

A COMPUTER PROGRAM INCORPORATING PITZER'S EQUATIONS FOR CALCULATION OF GEOCHEMICAL REACTIONS IN BRINES

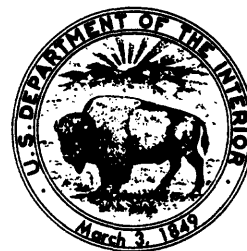
By L.N. Plummer, D.L. Parkhurst, G.W. Fleming, and S.A. Dunkle

U.S. GEOLOGICAL SURVEY

Water-Resources Investigations Report 88-4153

Reston, Virginia

1988



DEPARTMENT OF THE INTERIOR

DONALD PAUL HODEL, Secretary

U.S. GEOLOGICAL SURVEY

Dallas L. Peck, Director

For additional information write to:

**Chief, Branch of Regional Research,
Northeastern Region
U.S. Geological Survey
432 National Center
12201 Sunrise Valley Drive
Reston, Virginia 22092**

Copies of this report can be purchased from:

**U.S. Geological Survey
Books and Open-File Reports Section
Federal Center, Bldg. 810
P.O. Box 25425
Denver, Colorado 80225**

Additional copies of the software described in this report are available from:

**WATSTORE PROGRAM OFFICE
U.S. Geological Survey
437 National Center
Reston, VA 22092**

CONTENTS

	page
Abstract -----	1
Introduction -----	1
Pitzer equations -----	2
Data base -----	5
Extensions of the data base -----	6
Internal consistency of the data base -----	6
Literature sources of Pitzer interaction parameters -----	11
Scale convention of activity coefficients -----	11
Precautions and limitations -----	14
Numerical method -----	15
Basic equations -----	15
Restrictions on the Newton-Raphson approach -----	17
Scaling the Newton-Raphson results -----	17
Description of input -----	18
Title and option line -----	19
Keywords -----	21
ELEMENTS -----	21
SPECIES -----	22
SOLUTION -----	23
MINERALS -----	25
LOOK MIN -----	27
MEAN GAM -----	27
TEMP -----	28
STEPS -----	28
REACTION -----	29
NEUTRAL -----	30
SUMS -----	31
END -----	31
Description of PHRQPITZ.DATA file -----	31
Description of PITZER.DATA file -----	32

Test problems -----	34
1 -- Speciate a brine sample and examine effects of changing activity- coefficient scale -----	34
2 -- Equilibration of pure water with a set of minerals accompanied by an irreversible reaction -----	59
3 -- The anhydrite-gypsum phase boundary in the system NaCl-H ₂ O at 25 °Celsius -----	75
4 -- Solubility with incremental temperature variation: halite-water system -----	83
5 -- Fresh water-brine mixing in a closed system -----	98
6 -- Simulation of reaction path accompanying the evaporation of sea water ---	120
Program source code -----	140
Interactive construction of input files for the computer program -----	140
Acknowledgment -----	141
References -----	141
Attachments:	
A Listing of the file PHRQPITZ.DATA -----	144
B Listing of the file PITZER.DATA -----	149
C Summary of published Pitzer interaction parameters -----	154
C.1 List of references cited in Attachments C.2, C.3, C.4, and C.6 -----	155
C.2 Summary of literature values of Pitzer interaction parameters for single salts -----	164
C.3 Summary of literature values of Pitzer interaction parameters for mixed-salt solutions -----	173
C.4 Temperature dependence of single-salt parameters -----	181
C.5 Summary of analytical expressions for temperature (and pressure) dependence of selected single-salt Pitzer interaction parameters -----	184
C.6 Sources of data in the file PITZER.DATA -----	189
D Listing of source code to program PHRQPITZ -----	194
E Listing of source code to the interactive input program PITZINPT -----	261

F	Listing of the file MINERALS.2.DATA read by interactive input program PITZINPT -----	303
G	Example of use of interactive input program PITZINPT to , construct the input file for test problem 3 -----	306

ILLUSTRATIONS

Figure 1.	Comparison of calculated solubility of nahcolite in aqueous solutions of Na_2CO_3 -----	9
-----------	---	---

TABLES

Table 1.	Temperature dependence of equilibria in program PHRQPITZ -----	7
2.	Comparison of two internally consistent sets of Pitzer interaction parameters for the system $\text{NaHCO}_3\text{-Na}_2\text{CO}_3\text{-H}_2\text{O}$ at 25 °Celsius -----	10
3.	Sea water in equilibrium with aragonite at a P_{CO_2} of 10^{-3} atmospheres and 25 °Celsius on two different activity-coefficient scales -----	12
4.	Comparison of individual-ion activities using the measured pH -----	13
5.	Analytical data for Canadian Shield brine T-93 -----	34
6.	Line images of input file to test problem 1 -----	35
7.	Listing of printout from test problem 1 -----	39
8.	Line images of input file to test problem 2 -----	60
9.	Listing of printout from test problem 2 -----	62
10.	Line images of input file to test problem 3 -----	76
11.	Listing of printout from test problem 3 -----	78
12.	Line images of input file to test problem 4 -----	84
13.	Listing of printout from test problem 4 -----	86
14.	Line images of input file to test problem 5 -----	99
15.	Listing of printout from test problem 5 -----	101
16.	Line images of input file to test problem 6 -----	121
17.	Listing of printout from test problem 6 -----	123

DISKETTES (Back Pocket)

- 1 Diskette containing a machine-readable copy of the PHRQPITZ source code, the files COMMON.BLOCKS, PITZER.DATA, and PHRQPITZ.DATA, and input files to test problems 1-6.
- 2 Diskette containing a machine-readable copy of the source code to the interactive input program PITZINPT, and the files MINERALS.2.DATA and PHRQPITZ.DATA.

A COMPUTER PROGRAM INCORPORATING PITZER'S EQUATIONS FOR CALCULATION OF GEOCHEMICAL REACTIONS IN BRINES¹

L. Niel Plummer, D.L. Parkhurst, G.W. Fleming, and S.A. Dunkle

ABSTRACT

The program named PHRQPITZ is a computer code capable of making geochemical calculations in brines and other electrolyte solutions to high concentrations using the Pitzer virial-coefficient approach for activity-coefficient corrections. Reaction-modeling capabilities include calculation of (1) aqueous speciation and mineral-saturation index, (2) mineral solubility, (3) mixing and titration of aqueous solutions, (4) irreversible reactions and mineral-water mass transfer, and (5) reaction path. The computed results for each aqueous solution include the osmotic coefficient, water activity, mineral saturation indices, mean activity coefficients, total activity coefficients, and scale-dependent values of pH, individual-ion activities and individual-ion activity coefficients. A data base of Pitzer interaction parameters is provided at 25 °C (Celsius) for the system: Na-K-Mg-Ca-H-Cl-SO₄-OH-HCO₃-CO₃-CO₂-H₂O, and extended to include largely untested literature data for Fe(II), Mn(II), Sr, Ba, Li, and Br with provision for calculations at temperatures other than 25 °C. An extensive literature review of published Pitzer interaction parameters for many inorganic salts is given. Also described is an interactive input code for PHRQPITZ called PITZINPT.

INTRODUCTION

PHRQPITZ² is a FORTRAN 77³ computer program that makes geochemical calculations in brines and other electrolyte solutions to high concentrations. PHRQPITZ has been adapted from the U.S. Geological Survey geochemical simulation computer code PHREEQE⁴ of Parkhurst and others (1980) in which the aqueous model of PHREEQE has been replaced with the Pitzer virial coefficient approach (Pitzer, 1973; Pitzer and Mayorga, 1973, 1974; Pitzer and Kim, 1974; Pitzer, 1975). The PHRQPITZ code contains most of the reaction-modeling capabilities of the original PHREEQE code, including calculation of (1) aqueous speciation and mineral-saturation index, (2) mineral solubility, (3) mixing or titration of aqueous solutions, (4) irreversible reactions and mineral-water mass transfer, and (5) reaction path. The computed results for each aqueous solution include the osmotic coefficient, water activity, mineral saturation indices, mean-activity coefficients, total-activity coefficients, and scale-dependent values of pH, individual-ion activities and individual-ion activity coefficients.

The Pitzer treatment of the aqueous model is based largely on the equations as presented by Harvie and Weare (1980) and Harvie and others (1984). An expanded data base of Pitzer interaction parameters is provided that is identical to the partially validated data base of Harvie and others (1984) at 25 °C (Celsius) for the system Na-K-Mg-Ca-H-Cl-SO₄-OH-HCO₃-CO₃-CO₂-H₂O, and extended to include largely untested literature data for Fe(II), Mn(II), Sr, Ba, Li, and Br with provision for calculations at temperatures other than 25 °C. An extensive

¹ Manuscript approved for publication October 5, 1988.

² PH-Redox-eQuilibrium-equations incorporating the PITZER equations.

³ The use of trade names in this report is for identification purposes only and does not constitute endorsement by the U.S. Geological Survey.

⁴ PH-REdox-EQuilibrium-Equations

literature review of published Pitzer interaction parameters for many inorganic salts is also given that may serve as a guide in selection of additional data for inclusion in PHRQPITZ. Some new data for the temperature dependence of mineral equilibrium constants accompanies the additional (untested) data. As with PHREEQE, the aqueous model and thermodynamic data of PHRQPITZ are user-definable and external to the code.

This report also describes an interactive input code for PHRQPITZ called PITZINPT, which is analogous to the PHREEQE input code, PHRQINPT (Fleming and Plummer, 1984). PITZINPT contains the mineral thermodynamic data base taken largely from Harvie and others (1984) and is used interactively to construct input data sets to PHRQPITZ.

Because most modeling aspects of PHRQPITZ are identical to the original PHREEQE treatment, the reader is referred to the PHREEQE documentation (Parkhurst and others, 1980) for background and modeling information. The current report focuses on extensions to the PHREEQE code including documentation of the Pitzer data base, explanation of the Pitzer equations as incorporated in PHRQPITZ, program limitations and instructions for use.

PITZER EQUATIONS

The osmotic coefficient, ϕ , and activity coefficients of the cations, γ_M , and anions, γ_X , are given by Harvie and Weare, 1980; Harvie and others, 1984),

$$\begin{aligned}
 (\phi - 1) = & \frac{2}{\left(\sum_i m_i\right)} - \frac{A^\phi I^{3/2}}{1 + b I^{1/2}} + \sum_c \sum_a m_c m_a (B_{ca}^\phi + Z C_{ca}) \\
 & + \sum_{c < c'} \sum_{c'} m_c m_{c'} \left(\Phi_{cc'}^\phi + \sum_a m_a \psi_{cc'a} \right) \\
 & + \sum_{a < a'} \sum_{a'} m_a m_{a'} \left(\Phi_{aa'}^\phi + \sum_c m_c \psi_{aa'c} \right) \quad (1)
 \end{aligned}$$

$$\begin{aligned}
 \ln \gamma_M = & z_M^2 F + \sum_a m_a (2B_{Ma} + Z C_{Ma}) + \sum_c m_c \left(2\Phi_{Mc} + \sum_a m_a \psi_{Mc a} \right) \\
 & + \sum_{a < a'} \sum_{a'} m_a m_{a'} \psi_{aa'M} + |z_M| \sum_c \sum_a m_c m_a C_{ca} \quad , \quad (2)
 \end{aligned}$$

and

$$\begin{aligned}
 \ln \gamma_X = & z_X^2 F + \sum_c m_c (2B_{cX} + Z C_{cX}) + \sum_a m_a \left(2\Phi_{Xa} + \sum_c m_c \psi_{Xac} \right) \\
 & + \sum_{c < c'} \sum_{c'} m_c m_{c'} \psi_{cc'X} + |z_X| \sum_c \sum_a m_c m_a C_{ca} \quad . \quad (3)
 \end{aligned}$$

Equations 1-3 use the notation of Harvie and Weare (1980). In equations 1-3 m_i denotes molality of the i th ion (moles per kilogram) where the subscripts M, c, and c' denote cations, and X, a, and a' denote anions. The double summations $c < c'$ and $a < a'$ refer to all pairs of dissimilar cations and anions. The term A^ϕ is defined by

$$A^\phi = \frac{1}{3} (2\pi N_o \rho_w / 1000)^{1/2} (e^2 / D k T)^{3/2} \quad , \quad (4)$$

where N_0 is Avagadro's number, ρ_w is the density of water, e is the absolute electronic charge, k the Boltzman constant, D the static dielectric constant of pure water and T is temperature in Kelvins. A^ϕ then becomes

$$A^\phi = 1400684 \left(\frac{\rho_w}{DT} \right)^{3/2} . \quad (5)$$

In PHRQPITZ, values of A^ϕ are computed over the temperature range 0-350 °C. The total pressure is taken to be 1 atm. (atmosphere) between 0 and 100 °C and that of the vapor pressure curve for pure water of Haar and others (1984) beyond 100 °C. The dielectric constant of pure water is calculated from Bradley and Pitzer (1979). Values of A^ϕ are reported by Bradley and Pitzer (1979) to three significant figures between 0 and 350 °C and are identical to those calculated in PHRQPITZ. Between 0 and 100 °C at 1 atm. total pressure, A^ϕ calculated in PHRQPITZ agrees with values calculated by Ananthaswamy and Atkinson (1984) within 0.00004 or better. The computed value of A^ϕ from equation 5 in PHRQPITZ at 25 °C is 0.39148, which compares with 0.39145 reported by Ananthaswamy and Atkinson (1984) and 0.391 reported by Bradley and Pitzer (1979). Uncertainties beyond the third significant figure in A^ϕ lead to differences in thermodynamic calculations that are beyond the reliability of the parameterization of the model. However, to avoid introducing inconsistency with the Harvie and others (1984) data base, A^ϕ is defined to be 0.392 at 25 °C and 1 atm. total pressure in PHRQPITZ (Harvie and Weare, 1980). In their later paper (Harvie and others, 1984) these authors state that A^ϕ is 0.39 at 25 °C although calculations with their data base indicate it is consistent with the value 0.392. Because A^ϕ is defined to be 0.392 at 25 °C, small inconsistencies in calculations with PHRQPITZ may be observed between results at 25 °C and those very near 25 °C. Uncertainty in the value of A^ϕ is often a source of inconsistency in selecting literature data of Pitzer parameters.

The ionic strength, I , is given by

$$I = \frac{1}{2} \sum_i m_i z_i^2 , \quad (6)$$

where z_i is the charge of the i th ion. Because few ion pairs are considered in the Pitzer treatment, values of ionic strength computed for a given water sample tend to be larger in the Pitzer model than would be calculated using an ion-pairing model.

The term F in equations 2 and 3 is defined by

$$F = -A^\phi \left[\frac{\sqrt{I}}{1+b\sqrt{I}} + \frac{2}{b} \ln(1+b\sqrt{I}) \right] + \sum_c \sum_a m_c m_a B'_{ca} \\ + \sum_{c < c'} \sum_{c'} m_c m_{c'} \Phi'_{cc'} + \sum_{a < a'} \sum_{a'} m_a m_{a'} \Phi'_{aa'} \quad (7)$$

where b is 1.2. The parameters $\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, and C^ϕ that define the variables B and C are fitted from single-salt data. For any salt containing a monovalent ion

$$B_{MX}^\phi = \beta_{MX}^{(0)} + \beta_{MX}^{(1)} e^{-\alpha\sqrt{I}} \quad (8)$$

$$B_{MX} = \beta_{MX}^{(0)} + \beta_{MX}^{(1)} g(\alpha\sqrt{I}) \quad (9)$$

$$B'_{MX} = \beta_{MX}^{(1)} g'(\alpha\sqrt{I}) / I , \quad (10)$$

where $\alpha=2$ (Pitzer, 1973). The functions g and g' are

$$g(x) = 2[1 - (1+x)e^{-x}]/x^2 \quad (11)$$

$$g'(x) = -2\left[1 - \left(1+x + \frac{1}{2}x^2\right)e^{-x}\right]/x^2, \quad (12)$$

where $x = \alpha\sqrt{I}$. For 2-2 electrolytes and higher valence types

$$B_{MX}^\Phi = \beta_{MX}^{(0)} + \beta_{MX}^{(1)}e^{-\alpha_1\sqrt{I}} + \beta_{MX}^{(2)}e^{-\alpha_2\sqrt{I}} \quad (13)$$

$$B_{MX} = \beta_{MX}^{(0)} + \beta_{MX}^{(1)}g(\alpha_1\sqrt{I}) + \beta_{MX}^{(2)}g(\alpha_2\sqrt{I}) \quad (14)$$

$$B'_{MX} = \beta_{MX}^{(1)}g'(\alpha_1\sqrt{I})/I + \beta_{MX}^{(2)}g'(\alpha_2\sqrt{I})/I, \quad (15)$$

where, for 2-2 electrolytes, $\alpha_1=1.4$ and $\alpha_2=12.0$, and for 3-2 and 4-2 electrolytes $\alpha_1=2.0$ and $\alpha_2=50$. (Pitzer and Silvester, 1978). The other variable used to define the thermodynamic properties of single-salt solutions, C_{MX} , is given by

$$C_{MX} = C_{MX}^\Phi / (2\sqrt{|z_M z_X|}) \quad (16)$$

and the coefficient to C_{MX} , Z , in equations 1 - 3 is given by

$$Z = \sum_i m_i |z_i| \quad (17)$$

The thermodynamic properties of aqueous solutions containing a single salt depend only on the interaction parameters $\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, and C^Φ .

The parameters Φ and ψ are determined from aqueous mixtures of two salts. Φ accounts for cation-cation and anion-anion interactions while the parameter ψ is defined for cation-cation-anion and anion-anion-cation interactions. Values of Φ_{ij} are given by

$$\Phi_{ij}^\Phi = \theta_{ij} + {}^E\theta_{ij}(I) + I {}^E\theta'_{ij}(I) \quad (18)$$

$$\Phi_{ij} = \theta_{ij} + {}^E\theta_{ij}(I) \quad (19)$$

$$\Phi'_{ij} = {}^E\theta'_{ij}(I), \quad (20)$$

where θ_{ij} is the only adjustable parameter and is defined for each pair of cations and each pair of anions. The terms ${}^E\theta_{ij}(I)$ and ${}^E\theta'_{ij}(I)$ account for electrostatic mixing effects of unsymmetrical cation-cation and anion-anion pairs as defined by Pitzer (1975). The higher-order electrostatic terms of equations 18-20 are calculated routinely in PHRQPITZ for all unsymmetrical pairs of cations or unsymmetrical pairs of anions using the Chebyshev approximation to the

integrals $J_0(x)$ and $J_1(x)$ (see Pitzer, 1975; Harvie and Weare, 1980). Test calculations showed, however, little difference from more simplified approximations to $J_0(x)$ and $J_1(x)$ given by Pitzer (1975). Values of $E\theta_{ij}(I)$ and $E\theta'_{ij}(I)$ depend only on ion charge and total ionic strength and are zero when ij cation or anion pairs have the same charge.

Caution should be exercised in using literature values of θ_{ij} (Reardon and Armstrong, 1987). Values of θ_{ij} must be compatible with the same single-salt data ($\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, and $C\phi$) used in the model and, for use in PHRQPITZ, their determination from mixed-salt solutions must include the higher-order electrostatic terms discussed above. Both types of θ_{ij} are reported in the literature (i.e., determined with and without provision for higher-order electrostatic terms). In PHRQPITZ the higher-order electrostatic terms are always included. This precaution applies to cation or anion pairs such as $\text{Ca}^{2+}\text{-Na}^+$ or $\text{SO}_4^{2-}\text{-Cl}^-$, but not to interactions such as $\text{Cl}^-\text{-F}^-$ or $\text{Ca}^{2+}\text{-Mg}^{2+}$ where, because of the identical charge, the higher-order electrostatic terms are zero. Values of θ_{ij} given by Pitzer and Mayorga (1974) and Pitzer (1979) do not include the higher-order electrostatic terms for unsymmetrical cation pairs and unsymmetrical anion pairs. The data of Harvie and Weare (1980) and Harvie and others (1984) include the higher-order electrostatic terms. Harvie and Weare (1980) found significant improvement in modeling the system $\text{Na-K-Mg-Ca-Cl-SO}_4\text{-H}_2\text{O}$ when the higher-order electrostatic terms were included.

Values of the parameters ψ_{ijk} are included for all different combinations of two cations and an anion or two anions and a cation. ψ is usually determined from the same two-salt mixture used to define θ_{ij} and is therefore internally consistent with that value of θ_{ij} as well as the individual single-salt interaction parameters.

Harvie and others (1984) have also included a model for calculation of the activity coefficients, γ_N , of neutral species in solution. In the Harvie and others (1984) data base, this calculation applies to the species $\text{CO}_2(\text{aq})$, CaCO_3° and MgCO_3° . According to Harvie and others (1984), γ_N is calculated from the relation

$$\ln \gamma_N = \sum_c m_c (2\lambda_{nc}) + \sum_a m_a (2\lambda_{na}) \quad , \quad (21)$$

where λ_{nc} and λ_{na} refer to interactions between neutral species and cations or anions.

DATA BASE

Two data files are required to run PHRQPITZ. The first file, PHRQPITZ.DAT, is analogous to the thermodynamic data base of PHREEQE (though much smaller) and contains data under the keywords ELEMENTS, SPECIES, LOOK MIN and MEAN GAM. The first three of these keywords are identical to their usage in PHREEQE (Parkhurst and others, 1980). The keyword MEAN GAM is used in PHRQPITZ to define the stoichiometries of selected salts for calculation of the mean activity coefficient, γ_\pm . The mean activity coefficient is defined

$$\gamma_\pm = \left(\gamma_+^{\nu_+} \gamma_-^{\nu_-} \right)^{\frac{1}{\nu_+ + \nu_-}} \quad , \quad (22)$$

where γ_+ and γ_- denote total-activity coefficients of cations (+) and anions (-) and ν_+ and ν_- are the stoichiometric coefficients of the cation and anion in the neutral salt.

Equilibrium constants at 25 °C for (1) the formation of OH^- , HCO_3^- , $\text{CO}_2^*(\text{aq})$, HSO_4^- , CaCO_3° , MgOH^+ , and MgCO_3° and (2) dissolution of minerals listed under LOOK MIN in PHRQPITZ.DAT are computed from the free energies given by Harvie and others (1984).

Expressions for the temperature dependence of $\log K$ are of the form,

$$\log K = A_1 + A_2 T + A_3 / T + A_4 \log T + A_5 / T^2 \quad , \quad (23)$$

where $A_1 - A_5$ are constants and T is temperature in Kelvins. The form of equation 23 corresponds to the model for the heat capacity of Maier and Kelly (1932), as used in PHREEQE. Expressions of this form were either taken from the literature or fitted to solubility data reported by Linke (1965). In all cases the A_1 term has been adjusted to agree with the Harvie and others (1984) data base at 25 °C. In some cases the temperature dependence of $\log K$ is calculated from the van't Hoff equation using a value of ΔH_r° at 25 °C. No data for the temperature dependence of the equilibrium constant are available for many of the minerals in the PHRQPITZ.DAT file and for these, PHRQPITZ computes the same value of the equilibrium constant at all temperatures. Table 1 summarizes data used to calculate the temperature dependence of the equilibria in PHRQPITZ. Line images of the file PHRQPITZ.DAT are listed in Attachment A. Coding format and description of input variables in PHRQPITZ.DAT are given in a later section of this report under "Description of Input".

The second data file read by PHRQPITZ is named PITZER.DAT and contains values of the interaction parameters to the Pitzer equations including values of $\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, C^ϕ , θ , λ , ψ , and limited data on the temperature dependence of $\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, and C^ϕ . The file PITZER.DAT is listed in Attachment B. For the chemical system Na-K-Mg-Ca-H-Cl-SO₄-OH-HCO₃-CO₃-CO₂-H₂O at 25 °C, the data bases to PHRQPITZ (PHRQPITZ.DAT and PITZER.DAT) are identical to that of Harvie and others (1984), and in verification procedures have reproduced calculations reported in Harvie and others (1984).

Extensions of the Data Base

Extensions beyond the Harvie and others (1984) data base are largely untested and include additions to PITZER.DAT for (1) calculation of the thermodynamic properties of aqueous solutions containing, in addition to the elements considered in the Harvie and others (1984) data base, Fe(II), Mn(II), Sr²⁺, Ba²⁺, Li⁺, and Br⁻; (2) estimation of the temperature dependence of many of the single-salt parameters from selected literature data for the first derivative with respect to temperature; and (3) calculation of the thermodynamic properties of NaCl solutions to approximately 300 °C following the vapor pressure curve of water beyond 100 °C. Except for the NaCl-H₂O system, the PHRQPITZ aqueous model should be checked carefully before applications outside the temperature interval 0 to 60 °C are attempted. Several recent evaluations of the temperature dependence of Pitzer interaction parameters to relatively high temperatures (Pitzer, 1987; Moller, 1988) have not yet been incorporated in the PHRQPITZ data base.

Extensions to PHRQPITZ.DAT beyond that of Harvie and others (1984) include estimates of ΔH_r° for the formation of the aqueous non-master species such as OH⁻, HCO₃⁻, MgOH⁺, etc., and calculation of equilibrium constants for mineral dissolution reactions at temperatures other than 25 °C. Values of ΔH_r° have been taken from the published literature, or, in some cases, calculated from speciation of mineral solubility data as a function of temperature such as from Linke (1965). The temperature dependence of the interaction parameters and ΔH_r° is only partially complete. No data exist for the temperature dependence of the solubility of many minerals included in PHRQPITZ.DAT. In all cases, the data to PHRQPITZ reproduce the Harvie and others (1984) data base at 25 °C. If changes in the temperature-dependence of the Pitzer interaction parameters are made to the file PITZER.DAT, the appropriate mineral equilibrium constants and their temperature dependence will need to be examined in the file PHRQPITZ.DAT for consistency.

Internal Consistency of the Data Base

It is important to stress the need to maintain the internal consistency of the Harvie and others (1984) data base. As an example, consider the solubility of nahcolite in Na₂CO₃ solutions. The results are given in Harvie and others (1984) (see their fig. 7c, p.734). Curve 1 of figure 1 shows the solubility of nahcolite in Na₂CO₃ solutions calculated from PHRQPITZ using the Harvie and others (1984) data base. The results are identical to those of Harvie and others (1984) and, of course, adequately reproduce the original experimental data. We will now examine the consequences of changing the Harvie and others (1984) data base as regards the solubility of nahcolite.

Table 1 -- Temperature dependence of equilibria in program PHRQPITZ
[--, no data; °C, degrees Celsius; kcal/mol, kilocalories per mole]

Reaction	Mineral name	$\log K(1)$ 25 °C	ΔH_f° 25 °C (kcal/mol)	A1 (1)	A2 (1)	A3 (1/1)	A4 (log T)	A5 (1/12)	Source [1]
<u>Aqueous Species</u>									
31 $H_2O = H^+ + OH^-$	--	-13.997	13.345	--	--	--	--	--	2
34 $CO_3^{2-} + H^+ = HCO_3^-$	--	10.339	--	107.8975	0.03252849	-5151.79	-38.92561	563713.9	3
35 $CO_3^{2-} + 2H^+ = CO_2(aq) + H_2O$	--	16.677	--	464.1925	.09344813	-26986.16	-165.75951	2248628.9	3
40 $SO_4^{2-} + H^+ = HSO_4^-$	--	1.979	--	-5.3585	.0183412	557.2461	--	--	4
76 $Ca^{2+} + CO_3^{2-} = CaCO_3^o$	--	3.151	--	-1228.806	-.299440	35512.75	485.818	--	3
85 $Mg^{2+} + H_2O = MgOH^+ + H^+$	--	-11.809	15.419	--	--	--	--	--	2
86 $Mg^{2+} + CO_3^{2-} = MgCO_3^o$	--	2.928	--	-32.225	--	1093.486	12.72433	--	3
<u>Solids</u>									
$CaSO_4 = Ca^{2+} + SO_4^{2-}$	Anhydrite	-4.362	--	422.950	--	-18431.	-147.708	--	5
$CaCO_3 = Ca^{2+} + CO_3^{2-}$	Aragonite	-8.220	--	-171.8607	-.077993	2903.293	71.595	--	3
$K_2SO_4 = 2K^+ + SO_4^{2-}$	Arcanite	-1.776	--	2.823	--	-1371.2	--	--	5
$MgCl_2 \cdot 6H_2O = Mg^{2+} + 2Cl^- + 6H_2O$	Bischofite	4.455	--	3.524	--	277.6	--	--	5
$Na_2Mg(SO_4)_2 \cdot 4H_2O = 2Na^+ + Mg^{2+} + 2SO_4^{2-} + 4H_2O$	Bloedite	-2.347	--	--	--	--	--	--	1
$Mg(OH)_2 = Mg^{2+} + 2OH^-$	Brucite	-10.884	.85	--	--	--	--	--	6
$Na_6CO_3(SO_4)_2 = 6Na^+ + CO_3^{2-} + 2SO_4^{2-}$	Burkeite	-.772	--	--	--	--	--	--	1
$CaCO_3 = Ca^{2+} + CO_3^{2-}$	Calcite	-8.406	--	-171.8329	-.077993	2839.319	71.595	--	3
$KMgCl_3 \cdot 6H_2O = K^+ + Mg^{2+} + 3Cl^- + 6H_2O$	Carnallite	4.330	--	--	--	--	--	--	1
$CaMg(CO_3)_2 = Ca^{2+} + Mg^{2+} + 2CO_3^{2-}$	Dolomite	-17.083	-9.436	--	--	--	--	--	2
$MgSO_4 \cdot 7H_2O = Mg^{2+} + SO_4^{2-} + 7H_2O$	Epsomite	-1.881	--	1.718	--	-1073.	--	--	5
$CaNa_2(CO_3)_2 \cdot 5H_2O = Ca^{2+} + 2Na^+ + 2CO_3^{2-} + 5H_2O$	Gaylussite	-9.421	--	--	--	--	--	--	1
$NaK_3(SO_4)_2 = Na^+ + 3K^+ + 2SO_4^{2-}$	Glaserite	-3.803	--	--	--	--	--	--	1
$Na_2Ca(SO_4)_2 = 2Na^+ + Ca^{2+} + 2SO_4^{2-}$	Glauberite	-5.245	--	--	--	--	--	--	1
$CaSO_4 \cdot 2H_2O = Ca^{2+} + SO_4^{2-} + 2H_2O$	Gypsum	-4.581	--	90.318	--	-4213.	-32.641	--	5
$NaCl = Na^+ + Cl^-$	Halite	1.570	--	-713.4616	-.1201241	37302.21	262.4583	-2106915.	7
$MgSO_4 \cdot 6H_2O = Mg^{2+} + SO_4^{2-} + 6H_2O$	Hexahydrate	-1.635	--	-62.666	--	1828.	22.187	--	5
$KMgClSO_4 \cdot 3H_2O = K^+ + Mg^{2+} + Cl^- + SO_4^{2-} + 3H_2O$	Kainite	-.193	--	--	--	--	--	--	1
$KHCO_3 = K^+ + H^+ + CO_3^{2-}$	Kalicinite	-10.058	--	--	--	--	--	--	1
$MgSO_4 \cdot H_2O = Mg^{2+} + SO_4^{2-} + H_2O$	Kieserite	-.123	--	--	--	--	--	--	1
$Na_4Ca(SO_4)_3 \cdot 2H_2O = 4Na^+ + Ca^{2+} + 3SO_4^{2-} + 2H_2O$	Labile Salt	-5.672	--	--	--	--	--	--	1
$MgSO_4 \cdot 4H_2O = Mg^{2+} + SO_4^{2-} + 4H_2O$	Leonhadite	-.887	--	--	--	--	--	--	1

Table 1 -- Temperature dependence of equilibria in program PHRQPITZ (continued)
[--, no data; °C, degrees Celsius; kcal/mol, kilocalories per mole]

Reaction	Mineral name	$\log K[1]$ 25 °C	ΔH_f° 25 °C (kcal/mol)	A ₁ (1)	A ₂ (T)	A ₃ (1/T)	A ₄ (log T)	A ₅ (1/T ²)	Source [1]
$K_2Mg(SO_4) \cdot 2 \cdot 4H_2O = 2K^+ + Mg^{2+} + 2SO_4^{2-} + 4H_2O$	Leonite	-3.979	--	--	--	--	--	--	1
$MgCO_3 = Mg^{2+} + CO_3^{2-}$	Magnesite	-7.834	-6.169	--	--	--	--	--	6
$Na_2SO_4 \cdot 10H_2O = 2Na^+ + SO_4^{2-} + 10H_2O$	Mirabilite	-1.214	--	-3862.234	-1.19856	93713.54	1577.756	--	8
$K_8H_6(SO_4)_7 = 8K^+ + 6H^+ + 7SO_4^{2-}$	Misenite	-10.806	--	--	--	--	--	--	1
$NaHCO_3 = Na^+ + H^+ + CO_3^{2-}$	Nahcolite	-10.742	--	--	--	--	--	--	1
$Na_2CO_3 \cdot 10H_2O = 2Na^+ + CO_3^{2-} + 10H_2O$	Natron	-8.25	--	--	--	--	--	--	1
$MgCO_3 \cdot 3H_2O = Mg^{2+} + CO_3^{2-} + 3H_2O$	Nesquehonite	-5.167	--	--	--	--	--	--	1
$CO_2(g) = CO_2(aq)$	PCO ₂	-1.468	--	108.3865	.01985076	-6919.53	-40.45154	669365.	3
$MgSO_4 \cdot 5H_2O = Mg^{2+} + SO_4^{2-} + 5H_2O$	Pentahydrate	-1.285	--	--	--	--	--	--	1
$Na_2Ca(CO_3)_2 \cdot 2H_2O = 2Na^+ + Ca^{2+} + 2CO_3^{2-} + 2H_2O$	Pirssonite	-9.234	--	--	--	--	--	--	1
$K_2MgCa_2(SO_4)_4 \cdot 2H_2O = 2K^+ + Mg^{2+} + 2Ca^{2+} + 4SO_4^{2-} + 2H_2O$	Polyhalite	-13.744	--	--	--	--	--	--	1
$Ca(OH)_2 = Ca^{2+} + 2OH^-$	Portlandite	-5.190	--	--	--	--	--	--	1
$K_2Mg(SO_4)_2 \cdot 6H_2O = 2K^+ + Mg^{2+} + 2SO_4^{2-} + 6H_2O$	Schoenite	-4.328	--	--	--	--	--	--	1
$KCl = K^+ + Cl^-$	Sylvite	.900	--	3.984	--	-919.55	--	--	5
$K_2Ca(SO_4)_2 \cdot H_2O = 2K^+ + Ca^{2+} + 2SO_4^{2-} + H_2O$	Syngenite	-7.448	--	--	--	--	--	--	1
$Na_3H(CO_3)_2 \cdot 2H_2O = 3Na^+ + H^+ + 2CO_3^{2-} + 2H_2O$	Trona	-11.384	--	--	--	--	--	--	1

[1] References to values of $\log K_T$:

1. Harvie and others (1984) unless otherwise indicated.
2. Parkhurst and others (1980).
3. Plummer and Busenberg (1982) with A₁ adjusted for consistency with (1).
4. Parkhurst and others (1980) with A₁ Adjusted for consistency with (1).
5. Solubility data from Linke (1965), speciated in PHRQPITZ, $\log K$ fitted to analytical expression as $f(T)$ with A₁ adjusted to be consistent with (1).
6. Plummer and others (1976).
7. Halite solubility from Pitzer and others (1984) treated as in 5 (above).
8. Solubility data from Linke (1965) and Plummer and Busenberg (unpub. data, 1987). The revised value of $\log K$ at 25 °C for mirabilite (-1.214) compares to -1.228 of Harvie and others (1984).

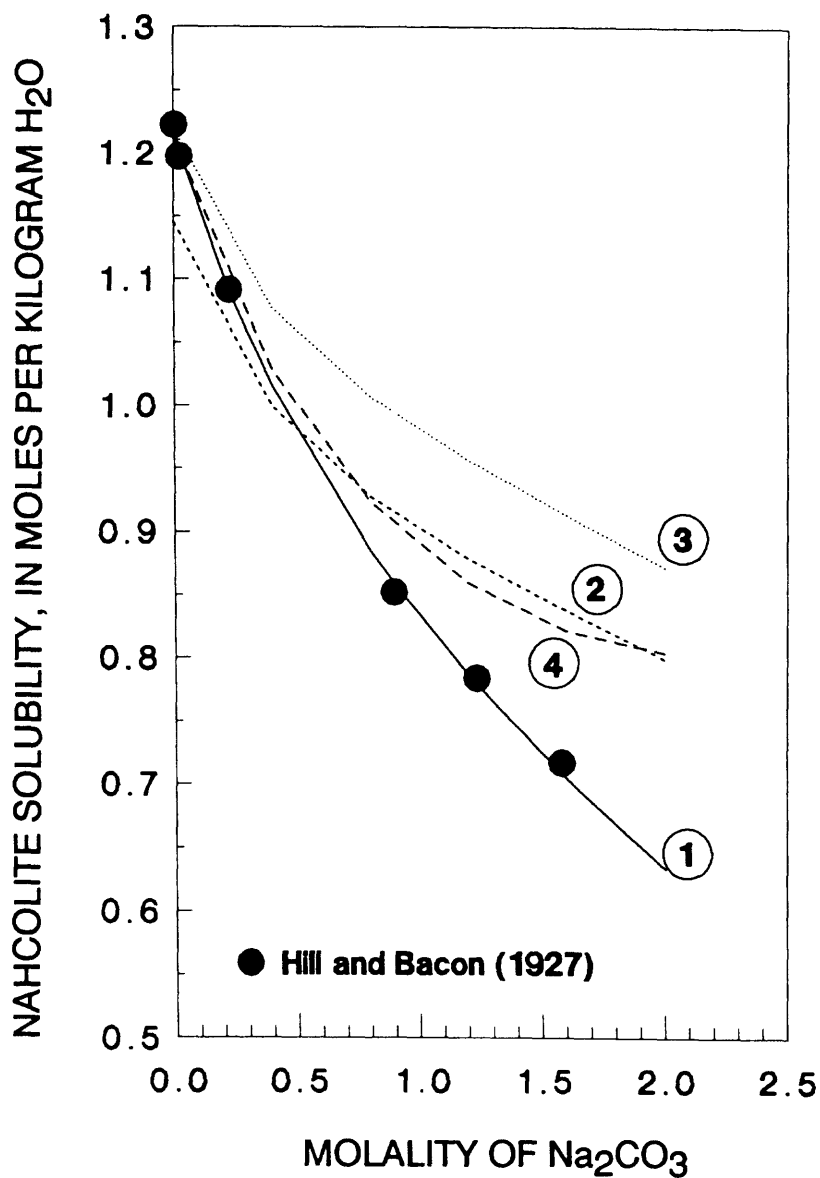


Figure 1 - Comparison of calculated solubility of nahcolite in aqueous solutions of Na₂CO₃ to 2 molal. (1) Calculated using the model of Harvie and others (1984), (2) Parameters β^0 , β^1 , and C^ϕ of NaHCO₃ changed to those of Sarbar and others (1982), (3) Same as curve 2 but with $K_{\text{Nahcolite}}$ changed to be consistent with Sarbar and others (1982), and (4) Same as curve 3 but with θ and ψ for NaHCO₃-Na₂CO₃ mixtures changed to the values of Roy and others (1984) reported for KHCO₃-K₂CO₃ mixtures. Solid points are the experimental data of Hill and Bacon (1927).

The $\text{Na}^+\text{-HCO}_3^-$ interaction parameters in the Harvie and others (1984) data base were originally taken from Pitzer and Peiper (1980) based on an evaluation of the original electrochemical measurements of Harned and Davis (1943) and Harned and Bonner (1945) in $\text{NaHCO}_3\text{-NaCl}$ aqueous solutions. Alternatively we have the Pitzer interaction parameters for NaHCO_3 of Sarbar and others (1982) determined from isopiestic measurements of mixed aqueous solutions of NaHCO_3 and Na_2CO_3 at 25 °C. Curve 2 of figure 1 shows the resulting calculated solubility of nahcolite in Na_2CO_3 solutions if the single-salt parameters of NaHCO_3 are arbitrarily changed from those of Pitzer and Peiper (1980) to those of Sarbar and others (1982). Curve 3 of figure 1 retains the Sarbar and others (1982) parameters for NaHCO_3 and adjusts the equilibrium constant of nahcolite to be internally consistent with the Sarbar and others (1982) parameters. Still the agreement is poor with the Harvie and others (1984) calculation (and experimental data).

One final adjustment is to change θ and ψ for $\text{CO}_3^{2-} - \text{HCO}_3^- - \text{Na}^+$ interactions. Because no other values of θ and ψ are known for $\text{NaHCO}_3\text{-Na}_2\text{CO}_3$ mixtures, we have substituted values of θ and ψ determined by Roy and others (1984) for $\text{KHCO}_3\text{-K}_2\text{CO}_3\text{-KCl}$ aqueous salt mixtures. Curve 4 of figure 1 shows marked improvement, but further adjustments in these parameters are clearly needed to obtain a revised data set for the $\text{NaHCO}_3\text{-Na}_2\text{CO}_3$ system consistent with the data of Sarbar and others (1982).

New values of θ and ψ for $\text{CO}_3^{2-} - \text{HCO}_3^- - \text{Na}^+$ interactions that are internally consistent with the NaHCO_3 data of Sarbar and others (1982) and the known solubility of nahcolite in Na_2CO_3 solutions (Hill and Bacon, 1927) are compared with the Harvie and others (1984) data base in table 2. Both sets of parameters in table 2 will reproduce the experimental data for the solubility of nahcolite in Na_2CO_3 solutions (curve 1, fig. 1). This example demonstrates the need for internal consistency of all Pitzer interaction parameters and thermodynamic data for the given chemical system when changes to the Harvie and others (1984) data base are contemplated.

Table 2. -- Comparison of two internally consistent sets of Pitzer interaction parameters for the system $\text{NaHCO}_3\text{-Na}_2\text{CO}_3\text{-H}_2\text{O}$ at 25 °Celsius ¹

Harvie and others (1984)	Alternate parameter set
NaHCO_3 ²	NaHCO_3 ³
$\beta^\circ = 0.0277$	$\beta^\circ = -0.04096$
$\beta^1 = .0411$	$\beta^1 = .5062$
$C^\phi = 0$	$C^\phi = .005250$
Na_2CO_3 ⁴	Na_2CO_3 ⁴
$\beta^\circ = .0399$	$\beta^\circ = .0399$
$\beta^1 = 1.389$	$\beta^1 = 1.389$
$C^\phi = .0044$	$C^\phi = .0044$
$\theta_{\text{CO}_3\text{-HCO}_3} = -.04$	$\theta_{\text{CO}_3\text{-HCO}_3} = .111$
$\psi_{\text{Na-CO}_3\text{-HCO}_3} = .002$	$\psi_{\text{Na-CO}_3\text{-HCO}_3} = -.025$
$\text{Log } K_{\text{Nahcolite}} = -10.742$	$\text{Log } K_{\text{Nahcolite}} = -10.696$

¹ $A^\phi = .392$. Higher-order electrostatic terms included in calculation of θ and ψ .

² Pitzer and Peiper (1980).

³ Sarbar and others (1982).

⁴ Harvie and others (1984).

Literature Sources of Pitzer Interaction Parameters

As an aid to future improvement of the Pitzer data base, and guide to interpretation of uncertainties of model results, a computer search of literature values of single and mixed inorganic salt data was conducted through March, 1988. The results are summarized in Attachment C, where Attachment C.1 lists a bibliography of 131 articles from which the data were selected. Attachment C.2 lists published values found for the single-salt parameters $\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, and C^ϕ . Attachment C.3 summarizes data for binary salt systems, θ and ψ . The asterisks assigned to values of θ on Attachment C.3 denote parameters for which higher-order electrostatic terms were not included in fitting to experimental data. Because PHRQPITZ automatically includes the higher-order electrostatic terms, these mixed-salt parameters cannot be used in PHRQPITZ without modification of the code. Attachment C.3 also summarizes values of the single-salt parameters used to define the mixed-salt parameters. Values of θ and ψ from Attachment C.3 should not be used in PHRQPITZ unless they are consistent with the same single-salt parameters found in PITZER.DATA.

Attachment C.4 summarizes published values of the temperature derivative of the single-salt parameters. Additional temperature-dependent data are given in Attachment C.5. Numbers in parentheses on Attachments C.2-C.5 refer to bibliographic citations listed in Attachment C.1. Several advanced formulations of the temperature dependence of single-salt parameters have been proposed for NaCl, CaCl₂, NaOH, KCl, CsCl, HCl, LiCl, MgSO₄, Na₂SO₄, K₂SO₄, MgCl₂, and CaSO₄. Attachment C.5 lists these equations, fitted constants and sources of data. The reader is also referred to the recent papers of Pabalan and Pitzer (1987), Pitzer (1987) and Moller (1988) for additional high temperature data.

PHRQPITZ currently calculates values of the single-salt interaction parameters either using the 25 °C value and its first derivative with respect to temperature, or utilizes an expression of the form given by Silvester and Pitzer (1977) for NaCl (only). For more extensive data on NaCl the reader is referred to Pitzer and others (1984) and Weres and others (1987).

Because of our dependence in PHRQPITZ on the Harvie and others (1984) data base and its extensive origins from Pitzer and Mayorga (1973, 1974), and Pitzer and Kim (1974), Attachments C.2-C.4 are arranged to list single-salt parameters at 25 °C in the order: Harvie and others (1984), Harvie and Weare (1980), followed by Pitzer sources, and then the remaining literature. Attachment C.6 lists the PITZER.DATA file read by PHRQPITZ (see Attachment B) indicating sources of parameters in parentheses keyed to citations given in Attachment C.1.

SCALE CONVENTION OF ACTIVITY COEFFICIENTS

As has long been recognized, individual-ion activities and activity coefficients cannot be measured independently. Equations 2 and 3 are valid only when individual-ion activity coefficients are combined to define properties of neutral combinations of ions such as in the calculation of mean activity coefficients, saturation indices, solubility, etc. Therefore, values of individual-ion activities and activity coefficients have meaning only in a relative sense and individual values depend on a particular choice of scale convention. The subject has received recent attention as applies to interpretation of pH in seawater (Bates, 1975; Bates and Culbertson, 1977; Millero, 1979; Millero and Schreiber, 1982; Plummer and Sundquist, 1982; Millero, 1983; Dickson, 1984; Covington, Bates and Durst, 1985).

PHRQPITZ offers two scaling conventions based on the work of Harvie and others (1984). In the first case (IOPT(10)=0, see Description of Input) no scaling is performed and individual-ion activity coefficients are as computed by equations 2 and 3. In the second case, all individual-ion activity coefficients are scaled according to the MacInnes (1919) convention (IOPT(10)=1, see Description of Input). In this case the activity coefficient of Cl is defined to be equal to the mean activity coefficient of KCl in a KCl solution of equivalent ionic strength, $\gamma_{Cl(Mac)} = \gamma_{\pm KCl}$. The scaling factor for the *i*th ion is computed from the term $(\gamma_{Cl(eqn.3)} / \gamma_{\pm KCl})^{z_i}$ and is multiplied through all other individual-ion activity coefficients computed from equations 2 and 3. That is

$$\gamma_{i(Mac)} = \gamma_i (\gamma_{Cl} / \gamma_{\pm KCl})^{z_i} \quad (24)$$

where $\gamma_{i(Mac)}$ is the individual ion activity coefficient of the i th ion converted to the MacInnes scale, γ_i is the activity coefficient of the ion consistent with some other convention, such as equations 2 and 3, γ_{Cl} is the activity coefficient of Cl^- according to the alternate convention, $\gamma_{\pm KCl}$ is the mean activity coefficient of KCl in a pure KCl solution of equivalent ionic strength and z_i is the charge of the i th ion (+ for cations, - for anions). The activity coefficients could be placed on other scales by substituting for $\gamma_{\pm KCl}$ or $\gamma_{Cl}/\gamma_{\pm KCl}$ in equation 24.

As an example, table 3 compares log values of the molality, activity, and activity coefficient of individual ions computed for sea water in equilibrium with aragonite at 25 °C and a CO_2 partial pressure of 10^{-3} atm. The results are presented for the problem computed on the MacInnes scale and without scaling as given by equations 2 and 3. In the problem, pH was calculated from the given equilibria and is therefore internally consistent with the aqueous model and respective scale. For this particular case, when all individual-ion activity coefficients are consistent with a single scale, the individual-ion molalities are independent of choice of scale, while the computed individual-ion activities and activity coefficients are scale-dependent. For example, the pH of sea water in equilibrium with aragonite at 25 °C and 10^{-3} atm. P_{CO_2} (using the Harvie and others (1984) data base) is 7.871 on the MacInnes scale and 7.828 without scaling equations 2 and 3, as shown in table 3.

Table 3. -- Sea water in equilibrium with aragonite at a P_{CO_2} of 10^{-3} atmospheres and 25 °Celsius on two different activity-coefficient scales

Species	MacInnes scale			Unscaled		
	Log molality	Log activity	Log gamma	Log molality	Log activity	Log gamma
H ⁺	-7.694	-7.871	-0.178	-7.694	-7.828	-0.134
H ₂ O	-.008	-.008	.000	-.008	-.008	.000
Ca ²⁺	-1.981	-2.795	-.814	-1.981	-2.708	-.726
Mg ²⁺	-1.259	-2.034	-.775	-1.259	-1.946	-.687
Na ⁺	-.313	-.552	-.239	-.313	-.508	-.195
K ⁺	-1.973	-2.248	-.274	-1.973	-2.204	-.230
Cl ⁻	-.246	-.363	-.118	-.246	-.407	-.162
CO ₃ ²⁻	-4.513	-5.424	-.911	-4.513	-5.512	-.999
SO ₄ ²⁻	-1.532	-2.419	-.887	-1.532	-2.507	-.975
OH ⁻	-5.938	-6.134	-.196	-5.938	-6.178	-.240
HCO ₃ ⁻	-2.781	-2.956	-.176	-2.781	-3.000	-.219
H ₂ CO ₃	-4.535	-4.482	.053	-4.535	-4.482	.053
HSO ₄ ⁻	-8.204	-8.312	-.108	-8.204	-8.356	-.152
CaCO ₃	-5.068	-5.068	.000	-5.068	-5.068	.000
MgOH ⁺	-5.890	-5.980	-.100	-5.890	-5.936	-.056
MgCO ₃	-4.530	-4.530	.000	-4.530	-4.529	.000

Combining the results of table 3 indicates that when different scales are used, which are in themselves internally consistent, all mean quantities such as γ_{\pm} , a_{\pm} , m_{\pm} and other neutral salt combinations such as ion activity products, saturation indices, and therefore solubility, are independent of scale. Clearly then, we will reach the same thermodynamic conclusion for a

given data base regardless of scale when all individual-ion values are internally consistent with a single scale. However, no significance can be assigned to the individual-ion activities, individual-ion activity coefficients and pH in comparing results on differing scales.

