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## ISOTOPIE REFERENCE MATERIAL CBNM-IRM 184

The IRM is supplied with with atomic isotope ratios certified as

$^{233}\text{U}/^{238}\text{U}$	$< 0.000\ 000\ 1$
$^{234}\text{U}/^{238}\text{U}$	$= 0.000\ 052\ 8 \pm 0.000\ 003\ 1$
$^{235}\text{U}/^{238}\text{U}$	$= 0.007\ 262\ 3 \pm 0.000\ 002\ 2$
$^{236}\text{U}/^{238}\text{U}$	$= 0.000\ 000\ 2 \pm 0.000\ 000\ 1$

This corresponds to an isotopic composition with following abundances :

	Isotopic atom %	Isotopic mass %
$^{233}\text{U}$	$< 0.000\ 01$	$< 0.000\ 01$
$^{234}\text{U}$	$0.005\ 24$	$0.005\ 15$
$^{235}\text{U}$	$0.720\ 96$	$0.711\ 92$
$^{236}\text{U}$	$0.000\ 02$	$0.000\ 02$
$^{238}\text{U}$	$99.273\ 78$	$99.282\ 91$
The atomic weight of the uranium is		$238.028\ 896 \pm 0.000\ 014$

## NOTES

1. All uncertainties indicated are accuracies, computed on a 2s basis.
2. Values for isotope ratios, isotopic compositions and for concentrations are valid for 1985.10.10.
3. The IRM solution has a molality of  $6\ \text{mol HNO}_3$  (i.e.  $6\ \text{mol HNO}_3 \cdot \text{kg}^{-1}$  of solvent) or a molarity of  $5\ \text{M HNO}_3$  (i.e.  $5\ \text{mol HNO}_3 \cdot \text{l}^{-1}$  of solution).

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