronously with those of the fan-drum. By their combined operation air is drawn in at one side of the apparatus at the ordinary pressure and compressed at the other to the pressure required, this pressure being in direct proportion to the velocity of the drum, as indicated in Table VI.

The blower is placed on a brick chamber, *M* (Fig. 2), connected with a sheet-iron pipe, *N*, through which air is drawn. This is the best arrangement, for by its means accidents which might result from the introduction of dust are prevented, and meanwhile concussion of air is avoided, which renders the machine comparatively noiseless. But the sheet-iron pipe *N* is often dispensed with, and air is simply drawn through the grating *R R′*, placed in front of the blower (Fig. 1). The blower is connected at *O* with the general system of blast pipes by means of galvanized sheet-iron pipes, which must be air-tight. When the apparatus is in full blast, a slight and regular pulsation is felt at two or three small holes placed at the rear of the case. Such is, in all its simplicity, the Baker's rotary blower, the most perfect apparatus of the kind ever used in smelting.

The Root positive-blast blower.—The Root blower, manufactured by Messrs. P. H. & F. M. Root, of Connersville, Ind., is shown in perspective (Fig. 3, Plate XLII) and in vertical section (Fig. 4). These sketches are copied from the maker's engraving of a No. 5 blower, delivering $23\frac{1}{4}$ cubic feet of blast per minute revolution. It is the Roots' new style of blower, formed entirely of metallic parts and less delicate in the details of its construction than the old style. This machine, like the Baker blower, is very simple and effective, and gives a positive blast in nearly every part of the case, proportionate to the number of revolutions. The external parts of the Root blower consist of two semicircular cast-iron cases, *A* and *B*, screwed and bolted to two cast-iron end plates, *S S′*, which serve also as supports to the whole machine and to the cast-iron blast-pipe *O* and air chamber *M;* of five cast-iron journals, *J;* five phosphor-bronze journal-boxes, *K;* two cut-gears, protected by the housing, *H*, and one driving pulley, *P*.

The internal-blast contrivance consists of two cast-iron revolvers, *C C′*, mounted on steel shafts, *I I′*. Each revolver acts as a fan-drum and drum-valve, with recess and abutments, forming the very simple and ingenious contrivance seen in Fig. 4. As with the Baker blower, the blast-pipe *O* is connected by means of an air-tight galvanized sheet-iron pipe with the general system of blast-pipes distributing the blast to the tuyeres of the furnaces.

Blast-pipes.—A glance at Table IV will show that at each smelter the number of blowers in use corresponds to the number of furnaces at work, and as a matter of course the capacity numbers of the blowers correspond to the smelting capacity of the furnaces. The furnaces of Leadville being always worked with several tuyeres, the blast is always distributed by branch pipes from an induction-pipe surrounding the furnace. A glance at any of the descriptive sketches of furnaces accompanying this report will show this arrangement, *I* always representing the induction pipes and *J* the branch pipes of the same. When the smelting works have only one furnace the induction-pipe is placed in direct communication with each blower by means of a branch pipe, *T*, which in this case acts as a main. At Smelter I, working with two furnaces, each furnace induction-pipe is similarly placed in direct communication with each blower; but the general system adopted in the camp, at smelters working with two or more furnaces, is to connect all the blowers with a main pipe, *R*, from which the branch pipes *T* distribute the blast to the induction-pipes of each furnace. The whole system of blast-pipes, including the tuyeres, but with the exception of the canvas wind-bags, is made of galvanized sheet iron.

Blast-pipe system.—The most complete and perfect system of blast-pipes, from which this description will be made, is applied at Smelter C. At Smelter F the arrangement is similar in every respect; at the other smelters the arrangement is very nearly the same, and if there are slight modifications of the general system there are no improvements on it. Two blowers, *A* and *B* (Fig. 1, Plate XXXI), communicate with the main pipe, *R′*, by means of the pipes *E*, each of which is provided with dampers or sliding-valves, *F*, regulating the draft, and with safety-valves, *S*, regulating the pressure. The safety-valves are set to a pressure of about nine-eighths of an inch of mercury. The draught is regulated in the main pipe, *R′*, by means of dampers, *F*, or sliding valves worked by a lever, *h h′* (see Fig. 5, Plate XXIX). A similar damper, *F″* (Fig. 1, Plate XXXI), allows the excess of blast to escape from the main pipe. The branch pipes *T′*, provided with dampers *F*, worked like the preceding, allow the introduction of the proper amount of blast required by each furnace; each pipe, *T′*, communicates at *Z″* with a manometer. The general arrangement of the induction-pipe *I* and a branch pipe, *J*, is clearly shown in the same figure, and a glance at Fig. 5, Plate XXIX, indicates the connection of the branch pipes *J* with the tuyeres *N* by means of the canvas hose *K*. Thus by the preceding disposition an even pressure of blast is secured in a main pipe, from which the proper amount required for each furnace is taken at will. As far as the distribution of blast is concerned, the blast-pipe system just described is absolutely perfect; but it lacks an important element, which would render invaluable services in smelting, namely, the means of ascertaining the volume of blast blown into each furnace during a certain lapse of time. This could easily be determined by the use of a meter similar to those used for the measurement of illuminating gas in cubic feet. This meter should be placed between the damper *F* of the branch pipes *T′* and the induction-pipes of each furnace. In this way atmospheric air might be considered as one of the elements of the smelting charges, and by this means weighed or measured with as much accuracy as the fuel itself, with which it bears the closest relation. This important point will be insisted upon in the discussion of the smelting reactions.

SMELTING OPERATIONS IN GENERAL.

Drying of the furnace.—When the furnace is new or when an old furnace has been recently relined, it is first of all carefully dried by means of a slow charcoal or wood fire kept steadily burning and slowly increasing in temperature for several days, every precaution being taken to prevent the escaping moisture from loosening the masonry.

When heat is perceptible on the outside of the walls of the furnace the drying is completed. The fire is allowed to burn out and the furnace left to cool. This done, the crucible is immediately lined either with steep in every part or with tamping in some parts only, viz, the dam, siphon, and siphon-tap. Steep or brasque is a mixture of one part fire-clay and one part coke-dust, but more generally two parts fire-clay and one part coke-dust. Tamping is a simple lining of fire-clay. It is only used for certain parts of crucibles entirely lined with fire-bricks.

Blowing-in or starting of the furnace.—The furnace, being ready for work, is filled up from the hearth to the throat with charcoal, which is set on fire at the hearth zone. The tuyere-holes of the water-jackets are left open, as well as the tymp-stone and the damper of the stack, in order to create a draft. The charcoal gradually becomes incandescent to the very throat, and when this zone has attained a low red heat the blowing-in begins. The tuyere-holes of the water-jackets, with the exception of from two to four of the holes nearest to the front, and in which the respective tuyeres are inserted (the number of tuyeres set in depending on the capacity of the furnace), are sealed with plugs of fire-clay and the wind-bags of the corresponding tuyeres are tied up with strings. The tymp-stone is set in and the blast is then turned on at full pressure. A long flame issues from the siphon-tap, and this is kept on steadily until the lead-pot becomes red hot. The clay stoppers of the tuyere-holes in the water-jackets are then removed and all the tuyeres let in. The blast at this point is regulated to the normal pressure, and the furnace is ready for the filling of the crucible.

Filling of the crucible.—Bars of bullion always kept in reserve for this purpose, and in amount from four to twelve tons, according to the capacity of the crucible, are thrown in at the feed-holes with more fuel. This is done gradually in the proportion of three bars of bullion, or 300 pounds, to eight shovels of charcoal, or about 14 per cent. of fuel. According to the capacity of the furnaces, from one hundred to two hundred and fifty bushels of charcoal are consumed in the preliminary operation constituting the blowing-in. When molten lead makes its appearance at the top of the siphon-tap a few pieces of live charcoal are placed upon it to prevent it from cooling, and the furnace is ready for charging.

Charging of the furnace.—Old slags are first of all thrown in the furnace, as a kind of test of the temperature of the furnace, which is not ready so long as the slags are not perfectly fluid. The head smelter or his assistant opens from time to time the tap-hole in the tymp-stone to ascertain their degree of fluidity, and the regular charging begins only when they run quite freely. This point being attained the charges are disposed inside of the furnace next to the walls, a depression being left in the center for the charging of the fuel. This mode of charging is the one generally adopted, but there are variations in the mode of mixing the materials forming the smelting charges. At some smelters fuel is first thrown in, then old slags, over the slags the fluxes, and above the fluxes the ore. At others fuel is mixed with old slags and fluxes are mixed with the ore. Lastly, and this is the mode of proceeding mostly adopted, the slags, fluxes, and ore are mixed together and the fuel is kept separate. At the most successful smelters the mixing of fuel and old slags, on the one hand, and of the fluxes and ore, on the other, is prevalent. Whatever mode of mixing the materials of the smelting charges is used, the manner in which they are distributed in the furnace is the same; that is, fuel is always thrown in the center of the furnace and the charge distributed on the sides next to the walls. This seems scarcely a good plan, as it favors the growth of accretions in the lower part of the shaft of the furnace immediately above the water-jackets, in the very place where their removal offers the greatest difficulties. It would seem that, if in each alternate charge the process was reversed and the fuel alternately distributed in the center and on the sides of the furnace, the lower part of the shaft would reach a higher temperature and prevent the formation of these accretions, which constitute the only real difficulty with which the smelters have to contend. The analyses which have been made of these products (see Section IV) show that they are formed of sublimated substances volatilized in the zones of higher temperature and deposited in the first cool zones which they encounter as they ascend in the form of vapor. Should the modification in the mode of charging which has been proposed prove practical, accretions would still be formed, but in the higher zones of the shaft, from which they could be detached oftener and with much more facility. In any case, they would interfere much less with the working of the furnace, which depends a great deal on the regularity with which the charges descend the shaft, and the dreaded hanging (i. e., the fall of the fuel to the tuyere-holes and suspension of the charge on the sides) would be in great part avoided.

Barring-out or barring-down of the furnace.—As it is, once per shift or once in 24 hours, as the case may be, or even once in two or three days, the furnace is barred out or down, i. e., the accretions are forcibly detached from the walls of the shaft by means of bars and sledges. The charges are allowed to descend to the level of the accretions, the blast is turned off, and long chisel pointed bars a little shorter than the height of the shaft are introduced from the feed-holes between the accretions and the walls by means of the sledge, and the accretions thus removed are left in the charge, by which they are fluxed down. When this operation is over the blast is turned on again, the charging of the furnace continues, and smelting is resumed. (The chisel-pointed bars used in barring down the furnace are represented in Fig. 7, Plate XLIV.)

Smelting of flue and chamber dusts.—Flue and chamber dusts are mixed in general with lime, and the mixture, either molded into bricks or not, is spread over the ore-beds, so that a little flue-dust enters into the composition of the smelting charges. This is evidently the best way of disposing of this rather troublesome product, and in the discussion of chamber-dust it will be shown that the admixture of lime is the best plan that can be devised for its treatment.

Running with dark top.—In Leadville, furnaces are always made to run with a dark top, and this is one of the best indications that the furnace is running properly. By this is meant that the zone of the throat is perfectly dark; that no flame issues from it; that the top part of the charge shows no signs of incandescence; and that all that is seen is a thick, black smoke ascending the chimney.

Tapping of slag.—As soon as the furnace begins to work with regularity it becomes necessary to draw out periodically the molten slag from the furnace. This is done on an average every fifteen or twenty minutes. To effect this, slag-pots, mounted on wheels and made entirely of cast iron (see Plates XXIII and XXXVII for the two styles of slag-pots used in Leadville), are brought close to the fore-hearth of the furnace and placed under the slag-gutter. A tap-hole is perforated at the middle of the base of the tymp-stone by means of a pointed steel bar about an inch thick, which is forced into the clay by gentle strokes of a light hammer. This operation is generally performed by the head smelter's assistant. The slag runs over the steep or clay with which the fore-hearth is covered, then along the slag-gutter, and thence into the slag-pot. As soon as the slag-pot fills, the head-smelter dexterously plugs the tap-hole with a small lump of soft tapping-clay stuck to the end of the peculiar iron rod shown in Fig. 6, Plate XLIV, and called the tapping-rod. During this operation showers of red-hot slag-sparks fly in every direction around the tap-hole. The tapping-pots or slag-pots are then wheeled away by the slag-wheelers to the slag-heap. The slag is either allowed to cool completely in the pot and the cake of slag thus formed is extracted bodily and broken up into fragments, or else the pot filled with molten slag is tipped over the edge of the slag-heap, where the slag runs down like lava.

Taking specimens of slag for assay.—Two or three times a day a specimen of slag is taken direct from the stream flowing from the furnace by means of a very small iron ladle provided with a long handle. The specimens thus obtained are forwarded to the assay office, where their specific gravity and their contents in lead and silver are determined. After every tapping some slag sticks to the fore-hearth and slag-gutter, from which it is easily detached by sprinkling a little cold water over it and knocking it off with an iron bar.

Matte and speiss.—In Leadville the little speiss and iron-and-lead matte formed during smelting are run into the slag-pots. At some works speiss, matte, and slag are thrown pell-mell over the slag-heap; at others the cakes of speiss or matte which have settled at the bottom of the slag-pot are knocked off the slag with the hammer. Speiss is kept in a separate heap, but no treatment has been found for it. The matte, separated from speiss and slag, is roasted in heaps and resmelted afterwards with the ore. In the study on mattes it will be seen that this roasting in heap appears to be a very bad operation.

Ladling-out of melted bullion.—From time to time bullion is ladled out of the lead-pot or siphon-tap by means of wrought-iron ladles, and poured into cast-iron molds placed in a row alongside the furnace on the lead-pot side. The molds bear in relief-letters the name of the smelting firm, so that each bar of bullion is branded with it. When cold the bars are taken out on the slag-heap, or under a shed near the engine-rooms, and then weighed, marked, and two small pieces — one from top and one from bottom — are detached by means of hammer and chisel and carefully kept for assay. When a car-load has been thus weighed the assay bits, all mixed up together in a tin can or copper pan, are forwarded to the assay office.

Watching the furnace.—Every part of the furnace requires constant watching in order to apply at a moment's notice the proper remedy for any accident that may happen. The siphon-tap requires some attention and its siphon must be kept constantly clear; this is effected by the introduction, from time to time, of a curved iron bar about two inches thick, previously heated to redness at the curved end. This bar is represented in Fig. 8, Plate XLIV. The water-jackets form perhaps the least troublesome part of the furnace, and yet it is necessary to insure the running of the water into them at such a rate that the temperature of the water issuing from them should be as nearly as possible 50° to 60° C.

The pressure at the induction-pipe manometer must be constantly watched and the pressure kept steady or modified according to momentary requirements.

The tuyeres must be kept perfectly clear from any chilled slag by the introduction of iron bars into the sliding valve, and the temperature and condition of the zone of fusion observed through the tuyeres. When black rings round the tuyeres indicate a beginning of chilling, a little more fuel is added, or the charge is somewhat diminished, the fuel remaining the same. If the temperature proves too high, fuel is diminished or the charge is slightly increased. If semi-fluid slags or raw ore form hearth accretions which do not disappear by an increase of the temperature, the blast must be shut off, the tymp-stone removed, and the hearth cleared from accretions by means of bars and sledges; a little fuel is then thrown in the hearth, the tymp-stone replaced, the blast turned on, and smelting resumed.

Blowing out of the furnace.—When a furnace needs repairing or when an accident interfering with a regular working of the furnace has occurred, the feeding is entirely suspended, but the blast is kept on until the contents of the furnace are entirely molten. The charge soon burns with a bright top and the furnace emits torrents of heavy white fumes. When the whole charge has reached the level of the tuyeres the furnace is emptied of its fluid contents, first at the tap-hole, then the breast is removed and the bullion taken out of the crucible.

Length of runs.—The length of runs or smelting campaigns is seldom less than three months, but often reaches six, eight, and even twelve and thirteen months. The lack of ore is one of the principal causes that shorten the smelting campaigns.

COST AND PROFITS OF SMELTING.

The discussion of profit and loss in smelting will be made for a smelter which stands in intermediate conditions between the smelters which produce most and those which produce least. This smelter works with two furnaces of a capacity of 35 to 40 tons each of ore per twenty-four hours. The discussion is based on data obtained at this smelter for the month of July, 1880, and which are to be found with the composition of smelting charges, the other data being all derived from Table IV. The cost of smelting per ton of ore, as estimated at each smelter, has been given also in the same table. The calculations are made on cost and profit per twenty-four hours and per ton of ore.

EXPENSES PER TWENTY-FOUR HOURS.

Power.

Cost of mechanical power per 24 hours, represented by 3¼ cords of pine wood burned under the boilers and driving engines, blowers, pumps, &c., at $4.75 per cord			$15 44

Labor.

Cost of manual labor per 24 hours:			
2 foremen, at $4 and $6		$10 00	
8 head smelters, at $4		32 00	
26 helpers, at $3		78 00	
65 laborers, at $2.50		162 50	
			282 50
Aggregate salary of staff per 24 hours			46 50

Ore.

Forty-eight tons ore smelted in 24 hours, contents 34 per cent. of lead, 41.5 ounces of silver to the ton, equivalent to 1,992 ounces, at $1.15 per ounce (New York quotations June 21, 1880)		2,290 80	
32,640 pounds of lead, equal to 1,632 units, at 15 cents per unit of 20 pounds		244 80	
		2,535 60	
Deducting 5 per cent. off price of silver	$114 54		
Deducting cost of treatment at $20 per ton	960 00		
		1,074 54	
			1,461 06

Flux.

Price paid for fluxes used in 24 hours:		
17 tons of dolomite, at $3.50 per ton	$59 50	
2¾ tons of hematite, at $9.50	26 12	
		$85 62

Fuel.

Price paid for fuel used in 24 hours:		
9¼ tons of charcoal, at $18.57 per ton	171 77	
7 tons of coke, at $37.50 per ton	262 50	
		434 27
General expenses		2, 325 39
Wear and tear and repairs of implements, say 5 per cent. of general expenses		116 27
Total expenses		2, 441 66

PROFITS PER TWENTY-FOUR HOURS.

Bullion obtained in 24 hours: 14 tons, assaying 136 ounces of silver to the ton, equivalent to 28,000 pounds, containing 1,904 ounces of silver, or 130.56 pounds (avoirdupois) of silver (New York quotations July 31, 1880):		
27,869.44 pounds of lead, at 4½ cents per pound	1, 254 12	
1,904 ounces of silver, at $1.14¼ per ounce	2, 175 32	
		3, 429 44
Deducting refiner's charges, at $14.50 per ton of bullion	203 00	
Deducting total expenses	2, 441 66	
		2, 644 66
Net profits per 24 hours		784 78
Total expenses per ton of ore		50 86
Cost of smelting per ton of ore		20 41
Profits per ton of ore		16 35

From the profits must be deducted a certain amount for the sinking fund of capital invested in plant and a certain amount for the interest on the working capital.

PLANT AND OPERATIONS OF INDIVIDUAL SMELTERS.

SMELTER A.

Disposition of works (see Plate XXV).—These works are erected on the northern bank of California gulch. Being one of the first smelters started in Leadville, its plant is somewhat antiquated, but, such as it is, it has rendered good service. The two furnaces *A A′* are the largest circular furnaces in the camp and are very clumsy. Their clumsiness is made more evident still when one hears that in spite of their large dimensions their smelting capacity is only equal to that of the smaller furnaces at present in use at the other smelters and when it is found, as has been pointed out in the composition of smelting charges, that they consume twice as much fuel as the smaller ones. On the furnace level there is a battery of three stamps. The weight of each stamp and stem is 400 pounds. These stamps are used chiefly for crushing the coke with which the steep of the furnaces is made. The furnace level communicates with the feeding-floor by means of a flight of steps, and also by means of an elevator, chiefly used to carry up barrowfuls of old slag used as flux in smelting. The boiler, engine, and blast room is next to the furnace-room, on the right facing the furnaces. The boilers are worked at a pressure of 60 pounds per square inch. The engine is of 40 horse-power and the blast apparatus consists of two No. 5 Baker blowers.

The offices, assay offices, and laboratory occupy detached buildings on a level with the foot of the slag-heap. On the feeding-floor is a large wooden trough in which the roasted flue-dust (at this smelter flue-dust is roasted before resmelting) is mixed with about 20 per cent. of milk of lime. The mixture is then spread over the ore-beds placed on this floor. The crushing machinery, placed also on this floor, consists of two large No. 5 Blake crushers (opening between the jaws, 15 by 9 inches).

Immediately outside of the main building, on the feeding-floor level, are the flues connecting the stack of the furnaces with the dust-chamber; this arrangement is the only one of its kind in Leadville. The upper part of the stacks *E*, *E′*, (Plate XXV) of the furnaces *A*, *A′*, are connected by means of the sheet-iron flues *H*, *H′*, with a main sheet-iron flue, *F′*, which enters the brick-dust chambers *D′*. Each of the flues *H H′* is provided with one, and flue *F′* with three, sliding doors, placed on the upper part of the flues and parallel with them (these doors are not visible in the sketch), and used for clearing the dust which accumulates periodically in the flues. The flue *F′* rests about half way on a small flue-dust chamber, *N*, made of bricks and provided with a sliding door, *d*, for the extraction of the flue-dust. Immediately at the rear of the dust-chamber *D′* are long rows of ore-bins, and immediately behind them is a large roasting-furnace. The level immediately above and at rear of the roasting-furnace is the fuel level, which communicates with the blast-furnaces by means of an elevated platform, *R′*, provided with a track of rails. The fuel, charged in light sheet-iron mining barrows, is thrown down next to the feed-holes along the chutes, *S*. This arrangement is capital and saves much labor; two fuel men are sufficient to supply all the fuel needed in smelting, but its great inconvenience is that of filling the whole feeding-floor with an ever-floating cloud of impalpable charcoal dust, very disagreeable to breathe and which must prove after a while most injurious to the lungs of those who live constantly in such an atmosphere. When in full blast these works employ 66 workmen per twenty-four hours.

Furnaces.—The two blast-furnaces at Smelter A are circular and identical in shape, dimensions, and capacity. Both furnaces are seen in perspective, Plate XXV, and one of them is drawn to scale in Plate XXIII. They are constructed on exactly the same general principles as all the other furnaces in the camp, but in detail differ a good deal.

The crucible *A* is very little larger than the water-jackets; it is framed in strong cast-iron plates, *a*, forming segments of a circle, six in number and firmly bolted together at the joints. The frame of the fore-hearth *X* is also made of cast-iron plates, and the projection *X″* of the fore-hearth, which exists only in this furnace, is similarly framed. The crucible, siphon-tap, fore-hearth, and fore-hearth projection are entirely lined with steep, made of one part fire-clay and one part finely pulverized coke. The projection of the fore-hearth is provided with two slag-spouts, *U*. The frame of the lead-pot is made of strong sheet-iron, *a′*, bolted to the cast-iron plates of the crucible. The system of water-jackets consists of six jackets of equal dimensions; four of these are made of strongly riveted, wrought-iron boiler-plates and two are made of cast iron. This principle is a bad one, owing to the unequal expansion and contraction of the two metals, and should be avoided; whenever this plan is adopted the water-jacket system frequently gets out of order. Each jacket is provided with a feeder, *R*, in which exists an outlet for the hot water, and a hole, *n*, for the introduction of the nozzle of a tuyere. Fig. 2 shows the disposition of the six tuyeres and of the jackets; the space between the water-jackets and the masonry above is filled as usual by fire-bricks, *b*. The pillars *P* have their capitals flange-shaped at *O''*. This flange rests on the pillars by means of brackets *t*. The main cast-iron plate-support *O* is also flanged at *o* (Fig. 4), and these flanges are connected with the circular and vertical flange *O'* of the plate by means of the brackets *r*. The masonry and stack are entirely surrounded by a wrought-iron casing or jacket, *J'*, surrounded at the base by the flange *O'*.

There is only one feed-hole, *H*, at the throat, but this feed-hole is twice as high as it is in most furnaces, and is divided into two sections by two hinged wrought-iron doors, *S' S''*. The upper door is only opened to bar out the furnace. The damper *G* of the stack is not single, as in all the other furnaces, but is made of two halves, *G G'*. The walls *C'* of this furnace are much thicker than the walls of most circular blast-furnaces.

The induction-pipe *I* is made, as usual, of galvanized sheet-iron. It has a peculiar shape; it forms a ring around the furnace, and this ring is square in vertical section, but the branch pipes *J* are cylindrical, as is always the case. Each furnace smelts from 17 to 20 tons in twenty-four hours; produces from 4 to 5 tons of bullion and from 13 to 25 tons of slag.[1] The length of run of these furnaces is about six months; they are barred out every twelve hours, at the beginning of each shift. The chief defect is that the diameter of the water-jackets at the tuyeres is rather too large. Contrary to the plan adopted at all the other smelters, periodical tapping of slag is not done here. The slag is allowed to flow in a constant stream, and the gutter in the steep of the forehearth and its projection are covered with live charcoal to prevent the chilling of the slag. The slag-pots used at these works are indicated by *B B'*; they are independent of the car *D'*, by means of which they are wheeled to the slag-heap.

The quantity of speiss resulting from the smelting of 10,241 tons of ore during the year ending June, 1880, was about 20 tons, assaying 49 ounces of silver to the ton and containing 980 ounces of silver; consequently the quantity of speiss formed amounts to about 0.2 of 1 per cent. of the ore.

Condensing chamber. — In Plate XXV the general disposition of the dust-chamber and its connection with the furnaces are seen in perspective. In Fig. 3, Plate XXIV, the same chamber is seen in horizontal section, divided into three parts by means of partition walls, *W'*, the arrows indicating the circulation of the fumes. Fig. 2, Plate XXIV, is a vertical section of the same chamber. Both sections are drawn to scale, and a glance at them is all that is necessary to understand its construction and its working. About 150 tons of dust were collected in this chamber in the space of six months.

Roasting-furnace. — This furnace is represented in elevation (Fig. 1, Plate XXIV), chiefly with a view of giving its dimensions, for it presents no peculiarity in construction. Its width (not indicated in the sketch) is 12 feet. The sketch shows the system of bracing by rails, the hinged cast-iron doors *d*, and the dotted lines indicate the in-

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[1] The municipality of Leadville uses most of this slag to macadamize the roads.

ternal disposition of the furnace. This roasting-furnace is used at Smelter A for roasting the chamber-dust previous to resmelting. In the study of metallurgical products it will be seen that it is an expensive and useless operation, and that it were better, on the contrary, to use it for the roasting of mattes and speiss. The former, being roasted in heaps, lose a great deal of their silver, and the latter is not treated in Leadville. The only point of interest in the roasting-furnace of Smelter A is the flue *C*, in which a good deal of the products volatilized during the roasting is condensed; so that this furnace is admirably adapted for the treatment of matte, accretions, and speiss, all products containing a good deal of silver. *S* represents the stack of the furnace; *G*, the damper of the stack. The ash-pit of the furnace is not visible, but is placed at *h*.

SMELTER B.

Disposition of works.—These works, the largest and most important in or near Leadville, are situated in Leadville proper, on the northern bank of California gulch and facing the gulch. With their 118 men at work in a somewhat limited space they present an unusual amount of bustle and activity and smelt about one hundred tons of ore in 12 hours.

The old slag-heap, placed immediately in front of the furnaces, is being entirely dug up, in order to be resmelted. These slags were made formerly of singulo-silicates, but now the slags contain a little more silica, and are called acid slags. The old slag-heap is placed in direct communication with the feeding-floor by means of an inclined tramway supported by timber trestle-work, on which a mine-wagon is run by a wire rope which winds over a drum placed on the feeding-floor. The new slags made at these works are allowed to solidify in the pots; they are detached while hot, lifted by means of a small crane, placed on a small iron truck running on a tramway resting partly on the ground in front of the works, partly on a timber bridge projecting over the gulch, and being taken to the end of the bridge are dumped into the gulch. The main smelting building at smelter B is 212 feet by 94 feet. The ore-bins are placed on the feeding-floor within this building, where are also placed the ore-beds, crushers, Cornish rolls, etc. The coke-room, at the back of the main building, is 200 feet by 20 feet and the charcoal bins are 150 by 18 feet. Large heaps of dolomite, hematite, and large dumps of low-grade ore fill up the open space at the rear of the works. Large Fairbanks scales, of a capacity of 20 tons, occupy a detached office at the entrance of the works. The offices, assay offices, and laboratory occupy a detached building. In the well-fitted laboratory, besides the current assays made, the specific gravity of slags is determined from day to day. The slags are considered fit to be thrown away when their specific gravity is 3.6.

Nine Baker blowers, standing in a row under the feeding-floor and at the back of the furnaces, supply the blast. These, as well as three Blake crushers, three sets of Cornish rolls, one small pulverizer, the slag-hoisting machine, and the pumps supplying the tanks from which the water-jackets are fed, are worked by two engines, 14 by 24 inches, of 60 horse-power each. Each engine is connected with two boilers, 44 inches by 14 feet, worked at a pressure of 70 pounds to the square inch. Both engines and boilers were manufactured by Messrs. Fraser & Chalmers. The engine and boiler rooms stand on the left of the works (facing the furnaces), next to the furnaces and on a level with them.

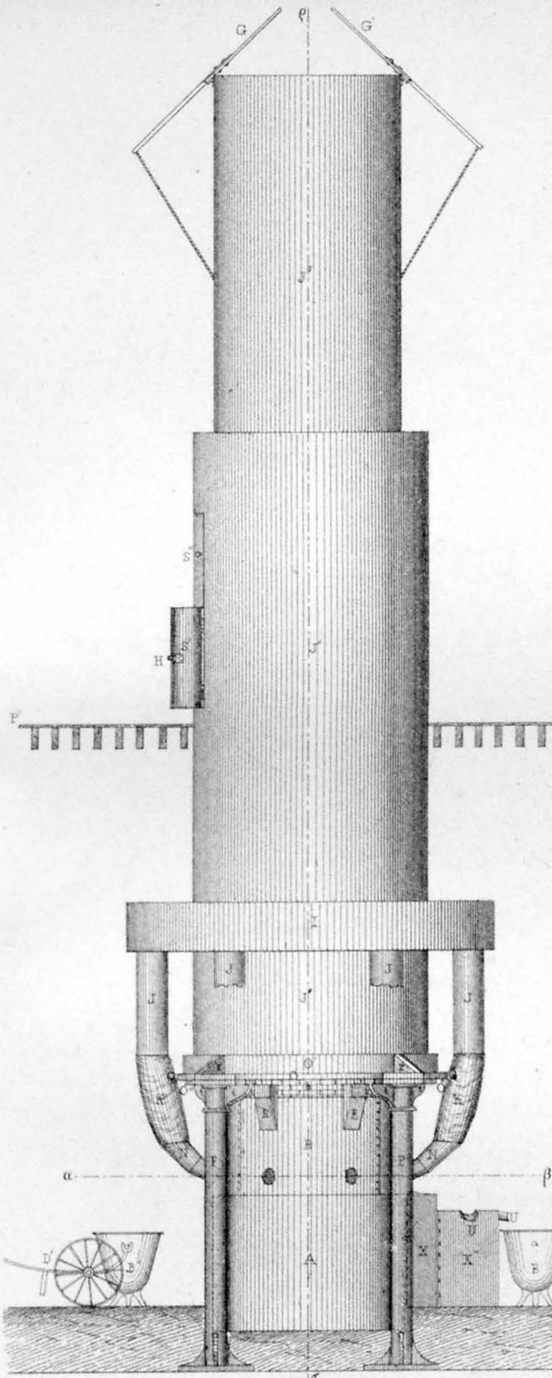


Fig. 1. ELEVATION

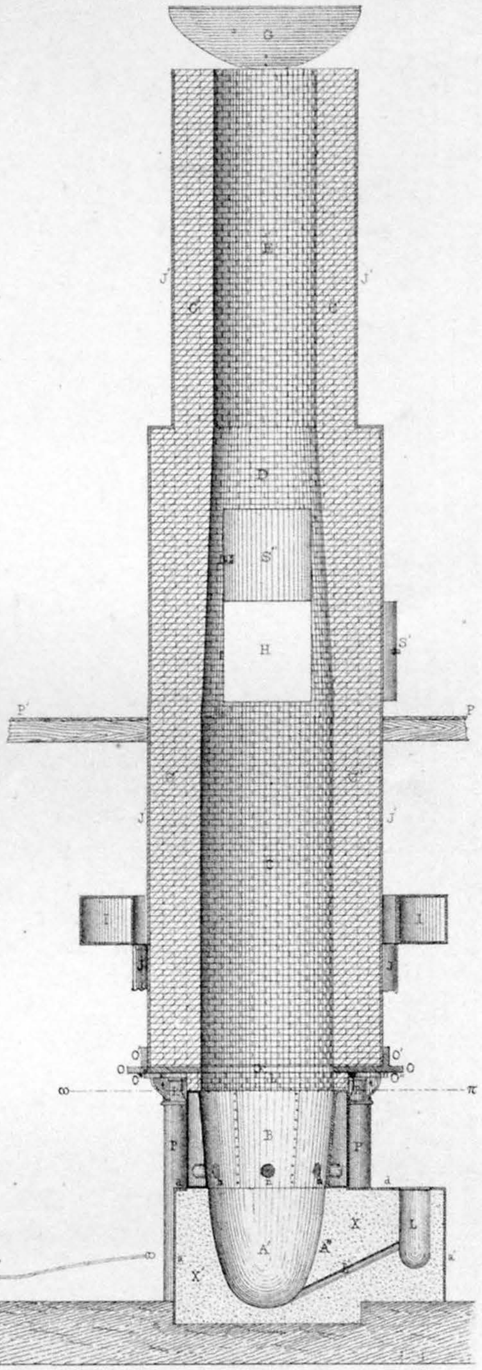


Fig. 3 SECTION ON ρ σ

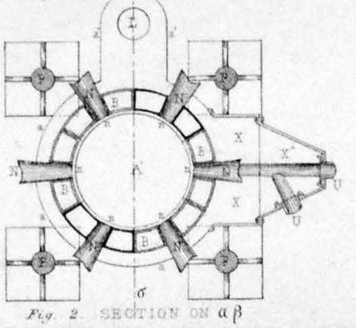


Fig. 2. SECTION ON α β

Cast Iron ▨ Wrought Iron ▨

Fire Brick ▨ Steep ▨

Scale 1 inch to 6 feet
or 1/72

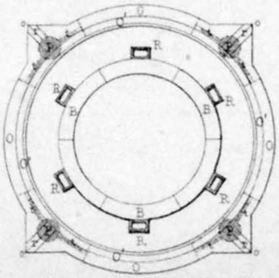


Fig. 4, SECTION ON ω π

M. Bien del.

A. Guyard, Metallurgist. Julius Bien & Co. lith. S. F. Emmons, Geologist-in-Charge.

CIRCULAR FURNACE.

SMELTER A.

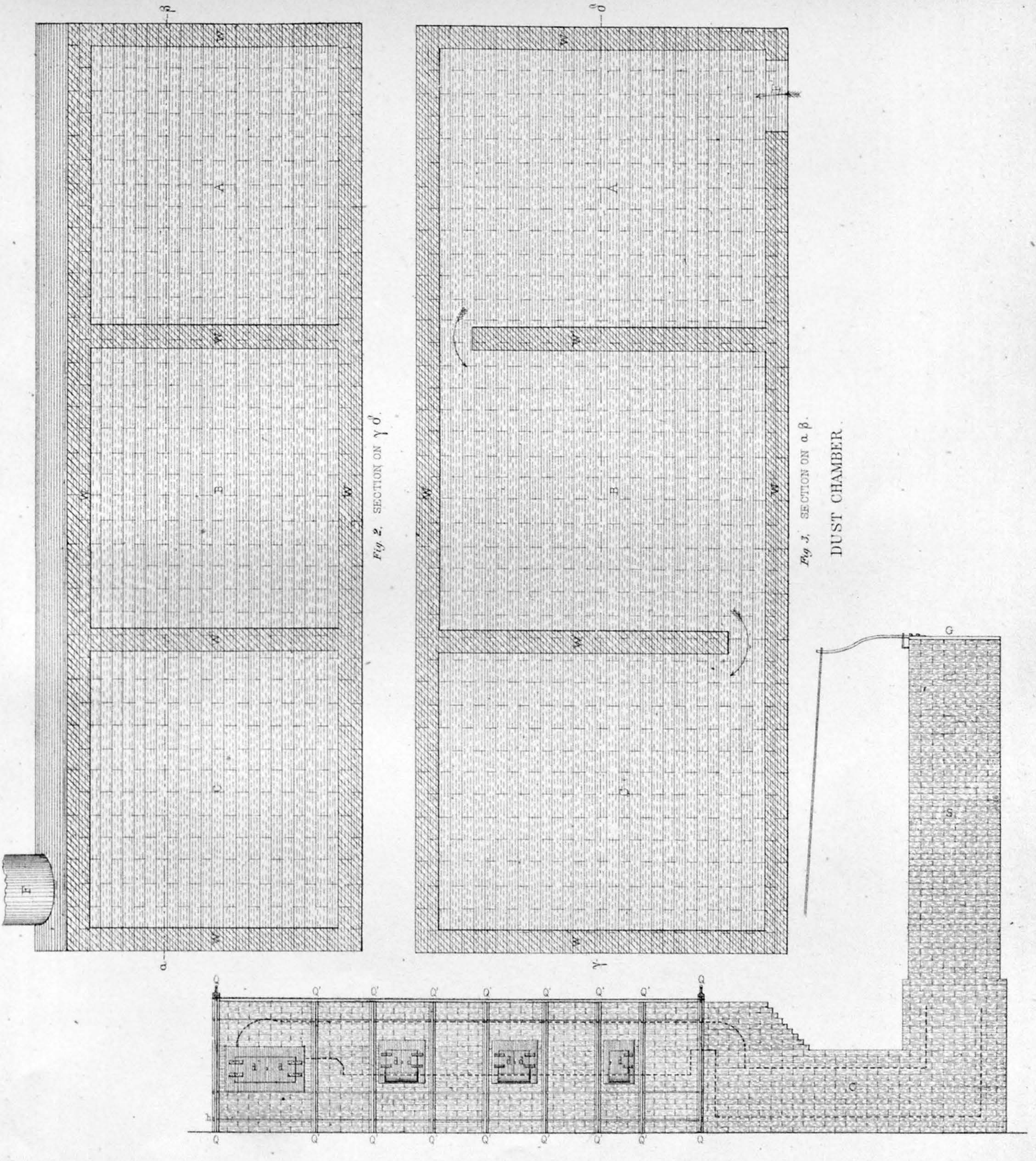


Fig. 2, SECTION ON γ δ.

Fig. 3, SECTION ON α β.

DUST CHAMBER

Fig. 1, REVERBERATORY FURNACE

Scale: 1 inch to 6 feet, or 1/72.

FEET 0 5 10 15 20 25

METERS 0 1 2 3 4 5 6 7 8

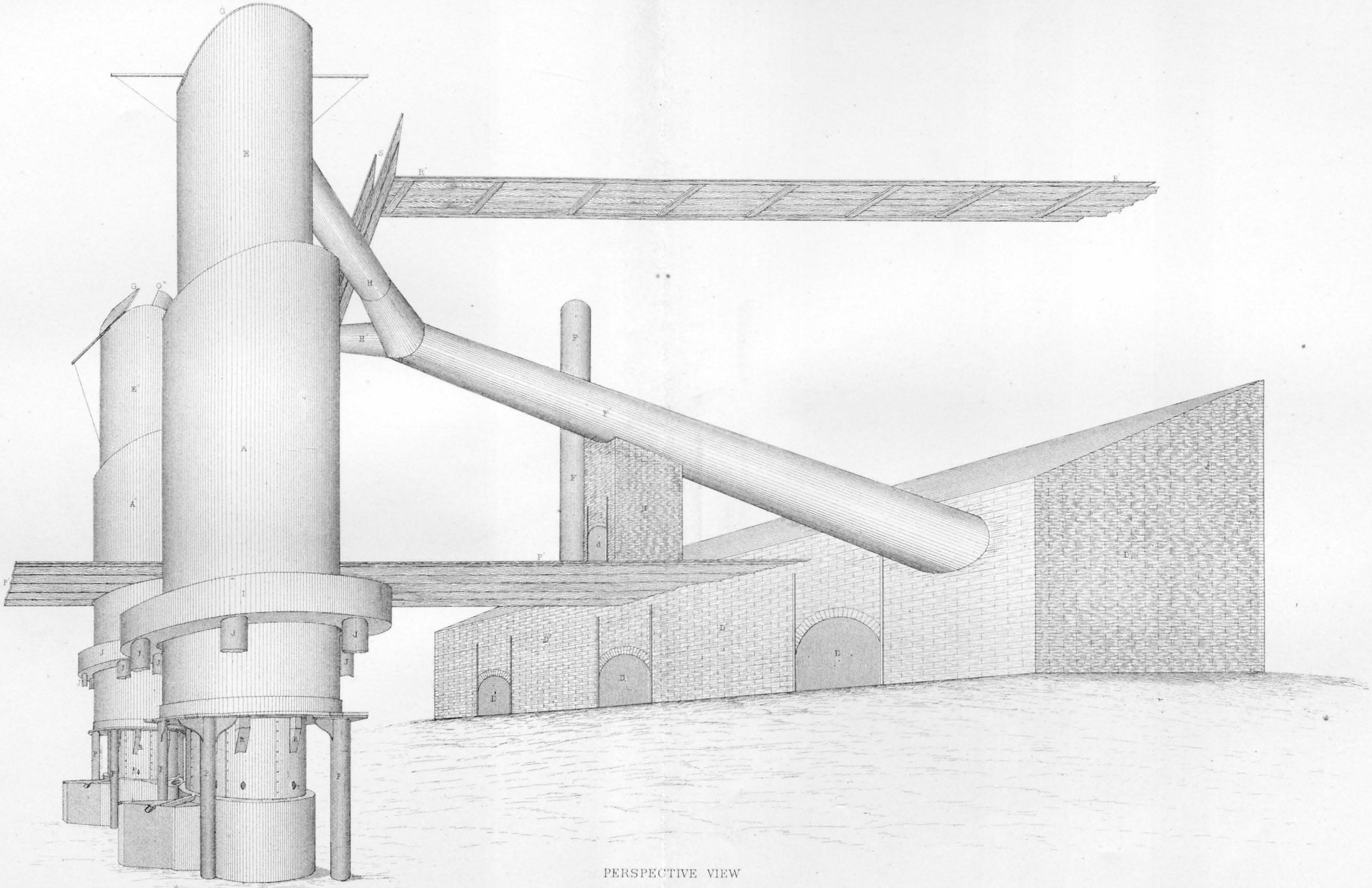
M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

REVERBERATORY FURNACE
AND DUST CHAMBER
SMELTER A.



PERSPECTIVE VIEW

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

FLUE ARRANGEMENT

SMELTER A.