The problem of activity-coefficient scale is more important when the measured pH is introduced in geochemical calculations. The measured pH is not likely to be on the same activity-coefficient scale as the aqueous model because the buffers used to define pH are conventional (Bates, 1973). Even if the measured pH were placed on the same scale as the aqueous model, uncertainties in the measurement of pH in brines, such as due to liquid-junction potentials (Bates, 1973; Wescott, 1978) will always introduce inconsistencies. Consequently, it is unlikely that the measured pH will be consistent with the particular scale used for the individual ions.

As an example of this problem, table 4 compares log activities, $\gamma_{\pm\text{CaCO}_3}$, $a_{\pm\text{CaCO}_3}$, SI_{calcite} and log P_{CO_2} for two water samples, including the measured pH (courtesy of R. Spencer, written Comm., 1985) from Great Salt Lake, Utah in both scaled and unscaled calculations. When the measured pH is introduced in the speciation calculation, all individual ion and mean properties, including saturation indices, P_{CO_2} , etc. are scale dependent. This dependency of thermodynamic properties on scale when the measured pH is used is particularly acute to the carbonate system and other chemical systems where the equilibria depend significantly on pH, but is not likely to be important to calculations involving chloride and sulfate minerals in most natural waters.

Table 4. -- Comparison of individual-ion activities using the measured pH^[1]
[--, carbonate alkalinity included in bicarbonate]

Great Salt Lake (Dilute)				Great Salt Lake (Evaporated)			
pH = 8.056				pH = 8.26			
$a_{\text{H}_2\text{O}} = 0.986$				$a_{\text{H}_2\text{O}} = 0.730$			
Ionic Strength = 0.6712				Ionic Strength = 8.5313			
Ion	Molality	Log activity		Ion	Molality	Log activity	
		MacInnes	Unscaled			MacInnes	Unscaled
Na	0.4651	-0.529	-0.521	Na	6.5204	0.613	0.793
K	.0204	-1.915	1.907	K	1.1119	-.462	-.282
Ca	.0096	-2.862	-2.845	Ca	.0011	-3.468	-3.109
Mg	.0170	-2.556	-2.539	Mg	.0009	-3.183	-2.824
Cl	.3263	-.672	-.680	Cl	5.8484	.955	.775
SO ₄	.1058	-1.899	-1.916	SO ₄	.8927	-1.516	-1.876
HCO ₃ ^[2]	.0010	-3.267	-3.276	HCO ₃ ^[2]	.0028	-2.970	-3.172
CO ₃	--	-5.551	-5.560	CO ₃	--	-5.049	-5.252
$\gamma_{\pm\text{CaCO}_3}$.1082	.1082	$\gamma_{\pm\text{CaCO}_3}$.1387	.1386
log $a_{\pm\text{CaCO}_3}$		-4.2063	-4.2023	log $a_{\pm\text{CaCO}_3}$		-4.2585	-4.1802
SI_{Calcite}		-.0064	.0014	SI_{Calcite}		-.1107	.0459
log P_{CO_2}		-3.51	-3.52	log P_{CO_2}		-3.29	-3.49

[1] All calculations at 25°C

[2] Carbonate alkalinity as HCO₃⁻

The problem of inconsistency of the measured pH with the adopted scale is less important at relatively low ionic strengths. For example, for dilute Great Salt Lake (ionic strength = 0.67), the calcite saturation index varies from 0.001 to -0.006 between unscaled and scaled (MacInnes) calculations (table 4.) But at relatively high ionic strengths, such as evaporated Great Salt Lake (ionic strength = 8.53), the inconsistency in pH scale leads to differences in the calculated calcite saturation index of 0.05 to -0.11 between unscaled and scaled (MacInnes)

calculations (table 4). Therefore, in addition to the relatively formidable task of measuring pH in brines, it is important to recognize the magnitude of error that can result simply from differences in activity-coefficient scale.

PRECAUTIONS AND LIMITATIONS

Because it was adapted from PHREEQE, PHRQPITZ retains some of the limitations of the original code. These are discussed in Parkhurst and others (1980) and reviewed here as they apply to geochemical reactions in brines.

All calculations are made relative to one kilogram (kg) of water. As there is no mass balance for the elements H and O, there is no formal provision for keeping track of the amount of water used in reactions such as hydration and dehydration of solids. This may be a source of error in simulation of the evaporation of brines where many of the minerals precipitated are often hydrated and remove water from solution.

PHRQPITZ retains the original logic of PHREEQE concerning oxidation-reduction reactions, but because an internally-consistent data base of Pitzer interaction parameters for multiple oxidation states is not currently available, geochemical redox reactions may not be attempted in PHRQPITZ. All redox equilibria used with PHREEQE have been removed from the PHRQPITZ data file PHRQPITZ.DAT. If PITZINPT is used interactively to construct input sets to PHRQPITZ, the pe is automatically set to 4.0 and the non-redox option is selected (IOPT(5)=0). Consequently, precautions in Parkhurst and others (1980) concerning possible formation of large amounts of H₂ or O₂ gas outside the stability field of H₂O are not relevant in PHRQPITZ if these guidelines are adhered to.

Many of the original convergence problems of PHREEQE were redox-related. Improvements to the convergence criteria of PHREEQE (noted in the January, 1985 version) have been incorporated in PHRQPITZ. Precautions and comments on the use of ion exchange, and titration/mixing reactions in PHREEQE, and the lack of uniqueness of modeled reaction paths, discussed in Parkhurst and others (1980), also apply to PHRQPITZ.

The activity of water, a_{H_2O} , in PHRQPITZ, is computed from the osmotic coefficient, ϕ , given by equation 1, and is

$$\ln a_{H_2O} = - \frac{\phi \sum_i m_i}{55.50837} \quad (25)$$

where m_i is the molality of the i th ion in solution and there are 55.50837 moles of water per kg of water. Equation 25 represents a substantial improvement over water activity calculations in PHREEQE and is now determined directly from the Pitzer model.

Other precautions have been discussed earlier as they apply specifically to the Pitzer data base. These include the likelihood of introducing error in calculated mean activities and saturation indices (particularly for carbonates) when the measured pH is on a different activity coefficient scale than the aqueous model, and the need for internal consistency between mixed-salt and single-salt interaction parameters.

The pH problem is particularly important to carbonate calculations in high ionic strength brines and could easily introduce uncertainties of ± 0.5 in the saturation index of calcite. Even larger errors in SI may be attributed to uncertainties in the measurement of pH in brines due to, for example, liquid-junction potentials (Bates, 1973; Westcott, 1978).

Although we have extended the Harvie and others (1984) data base to include Fe²⁺, Mn²⁺, Sr²⁺, Ba²⁺, Li⁺, and Br⁻, and limited temperature dependence, the data base has not been validated beyond that of Harvie and others (1984) at 25 °C. Any changes to the single-salt parameters of the Harvie and others (1984) data base will likely require extensive revision of the mixed-salt parameters and mineral free energies. Any new values of the mixed-salt parameters must be internally consistent with the single-salt parameters in the model and include the higher-order electrostatic terms, as discussed above. This requirement has been

followed in previously-discussed extensions to the Harvie and others (1984) data base given in PITZER.DATA, however, further testing is needed for calculations involving Fe^{2+} , Mn^{2+} , Sr^{2+} , Ba^{2+} , Li^+ , and Br^- . Any changes to the mixed-salt parameters, even if consistent with the given single-salt parameters, will likely require changes to the mineral free energies. Because of the lack of Pitzer interaction parameters for aqueous aluminum and silica species, we are unable to make calculations with aluminosilicates in PHRQPITZ.

The temperature range for equilibria in the PHRQPITZ.DATA file is variable and is generally 0 to 60 °C if ΔH_f° is known. The NaCl system is valid to approximately 350 °C, and the carbonate system is reliable to about 100 °C. The temperature dependence of the solubility of many of the minerals in PHRQPITZ.DATA is not known and large errors could result if calculations are made at temperatures other than 25 °C for these solids. Limited temperature-dependent data for single salt parameters are included in PITZER.DATA. If the sample temperature is outside the interval 0 to 55 °C, PHRQPITZ prints a nonfatal warning.

NUMERICAL METHOD

Much of the structure of PHREEQE is maintained in PHRQPITZ. PHRQPITZ still uses the Newton-Raphson approach to solve a set of algebraic equations by generating successively better estimates of the molalities and activity coefficients of the aqueous species. In effect, the only change in the results of simulations is caused by several new subroutines that calculate the activity coefficients and activity of water using the Pitzer equations. However, several major changes in the mechanics of the calculations have been made in order to deal with the large range of ionic strength for which the Pitzer formulations are applicable. At high ionic strengths, the activity coefficients and the activity of water have more extreme values. Thus, the numerical method has been changed to estimate changes in activity coefficients and the activity of water as part of refining the estimates of the unknowns.

Basic Equations

The most fundamental change from PHREEQE is that the master variables are the molalities of the master species instead of the activities of the master species. The Newton-Raphson equations rely primarily on the total differential of the molality of each aqueous species with respect to the master variables. The following equations define the total differential of an arbitrary aqueous species, m_i :

$$m_i = K_i (\gamma_1 m_1)^{c_{1,i}} (\gamma_2 m_2)^{c_{2,i}} a_{\text{H}_2\text{O}}^{c_{\text{H}_2\text{O},i}} / \gamma_i \quad (26)$$

where m_i is the molality of the i th aqueous species, m_1 and m_2 are the molalities of the master species in the mass action expression for the aqueous species, K_i is the equilibrium constant for the association reaction, γ is the activity coefficient of a species, c is the stoichiometric coefficient for a master species in the i th species, and $a_{\text{H}_2\text{O}}$ is the activity of water. The total differential is given by

$$dm_i = \frac{\partial m_i}{\partial \gamma_1} d\gamma_1 + \frac{\partial m_i}{\partial m_1} dm_1 + \frac{\partial m_i}{\partial \gamma_2} d\gamma_2 + \frac{\partial m_i}{\partial m_2} dm_2 + \frac{\partial m_i}{\partial a_{\text{H}_2\text{O}}} da_{\text{H}_2\text{O}} - \frac{\partial m_i}{\partial \gamma_i} d\gamma_i \quad (27)$$

An approximation is used at this point to reduce the activity coefficient terms in the equation to a single term which is dependent on ionic strength, I ,

$$d\gamma \sim \frac{d\gamma}{dI} dI \quad (28)$$

The derivative of the activity coefficients with respect to ionic strength is calculated based on the values of the activity coefficient and ionic strength at a given iteration and the values at the last time the Newton-Raphson procedure was used. Thus there is little computational expense in calculating these derivatives. In testing the program with a wide range of salts, this approximation appears to be adequate. The total differential of the molality of an aqueous species then becomes

$$dm_i = c_{1,i} m_i \frac{dm_1}{m_1} + c_{2,i} m_i \frac{dm_2}{m_2} + c_{H_2O,i} m_i \frac{da_{H_2O}}{a_{H_2O}} + \left(c_{1,i} m_i \frac{1}{\gamma_1} \frac{d\gamma_1}{dI} + c_{2,i} m_i \frac{1}{\gamma_2} \frac{d\gamma_2}{dI} - m_i \frac{1}{\gamma_i} \frac{d\gamma_i}{dI} \right) dI \quad (29)$$

This equation is the building block for the set of simultaneous equations which need to be solved. PHRQPITZ includes simultaneous equations for electroneutrality, mass balance for each element (except hydrogen and oxygen), and mineral mass action. Two equations that estimate the changes in the activity of water and the ionic strength are also included.

Electroneutrality error is defined

$$\sum_i z_i dm_i = E_{elect} \quad (30)$$

where z_i is the charge of the i th species and E_{elect} is the electrical imbalance calculated from the molalities of the aqueous species at the current iteration.

Mass-balance error is defined

$$\sum_i c_{j,i} dm_i = E_{mass(j)} \quad (31)$$

where $c_{j,i}$ is the stoichiometric coefficient of the master species j in the i th species, and $E_{mass(j)}$ is the error in the mass balance (the total concentration of the element j in solution minus the total concentration of j calculated from the molalities of the aqueous species at the current iteration).

Mineral mass action error is defined by

$$\sum_j c_{p,j} dm_j + \left(\sum_j c_{p,j} \frac{d\gamma_j}{dI} \right) dI = E_{IAP_p} \quad (32)$$

where $c_{p,j}$ is the stoichiometric coefficient of master species j in the mass action expression for mineral p and $E_{IAP(p)}$ is the error in the saturation state ($\log K/IAP$).

The change in the activity of water is estimated from

$$\sum_i DAW dm_i - da_{H_2O} = 0 \quad (33)$$

where DAW is the decrease in the activity of water per mole of solute, $DAW = (1 - a_{H_2O})/\sum m_i$, and $d(a_{H_2O})$ is the estimate of the change in the activity of water for the iteration. The activity of water is calculated explicitly in the Pitzer equations, so the utility of including this equation is to estimate the effect of the change in the activity of water on the molalities of all of the aqueous species and mineral saturation indices.

The change in the ionic strength is estimated from

$$\frac{1}{2} \sum_i z_i^2 dm_i - dI = 0 \quad , \quad (34)$$

where dI is the estimated change in the ionic strength for the iteration. Through the approximation stated above (eq. 28), this equation allows for the estimation of the effects of activity coefficients on the molalities of the aqueous species.

All of the equations presented are linear with respect to changes in the master variables: The molalities of the master species, the mass of minerals, ionic strength, and activity of water. The equations are solved simultaneously to find new values for the molalities of the master species.

Restrictions on the Newton-Raphson Approach

Though the Newton-Raphson approach is a very powerful technique for solving nonlinear equations, it is sensitive to poor estimates of the unknowns. Thus, PHRQPITZ applies the Newton-Raphson technique only after the estimates of the unknowns have satisfied certain criteria.

At each iteration, the sequence of calculations is as follows:

- (1) The molalities of the species are used to calculate activity coefficients,
- (2) the molalities of the master species and the activity coefficients are used to calculate the molalities of the aqueous complexes,
- (3) the molalities of the master species of the elements are estimated using the technique described for PHREEQE (Parkhurst and others, 1980), and
- (4) the molalities of the aqueous complexes are recalculated.

This sequence of calculations is repeated at least once to ensure that the molalities of the aqueous species are consistent with the activity coefficients and mass action expressions for the aqueous species. The sequence is repeated until the ionic strength has stabilized and the mass balance on the elements is satisfied within 10 percent. After these criteria are met, the Newton-Raphson equations are invoked. Until the electrical balance is less than 1.0 molal and the saturation indices of minerals are less than 0.5, no activity coefficient effects are included in the Newton-Raphson equations.

Scaling the Newton-Raphson Results

The result of solving the simultaneous equations is a vector of additive changes to the master variables. The magnitude of the changes can be extremely large. In these cases the vector is assumed to be in the direction of decreasing the error terms but must be scaled to much smaller values in order to approach the correct solution. Several criteria are used to ensure that the changes in the master variables are relatively small.

The pH is constrained to change less than 0.5 units per iteration. The maximum change in ionic strength is 0.5 molal. The maximum amount of a mineral that can dissolve or precipitate is 0.5 moles per iteration. Precipitation of minerals is limited so that no negative molalities result. These restrictions make the program more reliable over a range of chemical compositions but cause the program to require more iterations to find a solution to the equations.

DESCRIPTION OF INPUT

The user-defined input to PHRQPITZ is nearly identical to that of PHREEQE. Other input pertaining to the data file PITZER.DATA is new and described separately. Here we present the description of input section from Parkhurst and others (1980) which has been modified as needed to pertain to PHRQPITZ.

PHRQPITZ is designed to perform a sequence of simulations in a single computer run. Each simulation consists of two separate problems:

1. Process an initial solution or solutions and (or)
2. Model a reaction (starting from the initial solution(s)).

Many pathways for a simulation are accessible with a single input file; that is, no program modification should be necessary. Required input begins with a title line followed by a line of selected options. Depending on the options selected, additional data are supplied using various "Keyword" data blocks. A data block consists of a Keyword followed by appropriate data. The Keyword informs the program of the type and format of the data to follow. ELEMENTS and SPECIES, if they are used, should be the first two data blocks while the other keyword blocks may follow in any order. The keyword END denotes the end of the input data and is required once for each simulation. After the calculations for one simulation are completed, the program starts the data input process again, beginning with a new title and option line.

The general types of reactions that can be simulated are as follows:

1. Mixing of two solutions.
2. Titrating one solution with a second solution.
3. Adding or subtracting a net stoichiometric reaction (changing total concentrations of elements in proportion to a given stoichiometry).
4. Adding a net stoichiometric reaction until the phase boundary of a specified mineral is reached.
5. Equilibrating with mineral phases (mineral equilibrium can be specified with reaction types 1, 2, 3, 4, or 6 as well). Any condition which can be written in the form

$$\log(K_p) = \sum_i c_{p,i} \log(a_i) \quad ,$$

where a_i is the activity of the i th aqueous species, $c_{p,i}$ is the stoichiometric coefficient of the i th aqueous species in the p th phase, and K_p is the equilibrium constant for the p th phase, is considered a mineral phase. This definition of mineral equilibrium includes the following:

- a. Maintaining the aqueous phase in equilibrium with one or more minerals;
- b. Equilibration of a mineral-water system with a gas such as CO_2 ;
- c. Apparent ion exchange in the sense that a ratio of two aqueous ion activities is kept constant.

Any combination of the above can be included in the MINERALS keyword input provided the Gibbs Phase Rule is not violated.

6. Changing temperature.

These six types of reaction (processes) may be used in various combinations. For example, one could add a net stoichiometric reaction to a starting solution while maintaining mineral equilibrium and increasing temperature.

In each type of reaction, an initial solution must be specified. There are three ways to provide a starting solution for a reaction. (1) The total concentrations of elements (and other necessary information such as pH, and temperature) may be input using the SOLUTION n keyword. n is either 1 or 2 and indicates the number of the array where the solution data will be stored. Any stoichiometric reaction or simple mineral equilibration is performed on solution number 1 alone. Solution number 2 is required only for mixing and titrating. (2) The second method of providing a starting solution for a reaction is to save the final solution from the reaction step of the previous simulation (provided more than one simulation is made in a run). IOPT(7) is used to specify the solution number into which the final solution of a simulation will be saved. In the subsequent simulation no new solution should be read into that solution number. (3) Finally, if no reaction solution is saved and no new solutions are input, the solutions from the previous simulation remain in memory. Thus, a solution can be input once but can be used as the starting solution for multiple simulations.

One of the principle applications of PHRQPITZ is intended to be simulation of reactions based on observed water analyses, which will generally show an apparent electrical imbalance as a result of analytical errors. Because we use the electrical neutrality criterion in solving for pH, it is important to consider this apparent charge imbalance. Various options are available to achieve charge balance or to maintain a charge imbalance in the computations. Care should be taken in choosing the appropriate option and interpreting the results. Only perfect chemical analyses would produce electrical neutrality in an initial solution. Lacking these the solution may be left electrically unbalanced by setting IOPT(2) = 0. When a reaction is modeled the final calculated solution will have the same electrical imbalance as the initial solution. If IOPT(2) = 1, the pH of the initial solution will be adjusted to produce electrical neutrality in that solution. If the analytics are perfect, this places pH on the same activity coefficient scale as chosen for the aqueous model. It may be that the pH of the initial solution is relatively well known and it is more reasonable to add relatively inert ions like K⁺ or Cl⁻ to balance the charge. In this case set IOPT(2) = 2, and use the keyword NEUTRAL and associated input to specify K and Cl. The amount of K or Cl added will be listed in the output. One final alternative is to attribute the charge error to the most suspect analysis, e.g. carbon or sodium, or to a constituent known to be present for which one has no analytical data. Again, set IOPT(2) = 2 and use the NEUTRAL input. Lack of charge balance is a meaningful clue to the errors in analyses and a large error probably makes a solution unsuitable for reaction simulation. By using the various options of the program, one can investigate the significance of analytical errors and their effects on reaction simulations.

In the following description of the input, the Fortran format for each line is given. Any Fortran manual will provide a complete explanation of the symbols used in format statements. Format of input data sets is automatically controlled if PITZINPT is used to generate input data sets to PHRQPITZ.

A. Title and option lines.

1. TITLE LINE TITLE
 FORMAT (20A4)
 Eighty characters of titles or comments.
2. OPTION LINE (IOPT(I), I = 1,10), NSTEPS, NCOMPS, V0
 FORMAT (10I1,2I2, 6X, F10.5)

IOPT(1)	= 0,	No print of thermodynamic data or coefficients of aqueous species.
	= 1,	Print the aqueous model data (from PHRQPITZ.DATA) once during the entire computer run.
IOPT(2)	= 0,	Initial solutions are not to be charge balanced. Reaction solutions maintain the initial charge imbalance.
	= 1,	pH is adjusted in initial solution(s) to obtain charge balance. Caution: Large errors in pH can result from errors in analytical data, such as for water analyses that are not balanced in charge.

- = 2, The total concentration of one of the elements (except H or O) is adjusted to obtain electrical balance. NEUTRAL input is required.
- IOPT(3) = 0, No reactions are modeled. Only the initial solution is solved (speciated).
- = 1, Solution 1 is mixed (a hypothetical constant volume process) with solution 2 in specified reaction steps. STEPS input and a value of NSTEPS are required. MINERALS input may be included.
- = 2, Solution 1 is titrated with solution 2 in specified reaction steps. STEPS input, a value for NSTEPS, and a value for V0 are required. MINERALS input may be included.
- = 3, A stoichiometric reaction is added in specified reaction steps. REACTION input, STEPS input, a value for NSTEPS, and a value for NCOMPS are required. MINERALS input may be included.
- = 4, A net stoichiometric reaction is added in NSTEPS equal increments. REACTION input, STEPS input, a value of NSTEPS, and a value for NCOMPS are required. MINERALS input may be included. Only one value for the total reaction is read in STEPS.
- = 5, Solution number 1 is equilibrated with mineral phases only. No other reaction is performed. MINERALS input is required.
- = 6, A reaction is added to solution 1 until equilibrium is attained with the first phase in MINERALS input (equilibrium with other MINERALS phases is maintained throughout the reaction). REACTION input, a value for NCOMPS, and MINERALS input is required. No STEPS input is required. Note: there should be a common element in the reaction and the first phase in MINERALS, unless the phase boundary depends only on the activity of water (such as anhydrite-gypsum) at constant temperature and pressure. To evaporate or dilute solution number 1 to reach a phase boundary, define the REACTION a 1.0 H₂O (See description of REACTION input and Test Problem 6).
- IOPT(4) = 0, The temperature of the reaction solution is (a) the same as the initial solution if adding a reaction, or (b) calculated linearly from the end members if mixing or titrating. No TEMP input is required.
- = 1, The temperature is constant during the reaction steps and differs from that of the initial solution(s). One value is read in the TEMP input.
- = 2, The temperature is varied from T_o to T_f in NSTEPS equal increments during the reaction steps. A value for NSTEPS and two values of temperature, T_o and T_f, (in order) are required in the TEMP input, where T_o is the initial temperature and T_f is the final temperature.
- = 3, The temperature of each reaction step is specified in TEMP input, in order. NSTEPS values are read. Note: If a change in temperature is the sole "reaction", then a (null) stoichiometric reaction adding 0.0 moles must be defined using REACTION and STEPS input (see test problem 4).

IOPT(5)	= 0,	The pe of the initial solution (defined to be 4.0 if PIT-ZINPT is used to construct input sets) is held constant during all the reaction steps for the simulation. Redox reactions are currently not considered.
IOPT(6)	= 2,	Activity coefficients are calculated according to the Pitzer model. No other options are available.
IOPT(7)	= 0,	Do not save the aqueous phase composition at the end of a reaction for additional simulation.
	= 1,	Save the final reaction solution composition in solution number 1.
	= 2,	Save the final reaction solution composition in solution number 2.
IOPT(8)	= 0,	The debugging print routine is not called.
	= 1,	A long printout is output at each iteration in each problem. This print is to be used only if there are convergence problems with the program. (See discussion of subroutine PBUG given in Parkhurst and others (1980)).
IOPT(9)	= 0,	No printout of each array to be solved.
	= 1,	A long printout occurs of the entire array to be solved at each iteration. This print is used only if there are convergence problems. (See discussion of subroutine SLNQ given in Parkhurst and others (1980)).
IOPT(10)	= 0,	No scaling of the individual-ion activity coefficients is performed.
	= 1,	The individual-ion activity coefficients are scaled according to the MacInnes convention (see text).
NSTEPS		The number of reaction steps. A value is required if IOPT(3) = 1, 2, 3, or 4, or if IOPT(4) = 2 or 3. (Right justified.)
NCOMPS		The number of constituents in a net stoichiometric reaction. A constituent may be any element with an index number between 4 and 29 inclusive. No aqueous species with index numbers greater than 29 may be included as reaction constituents. A value of NCOMPS is required if IOPT(3) = 3, 4, or 6.
V0		The initial volume of solution 1 when modeling a titration. The unit of V0 must be the same as that of XSTEP (See STEPS input below) if IOPT(3) = 2. Otherwise, V0 is not required.

B. **Keyword data blocks.** Blocks are preceded by a keyword line. The keywords are numbered and underlined in the following text. Each keyword must begin in the first column of the line. The appropriate lines, which are lettered in the text, must follow in order directly after the keyword.

1. ELEMENTS FORMAT (A8)

This input defines the names and indices of all elements in the aqueous model data base (PHRQPITZ.DATA). One of line 1.a is read for each element. The index numbers of the elements do not need to be consecutive or sequential. This input block must be terminated with one blank line. Generally these data will be part of the PHRQPITZ.DATA file stored on disk and read by the program at the beginning of each run. Only changes to the data base need to be in the input data file.

- 1.a. TNAME, NELT, TGFW
 FORMAT (A8,2X,I2,3X,F10.0)
- TNAME Alphanumeric name of the element.
- NELT Index number assigned to the element. Number must be between 4 and 29, inclusive. (Right justified.)
- TGFW Gram formula weight of the species used to report the analytical data. If solution data include alkalinity, TGFW for the element carbon must be the equivalent weight of the reported alkalinity species. TGFW is not used if the concentrations are entered as molality (IUNITS = 0 in SOLUTION input line 3.b).
- 1.b. Blank line. (Denotes end of ELEMENTS input).
2. SPECIES FORMAT (A8)
- This input defines the names, index numbers and composition of all aqueous species in the aqueous model data base. Lines 2.a, 2.b, 2.c, and 2.d are read for each species. The index numbers for the species do not need to be sequential or consecutive. This input block must be terminated with one blank line. To eliminate a species (already in the PHRQPITZ.DATA file) from the aqueous model, only line 2.a followed by a blank line 2.b must be entered. More species changes could then follow or a second blank line would terminate this input block. All species must have association reactions which contain only master species (species numbers less than or equal to 29). Reactions containing non-master species must be converted to master species reactions and the appropriate association constants must be calculated before they can be entered into the program. These data are generally stored in the disk file PHRQPITZ.DATA which is read by the program at the beginning of each run and retained for the entire run. Only changes and additions to PHRQPITZ.DATA would appear in the input file.
- 2.a I
 FORMAT (I3)
- I The index number assigned to the aqueous species. Numbers 4 through 29 are reserved for master species. The maximum index number for an aqueous species is 250. (Right justified.)
- 2.b. SNAME, NSP, KFLAG, GFLAG, ZSP, THSP, DHA, ADHSP(1),
 ADHSP(2), ALKSP
 FORMAT (A8,2X,I3,2I1,6F10.3)
- SNAME Alphanumeric species name.
- NSP The total number of master species in the association reaction that forms this species; do not count the species itself unless the species is a master species (index number of 1 - 29). (Right justified.)
- KFLAG = 0. The van't Hoff equation is used to calculate the temperature dependence of the association constant for this species.
 = 1. An analytical expression is used to calculate temperature dependence of the association constant. Values of ASP are required on line 2.c.
- GFLAG Not used in PHRQPITZ. Read 0 or blank.
- ZSP The charge on this aqueous species.
- THSP Not used in PHRQPITZ. Read 0.0, or blank.

DHA Not used in PHRQPITZ. Read 0.0, or blank.
 ADHSP(1) Not used in PHRQPITZ. Read 0.0, or blank.
 ADHSP(2) Not used in PHRQPITZ. Read 0.0, or blank.
 ALKSP The alkalinity in equivalents assigned to this aqueous species. Such as 2.0 for carbonate species and 1.0 for bicarbonate. See discussion in Parkhurst and others (1980).

2.c. LKT0SP, DHSP, (ASP(I), I = 1,5)
 FORMAT (2F10.3,5E12.5)

Constants used to evaluate the association constant as a function of temperature. The analytical expression has the form of equation 23.

LKT0SP \log_{10} of the mass action association constant at 25 °C (used in van't Hoff equation).

DHSP Standard enthalpy of the association reaction at 25 °C (ΔH_r° , in Kcal/mole); used in the van't Hoff equation.

ASP(1) Constant term in equation 23 for calculation of the temperature dependence of $\log K$, ($= A_1$ in equation 3. The array ASP is used if KFLAG = 1.

ASP(2) Coefficient A_2 in equation 23.

ASP(3) Coefficient A_3 in equation 23.

ASP(4) Coefficient A_4 in equation 23.

ASP(5) Coefficient A_5 in equation 23.

2.d. (LSP(I), CSP(I), I = 1,NSP)
 FORMAT 6(I3,F7.3)

List of master species numbers and their coefficients in the mass action association reaction. NSP pairs of values, LSP and CSP, are read. One and only one of these lines is required for each species.

LSP(I) Index number of master species. (Right justified.)

CSP(I) Stoichiometric coefficient of master species in this aqueous species.

See Parkhurst and others (1980) for further information.

3. SOLUTION n FORMAT (A8,I1,I1)

This input is used to define a starting solution. n can be either 1 or 2 and indicates the solution number of the data following. Lines 3.a and 3.b are required. Line 3.c is not included if NTOTS = 0. There must be as many line 3.c's as necessary to read NTOTS total concentrations.

3.a. HEAD
 FORMAT (20A4)
 Title or comments about the solution.
 HEAD Alphanumeric heading

3.b. NTOTS, IALK, IUNITS, PH, PE, TEMP, SDENS
 FORMAT (I2,I3,I2,3X,4F10.3)

NTOTS The number of total concentrations to be read from line 3.c input. For example, if the starting solution is a MgCl_2 - NaHCO_3 solution, NTOTS = 4 (for Mg, Cl, Na, and C). (Right justified.)

IALK Flag which indicates whether total carbon or total alkalinity is to be input. (Right justified.)

- = 0 indicates the total concentration of carbon (not alkalinity) is input in the units specified by IUNITS (see below).
- = n $4 \leq n \leq 29$ where n is the index number for the element carbon, (in PHRQPITZ.DATA, we have defined $n = 15$) and indicates total alkalinity is being entered. If alkalinity is used ($n > 0$), then IOPT(2) cannot be equal to 1. It is theoretically impossible to use pH to achieve electrical neutrality if the alkalinity is fixed. ELEMENTS input may be required. The units of alkalinity are specified by IUNITS (below) and if IUNITS is greater than 0, the gram formula weight (GFW) of the element carbon is critically important. The GFW in the case of alkalinity must be the gram equivalent weight (grams/equivalent) of the chemical species in which the alkalinity is reported. The following is a list of species commonly used for reporting alkalinity and their corresponding equivalent weights:

CaCO_3 50.0446 g/eq
 HCO_3^- 61.0171 g/eq
 CO_3^{2-} 30.0046 g/eq.

(Note that: Alkalinity (mg HCO_3^-/L) = Alkalinity (mg CaCO_3/L) \times 1.21925.)

In PHRQPITZ.DATA 44.010 is the GFW of carbon which is suitable for entering carbon as total CO_2 . This GFW must be changed via ELEMENTS input if alkalinity is to be entered as mg/L or ppm (IUNITS = 2 or 3). If IUNITS = 0, alkalinity must be input as eq/kg H_2O and in this case the GFW need not be changed because no conversion of units is necessary. For a discussion of the contribution of different aqueous species to the total alkalinity see Parkhurst and others (1980).

- IUNITS** Flag describing units of input concentrations (right justified). The program makes all of its calculations in terms of molality and any other allowed concentration unit must be converted internally to molality before the calculations may begin. To make the conversions it is necessary to know the gram formula weight (GFW), in g/mole, of the chemical formula in which the elemental analyses are reported. The GFW is an input parameter under ELEMENTS input and must be in agreement with the analytical units for each solution data set. (If the units are molality, no conversion is necessary and the GFW's are not used.) Values of GFW currently in use are found in PHRQPITZ.DATA. Note: All elements must have the same units. It is not possible to enter mg/L of one element and molality of another.
- = 0 Concentration of elements entered as molality of each element, or for alkalinity, eq/kg H_2O (equivalents per kilogram).
 - = 1 Concentration of elements entered as mmol/L of each element, or for alkalinity, meq/L.
 - = 2 Concentration of elements entered as mg/L of the species which has a gram formula weight given in ELEMENTS input. (ELEMENTS input may be required.) For alkalinity see discussion under IALK above.

- = 3 Concentration of elements entered as ppm of the species which has a gram formula weight given in ELEMENTS input. (ELEMENTS input may be required.) For alkalinity see discussion under IALK above.
- = 4 Concentration of elements entered as millimoles per kilogram of solution, or for alkalinity as milliequivalents per kilogram of solution.
- PH The pH of the solution (the estimated pH if IOPT(2) = 1). Required for all solutions.
- PE The pe of the solution. It is suggested that the user enter 4.0 for this variable, as PHRQPITZ does not currently treat redox reactions.
- TEMP The temperature of the solution in °Celsius.
- SDENS The density of the solution. Required if concentrations are input in mmol/L or mg/L. If SDENS is omitted, 1.0 is assumed. Incorrect values of SDENS could lead to large errors in converting volume-based concentration units to molality.
- 3.c. (LT(I), DTOT(I), I = 1, NTOTS)
 FORMAT 5(I4,D11.3)
 Total concentrations of elements. Five values of LT. and DTOT are read on each line. The line may be repeated in order to enter all the elements desired. All data must appear consecutively in the fields, no blanks or zeros are allowed as values for LT. Omit this line if NTOTS is zero, the case of pure water.
- LT Index number of the element, consistent with ELEMENTS data in PHRQPITZ.DATA. (Right justified.)
- DTOT Total concentration of the element in molality, mmol/L, mg/L, ppm, or mmol/kg solution according to IUNITS.
4. MINERALS FORMAT (A8)
 This input defines the phases which will be maintained at equilibrium with each of the reaction solutions. Lines 4.a and 4.b are required for each mineral. Line 4.c is optional for each mineral depending on the value of MFLAG. Unlike SPECIES, MINERALS may be defined in terms of any aqueous species, not just the master species. A maximum of 19 minerals is allowed. The input expression for the equilibrium constant must correspond with the input mass action coefficients. Mineral reactions are written as dissociation reactions. MINERALS input must be terminated with a blank line.
- 4.a. MNAME, NMIN0, THMIN, LKT0M, DHMIN, MFLAG, SIMIN
 FORMAT (A8,2X,I2,3X,3F10.2,5X,I1,9X,F10.3)
 MNAME Alphanumeric name of mineral.
- NMIN0 Number of different species in the mineral dissociation reaction (including H⁺ and H₂O). NMIN0 must be less than or equal to 10. (right justified.)
- THMIN Not used in PHRQPITZ. Read 0.0.
- LKT0M Log₁₀ of the equilibrium constant at 25 °C for the reaction.
- DHMIN ΔH_r[°] (Kcal/mole) for the van't Hoff equation.
- MFLAG = 0. The van't Hoff equation is used to calculate the temperature dependence of the equilibrium constant.

- = 1. The analytical expression is used to calculate the temperature dependence of the equilibrium constant. Line 4.c is required.

SIMIN Saturation index ($\log_{10}(\text{Ion Activity Product}/K_{sp})$) desired in the final solution. SIMIN = 0.0 would result in equilibrium with the (one component, or fixed composition) mineral, while 1.0 would produce a solution 10 times supersaturated (SI = 1.0). This variable is also useful in specifying the partial pressure of a gas. The Henry's law constant for the gas would be entered using the equilibrium constant (LKT0M) or analytical expression (AMIN) and the log of the partial pressure would be entered for SIMIN.

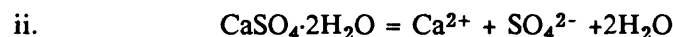
4.b. (LMIN(I), CMIN(I), I = 1,NMIN0)
FORMAT 5(I4,F11.3)

List of species index numbers and stoichiometric coefficients in the dissociation reaction for this mineral. NMIN0 pairs of numbers, LMIN and CMIN, are read. The maximum value of NMIN0 is 10. If NMIN0 is greater than 5, a second line 4.b is required.

LMIN(I) Index number of species (not necessarily master species) in the dissociation reaction for this mineral. (Right justified.)
CMIN(I) Stoichiometric coefficient of species in dissociation reaction. For example, using the data for aqueous species index numbers in PHRQPITZ.DATA,

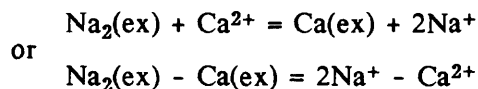


NMIN0 = 2
LMIN(1) = 4, CMIN(1) = 1.0
LMIN(2) = 15, CMIN(2) = 1.0.



NMIN0 = 3
LMIN(1) = 4, CMIN(1) = 1.0
LMIN(2) = 16, CMIN(2) = 1.0
LMIN(3) = 3, CMIN(3) = 2.0.

iii. $\text{Ca}^{2+} - \text{Na}^+$ ion exchange. (Assumes composition of exchanger does not change.)



NMIN0 = 2
LMIN(1) = 6, CMIN(1) = 2.0
LMIN(2) = 4, CMIN(2) = -1.0.

LKT0M = $a_{\text{Na}^+}^2 / a_{\text{Ca}^{2+}}$

- iv. Fix CO₂ partial pressure.⁵

$$\text{CO}_{2(\text{gas})} = \text{CO}_2^*(\text{aq})$$

where CO₂* (aq) represents the sum of CO₂^o and H₂CO₃^o

NMIN0 = 1
LMIN(1) = 35, CMIN(1) = 1.0
LKTO M = Henry's law constant for CO₂
SIMIN = Log P_{CO2} desired.

- 4.c. AMIN(I), I = 1,5
FORMAT (5E12.5)

Equilibrium constant expression of the form of equation 23 for the mineral dissociation reaction. This line is used only if MFLAG = 1.

AMIN(1) Coefficient A₁ of equation 23
AMIN(2) Coefficient A₂ of equation 23
AMIN(3) Coefficient A₃ of equation 23
AMIN(4) Coefficient A₄ of equation 23
AMIN(5) Coefficient A₅ of equation 23

- 4.d. Blank line.

5. LOOK MIN FORMAT (A8)

The purpose of this input is simply to provide information on the saturation state of the aqueous phase with respect to desired minerals. The minerals in this block of input do not affect the calculations of the initial solution or any of the reaction solutions. This input is never mandatory. The ion activity product (IAP) and saturation index (SI = log (IAP/K)) of each of these minerals is printed in the output following each solution description. Only the minerals which contain elements present in the solution are printed. The input following this line is identical to the input for MINERALS (see above). This input must be terminated with a blank line. The list of "look minerals" is maintained for the duration of the run and any new "look mineral" is simply added to the list. If a "look mineral" is added that has the identical 8 letter name as another mineral in the list, the new mineral replaces the old mineral. The word DELETE as a mineral name will eliminate all of the minerals in the list and new minerals may be added. Only 39 "look minerals" are allowed. LOOK MIN input is generally placed in PHRQPITZ.DATA and is read once for each run. The input file need not contain additions and changes to the PHRQPITZ.DATA list of "look minerals".

6. MEAN GAM FORMAT (A8)

This input allows the user to calculate the mean activity coefficient for any neutral salt combination of aqueous species. The input is optional. Data for calculation of 21 mean activity coefficients are included in PHRQPITZ.DATA. A total of 40 are allowed. MEAN GAM input can appear in the input data set as well as PHRQPITZ.DATA. The word DELETE is allowed and its usage is the same as in LOOK MIN input (see above).

⁵ The reaction is written for the predominant species, CO₂^o.

- 6.a. NMEANG, IMEANG, (LMEANG(I), CMEANG(I), I = 1,3)
FORMAT (A8,I2,3(I4,F6.0))

This input defines the stoichiometry of neutral salts for calculation of mean activity coefficients.

- NMEANG Alphanumeric name of salt or mineral.
IMEANG Number of different species in the salt. This number is usually 2 and may not exceed 3.
LMEANG(I) Index number of species (not necessarily a master species) in the salt. Index numbers are as defined under ELEMENTS and SPECIES input.
CMEANG(I) Stoichiometric coefficient of species in salt. For example, using SPECIES data from PHRQPITZ.DATA:
- i. NaCl
 IMEANG = 2
 LMEANG(1) = 6, CMEANG(1) = 1.0
 LMEANG(2) = 14, CMEANG(2) = 1.0
 - ii. NaHCO₃
 IMEANG = 2
 LMEANG(1) = 6, CMEANG(1) = 1.0
 LMEANG(2) = 34, CMEANG(2) = 1.0

7. TEMP FORMAT (A8)

This input varies the temperature during the reaction steps. It is required input if IOPT(4) is greater than 0. Only one line 7.a. is necessary unless IOPT(4) = 3. In that case as many lines as necessary to input NSTEPS values are required. If TEMP input is included, REACTION input is required, even for a null reaction (see test problem 4).

- 7.a. XTEMP
 FORMAT (8F10.1)

XTEMP Temperature in degrees Celsius.
 If IOPT(4)=1, one value of XTEMP is read.
 If IOPT(4)=2, two values of XTEMP are read,
 T₀ and T_f (in order).
 If IOPT(4)=3, NSTEPS values of XTEMP are read
 (no blank fields permitted).

8. STEPS FORMAT (A8)

This input defines the steps of the reaction progress. The input has a different meaning depending on the value of IOPT(3) (option line).

- IOPT(3) = 1, XSTEP is the fraction of solution 1 to be mixed with solution 2. NSTEPS values are read.
IOPT(3) = 2, XSTEP is the volume of solution 2 to be titrated into solution 1. XSTEP must have the same units as V0 (option line). NSTEPS values are read.

IOPT(3) = 3, XSTEP is the number of moles of reaction to be added to solution 1. NSTEPS values are read.

IOPT(3) = 4, Only one value of XSTEP is read. XSTEP is the total number of moles of reaction to be added in NSTEPS steps. NSTEPS reaction solutions will be calculated. The *I*th solution will have $I \cdot XSTEP / NSTEPS$ moles of reaction added to solution 1.

8.a. XSTEP
FORMAT (8F10.3)

XSTEP Reaction increments as defined above.

9. REACTION FORMAT (A8)

This input describes the stoichiometry of the elements to be added as a reaction. STEPS input (see above) defines the total number of moles of this reaction to be added. The REACTION input changes the total aqueous concentration of an element by the stoichiometric coefficient (CREAC) times the total moles of reaction (XSTEP). (However, the final total concentration in the reaction solution may also be altered by mass transfer to achieve equilibrium with minerals specified in MINERALS input). It is necessary to consider the charge balance of the reaction which is added. A charge imbalance by an input error, or by intent, is equivalent to adding acid or base. If the reaction is a simulation from a known solution to another known solution, it is possible to add an amount of inert electrical charge equal to the difference in the charge imbalance between the two solutions. In this case, set LREAC(I)=0, CREAC(I)= Δ charge imbalance (equivalents per kilogram H₂O). This will eliminate implicit addition of acid or base. However, to add acid or base intentionally, simply add the anion for an acid (such as Cl⁻ for HCl), or the cation for a base (such as Na⁺ for NaOH) as a reaction. For the special case of evaporation or dilution to reach a phase boundary, set IOPT(3) = 6, define the phase boundary and appropriate equilibria with MINERALS input and set LREACT(1) = 3, and CREAC(1) = 1.0 (see Test Problem 6). Line 8.a is repeated as often as necessary to read NCOMPS (option line) reaction constituents. REACTION input is required for all changes in temperature, even if no reaction is intended. In this case a null reaction is defined using REACTION and STEPS input (see test problem 4).

9.a. (LREAC(I), CREAC(I), THMEAN(I), I = 1,NCOMPS)
FORMAT 4(I4,2F8.3)

This input defines a net stoichiometric reaction. Four triples of numbers are read on each line. Enough lines must be included to read NCOMPS triples of numbers.

LREAC(I) Index number of element for the reaction.

CREAC(I) Stoichiometric coefficient of the element in the reaction.

THMEAN(I) The operational valence of the element in the reaction.
Not used in PHRQPITZ. Read 0.0.

The variables which affect a reaction simulation are IOPT(3), NSTEPS, and NCOMPS from the option line, REACTION input and STEPS input. The following examples use the species index numbers from PHRQPITZ.DATA.

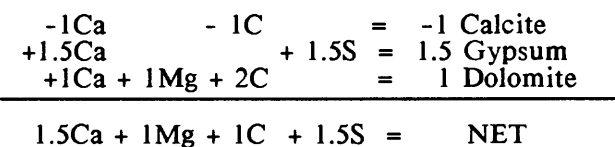
- i. Gypsum is added to the initial solution in 5 equal increments of 0.005 moles, to a total of 0.025 moles. Calcite equilibrium is maintained in each of the five steps.

IOPT(3) = 4 (net reaction linearly added).
 NSTEPS = 5 (5 reaction steps)
 NCOMPS = 2 (2 constituents, Ca and S)
 LREAC(1) = 4, CREAC(1) = 1.0 (Ca)
 LREAC(2) = 16, CREAC(2) = 1.0 (S)
 MINERALS input, calcite.
 XSTEP(1) = 0.025 (total moles of reaction to be added)

The total calcium at the completion of the first reaction step is given by:

$$\text{Ca}_{\text{tot}} = \text{Ca}_{\text{tot(initial)}} + 0.005 + \text{MIN}_{\text{calcite}}$$

- ii. Suppose mass balance between two solutions shows calcite, gypsum, and dolomite dissolving (+) and precipitating (-) in proportions of -1:1.5:1. The net reaction is written:



Three points along this path are modeled by (arbitrarily) adding 10^{-4} , 10^{-3} , and 10^{-2} moles of the net reaction.

IOPT(3) = 3 (add net reaction in specified steps)
 NSTEPS = 3 (number of steps)
 NCOMPS = 4 (number of constituents in reaction)
 LREAC(1) = 4 CREAC(1) = 1.5 (Ca)
 LREAC(2) = 5 CREAC(2) = 1.0 (Mg)
 LREAC(3) = 15 CREAC(3) = 1.0 (C)
 LREAC(4) = 16 CREAC(4) = 1.5 (S)
 XSTEP(1) = 10^{-4} , XSTEP(2) = 10^{-3} , XSTEP(3) = 10^{-2} .

(Note that reaction increments are in moles, reaction progress is always from the defined solution 1 and is not cumulative)

10. NEUTRAL FORMAT (A8)

This input defines the elements to be used to adjust the initial solution(s) to electrical neutrality. One element with a master species cation and one element with a master species anion are input. H^+ (and e^-) are not valid entries. A master cation and anion are required in order to add one or the other element according to the charge imbalance. Species are not subtracted, eliminating the possibility of negative total concentrations. This input is required only if IOPT(2) = 2. (Note that this is not equivalent to adding or subtracting charge as discussed in REACTION; remember that IOPT(2) = 0 will maintain an original charge imbalance during a simulation.)

10.a. LPOS, LNEG FORMAT (2I5)

LPOS Index number of an element with a cation master species.
(Right justified.)

LNEG Index number of an element with an anion master species.
(Right justified.)

11. SUMS

FORMAT (A8)

This input sums molalities of aqueous species which are then printed in the output of the run. These sums do not affect the calculations in any way and are never mandatory. The "sums" are defined by lists of species numbers, so that each time a sequence number for a species is listed, the sum is incremented by the molality of that species. If the species has, for example, two carbonate ions and the total carbonate is the sum which is desired, then the species should be listed twice in that sum. Up to 10 different sums may be defined. Each sum may have up to 50 species. Lines 11.a and 11.b are required for each sum. This input, after all sums have been defined, must be terminated with one blank line.

As in LOOK MIN input, the sums are kept for the duration of the run, but it is possible to add or replace sums or delete the entire set in any single simulation. Any sum input in this data block will be added to the list of sums if the name (SUNAME) is different from all other sum names. A sum with the identical name will replace the sum already in the list. The word DELETE as a sum name will eliminate all the sums known to the computer.

- 11.a. SUNAME, NSUM
FORMAT (A8,2X,I2)

SUNAME Alphanumeric name to be printed to identify the sum.

NSUM The number of index numbers to be read on line(s) 11.b;
NSUM cannot exceed 50. (Right justified.)

- 11.b. (LSUM(I,J), J = 1,NSUM)
FORMAT (20I4)

List of species numbers to define the sum. Twenty index numbers are read on this line. The line may be repeated as many times as necessary to input NSUM index numbers.

LSUM Index numbers of species in sum. (Right justified.)

Note: repeat lines 11.a and 11.b for each sum.

- 11.c. Blank line. One blank line at the end of all sums is required to terminate this input.

12. END

FORMAT (A8)

This line terminates input operations for a single simulation. Initial solution(s) are computed as directed by the preceding input. Any computer run has at least one END line.

