



CERTIFICATE

ISOTOPIE REFERENCE MATERIAL CBHM-IRM 184

The IRM is supplied with with atomic isotope ratios certified as

- $^{233}\text{U}/^{238}\text{U} < 0.000\ 000\ 1$
- $^{234}\text{U}/^{238}\text{U} = 0.000\ 052\ 8 \pm 0.000\ 003\ 1$
- $^{235}\text{U}/^{238}\text{U} = 0.007\ 262\ 3 \pm 0.000\ 002\ 2$
- $^{236}\text{U}/^{238}\text{U} = 0.000\ 000\ 2 \pm 0.000\ 000\ 1$

This corresponds to an isotopic composition with following abundances :

	Isotopic atom %	Isotopic mass %
^{233}U	$< 0.000\ 01$	$< 0.000\ 01$
^{234}U	$0.005\ 24$	$0.005\ 15 \pm 0.000\ 31$
^{235}U	$0.720\ 96$	$0.711\ 92 \pm 0.000\ 22$
^{236}U	$0.000\ 02$	$0.000\ 02 \pm 0.000\ 01$
^{238}U	$99.273\ 78$	$99.282\ 91 \pm 0.000\ 38$
The atomic weight of the uranium is		$238.078\ 896 \pm 0.000\ 014$

NOTES

1. All uncertainties indicated are accuracies, computed on a 2s basis.
2. Values for isotope ratios, isotopic compositions and for concentrations are valid for 1985.10.10.
3. The IRM solution has a molality of $6\ \text{m}\ \text{HNO}_3$ (i.e. $6\ \text{mol}\ \text{HNO}_3 \cdot \text{kg}^{-1}$ of solvent) or a molarity of $5\ \text{M}\ \text{HNO}_3$ (i.e. $5\ \text{mol}\ \text{HNO}_3 \cdot \text{l}^{-1}$ of solution).

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