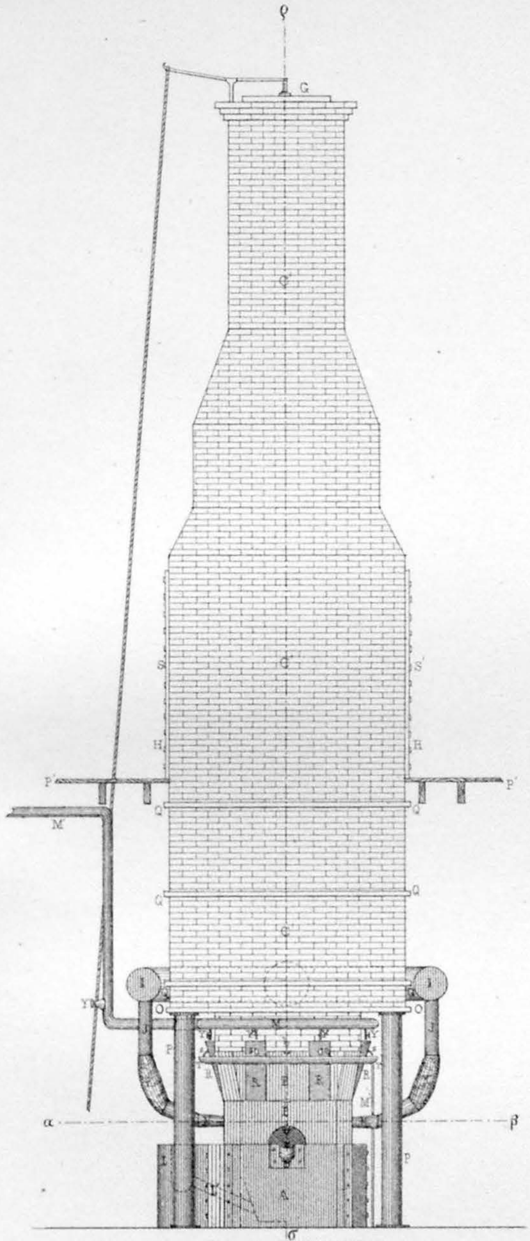


Fig. 1. ELEVATION

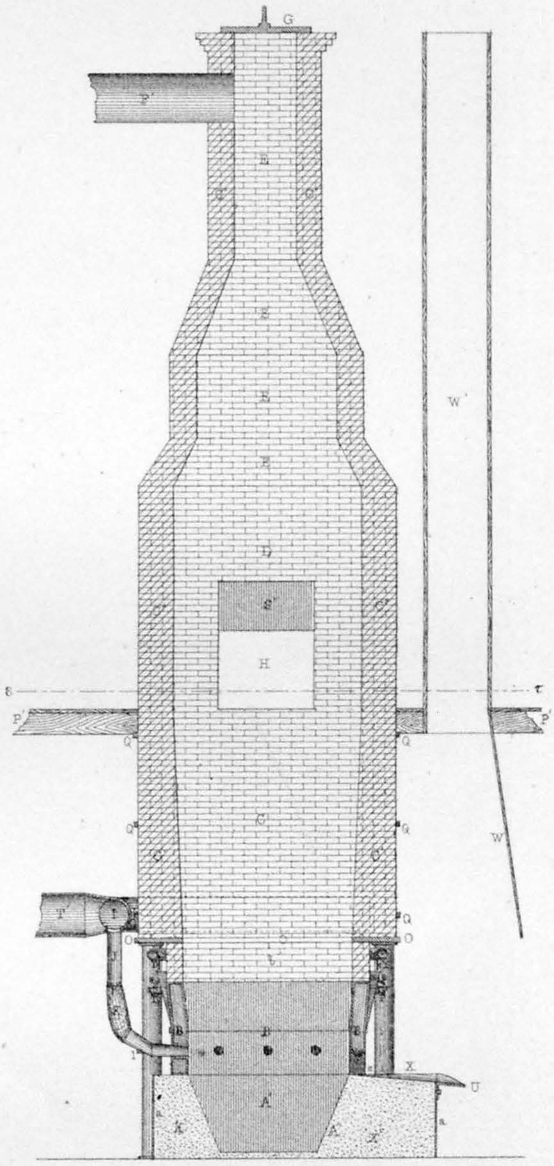


Fig. 3. SECTION ON ρσ

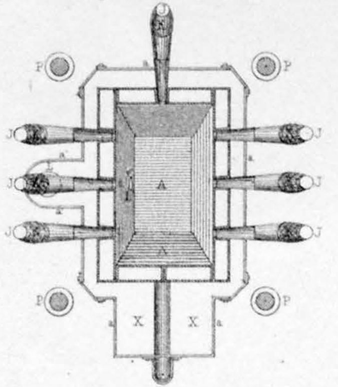


Fig. 2. SECTION ON αβ

Cast Iron ▨

Wrought Iron ▨

Fire Brick ▦

Steep ▒

Scale 1 inch to 6 feet
or 1/72.

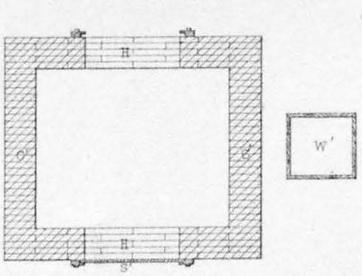


Fig. 4. SECTION ON δτ.

M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge

RECTANGULAR FURNACE

SMELTER B.

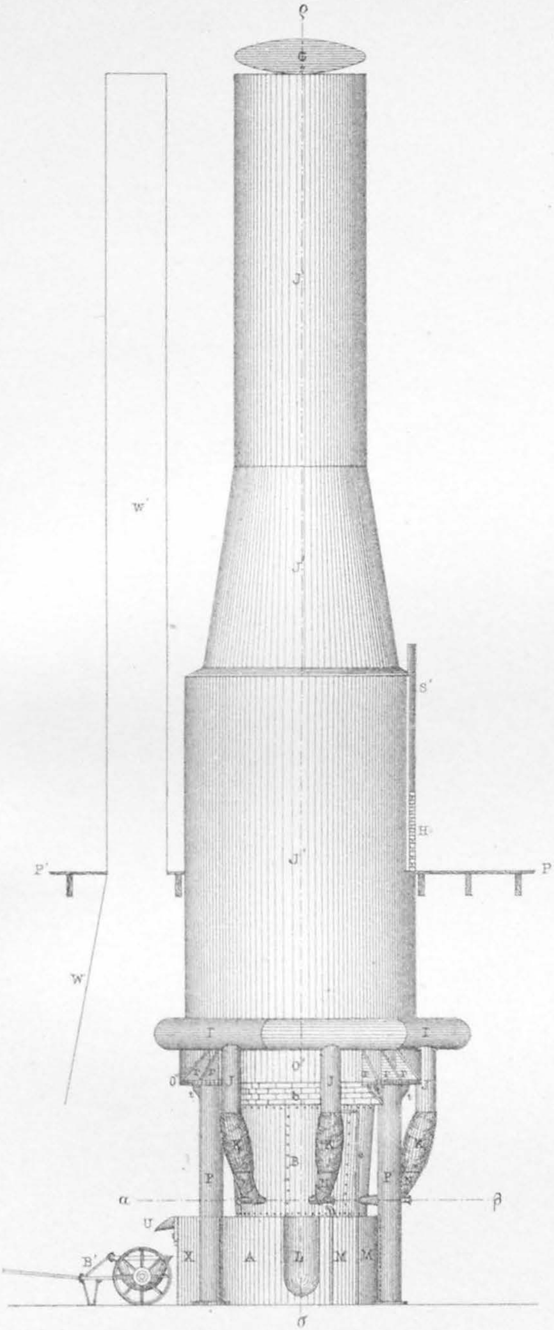


Fig. 1. ELEVATION

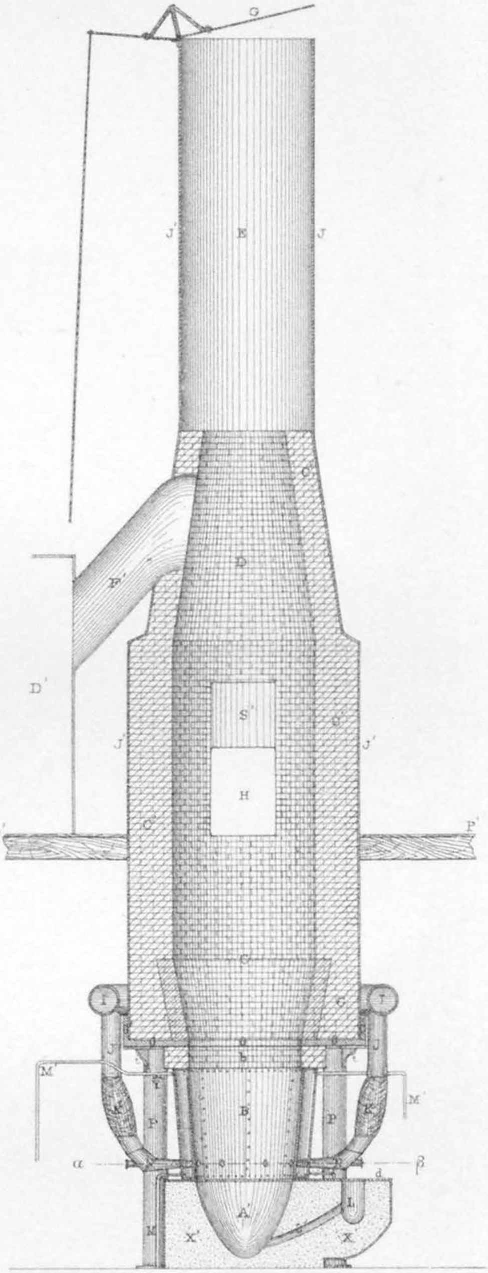


Fig. 3. SECTION ON ρσ

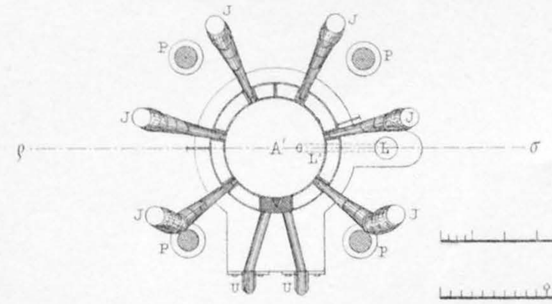


Fig. 2. SECTION ON αβ.

Cast Iron	Wrought Iron	Clay	Steep
▨	▨	▨	▨

Scale: 1 inch to 6 feet or 1:72

FEET 0 5 10 15

METERS 0 1 2 3 4 5

M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith. N.Y.

S. F. Emmons, Geologist-in-Charge.

CIRCULAR FURNACE.
SMELTER B.

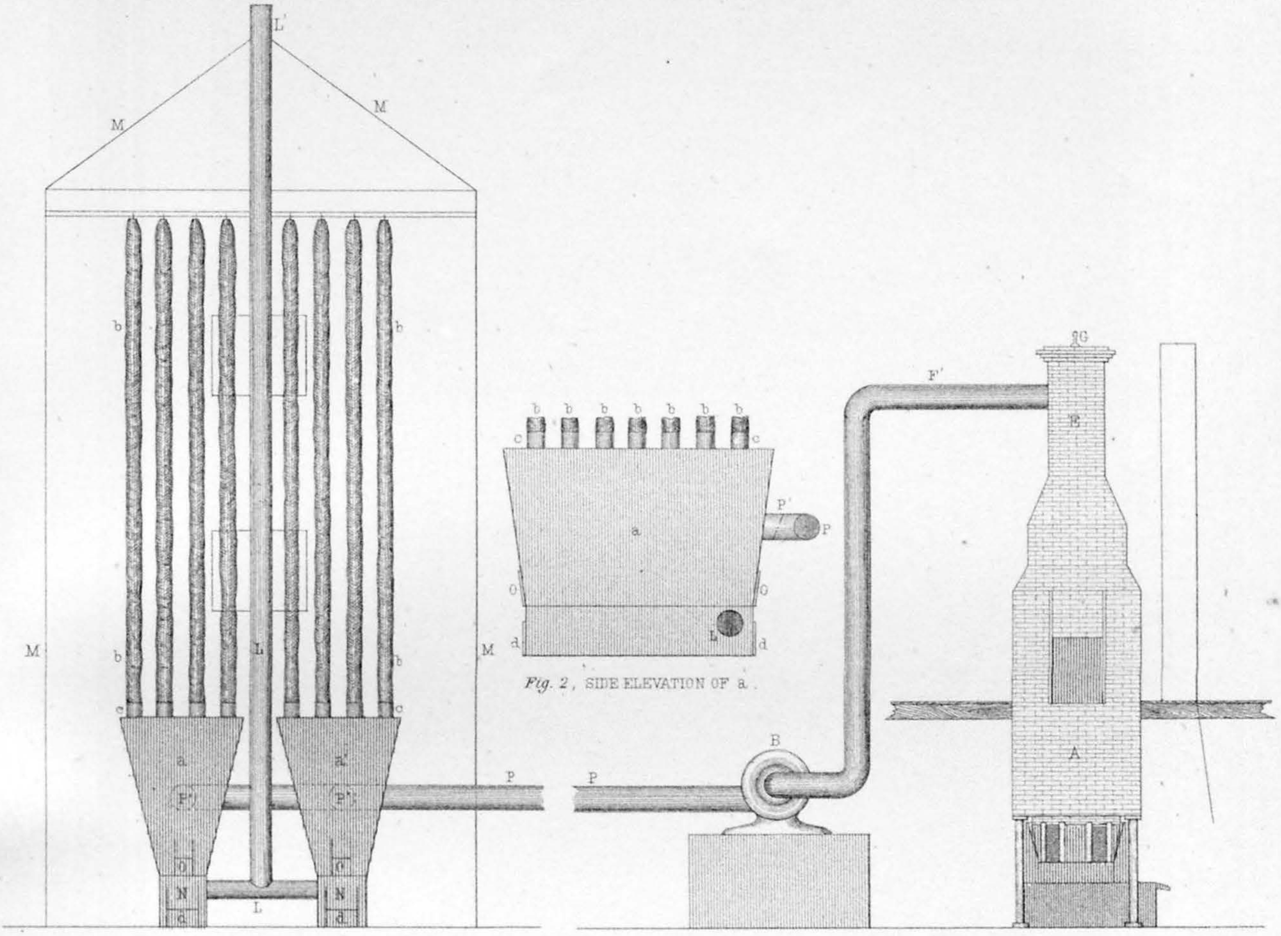


Fig. 2, SIDE ELEVATION OF a.

Fig. 1, ELEVATION.

BARTLETT'S SMOKE FILTER, SMELTER B.

Scale: 12 ft. = 1 in.

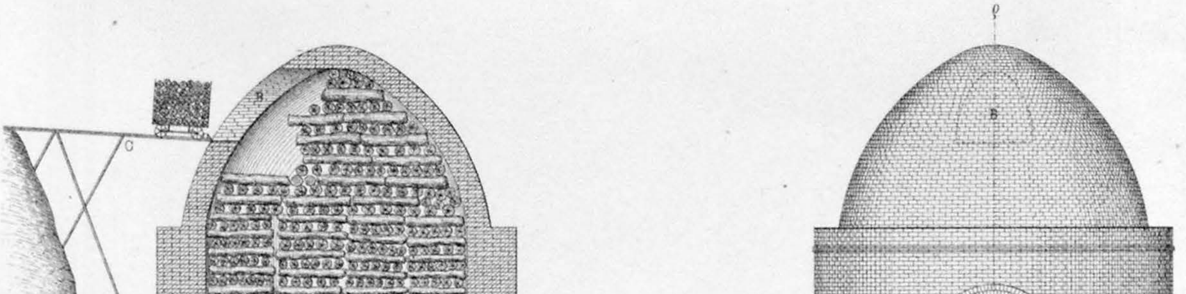


Fig. 3, SECTION ON ρ σ.

Fig. 4 ELEVATION.

M^c ALLISTER CHARCOAL KILN.

Scale: 12 ft. - 1 in.

M. Bien del.

A. Guyard Metallurgist.

Julius Bien & Co. lith.

S.F. Emmons, Geologist-in-Charge.

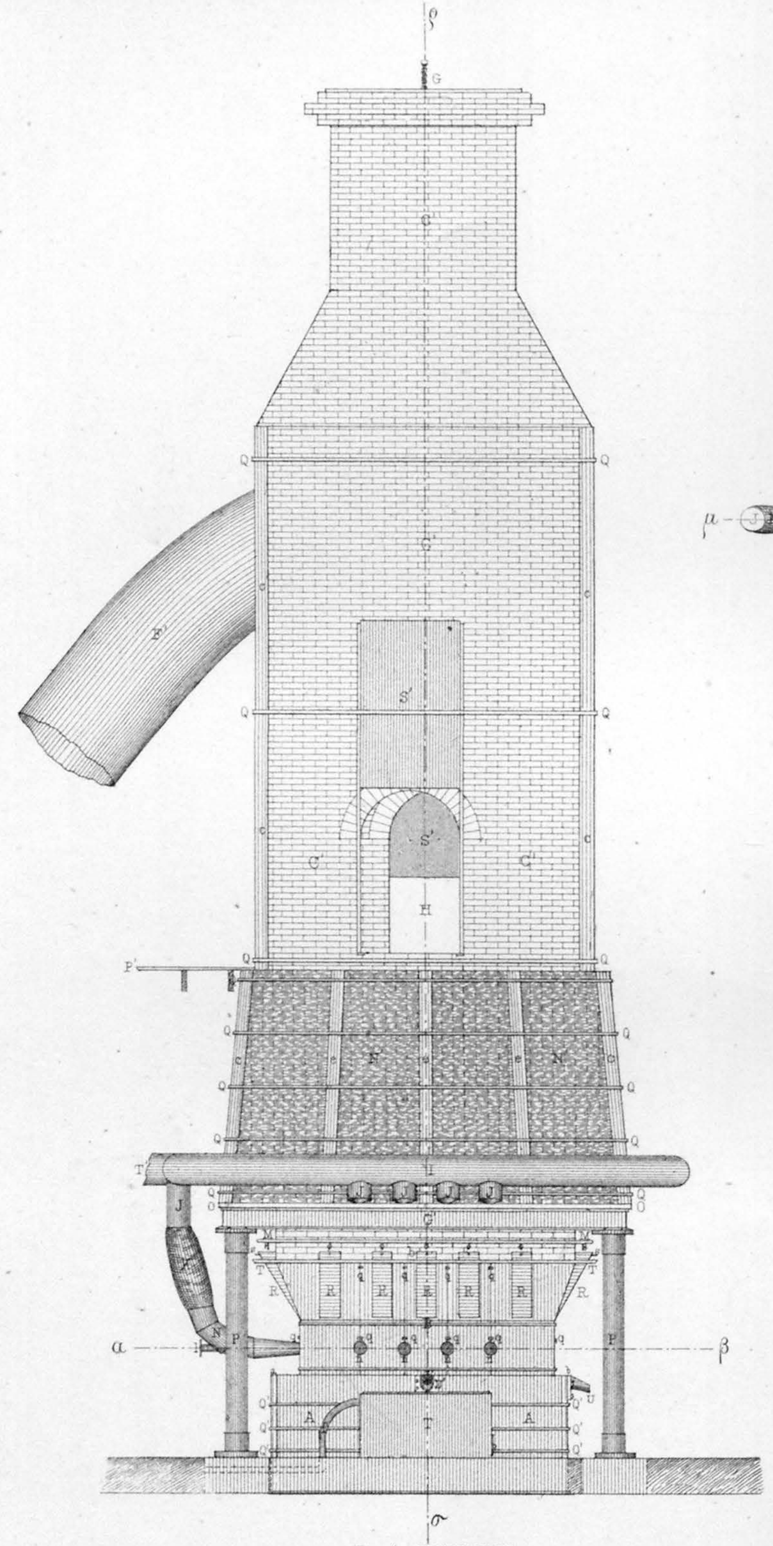


Fig. 1 ELEVATION

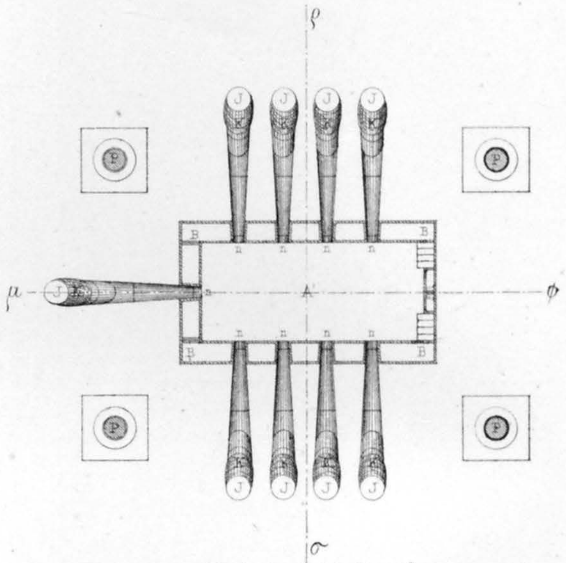


Fig. 2 SECTION ON αβ

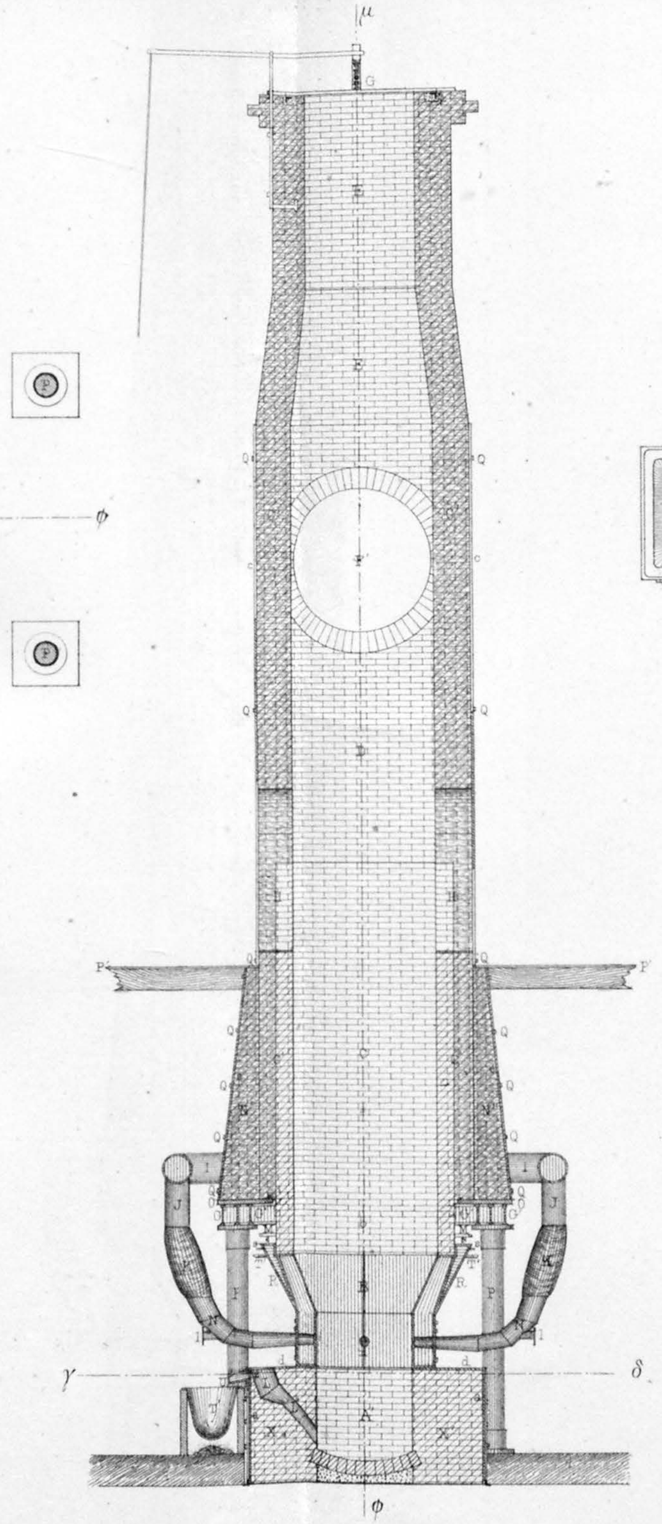


Fig. 3 SECTION ON ρσ

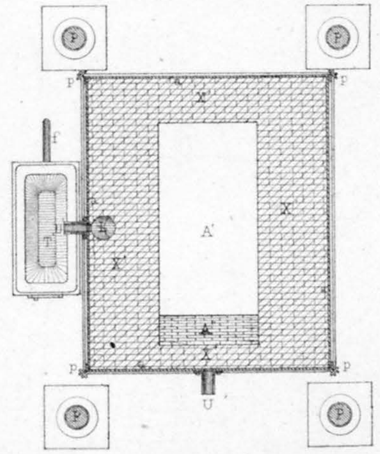


Fig. 4 SECTION ON γδ

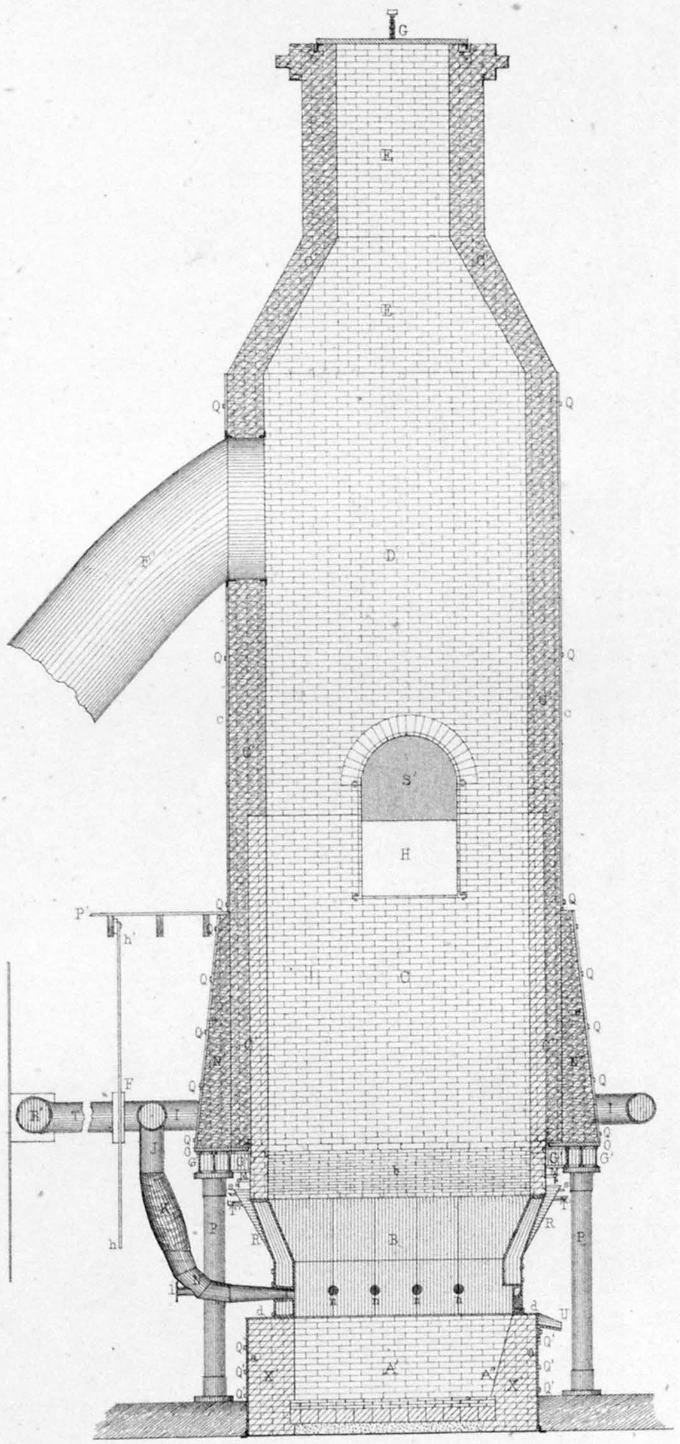


Fig. 5 SECTION ON μφ

Red Brick ▨
Fire Brick ▤
Clay ▩

Steep ▒
Wrought Iron ▨
Cast Iron ▧

FEET 0 5 10 15 20 25

Scale: 1 inch to 6 feet, or 1/72

METERS 0 1 2 3 4 5 6 7 8

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S.F. Emmons, Geologist-in-Charge.

RECTANGULAR FURNACE
SMELTER C.

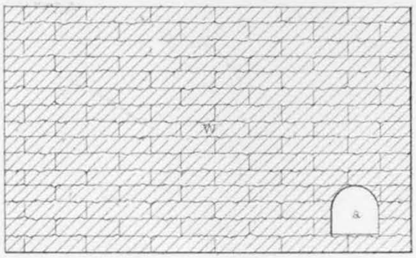


Fig. b SECTION ON 1-2

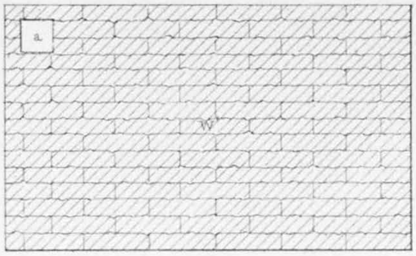


Fig. c SECTION ON 3-4

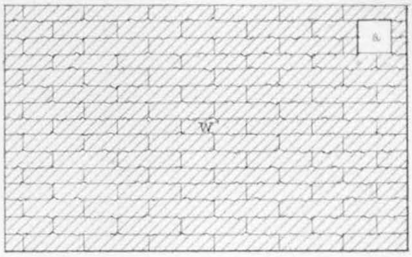


Fig. d SECTION ON 5-6

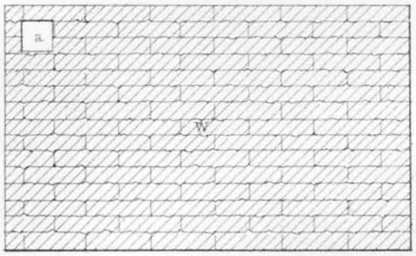


Fig. e SECTION ON 7-8

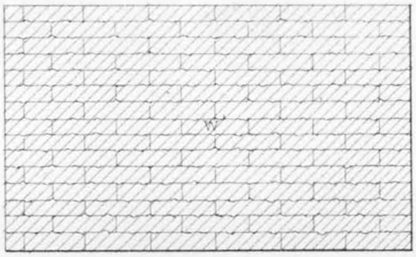


Fig. f SECTION ON 9-10

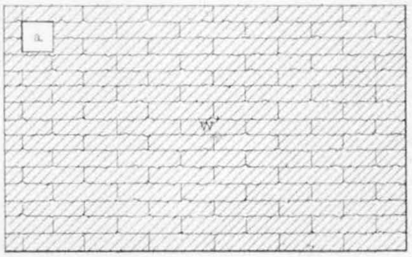


Fig. g SECTION ON 11-12

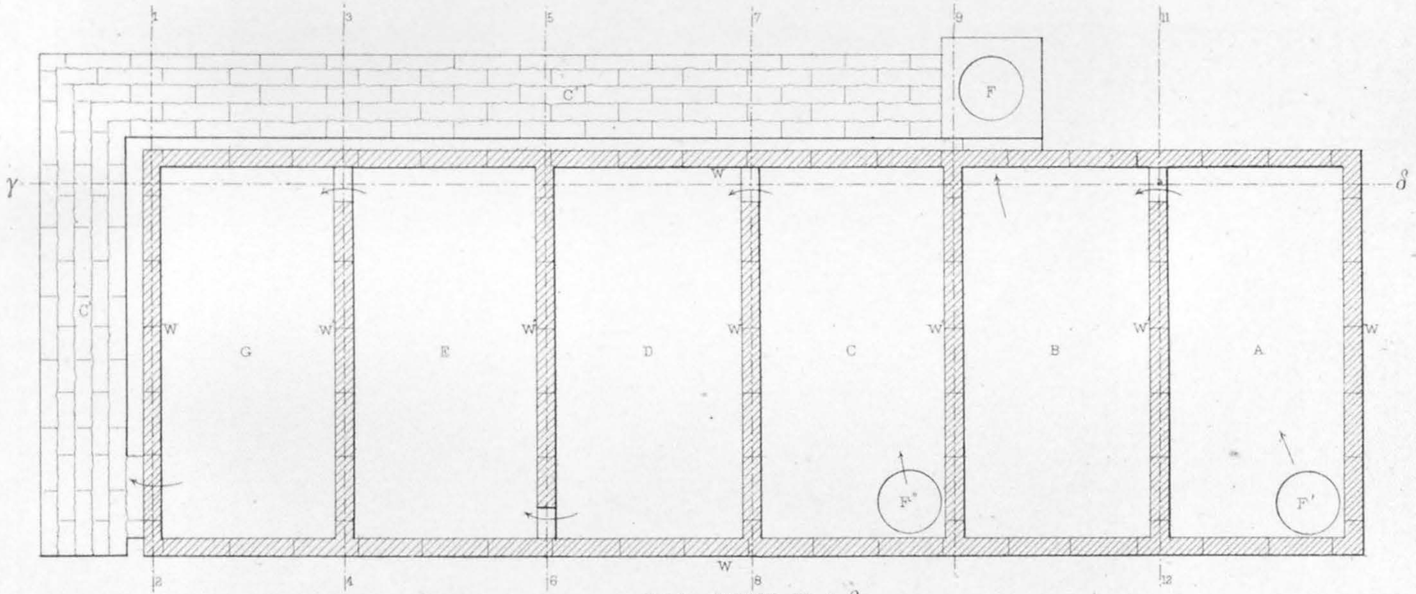


Fig. 2 SECTION ON α β

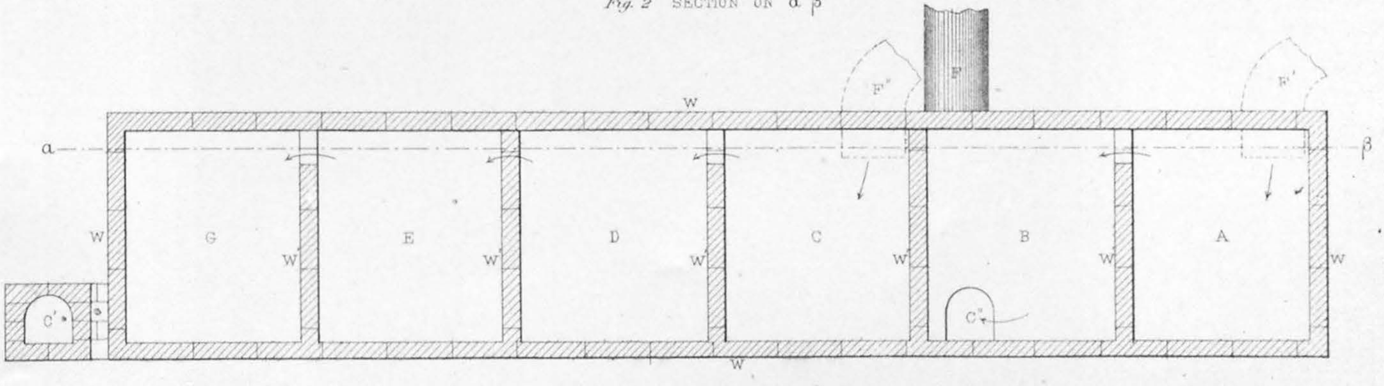


Fig. 1 SECTION ON γ δ

Scale: 1 inch to 12 feet, or 1/144.

M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

DUST CHAMBER

SMELTER C.

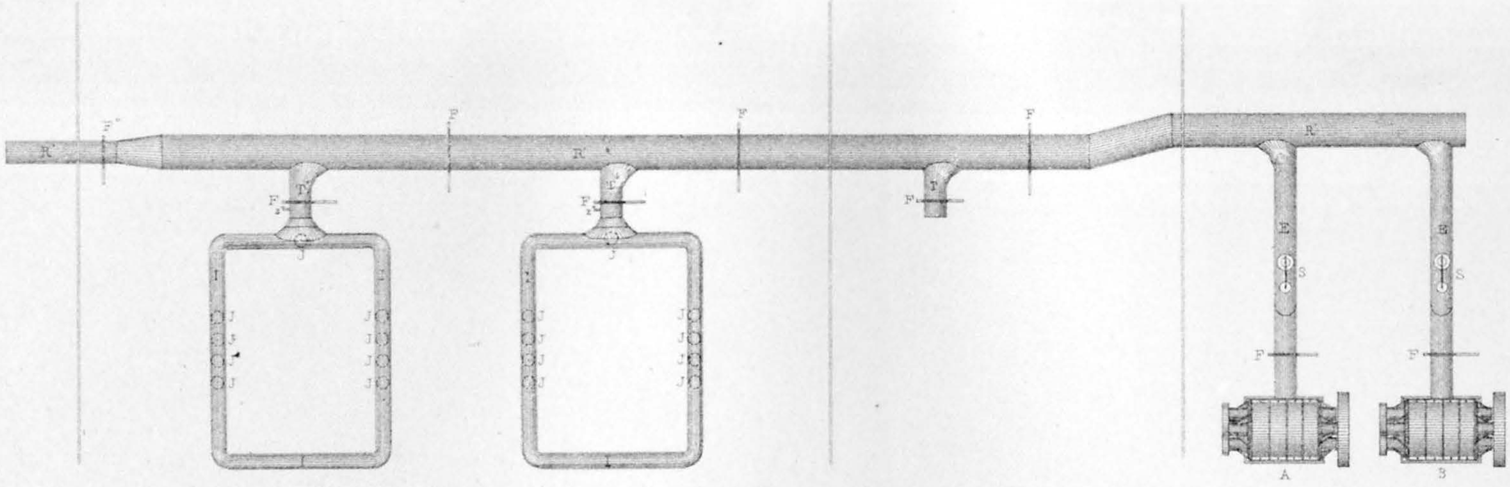


Fig. 1, BLAST ARRANGEMENT

Scale 1 inch to 12 feet, or 1/144.

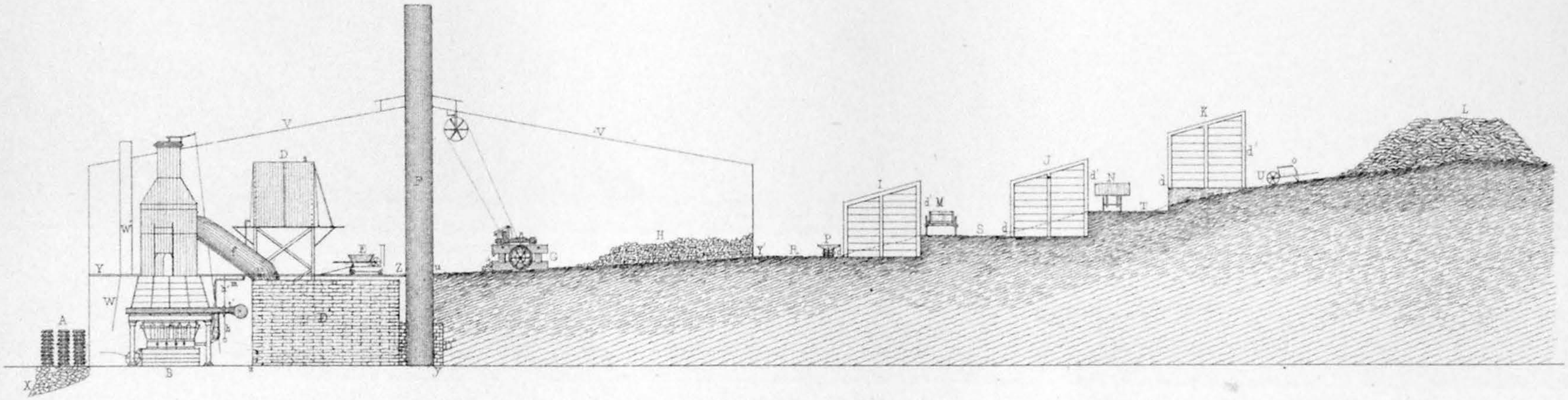


Fig. 2, ELEVATION OF WORKS

Scale 1 inch to 30 feet, or 1/360.

M Bien del.

A. Guyard, Metallurgist.
Julius Bien & Co. lith.
S.F. Emmons, Geologist-in-Charge

SMELTER C.

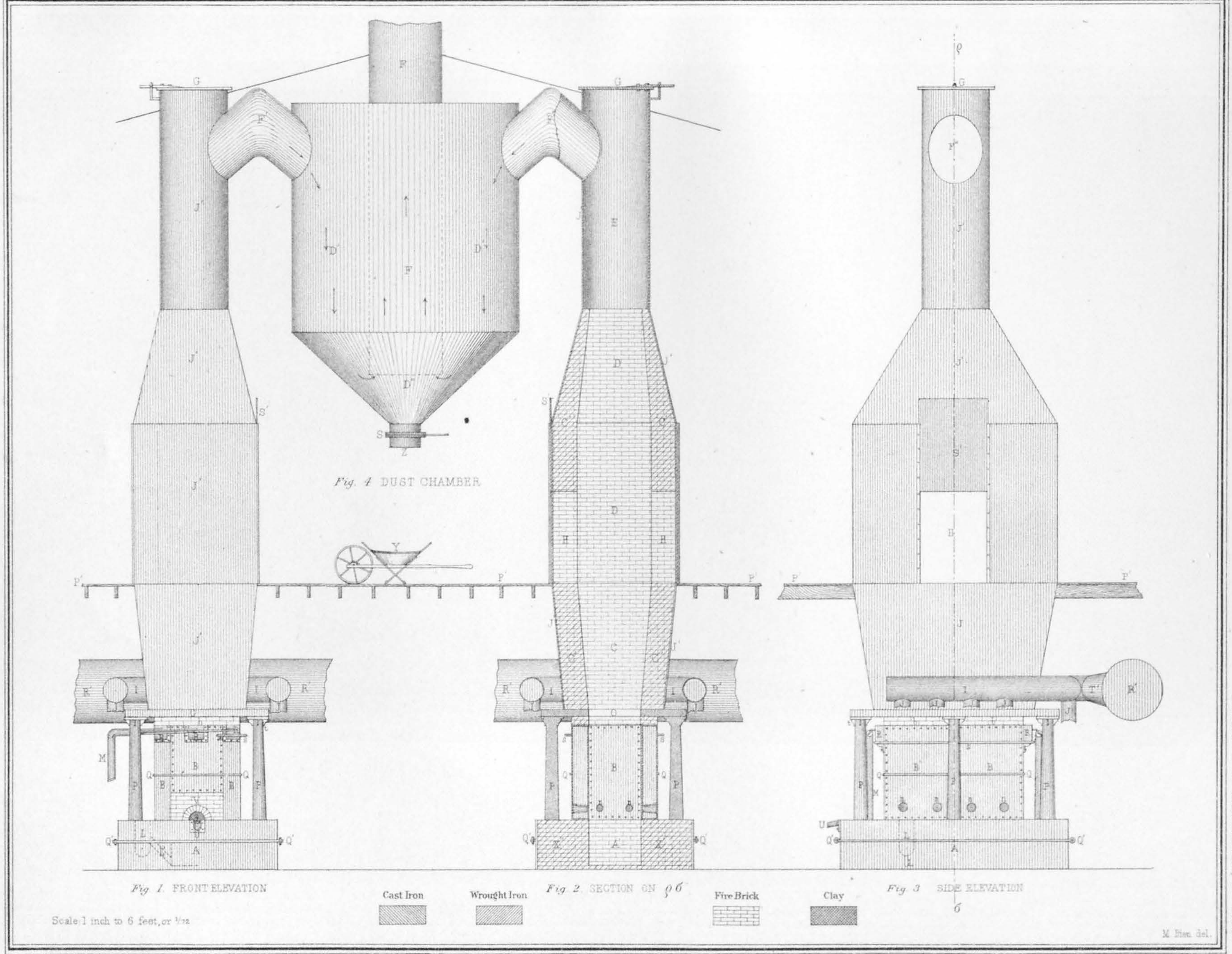


Fig. 4. DUST CHAMBER

Fig. 1. FRONT ELEVATION

Fig. 2. SECTION ON ρ σ

Fig. 3. SIDE ELEVATION

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

FURNACE AND DUST CHAMBER

SMELTER D.