DESCRIPTION OF PHRQPITZ.DATA FILE

The file PHRQPITZ.DATA contains blocks of KEYWORD information described above under "Description of Input". There is no title or option line in PHRQPITZ.DATA. The file begins with appropriate keywords such as ELEMENTS, SPECIES, LOOK MIN, and MEAN GAM, each with its associated input described above. The PHRQPITZ.DATA file (Attachment A) can be modified as needed to define any other aqueous model or chemical system that can be treated by the specific interaction approach. Data entered in PHRQPITZ.DATA are read automatically by PHRQPITZ and need not be included in the user-defined input file.

The PHRQPITZ.DATA file is read once for a given run (which may contain multiple, successive simulations) and remains in effect throughout the run, unless changed by user-defined input. The PHRQPITZ.DATA file may be changed, deleted, or supplemented by inclusion of appropriate keyword blocks and accompanying modifications in the user-defined input set. See "Description of Input" (above) and Parkhurst and others (1980) for further details.

DESCRIPTION OF PITZER.DATA FILE

The file PITZER.DATA contains values of all the interaction parameters for the Pitzer model. The current file used with PHRQPITZ is listed in Attachment B. The Pitzer interaction parameters are defined by equations 1-21. The PITZER.DATA file defines all known interactions for the species included in the PHRQPITZ.DATA file (Attachment A). Interactions may be entered in PITZER.DATA only for species defined with appropriate ELEMENTS and SPECIES input in PHRQPITZ.DATA. Unknown values of interaction parameters are defined internally in PHRQPITZ as zero and may not be included in PITZER.DATA. Inclusion of zero values for any interaction parameter (except ψ , which is the last block of data read) will result in a read error. Species names used in PITZER.DATA must be exactly as they appear in PHRQPITZ.DATA. The input to PITZER.DATA is defined as follows.

1. SPEC1, IS
FORMAT (A8,I3)

 SPEC1 Type the word SPECIES for this variable to identify the input that follows. This variable is not used.
 IS Total number of individual species in the Pitzer model.
 1.a. (NS(I), I = 1,IS)
 FORMAT 18(1X,I3)

 NS Index number assigned to the aqueous species as defined in PHRQPITZ.DATA. IS values are read and may appear in any order. Lines 1.a. are repeated as needed, provided no blank or zero values are read.
2. This input type contains all single salt interaction parameters and constants for defining their temperature dependence. The first line of this input requires a zero or blank entry for VALUE(1) in columns 21-32 (see description of input given below). The $\beta^{(0)}$ values must begin on the second line of input and continue for as many non-zero values of $\beta^{(0)}$ as to be defined. The second block of parameters contains values for $\beta^{(1)}$ followed by blocks of values for $\beta^{(2)}$ and then $C\phi$. Each of these blocks of parameters must begin with a line containing a blank or zero value for VALUE(1) in columns 21-32. The names B0, B1, B2, and C0 in Attachment B are for identification purposes only and mark lines where zero or blank entries for VALUE(1) are given. The order of data blocks of this type must be $\beta^{(0)}$, $\beta^{(1)}$, $\beta^{(2)}$, and $C\phi$. Zero values of the interaction parameters must not be entered unless to denote the beginning of a new data block.
 2.a SPEC1, SPEC2, (VALUE(I), I = 1,5)
 FORMAT 2(2X,A8),5(1X,F11.0)

 SPEC1 Name of the *i*th species exactly as used in PHRQPITZ.DATA for a particular ij cation-anion or anion-cation interaction.
 SPEC2 Name of the *j*th species exactly as used in PHRQPITZ.DATA for a particular ij cation-anion or anion-cation interaction.

VALUE(I) Coefficients for the temperature dependence of the p th single-salt interaction parameter in the form

$$P_{ij} = c_1 + c_2 \left(\frac{1}{T} - \frac{1}{T_R} \right) + c_3 \ln \left(\frac{T}{T_R} \right) + c_4 (T - T_R) + c_5 (T^2 - T_R^2)$$

where $T_R=298.15$, T is temperature in °K and c_1 - c_5 are read as VALUE(I), $I=1,5$. VALUE(1) ($= c_1$) contains the value of the interaction parameter at the reference temperature. In many cases we only know the first derivative of an interaction parameter with respect to temperature and this then appears in VALUE(4) ($= c_4$).

Subsequent data blocks contain values of θ , λ , and ψ and, as before, each of these blocks begins with a zero or blank entry for the interaction parameter in VALUE(1). The order of these data blocks must appear as θ , λ , and ψ . Zero or blank values for a given interaction are never required and are not permitted for θ and λ data blocks.

3. Input format for θ_{ij} and λ_{ij} . The subscripts ij refer to dissimilar cation-cation or anion-anion pairs for θ_{ij} and ion-neutral species interactions for λ_{ij} . PHRQPITZ does not treat θ , λ and ψ as functions of temperature. The Block of parameters for θ must precede the block for λ . Each block begins with a blank or zero value for VALUE(1).

3.a SPEC1, SPEC2, VALUE(1)
FORMAT 2(2X,A8),1X,F11.0

SPEC1 Name of i th species exactly as used in PHRQPITZ.DATA for a particular ij interaction.

SPEC2 Name of j th species exactly as used in PHRQPITZ.DATA for a particular ij interaction.

VALUE(1) Value of the parameter at 25 °C. Read a zero or blank entry for VALUE(1) for the first line of each block of input. Caution: as discussed in the text, θ_{ij} values must be internally consistent with the ψ_{ijk} values, higher-order electrostatic terms and the single salt interaction parameters.

4. This data block contains entries for ψ_{ijk} , where ijk refer to dissimilar cation-cation-anion and anion-anion-cation interactions. The first line of data must contain a blank or zero entry for VALUE(1). Zero values of ψ are permitted elsewhere in this block but are not required.

4.a SPEC1, SPEC2, SPEC3, VALUE(1)
FORMAT 3(2X,A8),2X,F10.0

SPEC1 Name of the i th species exactly as used in PHRQPITZ.DATA for the particular ijk interaction.

SPEC2 Name of the j th species exactly as used in PHRQPITZ.DATA for the particular ijk interaction.

SPEC3 Name of the k th species exactly as used in PHRQPITZ.DATA for the particular ijk interaction.

VALUE(1) Value of ψ_{ijk} at 25 °C. Caution: values of ψ_{ijk} must be internally consistent with values of θ_{ij} , higher-order electrostatic terms, and the single-salt parameters (see text).

The final line of the file PITZER.DATA is the last entry of ψ_{ijk} .

TEST PROBLEMS

The test problems that follow are designed to illustrate ways of setting up input sets for a wide range of problems that can be considered by PHRQPITZ. The problems also serve as a basis of comparison when transporting the code to other machines. For each problem the input set is given along with the printed output and a brief explanation. An example illustrating how to make temporary changes to the thermodynamic data base is given elsewhere (Parkhurst and others, 1980).

Test problem 1: Speciate a brine sample and examine effects of changing activity-coefficient scale.

This problem demonstrates how to speciate a brine solution for purposes of calculating mineral-water saturation state, activity coefficients, etc. The brine analysis is sample T-93 taken from Frape and others (1984) and is listed in table 5.

Table 5. -- Analytical data for Canadian Shield brine T-93 (mg/L)^[1]

Input #	Parameter	Value
---	Temperature °C	18.0
---	pH	5.00
---	Density	1.204
4	Calcium	64,000.
5	Magnesium	5,100.
6	Sodium	45,000.
7	Potassium	199.
12	Strontium	1,080.
14	Chloride	207,000.
15	Alkalinity (HCO ₃ ⁻)	19.
16	Sulfate	284.
22	Bromide	1,760.

^[1]From Frape and others (1984). Reported silica (4.80 mg/L SiO₂) not included.

In addition to examining the speciation calculation, this test problem demonstrates the significance of uncertainties in activity-coefficient scale on the predicted saturation state, particularly as they apply to the carbonate system. The problem is in 4 parts. In parts 1a and 1b the analysis of table 5 is speciated on the MacInnes scale, and unscaled, respectively, using the measured pH. Parts 1c and 1d attempt to resolve uncertainties in the carbonate system by assuming the water of table 5 is in equilibrium with calcite, on the MacInnes scale and unscaled, respectively. This procedure is preferable to alternate means of defining pH, such as through charge balance (using IOPT(2) = 1) owing to uncertainties in the analytical data. However, as we shall see, the analytical data are insufficient to fully resolve uncertainties in the carbonate system due to activity-coefficient scale.

The input data set for all four parts is listed in table 6. In all cases, the starting solution is as given in table 5, and IOPT(2) = 0 which retains the analytical charge imbalance (-0.30305 eq/kg H₂O) throughout the calculations. Consequently, it is not necessary to adjust the initial solution composition to obtain charge balance. Because total inorganic carbon is calculated from pH and alkalinity, it is theoretically impossible to adjust pH to obtain charge balance. In parts 1a and 1b IOPT(3) = 0 which indicates that SOLUTION 1 will be speciated only. IOPT(3) is 6 in parts 1c and 1d indicating that after speciation, an irreversible reaction will be added (or removed) to solution 1 to obtain equilibrium with one or more minerals. IOPT(6) is set to 2 indicating that the Pitzer equations will be used (no other choices are currently available). In parts 1a and 1c IOPT(10) is 1 which causes all activity coefficients to be scaled according to the MacInnes convention. IOPT(10) is 0 in parts 1b and 1d for unscaled activity

Table 6. -- Line images of the input file to test problem 1

Test Problem 1a: Canadian Shield Brine T-93; Frape et al., 1984, Alk+Mac Scale
0000020001 0 0 0.0

ELEMENTS

C 15 61.0171

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.	4.0	18.	1.204		
464000.	645000.	55100.	7199.	121080.	
14207000.	221760.	16284.	1519.		

END

Test Problem 1b: Canadian Shield Brine T-93; Frape et al., 1984, Alk + No scale

0000020000 0 0 0.0

ELEMENTS

C 15 61.0171

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.	4.0	18.	1.204		
464000.	645000.	55100.	7199.	121080.	
14207000.	221760.	16284.	1519.		

END

Test Problem 1c: Canadian Shield Brine T-93; Frape et al., 1984, calcite + Mac

0060020001 0 1 0.0

ELEMENTS

C 15 61.0171

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.	4.0	18.	1.204		
464000.	645000.	55100.	7199.	121080.	
14207000.	221760.	16284.	1519.		

MINERALS

CALCITE 2 4.00 -8.406 1

15 1.0 4 1.0
-171.8329 -.077993 2839.319 71.595

REACTION

151.0 4.0

END

Test Problem 1d: Canadian Shield Brine T-93, calcite equil. No scale

0060020000 0 1 0.0

ELEMENTS

C 15 61.0171

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.	4.0	18.	1.204		
464000.	645000.	55100.	7199.	121080.	
14207000.	221760.	16284.	1519.		

MINERALS

CALCITE 2 4.00 -8.406 1

15 1.0 4 1.0
-171.8329 -.077993 2839.319 71.595

REACTION
151.0 4.0
END

coefficients. ELEMENTS input is used to redefine the molecular weight of the carbon input species from that of CO_2 (default in PHRQPITZ.DATA) to that of HCO_3^- for input of total alkalinity as HCO_3^- (table 5). The new molecular weight of the carbon input species is used internally during execution and not retained in PHRQPITZ.DATA. Permanent changes to PHRQPITZ.DATA must be made through normal editing procedures. MINERALS input is included in parts 1c and 1d to establish equilibrium with calcite, and REACTION input is used to adjust the total inorganic carbon content to establish calcite equilibrium.

The output for test problems 1a - 1d is listed in table 7. The output formats have been adjusted to accommodate screen widths of 80 characters. The printed output first lists all input to PHRQPITZ. The first 4 keywords printed (table 7), ELEMENTS, SPECIES, LOOK MIN, MEAN GAM were printed as appropriate data from PHRQPITZ.DATA were read. Computed output begins with a print of the total molalities of all elements in SOLUTION 1 followed by a description of the solution and distribution of species. Computed variables under "Description of Solution" include the activity of H_2O , osmotic coefficient, ionic strength, pressure (in atmospheres) along the vapor pressure line of pure water (if temperature is greater than 100°C), the density of pure water at the given temperature and pressure, the electrical imbalance (in equivalents per kilogram H_2O), and the total concentration of inorganic carbon (moles per kg H_2O).

The molality, activity, and activity coefficient (and log values) of all aqueous species in the PHRQPITZ.DATA file for which an input concentration was given are listed under the "Distribution of Species". This output is followed by tables of total molalities, activities, activity coefficients, mean activity coefficients, and saturation indices. The saturation indices are given under the heading LOOK MIN IAP as $\log(\text{IAP}/K_T)$. This output is repeated in part 1b for unscaled results.

In parts 1c and 1d the equilibration step (with calcite) appears under the heading "STEP NUMBER 1". The total molalities of the elements are printed first before the total molality of inorganic carbon is adjusted to bring the solution to equilibrium with calcite. The calcite mass transfer given under the heading "PHASE BOUNDARIES" is zero indicating that calcite has neither been dissolved or precipitated. The moles of inorganic carbon added or removed from the solution to reach calcite saturation follows. This is followed by an updated print of the saturation indices, revised total molalities, "Description of Solution", etc.

The results of test problem 1 are useful in illustrating some of the problems in making speciation calculations with the measured pH. Examination of the output from parts 1a and 1b (table 7) shows that for the same brine analysis, the saturation indices of calcite and dolomite vary, respectively, from -0.31 and -1.08 (MacInnes scale) to +0.16 and -0.13 (unscaled). The computed total concentration of inorganic carbon and $\log \text{PCO}_2$, which is based on the pH and alkalinity measurements, varies from 32.43 mmol/kg H_2O and 0.379 (MacInnes scale) to 10.25 mmol/kg H_2O and -0.13 (unscaled), respectively. This is a change of more than 300 percent for these variables. Because total inorganic carbon, total alkalinity, and PCO_2 can be measured within ten percent or better in many brines, the thermodynamics of the carbonate system in this brine would be better defined by measuring total alkalinity and either total inorganic carbon or PCO_2 . If such data are available, the carbonate system is unambiguously defined and independent of the measured pH and choice of scale convention. A value of pH would then be defined by the aqueous model according to any particular choice in activity-coefficient scale.

As a means of examining possible variations in the pH of brine T-93 (table 5), and in the absence of total inorganic carbon or PCO_2 data, the solution was equilibrated with calcite on both activity-coefficient scales (parts 1c and 1d) by adjusting the total inorganic carbon content. This caused outgassing (MacInnes scale) and ingassing (unscaled) of CO_2 to reach calcite saturation. The results suggest that pH could vary as much as 0.5 between the two scales. The carbonate system is more closely defined, as evidenced by the similarity in the calculated PCO_2 and total inorganic carbon. The total inorganic carbon and PCO_2 are not in complete agreement because of the inconsistency of the measured pH (5.00) with either scale. This inconsistency can be resolved only if total inorganic carbon or PCO_2 is measured in conjunction with the total alkalinity measurement.

Table 7. -- Listing of printout from test problem 1

DATA READ FROM DISK

ELEMENTS
SPECIES
LOOK MIN
MEAN GAM

1Test Problem 1a: Canadian Shield Brine T-93; Frape et al., 1984, Alk+Mac Scale
0000020001 0 0 0.00000

ELEMENTS

C 15 0.61017E+02
0 0.00000E+00

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.00 4.00 18.0 1.20
4 6.400D+04 6 4.500D+04 5 5.100D+03 7 1.990D+02 12 1.080D+03
14 2.070D+05 22 1.760D+03 16 2.840D+02 15 1.900D+01

1SOLUTION NUMBER 1

Canadian Shield Brine T-93.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
TOT ALK	3.540280D-04	-3.4510
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

----DESCRIPTION OF SOLUTION----

PH = 5.0000
ACTIVITY H2O = 0.7208
OSMOTIC COEFFICIENT = 1.6522
IONIC STRENGTH = 8.5902
TEMPERATURE = 18.0000
PRESSURE = 1.0000 ATM
DENSITY OF H2O = 0.9986 G/CC
ELECTRICAL BALANCE = -3.0305D-01
TOTAL ALKALINITY = 3.5403D-04
ITERATIONS = 5
TOTAL CARBON = 3.2432D-02

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	MACINNES ACTIVITY	SCALE LOG ACT	MACINNES GAMMA	SCALE LOG GAM
1	H+	1.0	7.997E-06	-5.097	1.000E-05	-5.000	1.251E+00	0.097
3	H2O	0.0	7.208E-01	-0.142	7.208E-01	-0.142	1.000E+00	0.000
4	CA+2	2.0	1.815E+00	0.259	1.783E-01	-0.749	9.823E-02	-1.008
5	MG+2	2.0	2.386E-01	-0.622	4.144E-02	-1.383	1.737E-01	-0.760
6	NA+	1.0	2.225E+00	0.347	6.252E-01	-0.204	2.809E-01	-0.551
7	K+	1.0	5.787E-03	-2.238	5.991E-04	-3.222	1.035E-01	-0.985
12	SR+2	2.0	1.401E-02	-1.853	7.035E-04	-3.153	5.020E-02	-1.299
14	CL-	-1.0	6.638E+00	0.822	4.263E+01	1.630	6.422E+00	0.808
15	CO3-2	-2.0	3.764E-07	-6.424	1.169E-08	-7.932	3.106E-02	-1.508
16	SO4-2	-2.0	3.361E-03	-2.474	5.503E-04	-3.259	1.637E-01	-0.786
22	BR-	-1.0	2.504E-02	-1.601	2.389E-01	-0.622	9.539E+00	0.979
31	OH-	-1.0	1.797E-09	-8.745	4.213E-10	-9.375	2.345E-01	-0.630
34	HCO3-	-1.0	3.473E-04	-3.459	2.980E-03	-2.526	8.581E+00	0.934
35	H2CO3	0.0	3.208E-02	-1.494	9.948E-02	-1.002	3.101E+00	0.491
40	HSO4-	-1.0	9.753E-08	-7.011	4.326E-07	-6.364	4.436E+00	0.647
76	CACO3	0.0	2.615E-06	-5.583	2.615E-06	-5.583	1.000E+00	0.000
85	MGOH+	1.0	3.780E-08	-7.423	2.480E-09	-8.605	6.563E-02	-1.183
86	MGCO3	0.0	3.718E-07	-6.430	3.718E-07	-6.430	1.000E+00	0.000

SPECIES	TOTAL MOL	MACINNES ACTIVITY	MACINNES TOTAL GAMMA
H+	8.0943D-06	1.0000D-05	1.2354D+00
CA+2	1.8155D+00	1.7834D-01	9.8232D-02
MG+2	2.3857D-01	4.1441D-02	1.7371D-01
NA+	2.2254D+00	6.2518D-01	2.8092D-01
K+	5.7867D-03	5.9913D-04	1.0354D-01
SR+2	1.4014D-02	7.0352D-04	5.0202D-02
CL-	6.6382D+00	4.2630D+01	6.4219D+00
CO3-2	3.3633D-06	1.1691D-08	3.4761D-03
SO4-2	3.3612D-03	5.5030D-04	1.6372D-01
BR-	2.5043D-02	2.3887D-01	9.5386D+00
OH-	3.9592D-08	4.2135D-10	1.0642D-02
HCO3-	3.4726D-04	2.9799D-03	8.5812D+00
H2CO3	3.2081D-02	9.9484D-02	3.1010D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.5941D+00
CASO4	1.2682D-01
CACO3	1.8479D-02
CA(OH)2	2.2324D-02
MGCL2	1.9277D+00

MGSO4	1.6864D-01
MGCO3	2.4573D-02
MG(OH)2	2.6995D-02
NACL	1.3432D+00
NA2SO4	2.3465D-01
NAHCO3	1.5526D+00
NA2CO3	6.4976D-02
NAOH	5.4677D-02
KCL	8.1542D-01
K2SO4	1.2062D-01
KHCO3	9.4258D-01
K2CO3	3.3401D-02
KOH	3.3194D-02
HCL	2.8167D+00
H2SO4	6.2987D-01
HBR	3.4329D+00

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.0082	-4.3239	0.3157
ARAGONIT	-8.6809	-8.1781	-0.5028
ARCANITE	-9.7044	-1.8866	-7.8178
BISCHOFI	1.0238	4.4775	-3.4536
BLOEDITE	-8.8781	-2.3470	-6.5311
BRUCITE	-20.1333	-10.8990	-9.2343
BURKEITE	-15.6749	-0.7720	-14.9029
CALCITE	-8.6809	-8.3700	-0.3109
CARNALLI	-0.5689	4.3300	-4.8989
DOLOMITE	-17.9956	-16.9167	-1.0789
EPSOMITE	-5.6372	-1.9674	-3.6698
GAYLUSSI	-17.7319	-9.4210	-8.3109
GLASERIT	-16.3902	-3.8030	-12.5872
GLAUBERI	-7.6756	-5.2450	-2.4306
GYPSUM	-4.2925	-4.5834	0.2909
HALITE	1.4257	1.5574	-0.1317
HEXAHYDR	-5.4950	-1.7161	-3.7789
KAINITE	-6.6613	-0.1930	-6.4683
KALICINI	-16.1546	-10.0580	-6.0966
KIESERIT	-4.7841	-0.1230	-4.6611
LABILE S	-11.6273	-5.6720	-5.9553
LEONHARD	-5.2107	-0.8870	-4.3237
LEONITE	-14.9150	-3.9790	-10.9360
MAGNESIT	-9.3147	-7.7253	-1.5894
MIRABIL	-5.0891	-1.5459	-3.5432
MISENITE	-78.5956	-10.8060	-67.7896
NAHCOLIT	-13.1361	-10.7420	-2.3941
NATRON	-9.7619	-0.8250	-8.9369
NESQUEHO	-9.7412	-5.1670	-4.5742
PCO2	-1.0022	-1.3811	0.3788
PENTAHYD	-5.3528	-1.2850	-4.0678
PIRSSONI	-17.3054	-9.2340	-8.0714
POLYHALI	-22.6470	-13.7440	-8.9030
PORTLAND	-19.4995	-5.1900	-14.3095

SCHOENIT -15.1994 -4.3280 -10.8714
 SYLVITE -1.5928 0.8257 -2.4184
 SYNGENIT -13.8547 -7.4480 -6.4067
 TRONA -21.7606 -11.3840 -10.3766
 1Test Problem 1b: Canadian Shield Brine T-93; Frape et al., 1984, Alk + No scal
 0000020000 0 0 0.00000
 ELEMENTS
 C 15 0.61017E+02
 0 0.00000E+00
 SOLUTION 1
 Canadian Shield Brine T-93.
 9 15 2 5.00 4.00 18.0 1.20
 4 6.400D+04 6 4.500D+04 5 5.100D+03 7 1.990D+02 12 1.080D+03
 14 2.070D+05 22 1.760D+03 16 2.840D+02 15 1.900D+01
 1SOLUTION NUMBER 1
 Canadian Shield Brine T-93.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
TOT ALK	3.540280D-04	-3.4510
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

----DESCRIPTION OF SOLUTION----

PH = 5.0000
 ACTIVITY H2O = 0.7214
 OSMOTIC COEFFICIENT = 1.6512
 IONIC STRENGTH = 8.5902
 TEMPERATURE = 18.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9986 G/CC
 ELECTRICAL BALANCE = -3.0305D-01
 TOTAL ALKALINITY = 3.5403D-04
 ITERATIONS = 5
 TOTAL CARBON = 1.0249D-02

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	2.570E-06	-5.590	1.000E-05	-5.000	3.891E+00	0.590
3	H2O	0.0	7.214E-01	-0.142	7.214E-01	-0.142	1.000E+00	0.000
4	CA+2	2.0	1.815E+00	0.259	1.713E+00	0.234	9.435E-01	-0.025
5	MG+2	2.0	2.386E-01	-0.622	3.980E-01	-0.400	1.668E+00	0.222
6	NA+	1.0	2.225E+00	0.347	1.937E+00	0.287	8.703E-01	-0.060
7	K+	1.0	5.787E-03	-2.238	1.860E-03	-2.730	3.214E-01	-0.493
12	SR+2	2.0	1.401E-02	-1.853	6.812E-03	-2.167	4.861E-01	-0.313
14	CL-	-1.0	6.638E+00	0.822	1.370E+01	1.137	2.064E+00	0.315
15	CO3-2	-2.0	1.126E-06	-5.948	3.613E-09	-8.442	3.208E-03	-2.494
16	SO4-2	-2.0	3.361E-03	-2.473	5.659E-05	-4.247	1.684E-02	-1.774
22	BR-	-1.0	2.504E-02	-1.601	7.677E-02	-1.115	3.065E+00	0.486
31	OH-	-1.0	5.596E-09	-8.252	4.217E-10	-9.375	7.536E-02	-1.123
34	HCO3-	-1.0	3.339E-04	-3.476	9.209E-04	-3.036	2.758E+00	0.441
35	H2CO3	0.0	9.905E-03	-2.004	3.072E-02	-1.513	3.101E+00	0.491
40	HSO4-	-1.0	3.120E-08	-7.506	4.449E-08	-7.352	1.426E+00	0.154
76	CACO3	0.0	7.762E-06	-5.110	7.762E-06	-5.110	1.000E+00	0.000
85	MGOH+	1.0	1.168E-07	-6.933	2.384E-08	-7.623	2.042E-01	-0.690
86	MGCO3	0.0	1.103E-06	-5.957	1.103E-06	-5.957	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	2.6011D-06	1.0000D-05	3.8446D+00
CA+2	1.8155D+00	1.7128D+00	9.4345D-01
MG+2	2.3857D-01	3.9802D-01	1.6684D+00
NA+	2.2254D+00	1.9368D+00	8.7029D-01
K+	5.7867D-03	1.8601D-03	3.2145D-01
SR+2	1.4014D-02	6.8121D-03	4.8610D-01
CL-	6.6382D+00	1.3703D+01	2.0642D+00
CO3-2	9.9913D-06	3.6128D-09	3.6160D-04
SO4-2	3.3613D-03	5.6590D-05	1.6836D-02
BR-	2.5043D-02	7.6765D-02	3.0654D+00
OH-	1.2235D-07	4.2171D-10	3.4466D-03
HCO3-	3.3392D-04	9.2086D-04	2.7577D+00
H2CO3	9.9054D-03	3.0716D-02	3.1010D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.5901D+00
CASO4	1.2603D-01
CACO3	1.8470D-02
CA(OH)2	2.2379D-02
MGCL2	1.9228D+00
MGSO4	1.6760D-01
MGCO3	2.4562D-02
MG(OH)2	2.7062D-02
NACL	1.3403D+00

NA2SO4	2.3363D-01
NAHCO3	1.5492D+00
NA2CO3	6.4941D-02
NAOH	5.4768D-02
KCL	8.1458D-01
K2SO4	1.2027D-01
KHCO3	9.4152D-01
K2CO3	3.3431D-02
KOH	3.3285D-02
HCL	2.8171D+00
H2SO4	6.2899D-01
HBR	3.4329D+00

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.0136	-4.3239	0.3103
ARAGONIT	-8.2084	-8.1781	-0.0303
ARCANITE	-9.7082	-1.8866	-7.8216
BISCHOFI	1.0227	4.4775	-3.4548
BLOEDITE	-8.8877	-2.3470	-6.5407
BRUCITE	-19.1501	-10.8990	-8.2511
BURKEITE	-15.2142	-0.7720	-14.4422
CALCITE	-8.2084	-8.3700	0.1616
CARNALLI	-0.5709	4.3300	-4.9009
DOLOMITE	-17.0507	-16.9167	-0.1340
EPSOMITE	-5.6400	-1.9674	-3.6726
GAYLUSSI	-16.7855	-9.4210	-7.3645
GLASERIT	-16.3988	-3.8030	-12.5958
GLAUBERI	-7.6867	-5.2450	-2.4417
GYPSUM	-4.2972	-4.5834	0.2863
HALITE	1.4239	1.5574	-0.1336
HEXAHYDR	-5.4982	-1.7161	-3.7821
KAINITE	-6.6664	-0.1930	-6.4734
KALICINI	-16.1726	-10.0580	-6.1146
KIESERIT	-4.7892	-0.1230	-4.6662
LABILE S	-11.6434	-5.6720	-5.9714
LEONHARD	-5.2146	-0.8870	-4.3276
LEONITE	-14.9228	-3.9790	-10.9438
MAGNESIT	-8.8423	-7.7253	-1.1170
MIRABIL	-5.0911	-1.5459	-3.5452
MISENITE	-81.5745	-10.8060	-70.7685
NAHCOLIT	-13.1551	-10.7420	-2.4131
NATRON	-9.2860	-0.8250	-8.4610
NESQUEHO	-9.2677	-5.1670	-4.1007
PCO2	-1.5126	-1.3811	-0.1316
PENTAHYD	-5.3564	-1.2850	-4.0714
PIRSONI	-16.3601	-9.2340	-7.1261
POLYHALI	-22.6662	-13.7440	-8.9222
PORTLAND	-18.5163	-5.1900	-13.3263
SCHOENIT	-15.2064	-4.3280	-10.8784
SYLVITE	-1.5936	0.8257	-2.4193
SYNGENIT	-13.8635	-7.4480	-6.4155
TRONA	-21.3067	-11.3840	-9.9227

1Test Problem 1c: Canadian Shield Brine T-93; Frape et al., 1984, calcite + Mac
0060020001 0 1 0.00000

ELEMENTS

C 15 0.61017E+02
0 0.00000E+00

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.00 4.00 18.0 1.20
4 6.400D+04 6 4.500D+04 5 5.100D+03 7 1.990D+02 12 1.080D+03
14 2.070D+05 22 1.760D+03 16 2.840D+02 15 1.900D+01

MINERALS

CALCITE 2 4.0 -8.4 0.00 1 0.000
15 1.00 4 1.00
-1.7183E+02 -7.7993E-02 2.8393E+03 7.1595E+01 0.0000E-01
0 0.00 0.00 0.00 0 0.000

REACTION

15 1.000 4.000

1SOLUTION NUMBER 1

Canadian Shield Brine T-93.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
TOT ALK	3.540280D-04	-3.4510
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

----DESCRIPTION OF SOLUTION----

PH = 5.0000
ACTIVITY H2O = 0.7208
OSMOTIC COEFFICIENT = 1.6522
IONIC STRENGTH = 8.5902
TEMPERATURE = 18.0000
PRESSURE = 1.0000 ATM
DENSITY OF H2O = 0.9986 G/CC
ELECTRICAL BALANCE = -3.0305D-01
TOTAL ALKALINITY = 3.5403D-04
ITERATIONS = 5
TOTAL CARBON = 3.2432D-02

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	MACINNES ACTIVITY	SCALE LOG ACT	MACINNES GAMMA	SCALE LOG GAM
1	H+	1.0	7.997E-06	-5.097	1.000E-05	-5.000	1.251E+00	0.097
3	H2O	0.0	7.208E-01	-0.142	7.208E-01	-0.142	1.000E+00	0.000
4	CA+2	2.0	1.815E+00	0.259	1.783E-01	-0.749	9.823E-02	-1.008
5	MG+2	2.0	2.386E-01	-0.622	4.144E-02	-1.383	1.737E-01	-0.760
6	NA+	1.0	2.225E+00	0.347	6.252E-01	-0.204	2.809E-01	-0.551
7	K+	1.0	5.787E-03	-2.238	5.991E-04	-3.222	1.035E-01	-0.985
12	SR+2	2.0	1.401E-02	-1.853	7.035E-04	-3.153	5.020E-02	-1.299
14	CL-	-1.0	6.638E+00	0.822	4.263E+01	1.630	6.422E+00	0.808
15	CO3-2	-2.0	3.764E-07	-6.424	1.169E-08	-7.932	3.106E-02	-1.508
16	SO4-2	-2.0	3.361E-03	-2.474	5.503E-04	-3.259	1.637E-01	-0.786
22	BR-	-1.0	2.504E-02	-1.601	2.389E-01	-0.622	9.539E+00	0.979
31	OH-	-1.0	1.797E-09	-8.745	4.213E-10	-9.375	2.345E-01	-0.630
34	HCO3-	-1.0	3.473E-04	-3.459	2.980E-03	-2.526	8.581E+00	0.934
35	H2CO3	0.0	3.208E-02	-1.494	9.948E-02	-1.002	3.101E+00	0.491
40	HSO4-	-1.0	9.753E-08	-7.011	4.326E-07	-6.364	4.436E+00	0.647
76	CACO3	0.0	2.615E-06	-5.583	2.615E-06	-5.583	1.000E+00	0.000
85	MGOH+	1.0	3.780E-08	-7.423	2.480E-09	-8.605	6.563E-02	-1.183
86	MGCO3	0.0	3.718E-07	-6.430	3.718E-07	-6.430	1.000E+00	0.000

SPECIES	TOTAL MOL	MACINNES ACTIVITY	MACINNES TOTAL GAMMA
H+	8.0943D-06	1.0000D-05	1.2354D+00
CA+2	1.8155D+00	1.7834D-01	9.8232D-02
MG+2	2.3857D-01	4.1441D-02	1.7371D-01
NA+	2.2254D+00	6.2518D-01	2.8092D-01
K+	5.7867D-03	5.9913D-04	1.0354D-01
SR+2	1.4014D-02	7.0352D-04	5.0202D-02
CL-	6.6382D+00	4.2630D+01	6.4219D+00
CO3-2	3.3633D-06	1.1691D-08	3.4761D-03
SO4-2	3.3612D-03	5.5030D-04	1.6372D-01
BR-	2.5043D-02	2.3887D-01	9.5386D+00
OH-	3.9592D-08	4.2135D-10	1.0642D-02
HCO3-	3.4726D-04	2.9799D-03	8.5812D+00
H2CO3	3.2081D-02	9.9484D-02	3.1010D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.5941D+00
CASO4	1.2682D-01
CACO3	1.8479D-02
CA(OH)2	2.2324D-02
MGCL2	1.9277D+00
MGSO4	1.6864D-01

MGCO3	2.4573D-02
MG(OH)2	2.6995D-02
NACL	1.3432D+00
NA2SO4	2.3465D-01
NAHCO3	1.5526D+00
NA2CO3	6.4976D-02
NAOH	5.4677D-02
KCL	8.1542D-01
K2SO4	1.2062D-01
KHCO3	9.4258D-01
K2CO3	3.3401D-02
KOH	3.3194D-02
HCL	2.8167D+00
H2SO4	6.2987D-01
HBR	3.4329D+00

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.0082	-4.3239	0.3157
ARAGONIT	-8.6809	-8.1781	-0.5028
ARCANITE	-9.7044	-1.8866	-7.8178
BISCHOFI	1.0238	4.4775	-3.4536
BLOEDITE	-8.8781	-2.3470	-6.5311
BRUCITE	-20.1333	-10.8990	-9.2343
BURKEITE	-15.6749	-0.7720	-14.9029
CALCITE	-8.6809	-8.3700	-0.3109
CARNALLI	-0.5689	4.3300	-4.8989
DOLOMITE	-17.9956	-16.9167	-1.0789
EPSOMITE	-5.6372	-1.9674	-3.6698
GAYLUSSI	-17.7319	-9.4210	-8.3109
GLASERIT	-16.3902	-3.8030	-12.5872
GLAUBERI	-7.6756	-5.2450	-2.4306
GYPSUM	-4.2925	-4.5834	0.2909
HALITE	1.4257	1.5574	-0.1317
HEXAHYDR	-5.4950	-1.7161	-3.7789
KAINITE	-6.6613	-0.1930	-6.4683
KALICINI	-16.1546	-10.0580	-6.0966
KIESERIT	-4.7841	-0.1230	-4.6611
LABILE S	-11.6273	-5.6720	-5.9553
LEONHARD	-5.2107	-0.8870	-4.3237
LEONITE	-14.9150	-3.9790	-10.9360
MAGNESIT	-9.3147	-7.7253	-1.5894
MIRABIL	-5.0891	-1.5459	-3.5432
MISENITE	-78.5956	-10.8060	-67.7896
NAHCOLIT	-13.1361	-10.7420	-2.3941
NATRON	-9.7619	-0.8250	-8.9369
NESQUEHO	-9.7412	-5.1670	-4.5742
PCO2	-1.0022	-1.3811	0.3788
PENTAHYD	-5.3528	-1.2850	-4.0678
PIRSSONI	-17.3054	-9.2340	-8.0714
POLYHALI	-22.6470	-13.7440	-8.9030
PORTLAND	-19.4995	-5.1900	-14.3095
SCHOENIT	-15.1994	-4.3280	-10.8714

SYLVITE	-1.5928	0.8257	-2.4184
SYNGENIT	-13.8547	-7.4480	-6.4067
TRONA	-21.7606	-11.3840	-10.3766

1STEP NUMBER 1
0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
TOT ALK	3.243193D-02	-1.4890
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	0.000000D-01	-8.3700	-8.3700	0.0000

+ **

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

** -1.752259D-02 MOLES OF REACTION HAVE BEEN ADDED TO THE SOLUTION
TO REACH THE CALCITE PHASE BOUNDARY.

REACTION IS:

1.000000 MOLES OF C VALENCE = 4.000

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.0124	-4.3239	0.3115
ARAGONIT	-8.3700	-8.1781	-0.1919
ARCANITE	-9.7074	-1.8866	-7.8208
BISCHOFI	1.0229	4.4775	-3.4545
BLOEDITE	-8.8857	-2.3470	-6.5387
BRUCITE	-19.4792	-10.8990	-8.5803

BURKEITE	-15.3733	-0.7720	-14.6013
CALCITE	-8.3700	-8.3700	0.0000
CARNALLI	-0.5705	4.3300	-4.9005
DOLOMITE	-17.3739	-16.9167	-0.4572
EPSOMITE	-5.6394	-1.9674	-3.6720
GAYLUSSI	-17.1090	-9.4210	-7.6880
GLASERIT	-16.3970	-3.8030	-12.5940
GLAUBERI	-7.6843	-5.2450	-2.4393
GYPSUM	-4.2962	-4.5834	0.2873
HALITE	1.4243	1.5574	-0.1332
HEXAHYDR	-5.4975	-1.7161	-3.7814
KAINITE	-6.6653	-0.1930	-6.4723
KALICINI	-16.1699	-10.0580	-6.1119
KIESERIT	-4.7881	-0.1230	-4.6651
LABILE S	-11.6400	-5.6720	-5.9680
LEONHARD	-5.2138	-0.8870	-4.3268
LEONITE	-14.9211	-3.9790	-10.9421
MAGNESIT	-9.0038	-7.7253	-1.2786
MIRABIL	-5.0907	-1.5459	-3.5448
MISENITE	-80.5809	-10.8060	-69.7749
NAHCOLIT	-13.1521	-10.7420	-2.4101
NATRON	-9.4483	-0.8250	-8.6233
NESQUEHO	-9.4295	-5.1670	-4.2625
PCO2	-1.3451	-1.3811	0.0359
PENTAHYD	-5.3556	-1.2850	-4.0706
PIRSSONI	-16.6833	-9.2340	-7.4493
POLYHALI	-22.6622	-13.7440	-8.9182
PORTLAND	-18.8454	-5.1900	-13.6554
SCHOENIT	-15.2049	-4.3280	-10.8769
SYLVITE	-1.5935	0.8257	-2.4191
SYNGENIT	-13.8617	-7.4480	-6.4137
TRONA	-21.4654	-11.3840	-10.0814

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
C	1.490935D-02	-1.8265
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

----DESCRIPTION OF SOLUTION----

PH = 5.3282
 ACTIVITY H2O = 0.7213
 OSMOTIC COEFFICIENT = 1.6515
 IONIC STRENGTH = 8.5902
 TEMPERATURE = 18.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9986 G/CC
 ELECTRICAL BALANCE = -3.0305D-01
 TOTAL ALKALINITY = 3.4974D-04
 ITERATIONS = 5

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	MACINNES ACTIVITY	SCALE LOG ACT	MACINNES GAMMA	SCALE LOG GAM
1	H+	1.0	3.756E-06	-5.425	4.697E-06	-5.328	1.250E+00	0.097
3	H2O	0.0	7.213E-01	-0.142	7.213E-01	-0.142	1.000E+00	0.000
4	CA+2	2.0	1.815E+00	0.259	1.771E-01	-0.752	9.757E-02	-1.011
5	MG+2	2.0	2.386E-01	-0.622	4.116E-02	-1.386	1.725E-01	-0.763
6	NA+	1.0	2.225E+00	0.347	6.229E-01	-0.206	2.799E-01	-0.553
7	K+	1.0	5.787E-03	-2.238	5.980E-04	-3.223	1.033E-01	-0.986
12	SR+2	2.0	1.401E-02	-1.853	7.033E-04	-3.153	5.019E-02	-1.299
14	CL-	-1.0	6.638E+00	0.822	4.264E+01	1.630	6.424E+00	0.808
15	CO3-2	-2.0	7.750E-07	-6.111	2.408E-08	-7.618	3.107E-02	-1.508
16	SO4-2	-2.0	3.361E-03	-2.473	5.486E-04	-3.261	1.632E-01	-0.787
22	BR-	-1.0	2.504E-02	-1.601	2.389E-01	-0.622	9.540E+00	0.980
31	OH-	-1.0	3.828E-09	-8.417	8.977E-10	-9.047	2.345E-01	-0.630
34	HCO3-	-1.0	3.359E-04	-3.474	2.883E-03	-2.540	8.583E+00	0.934
35	H2CO3	0.0	1.457E-02	-1.837	4.517E-02	-1.345	3.101E+00	0.491
40	HSO4-	-1.0	4.566E-08	-7.341	2.026E-07	-6.693	4.437E+00	0.647
76	CACO3	0.0	5.350E-06	-5.272	5.350E-06	-5.272	1.000E+00	0.000
85	MGOH+	1.0	8.000E-08	-7.097	5.249E-09	-8.280	6.561E-02	-1.183
86	MGCO3	0.0	7.606E-07	-6.119	7.606E-07	-6.119	1.000E+00	0.000

SPECIES	TOTAL MOL	MACINNES ACTIVITY	MACINNES TOTAL GAMMA
H+	3.8021D-06	4.6967D-06	1.2353D+00
CA+2	1.8155D+00	1.7713D-01	9.7569D-02
MG+2	2.3857D-01	4.1162D-02	1.7254D-01
NA+	2.2254D+00	6.2288D-01	2.7989D-01
K+	5.7867D-03	5.9796D-04	1.0333D-01
SR+2	1.4014D-02	7.0328D-04	5.0185D-02
CL-	6.6382D+00	4.2645D+01	6.4241D+00
CO3-2	6.8857D-06	2.4080D-08	3.4971D-03
SO4-2	3.3613D-03	5.4863D-04	1.6322D-01
BR-	2.5043D-02	2.3891D-01	9.5403D+00
OH-	8.3826D-08	8.9771D-10	1.0709D-02
HCO3-	3.3588D-04	2.8827D-03	8.5827D+00
H2CO3	1.4567D-02	4.5171D-02	3.1010D+00

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
CACL2	1.5909D+00
CASO4	1.2620D-01
CACO3	1.8472D-02
CA(OH)2	2.2367D-02
MGCL2	1.9238D+00
MGSO4	1.6781D-01
MGCO3	2.4564D-02
MG(OH)2	2.7048D-02
NACL	1.3409D+00
NA2SO4	2.3384D-01
NAHCO3	1.5499D+00
NA2CO3	6.4948D-02
NAOH	5.4749D-02
KCL	8.1476D-01
K2SO4	1.2034D-01
KHCO3	9.4174D-01
K2CO3	3.3425D-02
KOH	3.3266D-02
HCL	2.8170D+00
H2SO4	6.2917D-01
HBR	3.4329D+00

1Test Problem 1d: Canadian Shield Brine T-93, calcite equil. No scale

0060020000 0 1 0.00000

ELEMENTS

C 15 0.61017E+02
0 0.00000E+00

SOLUTION 1

Canadian Shield Brine T-93.