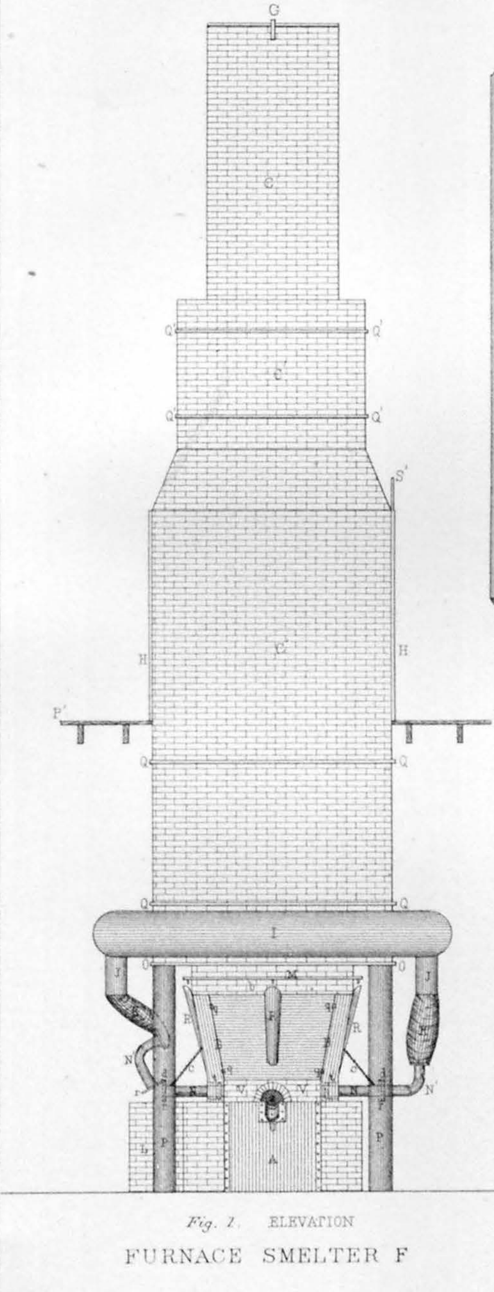


Fig. 1. ELEVATION

FURNACE SMELTER F

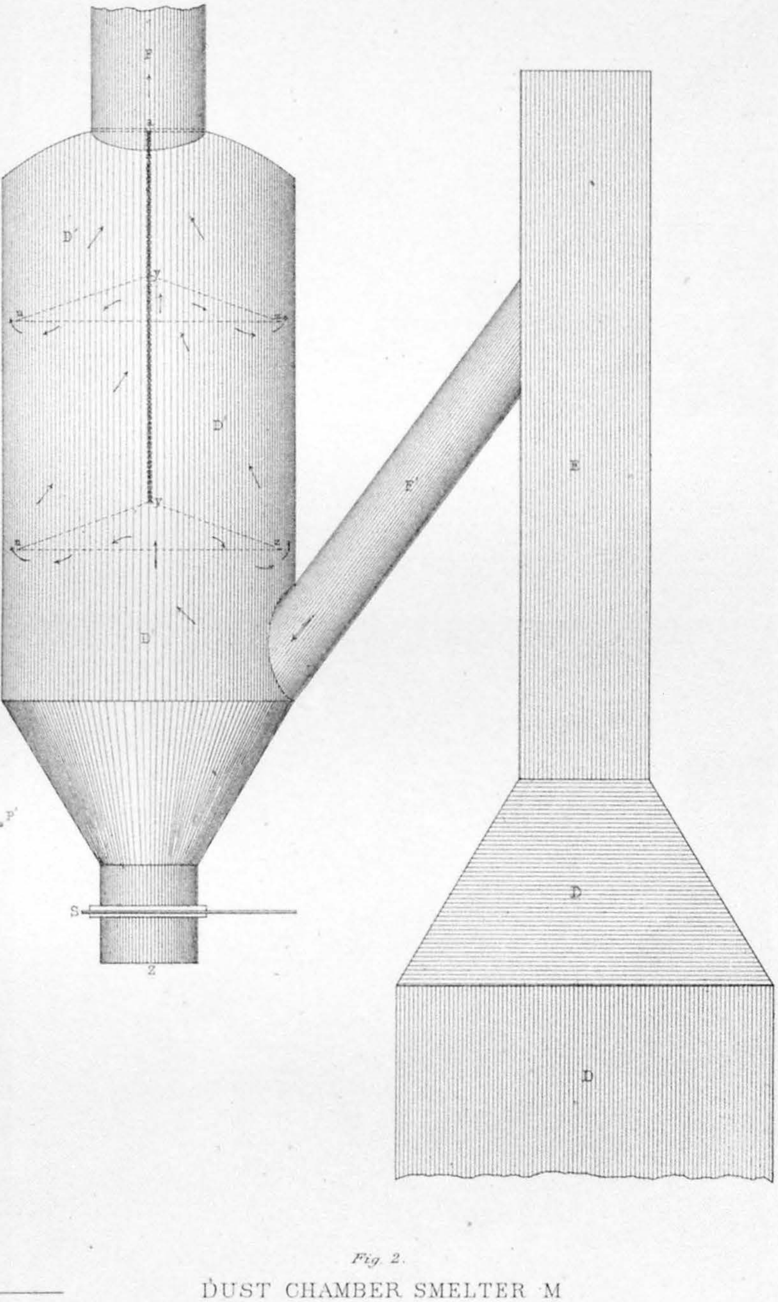
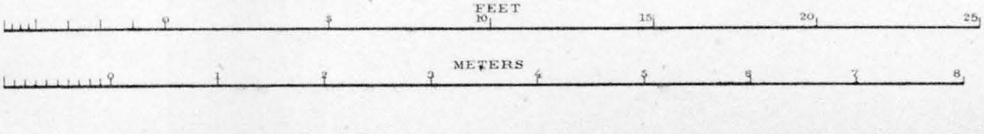


Fig. 2.

DUST CHAMBER SMELTER M

Scale: 1 inch to 6 feet, or 1:72



A. Guyard, Metallurgist. Julius Bien & Co. lith. S.F. Emmons, Geologist-in-Charge.

FURNACE AND DUST CHAMBER

SMELTERS F & M.

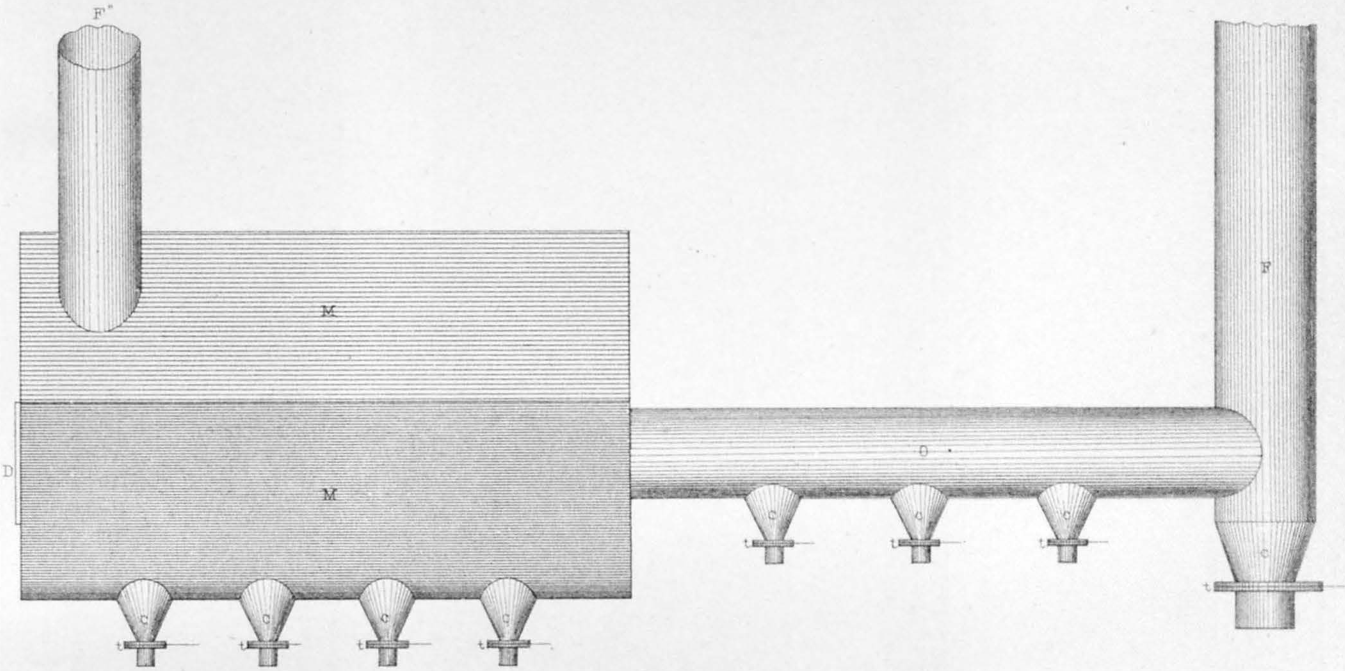


Fig. 1, SIDE ELEVATION.
Scale: 1 inch to 6 feet, or 1/72.

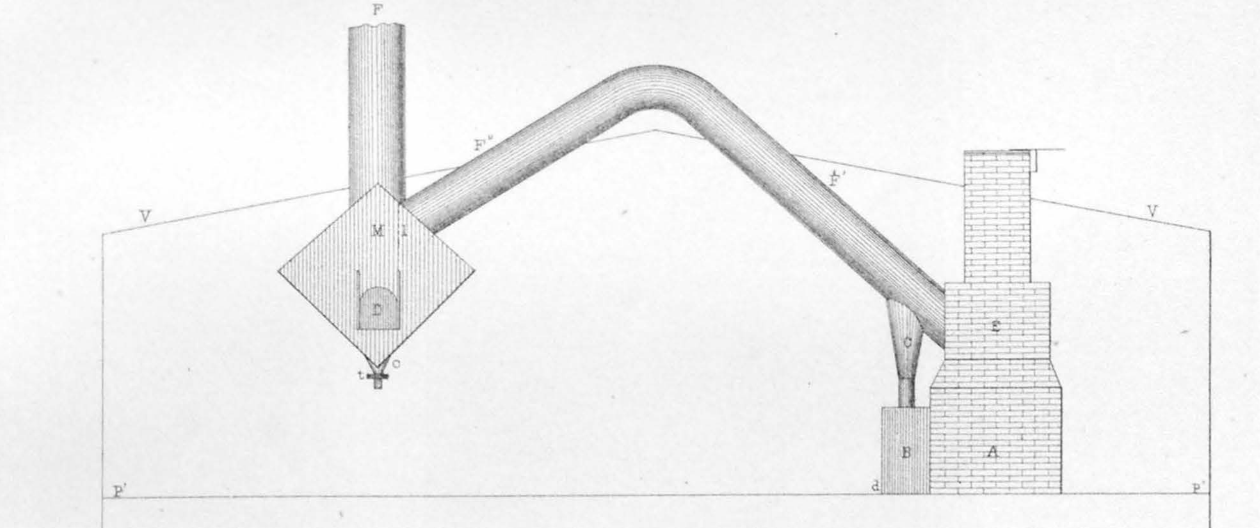


Fig 2, FRONT ELEVATION.
Scale: 1 inch to 12 feet, or 1/144.

M. Bien del.

A. Guyard, Metallurgist.
Julius Bien & Co. lith.
S. F. Emmons, Geologist-in-Charge.

DUST CHAMBER
SMELTER F.

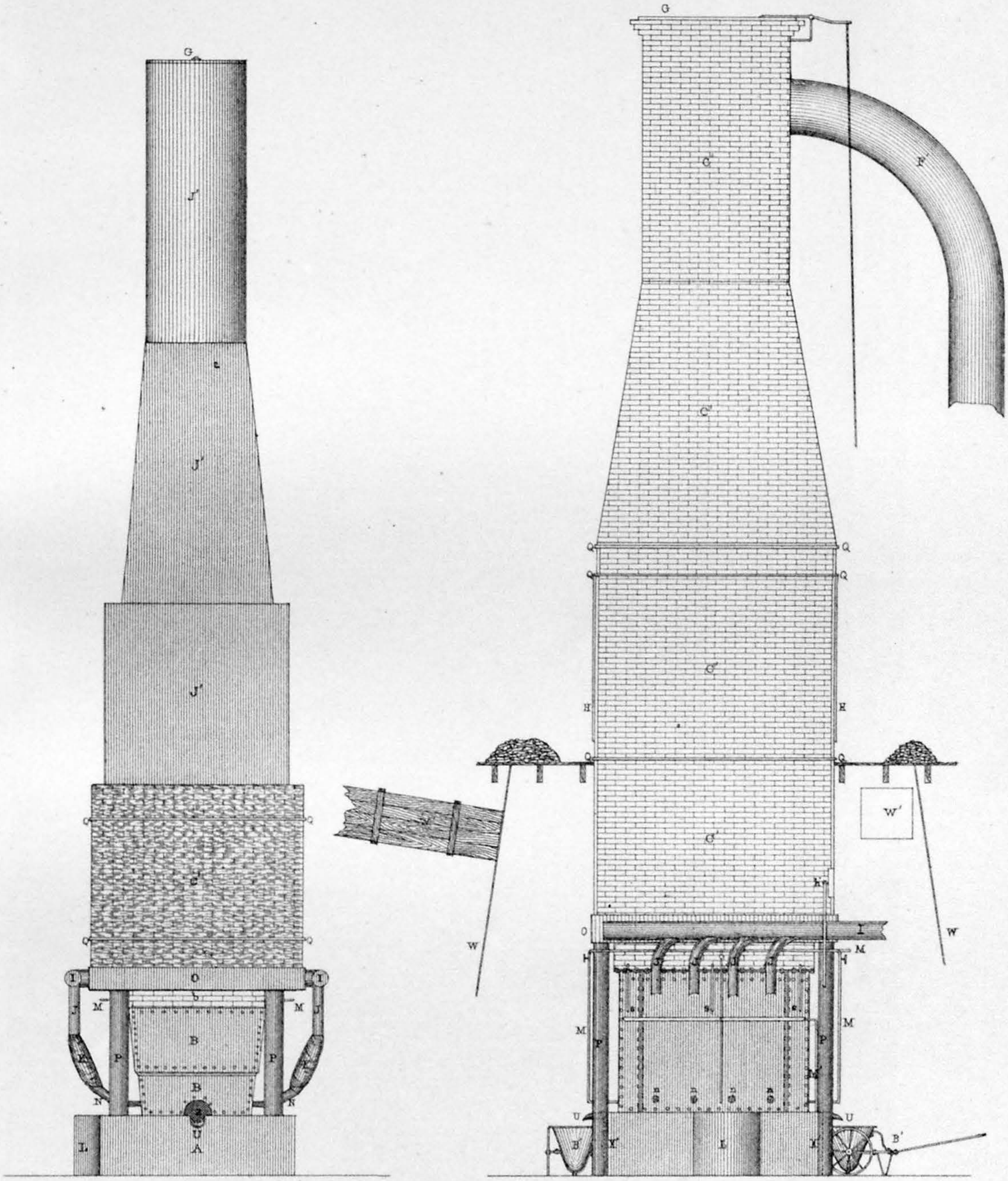


Fig 1. SMALLER FURNACE
ELEVATION

Fig. 2. LARGER FURNACE
ELEVATION

Red Brick ▨ Fire Brick ▦ Clay ■

Scale: 1 inch to 6 feet or 1:72

FEET
0 5 10 15 20 25

METERS
0 1 2 3 4 5 6 7 8

M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith. N.Y.

S.F. Emmons, Geologist-in-Charge

SQUARE FURNACES.
SMELTER G.

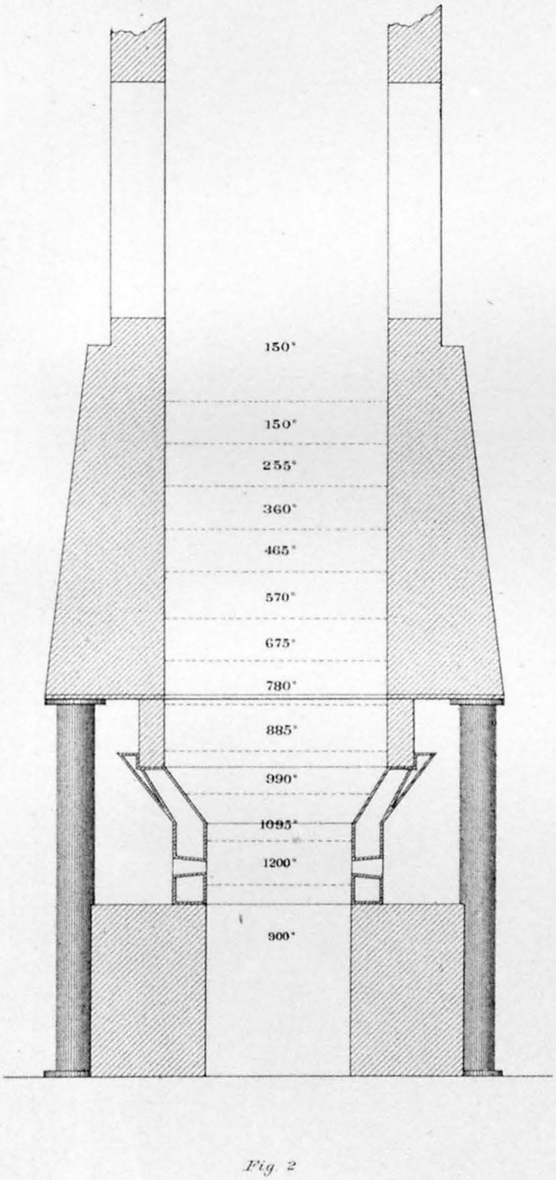


Fig. 2

Scale: 1 inch to 4 feet, or 1/48

SECTION OF FURNACE SHOWING
ZONES OF TEMPERATURE

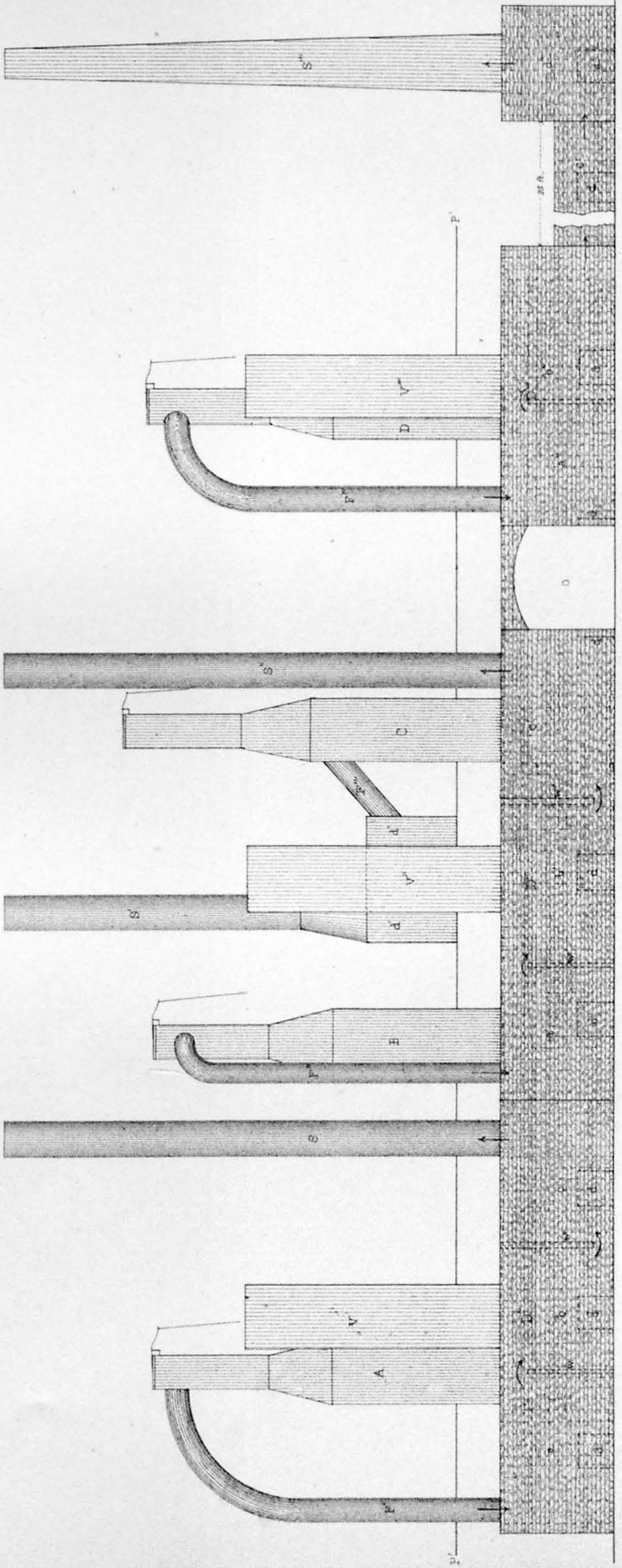


Fig. 1

ELEVATION

Scale 1 inch to 15 feet, or 1/180.

M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

SMELTER G.

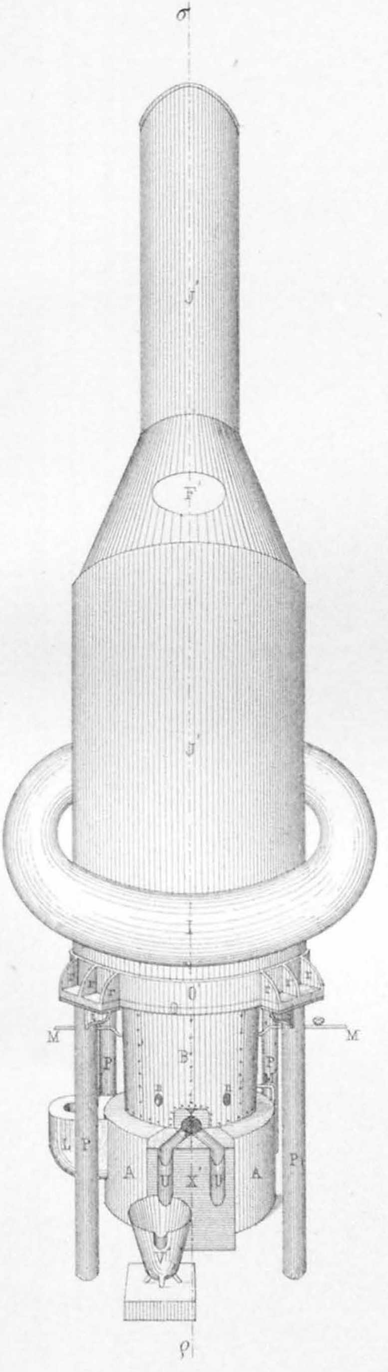


Fig. 1 PERSPECTIVE VIEW

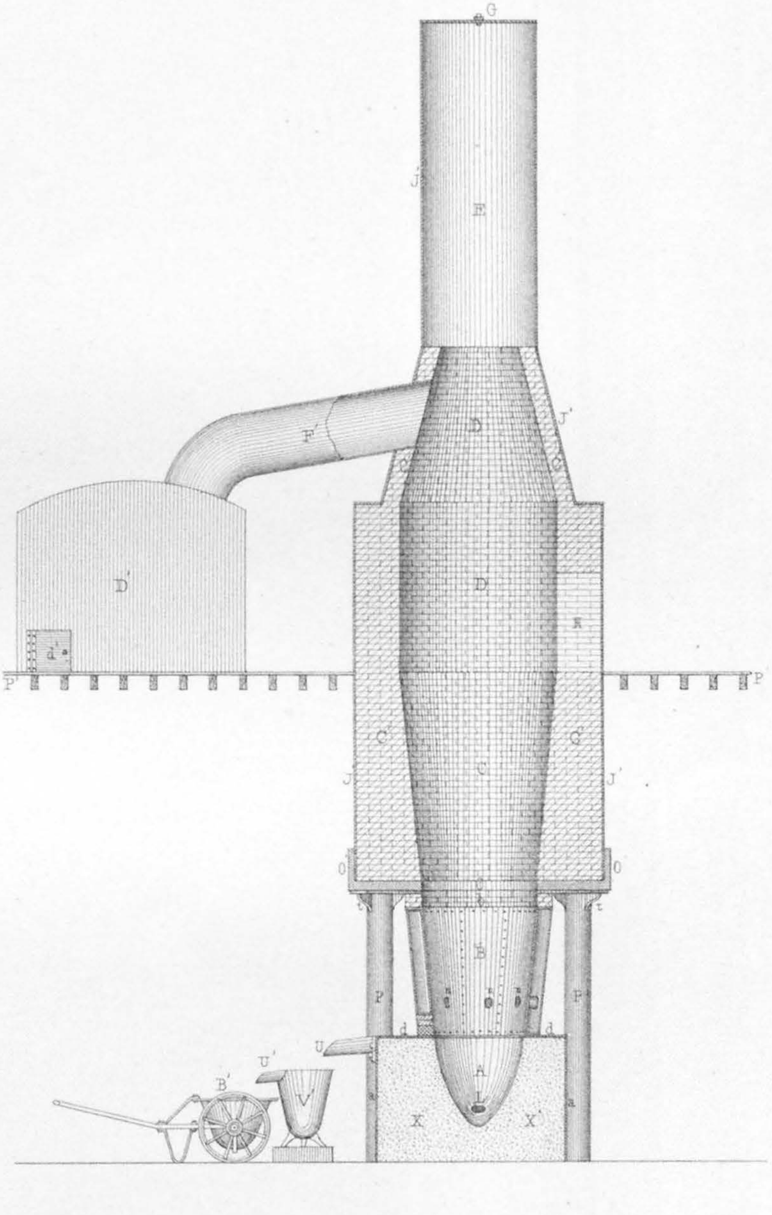
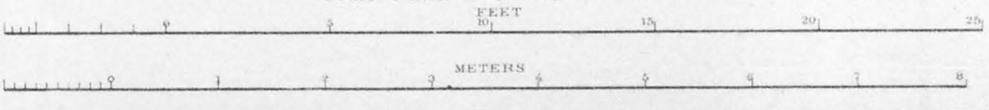


Fig. 2 SECTION ON ρσ

Cast Iron	Wrought Iron	Fire Brick	Clay	Steep
▨	▨	▦	▩	░

Scale 1 inch to 6 feet or 1:72



M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith. N.Y.

S. F. Emmons, Geologist-in-Charge.

CIRCULAR FURNACE
SMELTER H.

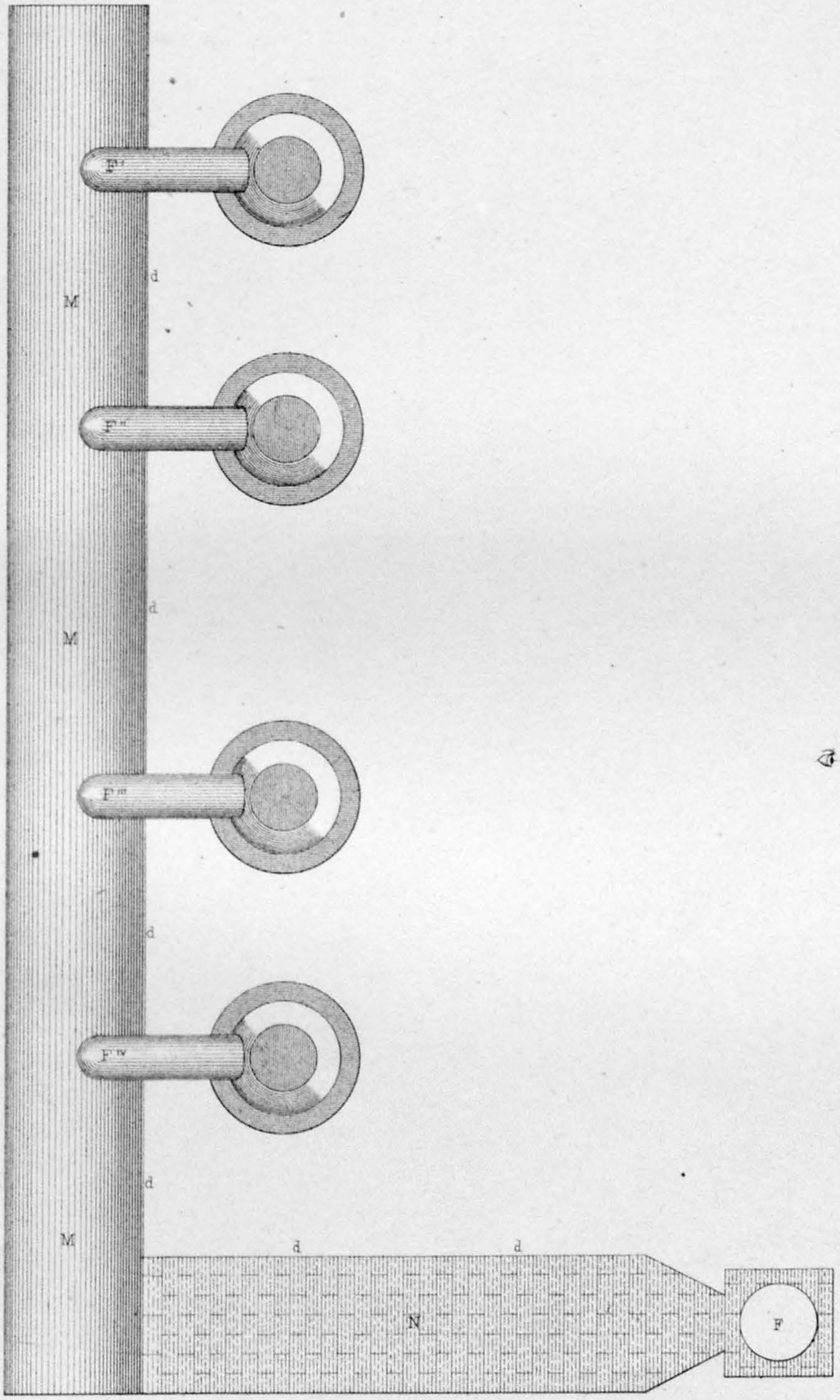


Fig. 1 DUST CHAMBER SMELTER H

Scale 10 ft = 1 in

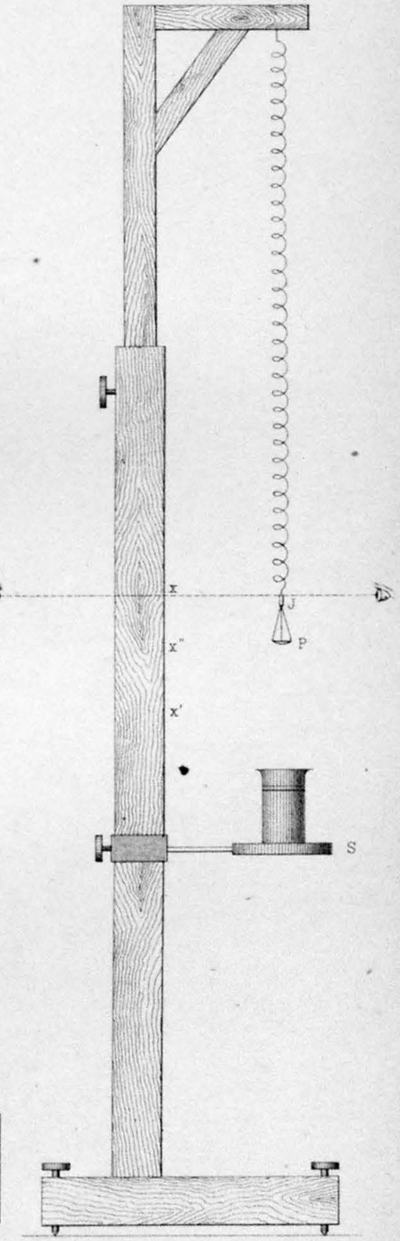


Fig. 2 JOLLY'S SPECIFIC GRAVITY SPRING BALANCE

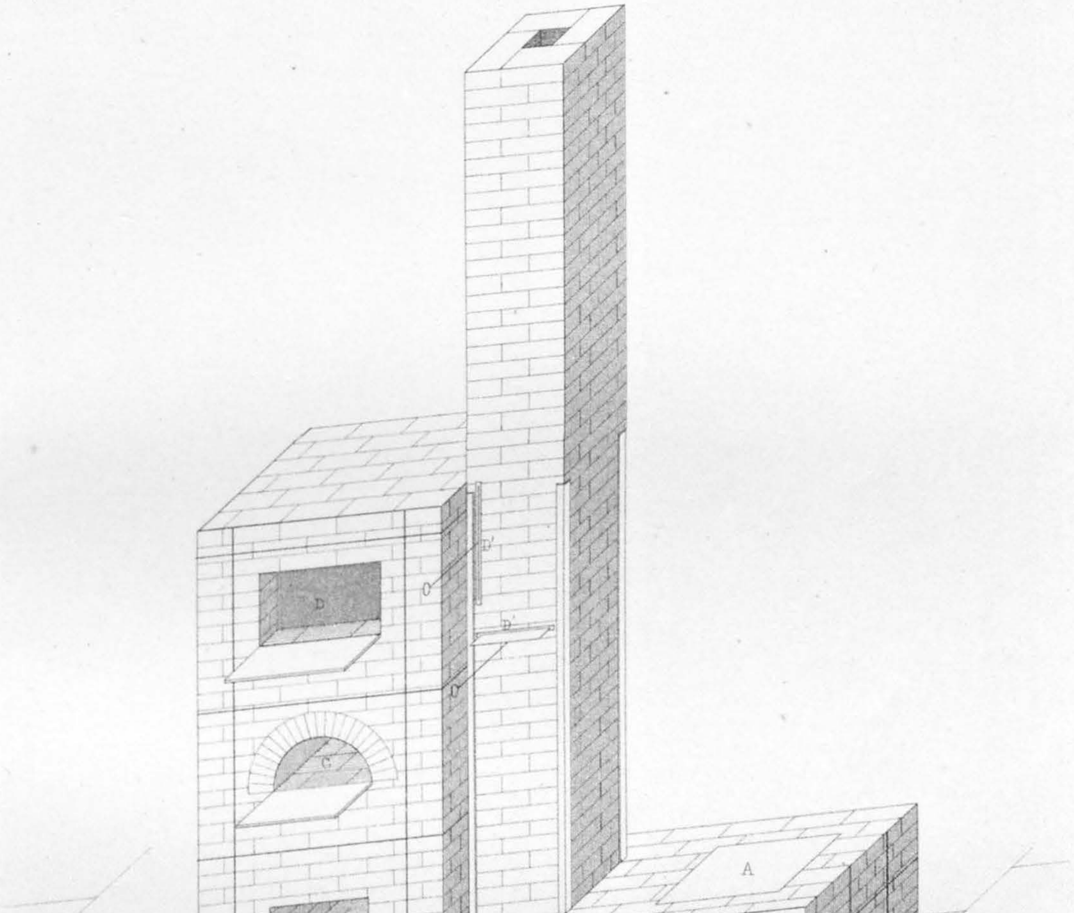
Scale: 1½ ft = 1 in.

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist in Charge

SMELTER H.



ISOMETRIC PROJECTION

Scale on vertical lines, 2 ft. to 1m. or $^1/_{24}$

M. Bien del.

A. Guyard, Metallurgist.

Julius Bien & Co. lith.

S. F. Emmons, Geologist-in-Charge.

ASSAY FURNACE
SMELTER H.

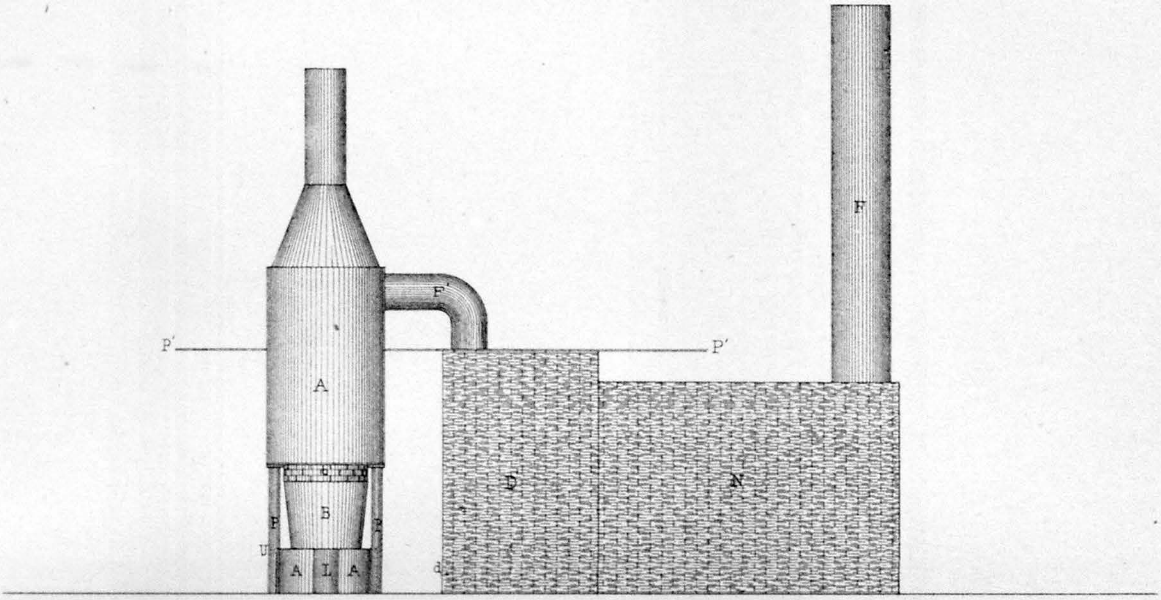


Fig. 1 SIDE ELEVATION

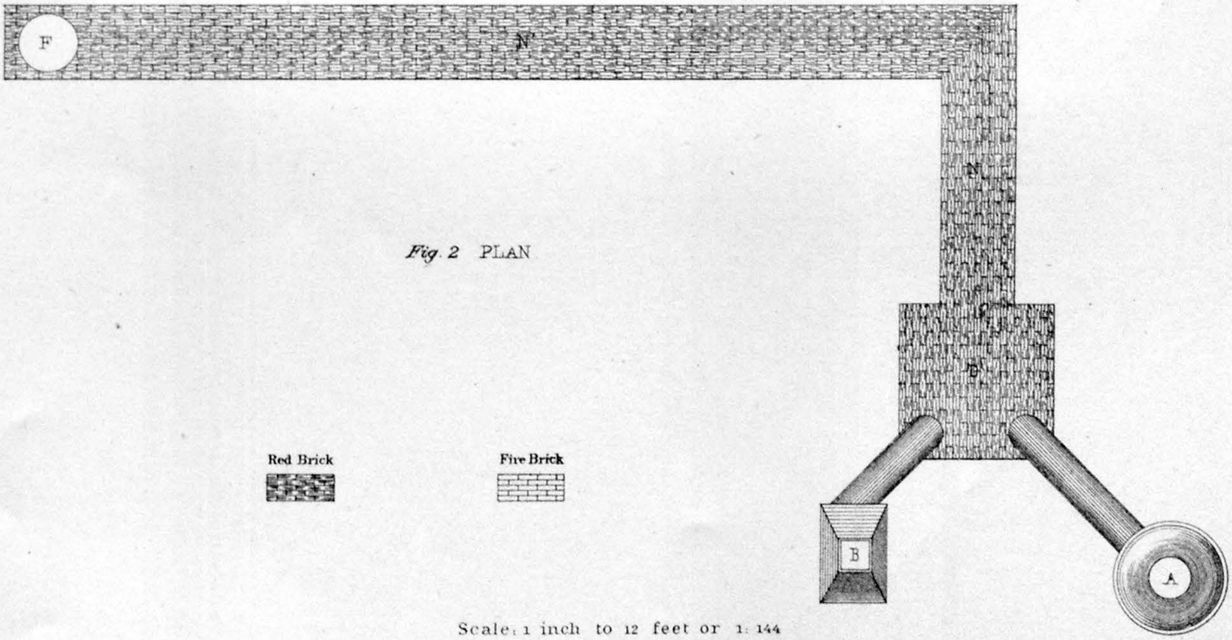


Fig. 2 PLAN

Scale: 1 inch to 12 feet or 1: 144

M. Bien del.

Julius Bien & Co. lith. N.Y.

A. Guyard, Metallurgist. S.F. Emmons, Geologist-in-Charge.

DUST CHAMBER.
SMELTER J.

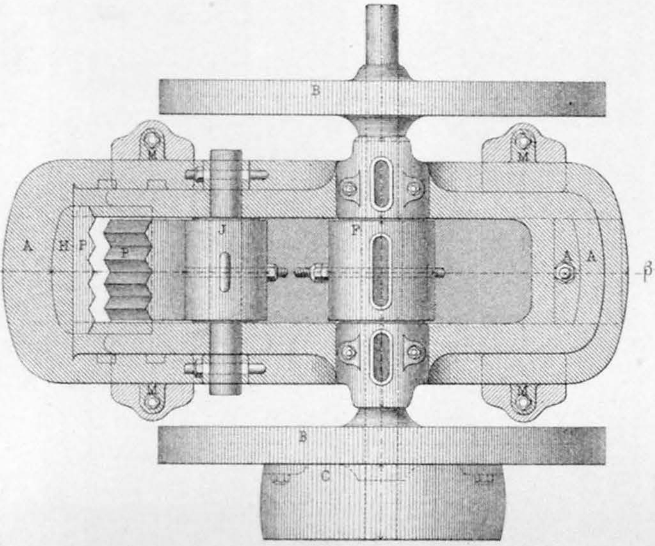


Fig. 1, HORIZONTAL PROJECTION

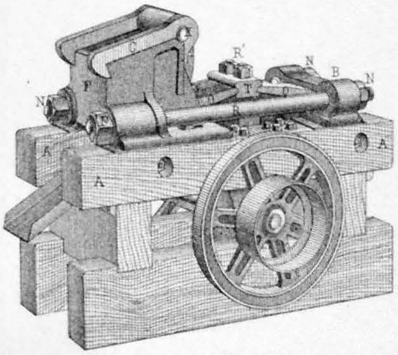


Fig. 3, PERSPECTIVE VIEW.

ECCENTRIC CRUSHER.

Scale: ⅛ in = 1ft.

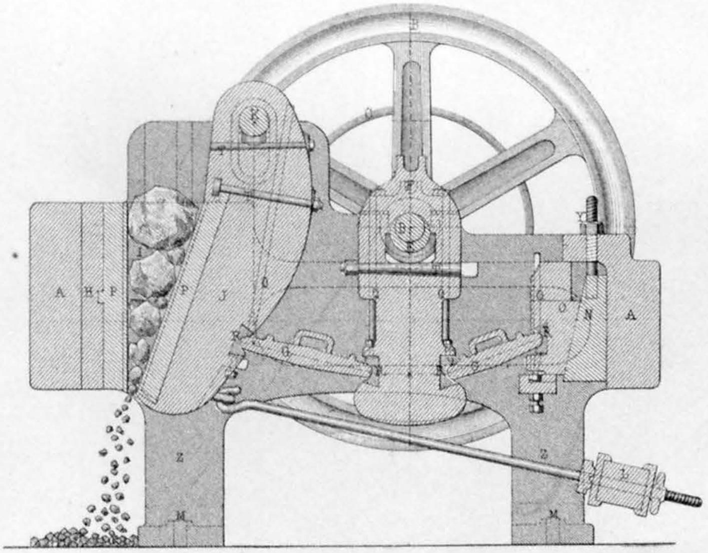


Fig. 2, SECTION ON α β.

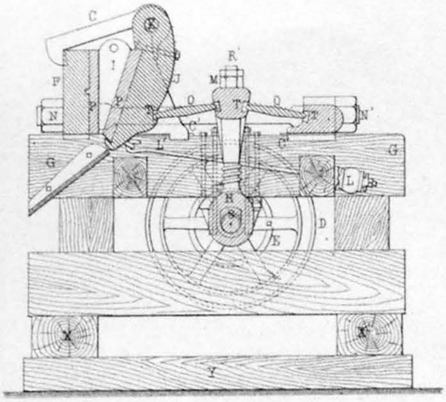


Fig. 4, VERTICAL SECTION.

CHALLENGE CRUSHER.

Scale ⅔ in = 1ft.

From Company's Catalogue

A. Guyard, Metallurgist. Julius Bien & Co. lith. S.F. Emmons, Geologist-in-Charge.

BLAKE CRUSHERS

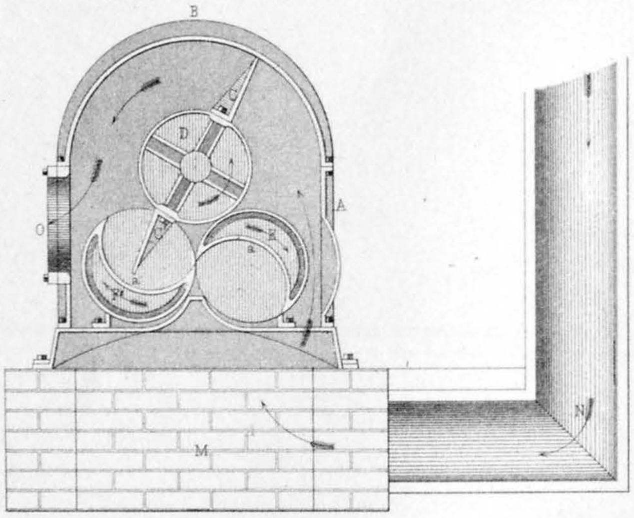


Fig. 2. SECTION
Scale: 7/16 in. = 1 ft.

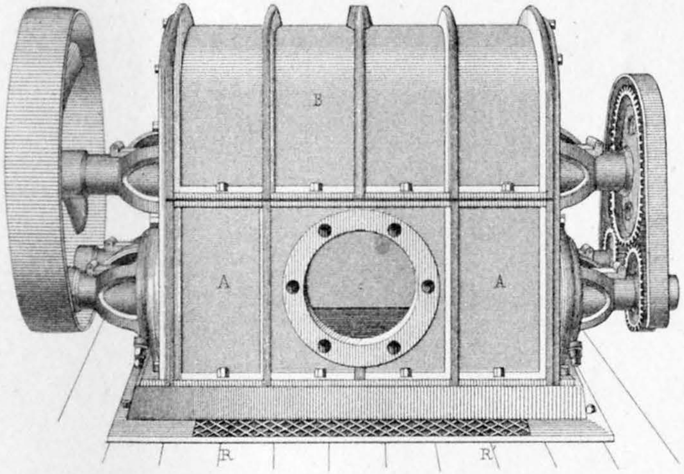


Fig. 1 PERSPECTIVE VIEW

BAKER'S ROTARY BLOWER, FORCED BLAST.

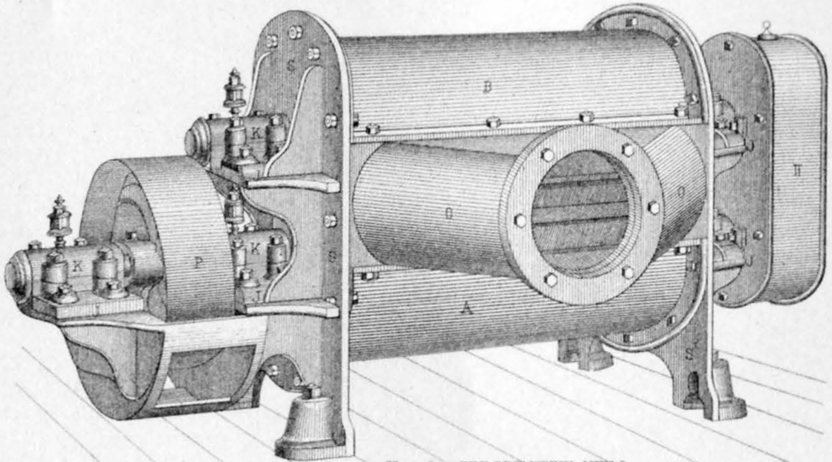


Fig. 3 PERSPECTIVE VIEW

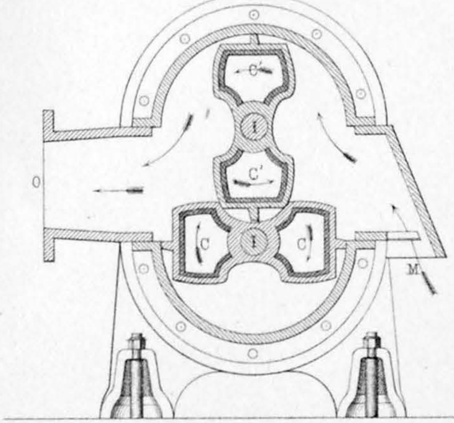


Fig. 4 SECTION
Scale 1/24

ROOT'S POSITIVE BLAST BLOWER

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A. Guyard Metallurgist.

Julius Bien & Co. lith.

S.F. Emmons, Geologist-in-Charge.

BLOWERS.

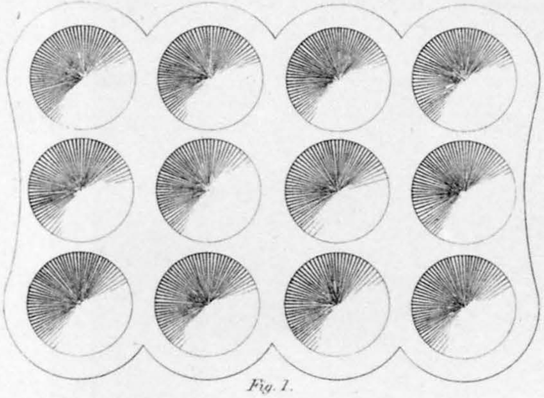

Fig. 1.

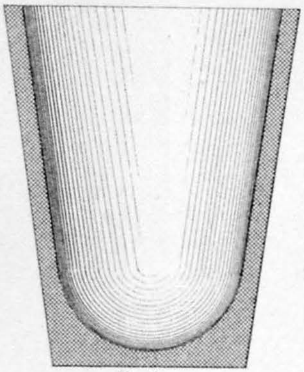

Fig. 3.
Scale, ½.

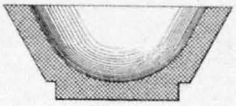

Fig. 4.
Scale, ½.

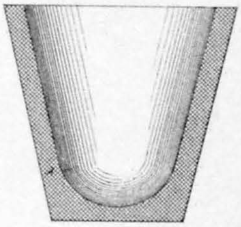

Fig. 5.
Scale, ½.

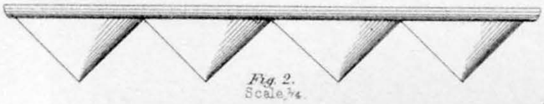

Fig. 2.
Scale, ¼.

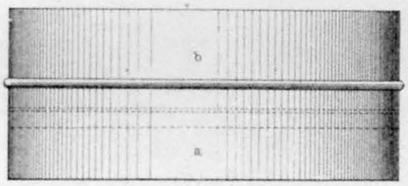

Fig. 8.
Scale, ¼.

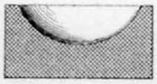

Fig. 6.
Scale, ½.

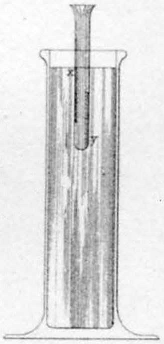

Fig. 7.
Scale, ⅛

Fig. 9
Scale, 1/12.

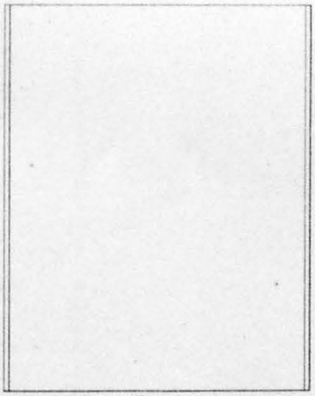

Fig. 10.
Scale, 1/12.

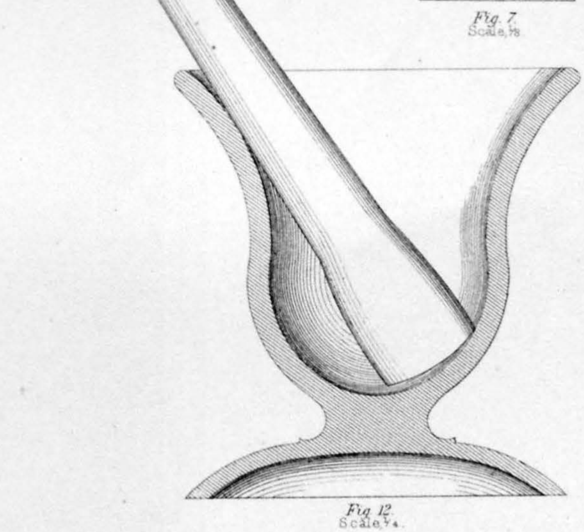

Fig. 12.
Scale, ¼.

Fig. 11.
Scale, ¼.

M. Bien del.

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ASSAY IMPLEMENTS

Scale 1/12

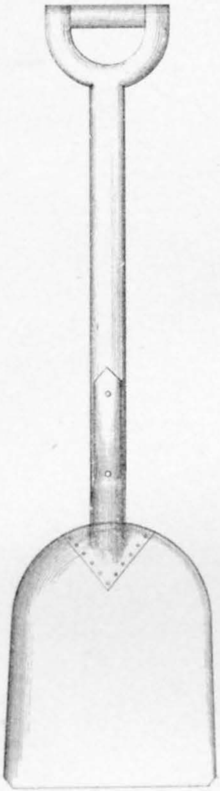
Fig. 1.

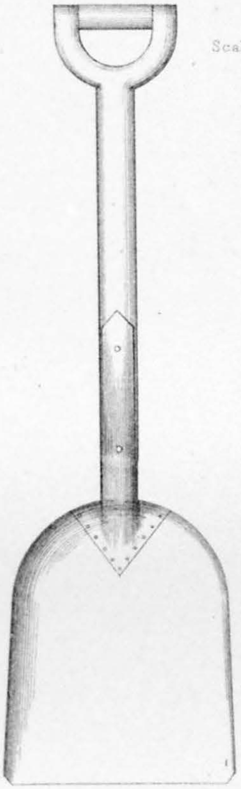
Fig. 2.

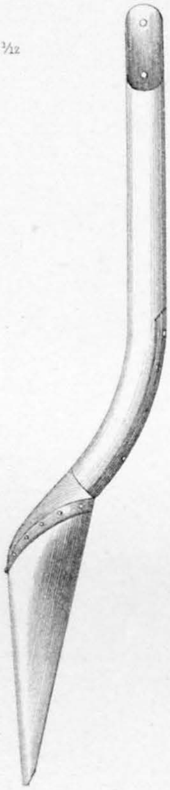
Fig. 3.

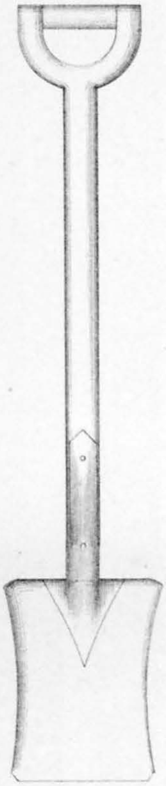
Fig. 4.


Fig. 5.

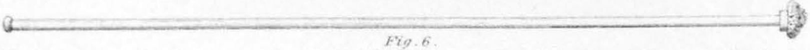
Fig. 6.
Scale, 1/24.


Fig. 7.
Scale. 1/24


Fig. 8.
Scale 1/24.

Fig. 9.
Scale 1/12.


Fig. 10.
Scale. 1/12.


Fig. 11.
Scale. 1/12.


Fig. 12.
Scale. 1/4.

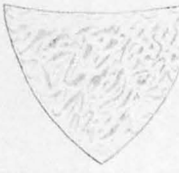
Fig. 13.
Scale, 1.


Fig. 15.
Scale, 1/2.

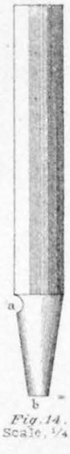

Fig. 14.
Scale, 1/4.

M. Bien del.

A. Guyard, Metallurgist. Julius Bien & Co. lith. S.F. Emmons, Geologist-in-Charge.

SMELTER'S IMPLEMENTS

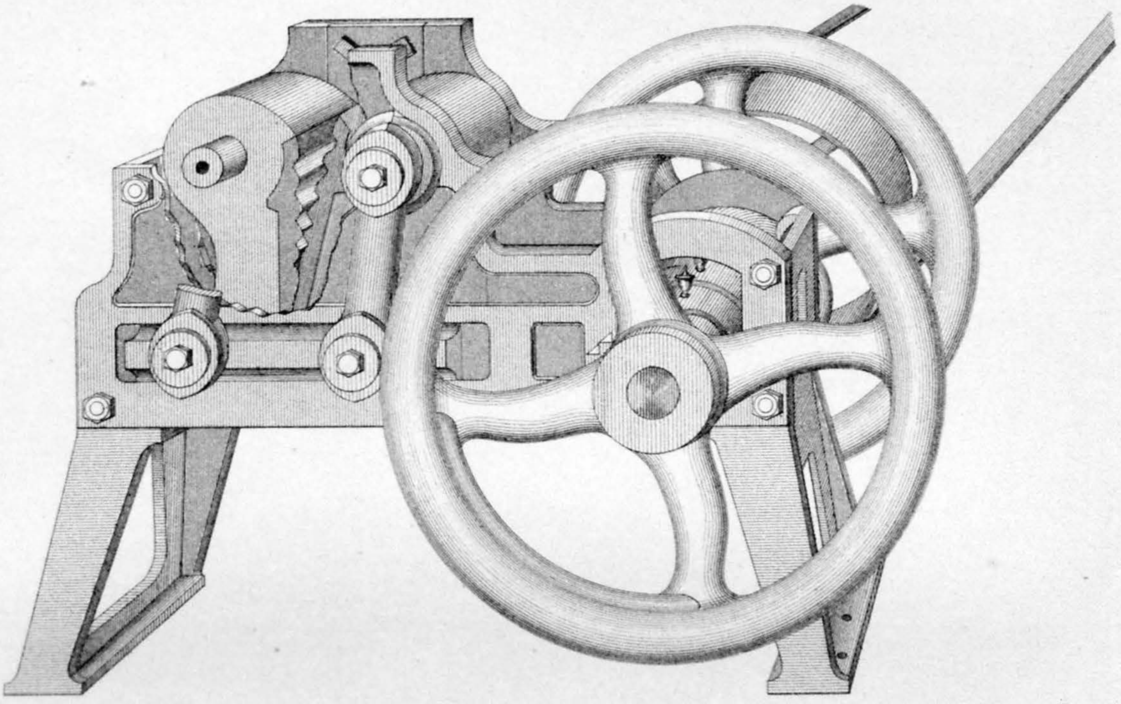

Fig. I ALDEN ORE CRUSHER.

From Company's Catalogue

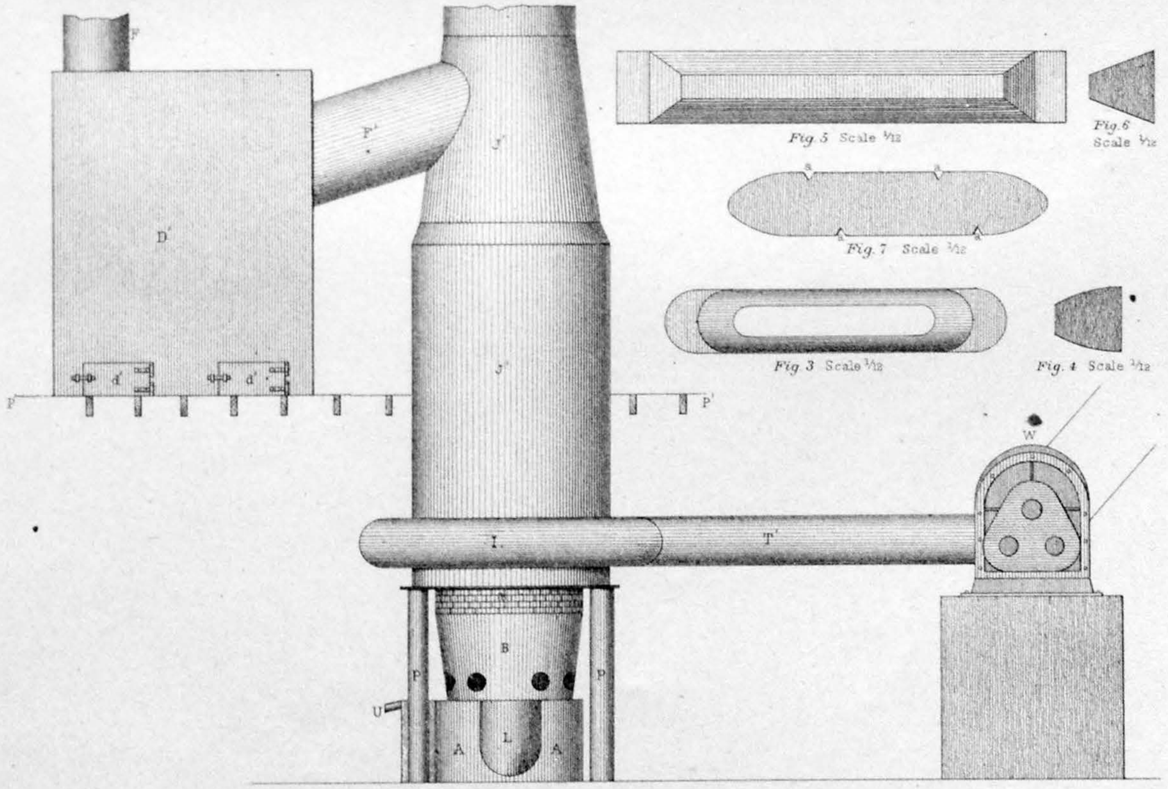

Fig. 5 Scale 1/12

Fig. 6 Scale 1/12

Fig. 7 Scale 1/12

Fig. 3 Scale 1/12

Fig. 4 Scale 1/12

Fig. 2 FURNACE AND DUST CHAMBER SMELTER I

Scale 1 inch to 6 feet, or 1/72.

M. Bien del.

A. Guyard, Metallurgist. Julius Bien & Co. lith. S. F. Emmons, Geologist-in-Charge.

Furnaces.—The smelting plant consists of three large square furnaces, made of brick. With the exception of the water-jackets they are exactly similar in shape, dimensions, and capacity. They are represented in elevation and in section on Plate XXVI. There are besides six circular furnaces, manufactured by Messrs. Fraser & Chalmers, of Chicago. These furnaces are represented in elevation and section on Plate XXVII.

A full description of both the square and the round furnaces at this smelter has already been given at the commencement of this section, to which the reader is referred for further details concerning their construction. The capacity of the square furnaces is 35 to 40 tons of ore per 24 hours, and that of the circular ones 20 to 22 of ore in the same time. The heat radiated in front of these nine furnaces, closely packed in a narrow space, is so great that the men are obliged to constantly water the ground in front of them. The crucibles of both square and round furnaces are made of steep, the mixture preferred at these works being two parts of fire-clay and one part of coke-dust. The length of runs is various and averages three months.

Condensing chambers.—The condensation of lead fumes at this smelter is of the poorest description. Four furnaces are totally without any condensing apparatus; three furnaces are connected with an oblong sheet-iron box, 50 feet long, 6 feet high, and 6 feet wide, placed on the feeding-floor; and one furnace communicates with a small chamber, 9-foot cube, placed on the same floor. Each chamber is provided with a stack made of sheet-iron. With such inadequate arrangements the works are perpetually enveloped in a thick atmosphere of smoke and lead fumes, and "leading" is of frequent occurrence.

Bartlett filter.—During the collection of notes for this report, a most interesting and valuable experiment was being made at these works with a view to the total condensation of lead fumes. This experiment was carried on at great expense with the elaborate apparatus known as the Bartlett smoke-catcher or filter. It was so successful and the results derived from it are so interesting that the whole deserves a full description. The arrangement adopted is shown in Fig. 1, Plate XXVIII. The stack *E* of one of the square furnaces *A* was connected by means of the sheet-iron flue *F'* with a Sturtevant fan, *B*, drawing the fumes from the furnace and blowing them through the sheet-iron pipe *P*, about one hundred and fifty feet long, where they parted with their dust, as in an ordinary flue. The pipe *P* is connected by means of two sheet-iron branch pipes, *P'*, with two thin sheet-iron boxes, *a a'*. Each branch pipe *P'* is provided with a damper or valve, exactly similar to those used in common stove-pipes for the regulation of draft, so that the fumes can be distributed evenly in the boxes or shut off from one and allowed to enter only the other. Each box is formed of two parts, the dust-chamber *a a'* and the fireplace *N*. The dust-chamber is provided with sliding doors *O*, placed at each extremity, and the fireplace with doors *d*, placed in front, and sheet-iron pipes *L* at the back which communicate with a stack, *L'*. The chambers *a a'* are provided at the top with 28 apertures, to each of which is fastened a cloth bag, *b*, 30 feet high, suspended to the beams of the light wooden structure in which the apparatus is inclosed. This building *M* is provided with very large openings for ventilation. When the apparatus is at work the fumes, blown in through pipes *P P'*, distribute themselves in the dust-chambers *a a'* and ascend the cloth bags, through which they filter. The gases come out perfectly colorless and are entirely deprived of any lead dust or even soot. The wind, entering freely through the apertures provided in the light building, shakes the bags, and the dust with which they are charged falls into the dust-chambers. When a sufficient quantity of this dust has accumulated there, the doors *O* are opened and a light wood fire is placed through doors *d* in the fireplace *N*. The soot of the dust soon catches fire, and the dust, which was quite black, like lampblack, becomes white; it becomes also denser by this operation and is more easily manipulated. When the smoke has thus been calcined it is shoveled out through doors *O*.

During a run of five days 3,030 pounds of this calcined dust was caught in the Bartlett filter from one furnace, but the experiment was not altogether satisfactory for the reason that the furnace was worked with an open feed-hole, as with an ordinary dust-chamber, and that the Sturtevant fan was drawing as much air as smoke, so that the damper *G* of the furnace had to be left half open and about half the smoke was lost. In the conditions in which the experiment was made this could not be avoided, but this is only an experimental defect, impairing in no way the value of the filter, which does its work to perfection; the writer estimates at 7,000 pounds the quantity of dust which would have been caught in five days had the experiment been made with closed feed-hole and damper, or say at 1,500 pounds per twenty-four hours. The calcined dust has been assayed by Dr. M. W. Iles, and found to contain 70 per cent. of lead and 6 ounces of silver to the ton; so that with a furnace of 35 to 40 tons of ore capacity per twenty-four hours one-half a ton of lead is lost in the air, as well as 4.5 ounces of silver, in twenty-four hours. The result of this is that the quantity of lead lost in the air is greater than the quantity of dust condensed in the dust-chambers. At smelter A, where the dust-chamber arrangement is of the best kind, 150 tons of dust were collected in 182 days, giving 1,648 pounds of dust per twenty-four hours for two furnaces whose joint smelting capacity in tons of ore is equal to that of the furnace connected with the Bartlett filter. This dust, assaying 35 per cent of lead, represents 577 pounds of lead, or a little over a quarter of a ton. These are indeed important results and are worth considering.

At smelter B, as at most smelters, chamber-dust is mixed with milk of lime; the mixture is spread over ore-beds and resmelted in this way. The composition of the dust caught in the Bartlett filter is extremely remarkable. It has been analyzed, such as it is previous to calcining, and the results will be found in the study of lead fumes, Analysis XXXVI.

SMELTER C.

Disposition of works.—These works, like all the other smelters situated in California gulch, are erected on the southern slope of the northern bank of the gulch, the gentle slope of this hill favoring singularly the construction of similar establishments. A glance at the vertical section through these works (Fig. 2, Plate XXXI) will show their general disposition. A deep cutting, *u y z*, in the bank is the only one needed for the erection of the furnaces *B* and dust-chamber *D′*. In front of the furnaces and extending some little distance is the slag-heap *X* upon which are seen piles of bullion in bars, *A*, ready for shipment. At the foot of the slag-heaps runs the lower road of California gulch. The furnace *B* is connected by means of the sheet-iron flue *f* with the dust-chamber *D′*. The water-jackets of the furnace are supplied with water from the main water-pipe *m*, connected with the tank *D*, placed on the feeding-floor *YZ*, and constantly filled with water by means of pumps worked by machinery. The sheet-iron stack *F*, communicating by a flue, *D''*, with the dust-chamber *D'*, carries off the smoke. On the feeding-floor, and in close proximity with the feed-holes of the furnaces, are disposed the fuel and old slag used as flux, as well as the mixtures of ore, dolomite, and hematite entering into the composition of smelting charges. On the portion *Z Y'* of the feeding floor are seen the crusher *G* and the ore-beds *H*, on the top of which are placed bricks of flue-dust and lime, specially molded in this form previous to resmelting with the ore. To the right and left of the ore-beds are rows of ore-bins, supplied from the wagon-road, *R*. The scales, *E*, placed on the feeding-floor, are used for weighing the different elements of the smelting charges, special ores and fluxes, which are classified and distributed right and left on each side of these scales. The portion of the works placed between *Y* and *Y'* is inclosed in a light timber construction, *V*, protecting the workman, the plant, and the works against rain, wind, and cold.[1]

I represents a second row of ore-bins supplied from the wagon-road *S*. These bins are made of light timber; they are opened in *d* for the removal of the ore, which is wheeled to the crushers in the wooden ore-barrows *P*. The bins are supplied from the ore-wagons *M* through the aperture *d'*, which can be closed by hinged wooden doors. *J* is the third row of ore-bins, exactly similar to the preceding. On road *T* is laid a railroad track, a siding of the Denver and Rio Grande Railway. *N*, represents a car. In *K* is a fourth row of ore-bins, similar to those previously described, and at *U* a wagon-road placed in direct communication with the upper level district road. At the extreme wings of the row of bins *K* are to be seen huge fuel-bins for both coke and charcoal. *L* represents a heap of charcoal placed on the upper level of the works; on this level are also ore-dumps and heaps of dolomite and hematite. Inclined ways run through the rows of ore-bins, connecting the different levels, which allow the wheeling down of ores, fluxes, and fuel. The charcoal-barrows, made of thin sheet iron, are represented in *O*; these barrows, in general use in the camp, hold about eight bushels of charcoal. The feeding-floor is connected with the furnace floor by means of a flight of steps placed outside the main building, and also by zig-zag inclined ways for the wheeling up of old slag to be resmelted.

Standing on the slag-heap and facing the furnace, but not shown in the sketch, are, on the left, the scales upon which the bars of bullion are weighed, and on the right, the boiler, engine, and blast-apparatus rooms. Farther on the right stands the shed in which flue-dust is mixed with lime, molded into bricks, and desiccated on driers artificially heated.

The offices, staff apartments, assay offices, and laboratory occupy a detached building situated a short distance from the works and on a level with the lower road of California gulch. A small office, provided with large Fairbanks scales, is placed at the entrance of the works at one of the upper branch roads. The assay offices and laboratory are well fitted up. The muffle furnace and crucible furnace are separate. In the assay of ores and slag for lead, iron rods are always inserted in the crucible. Experiments on the fusibility of mixtures of ores and fluxes are also made. The assays for silica are always evaporated in order to obtain the percentage both of gangue and of soluble silica.

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[1] This building is not correctly represented in the drawing, the ridge-pole and ventilator being at the top of the furnace *B*, instead of at the stack *F*, as there shown. (S. F. E.)

Furnaces.—Smelter C has three furnaces of equal dimensions and capacities, constructed on the plan of Messrs. Keyes & Arent's patent. At the time this report was made, however, only two furnaces were running, and all that has been said or will be said of this smelter refers to the work done by the two furnaces.

They may be considered the model furnaces of Leadville, both as regards appearance and working qualities, and will therefore be described in considerable detail. One of them is represented on Plate XXIX, in elevation (Fig. 1), transverse and longitudinal sections (Figs. 2 and 3). The crucible *A* of this furnace differs essentially from nearly all the others used in Leadville in that the lead-well *L* does not project outside of the crucible frame, but, together with the crucible, is confined within the frame formed by the hearth-plates. A glance at this furnace and at any of the others, with exception of the furnace shown in Plate XXXII, will show the difference. Another peculiarity is that lead is not ladled out of the lead-well, as at other smelters, but a tap hole, *z'*, is made in its clay lining, and the bullion is drawn periodically into the cast-iron lead-pot *T'*, mounted on a small cast-iron stove, in which a slow fire is placed in order to keep the bullion molten. The chimney *f* of this stove communicates under ground with the dust-chamber. The advantage of this disposition is twofold: the lead-well may remain constantly covered and its bullion be kept at a high temperature, thus assisting the clearing of the siphon when this is necessary. On the other hand, the bullion accumulating in the lead-pot, being kept molten, can be ladled rapidly into the molds, and the bars thus obtained are of a good shape and uniform composition. The entire hearth rests on a bed-plate of boiler-iron, with an angle-iron rim, which incloses the base of the hearth-walls or lining of the crucible. These walls *X* are entirely of fire-brick, but the dam of the crucible *A''* is protected by tamping or pressed fire-clay, as are the siphon *L'* and the lead-well *L*. The bottom of the crucible is formed by an inverted fire-brick arch, with quartz-brasque beneath, separating it from the bed-plate.

The hearth is inclosed on the sides by four cast-iron plates, *a*, each 1 inch thick, of which the front and back ones have each 2-inch flanges lapping over the ends of the side plates. These plates are firmly held together or inclosed by three rows of bars or rails, *Q'*, which are fastened at each corner by wrought-iron rings *p*. The top of the hearth is also covered by iron plates, *d*. To the front hearth-plate is screwed and bolted the slag gutter or spout *U*, and to the side plate the lead-gutter, *U'*, leading from the lead-tap *z*.

These are the only furnaces in which the fore-hearth does not project outside of the frame of the hearth. The crucible, though constructed on the same general principles as those of other furnaces in Leadville, differs essentially in the arrangement of details.

The water-jackets are thirteen in number, one in front, two at the back, and five on each side. They are screwed together and wedged at *q*, and braced by tie-rods; the tie-rods under the hot-water outlets *S* are not indicated on the drawing. The front water-jacket does not extend down to the hearth-plate, but rests on a fire-brick wall, in the middle of which is a small water-cooled cinder-block inclosing the slag-hole (see Fig. 2), instead of the ordinary tymp-stone shown in Fig. 1, Plate XXXIII.

The water-jackets are otherwise similar in construction and appearance to those already described. Cold water is brought by the pipe *M*, and passes through short inlet pipes into the water-feeder *R* of each jacket. The heated water passes out through the outlet pipes *S* into the gutter *T* (wrongly indicated as *T′* in Figs. 3 and 5). The tuyere-holes *n* are placed at the junction of the water-jackets.

The shaft of the furnace from the water-jackets up to near the top of the feed-holes is lined with fire-brick. The rest of the masonry walls *C′* are made of common brick and rest on a cast-iron carrier-plate, *O*, which in turn rests freely upon iron girders *G′*, three on each side of the furnace, which are firmly screwed and bolted to the capitals of the hollow cast-iron supporting pillars *P*. The plate *O* is in no way fastened to the girders, so that its expansion and contraction are absolutely free. The obvious advantage of this arrangement is to render the masonry absolutely independent of the pillars, so that both keep their relative position unaltered by any lateral motion of the main support. When the main cast-iron support is fixed to the pillars, as is the case in most furnaces, the pillars are not unfrequently cracked by the irresistible dilation of the main support. Upon the cast-iron plate *O*, an outer wall, *N′*, extending up to the charging-floor, is built of common bricks. It is strongly braced by five rows of rails, *Q*, inclosing flat vertical irons *e* and corner irons *c*. It is said that by the use of the outer wall, which naturally protects the hotter parts of the furnace against external radiation, a saving of 15 per cent. of fuel is effected; so that it must be considered as an important part of the furnace. Each side wall of the furnace is provided with two feed-holes, *H*; but, while in most furnaces these holes are placed in the middle of the wall and are directly opposite each other, here they are placed to the left of the middle on either side and thus are not directly opposite. This arrangement is necessitated by the large dimensions of the furnaces and enables the feeders to distribute properly the charge without effort. The feed-holes are provided, as usual, with sliding sheet-iron doors, *S′*, and have cast-iron door-frames. The chimney *D* is braced at *Q*, and provided with corner irons *c*. It is connected with the dust-chambers by means of the large sheet-iron flue *F′*, which fits into a circular brick ring in the furnace wall, where it is held in place by angle-iron rings. The damper *G* has riveted to it on its lower face an angle iron rim, which rests in grooves filled with sand in the top of the furnace wall, thus providing against any escape of fumes when it is closed.

The arrangement for ventilating or carrying off the smoke of the tap-holes and slags differs in these works from that ordinarily adopted. Instead of the hood *W* and chimney *W′* in front of each furnace (as has been by mistake indicated in the elevation, Fig. 2, Plate XXXI), the whole front part of the building in which the furnaces stand is made one big chimney by a partition wall, extending the length of the building, running up from the charging-floor on the line of the front of the furnace, and slanting a little backwards, so as to reach the middle of the ventilator at the ridge of the roof and where the top of the furnace projects above it. Excellent ventilation is accomplished by this simple method.

Barring-down.—The barring-down of these furnaces is effected in the following manner: To the upper end of the long, chisel-pointed bar (Fig. 7, Plate XLIV) is fastened a long and strong rope; the bar is then introduced through one of the feed-holes into the furnace, and the chisel point is forced by blows of a sledge-hammer between wall and accretions. The rope is then thrown across the furnace to the other feed-hole, and five or six men pull at the rope; voluminous masses of accretions are often detached in this manner and are left in the furnace, where they are subsequently fluxed down without serious stoppings and blowings-out. Even the dreaded sows float on the lead bath, remaining constantly at the same level, exposed to the oxidizing influence of the blast and to the sulphurizing influence of mattes and unreduced sulphurets; they thus rather assist than hinder the completion and perfection of smelting, and very soon disappear, being fluxed down in this way.

Sampling of bullion.—The chisel used to detach pieces of bullion from each bar at smelter C is represented in Fig. 14, Plate XLIV. It is a hollow conical chisel or punch, provided with two apertures, a and b. Samples of lead or assay bits (Fig. 15, Plate XLIV) are obtained from the top and bottom of each bar by hammering the chisel perpendicularly to the bar. The punch being driven by the hammer into the surface of the bar, the cylinder of lead is forced in at b and out through a, being detached by striking the butt-end of the punch.

Slags.—A uniform treatment of slags has been adopted at smelter C, which presents two advantages. The slag is left in the pot until a solid crust about two inches thick is formed on the sides and surface of the mass. The upper crust is pierced by two holes, the slag-pot is reversed on the slope of the slag-heap and its molten contents poured out. The thin shells of slag thus obtained are broken up and kept for smelting. It will be seen in the assays of slag that the side shell is a little richer in silver than the portion of the slags poured out; but that the upper crust is poorer in silver than either the poured out portion or the side shell, so that this portion of the slag might be thrown away as useless.

Bars of bullion.—At smelter C the bars of bullion weigh about ninety-eight pounds. They have the shape and dimensions indicated in Figs. 3 and 4, Plate XLV.

Length of run.—The average length of run of the admirable furnaces above described is 12 months, and runs of even 13 months have been obtained. The remarkable length of these campaigns is due not only to the mechanical perfection of the apparatus, but also to the great care bestowed upon every detail of the smelting operations.

Dust-chambers.—The dust-chambers are built of small blocks of limestone (Leadville dolomite), cemented by a mortar of sand and lime. They were constructed at a time when limestone was less expensive than bricks. They form a parallelopipedic construction 75 feet long, 15 feet high, and 25 feet wide, with walls about one foot thick, and are placed immediately below the feeding-floor at the rear of the furnace. In Plate XXX are seen the vertical sections (Fig. 1) and the horizontal sections (Fig. 2) of these chambers. Figs. b, c, d, e, f, g represent elevations of the partition walls W′, showing the disposition of the apertures a through which the fumes circulate. At the time this report was made the chambers were connected with only two furnaces, although they are constructed to condense the fumes from three; this explains the unequal allowance of condensing space provided for each furnace. The first furnace is connected with a chamber divided into two sections, A and B, by a wall, W', provided with an aperture, a (see also Fig. g); the fumes enter the chambers at F' and reach the sheet-iron stack through C''. The second furnace is connected by the flue F'' with the chamber, divided into sections C, D, E, G. The fumes circulate alternately up and down and from right to left, until they reach the flue C', which takes them to the stack F.

The whole arrangement is far from perfect, but the fumes are made to strike walls, and this seems to be one of the conditions essential to deprive them of their dust.

The lighter portions of the fumes are carried away into the air and fall back occasionally on the roof of the building, which is covered with an impalpable yellowish-white dust. In each section of the chambers is a door for the extraction of the dust, which is moistened with water before being wheeled away.

Treatment of flue-dust.—From four to five tons of flue-dust are collected weekly in the chambers just described. The dust is mixed with milk of lime and molded into bricks in the molds represented in Figs. 9 and 10, Plate XLIV, which are also used for common bricks. The bricks thus obtained are then dried under a shed and afterwards on driers; they are then laid on the ore-beds and resmelted.

One of these bricks was examined to determine the actual quantity of lime with which they are mixed. Their contents in lime and magnesia are: Lime, 6.9 per cent.; magnesia, 5.1 per cent., or 12 per cent. in all; but the original dust contained already 4 per cent. of a mixture of lime and magnesia, leaving 8 per cent. for the lime and magnesia thus introduced.

Treatment of mattes and accretions.—Mattes and accretions are placed in heaps in alternate layers with wood, and thus slowly roasted by slow combustion; but it will be seen in the study of mattes and accretions that this is a very imperfect mode of treatment, by which a great deal of silver is lost. Speiss is kept separate from other products at smelter C, but is not treated. It will be seen in the analytical study on speiss that this, as well as all the other speiss at the camp, contains a small quantity of molybdenum, which is entirely concentrated there.

Steam-power.—The boiler is worked at a pressure of 60 pounds to the square inch, and the engine is of 50 horse-power. The machinery driven by this engine consists of two No. 5½ Baker blowers, three Blake crushers, and the pumps feeding the water tanks.

The blast arrangement adopted at this smelter is the one chosen for the general description at the commencement of this section, to which the reader is referred. The normal pressure used is eight-eighths to nine-eighths inch of mercury. It will be seen in the discussion of the blast furnace that the weight of atmospheric air needed to work in the best conditions is about four-fifths the weight of the smelting charges; it will also be seen that the volume of blast in Leadville is much greater than at lower altitudes.

Smelter C, with two furnaces, smelts from 80 to 100 tons of ore per 24 hours.

SMELTER D.

Smelter D is a neat and compact little smelter situated on the northern bank of California gulch, and so like the similarly situated smelters just described, in its general arrangement, that its description will not be given in detail. The pressure of steam in the boilers is 65 pounds to the square inch; they supply a 40 horse-power engine, which drives two No. 5 Baker blowers, 1 Blake crusher, one set of Cornish rolls, and the pumps feeding the water-tank, which supplies the water-jackets of two furnaces. The diameter of the pipe supplying both is 2½ inches. The pressure of blast used at this smelter is the lowest in the camp and averages from four-eighths to six-eighths inch of mercury.

Furnaces.—The two furnaces used at this smelter, and which are equal in dimensions and capacity, are represented in elevation (Fig. 1) and in vertical section (Fig. 2,

Plate XXXII), and are connected with the dust-chamber represented in Fig. 4. A longitudinal or side elevation of the furnace, shown in front elevation in Fig. 1, is seen in Fig. 3. The hearth *A* is very similar to that of the furnaces at Smelter C, already described; it is lined with fire-brick, and the siphon-tap *L* is confined within the hearth-plates; but here the bullion is ladled out direct from the siphon-tap. The hearth-plates are braced by one row of braces, *Q'*. The hearth is also confined within the hearth-plates, as was the case at smelter C, and does not project out, as in other furnaces. The water-jackets *B* are made of riveted steel boiler-plates and are braced by tie-rods *Q*. There are only four jackets: one in front, one at back, and one large one on each side. The circulation of water in these jackets is similar to the one adopted with cast-iron jackets; the water is introduced by means of the pipe *M*, and the hot water comes out at the open outlets *R*, provided with outlet-pipes *S*. The back water-jacket is provided with two tuyere-holes, which are not used, and the side jackets with four holes, in each of which a tuyere is placed; so that each furnace is worked with eight tuyeres.

The pillars *P* do not rest on the ground, as is the case with all the furnaces thus far described, but on the lining of the crucible. Another peculiarity is that there are six of these pillars, instead of four as in most furnaces. The capitals are supported on the pillars by means of brackets *t*. The main cast-iron plate support is of unusual thickness, being four inches thick. The use of a plate of such unusual dimensions is necessitated by the fact that the masonry does not rest directly on the pillars, as in other furnaces.

The masonry consists of fire-bricks, as usual, but is entirely surrounded by a wrought iron jacket, *J'*. At the throat there are two feed-holes *H*, provided with sliding doors *S'*.

Dust-chamber.—The sheet iron stack of each furnace, which is a prolongation of the jacket, is connected by means of the flues *F' F''* with the sheet-iron dust-chamber, formed of a cylindrical portion *D'* and a conical portion *D''*. The fumes escape through the sheet-iron stack. The dust is withdrawn from this chamber by means of sliding valve *S*, and falls from the aperture *Z* into a wheelbarrow, *Y*. At smelter D flue-dust is not mixed with lime, as at most smelters, nor spread over ore-beds or mixed with smelting charges; it is simply moistened with water and thrown in the furnace in the proportion of one shovelful to every two smelting charges. The smelting capacity of each furnace is 24 tons of ore per 24 hours, or one ton per hour. The pressure in blast-pipe *R'* is regulated by a damper placed at its extremity. The length of runs at this smelter is about two months.

SMELTER E.

Disposition of works.— These works, situated on Big Evans gulch, are, like all the smelters erected on this gulch, divided into two levels only. This smelter is small, but well managed, and is one of the most successful of its size. The pressure of the steam in the boilers is 70 pounds to the square inch; they supply a 40 horse-power engine, which drives two Baker blowers, a set of Cornish rolls, two Blake crushers, and the pumps feeding the water-tanks. The ore and fuel bins are inclosed in the main building, through which runs a wagon-road, and fuel reserves are placed at the back, outside of the works. The offices, provided with Fairbanks scales, and the laboratory are situated in a detached building a short distance back of the main building.

Furnaces.—The smelting capacity of these works is 40 to 50 tons of ore per 24 hours. Smelting is carried on in two blast furnaces of unequal capacity, constructed by Messrs. Fraser & Chalmers, of Chicago. The small furnace is circular and jacketed all over. Its smelting capacity is 15 tons of ore per 24 hours. The crucible is lined with steep, made of two parts fire-clay and one part coke dust. The distance between tuyere and feed-hole is 10 feet 9 inches. The depth of the crucible is 28 inches. The diameter of the riveted wrought-iron-plate water-jacket is 36 inches. This jacket is single. The furnace is worked with five tuyeres, each 2¼ inches in diameter at the nozzle. The average pressure of blast used at smelter E is one inch of mercury. The large furnace presents the same general appearance as the small one, and is jacketed all over; but it is elliptical in section, the axes at the feed-hole being 69 inches and 54 inches, respectively, and the axes of the riveted wrought-iron-plate water-jackets 52 inches and 34 inches. The water-jacket is made in four sections. The furnace is worked with seven tuyeres 2½ inches' diameter at the nozzle. The depth of the crucible is the same as that of the small furnace, namely, 28 inches and it is lined with the same kind of steep. The furnaces have projecting fore hearths and lead siphon-taps, and are constructed on identically the same principles as the circular furnaces at smelter B, made by the same firm. The furnaces are barred out every four days, and the length of run is about three months.

Dust-chamber.—Both furnaces are connected by means of sheet-iron flues with a brick chamber, flue-shaped and placed under the feeding-floor. This chamber is 75 feet long, 4 feet wide, and 6 feet high. It is provided, as usual, with a sheet-iron stack about 30 feet high, placed at the extreme end. The chamber is not divided into sections by internal walls, so that the condensation of fumes is rather imperfect; doors placed at short intervals allow the clearing away of the dust.

At smelter E flue-dust is mixed with lime in the proportion of one ton of dust for 300 pounds of lime (dolomitic), or about 15 per cent. The mixture, after moistening, is molded into bricks, which are then dried in the air and laid over ore-beds to be resmelted.

SMELTER F.

Disposition of works.—Smelter F is the model smelter of Big Evans gulch. The smelting building is capacious and well distributed, and everything in the construction of details of plant indicates extensive previous experience in smelting and no small amount of forethought. Like all the smelters in Big Evans gulch, smelter F is divided into two levels. The boiler, engine, and blast rooms are placed on the left of the furnace room (facing the furnaces) and on a level with them. There are two fine boilers, worked at a pressure of 65 pounds to the square inch, and the engine is of 50 horse-power. The machinery driven by this engine consists of two Baker blowers, two Blake crushers, a small grinding-mill, and the pumps feeding the water-tank. The system of blast-pipes is identical with the one used at smelter C. It is provided with safety-valves and dampers, and the excess of blast is ejected by means of a damper placed at the extreme end of the main pipe. Several rows of ore-bins and fuel-bins are placed on the feeding floor of the main building, and reserves of fuel stand at the back and outside the works. A broad wagon-road runs through the entire length of the works. The offices occupy a detached building placed near the entrance of the works. The Fairbanks scales are placed immediately at the entrance of the main building and are connected with a small office. The laboratory occupies a detached construction distinct from the offices.