9 15 2 5.00 4.00 18.0 1.20
4 6.400D+04 6 4.500D+04 5 5.100D+03 7 1.990D+02 12 1.080D+03
14 2.070D+05 22 1.760D+03 16 2.840D+02 15 1.900D+01

MINERALS

CALCITE 2 4.0 -8.4 0.00 1 0.000
15 1.00 4 1.00
-1.7183E+02 -7.7993E-02 2.8393E+03 7.1595E+01 0.0000E-01
0 0.00 0.00 0.00 0 0.000

REACTION

15 1.000 4.000

1SOLUTION NUMBER 1

Canadian Shield Brine T-93.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224

NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401300D-02	-1.8534
CL	6.638239D+00	0.8221
TOT ALK	3.540280D-04	-3.4510
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

----DESCRIPTION OF SOLUTION----

PH = 5.0000
 ACTIVITY H2O = 0.7214
 OSMOTIC COEFFICIENT = 1.6512
 IONIC STRENGTH = 8.5902
 TEMPERATURE = 18.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9986 G/CC
 ELECTRICAL BALANCE = -3.0305D-01
 TOTAL ALKALINITY = 3.5403D-04
 ITERATIONS = 5
 TOTAL CARBON = 1.0249D-02

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	2.570E-06	-5.590	1.000E-05	-5.000	3.891E+00	0.590
3	H2O	0.0	7.214E-01	-0.142	7.214E-01	-0.142	1.000E+00	0.000
4	CA+2	2.0	1.815E+00	0.259	1.713E+00	0.234	9.435E-01	-0.025
5	MG+2	2.0	2.386E-01	-0.622	3.980E-01	-0.400	1.668E+00	0.222
6	NA+	1.0	2.225E+00	0.347	1.937E+00	0.287	8.703E-01	-0.060
7	K+	1.0	5.787E-03	-2.238	1.860E-03	-2.730	3.214E-01	-0.493
12	SR+2	2.0	1.401E-02	-1.853	6.812E-03	-2.167	4.861E-01	-0.313
14	CL-	-1.0	6.638E+00	0.822	1.370E+01	1.137	2.064E+00	0.315
15	CO3-2	-2.0	1.126E-06	-5.948	3.613E-09	-8.442	3.208E-03	-2.494
16	SO4-2	-2.0	3.361E-03	-2.473	5.659E-05	-4.247	1.684E-02	-1.774
22	BR-	-1.0	2.504E-02	-1.601	7.677E-02	-1.115	3.065E+00	0.486
31	OH-	-1.0	5.596E-09	-8.252	4.217E-10	-9.375	7.536E-02	-1.123
34	HCO3-	-1.0	3.339E-04	-3.476	9.209E-04	-3.036	2.758E+00	0.441
35	H2CO3	0.0	9.905E-03	-2.004	3.072E-02	-1.513	3.101E+00	0.491
40	HSO4-	-1.0	3.120E-08	-7.506	4.449E-08	-7.352	1.426E+00	0.154
76	CACO3	0.0	7.762E-06	-5.110	7.762E-06	-5.110	1.000E+00	0.000
85	MGOH+	1.0	1.168E-07	-6.933	2.384E-08	-7.623	2.042E-01	-0.690
86	MGCO3	0.0	1.103E-06	-5.957	1.103E-06	-5.957	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
---------	-----------	----------------------	-------------------------

H+	2.6011D-06	1.0000D-05	3.8446D+00
CA+2	1.8155D+00	1.7128D+00	9.4345D-01
MG+2	2.3857D-01	3.9802D-01	1.6684D+00
NA+	2.2254D+00	1.9368D+00	8.7029D-01
K+	5.7867D-03	1.8601D-03	3.2145D-01
SR+2	1.4014D-02	6.8121D-03	4.8610D-01
CL-	6.6382D+00	1.3703D+01	2.0642D+00
CO3-2	9.9913D-06	3.6128D-09	3.6160D-04
SO4-2	3.3613D-03	5.6590D-05	1.6836D-02
BR-	2.5043D-02	7.6765D-02	3.0654D+00
OH-	1.2235D-07	4.2171D-10	3.4466D-03
HCO3-	3.3392D-04	9.2086D-04	2.7577D+00
H2CO3	9.9054D-03	3.0716D-02	3.1010D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.5901D+00
CASO4	1.2603D-01
CACO3	1.8470D-02
CA(OH)2	2.2379D-02
MGCL2	1.9228D+00
MGSO4	1.6760D-01
MGCO3	2.4562D-02
MG(OH)2	2.7062D-02
NACL	1.3403D+00
NA2SO4	2.3363D-01
NAHCO3	1.5492D+00
NA2CO3	6.4941D-02
NAOH	5.4768D-02
KCL	8.1458D-01
K2SO4	1.2027D-01
KHCO3	9.4152D-01
K2CO3	3.3431D-02
KOH	3.3285D-02
HCL	2.8171D+00
H2SO4	6.2899D-01
HBR	3.4329D+00

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.0136	-4.3239	0.3103
ARAGONIT	-8.2084	-8.1781	-0.0303
ARCANITE	-9.7082	-1.8866	-7.8216
BISCHOFI	1.0227	4.4775	-3.4548
BLOEDITE	-8.8877	-2.3470	-6.5407
BRUCITE	-19.1501	-10.8990	-8.2511
BURKEITE	-15.2142	-0.7720	-14.4422
CALCITE	-8.2084	-8.3700	0.1616
CARNALLI	-0.5709	4.3300	-4.9009

DOLOMITE	-17.0507	-16.9167	-0.1340
EPSOMITE	-5.6400	-1.9674	-3.6726
GAYLUSSI	-16.7855	-9.4210	-7.3645
GLASERIT	-16.3988	-3.8030	-12.5958
GLAUBERI	-7.6867	-5.2450	-2.4417
GYPSUM	-4.2972	-4.5834	0.2863
HALITE	1.4239	1.5574	-0.1336
HEXAHYDR	-5.4982	-1.7161	-3.7821
KAINITE	-6.6664	-0.1930	-6.4734
KALICINI	-16.1726	-10.0580	-6.1146
KIESERIT	-4.7892	-0.1230	-4.6662
LABILE S	-11.6434	-5.6720	-5.9714
LEONHARD	-5.2146	-0.8870	-4.3276
LEONITE	-14.9228	-3.9790	-10.9438
MAGNESIT	-8.8423	-7.7253	-1.1170
MIRABIL	-5.0911	-1.5459	-3.5452
MISENITE	-81.5745	-10.8060	-70.7685
NAHCOLIT	-13.1551	-10.7420	-2.4131
NATRON	-9.2860	-0.8250	-8.4610
NESQUEHO	-9.2677	-5.1670	-4.1007
PCO2	-1.5126	-1.3811	-0.1316
PENTAHYD	-5.3564	-1.2850	-4.0714
PIRSSONI	-16.3601	-9.2340	-7.1261
POLYHALI	-22.6662	-13.7440	-8.9222
PORTLAND	-18.5163	-5.1900	-13.3263
SCHOENIT	-15.2064	-4.3280	-10.8784
SYLVITE	-1.5936	0.8257	-2.4193
SYNGENIT	-13.8635	-7.4480	-6.4155
TRONA	-21.3067	-11.3840	-9.9227

1STEP NUMBER 1
0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
TOT ALK	1.024929D-02	-1.9893
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	0.000000D-01	-8.3700	-8.3700	0.0000

+ **

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

** 5.154798D-03 MOLES OF REACTION HAVE BEEN ADDED TO THE SOLUTION
TO REACH THE CALCITE PHASE BOUNDARY.

REACTION IS:

1.000000 MOLES OF C VALENCE = 4.000

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.0123	-4.3239	0.3116
ARAGONIT	-8.3700	-8.1781	-0.1919
ARCANITE	-9.7073	-1.8866	-7.8207
BISCHOFI	1.0230	4.4775	-3.4545
BLOEDITE	-8.8855	-2.3470	-6.5385
BRUCITE	-19.4936	-10.8990	-8.5946
BURKEITE	-15.3731	-0.7720	-14.6011
CALCITE	-8.3700	-8.3700	0.0000
CARNALLI	-0.5705	4.3300	-4.9005
DOLOMITE	-17.3739	-16.9167	-0.4572
EPSOMITE	-5.6393	-1.9674	-3.6720
GAYLUSSI	-17.1090	-9.4210	-7.6880
GLASERIT	-16.3968	-3.8030	-12.5938
GLAUBERI	-7.6841	-5.2450	-2.4391
GYPSUM	-4.2961	-4.5834	0.2874
HALITE	1.4243	1.5574	-0.1331
HEXAHYDR	-5.4974	-1.7161	-3.7814
KAINITE	-6.6652	-0.1930	-6.4722
KALICINI	-16.1627	-10.0580	-6.1047
KIESERIT	-4.7880	-0.1230	-4.6650
LABILE S	-11.6396	-5.6720	-5.9676
LEONHARD	-5.2137	-0.8870	-4.3267
LEONITE	-14.9210	-3.9790	-10.9420
MAGNESIT	-9.0038	-7.7253	-1.2786
MIRABIL	-5.0907	-1.5459	-3.5447
MISENITE	-80.5372	-10.8060	-69.7312
NAHCOLIT	-13.1450	-10.7420	-2.4030
NATRON	-9.4484	-0.8250	-8.6234
NESQUEHO	-9.4295	-5.1670	-4.2625
PCO2	-1.3308	-1.3811	0.0503
PENTAHYD	-5.3556	-1.2850	-4.0706
PIRSSONI	-16.6833	-9.2340	-7.4493
POLYHALI	-22.6618	-13.7440	-8.9178
PORTLAND	-18.8598	-5.1900	-13.6698
SCHOENIT	-15.2047	-4.3280	-10.8767
SYLVITE	-1.5934	0.8257	-2.4191

SYNGENIT	-13.8615	-7.4480	-6.4135
TRONA	-21.4583	-11.3840	-10.0743

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.815465D+00	0.2590
MG	2.385668D-01	-0.6224
NA	2.225424D+00	0.3474
K	5.786697D-03	-2.2376
SR	1.401380D-02	-1.8534
CL	6.638239D+00	0.8221
C	1.540409D-02	-1.8124
S	3.361331D-03	-2.4735
BR	2.504262D-02	-1.6013

----DESCRIPTION OF SOLUTION----

PH =	4.8279
ACTIVITY H2O =	0.7213
OSMOTIC COEFFICIENT =	1.6515
IONIC STRENGTH =	8.5902
TEMPERATURE =	18.0000
PRESSURE =	1.0000 ATM
DENSITY OF H2O =	0.9986 G/CC
ELECTRICAL BALANCE =	-3.0305D-01
TOTAL ALKALINITY =	3.5529D-04
ITERATIONS =	7

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	3.819E-06	-5.418	1.486E-05	-4.828	3.891E+00	0.590
3	H2O	0.0	7.213E-01	-0.142	7.213E-01	-0.142	1.000E+00	0.000
4	CA+2	2.0	1.815E+00	0.259	1.716E+00	0.235	9.452E-01	-0.024
5	MG+2	2.0	2.386E-01	-0.622	3.988E-01	-0.399	1.672E+00	0.223
6	NA+	1.0	2.225E+00	0.347	1.939E+00	0.288	8.712E-01	-0.060
7	K+	1.0	5.787E-03	-2.238	1.861E-03	-2.730	3.216E-01	-0.493
12	SR+2	2.0	1.401E-02	-1.853	6.812E-03	-2.167	4.861E-01	-0.313
14	CL-	-1.0	6.638E+00	0.822	1.370E+01	1.137	2.064E+00	0.315
15	CO3-2	-2.0	7.748E-07	-6.111	2.486E-09	-8.605	3.208E-03	-2.494
16	SO4-2	-2.0	3.361E-03	-2.473	5.665E-05	-4.247	1.685E-02	-1.773
22	BR-	-1.0	2.504E-02	-1.601	7.677E-02	-1.115	3.065E+00	0.486

31	OH-	-1.0	3.765E-09	-8.424	2.837E-10	-9.547	7.535E-02	-1.123
34	HCO3-	-1.0	3.414E-04	-3.467	9.416E-04	-3.026	2.758E+00	0.441
35	H2CO3	0.0	1.506E-02	-1.822	4.669E-02	-1.331	3.101E+00	0.491
40	HSO4-	-1.0	4.642E-08	-7.333	6.618E-08	-7.179	1.426E+00	0.154
76	CACO3	0.0	5.350E-06	-5.272	5.350E-06	-5.272	1.000E+00	0.000
85	MGOH+	1.0	7.869E-08	-7.104	1.607E-08	-7.794	2.042E-01	-0.690
86	MGCO3	0.0	7.606E-07	-6.119	7.606E-07	-6.119	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	3.8658D-06	1.4862D-05	3.8445D+00
CA+2	1.8155D+00	1.7160D+00	9.4524D-01
MG+2	2.3857D-01	3.9877D-01	1.6715D+00
NA+	2.2254D+00	1.9388D+00	8.7118D-01
K+	5.7867D-03	1.8611D-03	3.2162D-01
SR+2	1.4014D-02	6.8121D-03	4.8610D-01
CL-	6.6382D+00	1.3702D+01	2.0641D+00
CO3-2	6.8856D-06	2.4856D-09	3.6099D-04
SO4-2	3.3613D-03	5.6646D-05	1.6853D-02
BR-	2.5043D-02	7.6765D-02	3.0654D+00
OH-	8.2458D-08	2.8369D-10	3.4404D-03
HCO3-	3.4144D-04	9.4159D-04	2.7577D+00
H2CO3	1.5056D-02	4.6688D-02	3.1010D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.5910D+00
CASO4	1.2621D-01
CACO3	1.8472D-02
CA(OH)2	2.2366D-02
MGCL2	1.9240D+00
MGSO4	1.6784D-01
MGCO3	2.4564D-02
MG(OH)2	2.7047D-02
NACL	1.3410D+00
NA2SO4	2.3386D-01
NAHCO3	1.5500D+00
NA2CO3	6.4949D-02
NAOH	5.4747D-02
KCL	8.1478D-01
K2SO4	1.2035D-01
KHCO3	9.4177D-01
K2CO3	3.3424D-02
KOH	3.3264D-02
HCL	2.8170D+00
H2SO4	6.2919D-01
HBR	3.4329D+00

Test problem 2: Equilibration of pure water with a set of minerals accompanied by an irreversible reaction

Using the computer code SNORM (Bodine and Jones, 1986) and the mineral stability data of Harvie and others (1984), it can be shown that the final equilibrium mineral assemblage upon evaporation of sea water to dryness in an environment open to air ($\log P_{CO_2} = -3.5$) at 25 °C is anhydrite ($CaSO_4$), bischofite ($MgCl_2 \cdot 6H_2O$), carnallite ($KMgCl_3 \cdot 6H_2O$), halite ($NaCl$), kieserite ($MgSO_4 \cdot H_2O$), and magnesite ($MgCO_3$). According to the mineralogic phase rule this is an invariant system. That is, once evaporation of sea water reaches saturation with these phases, further evaporation of that solution or other inputs of calcium, magnesium, sodium, potassium, carbon, chloride or sulfate would change only the masses of the solids formed without altering the composition of the equilibrium solution. This test problem uses PHRQPITZ to test this reasoning and at the same time evaluate the internal consistency of SNORM and PHRQPITZ. Because both the SNORM and PHRQPITZ codes use the Harvie and others (1984) data base (though in very different ways) we expect to find with PHRQPITZ that no other phases than those listed above could precipitate from the final equilibrium solution, that is, that the saturation indices of all other minerals in the Harvie and others (1984) data base are less than zero in the equilibrium solution.

The input file for test problem 2 is listed in table 8. The starting solution was taken as pure water at 25 °C. IOPT(2) was arbitrarily selected to define the starting pH via electrical balance. The problem combines irreversible reaction with mineral equilibration (IOPT(3) = 3). An irreversible reaction (the addition of 0.0, 0.1 and 1.0 moles of "sea salt") is used as one means of modeling evaporation (of sea water). The second line of the input file indicates that there will be 3 increments of reaction (NSTEPS = 3) and there are 7 components in the reaction (NCOMPS = 7). The moles of each reaction increment and the stoichiometry of the reactant (sea salt) are defined under STEPS and REACTION input.

The computed results are listed in table 9. Following a print of the input data set, the starting solution, pure water, is speciated and the pH consistent with the thermodynamic data of PHRQPITZ.DATA is 6.9990. The calculations are performed in 3 similar steps, with each case beginning from the pure water starting point (solution 1). In step 1 zero moles of sea salt were added so that pure water could be equilibrated with the mineral set alone. The mass transfers of anhydrite, bischofite, carnallite, halite, kieserite, magnesite, and CO_2 appear under the heading "PHASE BOUNDARIES". The results show that all 6 minerals would dissolve and be accompanied by a small loss of CO_2 to the atmosphere. Examination of the saturation indices shows that all other minerals of the Harvie and others (1984) data base are undersaturated in the equilibrium solution. This confirms the prediction from SNORM and constitutes a partial validation of the two codes. The final equilibrium solution is predominantly a $MgCl_2$ solution of ionic strength 17.3873, pH 6.01 and water activity of 0.3382.

In steps 2 and 3 (table 9) we examine the capability of PHRQPITZ to perform calculations in invariant systems by irreversibly adding 0.1 and 1.0 moles of sea salt to pure water and again equilibrating with the same set of minerals. Comparison of the solution compositions from steps 2 and 3 with that from step 1 shows that all three are identical, as expected. The required mineral mass transfers are summarized under "PHASE BOUNDARIES".

Table 8. -- Line images of the input file to test problem 2

Test Problem 2: Sea Water Invariant Point (equilibration with reaction).

0130020000 3 7 0.0

SOLUTION 1

Pure water

0 0 0 7.0 4.0 25.0 1.0

MINERALS

ANHYDRIT 2 6.00 -4.362 1

4 1.000 16 1.000

422.950 0.0 -18431. -147.708

BISCHOFI 3 0.00 4.455 1

5 1.000 14 2.000 3 6.000

3.524 0.0 277.6

CARNALLI 4 0.00 4.330 0

7 1.000 5 1.000 14 3.000 3 6.000

HALITE 2 0.00 1.570 1

6 1.000 14 1.000

-713.4616 -.1201241 37302.21 262.4583 -2106915.

KIESERIT 3 6.00 -0.123 0

5 1.000 16 1.000 3 1.000

MAGNESIT 2 4.00 -7.834 -6.169 0

5 1.000 15 1.000

PCO2 1 4.0 -1.468 -4.776 1 -3.5

35 1.0

108.3865 0.01985076 -6919.53 -40.45154 669365.0

STEPS

0. .1 1.0

REACTION

4.186 0.0 5.976 0.0 68.567 0. 7.181 0.

149.958 0. 15.043 4. 16.514 6.

END

Table 9. -- Listing of printout from test problem 2

DATA READ FROM DISK

ELEMENTS

SPECIES

LOOK MIN

MEAN GAM

1Test Problem 2: Sea Water Invariant Point (equilibration with reaction).

0130020000 3 7 0.00000

SOLUTION 1

Pure water

0 0 0 7.00 4.00 25.0 1.00

MINERALS

ANHYDRIT 2 6.0 -4.4 0.00 1 0.000

4 1.00 16 1.00

4.2295E+02 0.0000E-01 -1.8431E+04 -1.4771E+02 0.0000E-01

BISCHOFI 3 0.00 4.5 0.00 1 0.000

5 1.00 14 2.00 3 6.00

3.5240E+00 0.0000E-01 2.7760E+02 0.0000E-01 0.0000E-01

CARNALLI 4 0.00 4.3 0.00 0 0.000

7 1.00 5 1.00 14 3.00 3 6.00

HALITE 2 0.00 1.6 0.00 1 0.000

6 1.00 14 1.00

-7.1346E+02 -1.2012E-01 3.7302E+04 2.6246E+02 -2.1069E+06

KIESERIT 3 6.0 -0.12 0.00 0 0.000

5 1.00 16 1.00 3 1.00

MAGNESIT 2 4.0 -7.8 -6.2 0 0.000

5 1.00 15 1.00

PCO2 1 4.0 -1.5 -4.8 1 -3.500

35 1.00

1.0839E+02 1.9851E-02 -6.9195E+03 -4.0452E+01 6.6936E+05

0 0.00 0.00 0.00 0 0.000

STEPS

0.000 0.100 1.00

REACTION

4 0.186 0.000 5 0.976 0.000 6 8.567 0.000 7 0.181 0.00

14 9.958 0.000 15 0.043 4.000 16 0.514 6.000

1SOLUTION NUMBER 1

Pure water

TOTAL MOLALITIES OF ELEMENTS

ELEMENT MOLALITY LOG MOLALITY

PURE WATER

----DESCRIPTION OF SOLUTION----

PH = 6.9990
 ACTIVITY H2O = 1.0000
 OSMOTIC COEFFICIENT = 0.9999
 IONIC STRENGTH = 0.0000
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -2.2785D-19
 TOTAL ALKALINITY = 1.0027D-07
 ITERATIONS = 5

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.003E-07	-6.999	1.002E-07	-6.999	9.996E-01	0.000
3	H2O	0.0	1.000E+00	0.000	1.000E+00	0.000	1.000E+00	0.000
31	OH-	-1.0	1.003E-07	-6.999	1.002E-07	-6.999	9.996E-01	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0027D-07	1.0023D-07	9.9963D-01
OH-	1.0027D-07	1.0023D-07	9.9963D-01

1STEP NUMBER 1

0-----

0 0.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

0.186000 MOLES OF CA	VALENCE =	0.000
0.976000 MOLES OF MG	VALENCE =	0.000
8.567000 MOLES OF NA	VALENCE =	0.000
0.181000 MOLES OF K	VALENCE =	0.000
9.958000 MOLES OF CL	VALENCE =	0.000
0.043000 MOLES OF C	VALENCE =	4.000
0.514000 MOLES OF S	VALENCE =	6.000

 TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	9.307528D-04	-4.3617	-4.3617	0.0000
BISCHOFI	5.651543D+00	4.4551	4.4551	0.0000
CARNALLI	2.151592D-02	4.3300	4.3300	0.0000
HALITE	9.313089D-02	1.5700	1.5700	0.0000
KIESERIT	6.246533D-02	-0.1230	-0.1230	0.0000
MAGNESIT	2.080462D-04	-7.8340	-7.8340	0.0000
PCO2	-1.508047D-04	-4.9679	-1.4679	-3.5000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.3617	-4.3617	0.0000
ARAGONIT	-12.5435	-8.2195	-4.3241
ARCANITE	-7.1823	-1.7760	-5.4063
BISCHOFI	4.4551	4.4551	0.0000
BLOEDITE	-5.3276	-2.3470	-2.9806
BRUCITE	-14.6562	-10.8840	-3.7722
BURKEITE	-19.5583	-0.7720	-18.7863
CALCITE	-12.5435	-8.4062	-4.1373
CARNALLI	4.3300	4.3300	0.0000
DOLOMITE	-20.3775	-17.0830	-3.2945
EPSOMITE	-2.9479	-1.8809	-1.0670
GAYLUSSI	-26.8716	-9.4210	-17.4506
GLASERIT	-12.6695	-3.8030	-8.8665
GLAUBERI	-8.1539	-5.2450	-2.9089
GYPSUM	-5.3033	-4.5805	-0.7228
HALITE	1.5700	1.5700	0.0000
HEXAHYDR	-2.4771	-1.6346	-0.8425
KAINITE	-1.1897	-0.1930	-0.9967
KALICINI	-18.7398	-10.0580	-8.6818
KIESERIT	-0.1230	-0.1230	0.0000
LABILE S	-12.8877	-5.6720	-7.2157
LEONHARD	-1.5354	-0.8870	-0.6484
LEONITE	-8.7178	-3.9790	-4.7388
MAGNESIT	-7.8340	-7.8340	0.0000
MIRABIL	-8.5003	-1.2135	-7.2868
MISENITE	-70.5302	-10.8060	-59.7242
NAHCOLIT	-17.0447	-10.7420	-6.3027
NATRON	-16.6821	-0.8250	-15.8571
NESQUEHO	-9.2464	-5.1670	-4.0794
PCO2	-4.9679	-1.4679	-3.5000
PENTAHYD	-2.0063	-1.2850	-0.7213
PIRSSONI	-25.4591	-9.2340	-16.2251
POLYHALI	-16.4995	-13.7440	-2.7555
PORTLAND	-19.3657	-5.1900	-14.1757
SCHOENIT	-9.6594	-4.3280	-5.3314

SYLVITE	-0.1251	0.8998	-1.0249
SYNGENIT	-12.0148	-7.4480	-4.5668
TRONA	-29.9603	-11.3840	-18.5763

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	9.307528D-04	-3.0312
MG	5.735732D+00	0.7586
NA	9.313089D-02	-1.0309
K	2.151592D-02	-1.6672
CL	1.146076D+01	1.0592
C	5.724154D-05	-4.2423
S	6.339608D-02	-1.1979

----DESCRIPTION OF SOLUTION----

PH = 6.0119
 ACTIVITY H2O = 0.3382
 OSMOTIC COEFFICIENT = 3.4633
 IONIC STRENGTH = 17.3873
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = 2.6802D-10
 TOTAL ALKALINITY = 4.1616D-04
 ITERATIONS = 66

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	5.297E-08	-7.276	9.730E-07	-6.012	1.837E+01	1.264
3	H2O	0.0	3.382E-01	-0.471	3.382E-01	-0.471	1.000E+00	0.000
4	CA+2	2.0	9.308E-04	-3.031	3.533E-03	-2.452	3.796E+00	0.579
5	MG+2	2.0	5.735E+00	0.759	1.810E+02	2.258	3.156E+01	1.499
6	NA+	1.0	9.313E-02	-1.031	1.145E-01	-0.941	1.230E+00	0.090
7	K+	1.0	2.152E-02	-1.667	2.311E-03	-2.636	1.074E-01	-0.969
14	CL-	-1.0	1.146E+01	1.059	3.245E+02	2.511	2.831E+01	1.452
15	CO3-2	-2.0	1.448E-07	-6.839	8.097E-11	-10.092	5.590E-04	-3.253
16	SO4-2	-2.0	6.340E-02	-1.198	1.231E-02	-1.910	1.941E-01	-0.712
31	OH-	-1.0	4.868E-08	-7.313	3.492E-09	-8.457	7.174E-02	-1.144
34	HCO3-	-1.0	4.325E-05	-4.364	1.721E-06	-5.764	3.979E-02	-1.400
35	H2CO3	0.0	1.431E-06	-5.844	1.077E-05	-4.968	7.523E+00	0.876

40	HSO4-	-1.0	1.462E-08	-7.835	1.141E-06	-5.943	7.805E+01	1.892
76	CAC03	0.0	4.053E-10	-9.392	4.053E-10	-9.392	1.000E+00	0.000
85	MGOH+	1.0	3.477E-04	-3.459	9.767E-05	-4.010	2.809E-01	-0.551
86	MGC03	0.0	1.242E-05	-4.906	1.242E-05	-4.906	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	6.7585D-08	9.7296D-07	1.4396D+01
CA+2	9.3075D-04	3.5331D-03	3.7959D+00
MG+2	5.7357D+00	1.8099D+02	3.1555D+01
NA+	9.3131D-02	1.1451D-01	1.2296D+00
K+	2.1516D-02	2.3108D-03	1.0740D-01
CL-	1.1461D+01	3.2445D+02	2.8310D+01
C03-2	1.2564D-05	8.0973D-11	6.4446D-06
S04-2	6.3396D-02	1.2307D-02	1.9413D-01
OH-	3.4778D-04	3.4921D-09	1.0041D-05
HC03-	4.3246D-05	1.7206D-06	3.9787D-02
H2C03	1.4310D-06	1.0766D-05	7.5233D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.4490D+01
CAS04	8.5843D-01
CAC03	4.9460D-03
CA(OH)2	7.2604D-04
MGCL2	2.9353D+01
MGS04	2.4750D+00
MGC03	1.4260D-02
MG(OH)2	1.4708D-03
NACL	5.8999D+00
NA2S04	6.6455D-01
NAHCO3	2.2118D-01
NA2CO3	2.1358D-02
NAOH	3.5137D-03
KCL	1.7437D+00
K2S04	1.3083D-01
KHCO3	6.5370D-02
K2CO3	4.2047D-03
KOH	1.0385D-03
HCL	2.0188D+01
H2S04	3.4266D+00

1STEP NUMBER 2

0-----

0 1.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

0.186000 MOLES OF CA	VALENCE =	0.000
0.976000 MOLES OF MG	VALENCE =	0.000

8.567000 MOLES OF NA	VALENCE = 0.000
0.181000 MOLES OF K	VALENCE = 0.000
9.958000 MOLES OF CL	VALENCE = 0.000
0.043000 MOLES OF C	VALENCE = 4.000
0.514000 MOLES OF S	VALENCE = 6.000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.860000D-02	-1.7305
MG	9.760000D-02	-1.0106
NA	8.567000D-01	-0.0672
K	1.810000D-02	-1.7423
CL	9.958000D-01	-0.0018
C	4.300000D-03	-2.3665
S	5.140000D-02	-1.2890

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-1.766925D-02	-4.3617	-4.3617	0.0000
BISCHOFI	5.609147D+00	4.4551	4.4551	0.0000
CARNALLI	3.415878D-03	4.3300	4.3300	0.0000
HALITE	-7.635695D-01	1.5700	1.5700	0.0000
KIESERIT	2.966496D-02	-0.1230	-0.1230	0.0000
MAGNESIT	-4.091954D-03	-7.8340	-7.8340	0.0000
PCO2	-1.508049D-04	-4.9679	-1.4679	-3.5000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.3617	-4.3617	0.0000
ARAGONIT	-12.5435	-8.2195	-4.3241
ARCANITE	-7.1823	-1.7760	-5.4063
BISCHOFI	4.4551	4.4551	0.0000
BLOEDITE	-5.3276	-2.3470	-2.9806
BRUCITE	-14.6562	-10.8840	-3.7722
BURKEITE	-19.5583	-0.7720	-18.7863
CALCITE	-12.5435	-8.4062	-4.1373
CARNALLI	4.3300	4.3300	0.0000
DOLOMITE	-20.3775	-17.0830	-3.2945
EPSOMITE	-2.9479	-1.8809	-1.0670
GAYLUSSI	-26.8716	-9.4210	-17.4506
GLASERIT	-12.6695	-3.8030	-8.8665

GLAUBERI	-8.1539	-5.2450	-2.9089
GYPSUM	-5.3033	-4.5805	-0.7228
HALITE	1.5700	1.5700	0.0000
HEXAHYDR	-2.4771	-1.6346	-0.8425
KAINITE	-1.1897	-0.1930	-0.9967
KALICINI	-18.7398	-10.0580	-8.6818
KIESERIT	-0.1230	-0.1230	0.0000
LABILE S	-12.8877	-5.6720	-7.2157
LEONHARD	-1.5354	-0.8870	-0.6484
LEONITE	-8.7178	-3.9790	-4.7388
MAGNESIT	-7.8340	-7.8340	0.0000
MIRABIL	-8.5003	-1.2135	-7.2868
MISENITE	-70.5302	-10.8060	-59.7242
NAHCOLIT	-17.0447	-10.7420	-6.3027
NATRON	-16.6822	-0.8250	-15.8572
NESQUEHO	-9.2464	-5.1670	-4.0794
PCO2	-4.9679	-1.4679	-3.5000
PENTAHYD	-2.0063	-1.2850	-0.7213
PIRSSONI	-25.4591	-9.2340	-16.2251
POLYHALI	-16.4995	-13.7440	-2.7555
PORTLAND	-19.3657	-5.1900	-14.1757
SCHOENIT	-9.6594	-4.3280	-5.3314
SYLVITE	-0.1251	0.8998	-1.0249
SYNGENIT	-12.0148	-7.4480	-4.5668
TRONA	-29.9604	-11.3840	-18.5764

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	9.307534D-04	-3.0312
MG	5.735736D+00	0.7586
NA	9.313049D-02	-1.0309
K	2.151588D-02	-1.6672
CL	1.146077D+01	1.0592
C	5.724158D-05	-4.2423
S	6.339571D-02	-1.1979

----DESCRIPTION OF SOLUTION----

PH = 6.0119
 ACTIVITY H2O = 0.3382
 OSMOTIC COEFFICIENT = 3.4633
 IONIC STRENGTH = 17.3873
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -7.8562D-11
 TOTAL ALKALINITY = 4.1616D-04

ITERATIONS = 26

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	5.297E-08	-7.276	9.730E-07	-6.012	1.837E+01	1.264
3	H2O	0.0	3.382E-01	-0.471	3.382E-01	-0.471	1.000E+00	0.000
4	CA+2	2.0	9.308E-04	-3.031	3.533E-03	-2.452	3.796E+00	0.579
5	MG+2	2.0	5.735E+00	0.759	1.810E+02	2.258	3.156E+01	1.499
6	NA+	1.0	9.313E-02	-1.031	1.145E-01	-0.941	1.230E+00	0.090
7	K+	1.0	2.152E-02	-1.667	2.311E-03	-2.636	1.074E-01	-0.969
14	CL-	-1.0	1.146E+01	1.059	3.245E+02	2.511	2.831E+01	1.452
15	CO3-2	-2.0	1.448E-07	-6.839	8.097E-11	-10.092	5.590E-04	-3.253
16	SO4-2	-2.0	6.340E-02	-1.198	1.231E-02	-1.910	1.941E-01	-0.712
31	OH-	-1.0	4.868E-08	-7.313	3.492E-09	-8.457	7.174E-02	-1.144
34	HCO3-	-1.0	4.325E-05	-4.364	1.721E-06	-5.764	3.979E-02	-1.400
35	H2CO3	0.0	1.431E-06	-5.844	1.077E-05	-4.968	7.523E+00	0.876
40	HSO4-	-1.0	1.462E-08	-7.835	1.141E-06	-5.943	7.805E+01	1.892
76	CACO3	0.0	4.053E-10	-9.392	4.053E-10	-9.392	1.000E+00	0.000
85	MGOH+	1.0	3.477E-04	-3.459	9.767E-05	-4.010	2.809E-01	-0.551
86	MGCO3	0.0	1.242E-05	-4.906	1.242E-05	-4.906	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	6.7585D-08	9.7296D-07	1.4396D+01
CA+2	9.3075D-04	3.5331D-03	3.7959D+00
MG+2	5.7357D+00	1.8099D+02	3.1555D+01
NA+	9.3130D-02	1.1451D-01	1.2296D+00
K+	2.1516D-02	2.3108D-03	1.0740D-01
CL-	1.1461D+01	3.2446D+02	2.8310D+01
CO3-2	1.2564D-05	8.0972D-11	6.4445D-06
SO4-2	6.3396D-02	1.2307D-02	1.9413D-01
OH-	3.4779D-04	3.4921D-09	1.0041D-05
HCO3-	4.3246D-05	1.7206D-06	3.9787D-02
H2CO3	1.4310D-06	1.0766D-05	7.5233D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.4490D+01
CASO4	8.5843D-01
CACO3	4.9460D-03
CA(OH)2	7.2604D-04
MGCL2	2.9353D+01
MGSO4	2.4750D+00
MGCO3	1.4260D-02
MG(OH)2	1.4708D-03

NaCl	5.8999D+00
Na2SO4	6.6455D-01
NaHCO3	2.2118D-01
Na2CO3	2.1358D-02
NaOH	3.5137D-03
KCl	1.7437D+00
K2SO4	1.3083D-01
KHCO3	6.5369D-02
K2CO3	4.2047D-03
KOH	1.0385D-03
HCl	2.0188D+01
H2SO4	3.4266D+00

1STEP NUMBER 3

0-----

0 1.000D+00 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

0.186000 MOLES OF CA	VALENCE = 0.000
0.976000 MOLES OF MG	VALENCE = 0.000
8.567000 MOLES OF NA	VALENCE = 0.000
0.181000 MOLES OF K	VALENCE = 0.000
9.958000 MOLES OF CL	VALENCE = 0.000
0.043000 MOLES OF C	VALENCE = 4.000
0.514000 MOLES OF S	VALENCE = 6.000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.860000D-01	-0.7305
MG	9.760000D-01	-0.0106
NA	8.567000D+00	0.9328
K	1.810000D-01	-0.7423
CL	9.958000D+00	0.9982
C	4.300000D-02	-1.3665
S	5.140000D-01	-0.2890

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-1.850692D-01	-4.3617	-4.3617	0.0000
BISCHOFI	5.227546D+00	4.4551	4.4551	0.0000
CARNALLI	-1.594841D-01	4.3300	4.3300	0.0000
HALITE	-8.473869D+00	1.5700	1.5700	0.0000
KIESERIT	-2.655349D-01	-0.1230	-0.1230	0.0000
MAGNESIT	-4.279195D-02	-7.8340	-7.8340	0.0000
PCO2	-1.508048D-04	-4.9679	-1.4679	-3.5000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.3617	-4.3617	0.0000
ARAGONIT	-12.5435	-8.2195	-4.3241
ARCANITE	-7.1823	-1.7760	-5.4063
BISCHOFI	4.4551	4.4551	0.0000
BLOEDITE	-5.3276	-2.3470	-2.9806
BRUCITE	-14.6562	-10.8840	-3.7722
BURKEITE	-19.5583	-0.7720	-18.7863
CALCITE	-12.5435	-8.4062	-4.1373
CARNALLI	4.3300	4.3300	0.0000
DOLOMITE	-20.3775	-17.0830	-3.2945
EPSOMITE	-2.9479	-1.8809	-1.0670
GAYLUSSI	-26.8716	-9.4210	-17.4506
GLASERIT	-12.6695	-3.8030	-8.8665
GLAUBERI	-8.1539	-5.2450	-2.9089
GYPSUM	-5.3033	-4.5805	-0.7228
HALITE	1.5700	1.5700	0.0000
HEXAHYDR	-2.4771	-1.6346	-0.8425
KAINITE	-1.1897	-0.1930	-0.9967
KALICINI	-18.7398	-10.0580	-8.6818
KIESERIT	-0.1230	-0.1230	0.0000
LABILE S	-12.8877	-5.6720	-7.2157
LEONHARD	-1.5354	-0.8870	-0.6484
LEONITE	-8.7178	-3.9790	-4.7388
MAGNESIT	-7.8340	-7.8340	0.0000
MIRABIL	-8.5003	-1.2135	-7.2868
MISENITE	-70.5302	-10.8060	-59.7242
NAHCOLIT	-17.0447	-10.7420	-6.3027
NATRON	-16.6821	-0.8250	-15.8571
NESQUEHO	-9.2464	-5.1670	-4.0794
PCO2	-4.9679	-1.4679	-3.5000
PENTAHYD	-2.0063	-1.2850	-0.7213
PIRSSONI	-25.4591	-9.2340	-16.2251
POLYHALI	-16.4995	-13.7440	-2.7555
PORTLAND	-19.3657	-5.1900	-14.1757
SCHOENIT	-9.6594	-4.3280	-5.3314
SYLVITE	-0.1251	0.8998	-1.0249
SYNGENIT	-12.0148	-7.4480	-4.5668
TRONA	-29.9603	-11.3840	-18.5763

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	9.307531D-04	-3.0312

MG	5.735735D+00	0.7586
NA	9.313064D-02	-1.0309
K	2.151589D-02	-1.6672
CL	1.146077D+01	1.0592
C	5.724155D-05	-4.2423
S	6.339584D-02	-1.1979

-----DESCRIPTION OF SOLUTION-----

PH = 6.0119
 ACTIVITY H2O = 0.3382
 OSMOTIC COEFFICIENT = 3.4633
 IONIC STRENGTH = 17.3873
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = 1.1675D-10
 TOTAL ALKALINITY = 4.1616D-04
 ITERATIONS = 23

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	5.297E-08	-7.276	9.730E-07	-6.012	1.837E+01	1.264
3	H2O	0.0	3.382E-01	-0.471	3.382E-01	-0.471	1.000E+00	0.000
4	CA+2	2.0	9.308E-04	-3.031	3.533E-03	-2.452	3.796E+00	0.579
5	MG+2	2.0	5.735E+00	0.759	1.810E+02	2.258	3.156E+01	1.499
6	NA+	1.0	9.313E-02	-1.031	1.145E-01	-0.941	1.230E+00	0.090
7	K+	1.0	2.152E-02	-1.667	2.311E-03	-2.636	1.074E-01	-0.969
14	CL-	-1.0	1.146E+01	1.059	3.245E+02	2.511	2.831E+01	1.452
15	CO3-2	-2.0	1.448E-07	-6.839	8.097E-11	-10.092	5.590E-04	-3.253
16	SO4-2	-2.0	6.340E-02	-1.198	1.231E-02	-1.910	1.941E-01	-0.712
31	OH-	-1.0	4.868E-08	-7.313	3.492E-09	-8.457	7.174E-02	-1.144
34	HCO3-	-1.0	4.325E-05	-4.364	1.721E-06	-5.764	3.979E-02	-1.400
35	H2CO3	0.0	1.431E-06	-5.844	1.077E-05	-4.968	7.523E+00	0.876
40	HSO4-	-1.0	1.462E-08	-7.835	1.141E-06	-5.943	7.805E+01	1.892
76	CACO3	0.0	4.053E-10	-9.392	4.053E-10	-9.392	1.000E+00	0.000
85	MGOH+	1.0	3.477E-04	-3.459	9.767E-05	-4.010	2.809E-01	-0.551
86	MGCO3	0.0	1.242E-05	-4.906	1.242E-05	-4.906	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	6.7585D-08	9.7296D-07	1.4396D+01
CA+2	9.3075D-04	3.5331D-03	3.7959D+00
MG+2	5.7357D+00	1.8099D+02	3.1555D+01

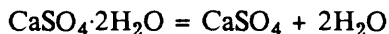
NA+	9.3131D-02	1.1451D-01	1.2296D+00
K+	2.1516D-02	2.3108D-03	1.0740D-01
CL-	1.1461D+01	3.2446D+02	2.8310D+01
CO3-2	1.2564D-05	8.0972D-11	6.4445D-06
SO4-2	6.3396D-02	1.2307D-02	1.9413D-01
OH-	3.4779D-04	3.4921D-09	1.0041D-05
HCO3-	4.3246D-05	1.7206D-06	3.9787D-02
H2CO3	1.4310D-06	1.0766D-05	7.5233D+00

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
CACL2	1.4490D+01
CASO4	8.5843D-01
CACO3	4.9460D-03
CA(OH)2	7.2604D-04
MGCL2	2.9353D+01
MGSO4	2.4750D+00
MGCO3	1.4260D-02
MG(OH)2	1.4708D-03
NACL	5.8999D+00
NA2SO4	6.6455D-01
NAHCO3	2.2118D-01
NA2CO3	2.1358D-02
NAOH	3.5137D-03
KCL	1.7437D+00
K2SO4	1.3083D-01
KHCO3	6.5369D-02
K2CO3	4.2047D-03
KOH	1.0385D-03
HCL	2.0188D+01
H2SO4	3.4266D+00

**Test problem 3: The anhydrite-gypsum phase boundary in the system
NaCl-H₂O at 25 °Celsius**

At constant temperature and pressure equilibrium between anhydrite and gypsum occurs at a single value of the activity of H₂O. This follows from the dehydration reaction



where the solids are in their standard state and the equilibrium constant for the reaction is equal to the square of the activity of H₂O in solution. Using the standard free energy data of Harvie and others (1984), the anhydrite-gypsum phase boundary occurs at a water activity of 0.7773 at 25 °C and 1 atmosphere total pressure. In this test problem we locate the anhydrite-gypsum phase boundary in the NaCl-H₂O system. This is accomplished by adding NaCl as an irreversible reaction until equilibrium with both anhydrite and gypsum occurs. Adding an irreversible reaction to reach a phase boundary is a special case of PHRQPITZ (and PHREEQE) made possible by setting IOPT(3) to 6. The program then adds (or subtracts) the irreversible reaction to solution 1 until saturation is just reached with the first mineral listed under MINERALS input. This first mineral will have zero mass transfer but will be just saturated as the reaction reaches the phase boundary. The mass transfers of other minerals in MINERALS input will vary with extent of reaction.

The input file for test problem 3 is listed in table 10. The second line of input selects IOPT(3) = 6 and there are 2 components in the reaction (Na and Cl) defined under REACTION input. The starting solution is pure water. By placing anhydrite first under MINERALS input we find the amount of NaCl to be dissolved in water, while maintaining equilibrium with gypsum to just reach saturation with anhydrite (though no solid anhydrite is yet present). More advanced problems of this type can be set up to define reaction paths in general (see test problem 5 of Parkhurst and others (1980), and test problem 6 of this report).

The computed results are listed in table 11 showing that dissolution of 5.5737 moles of halite per kg H₂O is required to reach the gypsum-anhydrite phase boundary corresponding to a water activity of 0.7773. The solubility of gypsum in this NaCl solution is 48.01561 mmol per kg H₂O.

Table 10. -- Line images of the input file to test problem 3

Test Problem 3: Find the anhydrite-gypsum phase boundary with NaCl addition.

0060020000 0 2 0.0

SOLUTION 1

pure water

0 0 0 7. 4.0 25. 1.0

MINERALS

ANHYDRIT 2 6.00 -4.362 1

4 1.000 16 1.000
422.950 0.0 -18431. -147.708

GYPSUM 3 6.00 -4.581 1

4 1.0 16 1.0 3 2.0
90.318 0.0 -4213. -32.641

REACTION

61. 0. 141. 0.

END

Table 11. -- Listing of printout from test problem 3

DATA READ FROM DISK

ELEMENTS

SPECIES

LOOK MIN

MEAN GAM

1Test Problem 3: Find the anhydrite-gypsum phase boundary with NaCl addition.

0060020000 0 2 0.00000

SOLUTION 1

pure water

0 0 0 7.00 4.00 25.0 1.00

MINERALS

ANHYDRIT 2 6.0 -4.4 0.00 1 0.000

4 1.00 16 1.00

4.2295E+02 0.0000E-01 -1.8431E+04 -1.4771E+02 0.0000E-01

GYPSUM 3 6.0 -4.6 0.00 1 0.000

4 1.00 16 1.00 3 2.00

9.0318E+01 0.0000E-01 -4.2130E+03 -3.2641E+01 0.0000E-01

0 0.00 0.00 0.00 0 0.000

REACTION

6 1.000 0.000 14 1.000 0.000

1SOLUTION NUMBER 1

pure water

TOTAL MOLALITIES OF ELEMENTS

ELEMENT MOLALITY LOG MOLALITY

PURE WATER

----DESCRIPTION OF SOLUTION----

PH = 7.0000
 ACTIVITY H2O = 1.0000
 OSMOTIC COEFFICIENT = 0.9999
 IONIC STRENGTH = 0.0000
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6175D-10
 TOTAL ALKALINITY = 1.0050D-07
 ITERATIONS = 5

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.000E-07	-7.000	1.000E-07	-7.000	9.996E-01	0.000
3	H2O	0.0	1.000E+00	0.000	1.000E+00	0.000	1.000E+00	0.000
31	OH-	-1.0	1.005E-07	-6.998	1.005E-07	-6.998	9.996E-01	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0004D-07	1.0000D-07	9.9963D-01
OH-	1.0050D-07	1.0046D-07	9.9963D-01

1STEP NUMBER 1
0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

----PHASE BOUNDARIES----

	PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
	ANHYDRIT	0.000000D-01	-4.3617	-4.3617	0.0000
+	** GYPSUM	4.801542D-02	-4.5805	-4.5805	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

** 5.573737D+00 MOLES OF REACTION HAVE BEEN ADDED TO THE SOLUTION
TO REACH THE ANHYDRIT PHASE BOUNDARY.

REACTION IS:

1.000000 MOLES OF NA	VALENCE =	0.000
1.000000 MOLES OF CL	VALENCE =	0.000

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
-------	---------	--------	------------

ANHYDRIT	-4.3617	-4.3617	0.0000
GLAUBERI	-5.9695	-5.2450	-0.7245
GYP SUM	-4.5805	-4.5805	0.0000
HALITE	1.4407	1.5700	-0.1293
LABILE S	-7.7962	-5.6720	-2.1242
MIRABIL	-2.7018	-1.2135	-1.4883
PORTLAND	-16.0990	-5.1900	-10.9090

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	4.801542D-02	-1.3186
NA	5.573737D+00	0.7461
CL	5.573737D+00	0.7461
S	4.801542D-02	-1.3186

----DESCRIPTION OF SOLUTION----

PH = 6.7203
 ACTIVITY H2O = 0.7773
 OSMOTIC COEFFICIENT = 1.2436
 IONIC STRENGTH = 5.7658
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6196D-10
 TOTAL ALKALINITY = 7.9772D-08
 ITERATIONS = 28

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	4.983E-08	-7.302	1.904E-07	-6.720	3.821E+00	0.582
3	H2O	0.0	7.773E-01	-0.109	7.773E-01	-0.109	1.000E+00	0.000
4	CA+2	2.0	4.802E-02	-1.319	4.734E-02	-1.325	9.859E-01	-0.006
6	NA+	1.0	5.574E+00	0.746	5.182E+00	0.715	9.298E-01	-0.032
14	CL-	-1.0	5.574E+00	0.746	5.323E+00	0.726	9.549E-01	-0.020
16	SO4-2	-2.0	4.802E-02	-1.319	9.186E-04	-3.037	1.913E-02	-1.718
31	OH-	-1.0	7.977E-08	-7.098	4.101E-08	-7.387	5.141E-01	-0.289
40	HSO4-	-1.0	2.948E-08	-7.531	1.666E-08	-7.778	5.653E-01	-0.248

UNSCALED UNSCALED

SPECIES	TOTAL MOL	ACTIVITY	TOTAL GAMMA
H+	7.9310D-08	1.9042D-07	2.4009D+00
CA+2	4.8015D-02	4.7336D-02	9.8586D-01
NA+	5.5737D+00	5.1824D+00	9.2979D-01
CL-	5.5737D+00	5.3226D+00	9.5494D-01
SO4-2	4.8015D-02	9.1857D-04	1.9131D-02
OH-	7.9772D-08	4.1010D-08	5.1410D-01

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
CACL2	9.6514D-01
CASO4	1.3733D-01
CA(OH)2	6.3871D-01
NACL	9.4228D-01
NA2SO4	2.5478D-01
NAOH	6.9138D-01
HCL	1.5142D+00
H2SO4	4.7954D-01

Test problem 4: Solubility with incremental temperature variation: halite-water system

PHRQPITZ does not allow the temperature to be varied incrementally without definition of a corresponding increment of reaction. Therefore, in order to calculate the solubility of a mineral at various temperatures, it is necessary to define a null irreversible reaction. In this problem we examine the solubility of halite in pure water between 0 and 300 °C. Halite is the only mineral in the PHRQPITZ data base for which the thermodynamic data are adequate for calculations beyond 100 °C. The input file is listed in table 12. IOPT(3) = 3 indicating that an irreversible reaction (null in this case) will be added at specified increments. IOPT(4) = 3 indicating that the temperature will be defined for each reaction step given under the keyword TEMP. There are 5 steps in the reaction (NSTEPS = 5) and 2 components in the reaction (NCOMPS = 2). We have arbitrarily defined the reaction to add NaCl (REACTION input) with the mass of NaCl added for each increment defined to be 0.0 (STEPS input). The starting solution is pure water and halite equilibrium is obtained by inclusion of the appropriate MINERALS data.

The computed results are listed in table 13. For each increment of null reaction a new temperature is defined and solution 1 (pure water) is equilibrated with halite at the new temperature. The computed solubilities all agree within 2 percent of values given in Pitzer, Peiper and Busey (1984) reflecting the accuracy of the fit of log K to these data (table 1). At temperatures greater than 100 °C the computed vapor pressure of pure water is identical to values given by Pitzer, Peiper and Busey (1984) reaching 84.7 atmospheres (85.5 bars) at 300 °C.

Table 12. -- Line images of the input file to test problem 4

Test Problem 4: Solubility with incremental temperature variation: halite-water

0033020000 5 2 0.0

SOLUTION 1

Pure water

0 0 0 7.0 4.0 25.0 1.0

MINERALS

HALITE 2 0.00 1.570 1

6 1.000 14 1.000

-713.4616 -.1201241 37302.21 262.4583 -2106915.

TEMP

0. 25. 100. 200. 300.

STEPS

0. 0. 0. 0.

REACTION

61. 0. 141. 0.

END

Table 13. -- Listing of printout from test problem 4

DATA READ FROM DISK

ELEMENTS
 SPECIES
 LOOK MIN
 MEAN GAM
 1Test Problem 4: Solubility with incremental temperature variation: halite-water
 0033020000 5 2 0.00000
 SOLUTION 1
 Pure water
 0 0 0 7.00 4.00 25.0 1.00
 MINERALS
 HALITE 2 0.00 1.6 0.00 1 0.000
 6 1.00 14 1.00
 -7.1346E+02 -1.2012E-01 3.7302E+04 2.6246E+02 -2.1069E+06
 0 0.00 0.00 0.00 0 0.000
 TEMP
 0.000 25.0 100. 200. 300.
 STEPS
 0.000 0.000 0.000 0.000 0.000
 REACTION
 6 1.000 0.000 14 1.000 0.000
 1SOLUTION NUMBER 1
 Pure water

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

-----DESCRIPTION OF SOLUTION-----

PH = 7.0000
 ACTIVITY H2O = 1.0000
 OSMOTIC COEFFICIENT = 0.9999
 IONIC STRENGTH = 0.0000
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6175D-10
 TOTAL ALKALINITY = 1.0050D-07
 ITERATIONS = 5

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.000E-07	-7.000	1.000E-07	-7.000	9.996E-01	0.000
3	H2O	0.0	1.000E+00	0.000	1.000E+00	0.000	1.000E+00	0.000
31	OH-	-1.0	1.005E-07	-6.998	1.005E-07	-6.998	9.996E-01	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0004D-07	1.0000D-07	9.9963D-01
OH-	1.0050D-07	1.0046D-07	9.9963D-01

1STEP NUMBER 1

0-----

0 0.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

1.000000 MOLES OF NA	VALENCE =	0.000
1.000000 MOLES OF CL	VALENCE =	0.000
0.00 = NEW TEMPERATURE (C).		

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
HALITE	6.093272D+00	1.5046	1.5046	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
HALITE	1.5046	1.5046	0.0000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
NA	6.093272D+00	0.7849
CL	6.093272D+00	0.7849

-----DESCRIPTION OF SOLUTION-----

PH = 7.0322
 ACTIVITY H2O = 0.7577
 OSMOTIC COEFFICIENT = 1.2641
 IONIC STRENGTH = 6.0933
 TEMPERATURE = 0.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9998 G/CC
 ELECTRICAL BALANCE = -4.6168D-10
 TOTAL ALKALINITY = 1.7906D-08
 ITERATIONS = 24

-----DISTRIBUTION OF SPECIES-----

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.744E-08	-7.758	9.284E-08	-7.032	5.322E+00	0.726
3	H2O	0.0	7.577E-01	-0.121	7.577E-01	-0.121	1.000E+00	0.000
6	NA+	1.0	6.093E+00	0.785	5.653E+00	0.752	9.278E-01	-0.033
14	CL-	-1.0	6.093E+00	0.785	5.653E+00	0.752	9.278E-01	-0.033
31	OH-	-1.0	1.791E-08	-7.747	1.043E-08	-7.982	5.827E-01	-0.235

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.7444D-08	9.2844D-08	5.3223D+00
NA+	6.0933D+00	5.6534D+00	9.2781D-01
CL-	6.0933D+00	5.6534D+00	9.2781D-01
OH-	1.7906D-08	1.0433D-08	5.8267D-01

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
---------	------------

NACL	9.2781D-01
NAOH	7.3526D-01
HCL	2.2222D+00

1STEP NUMBER 2

0-----

0 0.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

1.000000 MOLES OF NA	VALENCE =	0.000
1.000000 MOLES OF CL	VALENCE =	0.000
25.00 = NEW TEMPERATURE (C).		

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
---------	----------	--------------

PURE WATER

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
HALITE	6.099676D+00	1.5700	1.5700	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
HALITE	1.5700	1.5700	0.0000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
---------	----------	--------------

NA	6.099676D+00	0.7853
CL	6.099676D+00	0.7853

-----DESCRIPTION OF SOLUTION-----

PH = 6.5971
 ACTIVITY H2O = 0.7546
 OSMOTIC COEFFICIENT = 1.2813
 IONIC STRENGTH = 6.0997
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6169D-10
 TOTAL ALKALINITY = 5.5149D-08
 ITERATIONS = 24

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	5.469E-08	-7.262	2.529E-07	-6.597	4.624E+00	0.665
3	H2O	0.0	7.546E-01	-0.122	7.546E-01	-0.122	1.000E+00	0.000
6	NA+	1.0	6.100E+00	0.785	6.095E+00	0.785	9.993E-01	0.000
14	CL-	-1.0	6.100E+00	0.785	6.095E+00	0.785	9.993E-01	0.000
31	OH-	-1.0	5.515E-08	-7.258	2.998E-08	-7.523	5.436E-01	-0.265

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	5.4688D-08	2.5286D-07	4.6237D+00
NA+	6.0997D+00	6.0953D+00	9.9929D-01
CL-	6.0997D+00	6.0953D+00	9.9929D-01
OH-	5.5149D-08	2.9980D-08	5.4361D-01

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
NACL	9.9929D-01
NAOH	7.3704D-01
HCL	2.1495D+00

1STEP NUMBER 3

0-----

0 0.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

1.000000 MOLES OF NA	VALENCE =	0.000
1.000000 MOLES OF CL	VALENCE =	0.000

100.00 = NEW TEMPERATURE (C).