Furnaces.—There are two blast furnaces, of equal shape, dimensions, and capacity. The capacity of each furnace is 30 tons of ore per 24 hours. Fig 1, Plate XXXIII, represents the front elevation of one of these furnaces. They are square (8 by 7 feet outside measurement at feeding door, and 5 by 5 feet inside measurement of crucible), and their masonry is entirely made of bricks, braced at *Q*. They are provided with two feed-holes *H*, opened or closed by means of sliding doors *S'*. The masonry rests on a main cast-iron plate, *O*, supported on four cast-iron-pillars, *P*. The space *b* between the masonry and the water-jackets is filled with fire-brick.

The water-jackets, which are entirely made of cast iron, are similarly disposed in every respect to those of the same kind previously described. They consist of one jacket in front, one at the back, and two on each side. They are provided with feeders, outlet-pipes, and supply-pipes. The fire-brick breast *V*, placed between the hearth and front jacket, is seen in this furnace, and corresponds to a similar arrangement in all square furnaces in which the water-jackets are entirely made of cast iron. The water-jackets are provided with seven tuyere-holes, three on each side and one at the back, and the furnace is worked with seven tuyeres.

Patent tuyeres.—Fig. 1, Plate XXXIII, was specially drawn to show the system of tuyeres at this smelter, which differ in every respect from the thin sheet-iron galvanized tuyeres in general use in the camp. The tuyeres were patented December 6, 1875, by Mr. August Werner. They are made of cast iron, three fourths of an inch thick, and their internal diameter is 2½ inches. They are divided into two parts, the nozzle *N*, and the elbow *N'*. Both the nozzle and the elbow are flanged at *r*, the flanges being faced so as to fit closely and allow no escape of blast. The nozzle and elbow are hinged at *d*, and to the nozzle are fixed three small chains, *c*, hooked to the water-jackets. By means of these the direction of the tuyere can be changed at will so as to send the blast up and down or right and left. At this end the nozzle terminates in a wrought-iron spherical ring or ball, which works freely in a socket of the same metal, wedged in the tuyere-hole of the water-jacket. In other words, the tuyere works in a ball-and-socket joint. To stop the blast in any point of the furnace or to observe what is going on there, the elbow is lifted, as indicated in Fig. 1. The tuyeres are connected, as usual, with the blast-pipes by means of canvas wind-bags *K*. When the blast is turned off for the purpose of barring down the accretions of the furnace or clearing the hearth of accretions, a piece of paper is inserted between the flanges *r*, and should back flow of gases exert any pressure in the furnace the piece of paper would burst, the elbow of the tuyere be lifted, and the tuyere would thus act as a safety-valve. But this accident, so far as known, has never occurred in Leadville.

The normal pressure of blast used at these works is seven-eighths of an inch mercury. The crucible of the furnace is provided with a projecting fore-hearth and lead siphon-tap and is lined with fire-brick.

Dust-chambers.—The apparatus devised for the condensation of lead fumes at smelter F is the most elaborate of its kind used in Leadville, and is certainly the most efficient. Each furnace is connected with a separate condenser, placed above the feeding floor, and is identical with the one shown in side and front elevation, Plate XXXIV.

The chimney of the furnace *A* is connected by means of the angular sheet-iron flue *F′ F″*, which projects above the roof *V* of the building, with the lozenge-shaped sheet-iron chamber *M*. The fumes strike against the sheet-iron apron *I*, hinged to the upper part of the chamber, which may be also used to regulate the draft by means of a chain which passes through the wall of the chamber. After leaving the chamber *M*, Fig. 1, the fumes circulate through the sheet-iron flue *O*, and then escape through the sheet-iron stack *F*.

In close proximity to the wall of the furnace, the flue *F′* is provided with a sheet-iron branch, *C*, through which the flue-dust falls into the wooden box *B*, from which it is extracted at the door *d*. Flues *F′*, *F″* are provided with sliding doors, not seen in the sketch, for clearing them of their dust. The chamber *M* has also a large sliding door, *D*. This chamber, as well as the horizontal flue *O* and the stack *F*, are cleared of their dust through the branches *C*, provided with sliding valves *t*. The principle of ascending flues is in itself excellent, and the smoke which comes out at the stack is remarkably free from lead fumes. About ten tons of flue-dust is collected monthly in each of these chambers. The dust is mixed with milk of lime, the mixture spread over ore-beds, and then resmelted.

Smelter F is the only one in Leadville where a little metallic iron (old horseshoes) is added to the usual smelting charges when they contain more than a certain percentage of galena. The smelting campaigns have an average length of nine weeks. The furnaces are provided, as usual, with large hoods in front, above the slag-gutter.

SMELTER G.

Disposition of works.—This important smelter is situated on the northern bank of California gulch, and, like all the smelters situated on this gulch, is divided into several levels communicating with the upper and lower roads. One of the main features at these works is that the fuel storage, which is placed at the back of the works, is nearly on a level with the upper part of the stacks of the furnace and is connected with the furnaces by an elevated trestle-work having two branches, the one leading to the furnaces the other to the boiler-room. The fuel is transported in light sheet-iron mine cars, running on a light iron tramway, and dumped into chutes adjoining the feed-holes of the furnaces and the boiler. In the boiler-room a saving of 50 per cent. of the wood burned is effected by using the screenings of the fuel, which are usually wasted. A great saving of labor also results, since two fuel men are sufficient to supply all the fuel needed. In principle this arrangement is similar to the one adopted at smelter A. The charcoal sheds have an area of 30 by 325 feet and 35 by 100 feet, respectively, and hold about two hundred thousand bushels of charcoal. Coke is stored in sheds and bins of 500 tons capacity. The main smelting building is 360 by 110 feet, and the ore-room, placed on one side of the main building, is 60 by 210 feet. The storage capacity of this room, through which a wagon-road runs, is 7,000 tons of ore. The large dimensions of this room allow the preparation of numerous ore-beds, which insures great regularity in smelting. The offices, laboratory, Fairbanks scales, staff-houses, and 22 dwelling-houses for the workmen and their families, are distributed around the works. Particular attention is paid to the welfare of workmen, who are entitled to free medical attendance at the hospital, and for whom a bath-room and a reading and recreation room have been constructed.

Two boilers, 40 inches by 16 feet, worked at a pressure of 60 pounds to the square inch, supply a 70-horse power engine (cylinder 14 by 24 inches), and a second engine of 50 horse-power, with its boilers, is kept ready for use in case of need. This engine was the one formerly used at these works before they had attained their present smelting capacity.

The 70 horse-power engine drives three Baker blowers, one Root blower, two large Blake crushers, a set of Cornish rolls, and a slag-hoisting machine. The furnace-room is 120 by 40 feet, and contains four furnaces, smelting about one hundred and twenty tons of ore in twenty-four hours. The ventilation of this room will be shown in the description of the dust-chambers. The slag-heap is connected with the feeding-floor by an inclined-plane hoisting-machine, similar to the one used at smelter B, and used also to carry the slag up to be resmelted.

Furnaces. — Smelter G has three furnaces of equal shape and dimensions, similar to the one shown in front elevation (Fig. 1, Plate XXXV), and one larger furnace, shown in Fig. 2, Plate XXXV. Although built on the same general principles as the other furnaces of the camp, they offer a few interesting peculiarities in construction. The small furnaces are square (3 by 4 feet at the tuyeres), and their cast-iron pillars rest on the fire-brick lining of the crucible. The water-jackets *B* are made of riveted boiler-plates and are only four in number. Each side jacket is provided with two tuyere-holes and the back jacket with one; but the furnace is worked with the four side tuyeres only.

The main cast-iron plate support has a broad vertical flange, *O*, which confines the base of the outer walls *C′* of the furnace shaft, the shaft itself being, as usual, lined with fire-bricks; the outer wall is made of red brick, braced at *Q*.

These furnaces are fed through a single feed-hole placed at the back of each, and provided with sheet-iron sliding doors. The whole portion of the furnace comprised between the feeding-floor and the damper of the stack is surrounded by a sheet-iron jacket, J′.

The crucible of the furnace is framed in strong cast-iron plates, and the frame of the siphon-tap, lined with steep, is made of strong sheet iron. The smelting capacity of each of these furnaces is 26 to 28 tons of ore per twenty-four hours, and the length of runs is about 118 days.

The large furnace represented in Fig. 2 is the only one of its kind used in Leadville. The lead siphon tap *L* is placed in front of the furnace, and on each side of the furnace there are a fore-hearth, *X′*, and a slag-spout, *U*, alternately used for the tapping of slag. In *B′* are seen the slag-pots, mounted on wheels.

The water-jacket system is formed of four large water-jackets made of riveted boiler-plates. The front and back jackets are each provided with four tuyere-holes, but the furnace is worked with only six tuyeres. The dimensions at the tuyeres are 3 by 5 feet. The main cast-iron plate support has a broad, vertical flange, *O*, incasing the base of the masonry.

The furnace is fed from two feed-holes, *H*, opened or closed by sheet-iron sliding doors. The feed-holes are placed in the side walls of this furnace, which correspond to the front and back walls of other furnaces. The pressure of blast used at smelter G varies from five-eighths of an inch to ten eighths of an inch of mercury. The capacity of the large Raschette furnace, which has just been described, is 38 to 40 tons of ore per twenty-four hours. The manipulations of either furnace do not differ from those in use at other smelters. The shovels used are represented in Figs. 2 and 3, Plate XLIV, and the bars of bullion in Figs. 5 and 6, Plate XLV.

Dust-chambers and ventilation.—In Fig. 1, Plate XXXVI, is shown the general system of condensation of lead fumes and of ventilation of the furnace-room. The ventilators V' V'' V''' consist of large rectangular sheet-iron chimneys resting on the brick dust-chambers D' D'' D'''. They are open at their base on the side towards the furnaces to allow hot air to escape through them. Besides the ventilators each furnace is provided with a hood and chimney in front of the furnace towards the slag-gutters.

Furnace A is connected by means of sheet-iron flue F' with chamber D', divided into three sections, a, b, c, by means of partition walls w, and the smoke escapes through the sheet-iron stack S.

Furnace B is similarly connected through flue F'' with chamber D'', divided into three sections, a', b', c', by means of walls w', the smoke escaping through sheet-iron stack S.''

Furnace C is connected by means of sheet-iron flue F''' with a brick chamber, d', 8 feet high and 11 by 11 feet at base, resting on the feeding-floor P' P'. This chamber has an independent sheet-iron stack, S'.

Furnace D communicates by means of sheet-iron flue F''' with chamber D''', divided into two sections a'', b'', and the fumes circulate through the brick flue C'', 25 feet long, and then ascend the square brick stack S'''. Each section of the dust-chambers is provided with sliding doors d, for the extraction of the dust; those of sections e' and a'' of chambers D'' and D''' are in the arch-way O.

At smelter G flue-dust is mixed with argillaceous ores and resmelted.

Smelting charges.—The following figures show the smelting performed by the three small furnaces at smelter G from the 14th of June, 1879, to the 1st of January, 1880:

Ore smelted,	24,094,177 pounds	= 12,047 tons, assaying 73½ ounces silver; lead, 22 per cent.	
Dolomite,	1,521,085 pounds	= 6.31 per cent. of ore smelted.	
Hematite,	2,872,535 pounds	= 11.92 per cent. of ore smelted.	
Coke,	1,988,110 pounds	= 8.25 per cent. of ore smelted	Fuel = 24.5 per cent. of ore smelted.
Charcoal,	3,916,287 pounds	= 16.25 per cent. of ore smelted	

The contents were: Silver, 885,454 ounces; lead, 5,300,719 pounds. The products amounted to: Silver, 866,666 ounces; lead, 4,469,823 pounds.

The loss in silver was 2.12 per cent. and in lead 15.67 per cent.

The average price paid per ton of ore was $66.15.

The bullion produced each day, 12½ tons.

SMELTER H.

Disposition of works.—Smelter H is the most important smelter of Big Evans gulch. These works are in close proximity to the important mines of Fryer Hill, and the ores they receive comprise some of the richest in lead, gold, and silver. The bullion extracted there is also generally very rich in silver. The works are provided with a laboratory, in which the ores are assayed, chiefly by scorification, but some mines require also the crucible assay. The crucible assays of ore for silver and of ores and slags for lead are made only with reducing flux, and no iron rods are used for the reduction of sulphurets and arseniurets. The assays for iron and gangue are made as usual, although the solutions of ores are not evaporated to recover soluble silica and the estimation of moisture is made in a very rough way.

The specific gravity of slag is taken from day to day for each furnace by means of the Jolly specific-gravity balance, already described, which is figured in Plate XXXVIII. This operation is of no more advantage at smelter H than at smelter B, and here, as there, the slags thrown away are the richest of the camp, both in lead and silver.

The ore-beds are made to contain equal parts of iron and gangue, and the slags thus formed are slightly acid. This plan, which is recommended in Leadville, and which is gaining the confidence of smelters, should be condemned theoretically, and practice proves that theory is correct. It results, from the examination of slags made in the laboratory of the Survey, that the so-called acid slags are richer in lead and silver than the more basic ones. But the chief defect of this plan is that an insufficient quantity of iron is reduced, and that very large quantities of sulphuret accretions and unreduced galena are formed, interfering seriously with the working of the furnace. In the opinion of the writer the center of gravity of smelting operations, so to speak, should be periodically displaced, and alternate acid and basic charges should be used, for the inconveniences inherent in the use of these mixtures are precisely of an opposite character and calculated to counterbalance or destroy each other. At smelters C and G, where smelting is conducted on scientific principles, the mixtures are carefully made to correspond to singulo-silicate slags, which might be called neutral, so that the final result is the same as the one proposed.

The quantity of matte formed at smelter H is about 20 pounds per ton of ore, or 1 per cent. These mattes are roasted in heaps and resmelted; but it will be seen that this mode of treatment is bad, and that much silver is lost during the roasting.

The method of bullion assay, which is the one in general use in Leadville, is as follows:

At smelter H two assay bits of lead (one from the top and one from the bottom of each bar) are detached from each bar composing a car-load by the chisel represented in Fig. 12, Plate XLIV. By hammering in different directions triangular bits of lead are detached, such as are represented in Fig. 13. All the bits representing the car-load are melted together under live charcoal in a plumbago pot. The charcoal is then removed and the lead is skimmed by means of a small perforated ladle, and then poured into a bar-mold. A bar about one inch thick is thus obtained (see Fig. 7, Plate XLV), from which three or four assay bits are detached at *a* by means of an ordinary chisel and hammer. Half an assay ton is weighed from each bit, and the assay is made, as usual, by cupellation.

The offices, laboratory, and Fairbanks scales occupy a detached building at the entrance and rear of the works, and fuel is placed on the same level in the open space at the back of the main smelting building. The ore-bins are all placed within the building on the feeding-floor level, through which runs a wagon-road for the distribution of ore from wagons.

On the left of the furnace-room (facing the furnace) is the engine and blast room. Two boilers, worked at a pressure of 80 pounds to the square inch, supply a powerful engine of 100 horse-power, which drives four Baker blowers, one Blake crusher, one set of Cornish rolls, and the pump. The works are also provided with a small mechanics' shop. The slag-heap at this as well as at all the smelters on Big Evans gulch encroaches on the bed of the creek. The smelting capacity of the works is about sixty-five tons of ore per twenty-four hours.

Furnaces.—At smelter H there are three furnaces of the Piltz pattern, constructed by Messrs. Fraser & Chalmers, of Chicago. These furnaces, which have already been described in the general description of the furnaces, and which are also successfully at work at smelter B, are represented in perspective view, Fig. 1, Plate XXXVII. This sketch was drawn for the purpose of giving a correct idea of the general appearance of these furnaces, which cannot be obtained at a glance from the elevation and section alone. In this sketch the crucible *A*, with its frame of cast-iron plates, as well as the frame of the lead siphon-tap *L* and of the fore-hearth *X′*, is clearly seen. The cast-iron pillars *P*, with their capitals and brackets and the two slag-gutters *U*, are visible. Likewise the riveted wrought-iron boiler-plate water-jackets *B*, the fire-brick breast *V*, and the tymp-stone and tap-hole *Z*. The main cast-iron support *O*, with its vertical flange *O′*, supported by the brackets *r*, the induction-pipe *I*, and the wrought-iron casing *J′* around the masonry, are also visible.

The same furnaces are represented in vertical section in Fig. 2, Plate XXXVII, showing the steep-lining of the hearth and fore-hearth *X′*, the siphon *L′*, the space *b*, between the water-jackets and the masonry, filled with fire-brick, and the fire-brick lining *C′* of the furnace.

Fig. 2 shows also the arrangement adopted at smelter H for the tapping of slag. The slag runs into a cast-iron slag-pot, *V′*, provided with a spout, *U′*, and live charcoal in large pieces is kept over the molten slag to prevent it from cooling. Any bullion mechanically carried away falls at the bottom of the pot *V′*, which is cleared of its contents from time to time. The slag thus freed from bullion runs into the ordinary slag-pot *B′*, mounted on wheels.

This arrangement is evidently excellent, but is only necessitated by some defect in the lining of the dam, for in well-lined furnaces no bullion can escape, thus rendering the use of an intermediate slag-pot unnecessary; this is proved by the fact that slags never contain any metallic grains, no matter from what part of the cake the specimen is taken.

Fig. 2 shows also the connection, by means of the sheet-iron flue *F′* of the chimney *E*, of the furnace with the sheet-iron chamber *D′*, resting on the feeding-floor *P′*, used to catch lead-dust. At *d′* is seen one of the doors of this chamber, through which the dust is extracted. The small furnaces which have just been described are worked with six tuyeres, 2½ inches at the nozzle, and their smelting capacity is 16 to 18 tons of ore per twenty-four hours for each furnace.

Besides the three Piltz furnaces, smelter H has a large Raschette furnace, which was formerly run at smelter L. The smelting capacity of this furnace is 25 tons of ore in twenty-four hours. The internal dimensions of the crucible are 5 by 3 feet. The hearth is lined with steep and the furnace is supported on four cast-iron pillars, like the square furnaces of similar construction already described. The water-jacket system is rather complicated and is formed of one front and one back jacket, made of wrought-iron riveted boiler-plates, with five cast-iron water-jackets on each side. This plan, as has been observed before, is not good, and at the time this report was made the water-jacket system was under repair. The difference of dilation of the two metals is always a source of trouble and the plan should be condemned altogether. The furnace is worked with nine tuyeres, 3 inches in diameter at the nozzle, inserted, as usual, in the water-jackets; one of the tuyeres is placed in the back jacket and four on each side.

Each jacket is not only provided with inlet and outlet pipes for the circulation of water, but a general circulation has been established between all the jackets by means of pipes screwed into them at the base and communicating with one another.

Dust-chambers. —The system of condensation of lead fumes adopted at this smelter is poor, and "leading" is consequently of frequent occurrence. The four furnaces are connected by means of the sheet-iron flues F^{i} F^{ii} F^{iii} F^{iv} with the sheet-iron chamber M, connected by means of the brick flue N with the stack F (in Fig. 1, Plate XXXVIII). The sheet-iron chamber has already been seen in transverse elevation (Fig. 2, Plate XXXVII). Neither chamber M nor N is divided into sections, so that the condensation of fumes is very imperfect. Both chambers are provided, as usual, with sliding doors d for the extraction of the dust.

At smelter H, flue-dust is mixed with Hibernia ore (an argillaceous ore containing no lead) and introduced afterwards into the composition of ore-beds.

SMELTER I.

Smelter I is erected on the northern bank of California gulch, in a situation so similar to that of smelter C that the general description of the latter applies word for word to these works. The only peculiarity at Smelter I is that the furnace and feeding-floor levels are connected by a vertical elevator used for hoisting slags to be resmelted. This elevator is placed in the main building. The boilers are worked at a pressure of 70 pounds to the square inch. The machinery consists of a 60 horse-power engine, two Baker blowers, one Blake crusher, and the pump.

The slag-pots are independent of the cars and are identical with those which have been described at smelter A. The smelting plant consists of two Piltz furnaces, identical in capacity, shape, and dimensions, and constructed by Messrs. Fraser & Chalmers. These furnaces are similar to furnaces of the same pattern used at smelters B and H, but have only one slag-spout. The water-jackets B also are made in but two sections, and the frame of the crucible of four cast-iron plates, segments of a circle. One of these furnaces is shown in elevation in Fig. 2, Plate XLV. It may be seen that each Baker blower, W, is in direct communication with the induction-pipe I, the general system of connecting all the blowers with a main blast-pipe not being in use here.

The system of condensation of lead fumes consists of a sheet-iron box, D', 8 by 8 feet, and 10 feet high, provided with a sheet-iron stack, F. Each furnace is connected by means of a sheet-iron flue, F', with a similar chamber, from which the dust is extracted through hinged doors d'. The amount of flue-dust caught in both chambers is about 5 tons per week. The dust is mixed with milk of lime, the mixture is dried and then resmelted gradually with the smelting charge.

The smelting capacity of the works is about 40 tons per twenty-four hours.

SMELTER J.

Smelter J is a well-constructed smelter standing on the southwestern bank of Big Evans gulch, and is disposed exactly like smelter H, with this difference, that the offices and laboratory stand on one side of the main smelting building instead of being placed in the rear of the works. The works had ceased running at the time this report was made, but they deserve a description chiefly on account of the well-constructed brick dust-chamber with which the furnaces are connected.

The smelting plant consists of one Piltz furnace, worked with six tuyeres, and constructed by Messrs. Fraser & Chalmers. The diameter at the base of the water-jacket is 40 inches, and the furnace is similar in every respect to the Piltz furnaces used at other smelters. Besides the Piltz furnace there is a Raschette furnace, charged through two feed-holes, and almost identical in proportion and capacity with the Raschette furnace used at smelter H. The water-jackets, all made of cast iron, are thirteen in number: one in front, two at the back, and five on each side. The internal dimensions at the tuyeres are 5 by 3 feet and the furnace is worked with nine tuyeres.

Both furnaces are connected with a dust-chamber placed immediately below the feeding-floor. This chamber is built entirely of red brick.

The plan of this chamber is given in Fig. 2, Plate XL, showing the brick chamber *D′*, the brick flues *N N′*, communicating with the sheet-iron stack *F*. In this plan *A* represents the Piltz furnace and *B* the Raschette furnace. The same chamber is represented in elevation in Fig. 1. The two sliding doors through which dust is cleared away are placed at *d*. The capacity of the works is 40 tons per twenty-four hours. The machinery consists of a 50 horse-power engine, two Baker blowers, one Blake crusher, and one set of Cornish rolls. The charges are weighed, as at all the other smelters, on scales placed on the feeding floor. The slag-pots used are mounted on wheels.

SMELTER K.

Smelter K is the smallest smelter of Big Evans gulch, and has only one furnace, which was not running at the time this report was made. These works are, on a miniature scale, disposed exactly like smelters H and J, and they have this point in common with smelter H, that intermediate slag-pots are used for catching any bullion mechanically carried away. The furnace is a Piltz pattern furnace 40 inches in diameter at the base of the water-jackets, and worked with six tuyeres, 2¼ inches in diameter at the nozzle. The capacity of the works is 18 to 20 tons per twenty-four hours. This smelter affords an opportunity for showing the plant and manual labor required to work one furnace. At these works the manual labor was represented by—

	Pay per diem.	Length of shift.
		Hours.
1 foreman	$5 00	12
2 head smelters	4 00	12
2 feeders	3 00	12
2 helpers	3 00	12
2 engineers	3 00	12
8 day laborers	2 50	10
4 staff officers	....	....

The plant, supplied by Messrs. Fraser & Chalmers, of Chicago, consists of—

One Piltz furnace, 40 inches diameter at the water-jackets.

One tubular steam-boiler, 48 inches in diameter and 14 feet long. (This boiler is sufficient for two furnaces.)

One stationary steam engine (cylinder 12 by 18 inches).
One No. 5 Baker blower, with mercury gauge.
One complete set of blast and induction pipes, with hose and tuyeres
One Blake crusher (opening between the jaws, 10 by 7 inches).
One set of Cornish rolls, 16 by 10 inches.
Eight slag-pots, mounted on wheels or on independent cars.
Six lead-ladles.
Eighteen lead-molds, with name of smelting firm at bottom, for branding bullion.
Two No. 4½ sheet-iron mining-barrows for fuel.
Thirty-five steel furnace-bars, from ¾ inch to 1½ inches.

To this must be added either a sheet-iron dust-chamber, 8 by 8 by 10 feet, constructed by Messrs. Fraser & Chalmers, or else a convenient brick chamber placed under the feeding floor; a 10-ton Fairbanks platform scale, and several ore-barrows, shovels, etc., and a water-tank for feeding the water-jackets of the furnace.

SMELTER L.

Smelter L is the Little Chief smelter which stood on Fryer Hill, above the Little Chief mine, and which has been pulled down, owing to the sinking of the ground upon which it was erected. The Little Chief smelter ran only on Little Chief ore, and its ore-room was connected with the shaft of the mine by a railroad track, upon which the loaded mine-cars were run.

The capacity of smelter L in tons of ore per twenty-four hours was 35 tons, which were smelted with dolomite and hematite in the Raschette furnace now at work at smelter H. More sows were found at this smelter than at any other, showing that the slags were basic.

During the year ending June 1, 1880, 5,500 tons of ore were smelted, producing 760 tons of bullion. The plant consisted of a 40 horse-power engine (cylinder, 16 by 24 inches) and a boiler, 48 inches by 14 feet, constructed by Messrs. Fraser & Chalmers; one No. 5 Baker blower; one Blake crusher, 15 by 9 inches between the jaws; and the furnace previously described. The slags at this smelter are remarkable for the absence of titanic acid.

SMELTER M.

These works, the first that were erected in Leadville, were being pulled down at the time this report was made. They were situated immediately outside of the city of Leadville, at the junction of the upper and lower district roads of California gulch. The allotted space for the slag-heap between the works and the lower road was soon filled up, and the slags had to be wheeled up at the back of the works on a level with the feeding-floor. These slags are unlike any other produced in or near Leadville. They are coarse-grained, with a dull fracture, extremely dense, and contain an enormous quantity of lead and silver.

An opportunity was afforded of assaying a few of the ores which were in the bins after this smelter ceased running. Their contents in lead and silver were as follows:

Lead....per cent..	13.9	8.7	0.9	20.8
Silver.....ounces..	34.9	60.25	70.4	45.5

Smelter M had but one square furnace, of the Raschette pattern, entirely sheathed in an iron jacket. The water-jacket system was formed of four wrought-iron riveted-plate jackets, provided with seven tuyere-holes. The lining of the crucible was made of steep. The dimensions at the tuyere were 5 by 2 feet. In Fig. 2, Plate XXXIII, is seen the peculiar dust-chamber which was used at this smelter. The stack *E* of the furnace was connected by means of the sheet-iron flue *F* with the cylindrical sheet-iron chamber *D′*, placed high above the feeding-floor and outside of the main building. This chamber was provided with a stack *F*, and the fumes were compelled to circulate by means of the two sheet-iron cones *u y z*, suspended to the stack *F* by means of a chain *a y*. The chamber was provided with a sliding valve *S*, for the extraction of the dust, which fell through the pipe *z* into a wooden box, placed on a level with the feeding-floor.

The smelting capacity of the works was 30 tons of ore per twenty-four hours.

SMELTER N.

Smelter N, situated in Malta, at the end of California gulch, was the first smelter erected in Lake County, and was built in 1875. This smelter, which was not running at the time this report was made, has been started anew. The works are divided into several levels. The well from which the water was pumped into the water-jackets of the furnace stands in the furnace-room. The furnace, the only one used at these works, has the same shape and capacity as the furnace described at smelter M, and is also entirely sheathed in an iron jacket.

The engine is of 30 horse-power, driving one No. 4 Baker blower, one Blake crusher (10 by 7 feet), and the pumps.

Eighteen men and four officers are in charge of the works.

Smelter N is placed a few yards from the Denver and Rio Grande Railroad track, and is connected with it by a siding.

SMELTER O.

This smelter is also situated in Malta, at the end of California gulch, and a short distance south of smelter N.

Smelter O has two furnaces of the Piltz pattern, oval in shape and similar to the oval furnace used at smelter E. These works have long since ceased running.

SMELTER P.

Smelter P is the Adelaide smelter, which was situated near the mine in Stray Horse gulch, at the north end of Iron Hill. It also ceased running long ago for want of ore. Its furnace has since been purchased and removed, and is now running at one of the smelters in Big Evans gulch.

SECTION IV.

PRODUCTS OF SMELTING.

BULLION.

Sale of bullion.—In Leadville bullion is generally sold to agents of Eastern refineries, who pay for its transportation from the camp to the East. The cost of transportation varies, according to distance, from \$27 to \$35 per ton.

The price of lead in bullion at Leadville has varied during the year ending June 1, 1880, from \$30 to \$78 per ton. The average price has been from \$60 to \$72 per ton.

Sometimes bullion is paid for at New York quotations, with a deduction of 3 cents per ounce of silver and of \$14 to \$15 per ton of bullion for the refining charges. In other cases, the refiners' charges are 3 ounces of silver and 5 per cent., or 100 pounds of lead per ton of bullion.

When the smelting works of Leadville are branch establishments of large eastern refineries, private arrangements are made between the main works and its branch.

When bullion is shipped to refineries to be desilverized for account of the smelting firm, the smelters pay for the transportation of bullion from the camp to the refinery. In this instance the agreement between smelters and refiners is shown in the following model of bullion invoice.

NAME OF SMELTING FIRM.

Leadville, Colorado.

Invoice of bullion, No......
Car lot ...... bars ...... weighing ...... lbs. shipped .... 188—.
Assay per ton (2,000 lbs.)........ ......ozs. silver. New York quotations ....day shipped.
Total ozs. silver, ... less .... ozs. per ton in refining, ... ozs., @ .... per oz., \$ ....
Total lbs. lead, ..., less .... per cent. lost in refining, ... lbs., @ .... per lb., \$ ....
Value of lead and silver ....\$ ....
Deduct freight to ...., \$ ....; cost of refining per ton, \$........ \$ ....
Net value of shipment..\$ ....
Deduct ...., 10% of net value ....\$ ....
Amount for which draft may be made ....\$ ....

When the price of bullion is low it is frequently kept in reserve in Leadville, in the expectation of a rise in the New York price. During the month of August, 1880, one of the smelters presented the imposing sight of reserve piles of 14,625 bars of bullion, amounting to 1,453,250 pounds, or 716½ tons.

The bars of bullion in the camp belong to two principal types, shown in Figs. 3 and 4 and Figs. 5 and 6, Plate XLV. Their average weight is 100 pounds, so that car-loads weighing on an average 20,000 pounds, or 10 tons, are formed of 200 bars. In

Table VII will be found the weekly production of the different smelters, their weekly or monthly shipments, with the weight and average assay of bullion for the dates indicated.

TABLE VII.—*Shipment of bullion.*

Smelter.	Weekly production.		Weekly or monthly shipments.					
	Tons.	For week ending—	Number of bars.	Average weight of bars.	Total weight of bars.	Average tenor in silver.	For week ending—	For month of—
				Pounds.	*Tons.*	*Ozs. to ton.*		
Harrison Reduction Works.......	24	Dec. 18, 1879	6,125	85	260	324	....	July, 1880
Grant	145	Jan. 8, 1880	17,160	100	858	326	....	Do.
Leadville	21	Dec. 4, 1879	420	100	21	205	Dec. 4, 1879	
La Plata	89	Jan. 8, 1880	3,990	100	199½	178	....	Do.
American	15	....do	569	100	28½	258	Dec. 25, 1879	
Billing & Eilers	90	....do	5,300	98	259	136	....	Do.
California	20	....do	1,075	100	53¾	218	....	Do.
Malta	14	Dec. 18, 1879	218	102	11¼	314	Dec. 29, 1879	
Lizzie	22	Jan. 8, 1880	335	103	17¼	338	....do	
Little Chief	25	....do	1,045	100	52¼	370	....do	
Ohio and Missouri	20	....do	2,836	100	141½	443	....	Do.
Cumming & Finn	50	....do	3,600	100	180	314	....	Do.
Gage, Hagaman & Co.............	30	....do	475	100	31	445	Dec. 29, 1879	
Raymond, Sherman & McKay....	23	Dec. 18, 1879	470	100	23½	250	....do	
Elgin	40	Jan. 8, 1880	2,330	100	111½	297	....	Do.
Total and averages	628	....	....	99.20	....	294.4		

Composition of bullion.—The quality of bullion differs a good deal from smelter to smelter, and from day to day at each smelter, but the former difference is more sensible than the latter. At some works bullion is soft, with a clear surface; at others, more or less hard, with a scummy surface. The difference in the quality of bullion is due less to the difference in composition of the ores, which are sensibly the same, than to the care with which smelting is carried on. The same furnaces and the same ores will yield coarse or partly refined bullion, according to the rapidity with which the furnaces run, but chiefly according to the quantity of iron reduced during the operation, this metal being an excellent refining agent.

The charges for refining bullion being greater for coarse than for soft metal, it is quite evident that the smelters have a direct interest in obtaining from their furnaces a metal as refined as possible. The best smelting works of Leadville obtain a bullion of very fair quality.

Analyses.—The writer has made in the laboratory of the Survey the two following analyses of bullion:

Analysis XXII.—Specimen of bullion taken from the furnace at the La Plata smelter. This bullion is soft, with a clean surface.

Analysis XXIII.—Mixture of equal parts of bullion from the following smelters:

Names of smelters.	Remarks.
Billing & Eilers	Sample from one car-load, weighing about 11 tons; soft, with a clean surface.
Cumming & Finn	Sample from one car-load, weighing about 10 tons; somewhat hard, with clean surface.
California	One specimen from one furnace.
Elgin	One specimen from one furnace.
Grant	Two specimens from two furnaces.
Gage, Hagaman & Co	Sample from one car load shipped in December, 1879.
Harrison	One specimen from one furnace.
La Plata	One specimen from one furnace.
Ohio and Missouri	One specimen from one furnace.

ANALYSES XXII AND XXIII. BULLION.

	XXII.	XXIII.
Lead (by difference)	99.0798240	98.492379
Silver	0.6112445	0.793417
Gold	0.0000888	0.000891
Copper	0.0479100	0.071450
Tin	Faint trace	0.000897
Bismuth	Faint trace	0.011791
Arsenic	0.0391365	0.219528
Antimony	0.2138940	0.347881
Iron	0.0063000	0.012600
Zinc	0.0016052	0.000232
Cadmium	Faint trace	Faint trace
Sulphur	None	0.048934
	100.0000000	100.000000
Ounces of silver to the ton	178.275	231.408
Ounces of gold to the ton	0.026	0.260

Discussion.—Analysis XXIII enabled the writer to detect the presence of a great number of metals, some of which, like tin, were not even suspected to exist in Leadville, inasmuch as the sample analyzed represents ores from nearly every mine in the region. While investigating this sample of bullion it was observed that part of the silver exists there in the state of sulphide. Some of the lead, as might be anticipated, is also in the state of sulphide. This is very easily demonstrated in the following manner: The bullion is dissolved in weak nitric acid; the unattacked residue is both yellow and black. The yellow portion is sulphur from sulphide of lead, which is easily attacked by weak nitric acid with separation of sulphur; the black portion is formed of sulphide of silver, which is not touched by weak nitric acid. Neither the relative proportion of silver existing in bullion in the metallic state, nor the amount of lead as sulphide, was determined, because this kind of research would have led too far; but it would seem to be sufficient to call the attention of smelters and refiners to the fact.

Assays of bullion.—The following assays show the varying proportions of gold and silver in the bullion. The specimens assayed are those which had been mixed for analyses:

Location.	Smelter.	Bullion.	
		Silver.	Gold.
		Ounces to ton.	*Ounces to ton.*
California gulch	Billing & Eilers	87.2817	0.1423
	California	216.2267	0.0233
	Grant	325.1550	Faint trace
	Grant	368.5000	Faint trace
	Harrison	132.9417	0.3933
	La Plata	178.2750	0.0260
Average		218.0633	0.081
Big Evans gulch	Cumming & Finn	366.8787	1.1223
	Elgin	245.9750	0.1500
	Gage, Hagaman & Co.	127.2517	0.0833
	Ohio and Missouri	265.5984	0.6566
Average		251.4256	0.503

Thus it will be noticed that the bullion produced in Big Evans gulch is generally richer in gold than that of California gulch.

Assays of bullion made at Messrs. Cumming & Finn's smelter in August, 1880.

[Each assay represents a car-load of 10 tons of bullion.]

	Silver.	Gold.
	Ounces to ton.	*Ounces to ton.*
No. 1	327.9	0.0415
No. 2	301.5	None
No. 3	338.25	None

(Hadelberg.)

Daily assays of bullion made at one of the smelters in Leadville

[Each assay represents the bullion extracted in 24 hours.]

Dates.	Silver.	Gold.
1880.	*Ounces to ton.*	*Ounces to ton.*
May 30	314.25	1.1
May 31	289.70	1.3
June 1	281.57	1.75
June 2	266.675	1.325
June 3	260.22	1.10
June 4	275.60	0.65
June 6	357.175	0.50
June 7	414.46	0.20
June 8	443.85	0.15
June 9	318.90	0.10
June 10	322.4625	0.0375
June 11	284.5	None
June 12	278.0	None

Skimmings.—The following is an analysis of skimmings collected in the siphon-tap or lead-pot of one of the furnaces at the Grant smelter; it is interesting because it contains, in concentrated form, the metals which exist only in small quantity in the bullion, and thus more certainly proves their existence:

Analysis XXIV. Skimmings.

Lead	97.9172
Silver	0.8657
Copper	0.0359
Bismuth	0.0160
Iron	0.4249
Cobalt	0.0087
Nickel	Faint trace
Zinc	0.0158
Arsenic	1.1875
Antimony	0.1147
Tin	0.0095
Sulphur	3.3400
Oxygen and loss (by difference)	1.0641
	100.0000

Silver, 252.5 ounces to the ton. Gold, not a trace.

Discussion.—In the skimmings, as in the bullion itself, part of the silver and some lead exist in the state of sulphides; in fact, the skimmings are peculiar alloys of metals, sulphides, and oxides. Although it was known from the analyses of the ores by Dr. W. F. Hillebrand, and of the hematites by the writer, that cobalt was present in the smelting charge, the writer was extremely surprised not to find this metal concentrated in the speiss or in any of the other furnace products, mattes, accretions, etc. The preceding analyses show that it is in the skimmings that it must be looked for. This curious fact illustrates a most interesting case of separation of nickel from cobalt by the dry way, and by a method hitherto unknown and unsuspected. Nickel, as will be seen, is concentrated in the speiss, and cobalt accompanies the bullion, from which it can easily be separated by the simple process of skimming. There would seem to be no reason why this simple process should not be used in the metallurgy of nickel and cobalt; for no cobalt is found either in speiss or bullion. When the skimmings are cupelled, the presence of cobalt is revealed by the formation of blue specks of phosphate of zinc and cobalt. This phenomenon is so rarely seen that it should not pass unnoticed here.

The skimmings are covered with a crystalline, yellowish-black scum, from which they cannot be separated. When they are broken to pieces, the pieces are crystalline, with a white metallic luster, similar to lead. These pieces flatten under the hammer, but the flattened portions are very brittle, with a crystalline structure and a blackish color, due to small but very distinct crystals of galena.

Losses.—The loss of lead and silver in smelting is thus estimated at the different smelters:

	A.	B.	C.	D.	E.	F.	G.	H.
	Per cent.	*Per cent.*	*Per cent.*	*Per cent.*	*Per cent.*	*Per cent.*	*Per cent.*	*Per cent.*
Loss of lead..........	10 to 15	10 to 15	5 to 15	13	7 to 10	12 to 15	12	9 to 14
Loss of silver.........	2.5 to 5	5	5 to 12	1.5	3	None	3	3 to 5

The average loss is: Lead, 11.68 per cent.; silver, 3.59 per cent.

Part of the loss in both lead and silver is recovered in the smelting of the lead fumes.

Bullion capacity of smelters.—The charges contain on an average 20 per cent. of lead, of which 88 per cent. is extracted in the state of bullion; hence it is easy to calculate the bullion capacity of each smelter, as is done in the following table:

TABLE VIII.—*Bullion capacity of smelters.*

Smelter.	No. of furnaces running.	No. of charges each per twenty-four hours.	Total No. of charges per twenty-four hours.	Weight of each charge.	Total weight of charges run in twenty-four hours.	Lead in charges (20 per cent. of whole).	Lead in bullion (88 per cent. of lead goes into bullion.)	
				Pounds.	*Pounds.*	*Pounds.*	*Pounds.*	*Tons.*
A........................	2	300	600	200	120,000	24,000	21,120	10½
B.....................	6	88	528	700	369,600			
	3	150	450	700	315,000			
Total..................	....	....	978	....	684,600	136,920	120,489	60¼
C........................	2	100	200	920	184,000	36,800	32,384	16⅕
D........................	2	60	120	1,000	120,000	24,000	21,120	10½
E........................	1	100	100	380	38,000			
	1	175	175	380	66,500			
Total..................	....	....	275	....	104,500	20,900	18,392	9⅕
F........................	2	65	130	780	101,400	20,280	17,846	$8\frac{9}{10}$
G........................	1	96	96	790	75,840			
	2	76	152	790	120,080			
	1	70	70	790	55,300			
Total..................	....	....	318	....	251,220	50,244	44,215	$22\frac{1}{10}$
H........................	3	62	186	700	130,200			
	1	100	100	700	70,000			
Total..................	....	....	286	....	200,200	40,040	35,235	17⅗

	Tons.
Total amount of bullion extracted during twenty-four hours (weight of silver and other impurities being neglected)..	155¼
It has already been seen (Table IV) that the smelting capacity of the smelters in tons of ore is........................	700
Also, that the daily output of the mines is from 700 to 800 tons of ore, giving on an average............................	750

Bullion production.— In Table IX are given the amount and value of gold, silver, and lead produced at the different smelters in the month of December 1880; also, in the last column, the value of the total product for the year 1880.

TABLE IX.—*Bullion production.*

Smelter.	Bullion shipments for December, 1880.	Silver contents.			Gold contents.			Lead contents.	Total value of product.	
		No. of ounces per ton.	Total weight.	Value.	No. of ounces per ton.	Total weight.	Value.	Value.	For December.	For year 1880.
	Pounds.	*Ounces.*	*Ounces.*		*Ounces.*	*Ounces.*				
American	380, 000	125	23, 750	$26, 720	....	....	....	$17, 033	$43, 753	$299, 126
Billing & Eilers.......	1, 624, 000	130	105, 560	118, 755	....	....	....	72, 693	191, 448	2, 105, 701
California	441, 720	197. 5	43, 738	49, 208	....	....	....	15, 255	64, 463	702, 826
Cumming & Finn....	553, 500	367. 5	101, 808	114, 534	....	....	....	24, 621	139, 155	1, 324, 213
Elgin	187, 575	175	16, 407	18, 458	....	....	....	8, 395	26, 853	462, 439
Gage, Hagaman & Co..	None	None	None	None	....	....	....	None	....	213, 697
Grant................	1, 180, 000	242	142, 780	160, 628	....	....	....	52, 698	213, 326	4, 018, 290
Harrison.............	322, 263	156	25, 116	28, 280	0. 40	64. 5	$1, 280	14, 431	43, 991	917, 304
La Plata..............	895, 920	162. 5	73, 010	78, 761	......	....	....	40, 120	118, 881	2, 316, 310
Leadville	None	None	None	None	....	....	....	None	....	14, 218
Little Chief	None	None	None	None	....	....	....	None	....	109, 072
Lizzie	None	None	None	None	....	....	....	None	.. ..	63, 691
Malta.................	127, 316	83. 7	5, 320	5, 985	1. 46	93. 0	1, 800	5, 715	13, 500	24, 362
Ohio and Missouri....	157, 057	327. 5	25, 700	28, 910	....	....	....	6, 995	35, 905	822, 656
Totals and averages.	5, 869, 351	191. 86	563, 189	630, 239	....	157. 5	3, 080	257, 956	891, 275	13, 493, 905

SLAG.