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
HALITE	6.624753D+00	1.5605	1.5605	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
HALITE	1.5605	1.5605	0.0000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
NA	6.624753D+00	0.8212
CL	6.624753D+00	0.8212

-----DESCRIPTION OF SOLUTION-----

PH = 5.6149
ACTIVITY H2O = 0.7442
OSMOTIC COEFFICIENT = 1.2375
IONIC STRENGTH = 6.6248
TEMPERATURE = 100.0000
PRESSURE = 1.0000 ATM
DENSITY OF H2O = 0.9584 G/CC
ELECTRICAL BALANCE = -3.8820D-10
TOTAL ALKALINITY = 6.6989D-07

ITERATIONS = 24

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	6.695E-07	-6.174	2.427E-06	-5.615	3.626E+00	0.559
3	H2O	0.0	7.442E-01	-0.128	7.442E-01	-0.128	1.000E+00	0.000
6	NA+	1.0	6.625E+00	0.821	6.029E+00	0.780	9.101E-01	-0.041
14	CL-	-1.0	6.625E+00	0.821	6.029E+00	0.780	9.101E-01	-0.041
31	OH-	-1.0	6.699E-07	-6.174	2.849E-07	-6.545	4.253E-01	-0.371

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	6.6950D-07	2.4274D-06	3.6256D+00
NA+	6.6248D+00	6.0293D+00	9.1012D-01
CL-	6.6248D+00	6.0293D+00	9.1012D-01
OH-	6.6989D-07	2.8490D-07	4.2529D-01

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
NACL	9.1012D-01
NAOH	6.2215D-01
HCL	1.8165D+00

1STEP NUMBER 4

0-----

0 0.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

1.000000 MOLES OF NA	VALENCE =	0.000
1.000000 MOLES OF CL	VALENCE =	0.000
200.00 = NEW TEMPERATURE (C).		

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
HALITE	7.837636D+00	1.2041	1.2041	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
HALITE	1.2041	1.2041	0.0000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
NA	7.837636D+00	0.8942
CL	7.837636D+00	0.8942

-----DESCRIPTION OF SOLUTION-----

PH = 4.6777
ACTIVITY H2O = 0.7415
OSMOTIC COEFFICIENT = 1.0589
IONIC STRENGTH = 7.8376
TEMPERATURE = 200.0000
PRESSURE = 15.3333 ATM
DENSITY OF H2O = 0.8647 G/CC
ELECTRICAL BALANCE = -4.4288D-10
TOTAL ALKALINITY = 7.5880D-06
ITERATIONS = 27

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	7.588E-06	-5.120	2.100E-05	-4.678	2.768E+00	0.442
3	H2O	0.0	7.415E-01	-0.130	7.415E-01	-0.130	1.000E+00	0.000
6	NA+	1.0	7.838E+00	0.894	4.000E+00	0.602	5.103E-01	-0.292

14 CL-	-1.0	7.838E+00	0.894	4.000E+00	0.602	5.103E-01	-0.292
31 OH-	-1.0	7.588E-06	-5.120	1.472E-06	-5.832	1.940E-01	-0.712

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	7.5876D-06	2.1002D-05	2.7680D+00
NA+	7.8376D+00	3.9997D+00	5.1032D-01
CL-	7.8376D+00	3.9997D+00	5.1032D-01
OH-	7.5880D-06	1.4718D-06	1.9397D-01

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
NACL	5.1032D-01
NAOH	3.1462D-01
HCL	1.1885D+00

1STEP NUMBER 5

0-----

0 0.000D-01 MOLES OF REACTION HAVE BEEN ADDED.

REACTION IS:

1.000000 MOLES OF NA	VALENCE =	0.000
1.000000 MOLES OF CL	VALENCE =	0.000
300.00 = NEW TEMPERATURE (C).		

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
HALITE	1.020741D+01	0.2888	0.2888	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
HALITE	0.2888	0.2888	0.0000

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
NA	1.020741D+01	1.0089
CL	1.020741D+01	1.0089

-----DESCRIPTION OF SOLUTION-----

PH = 3.6159
 ACTIVITY H2O = 0.7658
 OSMOTIC COEFFICIENT = 0.7256
 IONIC STRENGTH = 10.2075
 TEMPERATURE = 300.0000
 PRESSURE = 84.7091 ATM
 DENSITY OF H2O = 0.7124 G/CC
 ELECTRICAL BALANCE = -6.5533D-10
 TOTAL ALKALINITY = 1.0802D-04
 ITERATIONS = 32

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.080E-04	-3.967	2.421E-04	-3.616	2.242E+00	0.351
3	H2O	0.0	7.658E-01	-0.116	7.658E-01	-0.116	1.000E+00	0.000
6	NA+	1.0	1.021E+01	1.009	1.394E+00	0.144	1.366E-01	-0.865
14	CL-	-1.0	1.021E+01	1.009	1.394E+00	0.144	1.366E-01	-0.865
31	OH-	-1.0	1.080E-04	-3.967	1.569E-06	-5.805	1.452E-02	-1.838

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0802D-04	2.4214D-04	2.2417D+00
NA+	1.0207D+01	1.3944D+00	1.3660D-01
CL-	1.0207D+01	1.3944D+00	1.3661D-01
OH-	1.0802D-04	1.5685D-06	1.4521D-02

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
NACL	1.3660D-01
NAOH	4.4538D-02
HCL	5.5338D-01

Test problem 5: Fresh water - brine mixing in a closed system.

In this problem we use PHRQPITZ to calculate the solubility of calcite in mixtures of (1) a dilute water in equilibrium with calcite at $\log \text{PCO}_2$ of -2.0 and (2) a sodium chloride brine in equilibrium with halite and calcite at $\log \text{PCO}_2$ of -2.0. The mixing environment is assumed to be closed to CO_2 . All calculations are at 25 °C. The problem is in 3 parts (table 14). The end-member waters are constructed in parts a and b by equilibrating pure water with calcite at a $\log \text{PCO}_2$ of -2.0 (part a), and calcite-halite at $\log \text{PCO}_2 = -2.0$ (part b). The final solution of part a is stored in solution 2 (IOPT(7)=2) and the final solution of part b is stored in solution 1 (IOPT(7)=1). In part c IOPT(3) is set to 1 selecting the mixing case and NSTEPS is 5 indicating that 5 mixtures will be solved. The fraction of solution 1 for each mixture is defined under STEPS input. Each mixture is equilibrated with calcite in a system closed to CO_2 gas. The results are listed in table 15.

Solution 2 (defined in part a) is representative of a dilute calcium bicarbonate ground water in equilibrium with calcite. Solution 1 is a hypothetical NaCl brine in equilibrium with calcite at the same PCO_2 . The solubility of calcite is higher in the NaCl solution owing to the decrease in the mean activity coefficient of CaCO_3 . Although both end-members are saturated with calcite, their mixtures are undersaturated and dissolve calcite owing to the non-linear variation in activity coefficients with concentration (table 15).

Table 14. -- Line images of the input file to test problem 5

Test Problem 5 (part a): Calcite-CO2-H2O equilibrium at log PCO2 = -2.

0050022000 0 0 0.0

SOLUTION 1

Pure water

0 0 0 7.0 4.0 25.0 1.0

MINERALS

CALCITE 2 4.00 -8.406 1

15 1.0 4 1.0

-171.8329 -.077993 2839.319 71.595

PCO2 1 4.0 -1.468 -4.776 1 -2.

35 1.0

108.3865 0.01985076 -6919.53 -40.45154 669365.0

END

Test Problem 5 (part b): Calcite-Halite-CO2-H2O system at log PCO2 = -2.

0050021000 0 0 0.0

SOLUTION 1

Pure water

0 0 0 7.0 4.0 25.0 1.0

MINERALS

CALCITE 2 4.00 -8.406 1

15 1.0 4 1.0

-171.8329 -.077993 2839.319 71.595

PCO2 1 4.0 -1.468 -4.776 1 -2.

35 1.0

108.3865 0.01985076 -6919.53 -40.45154 669365.0

HALITE 2 0.00 1.570 1

6 1.000 14 1.000

-713.4616 -.1201241 37302.21 262.4583 -2106915.

END

Test Problem 5 (part c): Mix solns. 1 and 2 maintaining calcite saturation.

0010020000 5 0 0.0

MINERALS

CALCITE 2 4.00 -8.406 1

15 1.0 4 1.0

-171.8329 -.077993 2839.319 71.595

STEPS

.1 .3 .5 .7 .9

END

Table 15. -- Listing of printout from test problem 5

DATA READ FROM DISK

ELEMENTS

SPECIES

LOOK MIN

MEAN GAM

1Test Problem 5 (part a): Calcite-CO2-H2O equilibrium at log PCO2 = -2.

0050022000 0 0 0.00000

SOLUTION 1

Pure water

0 0 0 7.00 4.00 25.0 1.00

MINERALS

CALCITE 2 4.0 -8.4 0.00 1 0.000

15 1.00 4 1.00

-1.7183E+02 -7.7993E-02 2.8393E+03 7.1595E+01 0.0000E-01

PCO2 1 4.0 -1.5 -4.8 1 -2.000

35 1.00

1.0839E+02 1.9851E-02 -6.9195E+03 -4.0452E+01 6.6936E+05

0 0.00 0.00 0.00 0 0.000

1SOLUTION NUMBER 1

Pure water

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
---------	----------	--------------

PURE WATER

----DESCRIPTION OF SOLUTION----

PH = 7.0000
 ACTIVITY H2O = 1.0000
 OSMOTIC COEFFICIENT = 0.9999
 IONIC STRENGTH = 0.0000
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6175D-10
 TOTAL ALKALINITY = 1.0050D-07
 ITERATIONS = 5

----- DISTRIBUTION OF SPECIES -----

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.000E-07	-7.000	1.000E-07	-7.000	9.996E-01	0.000
3	H2O	0.0	1.000E+00	0.000	1.000E+00	0.000	1.000E+00	0.000
31	OH-	-1.0	1.005E-07	-6.998	1.005E-07	-6.998	9.996E-01	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0004D-07	1.0000D-07	9.9963D-01
OH-	1.0050D-07	1.0046D-07	9.9963D-01

1STEP NUMBER 1

0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
---------	----------	--------------

PURE WATER

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	1.741541D-03	-8.4062	-8.4062	0.0000
PCO2	2.071791D-03	-3.4679	-1.4679	-2.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
PCO2	-3.4679	-1.4679	-2.0000
PORTLAND	-16.2576	-5.1900	-11.0676

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.741541D-03	-2.7591
C	3.813332D-03	-2.4187

-----DESCRIPTION OF SOLUTION-----

PH = 7.3147
 ACTIVITY H2O = 0.9999
 OSMOTIC COEFFICIENT = 0.9576
 IONIC STRENGTH = 0.0052
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = 4.3136D-09
 TOTAL ALKALINITY = 3.4831D-03
 ITERATIONS = 10

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	5.274E-08	-7.278	4.845E-08	-7.315	9.188E-01	-0.037
3	H2O	0.0	9.999E-01	0.000	9.999E-01	0.000	1.000E+00	0.000
4	CA+2	2.0	1.736E-03	-2.760	1.286E-03	-2.891	7.406E-01	-0.130
15	CO3-2	-2.0	4.251E-06	-5.372	3.052E-06	-5.515	7.181E-01	-0.144
31	OH-	-1.0	2.250E-07	-6.648	2.073E-07	-6.683	9.214E-01	-0.036
34	HCO3-	-1.0	3.463E-03	-2.461	3.230E-03	-2.491	9.327E-01	-0.030
35	H2CO3	0.0	3.402E-04	-3.468	3.405E-04	-3.468	1.001E+00	0.000
76	CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	5.2736D-08	4.8455D-08	9.1882D-01
CA+2	1.7415D-03	1.2857D-03	7.3826D-01
CO3-2	9.8103D-06	3.0523D-06	3.1113D-01
OH-	2.2499D-07	2.0731D-07	9.2144D-01
HCO3-	3.4633D-03	3.2301D-03	9.3267D-01
H2CO3	3.4024D-04	3.4045D-04	1.0006D+00

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
CACO3	4.7926D-01

CA(OH)2 8.5582D-01

1Test Problem 5 (part b): Calcite-Halite-CO2-H2O system at log PCO2 = -2.

0050021000 0 0

0.00000

SOLUTION 1

Pure water

0 0 0 7.00 4.00 25.0 1.00

MINERALS

CALCITE 2 4.0 -8.4 0.00 1 0.000

15 1.00 4 1.00

-1.7183E+02 -7.7993E-02 2.8393E+03 7.1595E+01 0.0000E-01

PCO2 1 4.0 -1.5 -4.8 1 -2.000

35 1.00

1.0839E+02 1.9851E-02 -6.9195E+03 -4.0452E+01 6.6936E+05

HALITE 2 0.00 1.6 0.00 1 0.000

6 1.00 14 1.00

-7.1346E+02 -1.2012E-01 3.7302E+04 2.6246E+02 -2.1069E+06

0 0.00 0.00 0.00 0 0.000

1SOLUTION NUMBER 1

Pure water

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

----DESCRIPTION OF SOLUTION----

PH = 7.0000
 ACTIVITY H2O = 1.0000
 OSMOTIC COEFFICIENT = 0.9999
 IONIC STRENGTH = 0.0000
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6175D-10
 TOTAL ALKALINITY = 1.0050D-07
 ITERATIONS = 5

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.000E-07	-7.000	1.000E-07	-7.000	9.996E-01	0.000

3 H2O	0.0	1.000E+00	0.000	1.000E+00	0.000	1.000E+00	0.000
31 OH-	-1.0	1.005E-07	-6.998	1.005E-07	-6.998	9.996E-01	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0004D-07	1.0000D-07	9.9963D-01
OH-	1.0050D-07	1.0046D-07	9.9963D-01

1STEP NUMBER 1
0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
PURE WATER		

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	2.544505D-03	-8.4062	-8.4062	0.0000
PCO2	2.596886D-03	-3.4679	-1.4679	-2.0000
HALITE	6.098571D+00	1.5700	1.5700	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
GAYLUSSI	-13.3763	-9.4210	-3.9553
HALITE	1.5700	1.5700	0.0000
NAHCOLIT	-12.3127	-10.7420	-1.5707
NATRON	-5.5818	-0.8250	-4.7568
PCO2	-3.4679	-1.4679	-2.0000
PIRSSONI	-13.0093	-9.2340	-3.7753
PORTLAND	-16.3799	-5.1900	-11.1899
TRONA	-16.9157	-11.3840	-5.5317

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.544505D-03	-2.5944
NA	6.098571D+00	0.7852
CL	6.098571D+00	0.7852
C	5.141390D-03	-2.2889

-----DESCRIPTION OF SOLUTION-----

PH = 7.1698
 ACTIVITY H2O = 0.7545
 OSMOTIC COEFFICIENT = 1.2813
 IONIC STRENGTH = 6.1062
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.0491D-10
 TOTAL ALKALINITY = 5.0890D-03
 ITERATIONS = 24

-----DISTRIBUTION OF SPECIES-----

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.464E-08	-7.834	6.764E-08	-7.170	4.619E+00	0.665
3	H2O	0.0	7.545E-01	-0.122	7.545E-01	-0.122	1.000E+00	0.000
4	CA+2	2.0	2.539E-03	-2.595	3.320E-03	-2.479	1.308E+00	0.117
6	NA+	1.0	6.099E+00	0.785	6.089E+00	0.785	9.984E-01	-0.001
14	CL-	-1.0	6.099E+00	0.785	6.102E+00	0.785	1.001E+00	0.000
15	CO3-2	-2.0	4.863E-05	-4.313	1.182E-06	-5.927	2.430E-02	-1.614
31	OH-	-1.0	2.065E-07	-6.685	1.121E-07	-6.951	5.426E-01	-0.266
34	HCO3-	-1.0	4.980E-03	-2.303	1.746E-03	-2.758	3.506E-01	-0.455
35	H2CO3	0.0	1.068E-04	-3.972	3.405E-04	-3.468	3.189E+00	0.504
76	CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.4644D-08	6.7638D-08	4.6187D+00
CA+2	2.5445D-03	3.3202D-03	1.3048D+00
NA+	6.0986D+00	6.0888D+00	9.9840D-01
CL-	6.0986D+00	6.1018D+00	1.0005D+00
CO3-2	5.4192D-05	1.1820D-06	2.1811D-02
OH-	2.0652D-07	1.1206D-07	5.4261D-01

HCO3-	4.9804D-03	1.7460D-03	3.5058D-01
H2CO3	1.0676D-04	3.4046D-04	3.1888D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.0931D+00
CACO3	1.6870D-01
CA(OH)2	7.2696D-01
NACL	9.9947D-01
NAHCO3	5.9162D-01
NA2CO3	2.7910D-01
NAOH	7.3603D-01
HCL	2.1497D+00

1Test Problem 5 (part c): Mix solns. 1 and 2 maintaining calcite saturation.
0010020000 5 0 0.00000

MINERALS

CALCITE	2	4.0	-8.4	0.00	1	0.000
15	1.00	4	1.00			
-1.7183E+02	-7.7993E-02	2.8393E+03	7.1595E+01	0.0000E-01		
0	0.00	0.00	0.00	0		0.000

STEPS

0.100	0.300	0.500	0.700	0.900
-------	-------	-------	-------	-------

1STEP NUMBER 1

0-----

0 0.100 = FRACTION OF SOLUTION 1. 0.900 = FRACTION OF SOLUTION 2.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.821837D-03	-2.7395
NA	6.098571D-01	-0.2148
CL	6.098571D-01	-0.2148
C	3.946138D-03	-2.4038

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	3.052849D-04	-8.4062	-8.4062	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
GAYLUSSI	-14.3220	-9.4210	-4.9010
HALITE	-0.7779	1.5700	-2.3479
NAHCOLIT	-13.3443	-10.7420	-2.6023
NATRON	-5.9600	-0.8250	-5.1350
PCO2	-4.1316	-1.4679	-2.6637
PIRSSONI	-14.2954	-9.2340	-5.0614
PORTLAND	-15.6028	-5.1900	-10.4128
TRONA	-19.2335	-11.3840	-7.8495

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.127122D-03	-2.6722
NA	6.098571D-01	-0.2148
CL	6.098571D-01	-0.2148
C	4.251423D-03	-2.3715

----DESCRIPTION OF SOLUTION----

PH = 7.8627
 ACTIVITY H2O = 0.9798
 OSMOTIC COEFFICIENT = 0.9228
 IONIC STRENGTH = 0.6163
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6165D-10
 TOTAL ALKALINITY = 4.2543D-03
 ITERATIONS = 6

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.718E-08	-7.765	1.372E-08	-7.863	7.984E-01	-0.098
3	H2O	0.0	9.798E-01	-0.009	9.798E-01	-0.009	1.000E+00	0.000
4	CA+2	2.0	2.122E-03	-2.673	4.847E-04	-3.314	2.285E-01	-0.641
6	NA+	1.0	6.099E-01	-0.215	4.075E-01	-0.390	6.682E-01	-0.175

14 CL-	-1.0	6.099E-01	-0.215	4.092E-01	-0.388	6.710E-01	-0.173
15 CO3-2	-2.0	6.187E-05	-4.209	8.096E-06	-5.092	1.309E-01	-0.883
31 OH-	-1.0	1.137E-06	-5.944	7.176E-07	-6.144	6.310E-01	-0.200
34 HCO3-	-1.0	4.118E-03	-2.385	2.425E-03	-2.615	5.890E-01	-0.230
35 H2CO3	0.0	6.573E-05	-4.182	7.386E-05	-4.132	1.124E+00	0.051
76 CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.7181D-08	1.3718D-08	7.9842D-01
CA+2	2.1271D-03	4.8474D-04	2.2789D-01
NA+	6.0986D-01	4.0750D-01	6.6819D-01
CL-	6.0986D-01	4.0921D-01	6.7099D-01
CO3-2	6.7429D-05	8.0958D-06	1.2006D-01
OH-	1.1371D-06	7.1756D-07	6.3104D-01
HCO3-	4.1183D-03	2.4255D-03	5.8896D-01
H2CO3	6.5727D-05	7.3859D-05	1.1237D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	4.6815D-01
CACO3	1.6541D-01
CA(OH)2	4.4938D-01
NACL	6.6959D-01
NAHCO3	6.2733D-01
NA2CO3	3.7705D-01
NAOH	6.4935D-01
HCL	7.3194D-01

1STEP NUMBER 2

0-----

0 0.300 = FRACTION OF SOLUTION 1. 0.700 = FRACTION OF SOLUTION 2.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.982430D-03	-2.7028
NA	1.829571D+00	0.2623
CL	1.829571D+00	0.2623
C	4.211749D-03	-2.3755

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	3.191327D-04	-8.4062	-8.4062	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
GAYLUSSI	-13.5474	-9.4210	-4.1264
HALITE	0.1659	1.5700	-1.4041
NAHCOLIT	-12.9227	-10.7420	-2.1807
NATRON	-5.2809	-0.8250	-4.4559
PCO2	-4.1394	-1.4679	-2.6715
PIRSSONI	-13.4635	-9.2340	-4.2295
PORTLAND	-15.6140	-5.1900	-10.4240
TRONA	-17.9800	-11.3840	-6.5960

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.301563D-03	-2.6380
NA	1.829571D+00	0.2623
CL	1.829571D+00	0.2623
C	4.530882D-03	-2.3438

-----DESCRIPTION OF SOLUTION-----

PH =	7.8391
ACTIVITY H2O =	0.9377
OSMOTIC COEFFICIENT =	0.9748
IONIC STRENGTH =	1.8366
TEMPERATURE =	25.0000
PRESSURE =	1.0000 ATM
DENSITY OF H2O =	0.9971 G/CC
ELECTRICAL BALANCE =	-4.6177D-10
TOTAL ALKALINITY =	4.6031D-03
ITERATIONS =	5

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED		UNSCALED	
					ACTIVITY	LOG ACT	GAMMA	LOG GAM
1	H+	1.0	1.320E-08	-7.879	1.449E-08	-7.839	1.097E+00	0.040
3	H2O	0.0	9.377E-01	-0.028	9.377E-01	-0.028	1.000E+00	0.000
4	CA+2	2.0	2.296E-03	-2.639	5.751E-04	-3.240	2.505E-01	-0.601
6	NA+	1.0	1.830E+00	0.262	1.209E+00	0.082	6.607E-01	-0.180
14	CL-	-1.0	1.830E+00	0.262	1.212E+00	0.084	6.626E-01	-0.179
15	CO3-2	-2.0	1.167E-04	-3.933	6.824E-06	-5.166	5.846E-02	-1.233
31	OH-	-1.0	1.164E-06	-5.934	6.503E-07	-6.187	5.587E-01	-0.253
34	HCO3-	-1.0	4.357E-03	-2.361	2.159E-03	-2.666	4.954E-01	-0.305
35	H2CO3	0.0	5.120E-05	-4.291	7.254E-05	-4.139	1.417E+00	0.151
76	CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED	UNSCALED
		ACTIVITY	TOTAL GAMMA
H+	1.3203D-08	1.4485D-08	1.0971D+00
CA+2	2.3016D-03	5.7510D-04	2.4987D-01
NA+	1.8296D+00	1.2087D+00	6.6065D-01
CL-	1.8296D+00	1.2123D+00	6.6262D-01
CO3-2	1.2229D-04	6.8239D-06	5.5801D-02
OH-	1.1639D-06	6.5030D-07	5.5872D-01
HCO3-	4.3574D-03	2.1588D-03	4.9543D-01
H2CO3	5.1195D-05	7.2537D-05	1.4169D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	4.7872D-01
CACO3	1.1808D-01
CA(OH)2	4.2727D-01
NACL	6.6163D-01
NAHCO3	5.7211D-01
NA2CO3	2.8987D-01
NAOH	6.0755D-01
HCL	8.5262D-01

1STEP NUMBER 3

0-----

0 0.500 = FRACTION OF SOLUTION 1.

0.500 = FRACTION OF SOLUTION 2.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.143023D-03	-2.6690

NA	3.049286D+00	0.4842
CL	3.049286D+00	0.4842
C	4.477361D-03	-2.3490

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	2.584929D-04	-8.4062	-8.4062	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
GAYLUSSI	-13.3252	-9.4210	-3.9042
HALITE	0.6773	1.5700	-0.8927
NAHCOLIT	-12.6872	-10.7420	-1.9452
NATRON	-5.1695	-0.8250	-4.3445
PCO2	-3.9791	-1.4679	-2.5112
PIRSSONI	-13.1749	-9.2340	-3.9409
PORTLAND	-15.7965	-5.1900	-10.6065
TRONA	-17.4559	-11.3840	-6.0719

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.401516D-03	-2.6195
NA	3.049286D+00	0.4842
CL	3.049286D+00	0.4842
C	4.735854D-03	-2.3246

----DESCRIPTION OF SOLUTION----

PH =	7.6807
ACTIVITY H2O =	0.8910
OSMOTIC COEFFICIENT =	1.0489
IONIC STRENGTH =	3.0566
TEMPERATURE =	25.0000
PRESSURE =	1.0000 ATM
DENSITY OF H2O =	0.9971 G/CC

ELECTRICAL BALANCE = -4.6177D-10
 TOTAL ALKALINITY = 4.8030D-03
 ITERATIONS = 5

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.289E-08	-7.890	2.086E-08	-7.681	1.619E+00	0.209
3	H2O	0.0	8.910E-01	-0.050	8.910E-01	-0.050	1.000E+00	0.000
4	CA+2	2.0	2.396E-03	-2.621	8.678E-04	-3.062	3.622E-01	-0.441
6	NA+	1.0	3.049E+00	0.484	2.178E+00	0.338	7.143E-01	-0.146
14	CL-	-1.0	3.049E+00	0.484	2.184E+00	0.339	7.161E-01	-0.145
15	CO3-2	-2.0	1.196E-04	-3.922	4.522E-06	-5.345	3.782E-02	-1.422
31	OH-	-1.0	7.980E-07	-6.098	4.291E-07	-6.367	5.377E-01	-0.269
34	HCO3-	-1.0	4.552E-03	-2.342	2.060E-03	-2.686	4.526E-01	-0.344
35	H2CO3	0.0	5.873E-05	-4.231	1.049E-04	-3.979	1.786E+00	0.252
76	CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.2888D-08	2.0861D-08	1.6187D+00
CA+2	2.4015D-03	8.6777D-04	3.6134D-01
NA+	3.0493D+00	2.1782D+00	7.1433D-01
CL-	3.0493D+00	2.1837D+00	7.1615D-01
CO3-2	1.2512D-04	4.5224D-06	3.6143D-02
OH-	7.9797D-07	4.2910D-07	5.3774D-01
HCO3-	4.5520D-03	2.0604D-03	4.5264D-01
H2CO3	5.8731D-05	1.0492D-04	1.7865D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	5.7013D-01
CACO3	1.1428D-01
CA(OH)2	4.7100D-01
NACL	7.1524D-01
NAHCO3	5.6863D-01
NA2CO3	2.6420D-01
NAOH	6.1978D-01
HCL	1.0767D+00

1STEP NUMBER 4

0-----

0 0.700 = FRACTION OF SOLUTION 1.

0.300 = FRACTION OF SOLUTION 2.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.303616D-03	-2.6376
NA	4.269000D+00	0.6303
CL	4.269000D+00	0.6303
C	4.742973D-03	-2.3239

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	1.619806D-04	-8.4062	-8.4062	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
GAYLUSSI	-13.2781	-9.4210	-3.8571
HALITE	1.0707	1.5700	-0.4993
NAHCOLIT	-12.5115	-10.7420	-1.7695
NATRON	-5.2514	-0.8250	-4.4264
PCO2	-3.7782	-1.4679	-2.3102
PIRSSONI	-13.0504	-9.2340	-3.8164
PORTLAND	-16.0233	-5.1900	-10.8333
TRONA	-17.1557	-11.3840	-5.7717

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.465596D-03	-2.6081
NA	4.269000D+00	0.6303
CL	4.269000D+00	0.6303
C	4.904953D-03	-2.3094

----DESCRIPTION OF SOLUTION----

PH = 7.4844
 ACTIVITY H2O = 0.8396
 OSMOTIC COEFFICIENT = 1.1354
 IONIC STRENGTH = 4.2765
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6175D-10
 TOTAL ALKALINITY = 4.9312D-03
 ITERATIONS = 5

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.341E-08	-7.873	3.278E-08	-7.484	2.445E+00	0.388
3	H2O	0.0	8.396E-01	-0.076	8.396E-01	-0.076	1.000E+00	0.000
4	CA+2	2.0	2.460E-03	-2.609	1.431E-03	-2.844	5.818E-01	-0.235
6	NA+	1.0	4.269E+00	0.630	3.426E+00	0.535	8.026E-01	-0.095
14	CL-	-1.0	4.269E+00	0.630	3.434E+00	0.536	8.045E-01	-0.094
15	CO3-2	-2.0	9.420E-05	-4.026	2.742E-06	-5.562	2.910E-02	-1.536
31	OH-	-1.0	4.822E-07	-6.317	2.573E-07	-6.590	5.337E-01	-0.273
34	HCO3-	-1.0	4.731E-03	-2.325	1.963E-03	-2.707	4.149E-01	-0.382
35	H2CO3	0.0	7.399E-05	-4.131	1.667E-04	-3.778	2.252E+00	0.353
76	CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.3406D-08	3.2778D-08	2.4451D+00
CA+2	2.4656D-03	1.4313D-03	5.8052D-01
NA+	4.2690D+00	3.4264D+00	8.0262D-01
CL-	4.2690D+00	3.4344D+00	8.0449D-01
CO3-2	9.9762D-05	2.7417D-06	2.7483D-02
OH-	4.8220D-07	2.5734D-07	5.3367D-01
HCO3-	4.7312D-03	1.9628D-03	4.1486D-01
H2CO3	7.3992D-05	1.6666D-04	2.2524D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	7.2158D-01
CACO3	1.2631D-01
CA(OH)2	5.4885D-01
NACL	8.0355D-01
NAHCO3	5.7704D-01
NA2CO3	2.6063D-01
NAOH	6.5447D-01

HCL 1.4025D+00

1STEP NUMBER 5

0-----

0 0.900 = FRACTION OF SOLUTION 1. 0.100 = FRACTION OF SOLUTION 2.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.464208D-03	-2.6083
NA	5.488714D+00	0.7395
CL	5.488714D+00	0.7395
C	5.008585D-03	-2.3003

-----PHASE BOUNDARIES-----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
CALCITE	5.382800D-05	-8.4062	-8.4062	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ARAGONIT	-8.4062	-8.2195	-0.1868
CALCITE	-8.4062	-8.4062	0.0000
GAYLUSSI	-13.3265	-9.4210	-3.9055
HALITE	1.4113	1.5700	-0.1587
NAHCOLIT	-12.3719	-10.7420	-1.6299
NATRON	-5.4492	-0.8250	-4.6242
PCO2	-3.5699	-1.4679	-2.1020
PIRSSONI	-13.0092	-9.2340	-3.7752
PORTLAND	-16.2614	-5.1900	-11.0714
TRONA	-16.9748	-11.3840	-5.5908

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	2.518036D-03	-2.5989
NA	5.488714D+00	0.7395

CL	5.488714D+00	0.7395
C	5.062413D-03	-2.2956

-----DESCRIPTION OF SOLUTION-----

PH =	7.2754
ACTIVITY H2O =	0.7838
OSMOTIC COEFFICIENT =	1.2309
IONIC STRENGTH =	5.4963
TEMPERATURE =	25.0000
PRESSURE =	1.0000 ATM
DENSITY OF H2O =	0.9971 G/CC
ELECTRICAL BALANCE =	-4.6175D-10
TOTAL ALKALINITY =	5.0361D-03
ITERATIONS =	5

DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.421E-08	-7.847	5.305E-08	-7.275	3.732E+00	0.572
3	H2O	0.0	7.838E-01	-0.106	7.838E-01	-0.106	1.000E+00	0.000
4	CA+2	2.0	2.512E-03	-2.600	2.486E-03	-2.605	9.894E-01	-0.005
6	NA+	1.0	5.489E+00	0.739	5.072E+00	0.705	9.240E-01	-0.034
14	CL-	-1.0	5.489E+00	0.739	5.083E+00	0.706	9.261E-01	-0.033
15	CO3-2	-2.0	6.264E-05	-4.203	1.579E-06	-5.802	2.520E-02	-1.599
31	OH-	-1.0	2.758E-07	-6.559	1.484E-07	-6.828	5.382E-01	-0.269
34	HCO3-	-1.0	4.899E-03	-2.310	1.829E-03	-2.738	3.733E-01	-0.428
35	H2CO3	0.0	9.480E-05	-4.023	2.692E-04	-3.570	2.840E+00	0.453
76	CACO3	0.0	5.560E-06	-5.255	5.560E-06	-5.255	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.4215D-08	5.3045D-08	3.7318D+00
CA+2	2.5180D-03	2.4858D-03	9.8720D-01
NA+	5.4887D+00	5.0718D+00	9.2404D-01
CL-	5.4887D+00	5.0829D+00	9.2606D-01
CO3-2	6.8196D-05	1.5787D-06	2.3149D-02
OH-	2.7580D-07	1.4844D-07	5.3824D-01
HCO3-	4.8994D-03	1.8290D-03	3.7330D-01
H2CO3	9.4797D-05	2.6922D-04	2.8399D+00

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
---------	------------

CACL2	9.4601D-01
CAC03	1.5117D-01
CA(OH)2	6.5885D-01
NACL	9.2505D-01
NAHCO3	5.8732D-01
NA2CO3	2.7038D-01
NAOH	7.0523D-01
HCL	1.8590D+00

Test problem 6: Simulation of reaction path accompanying the evaporation of sea water

Two limiting cases have traditionally been used to examine reaction paths in geochemical simulation: the open model and the zoned or closed model. In the open model, products formed earlier may re-react with the aqueous phase and form new products as reaction proceeds. In the closed model, products once formed are isolated (zoned, layered, buried) from the system and not allowed to re-react. Test problem 5 of Parkhurst and others (1980) is an example of the logic used in open system simulation, as used in the code PHREEQE. In closed system simulation of reaction paths, the MINERALS input is adjusted so that products once formed are not allowed to re-react with the aqueous solution.

Usually when we simulate reaction paths, the irreversible reactant is an unstable mineral or a suite of unstable minerals; that is, the stoichiometry of the irreversible reaction is fixed. Evaporation poses a special problem in reaction path simulation because the stoichiometry of the irreversible reaction (defined by the aqueous solution composition) continually changes. In test problem 2, REACTION input was used to simulate evaporation of sea water by irreversible addition of "sea salt", that is, a hypothetical solid containing calcium, magnesium, sodium, potassium, chloride, sulfate and carbon in stoichiometric proportion to sea water. The approach used was valid as long as we did not wish to examine details of the reaction path to reach the final equilibrium invariant point because the composition of the reactant (evaporated sea water) changes with reaction progress. The reaction path during evaporation could be solved by changing the stoichiometry of the irreversible reactant (altered "sea salt") at each new phase boundary but this method would be extremely laborious.

As a means of simplifying simulation of evaporation (in open or closed systems) a new feature has been added in PHRQPITZ that allows the user to automatically define the composition of solution 1 as a salt, without the necessity of entering the salt composition under REACTION input. This does not eliminate the need to locate each new phase boundary encountered during evaporation, but it is not necessary to define the stoichiometry of each new reactant (altered sea water for example) as evaporation proceeds. To use this new feature of PHRQPITZ, one first defines the composition of the starting solution. Evaporation of this solution is accomplished by setting IOPT(3) = 6 (specifying that a reaction will be added until the first mineral of MINERALS input is just saturated) and defining the reaction under REACTION input as water (LREAC(1) = 3 and CREAC(1) = 1.0). PHRQPITZ recognizes this combination of input as the special case of evaporation (or dilution, if the starting solution is oversaturated with the first mineral of MINERALS input) and additional output is provided. After the print of mass transfer under the heading "PHASE BOUNDARIES" is printed, PHRQPITZ prints the evaporation factor necessary to reach saturation with the first mineral in MINERALS input. For example an evaporation factor of 2.0 would mean that the concentration of an inert component in the starting solution would double upon evaporation to reach the phase boundary. (Similar logic applies to dilution where the evaporation factor would be less than unity.) Following a print of the evaporation factor, PHRQPITZ gives the total moles of each element remaining referenced to the initial condition defined in solution 1 (the initial condition being moles per one kilogram of water) and the amount of water (kg) remaining. Following an evaporation step, the total moles of each element should not be confused with the molality of the element, both of which are printed.

In test problem 6 we consider several steps in the evaporation of sea water in a closed system. The problem of sea water evaporation is too complex to be considered fully here. For example purposes, the evaporation of sea water is taken to the halite phase boundary only. The problem is solved in three parts and the input file listed in table 16. The starting solution (sea water) is defined in part a and equilibrated with dolomite in contact with air. Table 17 shows that this equilibration step causes precipitation of dolomite and outgassing of CO₂. This modified starting solution is now undersaturated with all other minerals in the LOOK MIN data base. The modified sea water was saved in solution 1 at completion of part a for use as the starting point of evaporation in part b by setting IOPT(7) to 1 (table 16). In all subsequent evaporation simulations the final solution may be stored in solution 1 to eliminate the need to re-define the starting solution (IOPT(7) = 1). In parts b and c IOPT(3) = 6 indicating that a reaction will be run until a phase boundary (defined by the first mineral under MINERALS input) is reached. The evaporation reaction (or dilution) is defined as 1.0 H₂O under REACTION input and no STEPS input is needed. The first mineral under MINERALS input defines

Table 16. -- Line images of the input file to test problem 6

Test Problem 6 (part a): Equilibrate sea water with dolomite and air.

0050021000 0 0 0.0

ELEMENTS

C 15 61.0171

SOLUTION 1

SEAWATER FROM NORDSTROM ET AL. (1979) TEST CASE IN PPM.

8	15	3	8.22	8.451	25.0	1.023		
4	412.3		5	1291.8	6	10768.0	7	399.1
14	19353.0		15	141.682	16	2712.0	22	67.3

MINERALS

PCO2	1	4.0	-1.468	-4.776	1	-3.5
------	---	-----	--------	--------	---	------

35 1.0

108.3865	0.01985076	-6919.53	-40.45154	669365.0
----------	------------	----------	-----------	----------

DOLOMITE	3	8.00	-17.083	-9.436	0
----------	---	------	---------	--------	---

4	1.0	5	1.0	15	2.0
---	-----	---	-----	----	-----

END

#6 (part b): Evaporate final solution of part a to gypsum saturation.

0060021000 1 1 0.0

MINERALS

GYPSUM	3	6.00	-4.581	1
--------	---	------	--------	---

4	1.0	16	1.0	3	2.0
---	-----	----	-----	---	-----

90.318	0.0	-4213.	-32.641
--------	-----	--------	---------

PCO2	1	4.0	-1.468	-4.776	1	-3.5
------	---	-----	--------	--------	---	------

35 1.0

108.3865	0.01985076	-6919.53	-40.45154	669365.0
----------	------------	----------	-----------	----------

DOLOMITE	3	8.00	-17.083	-9.436	0
----------	---	------	---------	--------	---

4	1.0	5	1.0	15	2.0
---	-----	---	-----	----	-----

REACTION

3 1.0

END

#6 (part c): Evaporate final solution of part b to halite saturation.

0060021000 1 1 0.0

MINERALS

HALITE	2	0.00	1.570	1
--------	---	------	-------	---

6	1.000	14	1.000
---	-------	----	-------

-713.4616	-.1201241	37302.21	262.4583	-2106915.
-----------	-----------	----------	----------	-----------

GYPSUM	3	6.00	-4.581	1
--------	---	------	--------	---

4	1.0	16	1.0	3	2.0
---	-----	----	-----	---	-----

90.318	0.0	-4213.	-32.641
--------	-----	--------	---------

PCO2	1	4.0	-1.468	-4.776	1	-3.5
------	---	-----	--------	--------	---	------

35 1.0

108.3865	0.01985076	-6919.53	-40.45154	669365.0
----------	------------	----------	-----------	----------

DOLOMITE	3	8.00	-17.083	-9.436	0
----------	---	------	---------	--------	---

4	1.0	5	1.0	15	2.0
---	-----	---	-----	----	-----

REACTION

3 1.0

END

Table 17. -- Listing of printout from test problem 6

DATA READ FROM DISK

ELEMENTS
SPECIES
LOOK MIN
MEAN GAM

1Test Problem 6 (part a): Equilibrate sea water with dolomite and air.

0050021000 0 0 0.00000

ELEMENTS

C 15 0.61017E+02
0 0.00000E+00

SOLUTION 1

SEAWATER FROM NORDSTROM ET AL. (1979) TEST CASE IN PPM.

8 15 3 8.22 4.00 25.0 1.02
4 4.123D+02 5 1.292D+03 6 1.077D+04 7 3.991D+02 22 6.730D+01
14 1.935D+04 15 1.417D+02 16 2.712D+03

MINERALS

PCO2 1 4.0 -1.5 -4.8 1 -3.500
35 1.00
1.0839E+02 1.9851E-02 -6.9195E+03 -4.0452E+01 6.6936E+05
DOLOMITE 3 8.0 -17. -9.4 0 0.000
4 1.00 5 1.00 15 2.00
0 0.00 0.00 0.00 0 0.000

1SOLUTION NUMBER 1

SEAWATER FROM NORDSTROM ET AL. (1979) TEST CASE IN PPM.

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.066163D-02	-1.9722
MG	5.508555D-02	-1.2590
NA	4.854426D-01	-0.3139
K	1.057942D-02	-1.9755
CL	5.657614D-01	-0.2474
TOT ALK	2.406585D-03	-2.6186
S	2.926073D-02	-1.5337
BR	8.729404D-04	-3.0590

-----DESCRIPTION OF SOLUTION-----

PH = 8.2200
ACTIVITY H2O = 0.9813
OSMOTIC COEFFICIENT = 0.9035
IONIC STRENGTH = 0.7223
TEMPERATURE = 25.0000

PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6020D-05
 TOTAL ALKALINITY = 2.4066D-03
 ITERATIONS = 9
 TOTAL CARBON = 2.2098D-03

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	8.201E-09	-8.086	6.026E-09	-8.220	7.347E-01	-0.134
3	H2O	0.0	9.813E-01	-0.008	9.813E-01	-0.008	1.000E+00	0.000
4	CA+2	2.0	1.064E-02	-1.973	1.999E-03	-2.699	1.879E-01	-0.726
5	MG+2	2.0	5.499E-02	-1.260	1.130E-02	-1.947	2.054E-01	-0.687
6	NA+	1.0	4.854E-01	-0.314	3.096E-01	-0.509	6.378E-01	-0.195
7	K+	1.0	1.058E-02	-1.976	6.226E-03	-2.206	5.885E-01	-0.230
14	CL-	-1.0	5.658E-01	-0.247	3.900E-01	-0.409	6.893E-01	-0.162
15	CO3-2	-2.0	9.105E-05	-4.041	9.137E-06	-5.039	1.004E-01	-0.998
16	SO4-2	-2.0	2.926E-02	-1.534	3.100E-03	-2.509	1.060E-01	-0.975
22	BR-	-1.0	8.729E-04	-3.059	6.243E-04	-3.205	7.151E-01	-0.146
31	OH-	-1.0	2.841E-06	-5.547	1.636E-06	-5.786	5.759E-01	-0.240
34	HCO3-	-1.0	1.992E-03	-2.701	1.202E-03	-2.920	6.037E-01	-0.219
35	H2CO3	0.0	1.421E-05	-4.847	1.606E-05	-4.794	1.130E+00	0.053
40	HSO4-	-1.0	2.523E-09	-8.598	1.780E-09	-8.750	7.055E-01	-0.152
76	CACO3	0.0	2.588E-05	-4.587	2.588E-05	-4.587	1.000E+00	0.000
85	MGOH+	1.0	3.252E-06	-5.488	2.856E-06	-5.544	8.782E-01	-0.056
86	MGCO3	0.0	8.747E-05	-4.058	8.747E-05	-4.058	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	1.0724D-08	6.0256D-09	5.6187D-01
CA+2	1.0662D-02	1.9990D-03	1.8749D-01
MG+2	5.5086D-02	1.1297D-02	2.0508D-01
NA+	4.8544D-01	3.0960D-01	6.3777D-01
K+	1.0579D-02	6.2260D-03	5.8850D-01
CL-	5.6576D-01	3.8998D-01	6.8931D-01
CO3-2	2.0439D-04	9.1369D-06	4.4702D-02
SO4-2	2.9261D-02	3.1005D-03	1.0596D-01
BR-	8.7294D-04	6.2425D-04	7.1511D-01
OH-	6.0930D-06	1.6361D-06	2.6852D-01
HCO3-	1.9917D-03	1.2024D-03	6.0371D-01
H2CO3	1.4212D-05	1.6059D-05	1.1300D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	4.4662D-01

CASO4	1.4095D-01
CACO3	9.1550D-02
CA(OH)2	2.3822D-01
MGCL2	4.6017D-01
MGSO4	1.4741D-01
MGC03	9.5748D-02
MG(OH)2	2.4545D-01
NACL	6.6304D-01
NA2SO4	3.5061D-01
NAHCO3	6.2051D-01
NA2CO3	2.6296D-01
NAOH	4.1383D-01
KCL	6.3691D-01
K2SO4	3.3231D-01
KHCO3	5.9606D-01
K2CO3	2.4923D-01
KOH	3.9752D-01
HCL	6.2233D-01
H2SO4	3.2221D-01
HBR	6.3388D-01

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-5.2078	-4.3617	-0.8461
ARAGONIT	-7.7384	-8.2195	0.4811
ARCANITE	-6.9202	-1.7760	-5.1441
BISCHOFI	-2.8141	4.4551	-7.2692
BLOEDITE	-8.0154	-2.3470	-5.6684
BRUCITE	-13.5194	-10.8840	-2.6354
BURKEITE	-13.1115	-0.7720	-12.3395
CALCITE	-7.7384	-8.4062	0.6678
CARNALLI	-5.4289	4.3300	-9.7589
DOLOMITE	-14.7246	-17.0830	2.3584
EPSOMITE	-4.5130	-1.8809	-2.6321
GAYLUSSI	-13.8370	-9.4210	-4.4160
GLASERIT	-12.1437	-3.8030	-8.3407
GLAUBERI	-8.7347	-5.2450	-3.4897
GYPSUM	-5.2242	-4.5805	-0.6437
HALITE	-0.9182	1.5700	-2.4881
HEXAHYDR	-4.5048	-1.6346	-2.8702
KAINITE	-7.0949	-0.1930	-6.9019
KALICINI	-15.4650	-10.0580	-5.4070
KIESERIT	-4.4638	-0.1230	-4.3408
LABILE S	-12.2781	-5.6720	-6.6061
LEONHARD	-4.4884	-0.8870	-3.6014
LEONITE	-11.4085	-3.9790	-7.4295
MAGNESIT	-6.9862	-7.8340	0.8478
MIRABIL	-3.6089	-1.2135	-2.3954
MISENITE	-84.5263	-10.8060	-73.7203
NAHCOLIT	-13.7684	-10.7420	-3.0264
NATRON	-6.1396	-0.8250	-5.3146
NESQUEHO	-7.0108	-5.1670	-1.8438
PCO2	-4.7943	-1.4679	-3.3263

PENTAHYD	-4.4966	-1.2850	-3.2116
PIRSSONI	-13.8124	-9.2340	-4.5784
POLYHALI	-21.8077	-13.7440	-8.0637
PORTLAND	-14.2716	-5.1900	-9.0816
SCHOENIT	-11.4249	-4.3280	-7.0969
SYLVITE	-2.6147	0.8998	-3.5146
SYNGENIT	-12.1361	-7.4480	-4.6881
TRONA	-19.8424	-11.3840	-8.4584

1STEP NUMBER 1

0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.066163D-02	-1.9722
MG	5.508555D-02	-1.2590
NA	4.854426D-01	-0.3139
K	1.057942D-02	-1.9755
CL	5.657614D-01	-0.2474
TOT ALK	2.209828D-03	-2.6556
S	2.926073D-02	-1.5337
BR	8.729404D-04	-3.0590

----PHASE BOUNDARIES----

PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
PCO2	-7.853663D-04	-4.9679	-1.4679	-3.5000
DOLOMITE	-4.880043D-04	-17.0830	-17.0830	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

---- LOOK MIN IAP ----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-5.2262	-4.3617	-0.8645
ARAGONIT	-8.9263	-8.2195	-0.7069
ARCANITE	-6.9181	-1.7760	-5.1421
BISCHOFI	-2.8166	4.4551	-7.2716
BLOEDITE	-8.0143	-2.3470	-5.6673
BRUCITE	-14.5162	-10.8840	-3.6322
BURKEITE	-14.2751	-0.7720	-13.5031
CALCITE	-8.9263	-8.4062	-0.5201

CARNALLI	-5.4311	4.3300	-9.7611
DOLOMITE	-17.0830	-17.0830	0.0000
EPSOMITE	-4.5138	-1.8809	-2.6330
GAYLUSSI	-16.1923	-9.4210	-6.7713
GLASERIT	-12.1397	-3.8030	-8.3367
GLAUBERI	-8.7512	-5.2450	-3.5062
GYPSUM	-5.2426	-4.5805	-0.6621
HALITE	-0.9179	1.5700	-2.4879
HEXAHYDR	-4.5057	-1.6346	-2.8711
KAINITE	-7.0956	-0.1930	-6.9026
KALICINI	-16.1355	-10.0580	-6.0775
KIESERIT	-4.4647	-0.1230	-4.3417
LABILE S	-12.2925	-5.6720	-6.6205
LEONHARD	-4.4893	-0.8870	-3.6023
LEONITE	-11.4074	-3.9790	-7.4284
MAGNESIT	-8.1567	-7.8340	-0.3227
MIRABIL	-3.6068	-1.2135	-2.3933
MISENITE	-81.5306	-10.8060	-70.7246
NAHCOLIT	-14.4390	-10.7420	-3.6970
NATRON	-7.3069	-0.8250	-6.4819
NESQUEHO	-8.1812	-5.1670	-3.0142
PCO2	-4.9679	-1.4679	-3.5000
PENTAHYD	-4.4975	-1.2850	-3.2125
PIRSSONI	-16.1678	-9.2340	-6.9338
POLYHALI	-21.8435	-13.7440	-8.0995
PORTLAND	-15.2858	-5.1900	-10.0958
SCHOENIT	-11.4238	-4.3280	-7.0958
SYLVITE	-2.6145	0.8998	-3.5143
SYNGENIT	-12.1525	-7.4480	-4.7045
TRONA	-21.6804	-11.3840	-10.2964

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.017363D-02	-1.9925
MG	5.459755D-02	-1.2628
NA	4.854426D-01	-0.3139
K	1.057942D-02	-1.9755
CL	5.657614D-01	-0.2474
C	4.484536D-04	-3.3483
S	2.926073D-02	-1.5337
BR	8.729404D-04	-3.0590

-----DESCRIPTION OF SOLUTION-----

PH = 7.7226
ACTIVITY H2O = 0.9813
OSMOTIC COEFFICIENT = 0.9036

IONIC STRENGTH = 0.7196
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6020D-05
 TOTAL ALKALINITY = 4.5459D-04
 ITERATIONS = 7

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	2.575E-08	-7.589	1.894E-08	-7.723	7.355E-01	-0.133
3	H2O	0.0	9.813E-01	-0.008	9.813E-01	-0.008	1.000E+00	0.000
4	CA+2	2.0	1.017E-02	-1.993	1.911E-03	-2.719	1.879E-01	-0.726
5	MG+2	2.0	5.459E-02	-1.263	1.124E-02	-1.949	2.059E-01	-0.686
6	NA+	1.0	4.854E-01	-0.314	3.099E-01	-0.509	6.384E-01	-0.195
7	K+	1.0	1.058E-02	-1.976	6.233E-03	-2.205	5.891E-01	-0.230
14	CL-	-1.0	5.658E-01	-0.247	3.898E-01	-0.409	6.890E-01	-0.162
15	CO3-2	-2.0	6.151E-06	-5.211	6.201E-07	-6.208	1.008E-01	-0.997
16	SO4-2	-2.0	2.926E-02	-1.534	3.109E-03	-2.507	1.062E-01	-0.974
22	BR-	-1.0	8.729E-04	-3.059	6.239E-04	-3.205	7.147E-01	-0.146
31	OH-	-1.0	9.027E-07	-6.044	5.205E-07	-6.284	5.766E-01	-0.239
34	HCO3-	-1.0	4.252E-04	-3.371	2.565E-04	-3.591	6.032E-01	-0.220
35	H2CO3	0.0	9.531E-06	-5.021	1.077E-05	-4.968	1.130E+00	0.053
40	HSO4-	-1.0	7.955E-09	-8.099	5.609E-09	-8.251	7.050E-01	-0.152
76	CACO3	0.0	1.679E-06	-5.775	1.679E-06	-5.775	1.000E+00	0.000
85	MGOH+	1.0	1.027E-06	-5.988	9.044E-07	-6.044	8.802E-01	-0.055
86	MGCO3	0.0	5.908E-06	-5.229	5.908E-06	-5.229	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	3.3704D-08	1.8939D-08	5.6192D-01
CA+2	1.0174D-02	1.9109D-03	1.8783D-01
MG+2	5.4598D-02	1.1243D-02	2.0592D-01
NA+	4.8544D-01	3.0991D-01	6.3840D-01
K+	1.0579D-02	6.2326D-03	5.8913D-01
CL-	5.6576D-01	3.8978D-01	6.8895D-01
CO3-2	1.3738D-05	6.2007D-07	4.5137D-02
SO4-2	2.9261D-02	3.1085D-03	1.0624D-01
BR-	8.7294D-04	6.2389D-04	7.1470D-01
OH-	1.9302D-06	5.2055D-07	2.6968D-01
HCO3-	4.2519D-04	2.5648D-04	6.0322D-01
H2CO3	9.5307D-06	1.0766D-05	1.1296D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
---------	------------

CACL2	4.4673D-01
CASO4	1.4126D-01
CACO3	9.2076D-02
CA(OH)2	2.3905D-01
MGCL2	4.6064D-01
MGSO4	1.4791D-01
MGCO3	9.6409D-02
MG(OH)2	2.4649D-01
NACL	6.6320D-01
NA2SO4	3.5114D-01
NAHCO3	6.2056D-01
NA2CO3	2.6398D-01
NAOH	4.1493D-01
KCL	6.3709D-01
K2SO4	3.3283D-01
KHCO3	5.9613D-01
K2CO3	2.5022D-01
KOH	3.9859D-01
HCL	6.2221D-01
H2SO4	3.2251D-01
HBR	6.3373D-01

1#6 (part b): Evaporate final solution of part a to gypsum saturation.

0060021000 1 1 0.00000

MINERALS

GYPSUM	3	6.0	-4.6	0.00	1	0.000
4	1.00	16	1.00	3	2.00	
9.0318E+01	0.0000E-01	-4.2130E+03	-3.2641E+01	0.0000E-01		
PCO2	1	4.0	-1.5	-4.8	1	-3.500
35	1.00					
1.0839E+02	1.9851E-02	-6.9195E+03	-4.0452E+01	6.6936E+05		
DOLOMITE	3	8.0	-17.	-9.4	0	0.000
4	1.00	5	1.00	15	2.00	
0	0.00	0.00	0.00	0	0.00	0.000

REACTION

3 1.000 0.000

1STEP NUMBER 1

0-----

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.017363D-02	-1.9925
MG	5.459755D-02	-1.2628
NA	4.854426D-01	-0.3139
K	1.057942D-02	-1.9755
CL	5.657614D-01	-0.2474
C	4.484536D-04	-3.3483
S	2.926073D-02	-1.5337

BR

8.729404D-04

-3.0590

-----PHASE BOUNDARIES-----

MASS PRECIPITATED/DISSOLVED FROM INITIAL KILOGRAM WATER

	PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
	GYPSUM	0.000000D-01	-4.5805	-4.5805	0.0000
+	**				
	PCO2	-2.017398D-04	-4.9679	-1.4679	-3.5000
	DOLOMITE	-8.623117D-05	-17.0830	-17.0830	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

** 3.487807D+00 IS THE EVAPORATION FACTOR NECESSARY TO REACH THE
GYPSUM PHASE BOUNDARY.

MOLES OF ELEMENTS REMAINING AFTER REACTION

ELEMENT	MOLES	LOG MOLES
CA	1.008740D-02	-1.9962
MG	5.451131D-02	-1.2635
NA	4.854426D-01	-0.3139
K	1.057942D-02	-1.9755
CL	5.657614D-01	-0.2474
C	7.425143D-05	-4.1293
S	2.926073D-02	-1.5337
BR	8.729404D-04	-3.0590

2.867131D-01 KILOGRAMS OF WATER REMAINING

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.5181	-4.3617	-0.1564
ARAGONIT	-8.9579	-8.2195	-0.7385
ARCANITE	-5.8520	-1.7760	-4.0760
BISCHOFI	-1.1940	4.4551	-5.6491
BLOEDITE	-6.1270	-2.3470	-3.7800
BRUCITE	-14.5076	-10.8840	-3.6236
BURKEITE	-11.3906	-0.7720	-10.6186
CALCITE	-8.9579	-8.4062	-0.5517
CARNALLI	-2.7808	4.3300	-7.1108

DOLOMITE	-17.0830	-17.0830	0.0000
EPSOMITE	-3.9037	-1.8809	-2.0229
GAYLUSSI	-15.8708	-9.4210	-6.4498
GLASERIT	-9.9365	-3.8030	-6.1335
GLAUBERI	-6.8350	-5.2450	-1.5900
GYP SUM	-4.5805	-4.5805	0.0000
HALITE	0.1808	1.5700	-1.3892
HEXAHYDR	-3.8725	-1.6346	-2.2379
KAINITE	-5.3656	-0.1930	-5.1726
KALICINI	-15.9839	-10.0580	-5.9259
KIESERIT	-3.7164	-0.1230	-3.5934
LABILE S	-9.2143	-5.6720	-3.5423
LEONHARD	-3.8101	-0.8870	-2.9231
LEONITE	-9.6621	-3.9790	-5.6831
MAGNESIT	-8.1251	-7.8340	-0.2911
MIRABIL	-2.6291	-1.2135	-1.4156
MISENITE	-75.1160	-10.8060	-64.3100
NAHCOLIT	-14.2163	-10.7420	-3.4743
NATRON	-7.0690	-0.8250	-6.2440
NESQUEHO	-8.2187	-5.1670	-3.0517
PCO2	-4.9679	-1.4679	-3.5000
PENTAHYD	-3.8413	-1.2850	-2.5563
PIRSSONI	-15.7772	-9.2340	-6.5432
POLYHALI	-18.6357	-13.7440	-4.8917
PORTLAND	-15.3405	-5.1900	-10.1505
SCHOENIT	-9.7245	-4.3280	-5.3965
SYLVITE	-1.5868	0.8998	-2.4866
SYNGENIT	-10.4013	-7.4480	-2.9533
TRONA	-21.0356	-11.3840	-9.6516

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	3.518289D-02	-1.4537
MG	1.901250D-01	-0.7210
NA	1.693130D+00	0.2287
K	3.689898D-02	-1.4330
CL	1.973267D+00	0.2952
C	2.589747D-04	-3.5867
S	1.020558D-01	-0.9912
BR	3.044648D-03	-2.5165

-----DESCRIPTION OF SOLUTION-----

PH = 7.4400
ACTIVITY H2O = 0.9306
OSMOTIC COEFFICIENT = 0.9892
IONIC STRENGTH = 2.5080

TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6020D-05
 TOTAL ALKALINITY = 2.6794D-04
 ITERATIONS = 10

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	UNSCALED LOG ACT	UNSCALED GAMMA	UNSCALED LOG GAM
1	H+	1.0	3.532E-08	-7.452	3.630E-08	-7.440	1.028E+00	0.012
3	H2O	0.0	9.306E-01	-0.031	9.306E-01	-0.031	1.000E+00	0.000
4	CA+2	2.0	3.518E-02	-1.454	6.884E-03	-2.162	1.957E-01	-0.708
5	MG+2	2.0	1.901E-01	-0.721	4.685E-02	-1.329	2.464E-01	-0.608
6	NA+	1.0	1.693E+00	0.229	1.046E+00	0.020	6.178E-01	-0.209
7	K+	1.0	3.690E-02	-1.433	1.786E-02	-1.748	4.841E-01	-0.315
14	CL-	-1.0	1.973E+00	0.295	1.450E+00	0.161	7.347E-01	-0.134
15	CO3-2	-2.0	4.914E-06	-5.309	1.600E-07	-6.796	3.257E-02	-1.487
16	SO4-2	-2.0	1.021E-01	-0.991	4.406E-03	-2.356	4.318E-02	-1.365
22	BR-	-1.0	3.045E-03	-2.516	2.514E-03	-2.600	8.259E-01	-0.083
31	OH-	-1.0	5.637E-07	-6.249	2.575E-07	-6.589	4.569E-01	-0.340
34	HCO3-	-1.0	2.391E-04	-3.621	1.269E-04	-3.897	5.307E-01	-0.275
35	H2CO3	0.0	7.039E-06	-5.152	1.077E-05	-4.968	1.529E+00	0.185
40	HSO4-	-1.0	2.149E-08	-7.668	1.524E-08	-7.817	7.092E-01	-0.149
76	CACO3	0.0	1.561E-06	-5.807	1.561E-06	-5.807	1.000E+00	0.000
85	MGOH+	1.0	2.616E-06	-5.582	1.864E-06	-5.729	7.126E-01	-0.147
86	MGCO3	0.0	6.354E-06	-5.197	6.354E-06	-5.197	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	5.6808D-08	3.6304D-08	6.3907D-01
CA+2	3.5183D-02	6.8842D-03	1.9567D-01
MG+2	1.9012D-01	4.6852D-02	2.4643D-01
NA+	1.6931D+00	1.0459D+00	6.1775D-01
K+	3.6899D-02	1.7863D-02	4.8410D-01
CL-	1.9733D+00	1.4497D+00	7.3467D-01
CO3-2	1.2828D-05	1.6003D-07	1.2475D-02
SO4-2	1.0206D-01	4.4064D-03	4.3177D-02
BR-	3.0446D-03	2.5144D-03	8.2585D-01
OH-	3.1799D-06	2.5753D-07	8.0985D-02
HCO3-	2.3911D-04	1.2689D-04	5.3067D-01
H2CO3	7.0392D-06	1.0766D-05	1.5295D+00

----- MEAN ACTIVITY COEFFICIENT -----

FORMULA	MEAN GAMMA
CACL2	4.7268D-01

CASO4	9.1915D-02
CACO3	4.9406D-02
CA(OH)2	1.0867D-01
MGCL2	5.1045D-01
MGSO4	1.0315D-01
MGCO3	5.5445D-02
MG(OH)2	1.1735D-01
NACL	6.7368D-01
NA2SO4	2.5446D-01
NAHCO3	5.7256D-01
NA2CO3	1.6822D-01
NAOH	2.2367D-01
KCL	5.9637D-01
K2SO4	2.1629D-01
KHCO3	5.0685D-01
K2CO3	1.4299D-01
KOH	1.9800D-01
HCL	6.8520D-01
H2SO4	2.6028D-01
HBR	7.2648D-01

1#6 (part c): Evaporate final solution of part b to halite saturation.

0060021000 1 1 0.00000

MINERALS

HALITE	2	0.00	1.6	0.00	1	0.000
6	1.00	14	1.00			
		-7.1346E+02	-1.2012E-01	3.7302E+04	2.6246E+02	-2.1069E+06
GYPSUM	3	6.0	-4.6	0.00	1	0.000
4	1.00	16	1.00	3	2.00	
		9.0318E+01	0.0000E-01	-4.2130E+03	-3.2641E+01	0.0000E-01
PCO2	1	4.0	-1.5	-4.8	1	-3.500
35	1.00					
		1.0839E+02	1.9851E-02	-6.9195E+03	-4.0452E+01	6.6936E+05
DOLOMITE	3	8.0	-17.	-9.4	0	0.000
4	1.00	5	1.00	15	2.00	
	0	0.00	0.00	0.00	0	0.000

REACTION

3 1.000 0.000

'STEP NUMBER 1

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	3.518289D-02	-1.4537
MG	1.901250D-01	-0.7210
NA	1.693130D+00	0.2287
K	3.689898D-02	-1.4330
CL	1.973267D+00	0.2952

C	2.589747D-04	-3.5867
S	1.020558D-01	-0.9912
BR	3.044648D-03	-2.5165

-----PHASE BOUNDARIES-----

MASS PRECIPITATED/DISSOLVED FROM INITIAL KILOGRAM WATER

	PHASE	DELTA PHASE*	LOG IAP	LOG KT	LOG IAP/KT
	HALITE	0.000000D-01	1.5700	1.5700	0.0000
+	**				
	GYPSUM	-3.155443D-02	-4.5805	-4.5805	0.0000
	PCO2	-1.255065D-04	-4.9679	-1.4679	-3.5000
	DOLOMITE	-4.023108D-05	-17.0830	-17.0830	0.0000

* NEGATIVE DELTA PHASE INDICATES PRECIPITATION
AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.

** 3.057208D+00 IS THE EVAPORATION FACTOR NECESSARY TO REACH THE
HALITE PHASE BOUNDARY.

MOLES OF ELEMENTS REMAINING AFTER REACTION

ELEMENT	MOLES	LOG MOLES
CA	3.588229D-03	-2.4451
MG	1.900847D-01	-0.7211
NA	1.693130D+00	0.2287
K	3.689898D-02	-1.4330
CL	1.973267D+00	0.2952
C	5.300600D-05	-4.2757
S	7.050135D-02	-1.1518
BR	3.044648D-03	-2.5165

3.270959D-01 KILOGRAMS OF WATER REMAINING

----- LOOK MIN IAP -----

PHASE	LOG IAP	LOG KT	LOG IAP/KT
ANHYDRIT	-4.3204	-4.3617	0.0413
ARAGONIT	-9.5397	-8.2195	-1.3202
ARCANITE	-4.9697	-1.7760	-3.1937
BISCHOFI	1.0046	4.4551	-3.4504
BLOEDITE	-3.8133	-2.3470	-1.4663
BRUCITE	-14.0247	-10.8840	-3.1407

BURKEITE	-8.1263	-0.7720	-7.3543
CALCITE	-9.5397	-8.4062	-1.1334
CARNALLI	0.5743	4.3300	-3.7557
DOLOMITE	-17.0830	-17.0830	0.0000
EPSOMITE	-3.2344	-1.8809	-1.3536
GAYLUSSI	-16.3782	-9.4210	-6.9572
GLASERIT	-7.9391	-3.8030	-4.1361
GLAUBERI	-5.2894	-5.2450	-0.0444
GYPSUM	-4.5805	-4.5805	0.0000
HALITE	1.5700	1.5700	0.0000
HEXAHYDR	-3.1044	-1.6346	-1.4698
KAINITE	-3.1446	-0.1930	-2.9516
KALICINI	-15.9819	-10.0580	-5.9239
KIESERIT	-2.4541	-0.1230	-2.3311
LABILE S	-6.5185	-5.6720	-0.8465
LEONHARD	-2.8443	-0.8870	-1.9573
LEONITE	-7.8140	-3.9790	-3.8350
MAGNESIT	-7.5433	-7.8340	0.2907
MIRABIL	-2.2696	-1.2135	-1.0560
MISENITE	-69.5451	-10.8060	-58.7391
NAHCOLIT	-13.9815	-10.7420	-3.2395
NATRON	-7.4889	-0.8250	-6.6639
NESQUEHO	-7.9335	-5.1670	-2.7665
PCO2	-4.9679	-1.4679	-3.5000
PENTAHYD	-2.9743	-1.2850	-1.6893
PIRSSONI	-15.9881	-9.2340	-6.7541
POLYHALI	-16.1946	-13.7440	-2.4506
PORTLAND	-16.0211	-5.1900	-10.8311
SCHOENIT	-8.0741	-4.3280	-3.7461
SYLVITE	-0.4304	0.8998	-1.3302
SYNGENIT	-9.4201	-7.4480	-1.9721
TRONA	-20.4299	-11.3840	-9.0459

TOTAL MOLALITIES OF ELEMENTS

ELEMENT	MOLALITY	LOG MOLALITY
CA	1.096996D-02	-1.9598
MG	5.811285D-01	-0.2357
NA	5.176251D+00	0.7140
K	1.128079D-01	-0.9477
CL	6.032687D+00	0.7805
C	1.620503D-04	-3.7904
S	2.155373D-01	-0.6665
BR	9.308121D-03	-2.0311

-----DESCRIPTION OF SOLUTION-----

PH = 7.1144

ACTIVITY H2O = 0.7412
 OSMOTIC COEFFICIENT = 1.3694
 IONIC STRENGTH = 7.2808
 TEMPERATURE = 25.0000
 PRESSURE = 1.0000 ATM
 DENSITY OF H2O = 0.9971 G/CC
 ELECTRICAL BALANCE = -4.6020D-05
 TOTAL ALKALINITY = 2.3239D-04
 ITERATIONS = 14

 DISTRIBUTION OF SPECIES

I	SPECIES	Z	MOLALITY	LOG MOLAL	UNSCALED ACTIVITY	LOG ACT	UNSCALED GAMMA	LOG GAM
1	H+	1.0	1.815E-08	-7.741	7.684E-08	-7.114	4.234E+00	0.627
3	H2O	0.0	7.412E-01	-0.130	7.412E-01	-0.130	1.000E+00	0.000
4	CA+2	2.0	1.097E-02	-1.960	1.014E-02	-1.994	9.248E-01	-0.034
5	MG+2	2.0	5.811E-01	-0.236	1.006E+00	0.003	1.731E+00	0.238
6	NA+	1.0	5.176E+00	0.714	4.773E+00	0.679	9.221E-01	-0.035
7	K+	1.0	1.128E-01	-0.948	4.769E-02	-1.322	4.228E-01	-0.374
14	CL-	-1.0	6.033E+00	0.781	7.784E+00	0.891	1.290E+00	0.111
15	CO3-2	-2.0	2.102E-06	-5.677	2.845E-08	-7.546	1.354E-02	-1.869
16	SO4-2	-2.0	2.155E-01	-0.666	4.714E-03	-2.327	2.187E-02	-1.660
22	BR-	-1.0	9.308E-03	-2.031	1.695E-02	-1.771	1.821E+00	0.260
31	OH-	-1.0	2.328E-07	-6.633	9.690E-08	-7.014	4.163E-01	-0.381
34	HCO3-	-1.0	1.322E-04	-3.879	4.775E-05	-4.321	3.612E-01	-0.442
35	H2CO3	0.0	3.100E-06	-5.509	1.077E-05	-4.968	3.473E+00	0.541
40	HSO4-	-1.0	3.995E-08	-7.398	3.451E-08	-7.462	8.638E-01	-0.064
76	CACO3	0.0	4.089E-07	-6.388	4.089E-07	-6.388	1.000E+00	0.000
85	MGOH+	1.0	4.644E-05	-4.333	1.506E-05	-4.822	3.243E-01	-0.489
86	MGCO3	0.0	2.425E-05	-4.615	2.425E-05	-4.615	1.000E+00	0.000

SPECIES	TOTAL MOL	UNSCALED ACTIVITY	UNSCALED TOTAL GAMMA
H+	5.8098D-08	7.6843D-08	1.3226D+00
CA+2	1.0970D-02	1.0145D-02	9.2479D-01
MG+2	5.8113D-01	1.0060D+00	1.7311D+00
NA+	5.1763D+00	4.7732D+00	9.2214D-01
K+	1.1281D-01	4.7694D-02	4.2279D-01
CL-	6.0327D+00	7.7836D+00	1.2902D+00
CO3-2	2.6764D-05	2.8450D-08	1.0630D-03
SO4-2	2.1554D-01	4.7138D-03	2.1870D-02
BR-	9.3081D-03	1.6952D-02	1.8212D+00
OH-	4.6678D-05	9.6904D-08	2.0760D-03
HCO3-	1.3219D-04	4.7746D-05	3.6120D-01
H2CO3	3.1002D-06	1.0766D-05	3.4727D+00

---- MEAN ACTIVITY COEFFICIENT ----

FORMULA	MEAN GAMMA
CACL2	1.1547D+00
CASO4	1.4221D-01
CACO3	3.1353D-02
CA(OH)2	1.5855D-02
MGCL2	1.4231D+00
MGSO4	1.9457D-01
MGCO3	4.2897D-02
MG(OH)2	1.9540D-02
NACL	1.0908D+00
NA2SO4	2.6494D-01
NAHCO3	5.7713D-01
NA2CO3	9.6688D-02
NAOH	4.3754D-02
KCL	7.3858D-01
K2SO4	1.5753D-01
KHCO3	3.9079D-01
K2CO3	5.7490D-02
KOH	2.9626D-02
HCL	1.3063D+00
H2SO4	3.3696D-01
HBR	1.5520D+00

each new phase boundary to be reached in evaporation. Many more phase boundaries are encountered in the evaporation of sea water beyond halite saturation and are not considered here.

Table 17 (part b output) shows that an evaporation factor of 3.4878 is required to make sea water just saturated with gypsum (while maintaining equilibrium with dolomite in contact with air). Of the initial 1000 grams of water in the starting solution, 286.713 grams remain. Inspection of the LOOK MIN output shows that the modified sea water is saturated with gypsum and dolomite at a PCO_2 of $10^{-3.5}$ atm. and undersaturated with all other minerals in the Harvie and others (1984) data base. The modified sea water at completion of part b is stored in solution 1 (IOPT(7) = 1) to define the starting point in part c.

In part c we assume that the next phase encountered in evaporation will be halite. The starting solution is put back on a molal scale, that is moles per one kg of H_2O . An evaporation factor of 3.0572 is required to reach halite saturation. The cumulative evaporation factor over parts b and c is the product of both evaporation factors, that is, $3.4878 \times 3.0572 = 10.6629$ and the cumulative water remaining at the halite phase boundary is $1000/10.6629 = 93.78$ grams H_2O .

Our problem becomes more complicated at the end of part c (Table 17) where inspection of the final saturation indices shows that at the halite phase boundary the solution is oversaturated with anhydrite and magnesite. If our goal is to follow rigorously the thermodynamic path, it would be necessary to back up and locate at least the magnesite phase boundary and evaluate again whether anhydrite saturation would be reached before halite saturation. At the magnesite phase boundary we could treat the system open or closed to re-reaction of dolomite previously formed, by inclusion or exclusion of dolomite in MINERALS input, respectively. In an open system it would be necessary to locate the point in evaporation where all the dolomite previously formed had reacted to form magnesite. After locating the magnesite phase boundary the point where dolomite vanishes is found by placing dolomite first in the MINERALS input with magnesite equilibrium included. Beyond this point in evaporation the solution will be undersaturated with dolomite and the mineral would be removed from the MINERALS input while maintaining equilibrium with magnesite. If for kinetic reasons it is assumed that magnesite would not precipitate in this environment, MINERALS input for magnesite would not be included and subsequent computed supersaturation with respect to magnesite in solution would be disregarded. Seawater evaporation reaction paths could also be more realistically examined if, for example, formation of magnesian calcites were considered rather than dolomite which is well known for its irreversible behavior in low-temperature environments. Diagenetic reaction paths responsible for dolomitization in marine evaporite sequences can be considered in subsequent simulations.

PROGRAM SOURCE CODE

PHRQPITZ is written in FORTRAN 77 and currently runs on a Prime 850 mini-computer. The source code is listed in Attachment D along with the file COMMON.BLOCKS which is inserted in PHRQPITZ at the \$INSERT statements. A machine-readable listing of PHRQPITZ, associated data files and test cases is written on a 360K 5 1/4 inch floppy disk (diskette 1) located in the back pocket of this report. The contents of this disk can be viewed on an IBM-compatible PC by typing the command "list". Although PHRQPITZ does not currently run on a PC, the floppy disk can be used to transport the code to other machines. In transporting to other machines, it may be necessary to insert the file COMMON.BLOCKS (listed at the end of Attachment D and on the floppy disk) at each \$INSERT statement throughout the source code. Two data files are read internally by PHRQPITZ. These are PHRQPITZ.DATA (read from the main program as unit 11) and PITZER.DATA (read from SUBROUTINE INITPZ as unit 12). The file PHRQPITZ.DATA is listed in Attachment A and the file PITZER.DATA is listed in Attachment B. Both data files can be read from the floppy disk. The user is referred to Parkhurst and others (1980) for information on the construction and logic of PHREEQE. Comment statements in subroutines PITZER, BDK, PTEMP, and INITPZ define calculations based on the Pitzer model. References to equation numbers refer to equations in Harvie and Weare (1980), or to a lesser extent Harvie and others (1984). Subroutines BB, QQ, DFIND and Functions DW, BASE, PS, VLEST and DC are taken from Haar and others (1984) for calculation of the properties of water as a function of temperature and pressure.

INTERACTIVE CONSTRUCTION OF INPUT FILES FOR THE COMPUTER PROGRAM

PITZINPT is a FORTRAN 77 program that facilitates formulation of input data files to PHRQPITZ. It is nearly identical to the code PHRQINPT (Fleming and Plummer, 1983) which is used interactively to construct input sets to PHREEQE (Parkhurst and others, 1980). The source code to PITZINPT is listed in Attachment E. PITZINPT reads two data files (PHRQPITZ.DATA, Attachment A, and MINERALS.2.DATA, Attachment F). Machine-readable copies of PHRQPITZ and associated data files are written on a 360K 5 1/4 inch floppy disk located in the back pocket of this report (diskette 2). Type the command "list" to view the contents of this disk on an IBM-compatible PC.

PITZINPT, written in Prime Fortran 77, interactively asks the user at the terminal for values of variables required by PHRQPITZ (see "Description of Input"), explains the meaning and significance of each variable when required, and internally checks to make sure that values entered are valid. In some cases, PITZINPT automatically assigns values to certain variables, based on values previously entered, and keeps track of additional required information. PITZINPT contains an editor which allows the user to make corrections after each line has been completed.

The file MINERALS.2.DATA (Attachment F) contains pre-constructed MINERALS input for the Harvie and others (1984) data base from which the user may select from the screen for inclusion under MINERALS input. The list of pre-constructed minerals may be displayed by entering the command LIST during MINERALS input.

Attachment G contains a listing of an interactive session in which PITZINPT was used to construct the input file to test problem 3 (table 10). The user is first asked to enter the name of the file to be created. The user may next enter a (optional) reference input file which can be edited to produce the new file. If no such pre-constructed input file exists, a carriage return should be entered (as was done in the example of Attachment G) signifying that all new data will be entered from the screen. PITZINPT then prompts for entry of a title and the option line; lines 1 and 2 of all PHRQPITZ input files. Following completion of entry for each line, the line is displayed and the user is given an opportunity to edit what has been entered. Following acceptance of lines 1 and 2 the user is then prompted for entry of a keyword. Allowable entries are ELEMENTS, SPECIES, SOLUTION, MINERALS, LOOK

MIN, TEMP, STEPS, REACTION, NEUTRAL, SUMS, and END. Under each keyword the user is prompted for necessary information defined under "Description of Input". See Fleming and Plummer (1983) for additional details.

ACKNOWLEDGMENT

We benefited from discussions and correspondence with Charles E. Harvie and John H. Weare (University of California, San Diego) during code development. Eric C. Prestemon contributed computer calculations. The manuscript was improved by reviews from Eric J. Reardon (University of Waterloo), Ronald K. Stoessell (University of New Orleans) and Ronald J. Spencer (University of Calgary).

REFERENCES

- Ananthaswamy, J., and Atkinson, G., 1984, Thermodynamics of concentrated electrolyte mixtures. 4. Pitzer-Debye-Huckel Limiting slopes for water from 0 to 100 °C and from 1 atm to 1 kbar: *Journal of Chemical Engineering Data*, v. 29, p. 81-87.
- Bates, R.G., 1973, *Determination of pH, Theory and Practice* (2nd ed.): New York, John Wiley & Sons, 479 p.
- Bates, R.G., 1975, pH scales for sea water: in Goldberg, E.D. (ed), *The Nature of Seawater*, Dahlem Workshop Report, Berlin, Abakon Verlagsgesellschaft, p. 313-338.
- Bates, R.G., and Culberson, C.H., 1977, Hydrogen ions and the thermodynamic state of marine systems: in (Anderson, N.R., and Malahoff, A., eds), *The Fate of Fossil Fuel CO₂ in the Oceans*, p. 45-61, New York, Plenum Pub. Corp.
- Bodine, M.W., Jr., and Jones, B.F., 1986, THE SALT NORM: A quantitative chemical-mineralogical characterization of natural waters: U.S. Geological Survey, Water-Resources Investigations Report 86-4086, 130 p.
- Bradley, D.J., and Pitzer, K.S., 1979, Thermodynamics of electrolytes. 12. Dielectric properties of water and Debye-Huckel parameters to 350 °C and 1 kbar: *Journal Physical Chemistry*, v. 83, p. 1599-1603.
- Covington, A.K., Bates, R.G., and Durst, R.A., 1985, Definition of pH scales, standard reference values, measurement of pH and related terminology: *Pure & Applied Chemistry*, v. 57, p. 531-542.
- Dickson, A.G., 1984, pH scales and proton-transfer reactions in saline media such as sea water: *Geochimica Cosmochimica Acta*, v. 48, p. 2299-2308.
- Fleming, G.W., and Plummer, L.N., 1983, PHRQINPT - An interactive computer program for constructing input data sets to the geochemical simulation program PHREEQE: U.S. Geological Survey, Water-Resources Investigations Report 83-4236, 108 p.
- Frape, S.K., Fritz, P., and McNutt, R.H., 1984, Water-rock interaction and chemistry of groundwaters from the Canadian Shield: *Geochimica Cosmochimica Acta*, v. 48, p. 1617-1627.
- Haar, L., Gallagher, J.S., and Kell, G.S., 1984, NBS/NRC STEAM TABLES: Thermodynamic and transport properties and computer programs for vapor and liquid states of water in SI units: Washington, Hemisphere Pub. Corp., 320 p.
- Harned, H.S., and Bonner, F.T., 1945, The first ionization of carbonic acid in aqueous solutions of sodium chloride: *Journal American Chemical Society*, v. 67, p. 1026-1031.
- Harned, H.S., and Davis, R., Jr., 1943, The ionization constant of carbonic acid in water and the solubility of carbon dioxide in water and aqueous salt solutions from 0 to 50 °C: *Journal American Chemical Society*, v. 65, p. 2030-2037.

- Harvie, C.E., Moller, N., and Weare, J.H., 1984, The prediction of mineral solubilities in natural waters: The Na-K-Mg-Ca-H-Cl-SO₄-OH-HCO₃-CO₃-CO₂-H₂O system to high ionic strengths at 25 °C: *Geochimica Cosmochimica Acta*, v. 48, p. 723-751.
- Harvie, C.E., and Weare, J.H., 1980, The prediction of mineral solubilities in natural waters: the Na-K-Mg-Ca-Cl-SO₄-H₂O system from zero to high concentration at 25 °C: *Geochimica Cosmochimica Acta*, v. 44, p. 981-997.
- Hill, A.E., and Bacon, L.R., 1927, Ternary systems VI. Sodium carbonate, sodium bicarbonate and water: *Journal American Chemical Society*, v. 49, p. 2487-2495.
- Linke, W.F., 1965, Solubilities of inorganic and metal organic compounds: Washington, American Chemical Society, (4th ed.), v. 1, 1487 p., v. 2, 1914 p.
- MacInnes, D.A., 1919, The activities of the ions of strong electrolytes: *Journal American Chemical Society*, v. 41, 1086-1092.
- Maier, C.G., and Kelly, K.K., 1932, An equation for the representation of high temperature heat content data: *Journal American Chemical Society*, v. 54, p. 3243-3246.
- Millero F.J., 1979, The thermodynamics of the carbonate system in seawater: *Geochimica Cosmochimica Acta*, v. 43, p. 1651-1661.
- Millero, F.J., 1983, The estimation of the pK^*_{HA} of acids in seawater using the Pitzer equations: *Geochimica Cosmochimica Acta*, v. 47, p. 2121-2129.
- Millero, F.J., and Schreiber, D.R., 1982, Use of the ion pairing model to estimate activity coefficients of the ionic components of natural waters: *American Journal Science*, v. 282, p. 1508-1540.
- Moller, N., 1988, The prediction of mineral solubilities in natural waters: A chemical equilibrium model for the Na-Ca-Cl-SO₄-H₂O system, to high temperatures and concentration: *Geochimica Cosmochimica Acta*, v. 52, p. 821-837.
- Pabalan, R.T., and Pitzer, K.S., 1987, Thermodynamics of NaOH(aq) in hydrothermal solutions: *Geochimica Cosmochimica Acta*, v. 51, p. 829-837.
- Parkhurst, D.L., Thorstenson, D.C., and Plummer, L.N., 1980, PHREEQE - A computer program for geochemical calculations: *Water-Resources Investigations Report 80-96*, 210 p.
- Pitzer, K.S., 1973, Thermodynamics of electrolytes. 1. Theoretical basis and general equations: *Journal Physical Chemistry*, v. 77, p. 268-277.
- Pitzer, K.S., 1975, Thermodynamics of electrolytes. 5. Effects of higher-order electrostatic terms: *Journal Solution Chemistry*, v. 4, p. 249-265.
- Pitzer, K.S., 1979, Theory: Ion Interaction Approach: in R.M. Pytkowicz, (ed.), *Activity Coefficients in Electrolyte Solutions*, v. 1, CRC Press, Inc., Boca Raton, Florida, pp. 157-208.
- Pitzer, K.S., 1987, A thermodynamic model for aqueous solutions of liquid-like density. In *Reviews in Mineralogy*, v. 17, *Thermodynamic Modeling of Geological Materials: Minerals, Fluids and Melts*. I.S.E. Carmichael and H.P. Eugster, Eds. Mineralogical Society of America.
- Pitzer, K.S., and Kim, J.J., 1974, Thermodynamics of electrolytes. 4. Activity and osmotic coefficients for mixed electrolytes: *Journal American Chemical Society*, v. 96, p. 5701-5707.
- Pitzer, K.S., and Mayorga, G., 1973, Thermodynamics of electrolytes. 2. Activity and osmotic coefficients for strong electrolytes with one or both ions univalent: *Journal Physical Chemistry*, v. 77, p. 2300-2308.
- Pitzer, K.S., and Mayorga, G., 1974, Thermodynamics of electrolytes. 3. Activity and osmotic coefficients of 2-2 electrolytes: *Journal Solution Chemistry*, v. 3, p. 539-546.
- Pitzer, K.S., and Peiper, J.C., 1980, The activity coefficient of aqueous NaHCO₃: *Journal Physical Chemistry*, v. 84, p. 2396-2398.
- Pitzer, K.S., Peiper, J.C., and Busey, R.H., 1984, Thermodynamic properties of aqueous sodium chloride solutions: *Journal of Physical and Chemical Reference Data*, v. 13, p. 1-102.

- Pitzer, K.S., and Silvester, L.F., 1978, The thermodynamics of electrolytes. 11. Properties of 3:2, 4:2 and high valence types: *Journal Physical Chemistry*, v. 82, p. 1239-1242.
- Plummer, L.N., and Busenberg, E., 1982, The solubilities of calcite, aragonite and vaterite in $\text{CO}_2\text{-H}_2\text{O}$ solutions between 0 and 90 °C, and an evaluation of the aqueous model for the system $\text{CaCO}_3\text{-CO}_2\text{-H}_2\text{O}$: *Geochimica Cosmochimica Acta*, v. 46, 1011-1040.
- Plummer, L.N., and Sundquist, E.T., 1982, Total individual ion activity coefficients of calcium and carbonate in seawater at 25 °C and 35 ‰ salinity, and implications to the agreement between apparent and thermodynamic constants of calcite and aragonite: *Geochimica Cosmochimica Acta*, v. 46, p. 247-258.
- Reardon, E.J., and Armstrong, D.K., 1987, Celestite (SrSO_4) solubility in water, seawater and NaCl solution: *Geochimica Cosmochimica Acta*, v. 51, p. 63-72.
- Roy, R.N., Gibbons, J.J., Williams, R., Godwin, L., Baker, G., Simonson, J.M., and Pitzer, K.S., 1984, The thermodynamics of aqueous carbonate solutions. 2. Mixtures of potassium carbonate, bicarbonate, and chloride: *Journal Chemical Thermodynamics*, v. 16, p. 303-315.
- Sarbar, M., Covington, A.K., Nuttal, R.L., and Goldberg, R.N., 1982, The activity and osmotic coefficients of aqueous sodium bicarbonate solutions: *Journal Chemical Thermodynamics*, v. 14, p. 967-976.
- Silvester, L.F., and Pitzer K.S., 1977, Thermodynamics of electrolytes. 8. High-temperature properties, including enthalpy and heat capacity, with application to sodium chloride: *Journal Physical Chemistry*, v. 81, p. 1822-1828.
- Weres, O., Peiper, J.C., Pitzer, K.S., and Pabalan, R., 1987, Documentation for computer program NACL: LBL-21859, UC-11, 30p, Lawrence Berkeley Laboratory, University of California. Available from NTIS, U.S. Department of Commerce, 5285 Port Royal Road, Springfield, VA 22161.
- Westcott, C. Clark, 1978, pH Measurements: New York, Academic Press, 172 p.

Attachment A. -- Listing of the file PHRQPITZ.DATA

ELEMENTS

CA	4	40.08	CA+2
MG	5	24.305	MG+2
NA	6	22.9898	NA+
K	7	39.0983	K+
FE	8	55.847	FE+2
MN	9	54.9380	MN+2
BA	11	137.33	BA+2
SR	12	87.62	SR+2
CL	14	35.453	CL-
C	15	44.0098	CO2
S	16	96.06	SO4-2
LI	21	6.941	LI+
BR	22	79.904	BR-

SPECIES

1							
H+	100	1.0	0.0	9.0			0.0
0.0	0.0						
1 1.0							
2							
E-	100	-1.0	0.0	0.0			0.0
0.0	0.0						
2 1.0							
3							
H2O	100	0.0	0.0	0.0			0.0
0.0	0.0						
3 1.0							
4							
CA+2	101	2.0	0.0	6.0	5.0	0.165	0.0
0.0	0.0						
4 1.0							
5							
MG+2	101	2.0	0.0	8.0	5.5	0.20	0.0
0.0	0.0						
5 1.0							
6							
NA+	101	1.0	0.0	4.0	4.0	0.075	0.0
0.0	0.0						
6 1.0							
7							
K+	101	1.0	0.0	3.0	3.5	0.015	0.0
0.0	0.0						
7 1.0							
8							
FE+2	100	2.0	2.0	6.0			0.0
0.0	0.0						
8 1.0							
9							
MN+2	100	2.0	2.0	6.0			0.0
0.0	0.0						
9 1.0							
11							
BA+2	100	2.0	0.0	5.0			0.0

0.0	0.0						
11 1.0							
12							
SR+2	101 2.0	0.0	5.0		5.26	0.121	0.0
0.0	0.0						
12 1.0							
14							
CL-	101 -1.0	0.0	3.0	3.5		0.015	0.0
0.0	0.0						
14 1.0							
15							
C03-2	101 -2.0	4.0	4.5	5.4		0.0	2.0
0.0	0.0						
15 1.0							
16							
S04-2	101 -2.0	6.0	4.0	5.0		-0.04	0.0
0.0	0.0						
16 1.0							
21							
LI+	100 1.0	0.0	6.0				0.0
0.0	0.0						
21 1.0							
22							
BR-	100 -1.0	0.0	3.0				0.0
0.0	0.0						
22 1.0							
31							
OH-	200 -1.0	0.0	3.5				1.0
-13.998	13.345						
3 1.0	1 -1.0						
34							
HCO3-	211 -1.0	4.0	4.5	5.4		0.0	1.0
10.3393	-3.561	107.8975	0.03252849	-5151.79		-38.92561	563713.9
15 1.0	1 1.0						
35							
H2CO3	310 0.0	4.0	0.0				0.0
16.6767	-5.738	464.1925	0.09344813	-26986.16		-165.75951	2248628.9
15 1.0	1 2.0	3 -1.0					
40							
HSO4-	210 -1.0	6.0	0.0				0.0
1.979	4.91	-5.3585	0.0183412	557.2461			
16 1.0	1 1.0						
76							
CACO3	210 0.0	4.0	0.0				2.0
3.151	3.547	-1228.806	-0.299440	35512.75		485.818	
4 1.0	15 1.0						
85							
MGOH+	300 1.0	0.0	0.0				1.0
-11.809	15.419						
5 1.0	3 1.0	1 -1.0					
86							
MGCO3	210 0.0	4.0	0.0				2.0
2.928	2.535	-32.225	0.0	1093.486		12.72433	
5 1.0	15 1.0						

LOOK MIN

ANHYDRIT	2		6.00	-4.362		1		
4	1.000	16	1.000					
422.950	0.0		-18431.	-147.708				
ARAGONIT	2		4.00	-8.220		1		
15 1.0		4 1.0						
-171.8607	-.077993		2903.293	71.595				
ARCANITE	2		6.00	-1.776		1		
7	2.000	16	1.000					
2.823	0.0		-1371.2					
BISCHOFI	3		0.00	4.455		1		
5	1.000	14	2.000	3	6.000			
3.524	0.0		277.6					
BLOEDITE	4		12.00	-2.347		0		
6	2.000	5	1.000	16	2.000	3	4.000	
BRUCITE	2		0.00	-10.884.85		0		
5	1.000	31	2.000					
BURKEITE	3		16.00	-.772		0		
6	6.000	15	1.000	16	2.000			
CALCITE	2		4.00	-8.406		1		
15 1.0		4 1.0						
-171.8329	-.077993		2839.319	71.595				
CARNALLI	4		0.00	4.330		0		
7	1.000	5	1.000	14	3.000	3	6.000	
DOLOMITE	3		8.00	-17.083	-9.436	0		
4 1.0		5 1.0		15 2.0				
EPSOMITE	3		6.00	-1.881		1		
5	1.000	16	1.000	3	7.000			
1.718	0.0		-1073.					
GAYLUSSI	4		8.00	-9.421		0		
4	1.000	6	2.000	15	2.000	3	5.000	
GLASERIT	3		24.00	-3.803		0		
6	1.000	7	3.000	16	2.000			
GLAUBERI	3		12.00	-5.245		0		
6	2.000	4	1.000	16	2.000			
GYPSUM	3		6.00	-4.581		1		
4 1.0		16 1.0		3 2.0				
90.318	0.0		-4213.	-32.641				
HALITE	2		0.00	1.570		1		
6	1.000	14	1.000					
-713.4616	-.1201241		37302.21	262.4583	-2106915.			
HEXAHYDR	3		6.00	-1.635		1		
5	1.000	16	1.000	3	6.000			
-62.666	0.0		1828.	22.187				
KAINITE	5		6.00	-0.193		0		
7	1.000	5	1.000	14	1.000	16	1.000	3 3.000
KALICINI	3		4.00	-10.058		0		
7	1.000	1	1.000	15	1.000			
KIESERIT	3		6.00	-0.123		0		
5	1.000	16	1.000	3	1.000			
LABILE S	4		18.00	-5.672		0		
6	4.000	4	1.000	16	3.000	3	2.000	
LEONHARD	3		6.00	-0.887		0		
5	1.000	16	1.000	3	4.000			
LEONITE	4		12.00	-3.979		0		
7	2.000	5	1.000	16	2.000	3	4.000	
MAGNESIT	2		4.00	-7.834	-6.169	0		

5	1.000	15	1.000					
MIRABIL	3	6.	-1.214				1	
62.		161.		310.				
-3862.234	-1.19856		93713.54	1577.756		0.		
MISENITE	3	42.00	-10.806			0		
7	8.000	1	6.000	16	7.000			
NAHCOLIT	3	4.00	-10.742			0		
6	1.000	1	1.000	15	1.000			
NATRON	3	4.00	-0.825			0		
6	2.000	15	1.000	3	10.000			
NESQUEHO	3	4.00	-5.167			0		
5	1.000	15	1.000	3	3.000			
PCO2	1	4.0	-1.468		-4.776	1		
35	1.0							
108.3865	0.01985076	-6919.53		-40.45154		669365.0		
PENTAHYD	3	6.00	-1.285			0		
5	1.000	16	1.000	3	5.000			
PIRSSONI	4	8.00	-9.234			0		
6	2.000	4	1.000	15	2.000	3	2.000	
POLYHALI	5	24.00	-13.744			0		
7	2.000	5	1.000	4	2.000	16	4.000	3
PORTLAND	2	0.00	-5.190			0		
4	1.000	31	2.000					
SCHOENIT	4	12.00	-4.328			0		
7	2.000	5	1.000	16	2.000	3	6.000	
SYLVITE	2	0.00	.900			1		
7	1.000	14	1.000					
3.984	0.0		-919.55					
SYNGENIT	4	12.00	-7.448			0		
7	2.000	4	1.000	16	2.000	3	1.000	
TRONA	4	8.00	-11.384			0		
6	3.000	1	1.000	15	2.000	3	2.000	

MEAN GAM

CACL2	2	4	1.0	14	2.0
CASO4	2	4	1.0	16	1.0
CACO3	2	4	1.0	15	1.0
CA(OH)2	2	4	1.0	31	2.0
MGCL2	2	5	1.0	14	2.0
MGSO4	2	5	1.0	16	1.0
MGCO3	2	5	1.0	15	1.0
MG(OH)2	2	5	1.0	31	2.0
NACL	2	6	1.0	14	1.0
NA2SO4	2	6	2.0	16	1.0
NAHCO3	2	6	1.0	34	1.0
NA2CO3	2	6	2.0	15	1.0
NAOH	2	6	1.0	31	1.0
KCL	2	7	1.0	14	1.0
K2SO4	2	7	2.0	16	1.0
KHCO3	2	7	1.0	34	1.0
K2CO3	2	7	2.0	15	1.0
KOH	2	7	1.0	31	1.0
HCL	2	1	1.0	14	1.0
H2SO4	2	1	2.0	16	1.0
HBR	2	1	1.0	22	1.0
END					

Attachment B. -- Listing of the file PITZER.DATA

SPECIES 21

1 4 5 6 7 8 9 11 12 14 15 16 21 22 31 34 35 40 76 85 86

B0

NA+	CL-	0.0765	-777.03	-4.4706	0.008946	-3.3158E-6
K+	CL-	0.04835			5.794E-4	
MG+2	CL-	0.35235			-1.943E-4	
CA+2	CL-	0.3159			-1.725E-4	
MGOH+	CL-	-0.1				
H+	CL-	0.1775			-3.081E-4	
LI+	CL-	0.1494			-1.685E-4	
SR+2	CL-	0.28575			0.717E-3	
FE+2	CL-	0.335925				
MN+2	CL-	0.327225				
BA+2	CL-	0.2628			0.6405E-3	
NA+	BR-	0.0973			7.692E-4	
K+	BR-	0.0569			7.39E-4	
H+	BR-	0.1960			-2.049E-4	
MG+2	BR-	0.4327			-5.625E-5	
CA+2	BR-	0.3816			-5.2275E-4	
LI+	BR-	0.1748			-1.819E-4	
SR+2	BR-	0.331125			-0.32775E-3	
BA+2	BR-	0.31455			-0.33825E-3	
NA+	SO4-2	0.01958			2.367E-3	
K+	SO4-2	0.04995			1.44E-3	
MG+2	SO4-2	0.221			-0.69E-3	
CA+2	SO4-2	0.2				
H+	SO4-2	0.0298				
LI+	SO4-2	0.136275			0.5055E-3	
SR+2	SO4-2	0.220			-2.9E-3	
FE+2	SO4-2	0.2568				
MN+2	SO4-2	0.2065				
NA+	HSO4-	0.0454				
K+	HSO4-	-0.0003				
MG+2	HSO4-	0.4746				
CA+2	HSO4-	0.2145				
H+	HSO4-	0.2065				
FE+2	HSO4-	0.4273				
NA+	OH-	0.0864			7.00E-4	
K+	OH-	0.1298				
CA+2	OH-	-0.1747				
LI+	OH-	0.015				
BA+2	OH-	0.17175				
NA+	HCO3-	0.0277			1.00E-3	
K+	HCO3-	0.0296			0.996E-3	
MG+2	HCO3-	0.329				
CA+2	HCO3-	0.4				
SR+2	HCO3-	0.12				
NA+	CO3-2	0.0399			1.79E-3	
K+	CO3-2	0.1488			1.788E-3	