Fire assays of slag, made by J. E. Hardman in March, 1880.

Smelter.	Lead.	Silver.	Smelter.	Lead.	Silver.
	Percentage.	*Oz. to ton.*		*Percentage.*	*Oz. to ton.*
La Plata.........	1. 8	3	Leadville	11. 5	10
Do............	4. 0	3	Billing & Eilers...	2. 5	3
Do............	11. 8	9	Do............	3. 3	3
California.........	9. 5	4. 5	American.........	1. 1	1
Grant	13. 2	5	Do............	0. 75	0

Fire assays made at the California smelter on California smelter slag from February to May, 1880, by J. E. Hardman.

Lead, percentage ...	1. 6	2. 5	0. 8	1. 6	2. 8	0. 7	1. 5	0. 2	1. 4	2. 2
Silver, ounces	1. 0	0. 5	0. 5	0. 5	0. 75	0. 5	1. 0	0. 5	1. 25	0. 5
Lead, percentage ...	2. 3	1. 0	1. 3	0. 7	3. 0	1. 0	1. 6	2. 0	0. 5	1. 5
Silver, ounces.......	0. 75	0. 5	0. 5	0. 5	1. 0	0. 0	0. 5	0. 5	0. 5	1. 0

Silver assays of Malta smelter slag, made by Robert Bunsen.

Silver, from 10 to 17 ounces to the ton.

Fire assays of slag, by Dr. M. W. Iles, made at Grant smelter.

Smelter.	Lead.	Silver.
	Percentage.	*Oz. to ton.*
Grant	4.50	3.00
Billing & Eilers	3.75	3.50
Cumming & Finn	2.00	4.75

The following table is of assays of slag and of corresponding bullion, made daily by the writer at one of the Leadville smelters, and represents the work done in each furnace in twenty-four hours.

TABLE X.—*Daily slag assays.*

	Furnace No.	Lead.	Silver.	Silver contents of bullion.
		Percentage.	*Oz. per ton.*	*Oz. per ton.*
May 30, 1880	1	6.7	6.0	
	2	3.7	4.5	
	3	3.7	4.0	
Square furnace	4	4.9	3.5	
Average		4.7	4.5	314.25
May 31, 1880	1	5.6	6.0	
	2	2.0	3.5	
	3	3.6	4.0	
	4	3.8	3.0	
Average		3.7	4.1	289.70
June 1, 1880	1	6.7	6.0	
	2	3.7	3.5	
	3	3.7	4.0	
	4	4.9	3.0	
Average		4.7	4.1	281.57
June 2, 1880	1	5.6	6.0	
	2	2.0	3.5	
	3	3.6	4.0	
	4	3.8	3.0	
Average		3.7	4.1	266.67
June 3, 1880	1	5.2	4.5	
	2	2.0	3.0	
	3	3.0	3.0	
	4	1.0	1.5	
Average		2.8	3.0	260.22
June 4, 1880	1	4.0	3.5	
	2	2.0	3.5	
	3	2.5	3.5	
	4	2.5	3.0	
Average		2.7	3.4	275.60

TABLE X.—*Daily slag assays*—Continued.

	Furnace No.	Lead.	Silver.	Silver contents of bullion.
		Percentage.	*Oz. per ton.*	*Oz. per ton.*
June 5, 1880	1	2.1	2.7	
	2	2.5	1.9	
	3	2.9	3.3	
	4	3.0	3.3	
Average	....	2.6	2.8	....
June 8, 1880	1	3.5	3.2	
	2	2.5	2.8	
	3	2.7	2.5	
	4	3.1	2.8	
Average	....	2.95	2.8	443.85
June 9, 1880	1	3.4	2.9	
	2	4.2	3.5	
	3	3.1	2.8	
	4	2.8	2.2	
Average	....	3.37	2.8	318.90
June 10, 1880	1	3.5	6.1	
	2	2.5	3.85	
	4	2.5	3.85	
Average	....	2.8	4.6	322.46
June 11, 1880	4	3.5	6.1	284.5
June 12, 1880	1	3.8	3.3	
	2	2.5	3.0	
	4	4.7	4.1	
Average	....	3.7	3.5	278.0

The above figures show the influence of the numerous elements with which the smelter has to contend. The influence of the head smelter and of the furnace is indicated by the fact that Furnace No. 1 gives nearly always the slag richest in lead and silver. The influence of the silver contents of the bullion is clearly seen in the cases in which the richest slags correspond to the richest bullion. The overpowering influence of the composition of smelting charges is forcibly indicated in the cases in which poorer slags correspond to a richer bullion.

Specific gravity of slag.—At Messrs Cumming & Finn's smelter the specific gravity of slag from each furnace is determined daily. It results from a very great number of determinations made by the superintendent, Mr. MacFarlane, by means of the Jolly specific-gravity spring-balance, described in Section I, that the average specific gravity of slag varies between 3.7 and 3.8.

Dr. M. W. Iles has obtained an average of 3.691 from a hundred determinations made on unusually fine runs at the Grant smelter.

ANALYSES XXV–XXIX.—*Analyses of Leadville slag.*

	XXV.*a*	XXVI.*a*	XXVII.*a*	XXVIII.*a*	XXIX.*b*
Silica	28.50	26.45	39.100	30.200	28.20
Protoxide of iron	42.20	37.11	39.085	36.186	47.07
Peroxide of iron	9.50	7.88	....	....	....
Protoxide of manganese	5.21	5.20	5.023	3.813	7.13
Alumina	0.62	4.29	0.849	4.293	0.89
Lime	4.50	8.38	8.300	22.800	7.10
Magnesia	3.06	3.23	3.800	0.144	3.46
Oxide of lead	6.51	6.19	2.649	2.355	5.25
Sulphur	0.82	0.55	1.194	0.618	0.90
	100.92	99.28	100.000	100.409	100.00
Lead by fire assay, per cent	4.5	....	....	1.25	....
Silver, ounces to the ton	3.0	....	....	1.00	....
Specific gravity	3.8	....	....	3.58	....

a Iles. *b* Hardman.

Analyses Nos. XXV and XXVI, made by Dr. M. W. Iles, are of Grant's old slags of the singulo-silicate type.

Analysis No. XXVII, made by Dr. M. W. Iles, is of a slag now made at the Grant smelter. It belongs to the acid type.

Analysis No. XXVIII, made by Dr. M. W. Iles, is that of a slag of the singulo-silicate type, made at Messrs. Billing & Eilers's smelter, with pure arragonite instead of dolomite. It is remarkable for its large percentage of lime.

Analysis No. XXIX is of a slag of the singulo-silicate type, made at the California works, and analyzed by Mr. J. E. Hardman.

From the preceding figures it will be seen that the composition of slags is well understood in Leadville, although some obscure points, such as their magnetic properties and the state in which sulphur exists in them, need elucidation, and although some metals always present in slags, such as zinc, and substances such as phosphoric and titanic acids, are not reported.

Special researches on slags made in the laboratory of the Survey.—The word *slag* seems appropriate to designate the strange products which flow from the blast furnaces during the process of lead smelting. These products are sometimes masses of large intersected crystals, brittle, with a vitreous luster; sometimes fine-grained tough masses, with a dull fracture, but always dark colored and opaque. On the other hand, the word *scoriæ* ought to be adopted for translucent or transparent slags. Scoriæ are accidentally formed in the blast furnaces, having been found by the writer in the cavities of iron sows. There is no doubt that they are regularly formed during the process of smelting, but are soon transformed into slag, so that only slag flows from the furnaces.

A rough qualitative examination was made of the scoriæ found intimately mixed with iron sows; the color was that of pure blende; they were translucent—almost transparent—contained no sulphur, and consisted almost exclusively of silicate of protoxide of iron and manganese, with traces only of lime and magnesia. This accidental product, which probably no one else has ever perceived in Leadville, affords a means of studying the nature of the reactions which take place in the blast furnace, and which such accidents alone can reveal.

Slags are not scoriæ. They do not belong to the type of glasses, since they are opaque[1] and crystalline. They are not artificial minerals, since they contain large quantities of sulphurets. Instead of belonging to some well-known type, they form one. It is only after a careful study of their nature and properties that it will be possible to attempt to give a satisfactory definition of these products.

Properties of slag.—1. Pulverized slag treated with the magnet almost always shows the presence of a magnetic portion which adheres strongly to the magnet. A slag beautifully crystallized in detached rhomboidal laminæ, with a steel-gray color and an almost metallic luster, from the La Plata smelter, could be separated by the magnet into two portions.

	Parts.
A strongly magnetic portion, amounting to	38
A feebly magnetic portion, amounting to	62
	100

But a rough examination, both quantitative and qualitative, of these two portions showed no great difference in the composition, and the investigation was carried no further in this direction.

2. The same slag finely pulverized[2] and treated by weak sulphuric acid (acid 1, water 4) is rapidly attacked. Sulphureted hydrogen is evolved, showing the presence of sulphides easily attacked by weak acids. The slag is, moreover, thoroughly disintegrated after a few hours. A large proportion of silica, iron, lime, magnesia, manganese, and zinc is dissolved. An unattacked residue is left; it is treated with weak nitric acid, which dissolves some sulphide of lead, formed evidently during the reaction, for it has the aspect of artificial sulphide of lead formed in the wet way. The residue is then boiled with carbonate of soda, which dissolves some gelatinous silica. A residue is still left; it is attacked a second time by weak sulphuric acid, weak nitric acid, and carbonate of soda. It is interesting to observe that after each successive treatment sulphureted hydrogen is evolved, showing that the sulphides are undoubtedly combined with silica or with silicates. After each treatment, silica, iron, lime, etc., are dissolved. These treatments are repeated until the residue consists of intensely black, fine, brilliant crystals. It is formed of pure magnetic oxide of iron, which is resolved into octahedra under the microscope. This oxide was analyzed; it contained—

Protoxide of iron	40.3
Peroxide of iron	59.7
	100.00

Its formula is $Fe_7O_9=3FeO, 2Fe_2O_3$, instead of $Fe_6O_8=2FeO, 2Fe_2O_3=2Fe_3O_4$, the formula of ordinary magnetite. It contains one equivalent of protoxide of iron more than normal magnetite. In hematite has been seen a magnetite containing an excess of peroxide of iron; here is found a magnetite formed in the midst of protoxide of iron and containing an excess of this oxide.

---

[1] The opacity of some slags is such that thin sections, prepared by Mr. Whitman Cross for microscopical examination, proved totally opaque, even under the microscope.

[2] Sifted slag is also very easily attacked by weak acids.

One problem is solved—the slags are magnetic, because they contain free magnetite disseminated throughout their mass; the magnetite is not combined, since it can be thus isolated in a state of purity, and it is evidently to this substance that the intense black color of slags is chiefly due. To this substance also they partially owe their opacity.

Magnetite can be isolated by a process much more simple and more rapid than the one previously described. Sifted slag is attacked in a platinum vessel by a mixture of weak nitric and hydrofluoric acids; the solution is decanted, the residue is treated with a boiling solution of caustic potash, and this residue is washed with water and weak hydrochloric acid. In a few minutes pure magnetite is isolated.

3. The pulverized slag is treated by a boiling solution of caustic potash; after a few minutes ebullition the potash is charged with sulphide of potassium, and in a few minutes more it takes the rich yellow color of persulphide of potassium. Only one among the sulphides that can possibly exist in the slags is capable of producing this reaction; it is sulphide of calcium. The existence of this sulphide, which has long been suspected and often reported, is demonstrated here beyond a doubt. Whether all the sulphur of slag exists in the state of sulphide of calcium is another question. That most of the sulphur is in that condition there is no doubt, but from the general behavior of slag, the writer is almost inclined to think that small quantities of sulphides of iron, manganese, zinc, and even lead exist there also. A great number of experiments were made to ascertain this, but in every case the presence of metallic sulphides might be attributed to secondary reactions, so they will not be described.

4. The pulverized slag is treated by a strong solution of cold potash. A considerable quantity of oxide of lead is dissolved; consequently there can be no doubt as to the state in which lead exists in slags. It is in the state of silicate of oxide.

5. Slags contain always a little chlorine, whose quantity is proportionate to the quantity of silver found; hence there is little doubt that silver exists in the slag in the state of chloride which has escaped decomposition. This fact is important because it explains why there is no relation between the quantities of lead and silver found in slag. The slag in indistinct but large crystals behaves with reagents exactly like the distinctly crystalline one; magnetite can be extracted by the processes previously described, but this oxide, instead of being crystalline to the eye, forms an apparently amorphous powder. The non-crystalline, fine-grained slags possess the same properties as the former. They are more easily attacked under the same circumstances and yield only traces of magnetite; yet they contain almost as much peroxide of iron as the former, but in this case peroxide of iron exists in the state of silicate.

Most slags in Leadville belong to the two types just described: the lustrous crystalline slag known as acid slag, which may be defined as a silicate of sulphides and oxides, colored by magnetite, and the fine-grained, non-crystalline known as basic slag, and which may be termed a silicate of sulphides and oxides, colored by sulphide of iron.

Complete analyses of slags.—The writer made the following analyses of slags, which reveal a few points which had not been observed before, such as the presence of small quantities of carbonate of lime and of carbon:

ANALYSES XXX, XXXI, AND XXXII. SLAGS.

Elementary.

	XXX.	XXXI.	XXXII.
Silica	29. 0123	33. 845650	31. 10656
Titanic acid	0. 5285	0. 820000	0. 57300
Sulphuric acid	None	Faint trace	Trace
Carbonic acid	None	Trace	Trace
Phosphoric acid	0. 8788	0. 657828	0. 32204
Arsenious acid	Trace	0. 024147	Trace
Chlorine and traces Br and I (calculated)	0. 0031	0. 004686	0. 00184
Sulphur	1. 9110	0. 914418	1. 27277
Calcium (in the state of sulphide)	2. 3887	1. 143022	1. 59096
Silver	0. 0096	0. 014266	0. 00562
Gold	Marked trace	Trace	Trace
Protoxide of iron	44. 5226	36. 789600	34. 40836
Peroxide of iron	(*a*)	(*a*)	2. 71000
Magnetic oxide of iron (Fe_7O_9)	2. 9500	3. 419000	(*b*)
Oxide of lead	6. 3138	4. 326576	3. 31746
Oxide of antimony	0. 0140	Trace	Trace
Oxide of zinc	1. 8040	2. 353500	1. 04100
Protoxide of manganese	2. 8606	4. 259378	1. 15720
Suboxide of copper	None	None	0. 04950
Lime	1. 9618	6. 539400	10. 28020
Magnesia	2. 9814	3. 671935	9. 75340
Alumina	1. 8427	1. 182000	2. 33500
Alkalies	Trace	Trace	Trace
Carbon	0. 0195	0. 034500	0. 07500
Total	100. 0024	99. 999906	99. 99991

[a] In magnetic oxide.

[b] Reported with FeO and Fe_2O_3.

ANALYSES XXX, XXXI, AND XXXII. SLAGS — Continued.

Rational.

	XXX.	XXXI.	XXXII.
Silica	29. 0123	33. 845650	31. 10656
Titanic acid	0. 5285	0. 820000	0. 57300
Protoxide of iron	44. 5226	36. 789600	34. 40836
Peroxide of iron	None	None	2. 71000
Alumina	1. 8427	1. 182000	2. 33500
Lime	0. 9221	5. 761125	9. 89919
Magnesia	2. 9814	3. 671935	9. 75340
Alkalies	Trace	Trace	Trace
Oxide of zinc	1. 8040	2. 353500	1. 04100
Protoxide of manganese	2. 8606	4. 259378	1. 15720
Suboxide of copper	None	None	0. 04950
Oxide of lead	6. 3138	4. 326576	3. 31746
Arsenious acid	Trace	0. 024147	Trace
Antimonious acid	0. 0140	Trace	Trace
Phosphate of lime	1. 9185	1. 436103	0. 70305
Sulphide of calcium	4. 2997	2. 057440	2. 86373
Sulphate of lime	None	Faint trace	Trace
Carbonate of lime	None	Trace	Marked trace
Chloride of silver, with traces AgBr, AgI	0. 0127	0. 018952	0. 00746
Gold	Marked trace	Trace	Trace
Magnetic oxide of iron (Fe_7O_9)	2. 9500	3. 419000	(*a*)
Carbon	0. 0195	0. 034500	0. 07500
Total	100. 0024	99. 999906	99. 99991
Percentage of lead	5. 85	4. 000	3. 07
Silver (ounces to the ton)	2. 80	4. 161	1. 639
Gold (ounces to the ton) about	0 005	0. 0005	0. 0005

a Reported with FeO and Fe_2O_3.

No. XXX is the slag, in distinct detached crystals, from the La Plata smelter, which has already been described in the reactions of slags. No. XXXI is a sample of so-called acid slag from Cumming & Finn's smelter. This sample was made of 124 pieces of slag, each piece representing the day's work of one furnace and specimens from four furnaces being mixed together.

No. XXXII is a slag of the singulo-silicate type, taken from the heap at Messrs. Billing & Eilers's smelter. These three specimens have already been described in the investigation on the properties of slags.

Discussion.—A glance at the analyses shows:

1. That the quantity of lead is in an inverse ratio to the quantity of lime and magnesia existing in the state of silicates.

2. That there is no relation whatever between the quantities of lead and of silver left in the slag. This was shown also in the assays of slags from various sources already given. It is rendered very apparent in an assay of another slag from Billing & Eilers's smelter exactly similar to No. XXXII, which was examined very carefully by the wet way for lead and assayed for silver. It contained 2.95 per cent. of lead and 0.5833 ounce of silver, while No. XXXII contains lead, 3.07 per cent.; silver, 1.639 ounces.

This can scarcely be otherwise since lead exists in the slag in the state of combination and silver in the state of mixture.

None of the slags analyzed contains any baryta, but, as this substance has been found in some of the lead fumes condensed in the dust-chambers, it must be inferred that some of the Leadville slags contain baryta.

The slags were examined for chromium, tungsten, and vanadium, but the presence of these metals could not be detected.

Titanic acid could not be detected by the classical methods. The process used with success for its detection and estimation was the following: The slag is dissolved in a mixture of hydrofluoric, hydrochloric, and sulphuric acids, in a platinum vessel, and the whole evaporated until sulphuric acid goes off in fumes (this to expel silica). The product, dissolved in water, is treated by an excess of sulphureted hydrogen to precipitate any lead which might remain in solution. The solution is filtered and then boiled for the expulsion of sulphureted hydrogen, and then brought as nearly as possible to the neutral point by an alkali. Alumina and titanic acid are then precipitated by hyposulphite of soda, and separated and estimated as usual. With the exception of the preliminary operation needed for the preparation of the solution, the process is the same as the one which has been recommended in the analysis of hematite. The quantity of titanic acid in the hematite is insufficient to account for the relatively large proportion of this acid in the slags. In all probability the oxide of iron of the lead ore contains this substance, but some small quantities of titanate of lead may also exist in the ore, although this mineral is not known to exist. As has previously been stated, the slags from the Little Chief smelter were examined very carefully and only doubtful traces of titanic acid were detected.

Careful experiments revealed the presence of carbon in the slag, but it is not known yet in what state it occurs there. Two hypotheses are acceptable: either carbon exists in the state of graphite, a very easily mixed substance, and one which combines, so to speak, crystallographically, as in cast iron, for instance, or else in the state of carburet of iron; it is possibly part of the carburet of iron of the iron reduced in the furnace at a certain stage of the smelting operation, and subsequently fluxed down by sulphur, arsenic, and silica. That carbon is thus liberated there is no doubt, for the writer has been able to isolate graphite blown off with the lead fumes in the chamber-dust.

Strange as it may appear, there is no doubt whatever about the presence of small quantities of carbonate of lime in the slag. This quantity is undoubtedly proportionate to the rapidity with which the furnace is run.

The large proportion of phosphate of lime in the slag, derived from the pyromorphite of the ore, is one of the causes of its opacity, it being well known that phosphate-of-lime glasses are opaque. The presence of large quantities of sulphide of calcium in the crystalline portions of the slags is another very clear indication that sulphide of calcium is really combined with the silicates, either chemically in the state of sulpho-silicate, or possibly crystallographically.

Although we are in possession of a good many facts relative to slags, it seems necessary to postpone an attempt at a rational definition of these products until we have examined the mattes of Leadville. (See observations on mattes, and final definition of slag.)

Assays of slags made in the laboratory of the Survey.—In Table XI is given the following information concerning the slags collected in Leadville and assayed in the laboratory of the Survey:

1. Reference numbers for discussion.
2. Names of the smelters.
3. Character of the slag.
4. Color of the powdered slag.
5. Remarks as to whether the slags are normal or accidental.
6. Number of specimens mixed for assay.
7. Places where the slags were collected.
8. Portions of the cakes from which the specimens were taken.
9. Assays of the slag in silver, ounces to the ton.
10. Assays of the slag in gold, ounces to the ton.

TABLE XI.—*Assays of slag.*

	Where from.			Character of specimens.
	Name of smelter.	Part of works.	Part of slag-cake, &c.	General composition.
1	American	Slag-heap	Core of cake	Normal.
2	do	do	Slag-cake	Accidental.
3	do	do	do	Normal accident.
4	California	do	Core of cake	Normal.
5	do	do	Detached pieces	Accidental.
6	Cumming & Finn	do	Core of cake	Normal.
7	do	Laboratory	do	Do.
8	Elgin	Given by superintendent		Normal accident.
9	do	Slag-heap	Core of cake	Normal.
10	Gage, Hagaman & Co	do	do	Do.
11	Grant	New slag-heap	Cakes	Normal; new slag.
12	do	do	do	Do.
13	do	Large square furnace, running too fast	Tap-hole	Do.
14	do	Small round furnace, running well	do	Do.
15	do	Old slag-heap	Cakes	Normal; old slag.
16	do	Slag-pot	Shell of cake	Normal; new slag.
17	do	Same slag-pot	Interior of cake	Do.
18	Harrison	Slag-heap	Core of cake	Normal.
19	La Plata	do	Cakes	Normal accident.
20	do	do	do	Normal.
21	Leadville	do	do	Do.
22	Little Chief	do	do	Do.
23	do	do	Detached pieces	Accidental.
24	Malta	do	Cakes	Normal; old slag.
25	Ohio and Missouri	do	do	Normal.
26	do	Large furnace	Tap-hole	Do.
27	Raymond, Sherman & McKay	Slag-heap	Core of cake	Do.
28	Billing & Eilers	do	Shell of cake	Normal, dolomitic old slag.
29	do	Smelting charge	do	Do.
30	do	Slag-heap	Core of cake	Do.
31	do	Furnace	Tap-hole	Normal.
32	do	Given by A. Eilers	Shell of cake	Normal, calcitic new slag.
33	do	do	Interior of cake	Do.
34	do	Slag-pot	Top of cake	Do.
35	do	Same slag-pot	Interior of cake	Do.

TABLE XI.—*Assays of slag*—Continued.

Character of specimens.		Assays.			
External appearance.	Color of powder.	No. of specimens mixed.	Silver.	Gold.	
			Ozs. to ton.	*Ozs. to ton.*	
Compact; fine-grained, with dull fracture	Grayish	3	1.8569	Faint trace.	1
Minute prismatic crystals	Yellow-white	1	0.95	Do.	2
Large lamellar crystals	Yellow-gray	1	0.73	None.	3
Fine, compact, and indistinctly crystalline	Gray-black	2	0.670818	Faint trace.	4
Acicular	Yellow-gray	1	0.76	None.	5
Vitreous; large indistinct crystals	do	1	6.17347	Trace.	6
Vitreous; crystalline; also compact	Gray-black	124	4.161	Do.	7
Large distinct lamellæ	Gray-yellow	1	0.6	Faint trace.	8
Compact; fine-grained	Blackish	2	4.919332	Do.	9
Vitreous; compact and indistinctly crystalline	Grayish	3	2.24	Do.	10
Vitreous; large indistinct crystals	Gray-yellow	1	5.084606	Do.	11
Vitreous; distinct crystals	do	1	4.2	None.	12
Vitreous; indistinct crystals	Gray-black	1	4.6	Do.	13
do	do	1	2.87	Do.	14
Vitreous; compact	Grayish	2	3.87	aint trace.	15
Vitreous; indistinct crystals	Gray-black	1	2.615	Trace.	16
do	do	1	2.515	Do.	17
Vitreous; compact and indistinctly crystalline.	Yellow-gray	2	1.85	Do.	18
Distinct lamellar crystals	do	2	2.8	0.005	19
Indistinctly crystalline and compact.	do	2	1.4	Trace.	20
Very dense, mottled, dull, non-crystalline	do	1	10.05	Do.	21
Vitreous; indistinct crystals	do	1	3.99	None.	22
Crystalline; looks like hornblende	do	1	4.44	Do.	23
Vitreous; large, indistinct crystals	do	2	6.22	Trace.	24
Vitreous; indistinctly crystalline and compact	do	3	2.78	Do.	25
Compact; vitreous	Brilliant black	1	1.72	None.	26
Vitreous; indistinct crystals	Yellow-gray	1	4.52	Trace.	27
Compact; fine-grained; dull fracture	Grayish	1	0.554154	Faint trace.	28
do	do	1	0.5833	None.	29
do	Yellow-gray	1	1.639	Faint trace.	30
do	Gray-black	1	0.92	None.	31
Lustrous; indistinct crystals	do	9	1.47	Do.	32
Distinct prismatic crystals	Gray-white	8	0.60	Do.	33
Lustrous; indistinct crystals	Gray	1	0.15	Do.	34
do	do	1	0.36	Do.	35

NOTE.—Specimens 16 and 17 were specially prepared by Dr. M. W. Iles, and specimens 34 and 35 by Mr. A. Eilers. The slag represented by numbers 32, 33, 34, and 35 has been analyzed by Dr. Iles (see Analysis XXVIII).

Discussion.—Smelting is conducted so methodically in Leadville that when the writer collected the specimens of slag it was done at random, since he felt convinced that the examination which he was about to make would only furnish additional proof of the admirable method employed; but the preceding table shows that he was greatly mistaken in this anticipation. Laying aside accidental slags, and discussing only the normal ones found by thousands of tons on the refuse-slag heap, it is found that at the Grant smelter, for instance, the composition of smelting charges has been altered from the singulo-silicate type to the acid type, and with what results? The old slag-heap and the new slag-heap are several hundred feet apart—the first immediately in front of the furnaces and the second outside of the works at the bottom of California gulch—so that a mistake on the writer's part in the collection of the specimens was impossible. Now, Table XI shows that the new slag, No. 11, contains 5.1 ounces of silver, while the old slag, No. 15, contains only 3.8 ounces; and at these works it is the poorest slag that is resmelted, while the richest is carried away to an almost inaccessible spot. At Messrs. Billing & Eilers's smelter the same is remarked. It is the poorest slag, No. 29, containing only 0.5 ounce of silver, that is resmelted, while the richest, No. 30, containing more than three times as much silver, or 1.6 ounce, remains on the slag heap. The compact, fine-grained slags, which represent slags of the singulo-silicate type, are conspicuous throughout Table XI for their low contents in silver, and yet we see two of the largest smelters, those of Messrs. Cumming & Finn and Grant, adopting slags of the acid type, containing four and six ounces of silver, like Nos. 7 and 6, and four and five ounces, like Nos. 12 and 11.

The blame for this belongs somewhere, and it is probable that the superintendents are constantly misled and misguided by the assayers, chiefly for the reason that the scorification process is not to be depended upon in the assay of slag and that the crucible process ought to be substituted for it. The clearest result of an inspection of Table XI is that there is no relation whatever between either the appearance or even the composition of slag and its contents in silver, and that smelters ought to give special attention to the assay of these products.

The process of shelling out the slag, which has been described in smelter C, induced the writer to make a few experiments on the distribution of silver in the cakes of slag, and specimens 16 and 17 were prepared specially for this purpose from the same slag-pot by Dr. M. W. Iles. The shell, No. 16, contains 2.6 ounces of silver, and the poured-out portion 2.5, showing a difference of one-tenth of an ounce in favor of the shell. The difference is much larger between the top shell and the poured-out portion, as is shown by specimens 34 and 35, prepared specially from the same slag-pot by Mr. A. Eilers. The top shell contains 0.15 ounce silver and the poured-out portion 0.36 of an ounce. These two experiments seem to indicate that during the process of cooling the chloride of silver, which is only mechanically mixed in the slag, settles, in virtue of its higher specific gravity, by means of a sort of liquation. Specimens 32 and 33 point out the same results. The outer portion of the cakes of slag, in indefinite crystals, assays 1.47 ounces, and the distinct crystals, forming the inner portion or core of the cakes, 0.6. There can be no doubt about these results, since the crystals were detached from the outer portions immediately before assaying; nor can the differences be attributed to the presence of traces of bullion, for in no case did the slag contain even a trace of metallic grains. These results were not suspected by the smelters, who were under the impression that there was no difference between the different parts of the same cake of slag, and who used the ingenious process of shelling out as a convenient way of breaking up the slag in small pieces before resmelting.

A comparison between specimens 28, 29, 30, and 31, or dolomitic slags, and specimens 32, 33, 34, and 35, or calcitic slags, shows that both kinds of slag contain sensibly the same amount of silver.

CHAMBER-DUST.

The flue and chamber dust of Leadville is always in the form of a coarse reddish or blackish powder and full of very small particles of charcoal and coke.

Very little has been done in Leadville with regard to a thorough examination of these products, and all the information which could be obtained bears on the estimation of lead and silver, and occasionally of silica and iron. In the following table is condensed such information as could be obtained:

TABLE XII.—*Assays of chamber-dust.*

Smelter.	Lead.	Silver.	Silica.	Iron.	Assayer.
	Per cent.	*Oz. to ton.*	*Per cent.*	*Per cent.*	
California	30-35	20-25			J. E. Hardman.
Cumming & Finn	35	36	20	14	Hadelberg.
Elgin	40-60	50-150			
Grant	28	34.7	14.3	8.65	M. W. Iles.
Grant	20-28	34-44			Do.
Harrison	40-50	35-40			Th. Fluegger.
La Plata	30-40	40			
Little Chief	26	44.5	16.1	11.56	M. W. Iles.
Ohio and Missouri	30	40			
Utah	15-35	10-25			

In the description of each smelter the amount of flue-dust caught and the methods of treatment of flue and chamber dust have already been given.

Analysis.—An examination of the flue and chamber dust of the blast furnaces of Leadville afforded such a fine opportunity for detecting most of the substances disseminated in the camp that the writer carried on quite exhaustive researches on these products, and in order not to let anything escape he treated the dust with boiling water and made a careful analysis of the soluble portion, then a careful examination of the portion soluble in acids, and lastly a complete analysis of the portion insoluble in acids. The results thus obtained are extremely complicated and present an unnatural appearance, but such as they are they give a clearer idea of the form under which the different compounds exist in the fumes, and of the reactions to which they owe their origin, and no attempt was made to simplify the reports.

The labor expended on the examination of the lead fumes was rewarded by the discovery of a new metal which appears to be distributed widely, though sparingly, throughout the camp. In the elementary analyses the earthy and alkaline metals in combination with metalloids other than oxygen have been reported in the metallic state.

Analysis No. XXXIII is that of a sample of chamber and flue dust from the Ohio and Missouri smelter. This dust had a reddish-brown tinge, due to peroxide of iron.

Analysis No. XXXIV is that of a mixture of equal parts of flue and chamber dust from the American, California, Cumming & Finn, Grant, La Plata, and Billing & Eilers smelters. The sample thus formed had a blackish color, due chiefly to fine charcoal dust.

ANALYSES XXXIII AND XXXIV. CHAMBER DUST.

Elementary.

	XXXIII.	XXXIV.
Lead	25. 535772	38. 620520
Silver	0. 130233	0. 121200
Gold	0. 000100	0. 000066
Bismuth	0. 013460	0. 090640
Copper	0. 003992	0. 099000
Cadmium	0. 017500	0. 012200
Iron	16. 065305	13. 340060
Manganese	1. 478780	0. 598645
Zinc	3. 311400	1. 303740
Arsenic	0. 176715	0. 090910
Antimony	0. 083560	0. 087740
Tin	0. 001180	0. 001180
Aluminium	0. 011300	0. 003000
Calcium	0. 235420	0. 224700
Magnesium	0. 022000	0. 020900
Potassium	0. 026000	0. 071000
Sodium	0. 175000	0. 114000
Alumina	1. 955000	2. 627000
Lime	4. 197600	3. 214880
Magnesia	0. 835119	2. 275000
Al_2O_3, Fe_2O_3, PbO, ZnO, CaO, MgO, K_2O, Na_2O (combined with SiO_2).	2. 391447	....
Water	0. 585000	0. 912500
Sulphuric acid	2. 641850	2. 440820
Carbonic acid	1. 993000	3. 175120
Phosphoric acid	0. 185484	0. 487400
Titanic acid	0. 008000	0. 035000
Silica (from slag and silicate of lead)	15. 126553	9. 704500
Silica (from quartz and refractory silicates)	2. 407000	3. 780000
Oxygen	8. 972380	7. 919895
Sulphur	0. 444300	1. 174000
Selenium and tellurium	Traces	Marked traces
Chlorine	1. 321660	1. 200490
Bromine	0. 244530	0. 242330
Iodine	0. 012660	0. 012080
Carbon (coke, charcoal, graphite)	9. 240000	5. 063000
Indium, thallium, new metal	Traces	Traces
Baryta	....	0. 215000
Potash	....	0. 035000
Soda	....	0. 025000
Graphite	....	0. 665000
Loss	0. 150700	....
Total	100. 000000	100. 003516
Silver ounces to ton	37. 984	35. 349
Gold do	0. 03	0. 01925

ANALYSES XXXIII AND XXXIV. CHAMBER DUST—Continued.

Rational.

	XXXIII.	XXXIV.
Portion soluble in water.		
Protosulphate of iron	0. 086500	0. 014700
Sulphate of manganese	0. 145000	0. 158020
Sulphate of zinc	0. 020200	....
Oxychloride of lead	0. 255850	0. 206310
Oxybromide of lead	0. 055000	0. 045400
Oxyiodide of lead	0. 003000	0. 002990
Chloride of calcium	0. 553000	0. 523390
Bromide of calcium	0. 175000	0. 174660
Iodide of calcium	0. 009000	0. 008510
Chloride of zinc	0. 120000	0. 118500
Chloride of aluminium	0. 055500	0. 010700
Chloride of magnesium	0. 090000	0. 082900
Caustic magnesia	0. 150000	0. 195000
Chloride of potassium	0. 050000	0. 140000
Chloride of sodium	0. 450000	0. 290000
Water	0. 585000	0. 912500
Total	2. 803050	2. 883580
Portion soluble in acids.		
Oxide of lead	17. 816661	22. 451750
Sulphate of lead	7. 954000	8. 898730
Phosphate of lead	1. 059348	2. 783600
Sulphide of lead	1. 255000	8. 215000
Chloride of lead	1. 726700	1. 725620
Bromide of lead	0. 185790	0. 185080
Iodide of lead	0. 004900	0. 004880
Chloride of silver	0. 142712	0. 132730
Bromide of silver	0. 037670	0. 035040
Iodide of silver	0. 002580	0. 002540
Gold	0. 000100	0. 000066
Protosulphide of iron	0. 400000	0. 200000
Protoxide of iron	7. 855290	6. 100000
Peroxide of iron	13. 814050	11. 491000
Oxide of bismuth	0. 015000	0. 101000
Oxides of indium, thallium, new metal	Traces	Traces
Oxide of zinc	3. 795270	1. 454000
Oxide of cadmium	0. 020000	0. 014000
Sulphide of manganese	0. 082380	Trace
Sulphide of zinc	0. 300000	Trace
Oxide of manganese (Mn_3O_4)	1. 903500	0. 750000
Arsenious acid	0. 233256	0. 120000
Oxide of antimony	0. 100000	0. 105000
Portion soluble in acids—Cont'd.		
Oxide of tin	0. 001500	0. 001500
Selenious and tellurous acids	Trace	Marked trace
Oxide of copper	0. 005000	0. 012400
Titanic acid	0. 008000	0. 034000
Alumina	1. 955000	1. 620000
Carbonate of lime	4. 529600	5. 473000
Lime	1. 661024	....
Magnesia	0. 835119	1. 303000
Carbonate of magnesia	....	1. 464000
Total	67. 699450	74. 677936
Portion insoluble in acids.[1]		
Silica (from slag and silicate of lead)	15. 126553	9. 704500
Silica (from quartz and refractory silicates)	2. 407000	3. 780000
Al_2O_3, Fe_2O_3, PbO, ZnO, MgO, CaO, K_2O, Na_2O (combined with SiO_2).	2. 391447	....
Carbon (from charcoal, coke, and graphite)	9. 240000	5. 063000
Oxide of lead	....	0. 907000
Oxide of zinc	....	0. 100000
Arsenious acid and oxide of antimony.	....	Traces
Baryta	....	0. 215000
Sulphate of baryta	....	Trace
Lime	....	0. 150000
Magnesia	....	0. 080000
Alumina	....	1. 007000
Peroxide of iron	....	0. 600400
Oxide of manganese	....	Trace
Titanic acid	....	0. 001000
Potash	....	0. 035000
Soda	....	0. 025000
Graphite	....	0. 665000
Total	29. 165000	22. 332900
Portion soluble in water	2. 803050	2. 883580
Portion soluble in acids.	67. 699450	74. 677936
Portion insoluble in acids	29. 165000	22. 332900
Loss	0. 332500	0. 105584
Total	100. 000000	100. 000000

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[1] The quantitative analysis of the insoluble portion of Analysis XXXIII was only roughly made, but the qualitative analysis of this portion was done carefully. The bases combined with silica are the same as those combined with silica in the insoluble residue of Analysis XXXIV.

Discussion.—Both samples of chamber-dust were examined very carefully for fluorine. If this substance existed in the camp it is here that it would be concentrated; but no trace even was detected. Some estimations were made of both selenium and tellurium, but the results were considered as too high, and traces only are reported. The estimation of both these substances in such mixtures as those of the flue-dust is too complicated, when traces only are in question, to devote much time to it. However, both selenium and tellurium have been handled in small quantities, and the writer feels perfectly sure of their existence. He also feels assured that both exist in the state of selenious and tellurous acids. It is a somewhat singular fact that baryta was carried in the fumes in the state of silicate of baryta, a trace only of sulphate of baryta being found. One is almost tempted to come to the conclusion that certain silicates are volatile.

The sulphides of iron, zinc, and manganese were estimated by means of the quantity of sulphureted hydrogen evolved when the dust is treated by weak acids. No molybdenum could be detected, although it will be seen that it is constantly present in the blast furnaces. Cadmium could not be estimated by the classical methods, and it was in examining the zinc obtained in the course of the analysis that this metal was found. The formula of the oxychloride of lead found in the portion of the dust soluble in water is $3PbO, PbCl_2$, the oxybromide and iodide having the same formula.

In all probability the chloro-bromo-iodide of lead found in the portion of the fumes insoluble in water and soluble in acids exists in combination with a very large excess of oxide of lead, and also in combination with phosphate and sulphide of lead. The fumes were not examined spectroscopically for indium and thallium, so that there is some doubt about the presence of these metals. However, the writer is almost certain that he has perceived the characteristic oxysulphide of indium and observed in several instances the green flame of thallium.

The new metal which the writer was fortunate enough to observe and to trace out in all the fumes has only been seen in such minute quantities that further investigation on quite a large scale is absolutely necessary, in order to isolate it, to study its properties, and to place its existence beyond doubt. It has been possible, however, to find already three characteristic properties:

1. The oxide of the new metal gives a beautiful blue color to a bead of borax placed in the ordinary flame of the blow-pipe, and the bead becomes perfectly colorless in the reduction flame.

2. The sulphide of the new metal is slightly soluble in sulphide of ammonium, and the solution takes a characteristic blue tinge.

3. The iodide of the new metal has a fine rich pink color when in solution.

It may also be added that in one instance the sulphide of the new metal was obtained in a state of great purity, and that this sulphide was very fusible and had a deep brown color.

Although a great many substances are carried away physically in virtue of their volatility, and others mechanically by the force of the blast, it seems pretty clear from the inspection of the analyses that some very complicated reactions take place in the furnace, by means of which some substances are carried away in the state of volatile compounds and deposited in the dust in their original non-volatile form. In all probability copper, titanic acid, tin, aluminium, magnesium, and silicium are carried away in the state of chloro-bromo-iodides and sulphides formed by the action of chloro-bromo-iodide of lead and silver, and of sulphide of lead in presence of carbon, on the non-volatile oxides. The volatile chlorides and sulphides thus formed are afterwards decomposed by water, when the steam which accompanies the fumes condenses. These reactions, which might appear doubtful from the examination of the chamber-dust alone, are forcibly demonstrated by the analysis of that portion of the fumes which is not condensed and which escapes into the atmosphere. (See Analysis XXXVI.)

Roasted dust. — At one smelter the flue and chamber dusts are roasted, previous to resmelting, in the roasting furnace, which is spoken of in the description of this smelter. This operation is probably carried on with a view to getting rid of the large proportion of arsenic which is erroneously supposed to exist in the dust. Whatever may have been the object of the superintendent of this smelter in performing this costly operation, if the composition of the chamber-dust of this smelter is comparable to that of the others, the influence of roasting on chamber-dust can be easily seen by comparing Analysis XXXV of the roasted dust with Analyses XXXIII and XXXIV.

The sample analyzed in XXXV is chiefly formed of roasted dust, but it contains also a little unroasted dust, which had been spread over the roasted dust taken out of the roasting furnace. The specimens mixed for analysis were in friable whitish and reddish masses, containing scarcely any charcoal or coke dust. The elementary analysis was made like those of the chamber-dust, and the different portions analyzed are in the following proportions:

Portion soluble in water	1.16588
Portion soluble in acids	85.48129
Portion insoluble in acids	13.29600
Loss	0.05683
Total	100.00000

ANALYSIS XXXV. ROASTED CHAMBER DUST.

Elementary.

Lead	53.130100	Oxygen	8.242172
Silver	0.112000	Selenium and tellurium	Traces
Gold	0.000200	Sulphur	0.490000
Bismuth	0.051420	Chlorine	0.694840
Zinc	1.479920	Potassium	0.054590
Cadmium	0.004800	Sodium	0.051300
Iron	8.894900	Alumina	1.144400
Manganese	0.551270	Lime	0.760240
Copper	0.115600	Magnesia	0.512200
Arsenic	0.210000	Potash	0.025000
Antimony	0.092940	Bromine	0.142420
Tin	0.003478	Iodine	0.007780
Soda	0.015000	Carbon (from charcoal and coke)	1.505000
Hygroscopic water	0 225000	Graphite	0.290000
Sulphuric acid	10.929000	Indium, thallium, new metal	Traces
Phosphoric acid	0.319800	Loss	0.056830
Carbonic acid	Trace	Total	100.000000
Titanic acid	0.055800		
Silica (from slag and silicate of lead)	8.332000	Silver ounces to ton	32.666
Silica (from quartz and refractory silicates)	1.500000	Gold do	0.058332

ANALYSIS XXXV. ROASTED CHAMBER DUST — Continued,

Rational.