B1

NA+	CL-	0.2664			6.1608E-5	1.0715E-6
K+	CL-	0.2122			10.71E-4	
MG+2	CL-	1.6815			3.6525E-3	
CA+2	CL-	1.614			3.9E-3	

MGOH+	CL-	1.658			
H+	CL-	0.2945			1.419E-4
LI+	CL-	0.3074			5.366E-4
SR+2	CL-	1.66725			2.8425E-3
FE+2	CL-	1.53225			
MN+2	CL-	1.55025			
BA+2	CL-	1.49625			3.2325E-3
NA+	BR-	0.2791			10.79E-4
K+	BR-	0.2212			17.40E-4
H+	BR-	0.3564			4.467E-4
MG+2	BR-	1.753			3.8625E-3
CA+2	BR-	1.613			6.0375E-3
LI+	BR-	0.2547			6.636E-4
SR+2	BR-	1.7115			6.5325E-3
BA+2	BR-	1.56975			6.78E-3
NA+	SO4-2	1.113			5.6325E-3
K+	SO4-2	0.7793			6.6975E-3
MG+2	SO4-2	3.343			1.53E-2
CA+2	SO4-2	3.1973			5.46E-2
LI+	SO4-2	1.2705			1.41E-3
SR+2	SO4-2	2.88			27.0E-3
FE+2	SO4-2	3.063			
MN+2	SO4-2	2.9511			
NA+	HSO4-	0.398			
K+	HSO4-	0.1735			
MG+2	HSO4-	1.729			
CA+2	HSO4-	2.53			
H+	HSO4-	0.5556			
FE+2	HSO4-	3.48			
NA+	OH-	0.253			1.34E-4
K+	OH-	0.32			
CA+2	OH-	-0.2303			
LI+	OH-	0.14			
BA+2	OH-	1.2			
NA+	HCO3-	0.0411			1.10E-3
K+	HCO3-	-0.013			1.104E-3
MG+2	HCO3-	0.6072			
CA+2	HCO3-	2.977			
NA+	CO3-2	1.389			2.05E-3
K+	CO3-2	1.43			2.051E-3
B2					
MG+2	SO4-2	-37.23			-0.253
CA+2	SO4-2	-54.24			-0.516
SR+2	SO4-2	-41.8			-0.42
FE+2	SO4-2	-42.0			
MN+2	SO4-2	-40.0			
CA+2	OH-	-5.72			
C0					
NA+	CL-	0.00127	33.317	0.09421	-4.655E-5
K+	CL-	-0.00084			-5.095E-5
MG+2	CL-	0.00519			-1.64933E-4
CA+2	CL-	-0.00034			
H+	CL-	0.0008			6.213E-5
LI+	CL-	0.00359			-4.520E-5
SR+2	CL-	-0.00130461			
FE+2	CL-	-0.00860725			

MN+2	CL-	-0.0204972	
BA+2	CL-	-0.0193782	-1.53796E-4
NA+	BR-	0.00116	-9.30E-5
K+	BR-	-0.00180	-7.004E-5
H+	BR-	0.00827	-5.685E-5
MG+2	BR-	0.00312	
CA+2	BR-	-0.00257	
LI+	BR-	0.0053	-2.813E-5
SR+2	BR-	0.00122506	
BA+2	BR-	-0.0159576	
NA+	SO4-2	0.00497	-4.87904E-4
MG+2	SO4-2	0.025	0.523E-3
H+	SO4-2	0.0438	
LI+	SO4-2	-0.00399338	-2.33345E-4
SR+2	SO4-2	0.019	3.0E-3
FE+2	SO4-2	0.0209	
MN+2	SO4-2	0.01636	
NA+	OH-	0.0044	-18.94E-5
K+	OH-	0.0041	
K+	HCO3-	-0.008	
NA+	CO3-2	0.0044	
K+	CO3-2	-0.0015	
THETA			
K+	NA+	-0.012	
MG+2	NA+	0.07	
CA+2	NA+	0.07	
H+	NA+	0.036	
CA+2	K+	0.032	
H+	K+	0.005	
CA+2	MG+2	0.007	
H+	MG+2	0.1	
H+	CA+2	0.092	
SO4-2	CL-	0.02	
HSO4-	CL-	-0.006	
OH-	CL-	-0.05	
HCO3-	CL-	0.03	
CO3-2	CL-	-0.02	
OH-	BR-	-0.065	
OH-	SO4-2	-0.013	
HCO3-	SO4-2	0.01	
CO3-2	SO4-2	0.02	
CO3-2	OH-	0.1	
CO3-2	HCO3-	-0.04	
LAMDA			
NA+	H2CO3	0.1	
K+	H2CO3	0.051	
MG+2	H2CO3	0.183	
CA+2	H2CO3	0.183	
CL-	H2CO3	-0.005	
SO4-2	H2CO3	0.097	
HSO4-	H2CO3	-0.003	
PSI			
NA+	K+	CL-	-0.0018
NA+	K+	BR-	-0.0022
NA+	K+	SO4-2	-0.010
NA+	K+	HCO3-	-0.003

NA+	K+	CO3-2	0.003
NA+	CA+2	CL-	-0.007
NA+	CA+2	SO4-2	-0.055
NA+	MG+2	CL-	-0.012
NA+	MG+2	SO4-2	-0.015
NA+	H+	CL-	-0.004
NA+	H+	BR-	-0.012
NA+	H+	HSO4-	-0.0129
K+	CA+2	CL-	-0.025
K+	MG+2	CL-	-0.022
K+	MG+2	SO4-2	-0.048
K+	H+	CL-	-0.011
K+	H+	BR-	-0.021
K+	H+	SO4-2	0.197
K+	H+	HSO4-	-0.0265
CA+2	MG+2	CL-	-0.012
CA+2	MG+2	SO4-2	0.024
CA+2	H+	CL-	-0.015
MG+2	MGOH+	CL-	0.028
MG+2	H+	CL-	-0.011
MG+2	H+	HSO4-	-0.0178
CL-	BR-	K+	0.0000
CL-	SO4-2	NA+	0.0014
CL-	SO4-2	CA+2	-0.018
CL-	SO4-2	MG+2	-0.004
CL-	HSO4-	NA+	-0.006
CL-	HSO4-	H+	0.013
CL-	OH-	NA+	-0.006
CL-	OH-	K+	-0.006
CL-	OH-	CA+2	-0.025
CL-	HCO3-	NA+	-0.015
CL-	HCO3-	MG+2	-0.096
CL-	CO3-2	NA+	0.0085
CL-	CO3-2	K+	0.004
SO4-2	HSO4-	NA+	-0.0094
SO4-2	HSO4-	K+	-0.0677
SO4-2	HSO4-	MG+2	-0.0425
SO4-2	OH-	NA+	-0.009
SO4-2	OH-	K+	-0.050
SO4-2	HCO3-	NA+	-0.005
SO4-2	HCO3-	MG+2	-0.161
SO4-2	CO3-2	NA+	-0.005
SO4-2	CO3-2	K+	-0.009
OH-	CO3-2	NA+	-0.017
OH-	CO3-2	K+	-0.01
OH-	BR-	NA+	-0.018
OH-	BR-	K+	-0.014
HCO3-	CO3-2	NA+	0.002
HCO3-	CO3-2	K+	0.012

Attachment C. -- Summary of published Pitzer interaction parameters

**Attachment C.1. -- List of references cited in Attachments
C.2 through C.6**

Attachment C.1 - List of References Cited in Attachments C.2 - C.6 *

- 1) Ananthaswamy, J., and Atkinson, G., 1985, Thermodynamics of Concentrated Electrolyte Mixtures. 5. A Review of the Thermodynamic Properties of Aqueous Calcium Chloride in the Temperature Range 273.15-373.15 K: Journal of Chemical and Engineering Data, v.30, pp.120-128.
- 2) Andreu, R., Cejudo, J.A., and Marcos, E.S., 1985, Ionpit: A Full Implementation of Pitzer's Ion Interaction Treatment: Computers and Chemistry, v.9, no.3, pp.185-190.
- 3) Atlas, E., Culberson, C., and Pytkowicz, R.M., 1979, Phosphate Association With Na⁺, Ca²⁺ and Mg²⁺ in Seawater: Marine Chemistry, v.4, pp.243-254.
- 4) Bahia, A.M., Lilley, T.H., and Tasker, I.R., 1978, The Osmotic Coefficients of Aqueous CsCl and CsCl+KCl Mixtures at 298.15 K: Journal of Chemical Thermodynamics, v.10, pp.683-685.
- 5) Ball, F.X., Furst, W., and Renon, H., 1985, An NRTL Model for Representation and Prediction of Deviation from Ideality in Electrolyte Solutions Compared to the Models of Chen (1982) and Pitzer (1983): AIChE Journal, v.31, no.3, pp.392-399.
- 6) Boyd, G.E., 1977, Solute Activity Coefficients in Dilute Aqueous Electrolyte Mixtures.3. The Ternary System HClO₄+UO₂(ClO₄)₂+H₂O at 25 Degrees C: Journal of Solution Chemistry, v.6, no.11, pp.747-756.
- 7) Byrne, R.H.Jr., and Kester, D.R., 1974, Inorganic Speciation of Boron in Seawater: Journal of Marine Research, v.32, pp.119-127.
- 8) Chan, C.Y., and Khoo, K.H., 1979, Re-determination of Mean Ionic Activity Coefficients for the System HCl+KCl+Water at 298.15K and Correlations Between Harned and Pitzer Equations: Journal of the Chemical Society, Faraday Transactions 1, v.75, pp.1371-1379.
- 9) Chan, C.Y., Khoo, K.H., and Lim, T.K., 1979, Specific Ionic Interactions in the Quaternary Systems HCl-NaCl-KCl-Water and HCl-NH₄Cl-KCl-Water at 25 Degrees C: Journal of Solution Chemistry, v.8, no.1, pp.41-52.
- 10) Cohen, M.D., Flagan, R.C., and Seinfeld, J.H., 1987, Studies of Concentrated Electrolyte Solutions Using the Electrodynamic Balance. 1. Water Activities for Single-Electrolyte Solutions: Journal of Physical Chemistry, v. 91, no. 17, pp. 4563-4574.
- 11) de Lima, M.C.P., and Pitzer, K.S., 1983, Thermodynamics of Saturated Electrolyte Mixtures of NaCl with Na₂SO₄ and with MgCl: Journal of Solution Chemistry, v.12, pp.187-199.
- 12) Dickson, A.G., 1987, Standardization of the (AgCl + 1/2H₂ = Ag + HCl) Cell from 273.15 to 318.15 K: Journal of Chemical Thermodynamics, v.19, no.9, pp.993-1000.
- 13) Downes, C.J., 1975, Thermodynamics of Mixed Electrolyte Solutions- The System H₂O-NaCl-CoCl₂ and H₂O-CaCl₂-CoCl₂ at 25 Degrees C: Journal of Solution Chemistry, v.4, no.3, pp.191-204.
- 14) Downes, C.J., and Pitzer, K.S., 1976, Thermodynamics of Electrolytes. Binary Mixtures Formed from Aqueous NaCl, Na₂SO₄, CuCl₂, and CuSO₄ at 25 Degrees C: Journal of Solution Chemistry, v.5, no.6, pp.389-398.
- 15) Dyrssen, D., and Hansson, I., 1973, Ionic Medium Effects in Seawater- a Comparison of Acidity Constants of Carbonic Acid and Boric Acid in Sodium Chloride and Synthetic Seawater: Marine Chemistry, v.1, pp.137-149.
- 16) Edwards, T.J., Mauer, G., Newman, J., and Prausnitz, J.M., 1978, Vapor-Liquid-Equilibria in Multicomponent Aqueous Solutions of Volatile Weak Electrolytes: AIChE Journal- American Institute of Chemical Engineers, v.24, no.6, pp.966-976.

- 17) Elgquist, B., 1970, Determination of the Stability Constants of MgF^+ and CaF^+ Using a Fluoride Ion Selective Electrode: *Journal of Inorganic Nuclear Chemistry*, v.32, pp.937-944.
- 18) Felmy, A.R., and Weare, J.H., 1986, The Prediction of Borate Mineral Equilibria in Natural Waters- Application to Searles Lake, California: *Geochimica et Cosmochimica Acta*, v.50, pp.2771-2783.
- 19) Filippov, V.K., Anionova, V.A., Fohland, P., and Chajko, I.G., 1982, Thermodynamic Investigation of the System K_2SO_4 - $ZnSO_4$ - H_2O at 25 Degrees C: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*, no.4, pp.63-67.
- 20) Filippov, V.K., and Antonova, V.A., 1982, Thermodynamic Study of the Rb_2SO_4 - $CdSO_4$ - H_2O System at 25 Degrees C: *Journal of Applied Chemistry of the USSR*, v.55, no.6, pp.1157-1161.
- 21) Filippov, V.K., and Barkov, D.S., 1986, Thermodynamic Investigation of the System $Cu(NO_3)_2$ - $Zn(NO_3)_2$ - H_2O at 25 Degrees C: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*, no.2, pp.98-101.
- 22) Filippov, V.K., Barkov, D.S., and Fedorov, J.A., 1985, The Use of the Pitzer Equation for Calculating Solubility in the System $Cu(NO_3)_2$ - $Ni(NO_3)_2$ - H_2O at 25 Degrees C: *Zeitschrift fur Physikalische Chemie-Leipzig*, v.266, no.1, pp.129-134.
- 23) Filippov, V.K., Barkov, D.S., and Fedorov, Yu.A., 1986, Application of the Pitzer Equations to the Solubility of Ternary Aqueous Nitrate Solutions at 25 Degrees C: *Journal of Solution Chemistry*, v.15, no.7, pp.611-619.
- 24) Filippov, V.K., and Charykov, N.A., 1986, Phase Equilibria in the System $Na, Co||Cl, SO_4$ - H_2O at 25 Degrees C: *Journal of Applied Chemistry of the USSR*, v.59, no.11, pp.2255-2260.
- 25) Filippov, V.K., Charykov, N.A., Cheremnykh, L.M., Rumyantsev, A.V., 1986, Thermodynamic Calculation of the Phase Equilibria in the $Na, Mg||Cl, SO_4$ - H_2O System at 25 Degrees C: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*, no.3, pp.57-66.
- 26) Filippov, V.K., Charykov, N.A., and Fedorov, Y.A., 1986, $NaCl$ - $NiCl_2$ ($CuCl_2$)- H_2O Systems at 25 Degrees C: *Zhurnal Neorganicheskoi Khimii*, v.31, no.7, pp.1861-1866.
- 27) Filippov, V.K., Charykov, N.A., and Rumaintsev, A.V., 1987, Extension of the Pitzer Method to the Aqueous Salt Systems with Solution Phase Complexing: *Doklady Akademii Nauk SSSR*, v.296, no.3, pp.665-668.
- 28) Filippov, V.K., Charykov, N.A., and Solechnik, N.D., 1986, Thermodynamics of the System $Ni||Cl, SO_4$ - H_2O and $Co||Cl, SO_4$ - H_2O at 25 Degrees C: *Journal of Applied Chemistry of the USSR*, v.58, no.9, pp.1811-1814.
- 29) Filippov, V.K., Charykova, M.V., and Trofimov, Yu.M., 1986, Thermodynamics of the System $NH_4H_2PO_4$ -(NH_4) $2SO_4$ - H_2O at 25 Degrees C: *Journal of Applied Chemistry of the USSR*, v.58, no.9, pp.1807-1811.
- 30) Filippov, V.K., and Cheremnykh, L.M., 1983, Application of Pitzer Equation to the Solubility Calculation of Systems Na, Mg, Cl - H_2O and Na, Mg, SO_4 - H_2O at 25 Degrees C: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*, no.3, pp.32-37.
- 31) Filippov, V.K., and Cheremnykh, L.M., 1984, Calculation of Thermodynamic Functions of the System Na_2SO_4 - K_2SO_4 - H_2O at 25 Degrees: *Journal of Applied Chemistry of the USSR*, v.56, no.7, pp.1382-1385.
- 32) Filippov, V.K., and Cheremnykh, L.N., 1984, The Thermodynamic Study of the $Mg||Cl, SO_4$ - H_2O System at 25 Degrees C: *Ukrainskii Khimicheskii Zhurnal*, v.50, no.10, pp.1027-1032.
- 33) Filippov, V.K., and Cheremnykh, L.M., 1986, Application of Pitzer's Method to the Solubility Calculation in Na - Cl, SO_4 - H_2O Systems at 25 Degrees C: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*,

- no.1, pp.46-50.
- 34) Filippov, V.K., Dmitriev, G.V., and Iakovleva, S.I., 1980, Application of the Pitzer's Method for Calculation of Component Activities in Mixed-Solutions of Electrolytes Using Solubility Data: *Doklady Akademii Nauk SSSR*, v.252, no.1, pp.156-159.
 - 35) Filippov, V.K., Fedorov, Y.A., and Charykov, N.A., 1983, Solubility in Systems of $\text{NaCl-MnCl}_2\text{-H}_2\text{O}$ and $\text{NaCl-CoCl}_2\text{-H}_2\text{O}$ at 25 Degrees C: *Zhurnal Neorganicheskoi Khimii*, v.28, no.12, pp.3166-3170.
 - 36) Filippov, V.K., and Kalinkin, A.M., 1984, Calculation of Isotherms of the $\text{K}_2\text{SO}_4\text{-Cs}_2\text{SO}_4\text{-H}_2\text{O}$ Solubility System at 25 Degrees C: *Zhurnal Neorganicheskoi Khimii*, v.29, no.9, pp.2412-2415.
 - 37) Filippov, V.K., Kalinkin, A.M., and Vasin, S.K., 1987, Thermodynamics of Phase Equilibria of Aqueous (lithium sulfate + cesium sulfate), (sodium sulfate + cesium sulfate), and (potassium sulfate + cesium sulfate) at 298.15K Using Pitzer's Model: *Journal of Chemical Thermodynamics*, v.19, pp.185-193.
 - 38) Filippov, V.K., and Nokhrin, V.I., 1985, $\text{Li}_2\text{SO}_4\text{-MeSO}_4\text{-H}_2\text{O}$ (Me= Mn, Co, Cu) System at 25 Degrees C: *Zhurnal Neorganicheskoi Khimii*, v.30, no.2, pp.513-516.
 - 39) Filippov, V.K., and Nokhrin, V.I., 1985, Solubility in the $\text{NaSO}_4\text{-CuSO}_4\text{-H}_2\text{O}$ System at 25 Degrees C: *Zhurnal Neorganicheskoi Khimii*, v.30, no.11, pp.2963-2965.
 - 40) Filippov, V.K., Nokhrin, V.I., and Kryloya, I.L., 1985, Isopiestic Determination of Li, Mn|| $\text{SO}_4\text{-H}_2\text{O}$ Systems at 25 Degrees C: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*, no.3, pp.45-50.
 - 41) Filippov, V.K., Nokhrin, V.I., and Muzalevskaya, A.P., 1985, Thermodynamic Study of $\text{Na}_2\text{SO}_4\text{-ZnSO}_4\text{-H}_2\text{O}$ and $\text{Na}_2\text{SO}_4\text{-CdSO}_4\text{-H}_2\text{O}$ Systems at 25 Degrees C: *Zhurnal Neorganicheskoi Khimii*, v.30, no.9, pp.2407-2411.
 - 42) Filippov, V.K., Nokhrin, V.I., and Nikitin, S.N., 1984, Thermodynamic Study of the System $\text{Li}_2\text{SO}_4\text{-CdSO}_4\text{-H}_2\text{O}$ at 25 Degrees C: *Vestnik Leningradskogo, Seriya Fiziki i Khimii*, no.3, pp.36-41.
 - 43) Filippov, V.K., Veretennikova, G.A., and Nokhrin, V.I., 1985, Thermodynamic Study of the System $\text{Na}_2\text{SO}_4\text{-NiSO}_4\text{-H}_2\text{O}$ at 25 Degrees C: *Journal of Applied Chemistry of the USSR*, v.57, no.9, pp.1793-1796.
 - 44) Filippov, V.K., and Yakovleva, S.I., 1982, Chemistry and Thermodynamics of Solutions. Leningrad University, v.5, p.3.
 - 45) Filippov, V.K., Yakovleva, S.I., and Dmitriev, G.V., 1979, Calculation of Activity of $\text{Na}_2\text{SO}_4\text{-CoSO}_4\text{-H}_2\text{O}$ System Components By Pitzer's Method: *Vestnik Leningradskogo Universiteta, Seriya Fiziki i Khimii*, no.3, pp.58-63.
 - 46) Furst, W., and Renon, H., 1982, Effect of the Various Parameters in the Application of Pitzer's Model to Solid-Liquid Equilibrium. Preliminary Study for Strong 1-1 Electrolytes: *Industrial Engineering Chemical Process Design and Development*, v.21, pp.396-400.
 - 47) Gueddari, M., Monnin, C., Perret, D., Fritz, B., and Tardy, Y., 1983, Geochemistry of Brines of the Chott el Jerid in Southern Tunisia- Application of Pitzer's Equations: *Chemical Geology*, v.39, pp.165-178.
 - 48) Harvie, C.E., Moller, N., and Weare, J.H., 1984, The Prediction of Mineral Solubilities in Natural Waters: The $\text{Na-K-Mg-Ca-H-Cl-SO}_4\text{-OH-HCO}_3\text{-CO}_3\text{-CO}_2\text{-H}_2\text{O}$ System to High Ionic Strengths at 25 Degrees C: *Geochimica et Cosmochimica Acta*, v.48, pp.723-751.
 - 49) Harvie, C.E., and Weare, J.H., 1980, The Prediction of Mineral Solubilities in Natural Waters: The $\text{Na-K-Mg-Ca-Cl-SO}_4\text{-H}_2\text{O}$ System from Zero to High Concentration at 25 Degrees C: *Geochimica et Cosmochimica Acta*, v.44, pp.981-997.

- 50) Hershey, J.P., Fernandez, M., Milne, P.J., and Millero, F.J., 1986, The Ionization of Boric Acid in NaCl, Na-Ca-Cl and Na-Mg-Cl Solutions at 25 Degrees C: *Geochimica et Cosmochimica Acta*, v.50, no.1, pp.143-148.
- 51) Holmes, R.F., Baes, C.F.Jr., and Mesmer, R.E., 1979, Isopiestic Studies of Aqueous Solutions at Elevated Temperatures.2. NaCl+KCl Mixtures: *Journal of Chemical Thermodynamics*, v.11, pp.1035-1050.
- 52) Holmes, R.F., Baes, C.F.Jr., and Mesmer, R.E., 1981, Isopiestic Studies of Aqueous Solutions at Elevated Temperatures.3. $\{(1-y) \text{NaCl} + y\text{CaCl}_2\}$: *Journal of Chemical Thermodynamics*, v.13, pp.101-113.
- 53) Holmes, R.F., Busey, R.H., Simonson, J.M., Mesmer, R.E., Archer, D.G., and Wood, R.H., 1987, The Enthalpy of Dilution of HCl (aq) to 648K and 40 MPa, Thermodynamic Properties: *Journal of Chemical Thermodynamics*, v.19, pp.863-890.
- 54) Holmes, R.F., and Mesmer, R.E., 1983, Thermodynamic Properties of Aqueous-Solutions of the Alkali-Metal Chlorides to 250 Degrees C: *Journal of Physical Chemistry*, v.87, no.7, pp.1242-1255.
- 55) Holmes, R.F., and Mesmer, R.E., 1986, Isopiestic Studies of Aqueous Solutions at Elevated Temperatures.8. The Alkali-Metal Sulfates: *Journal of Chemical Thermodynamics*, v.18, pp.263-275.
- 56) Izaguirre, M., and Millero, F.J., 1987, The pK^* for the Dissociation of TRISH⁺ in NaClO₄ Media: *Journal of Solution Chemistry*, v.16, no.10, pp.827-834.
- 57) Johansson, O., and Wedborg, M., 1979, Stability Constants of Phosphoric Acid in Seawater of 5-40% Salinity and Temperatures of 5-25 Degrees C: *Marine Chemistry*, v.8, pp.57-69.
- 58) Khoo, K.H., Chan, C.Y., and Lim, T.K., 1977a, Activity Coefficients in Binary Electrolyte Mixtures HCl+MgCl₂+H₂O at 298.15K: *Journal of Solution Chemistry*, v.6, pp.855-864.
- 59) Khoo, K.H., Chan, C.Y., and Lim, T.K., 1977b, Thermodynamics of Electrolyte Solutions. The System HCl+CaCl₂+H₂O at 298.15 Degrees K: *Journal of Solution Chemistry*, v.6, no.10, pp.651-662.
- 60) Khoo, K.H., Chan, C.Y., and Lim, T.K., 1978, Activity Coefficients for the System HCl+BaCl₂+H₂O at 298.15K- Comparison of Scatchard's and Pitzer's Interpretations: *Journal of the Chemical Society, Faraday Transactions 1*, v.74, pp.837-845.
- 61) Khoo, K.H., Lim, T.K., and Chan, C.Y., 1978a, Activity Coefficients for HCl+NiCl₂+H₂O at 298.15 Degrees K and Effects of Higher-Order Electrostatic Terms: *Journal of Solution Chemistry*, v.7, no.4, pp.291-301.
- 62) Khoo, K.H., Lim, T.K., and Chan, C.Y., 1978b, Activity Coefficients for the System HCl+CoCl₂+H₂O at 298.15 Degrees K- Effects of Higher-Order Electrostatic Terms: *Journal of the Chemical Society, Faraday Transactions 1*, v.74, pp.2037-2044.
- 63) Khoo, K.H., Lim, T.K., and Chan, C.Y., 1979a, Activity Coefficients for the System HBr+CaBr₂+H₂O at 298.15K: *Journal of the Chemical Society, Faraday Transactions 1*, v.75, pp.1067-1072.
- 64) Khoo, K.H., Lim, T.K., and Chan, C.Y., 1979b, Ionic Interactions in the System HBr+BaBr₂+H₂O at 25 Degrees C: *Journal of Solution Chemistry*, v.8, no.4, pp.277-282.
- 65) Khoo, K.H., Lim, T.K., and Chan, C.Y., 1981, Activity Coefficients in Aqueous Mixtures of Hydrochloric Acid and Lanthinum Chloride at 25 Degrees C: *Journal of Solution Chemistry*, v.10, no.10, pp.683-691.
- 66) Khoo, K.H., Ramette, R.W., Culberson, C.H., and Bates, R.G., 1977, Determination of Hydrogen Ion Concentrations in Seawater from 5 to 40 Degrees C- Standard Potentials at Salinities from 20 to 450/oo: *Analytical Chemistry*, v.49, pp.29-33.
- 67) Kumar, A., 1987, Volume Properties of Aqueous Electrolytes.2. Application

- of the Pitzer Model in Estimating Apparent Molal Compressibility and Expansibility: *Journal of Chemical and Engineering Data*, v.32, pp.109-112.
- 68) Kuschel, F., and Seidel, J., 1985, Osmotic and Activity Coefficients of Aqueous K_2SO_4 - $MgSO_4$ and KCl - $MgCl_2$ at 25 Degrees C: *Journal of Chemical and Engineering Data*, v.30, pp.440-445.
 - 69) Lilley, T.H., and Tasker, I.R., 1983, The Osmotic Coefficients of Aqueous (lithium chloride + nickel chloride) and (tetramethylammonium chloride + nickel chloride) at 298.15K: *Journal of Chemical Thermodynamics*, v.15, pp.399-400.
 - 70) Lim, T.K., Khoo, K.H., and Chan, C.Y., 1980, Activity Coefficients for the System HBr - $SrBr_2$ - H_2O at 25 Degrees C: *Journal of Solution Chemistry*, v.9, no.10, pp.785-789.
 - 71) Macaskill, J.B., and Bates, R.G., 1978, Activity Coefficient of Hydrochloric Acid in the System HCl - KCl - H_2O at 25 Degrees C and Ionic Strengths from 0.1 to 3 Moles/Kg: *Journal of Solution Chemistry*, v.7, no.6, pp.433-442.
 - 72) Macaskill, J.B., White, D.R.Jr., Robinson, R.A., and Bates, R.G., 1978, Isopiestic Measurements on Aqueous Mixtures of Sodium Chloride and Strontium Chloride: *Journal of Solution Chemistry*, v.7, no.5, pp.339-347.
 - 73) Marcus, Y., 1977, The Activities of Potassium Chloride and of Water in Dead Sea Brine: *Geochimica et Cosmochimica Acta*, v.41, pp.1739-1744.
 - 74) Miller, D.J., Rard, J.A., Eppstein, L.B., and Robinson, R.A., 1980, Mutual Diffusion Coefficients, Electrical Conductances, Osmotic Coefficients, and Ionic Transport Coefficients i_{ij} for Aqueous $CaSO_4$ at 25 Degrees C: *Journal of Solution Chemistry*, v.9, no.7, pp.467-496.
 - 75) Millero, F.J., 1982, Use of Models to Determine Ionic Interactions in Natural Waters: *Thalassia Jugoslavica*, v.18(1-4), pp.253-291.
 - 76) Millero, F.J., and Thurmond, V., 1983, The Ionization of Carbonic Acid in Na-Mg-Cl Solutions at 25 Degrees C: *Journal of Solution Chemistry*, v.12, p.401.
 - 77) Moller, N., 1988, The Prediction of Mineral Solubilities in Natural Waters: A Chemical Equilibrium Model for the Na-Ca-Cl- SO_4 - H_2O System, to High Temperature and Concentration: *Geochimica et Cosmochimica Acta*, v.52, pp.821-837.
 - 78) Monnin, C., and Schott, J., 1984, Determination of the Solubility Products of Sodium Carbonate Minerals and an Application to Trona Deposition in Lake Magadi (Kenya): *Geochimica et Cosmochimica Acta*, v.48, pp.571-581.
 - 79) Pabalan, R.T., and Pitzer, K.S., 1987, Thermodynamics of Concentrated Electrolyte Mixtures and the Prediction of Mineral Solubilities to High Temperatures for Mixtures in the System Na-K-Mg-Cl- SO_4 -OH- H_2O : *Geochimica et Cosmochimica Acta*, v.51, no.9, pp.2429-2443.
 - 80) Pabalan, R.T., and Pitzer, K.S., 1987, Thermodynamics of $NaOH(aq)$ in Hydrothermal Solutions: *Geochimica et Cosmochimica Acta*, v.51, no.4, pp.829-837.
 - 81) Peiper, J.C., and Pitzer, K.S., 1982, Thermodynamics of Aqueous Carbonate Solutions Including Mixtures of Sodium Carbonate, Bicarbonate, and Chloride: *Journal of Chemical Thermodynamics*, v.14, pp.613-638.
 - 82) Phutela, R.C., and Pitzer, K.S., 1983, Thermodynamics of Aqueous Calcium Chloride: *Journal of Solution Chemistry*, v.12, no.3, pp.201-207.
 - 83) Pitzer, K.S., 1973, Thermodynamics of Electrolytes.1. Theoretical Basis and General Equations: *Journal of Physical Chemistry*, v.77, no.2, pp.268-277.
 - 84) Pitzer, K.S., 1975, Thermodynamics of Electrolytes.5. Effects of Higher-Order Electrostatic Terms: *Journal of Solution Chemistry*, v.4, no.3, pp.249-265.
 - 85) Pitzer, K.S., 1976, Theoretical Considerations of Solubility With Emphasis on Mixed Aqueous Electrolytes: *Pure and Applied Chemistry*, v.58, no.12,

pp.1599-1610.

- 86) Pitzer, K.S., 1979, Theory: Ion Interaction Approach, in R.M. Pytkowicz, (ed.), Activity Coefficients in Electrolyte Solutions, v.1, CRC Press, Inc., Boca Raton, Florida, pp.157-208.
- 87) Pitzer, K.S., 1981, Chemistry and Geochemistry of Solutions at High T and P: (eds.) Rickard and Wickman, 295, v.13-14, p.94.
- 88) Pitzer, K.S., and Kim, J.J., 1974, Thermodynamics of Electrolytes.4. Activity and Osmotic Coefficients for Mixed Electrolytes: Journal of the American Chemical Society, v.96, pp.5701-5707.
- 89) Pitzer, K.S., and Mayorga, G., 1973, Thermodynamics of Electrolytes.2. Activity and Osmotic Coefficients for Strong Electrolytes With One or Both Ions Univalent: Journal of Physical Chemistry, v.77, no.19, pp.2300-2308.
- 90) Pitzer, K.S., and Mayorga, G., 1974, Thermodynamics of Electrolytes.3. Activity and Osmotic Coefficients for 2-2 Electrolytes: Journal of Solution Chemistry, v.3, no.7, pp.539-546.
- 91) Pitzer, K.S., Olsen, J., Simonson, J.M., Roy, R.N., Gibbons, J.J., and Rowe, L., 1985, Thermodynamics of Aqueous Magnesium and Calcium Bicarbonates and Mixtures with Chloride: Journal of Chemical and Engineering Data, v.30, pp.14-17.
- 92) Pitzer, K.S., and Peiper, J.C., 1980, The Activity Coefficient of Aqueous NaHCO_3 : Journal of Physical Chemistry, v.84, pp.2396-2398.
- 93) Pitzer, K.S., Peiper, J.C., and Busey, R.H., 1984, Thermodynamic Properties of Aqueous Sodium Chloride Solutions: Journal of Physical and Chemical Reference Data, v.13, pp.1-102.
- 94) Pitzer, K.S., Roy, R.N., and Silvester, L.F., 1977, Thermodynamics of Electrolytes.7. Sulfuric Acid: Journal of the American Chemical Society, v.99, no.15, pp.4930-4936.
- 95) Pitzer, K.S., and Silvester, L.F., 1976, Thermodynamics of Electrolytes.6. Weak Electrolytes Including H_3PO_4 : Journal of Solution Chemistry, v.5, no.4, pp.269-278.
- 96) Pytkowicz, R.M., and Hawley, J.E., 1974, Bicarbonate and Carbonate Ion-Pairs and a Model of Seawater at 25 Degrees C: Limnology and Oceanography, v.19, pp.223-234.
- 97) Rard, J.A., and Miller, D.G., 1981a, Isopiestic Determination of Osmotic Coefficients of Aqueous Na_2SO_4 , MgSO_4 and Na_2SO_4 - MgSO_4 at 25 Degrees C: Journal of Chemical and Engineering Data, v.26, pp.33-38.
- 98) Rard, J.A., and Miller, D.G., 1981b, Isopiestic Determination of Osmotic and Activity Coefficients of Aqueous MgCl_2 Solutions at 25 Degrees C: Journal of Chemical and Engineering Data, v.26, pp.38-43.
- 99) Rard, J.A., and Miller, D.G., 1982a, Isopiestic Determination of the Osmotic and Activity Coefficients of Aqueous CsCl , SrCl_2 , and Mixtures of NaCl and CsCl at 25 Degrees C: Journal of Chemical and Engineering Data, v.27, no.2, pp.169-173.
- 100) Rard, J.A., and Miller, D.G., 1982b, Isopiestic Determination of the Osmotic and Activity Coefficients of Aqueous Mixtures of NaCl and SrCl_2 at 25 Degrees C: Journal of Chemical and Engineering Data, v.27, no.3, pp. 342-346.
- 101) Rard, J.A., and Miller, D.G., 1987, Isopiestic Determination of the Osmotic and Activity Coefficients of Aqueous Mixtures of NaCl and MgCl_2 at 25 Degrees C: Journal of Chemical and Engineering Data, v.32, pp.85-92.
- 102) Reardon, E.J., and Armstrong, D.K., 1987, Celestite ($\text{SrSO}_4(\text{s})$) Solubility in Water, Seawater and NaCl Solution: Geochimica et Cosmochimica Acta, v.51, no.1, pp.63-72.
- 103) Reardon, E.J., and Beekie, R.D., 1987, Modelling Chemical Equilibria of Acid Mine-Drainage- The FeSO_4 - H_2SO_4 - H_2O System: Geochimica et Cosmochimica Acta, v.51, no.9, pp.2355-2368.

- 104) Robinson, R.A., 1980, The Temperature Variation of the Harned Coefficient in the System HCl-NaCl: Journal of Solution Chemistry, v.9, no.7, pp. 449-454.
- 105) Robinson, R.A., and Macaskill, J.B., 1979, Osmotic Coefficients of Aqueous Sodium Carbonate Solutions at 25 Degrees C: Journal of Solution Chemistry, v.8, no.1, pp.35-40.
- 106) Rogers, P.S.Z., and Pitzer, K.S., 1981, High-Temperature Thermodynamic Properties of Aqueous Sodium Sulfate Solutions: The Journal Physical Chemistry, v.85, pp.2886-2895.
- 107) Rosenblatt, G.M., 1979, Estimation of Activity Coefficients in Concentrated Sulfite-Sulfate Solutions: Lawrence Berkeley Laboratory Report, LBL, 9671.
- 108) Rosenblatt, G.M., 1981, Estimation of Activity Coefficients In Concentrated Sulfite-Sulfate Solutions: AIChE Journal, v.27, no.4, pp.619-626.
- 109) Roy, R.N., Gibbons, J.J., Bliss, D.P.Jr., Casebolt, R.G., and Baker, B.K., 1980, Activity Coefficients for Ternary Systems.6. The System HCl+MgCl₂+H₂O at Different Temperatures, Application of Pitzer's Equations: Journal of Solution Chemistry, v.9, pp.911-930.
- 110) Roy, R.N., Gibbons, J.J., Ovens, L.K., Bliss, G.A., and Hartley, J.J., 1982, 7. Activity Coefficients for the Systems HCl+CaCl₂+H₂O at Various Temperatures: Journal of the Chemical Society, Faraday Transactions, v.78, pp.1405-1422.
- 111) Roy, R.N., Gibbons, J.J., Trower, J.K., and Lee, G.A., 1980, Application of Pitzer's Equations on the System HCl+MnCl₂+H₂O at Various Temperatures.5.: Journal of Solution Chemistry, v.9, no.7, pp.535-551.
- 112) Roy, R.N., Gibbons, J.J., Williams, R., Godwin, L., and Baker, G., Simonson, J.M., and Pitzer, K.S., 1984, The Thermodynamics of Aqueous Carbonate Solutions.2. Mixtures of Potassium Carbonate, Bicarbonate, and Chloride: Journal of Chemical Thermodynamics, v.16, pp.303-315.
- 113) Roy, R.N., Gibbons, J.J., Wood, M.D., Williams, R.W., Peiper, J.C., and Pitzer, K.S., 1983, The First Ionization of Carbonic Acid in Aqueous Solutions of Potassium Chloride Including the Activity Coefficients of Potassium Bicarbonate: Journal of Chemical Thermodynamics, v.15, pp.37-47.
- 114) Roy, R.N., Hufford, K., Lord, P.J., Mrad, D.R., Roy, L.N., and Johnson, D.A., 1988, The First Acidity Constant of Carbon Dioxide in Solutions of Ammonium Chloride from e.m.f. Measurements at 278.15 to 318.15 K: Journal of Chemical Thermodynamics, v.20, no.1, pp.63-77.
- 115) Ruaya, J.R., and Seward, T.M., 1987, The Ion-Pair Constant and Other Thermodynamic Properties of HCl up to 350 Degrees C: Geochimica et Cosmochimica Acta, v.51, pp.121-130.
- 116) Sabar, M., Covington, A.K., Nuttal, R.L., and Goldberg, R.N., 1982, The Activity and Osmotic Coefficients of Aqueous Sodium Bicarbonate Solutions: Journal of Chemical Thermodynamics, v.14, pp.967-976.
- 117) Sadowska, T., and Libus, W., 1982, Thermodynamic Properties and Solution Equilibria of Aqueous Bivalent Transition Metal Nitrates and Magnesium Nitrate: Journal of Solution Chemistry, v.11, no.7, pp.457-468.
- 118) Saluja, P.P.S., Pitzer, K.S., and Phutela, R.C., 1986, High-Temperature Thermodynamic Properties of Several 1-1 Electrolytes: Canadian Journal of Chemistry, v.64, no.7, pp.1328-1335.
- 119) Sankar, M., Macaskill, J.B., and Bates, R.G., 1981, Activity Coefficients of Hydrochloric Acid and Ionic Interactions in the System HCl-CsCl-H₂O from 5 to 50 Degrees C: Journal of Solution Chemistry, v.10, no.3, pp.169-179.
- 120) Silvester, L.F., and Pitzer, K.S., 1977, Thermodynamics of Electrolytes. 8. High-Temperature Properties, Including Enthalpy and Heat Capacity, with Application to Sodium Chloride: Journal of Physical Chemistry,

v.81, no.19, pp.1822-1828.

- 121) Silvester, L.F., and Pitzer, K.S., 1978, Thermodynamics of Electrolytes. 10. Enthalpy and the Effect of Temperature on the Activity Coefficients: *Journal of Solution Chemistry*, v.7, no.5, pp.327-337.
- 122) Simonson, J.M., Roy, R.N., Roy, L.N., and Johnson, D.A., 1987, The Thermodynamics of Aqueous Borate Solutions 1. Mixtures of Boric Acid with Sodium or Potassium Borate and Chloride: *Journal of Solution Chemistry*, v.16, no.10, pp.791-803.
- 123) Thiessen, D.B., and Wilson, A.J., 1987, An isopiestic Method for Measurement of Electrolyte Activity Coefficients: *AICHE Journal*, v.33, no.11, pp.1926-1929.
- 124) Thurmond, V., and Millero, F.J., 1982, Ionization of Carbonic Acid in Sodium Chloride Solutions at 25 Degrees C: *Journal of Solution Chemistry*, v.11, no.7, pp.447-456.
- 125) Tialowska-Mocharla, H., and Atkinson, G., 1985, Thermodynamics of Concentrated Electrolyte Mixture.6. Activity Coefficients of Aqueous CaCl_2 - CaBr_2 Mixtures at 25 Degrees C: *Journal of Physical Chemistry*, v.89, pp.4884-4887.
- 126) Torrent, J., Sanz, F., and Virgili, J., 1986, Activity Coefficients of Aqueous Perchloric Acid: *Journal of Solution Chemistry*, v.15, no.4, pp.363-375.
- 127) Weare, J.H., private communication (from Pitzer, K.S., Olsen, J., Simonson, J.M., Roy, R.N., Gibbons, J.J., and Rowe, L., 1985, Thermodynamics of Aqueous Magnesium and Calcium Bicarbonates and Mixtures with Chloride, *Journal of Chemical and Engineering Data*, v.30, pp.14-17.)
- 128) Weare, J.H., 1987, Models of Mineral Solubility in Concentrated Brines with Application to Field Observations: *Reviews in Mineralogy, Thermodynamic Modeling of Geological Materials: Minerals, Fluids, and Melts*, v.17, pp.143-176.
- 129) White, D.R., and Bates, R.G., 1980, Osmotic Coefficients and Activity Coefficients of Aqueous Mixtures of Sodium Chloride and Sodium Carbonate at 25 Degrees C: *Australian Journal of Chemistry*, v.33, pp.1903-1908.
- 130) White, D.R., Robinson, R.A., and Bates, R.G., 1980, Activity Coefficients of Hydrochloric Acid in HCl/MgCl_2 Mixtures and $\text{HCl}/\text{NaCl}/\text{MgCl}_2$ Mixtures from 5 to 45 Degrees C: *Journal of Solution Chemistry*, v.9, no.7, pp.457-465.
- 131) Whitfield, M., 1975, The Extension of Chemical Models for Seawater to Include Trace Components at 25 Degrees C and 1 atm. Pressure: *Geochimica et Cosmochimica Acta*, v.39, pp.1545-1557.