Portion soluble in water.	
Oxychloride of lead	0.105970
Oxybromide of lead	0.023270
Oxyiodide of lead	0.001620
Sulphate of zinc	0.016000
Sulphate of manganese	0.036100
Sulphate of lime	0.447440
Sulphate of magnesia	0.049050
Caustic magnesia	0.013650
Chloride of potassium	0.083080
Bromide of potassium	0.032870
Iodide of potassium	0.001370
Chloride of sodium	0.130460
Hygroscopic water	0.225000
Total	1.165880
Portion soluble in acids.	
Oxide of lead	22.092080
Phosphate of lead	1.826460
Sulphate of lead	37.443000
Sulphide of lead	3.676220
Chloride of lead	2.105760
Bromide of lead	0.236140
Iodide of lead	0.009230
Oxide of antimony	0.117000
Oxide of tin	0.004300
Oxide of copper	0.144800
Oxide of zinc	1.810000
Sulphate of lime	1.224000
Caustic magnesia	0.447200
Titanic acid	0.037800
Oxide of bismuth	0.057300
Chloride of silver	0.122640
Portion soluble in acids—Continued.	
Bromide of silver	0.032270
Iodide of silver	0.002390
Oxide of cadmium	0.005500
Gold	0.000200
Arsenious acid	0.277000
Peroxide of iron	12.020000
Oxide of manganese (Mn_3O_4)	0.746000
Alumina	1.044000
Carbonate of lime	Trace
Selenious and tellurous acids	Traces
Oxides of indium, thallium, new metal	Traces
Total	85.481290
Portion insoluble in acids.	
Silica (combined)	8.332000
Silica (quartz)	1.500000
Oxide of lead	0.690000
Arsenious acid	0.000200
Oxide of antimony	0.000300
Oxide of tin	0.000100
Titanic acid	0.018000
Peroxide of iron	0.687000
Alumina	0.100400
Lime	0.072000
Magnesia	0.035000
Oxide of zinc	0.026000
Oxide of magnesia	Trace
Potash	0.025000
Soda	0.015000
Carbon (from charcoal and coke)	1.505000
Graphite	0.290000
Total	13.296000

Discussion.—Everything indicates that the dust was roasted at a very high temperature, and this is proved beyond doubt by the fact that traces only of carbonic acid are detected.

On comparing this analysis with the average analysis (XXXIV) of unroasted chamber-dust, it is found that the percentage of lead is considerably increased, but that a considerable quantity of silver is lost. One notices also that the quantity of chlorine, bromine, and iodine is about half what it was in the average sample of dust. Coupling this with the loss of silver, it may justly be inferred that silver is lost in the state of chloro-bromo-iodide, and that some lead is also lost in the same form. The quantity of phosphate of lead has also diminished instead of increasing, showing that lead is also lost in this form. The percentage of arsenic and antimony is lower than in the unroasted dust, but this appears to be the only advantage gained by roasting, and a very slight one it is in these cases. Sulphur, instead of being driven off by roasting, is concentrated in the form of sulphate of lead, amounting to 37 per cent., and representing 5 per cent. of sulphur instead of 2 per cent., as in the chamber-dust. Lastly, about 9 per cent. of carbon is driven off at great expense. This carbon is so intimately mixed with the original dust that by simple heating in the blast furnace there would be more than enough of it to reduce all the lead of the fumes. In fact, everything indicates that roasting chamber-dust in Leadville is a useless and a ruinous operation, by which nothing is gained and by which many valuable substances, lead, silver, fuel, etc., are lost, at great expense. It is only fair to state once more that roasting of chamber dust has only been done at one smelter. The analyses show that the mischievous substances of chamber-dust are not arsenic and antimony, but chlorine, bromine, iodine, and phosphoric acid. Analysis XXXVI, of the portion of lead fumes lost in the atmosphere, will further demonstrate this statement. *A priori*, the best way of treating lead dust and fumes is to resmelt them with an excess of lime, in order to fix the volatile constituents by combination with calcium and lime. This is precisely what is done in practice at most of the smelters, the only improvement that might be suggested being the use of pure lime instead of the magnesian lime of Leadville; but the writer thinks the general use of caustic lime instead of limestone would be of great advantage to smelting at Leadville, since the chloro-bromo-iodo phosphates of lead are driven off in the state of fumes in the furnace, before limestone, by the loss of carbonic acid, has become caustic and thus acquired the power of acting chemically on the fumes with which it comes in contact.

Lead lost in fumes.—At the time the writer was collecting notes and specimens for this report, it happened fortunately that experiments were being made with the Bartlett smoke-filter (described on page 673), at the Grant smelter, for the purpose of condensing that portion of the lead fumes which escapes with the smoke into the atmosphere. A fine opportunity was thus had of making a thorough investigation of that part of the smelting products which is always lost at all the smelters. The analysis of these fumes proves to be the most interesting by far of all those made on lead dusts, since they are not only richer in lead than any of the others and contain more silver than the average Leadville slag, but they show also other remarkable peculiarities, as an inspection of the following analysis will show. They have exactly the appearance of soot or lampblack.

ANALYSIS XXXVI. FUMES FROM BARTLETT FILTER.

Elementary.

Lead	60.370150	Potassium	0.141300
Silver	0.014750	Sodium	0.163600
Gold	Faint trace	Arsenic	1.113672
Zinc	4.638410	Antimony	0.208900
Cadmium	0.001310	Tin	0.160370
Iron	1.227660	Lime	2.756990
Manganese	0.311420	Magnesia	1.856690
Aluminium	0.060600	Alumina	0.421000
Calcium	0.124000	Potash	0.007000
Magnesium	0.174600	Chlorine	3.953590
Soda	0.013000	Bromine	0.798230
Water	0.912500	Iodine	0.043630
Carbonic acid	3.057510	Carbon (soot)[a]	3.365000
Phosphoric acid	2.038400	Indium, thallium, new metal	Traces
Sulphuric acid	0.060000	Loss	0.009200
Titanic acid	0.000500	Total	100.000000
Silica	3.279300		
Oxygen	3.393178	Silver ... ounces to ton	4.302
Sulphur	5.320540	Gold ... do	Trace only
Selenium and tellurium	Traces		

[a] The soot contains also a trace of graphite.

ANALYSIS XXXVI. FUMES FROM BARTLETT FILTER—Continued.

Rational.

Portion soluble in water.	
Oxychloride of lead	3.203400
Oxybromide of lead	0.704600
Oxyiodide of lead	0.348340
Protosulphate of iron	0.057000
Protochloride of iron	0.537200
Protobromide of iron	0.184300
Protoiodide of iron	0.012400
Sulphate of zinc	0.060300
Chloride of zinc	0.623200
Bromide of zinc	0.176000
Iodide of zinc	0.007660
Sulphate of manganese	Trace
Sulphate of alumina	Trace
Chloride of aluminium	0.446600
Chloride of calcium	0.286380
Bromide of calcium	0.104000
Chloride of magnesium	0.680670
Bromide of magnesium	0.020240
Chloride of potassium	0.269930
Chloride of sodium	0.423930
Water	0.912500
Total	9.058650
Portion soluble in acids.	
Oxide of lead	28.176900
Phosphate of lead	11.641800
Sulphide of lead	18.899160
Chloride of lead	6.341150
Bromide of lead	0.750000
Iodide of lead	0.032460
Chloride of silver	0.017890
Bromide of silver	0.002060
Iodide of silver	0.000150
Gold	Faint trace
Portion soluble in acids—Continued.	
Protosulphide of iron	1.406680
Sulphide of zinc	6.357000
Sulphide of manganese	0.491960
Sulphide of cadmium	0.001500
Arsenious acid	1.470000
Protoxide of antimony	0.250000
Oxide of tin (SnO_2)	0.203850
Selenious and tellurous acids	Traces
Alumina	0.220200
Carbonate of lime	4.853200
Carbonate of magnesia	1.760300
Caustic magnesia	1.008490
Indium, thallium, new metal	Traces
Total	83.884750
Portion insoluble in acids.	
Silica	3.279300
Titanic acid	0.000500
Oxide of lead	0.142800
Oxide of zinc	0.004800
Peroxide of iron	0.034000
Alumina	0.200800
Lime	0.039200
Magnesia	0.010000
Potash	0.007000
Soda	0.013000
MnO_2, AsO_2, SnO_2, Sb_2O_4	Traces
Carbon (soot)	3.365000
Total	7.096400
Portion soluble in water	9.058650
Portion soluble in acids	83.884750
Portion insoluble in acids	7.096400
Total	100.039800

Discussion.—The combinations in which lead exists in these fumes and their volatility is most remarkable. To appreciate this fully one must remember that the fumes were filtered at a distance of nearly 200 feet from the blast furnace. The above analysis shows, besides 9 per cent. of chloro-bromo-iodide of lead, 18 per cent. of sulphides in combination with the chloro bromo-iodides, and 11 to 12 per cent. of phosphate of lead, proving a most surprising volatility for the two latter combinations. The percentage of lead as oxide is relatively low as compared with other dusts. It is most worthy of remark that the iron, manganese, zinc, and cadmium must be in the state of sulphides and also have an unusual degree of volatility. That these fumes should be richer in arsenious acid than the ordinary chamber dust might have been expected, but that they should have carried off tin and titanic acid is most remarkable. Silica being combined with the lead as silicate of lead shows that this combination is also volatile.

The low barometric pressure at Leadville, which is, on an average, nine inches to ten inches of mercury less than that at sea level, explains in a great measure the extraordinary volatilization of so many products; but the fact that non-volatile substances are carried away in the state of volatile compounds, and, so to speak, reprecipitated in a non-volatile form, is abundantly proved by the fact that this smoke contains chloro-bromo-iodides of calcium, magnesium, and aluminium, and still more by the fact that it contains more tin than any other. Most surprising of all is the small percentage of oxygen and the large percentage of sulphur, as compared with the amounts of both substances in flue and chamber dusts.

SPEISS.

The speiss formed in the blast furnaces of Leadville belongs to three types: 1. The white metallic-looking speiss, in large lamellar crystals, studded all over with very small, indistinct crystals. 2. The grayish sub-metallic looking speiss, in fine crystalline grains. 3. The vesicular speiss.

It will be seen that iron sows belong also to the speiss family, being, so to speak, embryonic speiss.

A specimen of type No. 1, taken from a cake at Messrs. Billing & Eilers's smelter and made from dolomitic smelting charges, contained only a few grains of metallic lead and none of metallic iron. The lead grains were separated by the sieve, and the speiss powder being analysed gave the composition reported in Analysis XXXVII. This speiss has such a characteristic appearance that it may be assumed that similar speiss found at the other smelters possesses the same composition.

No. XXXVIII is the analysis of a sample representing type No. 2, composed of equal parts of specimens taken from different smelters, as indicated below:

Smelter.	No. of specimens for analysis.
American	3
Cumming & Finn	1
Elgin	1
Grant	1
La Plata	1
Little Chief	1
Ohio and Missouri	2
Billing & Eilers	1
Total number of specimens mixed for analysis	11

By sifting, the speiss was separated into—

Speiss powder	98.21
Lead grains	1.22
Iron grains	0.57
Total	100.00

The non-combined iron grains did not contain any arsenic and were very tough; the lead was also very pure. The speiss powder only was analysed.

The sample of vesicular speiss was made up of equal parts of nine specimens, collected, three at the American smelter, two at the Harrison, four at the Ohio and Missouri.

By sifting, it was separated into—

Speiss powder	90.17
Iron grains	9.83
Lead grains	Trace
Total	100.00

The iron grains are also perfectly free from arsenic and very tough. The speiss powder of this sample was not analyzed, as in all probability its composition is the same as that of type No. 2; but it was assayed for gold and silver. (See speiss assay No. 5.)

Conclusions.—The sieve examination of speiss shows: 1. That speiss type No. 1 contains no free iron and is a non-saturated speiss. 2. That speiss type No. 2 contains just enough iron in excess to indicate that it is saturated with iron. 3. Lastly, vesicular speiss type No. 3 contains a very large excess of free iron. It is a supersaturated speiss, whose fusion has been prevented by this excess of infusible iron.

The writer made three comparative experiments on the fusibility of the three kinds of speiss. In each case the pulverized speiss was mixed with borax in a porcelain crucible, heated over the blast-lamp, properly regulated to operate as nearly as possible in the same conditions. Speiss types No. 1 and No. 2 melted easily, and no appreciable difference could be detected in their melting point. Speiss type No. 3 was melted with more difficulty than the preceding and formed a vesicular button, showing incomplete fusion.

ANALYSES XXXVII AND XXXVIII. SPEISS.

	XXXVII.	XXXVIII.
Sulphur	5.8191	4.4695
Arsenic	31.4725	21.8003
Antimony	Trace	0.1450
Iron	60.5780	70.4780
Zinc	Faint trace	Trace
Silver	0.0085	0.0301
Gold	Faint trace	0.0009
Lead	1.4935	2.5030
Copper	0.3628	0.2566
Nickel	0.0876	0.0981
Molybdenum	0.2110	0.2155
Loss		0.0030
Total	100.0330	100.0000
Silver ounces to ton	2.48	8.7822
Gold do	Trace	0.26

Discussion.—The formula of the large crystalline speiss analyzed in XXXVII is $(FeM)_6AsS$, M designating the small quantities of metals accompanying iron, and the rational formula is probably represented by $Fe_5As(FeM)S$.

The fine-grained speiss analyzed in XXXVIII is represented by the formula $(FeM)_9AsS$, and probably by the rational formula $Fe_8As(FeM)S$.

The powder of the vesicular speiss was not analyzed, as has been stated before, but it is probable that its composition and formula are the same as that of the preceding, since both are saturated speiss and contain an excess of non-combined iron. However, both cannot be considered as mere mixtures of definite arsenio-sulphurets of iron and iron, for both the iron and the arsenio-sulphuret are crystalline, and they are in all probability crystallographic compounds.[1] It has been seen that slags were similarly crystallographic compounds. It will be seen that mattes also are crystallographic compounds. A few remarks will be made with regard to slag-mattes, which form the highest expression of this class of substances, so that in almost every stage of the smelting operations crystallographic compounds are the regular products of the blast furnaces.

The speisses analyzed are remarkable on account of—

1. The presence of antimony in such very small quantities, whilst it exists in the smelting charges, and is formed in no inconsiderable quantity in the bullion, the fumes, etc.

2. The presence of molybdenum in each and sensibly in the same amount, showing how widely and evenly distributed this metal is in the camp. In the writer's opinion it is the first time molybdenum has been pointed out in speiss. It is so thoroughly concentrated in this product that it was not possible to detect it in either bullion, slag, or fumes.

3. The total absence of cobalt, which, as has been observed, is concentrated in the skimmings of the lead of the siphon-tap, and thus thoroughly separated from nickel, which, as the preceding analyses show, remains in the speiss.

Assays of speiss for gold and silver.—In order to complete the study of the Leadville speiss, the following assays were made of speiss belonging to the three different types in the Survey laboratory:

TABLE XIII.—*Assays of speiss.*

No.	Smelters.	Nature of speiss.	No. of specimens mixed for assay.	Silver	Gold.
				Oz. to ton.	*Oz. to ton.*
1	Billing & Eilers	Type No. 1.	1	2.48	Trace
2	do		3	11.95	Trace
3	Cumming & Finn		3	15.25	0.066
	Elgin		3		
	La Plata		3		
	Average assay of type, No. 1	....	....	9.893	0.022
4	Little Chief	Type No. 2	1	8.7822	0.26
	La Plata		1		
	Grant		1		
	American		3		
	Billing & Eilers		1		
	Cumming & Finn		1		
	Elgin		1		
	Ohio and Missouri		2		
5	American	Type No. 3	3	40.655	0.0479
	Harrison		2		
	Ohio and Missouri		4		

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[1] Exactly what Mr. Guyard means by the term *crystallographic compounds* I am not able to state. (S. F. E.)

To complete this examination the powder and metallic grains of vesicular speiss No. 5 were tried separately, with the following results:

	Silver.	Gold.
	Oz. to ton.	*Oz. to ton.*
Speiss powder...	43.705	0.045
Metallic grains...	12.675	0.075

Discussion.—It has been seen that types No. 1 and No. 2 contain sensibly the same amount of precious metals, but why type No. 3 should contain such a large quantity of silver as 40 ounces the writer is not prepared to explain. The fact that the silver is contained, not in the metallic grains, as might have been supposed, but in the speiss itself, renders the explanation more difficult yet.

Speiss No. 2 (type No. 1), from Messrs. Billing & Eilers's smelter, was made in connection with dolomitic slag. The difference in the contents of silver is very large, and is probably due to a corresponding difference in the assay of the bullion extracted at the same time.

IRON SOWS OR SALAMANDERS.

The sows, or small masses of reduced metallic iron, which are formed occasionally in the blast-furnaces, on being analyzed prove to be a variety of speiss. The writer expected some important results from their examination, for the reason that metallic iron reduced in the blast-furnaces becomes the center of the most important chemical reactions, and that here was an opportunity of observing the various components of the smelting charges in the very act of combination.

The sample submitted to analysis was prepared from a great variety of specimens—two from the Little Chief, three from Gage, Hagaman & Co., and four from the California smelter.

The specimens from the California smelter were taken from three sows of very tough metal, and measuring 1 foot 6 inches by 10 inches by 1 foot, and also 1 foot 6 inches by 10 inches by 5 inches. One of these sows was impregnated with the peculiar slag which has been described under the name of scoria. One of the specimens of reduced iron was extracted in small bits from the very midst of a slag, where it was found side by side with small bits of speiss, which were perfectly fused, but which had not had time to collect together. A similar specimen was extracted from a slag found at the Little Chief smelter. All the specimens of sows were full of large cavities, filled with charcoal, coke, slag, scoria, and regular speiss. The sample analyzed was prepared by pounding bits detached from the sows until no dust could be obtained, so that it is the toughest portion of the sows that has been analyzed. The metal forming the sows is sometimes rather brittle, but is also often very tough. Tough or brittle, however, when broken it exhibits a decidedly crystalline structure, the crystals being almost large enough to be measured, and the metal itself is perfectly white and bright.

ANALYSIS XXXIX. IRON SOW.

Iron	72.82830	Silver	0.11492
Lead	18.79340	Gold	0.00003
Arsenic	5.08330	Phosphorus	0.10905
Nickel	0.04500	Graphite	0.75000
Cobalt	Faintest trace	Combined carbon	0.55000
Manganese	0.01500	Silicium, slag, loss	0.90000
Antimony	Faint trace	Total	100.00000
Copper	Faint trace		
Zinc	Faintest trace	Silver ounces to ton	33.51750
Molybdenum	0.16100	Gold do	[a]0.00875
Sulphur	0.65000		

[a] In order to obtain a fair average assay of silver and gold the assay was made on 200 grams, which explains why gold is given to the $\frac{1}{100000}$ of an ounce to the ton.

Analysis XXXIX shows that the sows are an alloy of iron and lead combined with a quantity of arsenic sufficient to class them with speiss. Like the latter, sows contain nickel, molybdenum, and sulphur. Their graphite is liberated when their iron enters into combination with larger quantities of sulphur and arsenic, and is blown away with the lead dust where it has been found. Like supersaturated speiss, sows are very rich in silver. The fact that they contain traces of cobalt, which cannot be detected in speiss, shows that this metal is forcibly taken away by virtue of a reaction which is not yet understood, although one is inclined to presume that sulphide of cobalt is formed, and that it combines more readily with sulphide of lead than with arsenide of iron.

MATTES.

The examination of the mattes of Leadville proved so surprising to the writer that he accumulated all possible proofs in order to have no doubts as to their nature. It will be seen that they are formed of sulphides of iron and lead and magnetic oxide of iron. It was in order to demonstrate the presence of this oxide that experiments and reactions were varied in every possible way, which were discarded afterwards, and became useless when the magnetic oxide of iron was successfully isolated in the pure crystalline state. Thus mattes are formed of sulphides and oxides, and, like most furnace products, they are *crystallographic combinations of chemical compounds.*

A careful examination was made of the two typical kinds of mattes of Leadville. The so-called iron matte, with a fine crystalline structure and a brown luster, was collected at Messrs. Cumming & Finn's smelter (Analysis XL). The so called lead matte, blackish gray, with a decided crystalline structure, was found at the Elgin smelter (Analysis XLI). Both specimens were slightly swollen and split by a beginning of oxidation. Both yield strongly magnetic particles to the magnet, which were at first mistaken for magnetic sulphides. By rubbing the magnetic portion on filter paper, isolating with the magnet, repeating the operation several times, and finally treating it with nitric acid, in order to destroy some sulphides which adhere to it, the pure magnetic oxide is obtained. Whether the adhering sulphides are magnetic or not could not be decided. The pure magnetic oxide was not analyzed, but there is but little doubt that its formula is the same as that of the oxide isolated from slags, namely, Fe_7O_9.

ANALYSES XL AND XLI. MATTES.

Elementary.

	XL.	XLI.
Portion soluble in water.		
Sulphuric acid	0. 037895	0. 030
Protoxide of iron	0. 034105	0. 028
Water	Trace	Trace
Portion soluble in acids.		
Sulphur	24. 257340	19. 350
Oxygen	4. 500000	6. 230
Iron	48. 906300	40. 430
Nickel	None	0. 067
Lead	20. 100000	31. 970
Silver	0. 291666	0. 240
Portion soluble in acids.—Cont'd.		
Gold	Faint trace	Faint trace
Zinc	1. 143700	0. 220
Copper	0. 278277	0. 855
Arsenic	0. 219906	0. 131
Antimony	Trace	0. 270
Portion insoluble in acids.		
Slag	0. 137000	0. 140
Loss	0. 093811	0. 039
Total	100. 000000	100. 000
Silver...........ounces to ton	85. 067	69. 9984

Rational.

	XL.	XLI.
Portion soluble in water.		
Protosulphate of iron	0. 072000	9. 058
Portion soluble in acids.		
Protosulphide of iron	56. 151956	37. 446
Sulphide of lead	23. 192307	36. 912
Magnetic oxide of iron	16. 312500	22. 826
Sulphide of zinc	1. 706700	0. 330
Sulphide of nickel	None	0. 102
Sulphide of copper	0. 347846	1. 282
Sulphide of arsenic	0. 290275	0. 214
Portion soluble in acids.—Cont'd.		
Sulphide of antimony	Trace	0. 376
Sulphide of silver	0. 334875	0. 275
Sulphide of gold	Trace	Trace
Iron (combined with FeS_2)	1. 360730	Trace
Portion insoluble in acids.		
Slag	0. 137000	0. 140
Loss	0. 093811	0. 039
Total	100. 000000	100. 000

The writer has described in the Bulletin de la Société Chimique de Paris a peculiar mineral formed of arsenio sulphuret of nickel and chrome-iron oxide, which certainly presents a great analogy with the mattes.

Lead and iron mattes are not the only ones which form in the furnace. A third one, which is much more interesting, may be called the calcium matte. This matte is formed, like its congeners, of a sulphide, sulphide of calcium, and of magnetic oxide of iron, crystallographically combined. This matte has not been found in an isolated state, but it exists in combination with scoriæ, and the product thus formed is precisely the slag of Leadville. So that the best definition of slags that can be given is the following: Slags are compounds of scoriæ or silicates and of calcium mattes, and, like most of the furnace products, they are formed of chemical compounds crystallographically combined.[1]

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[1] Although Mr. Guyard's definition may perhaps seem somewhat obscure to the reader, I do not feel sufficiently certain of his meaning to attempt to modify it, and therefore leave it in his own words. (S. F. E.)

Assays of mattes. — Crude mattes generally contain variable quantities of metallic grains, so that whenever it was thought advisable the crude matte was assayed, as well as the powder and metallic grains separated by the sieve. The following assays were made in the laboratory of the Survey.

TABLE XIV. ASSAYS OF MATTES.

Matte.	No. 1.	No. 2.	No. 3.	No. 4.
Separates into—				
Powder	51.37	98.62	79.32	97.5
Metallic grains	48.63	1.38	20.68	2.5
	100.00	100.00	100.00	100.00
Silver contents.				
Crude matte . . . . ounces per ton	102.7344	Not assayed	99.0256	Not assayed
Powder . . . . do	45.47	74.35	86.63	66.25
Metallic grains . . . . do	163.22	Not assayed	146.6	Not assayed
Gold contents.				
Crude matte . . . . ounces per ton	0.02743	Not assayed	0.03	Not assayed
Powder . . . . do	0.025	Faint trace	0.03	Trace
Metallic grains . . . . do	0.030	Not assayed	0.03	Not assayed

No. 1. American smelter, lead matte.
No. 2. Billing & Eilers's smelter, lead matte, blackish gray.
No. 3. Cumming & Finn's smelter, lead and iron matte.
No. 4. Elgin smelter, lead matte, grayish.

Discussion. — It has been seen (Analysis XL) that the normal iron matte contains 85 ounces of silver, and that (assay of mattes No. 3) the powder of a similar matte from the same smelter contains 86½ ounces. The normal lead matte (Analysis XLI) contains 70 ounces of silver, and the powder from similar mattes at the same and at other smelters contains 66.25 ounces of silver (No. 4), 45.5 ounces (No. 1), 74.35 ounces (No. 2).

The relation between the contents of bullion (represented by the metallic grains) in silver and the contents of the mattes in silver is rather peculiar. In No. 1, the richest bullion, 163 ounces, corresponds to the poorest matte; and in No. 3, the poorest bullion, 146½ ounces, accompanies the richest matte, 86 ounces. Evidently there is an antagonism between matte and bullion, in which the latter has not always the advantage.

ACCRETIONS.

There are two kinds of accretions formed in the blast furnace which have nothing in common: the hearth accretions and the shaft accretions.

Hearth accretions. — Hearth accretions are, in general, very similar to mattes, but sometimes they look more like slags. From a very rough examination of these products, it results that they are chiefly formed of slag, and lead- and iron-matte, and it is thought that the word *slag-matte* is the most appropriate to designate them. They are formed of variable quantities of slag and matte, but often contain nearly equal parts of both, and represent in the highest degree those singular compounds, crystallographically combined, which have been described so often in this report.

The writer has not pursued further the investigations of these hearth accretions for the reason that the elements which constitute them have already been fully examined and discussed. The following assays were made however, in the same manner as those of the mattes:

TABLE XV. ASSAYS OF HEARTH ACCRETIONS FROM GRANT SMELTER.

Accretions.	No. 1.	No. 2.
Separates into—		
Slag-matte powder	77.51	63.06
Metallic grains	22.49	36.94
	100.00	100.00
Silver contents.		
Crude hearth accretion .....ounces per ton.	47.6916	81.6473
Slag-matte powder ....do.....	28.8	16.125
Metallic grains ....do......	112.8	193.50
Gold contents.		
Crude hearth accretion .....ounces per ton.	Trace	Trace
Slag-matte powder ....do......	Trace	Trace
Metallic grains ....do......	Trace	Trace

Discussion.—The assay of silver in the slag-matte indicates to some extent the relative proportion of slag and matte in the hearth accretion, it being evident that No. 1 contains more matte than No. 2. The fact that the poorest bullion, No. 1, corresponds to the richest accretion, and the richest bullion, No. 2, to the poorest accretion, confirms the similar observation made on the mattes, and seems to indicate that bullion is deprived of its silver by the matte.

Shaft accretions.— As has already been stated, the shaft accretions have nothing in common with hearth accretions. Shaft accretions generally result from the condensation of sublimated products. They form thick incrustations against the lower parts of the walls of the shaft of blast-furnaces, and occasionally line the whole of the shaft. At Gage, Hagaman & Co.'s the writer has seen a small round furnace entirely lined, from the top of the water-jackets to within six inches of the feed-hole, with accretions a foot thick. A very complete collection of those products was made, as it was expected that in them would be found a great concentration of the metals which occur in minute quantities in the ores.

Before describing normal accretions the writer would say a few words concerning some pretty yellowish semi-translucid crystals of chloro-bromide of lead found by Dr. M. W. Iles in one of the furnaces at Grant smelter, between the main cast-iron plate support of the furnace and the masonry. These crystals were analyzed by Dr. Iles and found to contain—

Chlorine	10.345
Bromine	25.321
Lead	63.927
	99.593

A small quantity of the crystals were kindly forwarded to the writer by Dr. Iles, and were examined qualitatively; in these were found, besides chloride and bromide of lead, a small quantity of iodide of lead. They appeared, moreover, to have the same composition as the chloro-bromo-iodide of lead which was found by the writer in the lead dust; in which case they would contain more chlorine than bromine and more bromine than iodine; but being crystallized it is very possible that they have the composition assigned to them by Dr. M. W. Iles and the formula Pb Br Cl. These crystals form quite an accidental accretion, and are found only in very small quantity.

The normal accretion, the one found in every smelter and in every furnace, is a light crystalline fibrous and porous mass with a luster like galena and full of cavities, spotted with a whitish-yellow and reddish crust. The fibrous portion is formed of alternate layers of galena-like and of yellowish-white crystals. No metallic grains can be seen in the mass of the accretion, even when broken into very small bits; but, while some parts are quite brittle, others have a toughness which is due to the presence of metallic lead. A similar accretion, collected at Messrs. Billing & Eilers's smelter, was separated by the sieve into powder, 81.25; lead grains, 18.75.

The whole accretion, powder and grains, was analyzed in the above proportion and found to contain, among other substances, an enormous quantity of metallic lead; more than twice as much as the quantity represented by the lead grains. This lead is evidently condensed vapor of this metal or sublimated lead, so intimately mixed with the sulphides that it passes through the sieve.

Another point of interest is the large percentage of sulphide of zinc found, which forms the yellowish white crust and layers visible in the accretion already mentioned.

The following is the analysis of a typical shaft accretion collected at Messrs. Billing & Eilers's smelter.

ANALYSIS XLII. SHAFT ACCRETION.

Portion soluble in acids

Elementary analysis.		Rational analysis.	
	Per cent.		*Per cent.*
Sulphur	8. 2912730	Sulphide of lead	7. 2049800
Lead	47. 4912340	Sulphide of bismuth	Trace
Bismuth	Trace	Sulphide of silver	0. 0866208
Silver	0. 0754498	Sulphide of gold	Trace
Gold	Trace only	Sulphide of copper	Trace
Copper	Trace	Sulphide of zinc	10. 4124000
Zinc	6. 9774000	Sulphide of cadmium	Marked trace
Cadmium	Marked trace	Protosulphide of iron	10. 6129000
Iron	6. 7536000	Sulphide of arsenic	0. 0644009
Arsenic	0. 0392689	Sulphide of antimony	Trace
Antimony	Trace	Metallic lead	41. 2469180
Sulphuric acid	0. 9653590	Sulphate of lime	1. 6411103
Phosphoric acid	2. 1663252	Phosphate of lime	4. 6937046
Carbonic acid	Trace	Carbonate of lime	Trace
Lime	4. 2616000	Caustic lime	1. 0584693
Magnesia	3. 2021540	Caustic magnesia	3. 2021540
Oxide of manganese	2. 8871860	Oxide of manganese	2. 8871860
Alumina	0. 0720000	Alumina	0. 0720000

ANALYSIS XLII. SHAFT ACCRETION—Continued.

Portion insoluble in acids.

Elementary analysis.	*Per cent.*	Rational analysis.	*Per cent.*
Silica	10.1000000	Silica	10.1000000
Oxide of lead	0.4050000	Oxide of lead	0.4050000
Oxide of zinc	0.3000000	Oxide of zinc	0.3000000
Oxide of iron	1.1000000	Oxide of iron	1.1000000
Oxide of manganese	Trace	Oxide of manganese	Trace
Alumina	1.6000000	Alumina	1.6000000
Lime	2.1000000	Lime	2.1000000
Magnesia	1.0950000	Magnesia	1.0950000
Loss	0.1171501	Loss	0.1171501
Total	100.0000000	Total	100.0000000

Silver, 22 ounces to the ton; gold, trace only.

Discussion.—The points of interest in this analysis are the concentration of phosphoric acid in the form of phosphate of lime, the presence of caustic lime, and the presence of 22.3 per cent. of metallic lead in an impalpable form. A glance at the whole analysis shows the close resemblance between accretions and chamber-dust, the former representing products of sublimation, the latter products of volatilization.

Assays of normal accretions.—The writer prepared a sample of accretions identical in appearance with the one reported above and made up of twelve specimens from the following smelters:

Cumming & Finn	7
Grant	1
Harrison	1
Gage, Hagaman & Co	3
Total number of specimens mixed for assay	12

The powder and grains in this mixture were separated, as usual, by the sieve.

The whole accretion assayed 21.1092 ounces of silver to the ton. The specimen analyzed assayed 22 ounces, showing a remarkable uniformity in the composition of these products from various sources. The accretion powder assayed 21 ounces and the accretion grains 23.6 ounces of silver to the ton. This last figure is very interesting, as showing the contents in silver of volatilized bullion. There can be no doubt about this, since the accretions assayed come precisely from smelters which run the richest and the poorest bullion; besides, if we consult the assays of lead grains found in other products (the mattes and hearth accretions, for instance), we find that in no instance do lead grains contain so little as 23 ounces of silver.

In connection with the normal accretions just described will be given the assay of common accretions, which differ entirely from the preceding. Instead of being light and porous, they are very heavy and compact, and consist either of galena which has escaped reduction or of artificial galena formed in the furnace by the reduction of sulphate of lead. Whatever their origin, they are formed of galena. The sample assayed was prepared from five specimens from the following smelters:

Cumming & Finn	2
Grant	2
Raymond, Sherman & McKay	1
Total	5

The whole accretion assayed 45 ounces of silver to the ton. It was separated by the sieve into—

Powder	91.25
Metallic lead grains	8.75
Total	100.00

The powder assayed 35 ounces and the grains 148.5 ounces to the ton.

When mattes and accretions are not thrown away pell-mell with the slags, they are always roasted in heaps in Leadville. A mixture of such roasted mattes and accretions from Billing & Eilers's smelter was assayed; it consisted of two fine specimens, yellowish-white and reddish on the surface, full of large cavities, and having a blackish fracture; free sulphur was visible throughout the mass. It was separated by the sieve into—

Powder	97.96
Grains	2.04
Total	100.00

The powder assayed 8.25 ounces of silver to the ton.[1]

If we consider that the products roasted assayed 22.45 ounces (accretions), 102, 74.5, 99, and 66 ounces (mattes), respectively, we may well ask, what becomes of the silver in this ruinous operation of roasting in heaps performed at several smelters? A definite answer is difficult, since if silver is volatilized it is not known in what form; but it seems probable that, since the roasting takes place in the open on the slag-heap and the roasting heap is periodically leached by rain, silver may be carried away in the state of sulphate.

Peculiar accretions.—The following analysis has been made of a very peculiar accretion found at Messrs. Cumming & Finn's smelter. This accretion is half black and half yellow and looks like a mixture of galena and orpiment. It was separated by the sieve into—

Powder	75.54
Lead grains	24.46
Total	100.00

The whole accretion, powder and grain, was analyzed (see Analysis XLIII), and more interesting results than those obtained could scarcely have been anticipated. The bright-yellow portion proved to be a peculiar Naples yellow formed of arsenio-antimonio-stannate of lead; and the whole of the constituents of that portion were in the state of oxides. Here we find a most remarkable instance of concentration, that of tin, which apparently exists only in traces throughout the camp, and that of anti-

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[1] All the assays of accretions were made in the laboratory of the Survey.

mony and arsenic, which exist only in small quantities in the ores. This accretion contains tin, 5.6 per cent.; antimony, 2.7 per cent.; and arsenic, 5 per cent.

The behavior of this accretion with reagents is noteworthy: treated by weak nitric acid the metallic grains and sulphides, as well as the excess of the oxide of lead, are dissolved and the yellowish compound is left untouched. This residue, treated by strong nitric acid, yields arsenite of lead, which dissolves in the state of arseniate. The residue, which is still yellow, is decomposed by sulphide of ammonium, which dissolves tin. The residue, treated by weak nitric acid, in order to dissolve the sulphide of lead formed, leaves behind a yellow powder, which proves to be antimoniate of lead. The separations in this instance are so remarkable that the relative proportions of arsenite, antimoniate, and stannate of lead forming the yellow compound could in this way be roughly estimated.

ANALYSIS XLIII. PECULIAR ACCRETION.

Elementary analysis.		Rational analysis.	
Lead	70.63119	Metallic grains:	
Silver	0.29700	Lead	23.72921
Gold	0.00075	Silver	0.19800
Arsenic	5.00925	Arsenic	0.02500
Antimony	2.69686	Antimony	0.05500
Tin	5.59344	Tin	0.23100
Iron	0.02800	Sulphide of lead	0.21913
Zinc	Trace	Sulphides—	
Sulphur	2.60067	Sulphide of iron	0.04400
Oxygen	6.25617	Sulphide of silver	0.11366
Peroxide of iron	3.45600	Sulphide of lead	18.96927
Alumina	0.60000	Oxide of lead compounds—	
Silica	2.83000	Oxide of lead	32.62459
Loss	0.00067	Oxide of zinc	Trace
Total	100.00000	Arsenious acid	6.57900
Silver, ounces to ton	86.623	Antimonic acid	3.50806
Gold, do	0.21875	Stannic acid	6.81666
		Oxide of gold	0.00075
		Silicic acid	2.83000
		Alumina	0.60000
		Peroxide of iron	3.45600
		Loss	0.00067
		Total	100.00000

Discussion.—The rational report may be expressed under the simple form:

Metallic grains	24.45734
Metallic sulphides	19.12693
Oxide of lead compounds	56.41506
Loss	0.00067
Total	100.00000

The writer has found that silicic acid, alumina, and peroxide of iron were combined with oxide of lead in the state of silicate, aluminate, and ferrite; as for gold, it is a well-known fact that this metal is oxidized by oxide of tin and combines with it.

Peculiar accretion found at Gage, Hagaman & Co.'s smelter.—This accretion was a thin, compact, yellowish-green mass of alternate layers of sulphide of lead and yellowish-green oxides; it contained only a very few grains of lead in fine particles. This accretion is remarkable on account of the enormous quantity of zinc concentrated in it (53 per cent.), chiefly in the state of oxide (65.5 per cent.). It is also the only accretion in which were found traces of chlorine, bromine, and iodine. The rational report of the analysis has been arranged so as to show as clearly as possible how the different substances are mixed or combined. Sulphide of zinc was estimated by means of the sulphureted hydrogen evolved on treating the accretion with weak sulphuric acid. No iron was present in the solution.

Traces of the silver reported as chloride exist in the state of sulphide.

ANALYSIS XLIV. PECULIAR ACCRETIONS.

Elementary analysis.	Per cent.	Rational analysis.	Per cent.
Lead	25.9530	Sulphide of lead	17.8005
Zinc	53.3294	Sulphide of zinc	1.0670
Silver	0.0944	Mixture—	
Chlorine with traces Br, I	0.0350	Oxide of zinc	65.5550
Sulphuric acid	0.2866	Sulphate of lead	1.0891
Arsenious acid	0.0715	Oxide of lead	10.5341
Antimonious acid	0.0565	Arsenious acid .. in combination	0.0715
Oxygen	13.7513	Antimonious acid	0.0565
Sulphur	2.7254	Chloride of lead, with traces PbBr, PbI	Traces
Peroxide of iron	1.3815	Chloride of silver, with traces AgBr, AgI	0.1294
Alumina	0.3285	Peroxide of iron (mixed, not combined)	1.3815
Lime	0.2007	Alumina	0.3285
Magnesia	0.1588	Lime .... partly combined with SiO_2	0.2007
Silica	1.5770	Magnesia	0.1588
Loss	0.0504	Silica	1.5770
Total	100.0000	Loss	0.0504
Silver ... ounces to ton	27.5327	Total	100.0000
Gold	Trace only		

SECTION V.

THEORETICAL DISCUSSION.

REACTIONS IN THE BLAST-FURNACES.

To form a correct conception of the metallurgical reactions in the blast-furnaces of Leadville we must take into consideration—

(1) The great altitude at which the smelting operations take place, which modifies to a considerable extent the volume of the blast and the volatility of volatile compounds.

(2) The manner in which the smelting charges are disposed in the furnace. It has been seen already that the ores and fluxes are placed between two layers of fuel, so that in all the zones of the furnace above those of agglomeration and fusion the reactions take place by actions of gases upon solid substances, and that in a very limited space only reactions by contact of solid matter can take place.

(3) The elements of the blast: Oxygen, nitrogen, moisture, and carbonic acid.

(4) The elements of the fuel: In coke, carbon, moisture, a little sulphide of iron, and a considerable quantity of ash, formed of silica, alumina, lime, and oxide of iron; in charcoal, carbon, moisture, and a little ash, composed of alumina and alkaline carbonates.

(5) The elements of dolomites: Carbonic acid, lime, magnesia, with small quantities of iron and other substances.

(6) The elements of hematite: Peroxide of iron, protoxide of iron, carbonate of iron, with small quantities of other substances.

(7) The elements of the ores: Carbonate of lead, sulphide of lead, sulphate of lead, pyrite, oxides of iron and manganese, chlorophosphate of lead, chloro-bromo-iodide of silver, gold, zinc, titanic and molybdic acids, and arsenic and antimonic acids, with small quantities of cobalt, nickel, and other substances.

The examination of the furnace products, which has already been made, affords means of pointing out with precision what becomes of the elements introduced in the furnace. The analyses of slag, bullion, speiss, dust, and mattes are fair representatives of the complete or normal reactions of the furnace, and those of hearth and shaft accretions, of incomplete or accidental reactions. But before entering into these considerations it is necessary to pass in review the principal reactions of lead, silver, and iron compounds, and to study their action upon each other and upon the chief ingredients either used in smelting or produced by smelting. At the same time stress will be laid upon the reactions that are represented by specimens found in the furnaces of Leadville and kept for reference in the collections of the Geological Survey, and also upon the reactions which were revealed by analysis.

REACTIONS OF LEAD COMPOUNDS.

No. 1. Reactions of carbonate of lead.—Carbonate of lead loses its carbonic acid between 170° C. and 200° C. (J. A. Phillips),[1] and is converted into protoxide of lead.

No. 2. Reactions of protoxide of lead.—Oxide of lead combines in the dry way with stannic acid, arsenious and arsenic acids, antimonious and antimonic acids, and with peroxide of iron and oxide of zinc (Berthier)[2] These reactions take place in the furnaces as is shown by analyses XLIII and XLIV of peculiar accretions.

No. 3.—Oxide of lead is partially reduced to the metallic state by magnetic oxide of iron with formation of peroxide of iron: $3Fe_3O_4+2PbO=Fe_3O_4+3Fe_2O_3+PbO+Pb$ (Berthier). The fact that some slags (see analyses of slags) contain peroxide of iron in the state of silicate seems to indicate that this reaction takes place.

No. 4.—Oxide of lead in excess is reduced to the metallic state by sulphur with formation of sulphurous acid: $2PbO+S=2Pb+SO_2$ (Berthier). This reaction undoubtedly occurs when the charges contain pyrites.

No. 5.—Oxide of lead is reduced by arsenic with formation of lead and arsenite of lead: $4PbO+As=PbO, AsO_3+3Pb$ (Berthier).

No. 6.—Conversely, metallic lead reduces arsenite of lead with formation of basic arsenite of lead and arseniuret of lead: $2(PbO,AsO_3)+4Pb=5PbO, AsO_3+PbAs$

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[1] Liebig und Kopp's Jahresb., 1851, p. 357.

[2] All the quotations from Berthier are taken from his Traité des essais par la voie sèche, Paris, 1834, and may also be found in Percy's Metallurgy of Lead, London, 1870.

(Berthier). The presence of arsenite of lead in all the oxide of lead compounds of the furnace and the presence of arsenic in bullion show that reactions 5 and 6 are of frequent occurrence.

No. 7.—Antimony acts on oxide of lead, and lead on oxide of antimony, in the same way as with arsenic (Berthier). The presence of antimonious acid in oxide of lead compounds and of antimony in bullion shows that these reactions are constantly taking place in the furnace.

No. 8.—Protoxide of lead is reduced to the metallic state by iron with formation of magnetic oxide of iron: $4PbO+Fe_3=Fe_3O_4+Pb$ (Berthier). To this reaction is undoubtedly due part of the magnetic oxide of iron found in slags and other furnace products.

No. 9. Protoxide of lead and silica combine easily at the temperature at which the oxide of lead becomes pasty. The silicate $3PbO,SiO_2$ is very fusible and very fluid. The silicate $2PbO, SiO_2$ is pasty (Percy-Beck).[1] The presence of silicate of lead in all the slags shows that this substance is formed in the furnace.

No. 10. Oxide of lead and galena.—In this well-known reaction sulphurous acid is evolved and lead is reduced to the metallic state (Berthier, Percy-Smith): $PbS+2PbO=3Pb+SO_2$. This is one of the fundamental reactions of blast-furnaces which has been proved too often to need demonstration.

No. 11.—Oxide of lead is completely reduced to the metallic state by charcoal, coke, oxide of carbon, hydrogen with formation of carbonic oxide, carbonic acid and water (Berthier, Percy, and others).

No. 12.—Oxide of lead is reduced by zinc to the metallic state by formation of oxide of zinc: $PbO+Zn=Pb+ZnO$ (Berthier). The oxide of zinc deposited in accretions and fumes is undoubtedly produced in this way by zinc reduced in the zone of agglomeration.

No. 13. Reactions of silicate of lead.—Silicate of lead behaves almost exactly like protoxide of lead in its reactions upon sulphur, iron scales, iron, carbon, carbonic oxide, galena, etc. (Percy-Beck).

No. 14.—Silicate of lead is completely reduced to the metallic state by mixtures of oxide of iron and carbon (Percy-Beck). This is undoubtedly one of the chief reactions of the furnace at the zones of agglomeration; but reactions No. 13 take place in most of the zones of the furnace.

No. 15. Reactions of sulphate of lead.—Sulphate of lead is decomposed by silica with evolution of sulphurous acid and oxygen and formation of silicate of lead (Berthier, Percy).

No. 16.—Sulphate of lead is reduced by lead to the state of oxide with evolution of sulphurous acid: $PbO, SO_3+Pb=2PbO+SO_2$ (Berthier, Percy-Smith).

No. 17.—Sulphate of lead is reduced by iron to the metallic state with formation of magnetic oxide of iron and sulphide of iron: $PbO,SO_3+4Fe=Fe_3O_4+FeS+Pb$. There is but little doubt that the mattes of Leadville owe their origin in great part to this reduction.

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[1] The quotations from Percy and his assistants, whose names follow Percy's, are taken from Percy's Metallurgy of Lead, London, 1870.

No. 18.—Sulphate of lead is reduced by carbon to the state of sulphide: $PbO, SO_3+C_2=PbS+2CO_2$ (Gay-Lussac),[1] and also by carbonic oxide (Rodwell).[2] The sulphide of lead of the mattes is produced partly by these reactions.

No. 19.—Sulphate of lead is decomposed by lime with formation of sulphate of lime and oxide of lead; sulphate of lime has been pointed out in the analyses of some furnace products.

No. 20. Reactions of sulphide of lead.—Sulphide of lead is somewhat volatile; it is sublimated at high temperatures (Berthier, Percy). This sublimated galena in fine distinct irisated crystals is one of the constituents of normal shaft accretions.

No. 21.—Sulphide and oxide of lead react upon each other with formation of metallic lead and sulphurous acid (see reaction No. 10).

No. 22.—Sulphide of lead and metallic lead combine together and form subsulphides and alloys. The analyses of bullion and skimmings prove this reaction. Moreover, the metallurgical collection of the Survey contains specimens of alloys highly charged with sulphide of lead.

No. 23.—Sulphide of lead is reduced by zinc: $PbS+Zn=Pb+ZnS$ (Percy-Smith). The sulphide of zinc found in normal accretions and also in lead fumes is certainly deposited in virtue of this reaction.

No. 24.—Sulphide of lead and sulphate of lead react upon each other with formation of metallic lead and sulphurous acid: $PbS+PbO,SO_3=2Pb+2SO_2$ (Berthier, Percy).

No. 25.—Sulphide of lead and iron produce one of the most important reactions of the blast furnace; lead is completely reduced to the metallic state and sulphide of iron is formed: $PbS+Fe=Pb+FeS$ (Berthier). Most of the sulphide of iron in mattes is produced in this way.

No. 26.—Sulphide of lead and oxide of carbon act slightly upon each other with formation of sulphide of carbon (Rodwell).[3] In all probability some of the silica found in that portion of the fumes which escapes in the air is volatilized in the state of sulphide of silicium by the sulphide of carbon thus produced.

No. 27.—Sulphide of lead mixed with lime and carbon is partly reduced with formation of sulphide of calcium and carbonic oxide: $2PbS+CaO+C=Pb+PbS,CaS+CO$ (Berthier). This important reaction, which undoubtedly takes place in the zone of agglomeration of the furnace, accounts for the sulphide of calcium in the slags.

No. 28.—Sulphide of lead, heated with oxide of iron and carbon, produces metallic lead and sulphide of iron: $4PbS+2Fe_2O_3+3C=4Pb+4FeS+3CO_2$.[4] This reaction is interesting as indicative of what actually takes place in the furnace.

No. 29.—Sulphide of lead and basic silicate of protoxide of iron react upon each other with formation of metallic lead, and iron and lead matte: $2(3FeO, SiO_3)+5PbS=2(2FeO, SiO_3)+2(PbS, FeS)+SO_2+Pb_3$ (Percy-Cloud). This important reaction is illustrated by the specimens of hearth accretions or slag-mattes in the collection of the Survey.

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[1] Ann. de Chim. et de Phys., 1836, 73, p. 435.

[2] Journal of the Chemical Society, new series, 1863, p. 42.

[3] Journal of the Chem. Soc., antea cit., p. 48.

[4] Jordan, Erdmann's Journal, 1831, 11, p. 334.

No. 30. Reactions of phosphate of lead.—Phosphate of lead forms, with chloride, bromide, and iodide of lead, very volatile compounds (A. Guyard). This is proved by the analysis of roasted chamber-dust and by that of the smoke caught in the Bartlett filter.

No. 31.—Phosphate of lead is reduced by carbon and iron to the metallic state, like oxide of lead (Percy-Cloud). Part of the phosphoric acid found in the slags comes from this reaction.

No. 32. Reactions of chloride of lead.—It is a well-known fact that chloride, bromide, and iodide of lead are volatile compounds; hence their constant presence in lead fumes of every kind in Leadville.

No. 33. Chloride of lead with lime and carbon.—Chloride of lead is reduced to the metallic state with formation of chloride of calcium and carbonic acid (Berthier): $2PbCl_2+2CaO+C=Pb_2+2CaCl_2+CO_2$. The analyses of slags show that this reaction does not take place in Leadville. It is chiefly due to the fact that chloride of lead is volatilized before carbonate of lime is decomposed, and it indicates that there would be an advantage in using caustic lime instead of raw limestone.

No. 34. Chloride of lead and galena.—These two substances form a very volatile chloro-sulphide of lead similar to galena (Berthier). This product has been found in the portion of the lead fumes lost in the air.

No. 35. Reactions of metallic lead.—Lead is somewhat volatile (all authors). It has been seen that normal accretions are chiefly formed of sublimated lead, and the contents of this sublimated lead in silver were also given.

REACTIONS OF SILVER COMPOUNDS.

No. 36. Reactions of metallic silver.—Silver is somewhat volatile (all authors). The assay of sublimated bullion found in normal shaft accretions gives an idea of the relative proportion of lead and silver volatilized and sublimated in the blast furnace.

No. 37. Reactions of sulphide of silver.—Sulphide of silver combines with metallic silver and with sulphides of lead and iron. The analyses of bullion, skimmings, and mattes show that these reactions take place in Leadville.

No. 38.—Sulphide of silver heated with oxide of lead is reduced to the metallic state, with formation of an alloy of lead and silver and sulphurous acid: $Ag_2S+2PbO=2(PbAg)+SO_2$ (Percy-Smith).[1]

No. 39.—Sulphide of silver is not completely reduced to the metallic state by metallic lead in excess (Percy).

No. 40.—Sulphide of silver is completely reduced to the metallic state by iron, with formation of sulphide of iron (Berthier, Percy, and others).

No. 41.—Sulphide of silver is not completely reduced by iron in presence of an excess of sulphide of iron (A. Guyard). The matte analyzed in XL, and which yielded 85.067 ounces of silver to the ton by scorification, gave only 80.16 ounces when it was treated directly with flux, litharge, and iron. This experiment throws light on many furnace reactions.

Reactions of chloride of silver.—What is said for chloride of silver is true for bromide and iodide.

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[1] Quotations from Percy and his assistants are taken from Percy's Metallurgy of Silver and Gold. Part 1. London, 1880.

No. 42.—It is a well-known fact that this compound is volatile; hence its presence in large quantities in the lead dust and even in the lost fumes.

No. 43.—Chloride of silver is reduced in the dry way by metallic lead and also by metallic iron. It is owing to these important reactions that so much chloro-bromo iodide of lead is formed and that so much silver is reduced in the bullion.

REACTIONS OF IRON COMPOUNDS.

No. 44. Reactions of carbonate of iron.—Carbonate of iron is reduced at red heat to the state of magnetic oxide of iron with formation of carbonic oxide (I. Lowthian Bell),[1] with formation of peculiar magnetic oxide of iron, containing an excess of protoxide of iron (Percy).[2] The writer has found a similar magnetic oxide of iron in slags and mattes.

No. 45. Reactions of peroxide of iron.—Under the influence of carbonic oxide, peroxide of iron begins to lose oxygen at the temperature of 200° C., protoxide of iron being formed as well as carbonic acid. The decomposition increases rapidly with the temperature until it reaches 417° C. The loss in oxygen is greater in the same lapse of time in a rapid current of carbonic oxide. At 410° C. peroxide of iron loses 36 per cent. of its oxygen in a slow current of carbonic oxide and 56 per cent. in a rapid current of the same gas (Bell). In the blast furnaces of Leadville the conditions are those of a rapid current. To form magnetic oxide, peroxide of iron must lose 11.1 per cent. of its oxygen, and to form protoxide of iron 33.3 per cent. Consequently at the temperature of 410°—i. e., below red heat—and in a rapid current of carbonic oxide, peroxide of iron losing more than 50 per cent. of its oxygen, some metallic iron is produced. This is an important fact, but one which is profoundly modified in the furnace, where carbonic oxide is diluted with nitrogen and carbonic acid.

No. 46.—At the temperature of 417° C.—that is, at the temperature at which metallic iron makes its appearance—it is rapidly attacked by carbonic acid, with formation of oxide of iron and oxide of carbon (Bell).

No. 47.—At the same temperature of 417° C. a mixture of equal volumes of carbonic acid and carbonic oxide exerts no action upon metallic iron, but at full red heat the carbonic acid of the mixture is rapidly decomposed and converted into carbonic oxide (Bell).

No. 48.—Mixtures of carbonic acid and oxide reduce peroxide of iron, but only to the state of protoxide, at the temperature of 417° C., with formation of carbonic acid (Bell).

No. 49.—A mixture of carbonic acid with an excess of oxide of carbon (CO_2 9 volumes, CO 100 volumes) oxidizes spongy iron, and carbon is deposited from reduced oxide of carbon, oxide of iron being formed. Pure spongy iron thus treated has for composition $Fe=91.42$, $C=0.33$, $O=8.25$ (Bell). In pure oxide of carbon spongy iron takes up as much as 23 per cent. of carbon (Bell).

The above considerations, which are purely theoretical, are interesting as showing the mechanism of the formation of cast iron in the blast furnaces, such as those of Leadville, in which the phenomena of lead and silver smelting take place jointly

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[1] All quotations from I. L. Bell are from his Chemical Phenomena of Iron Smelting. London, 1872.

[2] Percy's Metallurgy of Iron and Steel. London, 1864.

with those of iron smelting; but the following experiments, made by I. L. Bell in the iron blast furnaces in the presence of the gases actually produced in smelting, show with more accuracy the real process of the reduction of iron:

(*a*) Furnace working with raw limestone; pieces of calcined Cleveland ore or artificial hematite kept for two hours in the zone of the furnace below red point. Composition of gases: $CO=100$ volumes; $CO_2=25$ volumes; $N=190$ volumes. The ore loses 11.85 per cent. of its oxygen (mean of two experiments).

(*b*) Same experiment as in *a*, but in cherry-red zone. Composition of gases: $CO=100$ volumes; $CO_2=8\frac{1}{2}$ volumes; $N=172\frac{1}{2}$ volumes. The ore loses 76.2 per cent. of its oxygen, showing great reduction of iron.

(*c*) Same experiment as in *a* and *b*, but in bright-red zone. Composition of gases: $CO=100$ volumes; $CO_2=3\frac{1}{2}$ volumes; $N=169\frac{1}{2}$ volumes. The ore loses 73.8 per cent. of its oxygen.

(*d*) Same experiment as preceding, but in very bright-red zone. Composition of gases: $CO=100$ volumes; $CO_2=3$ volumes; $N=183\frac{1}{2}$ volumes. The ore loses 80 per cent. of its oxygen.

(*e*) Same experiment as preceding, but in intensely bright-red zone near tuyeres. Composition of gases: $CO=100$ volumes; $CO_2=5$ volumes; $N=172\frac{1}{2}$ volumes. The ore loses 71 per cent. of its oxygen.

To interpret correctly these experiments we must take into consideration that the ore does not remain exposed two hours to the influence of the gases of the same zone in the furnaces of Leadville, and that although the reducing power of the corresponding zones is sensibly the same the quantity of ore reduced is greatly diminished.

Reactions of sulphides of iron.—Pyrites existing in some ores and sulphide of iron being formed in the furnace, the following reactions are interesting:

No. 50.—Protosulphide of iron and peroxide of iron act upon each other with formation of magnetic oxide of iron and sulphurous acid: $FeS+10Fe_2O_3=7Fe_3O_4+SO_2$ (Percy-Hochstätter). To this reaction is probably due in part the magnetic oxide of mattes.

No. 51.—Iron pyrites and oxide of lead react upon each other, give off sulphurous acid, and form a magnetic mixture of sulphides and oxides of lead and iron (Percy). In this instance the origin of mattes is clearly indicated.

CHEMICAL DISCUSSION OF THE LEADVILLE FURNACES.

The object of this discussion is to illustrate the chemical and metallurgical reactions of the blast furnace, and it is based as much as possible on general averages obtained during the preparation of this report.

It has already been seen (Table IV):

(1) That the average proportion of fuel to ore is 32.83 per cent.

(2) That the average proportion of fuel to charge is 24.03 per cent.

(3) That the average composition of the fuel used in the camp is: charcoal, 57 per cent.; coke 43 per cent. = 100 per cent.

(4) That the average proportion of ash in coke is 22 per cent. and in charcoal 2.5 per cent., giving for the fuel under consideration an average of 10.88 per cent. of

ash. We will assume from the nature of this fuel that its average proportion of moisture is equal to 5 per cent. and of gases 3 per cent.

Raw materials.—The average composition of ore, hematite, dolomite, fuel, ash in fuel, and atmospheric air which will be adopted in this discussion is the following:

Component-parts.	Ore.	Hematite.	Dolomite.	Fuel.	Ash in fuel.	Atmospheric air.
Carbon				80.82		
Gases in fuel				3.00		
Nitrogen						77.50
Oxygen	9.53	22.50				21.41
Carbonic acid	5.58		45.83		3.50	0.04
Moisture	5.53	1.92	0.10	5.00		1.05
Lead	23.00					
Silver	0.31					
Metallic iron	18.10	52.50				
Alumina	3.99	0.50				
Peroxide of iron *a*			0.66		22.20	
Peroxide of manganese	4.03	0.42				
Lime	2.36	0.37	30.88		4.00	
Magnesia	3.04	0.63	19.57		0.50	
Alkalies	0.98				2.00	
Silica	22.59	16.30	2.40		56.80	
Ash				10.88		
Other constituents *b*	0.96	4.86	0.56	0.30	11.00	
Totals	100.00	100.00	100.00	100.00	100.00	100.00

a It will be assumed that this peroxide of iron escapes reduction.

b Sulphur, arsenic, antimony, chlorine, phosphoric acid, &c., are neglected in subsequent calculations.

The composition of old slags is not given here, it being unnecessary for this discussion.

Of the above chemical constituents the following proportional quantities enter into the smelting charge, calculated on the basis of averages given in Section V of Table IV:

Constituents of smelting charges.	From ore.	From hematite.	From dolomite.	From fuel.	From ash in fuel.
Carbon				26.533	
Gases in fuel				0.985	
Oxygen	9.53	1.867			
Carbonic acid	5.58		4.986		0.125
Moisture	5.53	0.159	0.011	1.641	
Lead	23.00				
Silver	0.31				
Metallic iron	18.10	4.357			
Alumina	3.99	0.041			
Peroxide of iron			0.072		0.793
Peroxide of manganese	4.03	0.035			
Lime	2.36	0.031	3.360		0.143
Magnesia	3.04	0.052	2.129		0.018
Alkalies	0.98				0.071
Silica	22.59	1.353	0.261		2.029

The above figures may be conveniently combined as follows:

Volatile portion.

Carbon from fuel		26.533	
Gases from fuel		0.985	
Carbonic acid from ore, dolomite, and ash		10.691	
Oxygen from Fe_2O_3 in ore and hematite, escaping as CO_2		3.208	
Oxygen from MnO_2 in ore and hematite, escaping as CO_2		0.747	
Oxygen from PbO escaping as CO_2		1.640	
Moisture from whole charge		7.341	
Dust and fumes carried away	Pb=1.040 Ag=0.011 Dust=0.890	1.941	
Total			53.086

Slag.

Silica	26.202	
Alkalies	1.051	
Magnesia	5.239	
Lime	5.894	
Alumina	4.031	
Protoxide of iron	28.873	
Peroxide of iron	0.865	
Protoxide of manganese	3.318	
Protoxide of lead	1.860	
Old slag	17.441	
Total	94.774	
Less dust	0.890	
Total slag produced		93.884

Bullion.

Lead	20.240	
Silver	0.299	
Total		20.539
Total of volatile portion, slag, and bullion		167.509

In these calculations speiss and matte have been purposely neglected. Besides the descending charge of solid matter thrown in the furnace at the feed-hole, there is the ascending charge of blast forced in at the tuyeres. The weight of the gaseous charge will be calculated after the weights of charges filling the furnace have been determined.

This discussion is carried on for the blast furnace of smelter C, which is shown in Fig. 2, Plate XXXVI, divided into zones of charges and temperatures in accordance with the working of this furnace. These zones will be designated hereafter by their temperatures. The zone 900° of the crucible is charged with 12 tons of bullion.[1]

It is assumed that the furnace is in full blast; that the weight of smelting charges, including fuel, is equal to 502.527 pounds, or three times the weight of the charge cal-

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[1] Temperatures are theoretical, not observed.

culated for 100 pounds of ore. This figure is calculated from the capacity of the furnace, and represents very nearly a semi-charge of smelter C, but as a great many furnaces run charges of this weight it will be adopted in this discussion. The furnace is capable of smelting this charge in about seven minutes, giving a cake of slag weighing 281.652 pounds, and it is assumed that at the time of the experiment the furnace is properly filled with eleven similar charges, each charge being divided into two layers, a layer of fuel and old slag and a layer of flux and ore, all these suppositions and all the figures adopted being in perfect accordance with the practical working of the furnace.

The composition in pounds of the charge in zone 150°, the zone in which it is thrown into the furnace, will be as follows:

Carbon	79.599
Gases from fuel	2.955
Carbonic acid (C=8.747)	32.073
Oxygen	33.792
Moisture	22.023
Lead (total CO_2 combined with PbO=14.667; PbO=68.774; O=4.934)	63.840
Silver	0.930
Metallic iron (Fe_3O_2=93.036; FeO=86.620; Fe_2O_3=96.244)	67.371
Oxide of lead in slag	5.580
Alumina	12.093
Peroxide of iron	2.595
Peroxide of manganese (MnO=9.952; O=2.243)	12.195
Silica	78.606
Lime (total CO_2 in combination=17.406)	17.682
Magnesia	15.717
Alkalies	3.153
Old slag	52.323
Total	502.527

Weight of blast.—It will be seen in the discussion of losses in each zone of the furnace that of the 79.599 pounds of carbon thrown in the furnace with each charge only 32.1257 pounds reach the zone of combustion. It is this quantity of carbon which will enable us to calculate the quantity of air necessary to convert it into carbonic acid in seven minutes, an excess of air being injurious and calculated to cool the furnace. As 32.1257 pounds of carbon require 85.6685 pounds of oxygen for their combustion into carbonic acid, it is deduced from the composition of air given previously that the air blown in the furnace in seven minutes will be composed of—

Oxygen	85.6685
Moisture	4.2023
Carbonic acid	0.1600
Nitrogen	310.1031
Total	400.1339

In other words, the weight of air strictly necessary to burn the carbon left in the smelting charge at the tuyeres is about four-fifths of the weight of the charge thrown in at the feed-hole. At sea-level the volume of air corresponding to the weight of 400.1339 would be 5,356 cubic feet (1 cubic foot=538.569 grains). In Leadville, at the normal pressure of 21 inches of mercury, the volume of the same weight of air is represented by 7,906.2 cubic feet.

It has been seen that the normal volume of blast delivered by the blower connected with this furnace is 2,550 cubic feet per minute, or 17,850 cubic feet in seven minutes, at the rate of 80 revolutions per minute; consequently, 9,943.8 cubic feet of blast must be shut out during the seven minutes. This is done, as before shown, by leaving open the damper placed at the extreme end of the main blast-pipe. However, there is hardly any doubt that in practice an excess of air passes through to the furnace, and it is to avoid this that the adoption of a meter to be placed at the induction pipe has been recommended.

Loss of weight of charges.—The data are now prepared which are necessary for the discussion of the loss of weight of smelting charges in every zone of the furnace in seven minutes, and for giving an idea of the chief reactions that take place in each. As the element of time is all-important in these discussions, the reactions have been described in zones of temperature higher than those indicated by theory; but, if there are a few errors of judgment in the position assigned to them, the final results remain unaltered.

The weight of gases[1] which pass through the uppermost zone in seven minutes is as follows:

Zone of gases 150° *C.*

Carbonic acid	27.455
Oxide of carbon	188.771
Vapor of water	26.225
Nitrogen	310.103
Gases from fuel	2.955
Dust and lead fumes	5.823
Total	561.322

Zone of desiccation 150° *C.*

Weight of charge entering zone		502.5270
It loses one-fourth of its moisture (chiefly from the ore), or	5.5058	
It loses one-eleventh of total loss in dust and fumes	0.5294	
		6.0352

Zone of desiccation 255° *C.*

Weight of charge entering zone		496.4918
It loses one-fourth of its moisture (chiefly from the ore)	5.5058	
It loses one-eleventh of total loss in dust and fumes	0.5294	
		6.0352

Zone of desiccation 360° *C.*

Weight of charge entering zone		490.4566
It loses one-fourth of its moisture (from ore and fuel)	5.5057	
It loses one-eleventh of total loss in dust and fumes	0.5294	
		6.0351

Zone of desiccation 465° *C.*

Weight of charge entering zone		484.4215
It loses the remaining one-fourth of its moisture (chiefly from fuel)	5.5057	
It loses one-eleventh of total loss in dust and fumes	0.5294	
		6.0351
		478.3864

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[1] As no analyses were made of the gases passing through the different zones of the furnaces at Leadville, the figures given by Mr. Guyard in the following tables cannot be assumed to represent their actual composition. In point of fact, their composition must be very different from that assumed by him, owing to the excess of blast used. His idea is evidently to represent what the theoretical conditions in the furnace should be when employing the least amount of blast for the given amount of fuel. (S. F. E.)

Zone of decomposition 570° *C.*

Weight of charge entering zone			478.3864
It loses:			
One-eleventh of the total loss in dust and fumes		0.5294	
One-half of the gases from fuel		1.4775	
One-half of the CO_2 from PbO, CO_2[a]		7.3333	
C oxidized by ⅓ CO_2 formed in zone 780° C. (2.2614 pounds CO_2)		0.6167	
C oxidized by ¼ CO_2 formed in zone 885° C. (5.4436 pounds CO_2)		1.4846	
C oxidized by ⅕ CO_2 formed in zone 990° C. (3.5288 pounds CO_2)		0.9624	
C oxidized by ⅙ CO_2 formed in zone 1,095° C. (0.8822 pounds CO_2)		0.2406	
C oxidized by $\frac{1}{11}$ CO_2 formed in zone 1,200° C. (10.7086 pounds CO_2)		2.9205	
C oxidized by ⅓ CO_2 expelled from PbO, CO_2 in zone 675° C. (2.4444 pounds CO_2)		0.6667	
C oxidized by ⅕ CO_2 expelled from carbonates in zone 990° C. (1.1604 pounds CO_2)		0.3164	
C oxidized by ⅙ CO_2 expelled from carbonates in zone 1,095° C. (0.9670 pounds CO_2)		0.2637	
C oxidized by ⅐ CO_2 expelled from carbonates in zone 1,200° C. (0.8288 pounds CO_2)		0.2260	17.0378
Total C oxidized in zone 570° C.	7.6976		
Total CO formed in same zone	35.9221		

Zone of reaction 675° *C.*

Weight of charge entering zone			461.3486
It loses:			
One-eleventh of the total loss in dust and fumes		0.5294	
The remaining ½ of gases from fuel		1.4775	
The remaining ½ of CO_2 from PbO, CO_2[b]		7.3334	
C oxidized by ⅓ CO_2 formed in zone 780° C. (2.2614 pounds CO_2)		0.6168	
C oxidized by ¼ CO_2 formed in zone 885° C. (5.4436 pounds CO_2)		1.4846	
C oxidized by ⅕ CO_2 formed in zone 990° C. (3.5288 pounds CO_2)		0.9624	
C oxidized by ⅙ CO_2 formed in zone 1,095° C. (0.8822 pounds CO_2)		0.2406	
C oxidized by $\frac{1}{11}$ CO_2 formed in zone 1,200° C. (10.7086 pounds CO_2)		2.9205	
C oxidized by ⅕ CO_2 expelled from carbonates in zone 990° C. (1.1604 pounds CO_2)		0.3165	
C oxidized by ⅙ CO_2 expelled from carbonates in zone 1,095° C. (0.9670 pounds CO_2)		0.2637	
C oxidized by ⅐ CO_2 expelled from carbonates in zone 1,200° C. (0.8288 pounds CO_2)		0.2260	16.3714
Total C oxidized in zone 675° C.	7.0311		
Total CO formed in same zone	32.8118		

Zone of reduction 780° *C.*

Weight of charge entering zone			444.9772
It loses:			
One-eleventh of the total loss in dust and fumes		0.5294	
One-half of the oxygen from total PbO reduced (this forms with CO 6.78425 pounds CO_2, of which ⅔ is reduced in zones 570° and 675°, while ⅓ escapes)		2.4670	
C oxidized by ¼ CO_2 formed in zone 885° C. (5.4437 pounds CO_2)		1.4847	
C oxidized by ⅕ CO_2 formed in zone 990° C. (3.5288 pounds CO_2)		0.9624	
C oxidized by ⅙ CO_2 formed in zone 1,095° C. (0.8822 pounds CO_2)		0.2406	
C oxidized by $\frac{1}{11}$ CO_2 formed in zone 1,200° C. (10.7086 pounds CO_2)		2.9205	
C oxidized by ⅕ CO_2 expelled from carbonates in zone 990° C. (1.1604 pounds CO_2)		0.3165	

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[a] It will be assumed that, owing to the velocity of blast, the admixture of inert gases from zones below, and the low temperatures of zones above, this CO_2 escapes reduction to CO.

[b] It will be assumed, for the reasons stated under zone 570°, that only ⅓ of the CO_2 driven off here is reduced to CO in upper zone 570°.

C oxidized by $\frac{1}{6}$ CO_2 expelled from carbonates in zone 1,095° C. (0.9670 pounds CO_2)		0. 2637	
C oxidized by $\frac{1}{7}$ CO_2 expelled from carbonates in zone 1,200° C. (0.8288 pounds CO_2)		0. 2260	
			9. 41
Total C oxidized in zone 780° C.	6. 4144		
Total CO formed in same zone	29. 9339		
Of this latter there is used for reduction of ore	4. 3172		
Leaving for escaping CO	25. 6167		

Zone of reduction 885° *C.*

Weight of charge entering zone			435, 5664
It loses:			
One-eleventh of the total loss in dust and fumes		0. 5293	
The remaining $\frac{1}{2}$ of oxygen from total PbO reduced		2. 4670	
Oxygen from MnO_2 reduced to MnO		2. 2430	
Oxygen from Fe_2O_3 reduced to Fe_3O_4		3. 2080	
(Total oxygen 7.9180. This forms with CO 21.7745 pounds CO_2, of which $\frac{3}{4}$ is reduced in upper zones, while $\frac{1}{4}$ escapes.)			
C oxidized by $\frac{1}{5}$ CO_2 formed in zone 990° C. (3.5288 pounds CO_2)		0. 9624	
C oxidized by $\frac{1}{6}$ CO_2 formed in zone 1,095° C. (0.8823 pounds CO_2)		0. 2406	
C oxidized by $\frac{2}{11}$ CO_2 formed in zone 1,200° C. (21.4171 pounds CO_2)		5. 8410	
C oxidized by $\frac{1}{5}$ CO_2 expelled from carbonates in zone 990° C. (1.1604 pounds CO_2)		0. 3165	
C oxidized by $\frac{1}{6}$ CO_2 expelled from carbonates in zone 1,095° C. (0.9670 pounds CO_2)		0. 2637	
C oxidized by $\frac{1}{7}$ CO_2 expelled from carbonates in zone 1,200° C. (0.8289 pounds CO_2)		8. 2261	
			16. 2976
Total C oxidized in zone 885° C.	7. 8503		
Total CO formed in same zone	36. 6347		
Of this latter there is used for reduction of ore	13. 8565		
Leaving for escaping CO	22. 7782		

Zone of semi-agglomeration 990° *C.*

Weight of charge entering zone			419. 2688
It loses:			
One-eleventh of total loss in dust and fumes		0. 5293	
One-third of CO_2 from carbonates		5. 8020	
(Of this $\frac{4}{5}$ is reduced in upper zones, while $\frac{1}{5}$ escapes.)			
Oxygen from Fe_3O_4 reduced to FeO		6. 4160	
(This forms with CO 17.6440 pounds CO_2, of which $\frac{4}{5}$ is reduced in upper zones, while $\frac{1}{5}$ escapes.)			
C oxidized by $\frac{1}{6}$ CO_2 formed in zone 1,095° C. (0.8823 pounds CO_2)		0 2407	
C oxidized by $\frac{3}{11}$ CO_2 formed in zone 1,200° C. (32.1256 pounds CO_2)		8. 7616	
C oxidized by $\frac{1}{6}$ CO_2 expelled from carbonates in zone 1,095° C. (0.9670 pounds CO_2)		0. 2637	
C oxidized by $\frac{1}{7}$ CO_2 expelled from carbonates in zone 1,200° C. (0.8289 pounds CO_2)		0. 2261	
			22. 2395
Total C oxidized in zone 990° C.	9. 4921		
Total CO formed in same zone	44. 2965		
Of this latter there is used for reduction of ore	11. 2280		
Leaving for escaping CO	33. 0685		

Zone of agglomeration 1,095° *C.*

Weight of charge entering zone			397, 0293
It loses:			
One-eleventh of the total loss in dust and fumes		0. 5293	
One-third of the CO_2 from carbonates		5. 8020	
(Of this $\frac{5}{6}$ is reduced in upper zones, while $\frac{1}{6}$ escapes.)			
Oxygen from $\frac{1}{10}$ FeO reduced to Fe		1. 9249	
(This forms with CO 5.2935 pounds CO_2, of which $\frac{5}{6}$ is reduced above, while $\frac{1}{6}$ escapes.)			
C oxidized by $\frac{1}{7}$ CO_2 expelled from carbonates in zone 1,200° C. (0.8289 pounds CO_2)		0. 2261	
C oxidized by $\frac{3}{11}$ CO_2 formed in zone 1,200° C. (32.1256 pounds CO_2)		8.7616	
			17. 2439
Total C oxidized in zone 1,095°C	8. 9877		
Total CO formed in same zone	41. 9426		
Of this latter there is used for reduction of ore	3. 3686		
Leaving for escaping CO	38. 5740		

Zone of combustion and reaction by contact of solid matter, 1,200° *C.*

Weight of charge entering zone		379. 7854
It loses:		
One-eleventh of the total loss in dust and fumes	0. 5293	
One-third of the CO_2 from carbonates	5. 8020	
(Of this $\frac{6}{7}$ is reduced in upper zones, while $\frac{1}{7}$ escapes.)		
C oxidized to CO_2	32. 1257	
		38. 4580
Remaining in furnace		341. 3274 [a]
Weight of bullion reaching zone of crucible	61. 6170	
Weight of slag produced	281. 6520	
		343. 2690

Chemical reactions of the different zones.—Some important reactions begin to take place in zone 570° C. and are continued in zone 675° C. Oxide of lead acts on galena, sulphide of silver, and pyrites, and some sulphurous acid is evolved. Mattes begin to form. Oxide of lead acts on silica, and some silicate of lead is formed. In practice dolomites lose here a portion of their carbonic acid, but the discussion has been carried on as if dolomite behaved like carbonate of lime, in order to get at extreme results.

In zone 780° C., the important reaction of reduction of lead taking place, several more reactions are produced in consequence. Metallic lead acts on arseniate and antimoniate of lead, forming arseniuret and antimoniuret of lead, with regeneration of oxide of lead. Metallic lead acts also on sulphate of lead, with regeneration of oxide and evolution of sulphurous acid; all the reactions which have escaped completion in the upper zones are completed here. Metallic lead acts on galena, and subsulphides are formed. Sulphide of lead acts on sulphate and oxide of lead, with evolution of sulphurous acid and reduction of lead. Sulphide of lead acts on silicate of lead, and mattes are produced. Metallic lead acts on chloro-bromo-iodide of silver and forms

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[a] The discrepancy between this number and that next below, representing the sum of bullion and slag, is mainly due to the fact that while Mr. Guyard expressly says (p. 739) he has neglected consideration of speiss and matte in this discussion, he nevertheless allows for the reduction of $\frac{1}{10}$ FeO in zone 1,095° C. Calculations carried out with a view to correcting the above discrepancy give 32.7272 as the amount of carbon burnt in zone 1,200°. Such correction involves, naturally, slight alterations in some of the figures, representing loss in all except the four highest of the other zones, and also alters the data relative to the amount of blast theoretically required. A second cause of difference between the above numbers is found in the fact that the figures for composition of charge (p. 740) are accurate in some cases only to the second decimal figure. The slight inaccuracies in the values of the third decimal necessarily involve errors in calculations carried to four decimal places. (S. F. E.)

chloro-bromo-iodide of lead, which is volatilized with the chlorophosphate of lead of the ore, and reduced silver alloys with lead, forming bullion. Sulphide of silver is partly acted upon by lead also, and some galena is regenerated. Chloro-bromo-iodide of lead acts on galena and volatilizes a portion of this substance.

Zone 885° C. is one of the most important with regard to reactions. Silicate of lead acts partly on the magnetic oxide of iron formed, reoxidizes it, and some peroxide of iron combines with silica. All the constituents of the charge are in a semi-fluid condition, and all possible compounds are formed here, some of which will be destroyed by thorough fusion in lower zones. Sulphate of lead is acted on energetically by silica; all the reactions of zone 780° C. are produced here also with even more energy. Sulphide of carbon is formed and produces sulphides of silicium and magnesium. Some volatile chlorides of non-volatile metals are also formed. All the reactions which generate mattes are to be observed in this zone. In this zone also the quantities of carbonic oxide and carbonic acid are nearly equal. Hematite loses completely the carbonic acid of its carbonate of iron. Zinc, reduced in zones below, acts on galena, and sulphide of zinc is formed.

Zone 990° C. is one of very important reactions. Lime and magnesia being set free act energetically on sulphide of lead and pyrites, forming the sulphide of calcium found in the slag. Silica combines with lime, magnesia, and protoxide of iron, and slag is formed. Oxide of lead is expelled from its silicate. Phosphate of lead which has escaped volatilization forms the phosphate of lime found in slags and accretions.

In zone 1,095° C., iron reduces arseniuret of lead, forming spiess, and sulphide of lead, forming matte. It acts also on oxide of lead expelled from silicate, forming magnetic oxide of iron, which enters the slag and the matte. This zone is the zone of refining of bullion. It is here also that molybdic oxide is reduced and that iron and speiss combine with it.

The preceding chemical discussion was carried on also with a view to ascertain the zones of absorption and of production of heat in the furnace, and it was the writer's intention to develop a complete thermic discussion of the different zones; but neither the time nor the means of determining with accuracy a few important data peculiar to the blast-furnaces in which lead is smelted could be had, and a discussion based on hypotheses would have lost all scientific or practical value.

CONCLUSIONS.

The chief conclusions arrived at in the preceding pages are:

1. That smelting in Leadville is a profitable operation, but that the aggregate smelting capacity of the working smelters is about equal to the present mining product of the camp.

2. That lead smelting in Leadville has, on the whole, been brought to a state of great perfection with regard both to the plant adopted, which is constructed on the most approved principles, and to the manner in which fuel, fluxes, and ores are mixed for smelting, giving slags which are remarkable for their fluidity and not too highly charged with either silver or lead (especially when it is remarked that the bullion produced is very rich), and from which by-products, such as speiss and matte, are easily detached.

3. That the quantity of by-products, other than lead fumes, resulting from smelting in Leadville amounts to but little.

4. That the camp is provided with the necessary plant to work profitably such by-products as are generally rich in silver and either completely neglected, or treated imperfectly and with a considerable loss of silver.

5. That the mode adopted at a great many smelters of mixing and resmelting with caustic lime the chamber-dust, formed in considerable quantity, is the best that could have been devised, and that it would be advisable to substitute pure lime for the dolomitic lime used in Leadville for this operation.

6. That the numerous imperfections noticeable at various smelters are mostly intentional and based on economical grounds, and not on ignorance, for smelting is conducted in Leadville by very clever superintendents and smelters.

7. That the smelting of lead ores in the presence of iron-stone has here been brought to a state of great practical perfection, and is carried on most successfully from one year's end to the other with the greatest regularity at a dozen smelters, and that superintendents of smelters do not hesitate to introduce in the charges sometimes very large quantities of galena, which are reduced with the greatest facility.

8. That, owing to the peculiar nature of the Leadville ores and to the great altitude at which smelting is performed, which increases the volatility of lead compounds, attempts ought to be made to substitute caustic lime free from magnesia for the raw dolomite used in Leadville, in order to avoid as much as possible the formation of volatile lead compounds.

9. That, cæteris paribus, dolomite forms as good a flux as calcitic limestone, so far as the actual working of the blast-furnaces is concerned, and that the fluidity of the slag thus formed is not only irreproachable but quite remarkable.

10. That, besides the substances existing in large quantities in the camp, such as silica, sulphur, carbonic acid, lime, magnesia, alumina, oxides of iron and manganese, lead, silver, chlorine, and phosphoric acid, the following substances exist in small quantities: Sulphuric acid, titanic acid, bromine, iodine, zinc, baryta, gold, nickel, molybdenum, arsenic, antimony, and copper; and that traces of the following substances may be detected: Tin, bismuth, cobalt, indium, selenium, tellurium, cadmium, and a new metal, which has been imperfectly studied as yet, and which appears to be intermediate between the metals of the iron group and those of the lead group.

11. That the ores of Leadville are either rich in lead and poor in silver, rich in silver and poor in lead, or equally rich in both silver and lead, and very variable in composition; but that, by judicious admixtures of various ores, ore-beds of sensibly the same composition are made at the smelters, which are needed to insure regularity in the smelting operations.

12. That the quantity of lead completely lost in the atmosphere is sensibly twice as large as the quantity of lead caught in the dust-chambers generally used.

13. That the crude bullion extracted in the blast-furnaces of Leadville by the process referred to in section 7 is of very fair quality, and that a little of its silver and some of its lead exist there in the state of sulphides.

14. That mattes (both iron and lead mattes), which had hitherto been considered as entirely formed of sulphides, are crystallographic compounds of sulphides of iron and lead and crystallized magnetic oxide of iron. (This last observation, however, interferes in no way with the fact that in various smelting operations mattes entirely formed of sulphides are produced.)

15. That slags cannot very well be compared with minerals, from which they differ essentially; that they contain minute quantities of carbonates which have escaped destruction, and small quantities of carbon or carburets, two products which hitherto had not been generally known to exist. That slags are formed of crystallographic compounds of silicates of iron, manganese, zinc, lead, lime, and magnesia on the one hand, and on the other of a peculiar matte which is designated by the name of calcium matte, and which, like its congeners, is formed of a sulphide (sulphide of calcium) and of magnetic oxide of iron, which can be isolated in the pure crystalline state.

16. That at least three distinct metallurgical kinds of speiss, containing two distinct chemical arsenio-sulphurets of iron, are formed in lead smelting; and that they always contain small quantities of nickel and molybdenum entirely concentrated in them, showing that the metallurgy of molybdenum could be conducted jointly with that of lead with ores containing only traces of molybdenum.

17. That a very curious and a hitherto unsuspected reaction takes place in the blast-furnaces of Leadville, by means of which cobalt is completely separated from nickel (nickel being concentrated in speiss and cobalt in the skimmings of the lead-pots of blast furnaces), and showing that the metallurgy of both metals and their separation could be effected in lead furnaces by operating under conditions similar to those observed in Leadville.

18. That iron sows are a variety of speiss and present a great analogy with the latter products.

19. That lead fumes are very complicated products, characterized in Leadville by the presence of no inconsiderable amount of chloro-bromo-iodide of lead and phosphate of lead, and that they contain, contrary to the opinion formed in Leadville, but small quantities of arsenic and antimony.

20. That the practice of roasting the dust in order to free it from arsenic and antimony, as adopted at one smelter, is a useless and costly one, which ought not to be generalized in Leadville.

21. That accretions are products of sublimation, and that these products, which line the shafts of the furnaces and interfere seriously with a regular run, might be to some extent avoided, or made less troublesome, by a slight modification of the manner of charging the furnaces and by the adoption of caustic lime instead of raw limestone in smelting.

22. That some accretions are characterized by the concentration, sometimes in large quantities, of metals such as tin, arsenic, antimony, and zinc, which exist but in small quantities in the ores.

23. That the charcoal used in smelting is of very good and the coke of bad quality; but that the fuel obtained by mixing them contains 10 per cent. of ash, and that it requires a maximum amount of 32 to 33 parts of this fuel for 100 parts of ore, and 24 parts for 100 parts of charges, to effect smelting; but that at several smelters these percentages are considerably lowered.

24. That for every 100 parts of carbon thrown in the furnaces with the smelting charges, only 40.36[1] parts reach the zone of combustion at the tuyeres, the balance being oxidized in the upper zones to carbonic oxide, chiefly by the carbonic acid formed in the zone of combustion, involving, as is well known, an absorption of heat.

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[1] This figure differs from that given in the abstract of this report published in the Second Annual report of the Director of the United States Geological Survey (1880-'81), several clerical errors in Mr. Guyard's calculations having been discovered since that appeared. (S. F. E.)

LIST OF METALLURGICAL PLATES.

INDEX OF LETTERS USED ON PLATES.

A, crucible.
A', interior of crucible.
A'', dam of crucible.
a, sheathing of crucible.
a', sheathing of lead-pot.
B, water-jackets.
B', slag-pot.
b, fire-brick above water-jacket.
C, masonry of shaft of furnace.
C', walls of shaft of furnace.
D, chimney.
D', dust-chamber.
d, sheathing on top of crucible.
E, stack.
F, stack to dust-chamber.
F', flue to dust-chamber.
G (G'), damper.
H, feed-holes.
I, induction blast-pipe.
J, branch blast-pipe from I.
J', iron jacket to furnace.
K, wind-bags.
L, lead-pot.
L', siphon.
l, sliding valve of tuyere.
M, cold-water conduits.
M', hot-water outlet.
N, tuyere.
n, tuyere-holes.
O, cast-iron bed-plate.
O', vertical flange of bed-plate.
O'', projecting horizontal flange on O (in circular furnaces).
P, cast iron pillars.
P', feeding-floor.
Q, braces on C.
q, water-jacket clamps.
R, water-jacket feeder.
R', main blast-pipe, feed of T'.
r, brackets on O'.
S' (S''), sliding-doors of feed-openings.
s, outlet-pipes of R.
T, water-gutter.
T', blast-pipe, feed of I.
t, brackets of pillars.
U, slag-gutter.
V, slag-outlet in water-jacket.
W, hood,
W', chimney of hood.
X, fore-hearth.
X', hearth.
Y, cold-water faucets.
Z, tap-hole in V.

INDEX.

[Small capitals indicate headings of chapters, sections, or subsections. The names of all mines or prospects mentioned in the work will be found under "Mines."]

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* Under this entry, s, signifies shaft; t, tunnel; i, incline; b h, bore hole. The letters and numbers (*e. g.*, M-36) denote the location on the Leadville map.