* Additional references are given in the text

**Attachment C.2. -- Summary of literature values of Pitzer interaction
parameters for single salts**

Attachment C.2 - Summary of Literature Values of Pitzer Interaction Parameters for Single Salts

INTERACTION	Na+	Cl-	BO	B1	CO	B2
			.0765 (48)	.2644 (48)	.00127 (48)	
			.07650 (49)	.2664 (49)	.00127 (49)	
			.0765 (89)	.2664 (89)	.00127 (89)	
			.07670 (83)	.26495 (83)	.00122 (83)	
			.0754 (93)	.2770 (93)	.0014 (93)	
			.0754 (85)	.2770 (85)	.0007 (85)	
			.07669 (129)	.26461 (129)	.0012193 (129)	
			.0757 (46)	.2795 (46)	.0014 (46)	
			.08305 (5)	.21900 (5)	----	
			.0765 (26)	.2665 (26)	.00062 (26)	
			.0765 (25)	.2664 (25)	.00121 (25)	
			.0765 (35)	.2664 (35)	.000126 (35)	
			.10820 (10)	.03127 (10)	.002469 (10)	
			.0765 (24)	.2664 (24)	.00124 (24)	
			.04835 (48)	.2122 (48)	----	
			.04835 (49)	.2122 (49)	.00084 (48)	
			.04835 (89)	.2122 (89)	.00084 (49)	
			.04827 (83)	.20887 (83)	.00080 (89)	
			.0481 (85)	.2187 (85)	.00082 (83)	
			.0464 (46)	.2274 (46)	.000394 (85)	
			.0453 (5)	.22839 (5)	.0004 (46)	
			.04808 (54)	.0476 (54)	----	
			.04626 (68)	.21844 (68)	.000788 (54)	
			.06577 (10)	.09351 (10)	.000398 (68)	
			.0504176 (123)	.195522 (123)	.002160 (10)	
			.1775 (48)	.2945 (48)	.001355442 (123)	
			.1775 (89)	.2945 (89)	.0008 (48)	
			.18352 (83)	.25503 (83)	.00080 (89)	
			.1775 (85)	.2945 (85)	.00059 (83)	
			.18029 (5)	.27837 (5)	.0004 (85)	
			.17527 (71)	.30197 (71)	----	
			.18104 (65)	.27779 (65)	.0014670 (71)	
			.3729 (61)	.3886 (61)	.00007307 (65)	
			.1494 (89)	.3074 (89)	.0008 (61)	
			.1485 (85)	.3070 (85)	.00359 (89)	
			.16832 (5)	.17721 (5)	.00186 (85)	
			.14847 (54)	.307 (54)	----	
			.16 (18)	----	.003710 (54)	
			.244 (50)	----	----	
			.12 (18)	----	----	
			.204 (50)	----	----	
			.28575 (89)	.1.66725 (89)	----	
			.2834 (85)	.1.626 (85)	.00130461 (89)	
			.27967 (5)	.1.72429 (5)	.000315 (85)	
			.28994 (72)	.1.5795 (72)	----	
			.27948 (72)	.1.6745 (72)	.003755 (72)	
			.28344 (99)	.1.625625 (99)	.0003532 (72)	
			.0300 (89)	.0558 (89)	.000890954 (99)	
			.03449 (83)	.01336 (83)	.00038 (89)	
			.0335 (85)	.0429 (85)	.00049 (83)	
			.03352 (54)	.0429 (54)	.000131 (85)	
			.03478 (99)	.03974 (99)	.000262 (54)	
			.03917 (99)	.002984 (99)	.000496 (99)	
			.03131 (4)	.065925 (4)	.001183 (99)	
			.03178 (5)	.02762 (5)	.00030705 (4)	
			.0522 (89)	.1918 (89)	----	
			.03603 (5)	.29619 (5)	.00301 (89)	
			.04568 (10)	.20431 (10)	----	
					.001731 (10)	

					.0 (49)	

Attachment C.2 - continued

INTERACTION	BO	B1	CO	B2
Cu+2	C1-	1.37625 (89)	-.0404323 (89)	
		1.4677 (26)	-.01395 (26)	.0023 (26)
		1.39125 (14)	.03602 (14)	
Cd+2	C1-	1.3468 (27)	-.00747 (27)	-5.829 (27)
Fe+2	C1-	1.33225 (89)	-.00860725 (89)	
		10.4224 (10)	-.010078 (10)	
Rb+	C1-	1.1483 (89)	-.00101 (89)	
		1.17280 (5)	-----	
UO+2	C1-	1.644 (89)	-.0368632 (89)	
Zn+2	C1-	1.6425 (89)	-.0879817 (89)	
Mn+2	C1-	1.55025 (89)	-.0204972 (89)	.00102 (35)
		1.4091 (35)	.0298 (35)	
Ba+2	C1-	2.31108 (10)	-.010540 (10)	
		1.49625 (89)	-.0193782 (89)	
		1.84591 (5)	-----	
Mg+2	C1-	1.9783 (60)	-.01938 (60)	
		1.6815 (48)	.00519 (48)	
		1.6815 (49)	.00519 (49)	.0 (49)
		1.6815 (89)	.00519193 (89)	
		1.651 (85)	.002301 (85)	
		1.21197 (5)	-----	
		1.65075 (98)	.006507 (98)	
		1.6508 (25)	.00610 (25)	
N1+2	C1-	-2.2231 (58)	.0051919 (58)	
		1.581 (89)	-.00372291 (89)	
		2.4076 (26)	-.00784 (26)	
		2.34470 (28)	.02156 (28)	
Co+2	C1-	-2.0903 (61)	.0037 (61)	
		1.47525 (89)	-.0152151 (89)	.5680 (26)
		1.625 (35)	.00179 (35)	-.53354 (28)
		2.06641 (28)	-.02429 (28)	
		1.44975 (13)	.012537 (13)	
		1.9504 (62)	-.01522 (62)	
		2.0664 (24)	.02430 (24)	
Ca+2	C1-	1.614 (48)	-.00034 (48)	
		1.6140 (49)	.00034 (49)	
		1.614 (89)	-.000339411 (89)	
		1.7083 (82)	.002153 (82)	
		1.708 (85)	.000761 (85)	
		1.64585 (5)	-----	
		1.53043 (10)	-.014250 (10)	
MgOH+	C1-	1.658 (48)	-----	
Al+3	C1-	5.8446667 (89)	-.00273279 (89)	
		5.63756 (5)	-----	
Cr+3	C1-	5.2553333 (89)	-.0451103 (89)	
Na+	B1-	6.14389 (5)	-----	
		-.2791 (89)	.00116 (89)	
		-.25274 (5)	-----	
		.02917 (10)	-.003478 (10)	
K+	B1-	.2212 (89)	-.00180 (89)	
		-.27445 (5)	-----	
		1.6046 (10)	-.001981 (10)	
H+	B1-	.3564 (89)	.00827 (89)	
		.23087 (5)	-----	
		.3975 (64)	.009687 (64)	
Mg+2	B1-	1.75275 (89)	.00312364 (89)	
		1.44995 (5)	-----	

[illegible]

Attachment C.2 - continued

INTERACTION	BO	B1	CO	B2
Co+2				
S04-2	.20 (90)	2.70 (90)	-.00926 (28,38)	-30.7 (90)
	.16310 (28,38)	3.3458 (28,38)	.03704 (38)	
	.1631 (38)	3.3460 (38)	.0182 (90)	
Mn+2	.2010 (90)	2.9511 (90)	.01636 (40)	-40.0 (40)
	.2065 (40)	2.9511 (40)	.0182 (86)	
	.201 (86)	2.980 (86)	-.014731 (10)	
	.33089 (10)	3.14630 (10)	-.0062 (90)	
Be+	.3170 (90)	2.914 (90)	.0176 (90)	
U02+2	.3220 (90)	1.827 (90)	-.019 (102)	
Sr+2	.220 (102)	2.88 (102)	.0209 (103)	-41.8+/- 9.2 (102)
S04-2	.2568 (103)	3.063 (103)	-.0059803 (89)	-42.0+/- 9.2 (103)
Fe+2	.0888 (89)	1.11075 (89)	-.00795 (55)	
CS+	.0882 (55)	1.105 (55)	.00291 (44)	
	.0714 (44)	1.2010 (44)	.00399338 (89)	
Li+	.136275 (89)	1.2705 (89)	-.000894 (85)	
S04-2	.1308 (85)	1.2913 (85)	.00333 (55)	
	.1339 (55)	1.282 (55)	-.000100762 (89)	
Rb+	.0579 (89)	1.11075 (89)	0.0 (20)	
	.05772 (20)	1.1105 (20)	-.000036 (34)	
NH4+	.0579 (34)	1.1108 (34)	-.00116142 (89)	
	.040875 (89)	.6585 (89)	-.00082 (29)	
	.03369 (5)	.74799 (5)	-.00116 (10)	
	.03905 (29)	.66379 (29)	-.001311 (10)	
	.04888 (10)	.6585 (10)	-.00118 (94)	
	.04763 (10)	.44459 (10)	-.00118 (94)	
Na+	.0454 (48)	.398 (48)	-.00082 (29)	
HS04-	.0554 (94)	.2755 (94)	-.00116 (10)	
K+	.0003 (48)	.1735 (48)	-.001311 (10)	
Mg+2	.4746 (48)	1.729 (48)	-.00082 (29)	
HS04-	.2145 (48)	2.53 (48)	-.00116 (10)	
Ca+2	.2065 (48)	.5556 (48)	-.001311 (10)	
H+	.2103 (94)	.4711 (94)	-.00082 (29)	
Fe+2	.4273 (103)	3.48 (103)	-.00116 (10)	
Na+	.0864 (48)	.253 (48)	-.00118 (94)	
OH-	.0864 (89)	.253 (89)	-.00082 (29)	
	.0864 (85)	.2530 (85)	-.00116 (10)	
	.10691 (5)	.17653 (5)	-.001311 (10)	
K+	.1298 (48)	.320 (48)	-.00082 (29)	
	.1298 (89)	.32 (89)	-.00116 (10)	
OH-	.14982 (5)	.22070 (5)	-.001311 (10)	
OH-	.1747 (48)	-.2303 (48)	-.00082 (29)	
Ca+2	.015 (89)	.14 (89)	-.00116 (10)	
Li+	.150 (89)	.30 (89)	-.001311 (10)	
CS+	-7.6 (15)	-.00082 (29)	-.00116 (10)	
OH-	.17175 (89)	1.2 (89)	-.00082 (29)	
OH-	.0277 (48)	.0411 (48)	-.00116 (10)	
OH-	.028 (81)	.044 (81)	-.001311 (10)	
Ba+2	-.04096 (116)	.5062 (116)	-.00082 (29)	
Na+	-.0490 (16)	-.00082 (29)	-.00116 (10)	
HC03-	.0277 (92)	.0411 (92)	-.00082 (29)	
	.0296 (48)	-.013 (48)	-.00116 (10)	
K+	-.0005 (92)	-.013 (92)	-.001311 (10)	
	-.022 (113)	.09 (113)	-.00082 (29)	
	-.0107 (112)	.0478 (112)	-.00116 (10)	
	-.011 (114)	.049 (114)	-.001311 (10)	

B2

-204.90 (76)
-46.0 (107)

-169-

Attachment C.2 - continued

INTERACTION		BO	B1	B2
Ca+2	F-	-2.3 (75)	----	1.9 (75)
		-2.3 (17)	----	1.92 (17)
		-2.3 (17)	----	.81 (17)
Cs+	F-	.1306 (89)	.2570 (89)	-.0043 (89)
		.1306 (85)	.2570 (85)	-.00215 (85)
Rb+	F-	.1141 (89)	.2842 (89)	-.0105 (89)
		-2.3 (75)	----	1.9 (75)
Na+2	F-	.0554 (89)	.2755 (89)	-.00118 (89)
		.040 (107)	.08 (107)	----
K+	C104-	.496125 (89)	2.0085 (89)	-.0095776 (89)
Mg+2	C104-	.451125 (89)	1.7565 (89)	-.005001 (89)
Ca+2	C104-	.1747 (89)	.2931 (89)	-.00819 (89)
H+	C104-	.1853 (126)	.2329 (126)	----
		.1998 (126)	.2695 (126)	----
		.1634 (6)	.3465 (6)	-.0117 (6)
		.21795 (5)	-.00041 (5)	----
		.1973 (89)	.3996 (89)	-.0008 (89)
		.0103 (89)	-.0194 (89)	----
Li+	C104-	.4269 (89)	1.56675 (89)	-.0131097 (89)
NH4+	C104-	.361425 (89)	1.57575 (89)	-.0312576 (89)
Sr+2	C104-	.506025 (89)	1.797 (89)	-.0113172 (89)
Ba+2	C104-	.333225 (89)	1.722 (89)	-.0088406 (89)
Zn+2	C104-	.27386 (5)	2.71678 (5)	----
Pb+2	C104-	.611325 (89)	2.14425 (89)	-.0216851 (89)
		.61785 (6)	2.06175 (6)	-.0153795 (6)
UO2+2	C104-	.0249 (89)	.2455 (89)	-.0004 (89)
Na+2	C103-	.0960 (89)	.2481 (89)	----
K+	C103-	.490 (107)	1.804 (107)	----
Mg+2	C103-	.438 (107)	1.76 (107)	----
Ca+2	C103-	.0068 (89)	.1783 (89)	-.00072 (89)
Na+2	C103-	.00661 (83)	.1739 (83)	-.00067 (83)
		.003 (46)	.2037 (46)	-.00013 (46)
		.00351 (5)	.18642 (5)	----
		.08155 (83)	.04939 (83)	-.00660 (89)
K+	N03-	.0802 (46)	.0722 (46)	-.0065 (46)
		.05771 (5)	-.05759 (5)	----
Mg+2	N03-	.367125 (89)	1.58475 (89)	-.0206245 (89)
		.29978 (23)	1.19260 (23)	-.00181 (23)
Ca+2	N03-	.3628425 (117)	1.5944175 (117)	-.0166947 (117)
		.210825 (89)	1.40925 (89)	-.0201419 (89)
H+	N03-	.12920 (23)	.79437 (23)	-.00263 (23)
		.1119 (89)	.3206 (89)	-.0010 (89)
		.11477 (5)	.30437 (5)	----
Li+	N03-	.1420 (89)	.2780 (89)	-.00551 (89)
		.12274 (5)	.36821 (5)	----
		.1421 (2)	.2772 (2)	----
Rb+	N03-	.0789 (89)	-.0172 (89)	-.0055 (2)
		.07885 (83)	-.01736 (83)	-.00529 (89)
Cs+	N03-	.0758 (89)	-.0669 (89)	-.00528 (83)
Ag+	N03-	.0856 (89)	.0025 (89)	----
		.05431 (5)	-.21042 (5)	-.00591 (89)
NH4+	N03-	.0154 (89)	.1120 (89)	----
		.01561 (5)	.10898 (5)	-.00003 (89)
Cu+2	N03-	.3168 (89)	1.43025 (89)	-.0219344 (89)
		.28076 (21)	1.73230 (21)	-.008372 (21)
		.34782 (117)	1.369185 (117)	-.0300697 (117)
Sr+2	N03-	.134625 (89)	1.38 (89)	-.0199245 (89)
Ba+2	N03-	-.03225 (89)	.8025 (89)	----

Attachment C.2 - continued

INTERACTION	B0	B1	CO	B2
Co+2 N03-	.311925 (89)	1.6905 (89)	-.00761554 (89)	
	.356025 (117)	1.577175 (117)	-.0195797 (117)	
Zn+2 N03-	.348075 (89)	1.69125 (89)	-.0156712 (89)	
	.3015 (21)	1.2012 (21)	-.005798 (21)	.26536 (21)
Co+2 N03-	.36915 (117)	1.4663475 (117)	-.0241247 (117)	
	.2865 (89)	1.668 (89)	-.0256467 (89)	
Pb+2 N03-	-.03615 (89)	.285 (89)	.00532881 (89)	
UO2+2 N03-	.460725 (89)	1.61325 (89)	-.031544 (89)	
Cr+5 N03-	.704 (89)	5.1846667 (89)	-.0590051 (89)	
Mn+2 N03-	.337005 (117)	1.54236 (117)	-.0155439 (117)	
Ni+2 N03-	.28790 (22)	1.43724 (22)	-.00189 (22)	.30331 (22)
Na+ N02-	.353475 (117)	1.60485 (117)	-.0194525 (117)	
	.0641 (89)	.1015 (89)	-.0049 (89)	
	.051 (46)	.1625 (46)	-.0023 (46)	
K+ N02-	.0151 (89)	.015 (89)	.0007 (89)	
Li+ N02-	.1336 (89)	.325 (89)	-.0053 (89)	
Rb+ N02-	.0269 (89)	-.1553 (89)	-.00366 (89)	
Cs+ N02-	.0427 (89)	.060 (89)	-.0051 (89)	
Na+ H2PO4-	-.0533 (89)	.0396 (89)	.00795 (89)	
K+ H2PO4-	-.0678 (89)	-.1042 (89)	-----	
Mg+2 H2PO4-	-.41 (3)	-----	-----	
Ca+2 H2PO4-	-.29 (3)	-----	-----	
NH4+ H2PO4-	-.07043 (29)	-.41564 (29)	.00669 (29)	
Na+ HP04-2	-.058275 (89)	1.4655 (89)	.0293802 (89)	
K+ HP04-2	.02475 (89)	1.27425 (89)	.0163872 (89)	
Mg+2 HP04-2	-3.6 (3)	-----	-----	
	-4.1 (57)	-----	-----	
Ca+2 HP04-2	-1.6 (3)	-----	-----	
	-2.0 (57)	-----	-----	
Na+ P04-3	.1781333 (89)	3.8513333 (89)	-.0515381 (89)	
K+ P04-3	.3729333 (89)	3.972 (89)	-.0867949 (89)	
Mg+2 P04-3	-62.0 (3)	-----	-----	
	-43.0 (57)	-----	-----	
Ca+2 P04-3	-67.0 (3)	-----	-----	
	-56.0 (57)	-----	-----	
Na+ H2AsO4-	-.0442 (89)	.2895 (89)	-----	
K+ H2AsO4-	-.0584 (89)	.0626 (89)	-----	
Na+ HAsO4-2	.030525 (89)	1.62975 (89)	-.00180312 (89)	
K+ HAsO4-2	.1286 (89)	1.6485 (89)	-.017819 (89)	
Na+ AsO4-3	.2388 (89)	3.93 (89)	-.0477276 (89)	
K+ AsO4-3	.4994 (89)	4.3406667 (89)	.1299423 (89)	
Na+ B(OH)4-	-.0526 (50)	.1104 (50)	.0154 (50)	
	-.0427 (18)	.089 (18)	.0114 (18)	
K+ B(OH)4-	.035 (18)	.14 (18)	-----	
	.20 (75)	-----	-----	
	.19 (7)	-----	-----	
Mg+2 B(OH)4-	-.21 (50)	-4.98 (50)	-.36 (50)	
	-.93 (7)	-----	-----	
Ca+2 B(OH)4-	-1.57 (50)	-4.49 (50)	-.17 (50)	
	-2.35 (7)	-----	-----	
	-2.35 (75)	-----	-----	
Sr+2 B(OH)4-	-2.35 (75)	-----	-----	
Na+ B3O3(OH)4	-.056 (18)	-.910 (18)	-----	
K+ B3O3(OH)4	-.13 (18)	-----	-----	
Na+ B4O5(OH)4	-.11 (18)	-.40 (18)	-----	
K+ B4O5(OH)4	-.022 (18)	-----	-----	
Na+ BrO3-	-.0205 (89)	.1910 (89)	.0059 (89)	
K+ BrO3-	-.1290 (89)	.2565 (89)	-----	
Na+ BrO2-	-.0526 (86)	.1104 (86)	.0154 (86)	

Attachment C.2 - continued

INTERACTION		B0	B1	CO	B2
Na+	C+04-2	.09375 (89)	1.3695 (89)	-.00215844 (89)	
K+	C+04-2	.075825 (89)	1.239 (89)	-.000779585 (89)	
Na+	HS03-	.0249 (107)	.2455 (107)	.0004 (107)	
K+	HS03-	.096 (107)	.2481 (107)	----	
Mg+2	HS03-	.490 (107)	1.804 (107)	----	
Ca+2	HS03-	.438 (107)	1.76 (107)	----	
Na+	SO3-2	.021 (108)	1.0 (108)	----	
K+	SO3-2	.065 (108)	3.00 (107)	----	
Mg+2	SO3-2	.200 (107)	3.60 (108)	----	
Ca+2	SO3-2	.24 (108)	2.38 (107)	----	-41.0 (107)
		.180 (107)	2.38 (108)	----	-33.0 (108)
		.18 (108)	1.27575 (89)	-.00373882 (89)	-61.3 (107)
Na+	S203-2	.06615 (89)	.9 (107)	----	-40.0 (108)
Na+	S205-2	.07 (107)	1.1 (107)	----	
K+	S205-2	.10 (107)	3.7 (107)	----	
Mg+2	S205-2	.24 (107)	3.0 (107)	----	-34.0 (107)
Ca+2	S205-2	.22 (107)		----	-50.0 (107)

**Attachment C.3. -- Summary of literature values of Pitzer interaction
parameters for mixed-salt solutions**

Attachment C.3 - Summary of Literature Values of Pitzer Parameters for Mixed Salt Solutions

INTERACTION	THETA	PSI	MEDIA	SINGLE SALT COMPATIBILITY DATA		
				B0	B1	B2
Na+ K+ Cl-	-.012 (48) -.012 (49) -.012 (88) -.0154 (51) -.0203 (46)	-.0018 (48) -.0018 (49) -.0018 (88) -.00078 (51) -.004 (46)	NaCl+KCl NaCl+KCl NaCl+KCl	.0757 .0464	.2795 .2274	.0014 -.0004
Na+ K+ Br-	-.012 (88)	-.0022 (88)				
Na+ K+ SO4-2	-.012 (48) -.012 (49) -.012 (88) .0 (47) -.0303 (31)	-.010 (48) -.010 (49) -.010 (88) -.0066 (31)	Na2SO4+K2SO4 Na2SO4+K2SO4	.0187 .050	1.0994 .779	.0055 .0
Na+ K+ HCO3-	-.0322 (31)	.0085 (31)	Na2SO4+K2SO4			
Na+ K+ CO3-2	-.012 (48)	-.003 (48)				
Na+ K+ NO3-2	-.012 (48)	.003 (48)	NaNO3+KNO3			
Na+ Ca+2 Cl-	-.012 (88) .07 (48) .07 (49) .0 (88)* -.0078 (52)* .05 (77) .07 (48) .07 (49) .0 (47)*	-.0012 (88) -.007 (48) -.014 (49) .0 (88) .0029 (52) -.003 (77) -.055 (48) -.023 (49) .0 (47)	NaCl+CaCl2 NaCl+CaCl2			
Na+ Ca+2 SO4-2	.05 (77) .07 (48) .07 (49) .0 (47)*	-.003 (77) -.055 (48) -.023 (49) .0 (47)	Na2SO4+CaSO4 Na2SO4+CaSO4	.01869 .15	1.0994 3.0	.0055488 .0
Na+ Mg+2 Cl-	.05 (77) .07 (48) .07 (49) .07 (84) .0 (88)* .050 (30)* .016130 (101)*	-.012 (77) -.0122 (48) -.0122 (49) -.010 (84) .0 (88) -.01 (30) -.004633 (101)	NaCl+MgCl2 NaCl+MgCl2 NaCl+MgCl2			
Na+ Mg+2 SO4-2	.07 (48) .07 (49) .07 (84) .0 (88)* .0 (25)* .050 (30)*	-.015 (48) -.015 (49) -.023 (84) .0 (88) .00566 (25) .0224 (30)	Na2SO4+MgSO4 Na2SO4+MgSO4 Na2SO4+MgSO4 Na2SO4+MgSO4	.0765 .35093	.2664 1.65075	.00127 .006507
Na+ H+ Cl-	.036 (48) .036 (88) .0279 (9)	-.004 (48) -.004 (88) .0105 (9)	NaCl+HCl HCl+NaCl			
Na+ H+ Br-	.036 (88)	-.012 (88)				
Na+ H+ ClO4-	.036 (56)	.0 (56)				
Na+ H+ HSO4-	.036 (48)	-.0129 (48)				
Na+ Cl- NO3-2	.016 (88) -.0279 (46)	-.006 (88) -.0035 (46)	NaCl+NaNO3 NaCl+NaNO3	.0757 .003	.2795 .2037	.0014 .00013
			Na2SO4 MgSO4 Na2SO4 MgSO4	.0187 .3150 .0187 .2150	1.0994 3.3646 1.0994 3.3646	.00550 .02797 .00550 .02797
			NaCl MgCl2	.0765 .35093	.2664 1.65075	.00127 .006507
			NaCl HCl	.0765 .1775	.2664 .2945	.00127 .00080
			NaCl NaNO3	.0757 .003	.2795 .2037	.0014 .00013
			Na2SO4 MgSO4	.0187 .3150 .0187 .2150	1.0994 3.3646 1.0994 3.3646	.00550 .02797 .00550 .02797
			NaCl HCl	.0765 .1775	.2664 .2945	.00127 .00080
			NaCl NaNO3	.0757 .003	.2795 .2037	.0014 .00013

Attachment C.3 - continued

INTERACTION			THETA	PSI	MEDIA	SINGLE SALT COMPATIBILITY DATA			
						B0	B1	CO	B2
Na+	Cl-	H2PO4-	.1004 (131)	.0 (131)	NaCl+NaH2PO4				
Na+	Cl-	H2AsO4-	.228 (75)	.0 (75)	NaCl+NaH2AsO4				
Na+	Cl-	HasO4-2	.122 (75)*	.0 (75)	NaCl+NaHasO4				
Na+	Cl-	AsO4-3	.060 (75)*	.0 (75)	NaCl+NaAsO4				
Na+	Cl-	Ac+	-.017 (75)	.0 (75)	NaCl+NaAc				
Na+	Cl-	B(OH)3-	.0 (75)*	.0 (75)	NaCl+NaB(OH)3				
Na+	Cl-	B(OH)4-	-.065 (18)	-.0073 (18)	NaCl+NaB(OH)4	NaCl			
						NaB(OH)4	-.0427	.089	.0114
Na+	Cl-	B3O3(OH)4-	.12 (18)	-.024 (18)	NaCl+NaB3O3(OH)4	NaCl			
Na+	Cl-	B4O5(OH)4-2	.074 (18)	.026 (18)	NaCl+NaB4O5(OH)4	NaB3O3(OH)4	-.056	-.910	.0
						NaCl			
						NaB4O5(OH)4	-.11	-.40	
Na+	Br-	Cl-	.0 (88)	.0 (88)	NaCl+NaBr				
Na+	Ba+2	Cl-	-.003 (88)*	.0 (88)	NaCl+BaCl2				
			.067 (84)	-.012 (84)	NaCl+BaCl2				
Na+	Cs+	Cl-	-.033 (88)	-.003 (88)	NaCl+CsCl				
Na+	Mn+2	Cl-	.0 (88)*	-.003 (88)	NaCl+MnCl2				
			.082 (84)	-.0174 (84)	NaCl+MnCl2				
			.03687 (35)*	-.01275 (35)	NaCl+MnCl2	NaCl	.0765	.2664	.000126
						MnCl2	.3420	1.4091	-.0298
Na+	Mn+2	S04-2	.04320 (40)*	-.02155 (40)					.00102
Na+	Co+2	Cl-	-.016 (13)*	-.001 (13)	NaCl+CoCl2	NaCl	Ref. (89)		
						CoCl2	.3582	1.44975	-.012537
						NaCl	.0765	.2664	.000126
						CoCl2	.3454	1.625	.00179
Na+	Co+2	S04-2	-.01939 (24)	-.00845 (24)	NaCl+CoCl2				-.00329
			-.001413 (45)*	-.008986 (45)	Na2S04+CoS04	Na2S04	.01957	1.113	.001758
						CoS04	.16310	3.3458	.00926
Na+	Cu+2	Cl-	-.01939 (24)	-.01738 (24)	Na2S04+CoS04				-.30.7
			.0 (14)*	-.014 (14)	NaCl+CuCl2	NaCl	Ref. (89)		
			.077 (14)	-.026 (14)	NaCl+CuCl2	CuCl2	.2966	1.391	-.03602
						NaCl	Ref. (89)		
						CuCl2	.2966	1.391	-.03602
						NaCl	.0765	.2665	.00062
			-.06404 (26)*	-.00358 (26)	NaCl+CuCl2	CuCl2	.2817	1.4677	-.01395
Na+	Cu+2	S04-2	.0 (14)*	-.011 (14)	Na2S04+CuS04	Na2S04	Ref. (89)		.0
						CuS04	.2340	2.527	.0044
									-.48.33
						Na2S04	Ref. (89)		
						CuS04	.2340	2.527	.0044
									-.48.33
Na+	Zn+2	Br-	.0 (88)*	----- (88)	NaBr+ZnBr2	Na2S04	.0187	1.0994	.00553
Na+	Zn+2	S04-2	.05073 (41)*	-.03089 (41)	Na2S04+ZnS04	ZnS04	.1949	2.8830	.0290
						CdCl2	-.0169	.3468	.00747
Na+	Cd+2	Cl-	-.00032 (27)	.14913 (27)	NaCl+CdCl2	Na2S04	.0187	1.0994	.00553
Na+	Cd+2	S04-2	-.00032 (41)*	-.01521 (41)	Na2S04+CdS04	CdS04	.2053	2.6170	.01140
						NaCl	.0765	.2665	.00062
Na+	Ni+2	Cl-	.05917 (26)*	.01152 (26)	NaCl+NiCl2	NiCl2	.4456	2.4076	-.00784
									.0
Na+	Cs+	S04-2	-.0153 (37)	-.0035 (37)	Na2S04+Cs2S04	Na2S04	.0187	1.0094	.00550
Na+	Ni+2	S04-2	.0027 (43)*	-.0128 (43)	Na2S04+NiS04	Cs2S04	.0714	1.2010	.00291
						Na2S04	.0187	1.0994	.0053
						NiS04	.1702	2.9070	.03660
									-.40.06

Attachment C.3 - continued

INTERACTION	THETA	PSI	MEDIA	NaNO3 NaNO2	SINGLE SALT COMPATIBILITY DATA			B2
					BO	B1	CO	
Na+ NO3- NO2-	-.0144 (46)	-.0012 (46)	NaNO3+NaNO2		.003	.2037	.00013	
K+ Ca+2 Cl-	.032 (48) .032 (49) -.040 (88)* .032 (49) -.040 (47)* .0 (48) .0 (49) .0 (47)* .0 (73)* .0 (48) .0 (49) .0 (47)*	-.025 (48) -.025 (49) -.015 (88) .0 (49) .0 (47) -.022 (48) -.022 (49) .0 (47) -.05 (73) -.048 (48) -.048 (49) .0 (47)	KCl+CaCl2 K2SO4+CaSO4		.051	.1625	-.0023	
K+ Mg+2 Cl-			KCl+MgCl2 KCl+MgCl2					
K+ Mg+2 SO4-2			K2SO4+MgSO4	K2SO4 MgSO4	.04995 .21499	.7793 3.3646	.0 .02797	-32.743
K+ H+ Cl-	.005 (48) .005 (88) .0049 (8) .0068 (8) .0074 (9) .0088 (9) .005 (88) .005 (48) .005 (48) .016 (88) -.000815 (46)	-.011 (48) -.007 (88) -.0081 (8) -.0106 (8) -.0105 (9) -.0131 (9) -.021 (88) .197 (48) -.0265 (48) -.006 (88) -.00023 (46)	HCl+KCl HCl+KCl HCl+KCl HCl+KCl HCl+KCl HCl+KCl KCl+KNO3 KCl+KNO3	Ref. (88) HCl KCl HCl KCl HCl KCl KCl KNO3	.1775 .04835 .1775 .04835 .1775 .04835 .1775 .04835	.2945 .2122 .2945 .2122 .2945 .2122 .2945 .2122	.00080 -.00084 .00080 -.00084 .00080 -.00084 .00080 -.00084	
K+ Cl- H2PO4-	.1004 (131) .1004 (95) .10 (86)	.06 (131) -.0105 (95) -.01 (86)	KCl+KH2PO4 KCl+KH2PO4 KCl+KH2PO4	Ref. (89) KCl KH2PO4	.0464 -.0802	.2274 .0722	-.0004 .0065	
K+ Br- Cl- Ba+2 Cl- Cs+ Cl- Cs+ SO4-2	.0 (88) -.072 (88)* .010 (84) .0 (88) -.0049 (36)	.0 (88) .0 (88) -.017 (84) -.0013 (88) -.0016 (36)	KCl+KBr KCl+BaCl2 KCl+BaCl2 KCl+CsCl K2SO4+Cs2SO4					
K+ Co+2 SO4-2	-.0982 (19)* -.0819 (19)* .007 (48) .007 (49) .010 (88) .007 (48) .007 (49) .01 (47)	-.0328 (19) -.0411 (19) -.012 (48) -.012 (49) .0 (88) .024 (48) .05 (49) .0 (47)	MgCl2+CaCl2	K2SO4 Cs2SO4	.0500 .0714	.7790 1.2010	.00291	
Ca+2 Mg+2 SO4-2				Ca2SO4 MgSO4	.200 .21499	2.65 3.3646	.0 .02797	-57.70 -32.743
Ca+2 Cl- NO3-	.016 (88)	-.017 (88)	CaCl2+Ca(NO3)2					

SINGLE SALT COMPATIBILITY DATA

	B0	B1	C0	B2
1				
2				
3				
4				
5				
6				
7				
8				
9				
10				
11				
12				
13				
14				
15				
16				
17				
18				
19				
20				
21				
22				
23				
24				
25				
26				
27				
28				
29				
30				
31				
32				
33				
34				
35				
36				
37				
38				
39				
40				
41				
42				
43				
44				
45				
46				
47				
48				
49				
50				
51				
52				
53				
54				
55				
56				
57				
58				
59				
60				
61				
62				
63				
64				
65				
66				
67				
68				
69				
70				
71				
72				
73				
74				
75				
76				
77				
78				
79				
80				
81				
82				
83				
84				
85				
86				
87				
88				
89				
90				
91				
92				
93				
94				
95				
96				
97				
98				
99				
100				

[illegible]

Attachment C.3 - continued

INTERACTION	THETA	PSI	MEDIA	SINGLE SALT COMPATIBILITY DATA		
				B0	B1	B2
H+ Mn+2 Cl-	.0 (88)* .0834 (111) .075 (84) -.007 (84) -.0006 (61) .0694 (61) .0666 (61)* .0818 (61) .0559 (61) -.0732 (6)*	.0 (88) -.0090 (111) -.007 (84) -.0006 (61) .0386 (61) .0051 (61) .0475 (61) .0434 (6)	HCl+MnCl2 HCl+MnCl2 HCl+MnCl2 HCl+MnCl2 HCl+MnCl2 HCl+NiCl2 HClO4+UO2(ClO4)2 NaCl+SrCl2	.1634 .61785 .0765 .2858	.3465 2.06175 .2664 1.667	.0117 .0153795 .00127 -.00130
Cl- Na+ Sr+2	.051 (102)	-.0021 (102)	NaCl+SrCl2	UO2(ClO4)2 NaCl SrCl2		
Cl- H+ Al+3	-.0084 (72)* .00940 (100)* .185 (84)	.0036 (72) -.00323 (100) .013 (84)	NaCl+SrCl2 NaCl+SrCl2 HCl+AlCl3			
Cl- SO4-2 Na+	.216 (2) .02 (48) .02 (49) .030 (14) -.02 (14)*	.004 (2) .0014 (48) .0014 (49) .000 (14) .004 (14)	HCl+AlCl3 NaCl+Na2SO4 NaCl+Na2SO4	NaCl Na2SO4	Ref. (89) Ref. (89)	
Cl- SO4-2 K+	-.035 (88)* -.045 (25)* .030 (79) -.035 (88)* .02 (84) -.045 (25)* -.045 (33)* -.02 (86)*	.007 (88) .007 (25) .0 (79) .0 (88) .0 (84) .014 (25) .018 (33) -.007 (86)	NaCl+Na2SO4 NaCl+Na2SO4 KCl+K2SO4 KCl+K2SO4 KCl+K2SO4 KCl+K2SO4	NaCl Na2SO4	.0765 .0187 .2664 1.0994	.00127 .00550
Cl- SO4-2 Ca+2	.030 (14) .030 (79) .02 (48) .02 (49) -.02 (47)*	-.005 (14) -5.0 (79) -.018 (48) .0 (49) .0 (47)	KCl+K2SO4 KCl+K2SO4 KCl+K2SO4	Pitz. Param. KCl K2SO4 Ref. (14)	.04835 .04995 .2122 .77925	-.00084
Cl- SO4-2 Mg+2	.02 (48) .02 (49) -.035 (88)* -.02 (86)* .02 (84) -.056 (32)* .030 (14) -.045 (20,27)*	-.004 (48) -.004 (49) .0 (88) -.007 (86) -.014 (84) .0129 (32) -.020 (14) .007 (20,27)	CaCl2+CaSO4 MgCl2+MgSO4 MgCl2+MgSO4 MgCl2+MgSO4 MgCl2+MgSO4 MgCl2+MgSO4 MgCl2+MgSO4	CaCl2 CaSO4 MgCl2 MgSO4 MgCl2 MgSO4 Ref. (14) MgCl2 MgSO4	.3159 .200 .35235 .2210 .3509 .2150 .3509 .3150	1.614 2.65 1.6815 3.343 1.6508 3.3646 1.6508 3.3646

Attachment C.3 - continued

INTERACTION	THETA	PSI	MEDIA	SINGLE SALT COMPATIBILITY DATA			
				B0	B1	C0	B2
C1- S04-2 Cu+2	.030 (14)	.031 (14)	CuCl2+CuS04				
C1- HS04- Na+	-.006 (48)	-.006 (48)					
C1- HS04- H+	-.006 (48)	.013 (48)					
C1- OH- Na+	-.050 (88)	-.006 (88)	NaCl+NaOH				
C1- OH- K+	-.050 (48)	-.006 (48)					
	-.050 (88)	-.008 (88)	KCl+KOH				
C1- OH- Ca+2	-.050 (48)	-.025 (48)					
C1- HC03- Na+	.03 (48)	-.015 (48)					
	.0359 (81)	-.0143 (81)	NaCl+NaHC03				
				.0765	.2664	.00127	
	.030 (92)	----- (92)	NaCl+NaHC03	.028	.044		
C1- HC03- K+	-.063 (112)	-.0037 (112)	KCl+KHC03				
				.0277	.0411		
C1- HC03- Mg+2	.03 (48)	-.096 (48)					
C1- C03-2 Na+	-.02 (48)	.0085 (48)					
	-.1605 (129)*	.0503 (129)					
			NaCl+Na2C03				
	-.2343 (47)*	.0431 (47)		.07669	.026461	.0012193	
			NaCl+Na2C03	.040822	1.4679	.0042374	
	-.053 (81)	.016 (81)		.07669	.26461	.0012193	
	-.092 (81)*	.006 (81)	NaCl+Na2C03	.040822	1.4679	.0042374	
				.0765	.2664	.00127	
	-.02 (48)	.004 (48)	NaCl+Na2C03	.0362	1.510	.0052	
	-.053 (112)	.024 (112)		.0765	.2664	.00127	
	-.092 (112)*	.0073 (112)	NaCl+Na2C03	.0362	1.510	.0052	
	-.18140 (28)*	.08718 (28)	KCl+K2C03	.04835	.2122	-.00084	
			KCl+K2C03	.1288	1.433	.00050	
	-.045 (24)	.02911 (24)	KCl+K2C03	.04835	.2122	-.00084	
	-.16242 (28)*	.07202 (28)	CoCl2+CoS04	.1288	1.433	.00050	
				.41908	2.06641	-.02429	-.35136
				Ref. (90)			
C1- N1+2 S04-2			N1Cl2+N12S04				
				.44104	2.34470	-.02156	-.53354
				Ref. (90)			
S04-2 HS04- Na+	.0 (48)	-.0094 (48)					
S04-2 HS04- K+	.0 (48)	-.0677 (48)					
S04-2 HS04- Mg+2	.0 (48)	-.0425 (48)					
S04-2 OH- Na+	-.013 (48)	-.009 (48)					
S04-2 OH- K+	-.013 (48)	-.050 (48)					
S04-2 HC03- Na+	.01 (48)	-.005 (48)					
S04-2 HC03- Mg+2	.01 (48)	-.161 (48)					
S04-2 C03-2 Na+	.02 (48)	-.005 (48)					
S04-2 C03-2 K+	.02 (48)	-.009 (48)					
OH- C03-2 Na+	.10 (48)	-.017 (48)					
	.0 (78)*	.0 (78)	NaOH+Na2C03				
				.0864	.253	.0044	
				.0362	1.51	.0052	
OH- C03-2 K+	.10 (48)	-.01 (48)					
OH- Br- Na+	-.065 (88)	-.018 (88)					
OH- Br- K+	-.065 (88)	-.014 (88)					

Attachment C.3 - continued

INTERACTION	THETA	PSI	MEDIA	SINGLE SALT COMPATIBILITY DATA			
				B0	B1	C0	B2
HC03- C03-2 Na+	-.04 (48)	.002 (48)					
HC03- C03-2 K+	-.04 (48)	.012 (48)					
	.089 (112)	-.036 (112)	KHC03+K2C03				
	-.063 (112)*	.013 (112)	KHC03+K2C03				
Li+ Na+ Cl-	.012 (88)	-.003 (88)	LiCl+NaCl				
Li+ Na+ NO3-2	.012 (88)	-.0072 (88)	LiNO3+NaNO3				
Li+ Na+ ClO4-	.012 (88)	-.0080 (88)	LiClO4+NaClO4				
Li+ K+ Cl-	-.022 (88)	-.010 (88)	LiCl+KCl				
Li+ Cl- NO3-2	.016 (88)	-.003 (88)	LiCl+LiNO3				
Li+ Ba+2 Cl-	-.070 (88)*	.019 (88)	LiCl+BaCl2				
	.0 (84)	.009 (84)	LiCl+BaCl2				
Li+ Cs+ Cl-	-.095 (88)	-.0094 (88)	LiCl+CsCl				
Li+ Cs+ SO4-2	-.1242 (37)	.0088 (37)	Li2SO4+Cs2SO4				
Li+ Cd+2 SO4-2	.0 (42)*	.0 (42)	Li2SO4+CdSO4				
Li+ Ni+2 Cl-	-.059 (69)*	.047 (69)	LiCl+NiCl2				
Li+ Mn+2 SO4-2	.0 (40)*	.0 (40)	Li2SO4+MnSO4				
Cu+2 Ni+2 NO3-2	.01314 (22)	-.003909 (22)	Cu(NO3)2+Ni(NO3)2				
Cu+2 Zn+2 NO3-2	.00664 (21)	-.00141 (21)	Cu(NO3)2+Zn(NO3)2				
Cs+ Ba+2 Cl-	-.150 (88)*	.0 (88)	CsCl+BaCl2				
	-.070 (84)	.015 (84)	CsCl+BaCl2				
Rb+ Cd+2 SO4-2	-.1226 (20)*	.3730 (20)	Rb2SO4+CdSO4				
Rb+ Co+2 SO4-2	-.3058 (34)*	.0466 (34)	Rb2SO4+CoSO4				
NH4+ SO4-2 H2PO4-	.01837 (29)*	-.00427 (29)	NH4H2PO4+(NH4)2SO4				

* - Higher order electrostatic terms not included. Values of theta and psi not consistent with PHRQPITZ.

Attachment C.4. -- Temperature dependence of single-salt parameters

Attachment C.4 - Temperature dependence of single salt parameters

		dB0/dT	dB1/dT	dCO/dT	dB2/dT
Na+	C1-	7.159E-4 (121)	7.005E-4 (121)	-10.54E-5 (121)	
Na+	Br-	.008946 (120)	6.160E-5 (120)	-4.655E-5 (120)	
Na+	S04-2	2.367E-3 (121)	10.79E-4 (121)	-9.30E-5 (121)	
Na+	I-	.002349 (106)	5.6325E-3 (121)	-4.879037 (121)	
Na+	F-	8.355E-4 (121)	.005958 (106)	- .00479 (106)	
Na+	OH-	9.163E-4 (118)	8.28E-4 (121)	-8.35E-5 (121)	
Na+	HC03-	5.361E-4 (121)	8.70E-4 (121)	-5.58E-5 (118)	
Na+	C03-2	7.00E-4 (121)	1.34E-4 (121)	-18.94E-5 (121)	
Na+	C104-	1.00E-3 (81)	1.10E-3 (81)		
Na+	N03-	12.96E-4 (121)	2.05E-3 (81)	-16.23E-5 (121)	
Na+	C103-	12.66E-4 (121)	20.60E-4 (121)	-23.16E-5 (121)	
Na+	Br03-	10.35E-4 (121)	19.07E-4 (121)	-9.29E-5 (121)	
Na+	I03-	20.66E-4 (121)	34.37E-4 (121)		
Na+		11.79E-4 (118)	60.57E-4 (121)		
K+	C1-	5.794E-4 (121)	59.38E-4 (118)	-5.095E-5 (121)	
K+	Br-	7.39E-4 (121)	10.71E-4 (121)	-7.004E-5 (121)	
K+	S04-2	1.44E-3 (121)	17.40E-4 (121)		
K+	I-	9.914E-4 (121)	11.86E-4 (121)	-9.44E-5 (121)	
K+	HC03-	2.14E-4 (121)	5.44E-4 (121)	-5.95E-5 (121)	
K+	C03-2	1.1E-3 (112)	1.104E-3 (81)		
K+	C104-	1.788E-3 (81)	4.36E-3 (112)		
K+	N03-	.60E-4 (121)	2.051E-3 (81)		
K+	C103-	2.06E-4 (118)	10.07E-3 (121)		
K+	H2P04-	19.87E-4 (121)	95.19E-4 (118)	39.7E-5 (121)	
K+	I03-	6.045E-4 (121)	64.5E-4 (121)	-10.11E-5 (121)	
H+	C1-	21.10E-4 (118)	31.8E-4 (121)		
H+	Br-	-3.081E-4 (121)	28.6E-4 (121)		
H+	I-	-2.049E-4 (121)	45.36E-4 (118)		
H+	C104-	4.905E-4 (121)	1.419E-4 (121)	6.213E-5 (121)	
Mg+2	C1-	-1.9425E-3 (121)	1.419E-4 (118)	3.106E-5 (118)	
Mg+2	Br-	-4.29E-3 (86)	4.467E-4 (121)	-5.685E-5 (121)	
Mg+2	S04-2	- .05625E-3 (121)	8.86E-4 (121)	-7.32E-5 (121)	
Mg+2	C104-	- .69E-3 (121)	19.31E-4 (121)	-11.77E-5 (121)	
Mg+2	N03-	.52275E-3 (121)	2.775E-3 (121)	-1.6493266 (121)	
Ca+2	C1-	-1.725E-3 (121)	3.6525E-3 (86)		
Ca+2	Br-	1.049E-4 (82)	3.8625E-3 (121)	-2.53E-1 (121)	
Ca+2	S04-2	- .52275E-3 (121)	1.53E-2 (121)	-3.5266952 (121)	
Ca+2	C104-	.8925E-3 (121)	4.5E-3 (121)		
Ca+2	N03-	.5295E-3 (121)	4.4925E-3 (121)	-1.565E-4 (82)	
NH4+	C1-	.779E-4 (121)	3.9E-3 (121)		
NH4+	HC03-	.5794E-3 (114)	3.54E-3 (82)		
NH4+	H2P04-	-12.6E-3 (114)	6.0375E-3 (121)	-5.16E-1 (121)	
NH4+	Br+	1.51E-4 (121)	5.46E-2 (121)		
Rb+	C1-	6.780E-4 (121)	5.0775E-3 (121)	-3.0918245 (121)	
Rb+	S04-2	.9375E-3 (121)	9.1875E-3 (121)		
Rb+	I-	8.578E-4 (121)	12.58E-4 (121)	2.10E-5 (121)	
			1.071E-3 (114)	-.05E-3 (114)	
			4.36E-3 (114)		
			-1.75E-3 (114)		
			22.8E-4 (121)		
			15.06E-4 (121)		
			20.35E-4 (121)		
			8.64E-3 (121)		
			23.83E-4 (121)		

Attachment C.4 - continued

	dB0/dT	dB1/dT	dC0/dT	dB2/dT
Rb+	F-	-1.76E-4 (121)		
Cs+	C1-	8.28E-4 (121)	-12.25E-5 (121)	
		7.030E-4 (118)	-4.06E-5 (118)	
Cs+	Br-	7.80E-4 (121)		
Cs+	SO4-2	.8925E-3 (121)		
Cs+	I-	9.75E-4 (121)		
Cs+	F-	.95E-4 (121)		
Ni+2	C104-	.666E-3 (121)	-3.8236801 (121)	
Cu+2	C1-	-2.715E-3 (121)		
Cu+2	SO4-2	-4.4E-3 (121)	4.80E-3 (121)	-4.73E-1 (121)
Zn+2	SO4-2	-3.66E-3 (121)	3.97E-3 (121)	-3.33E-1 (121)
Zn+2	C104-	.59625E-3 (121)	-3.8554999 (121)	
Co+2	SO4-2	-2.79E-3 (121)	2.61E-3 (121)	-5.22E-1 (121)
Co+2	C104-	.54525E-3 (121)	-3.5903348 (121)	
Cr+2	C1-	.717E-3 (121)	----- (121)	
Cr+2	Br-	-.32775E-3 (121)		
Cr+2	SO4-2	-2.9E-3 (102)	3.0E-3 (102)	-4.42 (102)
Cr+2	C104-	1.143E-3 (121)	-3.1077344 (121)	
Cr+2	N03-	.177E-3 (121)		
Ba+2	C1-	.6405E-3 (121)	-1.5379573 (121)	
Ba+2	Br-	-.33825E-3 (121)		
Ba+2	N03-	-2.91E-3 (121)		
Mn+2	C104-	.39675E-3 (121)	-3.3357764 (121)	
Li+2	C1-	-1.685E-4 (121)	-4.520E-5 (121)	
Li+2	Br-	-1.819E-4 (121)	-2.813E-5 (121)	
Li+2	SO4-2	.5055E-3 (121)	-2.3334525 (121)	
Li+2	C104-	.386E-4 (121)	-7.712E-5 (121)	

**Attachment C.5. -- Summary of analytical expressions for temperature
(and pressure) dependence of selected single-salt
Pitzer interaction parameters**

Attachment C.5 - Summary of analytical expressions for temperature (and pressure) dependence of selected single salt Pitzer interaction parameters *

Parameter Number	Solution NaCl (120,5)	CaCl2 (1,1)	Parameter Number NaOH (80,3)	
1	.0765	3.39701E-1	6	-2.8131778E-3
2	-777.03	5.10778E+3	7	3.7012540E-1
3	-4.4706	4.64785E+1	8	7.1788733E-6
4	.008946	-1.38971E-1	9	-5.8847404E-9
5	-3.3158E-6	6.82557E-5	10	-4.8217410E-3
6	.2664	1.50481	11	4.0534778E-4
7	-----	-4.98241E+2	12	-5.1714017E-2
8	-----	.0	13	-1.0553037E-6
9	6.1608E-5	-1.64685E-2	14	8.9893405E-10
10	1.0715E-6	2.34630E-5	17	-4.9359970E+1
11	.00127	-2.67882E-2	20	1.1931122E+1
12	33.317	-9.67633E+3	22	4.6286977E+2
13	.09421	-8.48338E+1	23	-1.0294181E+4
14	-4.655E-5	2.46219E-1	24	-8.5960581E+1
15	-----	-1.18352E-4	25	2.3905969E-1
16	41587.11		26	-1.0795894E-4
17	-315.90		29	2.9680772E-0
18	.8514		32	-6.8923899E-1
19	-8.3637E-4		33	-8.1156286E-2
			34	2.7682478E+2
			35	-7.3755443E+3
			36	1.0945106E-1
			37	-4.0218506E-5
			38	2.4824963E-0
			39	-1.6686897E+1
			40	4.5364961E+2
			41	-6.5161667E-3
			42	2.3765786E-6

	KCl (54,2)	CsCl (54,2)	HCl (115,4)	LiCl (54,2)
B0 1	.04808	.03352	.17416	.14847
2	-758.48	-1290.0	-773.62	.0
3	-4.7062	-8.4279	-4.5174	.0
4	.010072	.018502	8.1556E-3	-1.546E-4
5	-3.7599E-6	-6.7942E-6	-2.8525E-6	.0
6	.0	.0	.0	.0
B1 1	.0476	.0429	.28799	.307
2	303.9	-38.0	-374.50	.0
3	1.066	.0	-4.1319	.0
4	.0	.001306	1.0855E-2	6.36E-4
5	.0	.0	-9.2990E-7	.0
6	.0470	.0	.0	.0
C0 1	-7.88E-4	-2.62E-4	.0	.003710
2	91.270	157.13	.0	4.115
3	.58643	1.0860	.0	.0
4	- .0012980	- .0025242	.0	.0
5	4.9567E-7	9.840E-7	.0	-3.71E-9
6	.0	.0	.0	.0

Attachment C.5 - continued

B0	1	MgS04 (79,6)	Na2S04 (79,10)	K2S04 (79,7)	MgC12 (79,9)
		-1.0282	-1.727E-2	.0	5.93915E-7
		8.4790E-3	1.7828E-3	7.476E-4	-9.31654E-4
		-2.33667E-5	9.133E-6	.0	.576066
		2.1575E-8	.0	4.265E-3	-----
		6.8402E-4	-6.552	-3.088	-----
B1	7	.21499	.0	.0	-----
		-----	-96.90	.0	-----
		-2.9596E-1	.7534	.6179	2.60169E-5
		9.4564E-1	5.61E-3	6.85E-3	-1.09438E-2
		.0	-5.7513E-4	5.576E-5	2.60135
		.0	1.11068	-5.841E-2	-----
C0	7	1.1028E-2	-378.82	.0	-----
		3.3646	.0	-90	-----
		-----	1861.3	.0	-----
		1.0541E-1	1.1745E-2	9.1547E-3	2.41831E-7
		-8.9316E-4	-3.3038E-4	.0	-2.49949E-4
		2.51E-6	1.85794E-5	.0	5.95320E-2
B2	6	-2.3436E-9	-3.9200E-2	-1.81E-4	-----
		-8.7899E-5	14.2130	.0	-----
		.006993	.0	.0	-----
		-----	-24.950	.0	-----
		1.0541E-1	-----	-----	-----
		1.2121E-1	-----	-----	-----
B0	5	-2.7642E-4	-----	-----	-----
		.0	-----	-----	-----
		-2.1515E-1	-----	-----	-----
		-32.743	-----	-----	-----
		-----	-----	-----	-----
		-----	-----	-----	-----
B1	5	MgC12 (11)	HCl (53)	MgC12 (79,9)	MgC12 (79,9)
		5.93915E-7	.17690	.0	5.93915E-7
		-9.31654E-4	-9.140E-2	7.476E-4	-9.31654E-4
		.576066	.0	.0	.576066
		-----	-4.034E-4	4.265E-3	-----
		-----	6.02E-6	-3.088	-----
C0	5	2.60169E-5	.2973	.6179	2.60169E-5
		-1.09438E-2	16.147	6.85E-3	-1.09438E-2
		2.60135	-1.7631E-2	5.576E-5	2.60135
		-----	.0	-5.841E-2	-----
		-----	7.20E-5	.0	-----
		2.41831E-7 (2**3/2C)	.362E-3	.0	-----
B0	5	-2.49949E-4 (2**3/2C)	.0	-----	-----
		5.95320E-2 (2**3/2C)	.0	-----	-----
		-----	-3.036E-5	-----	-----
		-----	.0	-----	-----
		-----	-----	-----	-----
		-----	-----	-----	-----
B1	7	NaCl (79,8)	B1	K2S04 (79,7)	MgC12 (79,9)
		B0	B1	.0	5.93915E-7
		-656.81518	119.31966	7.476E-4	-9.31654E-4
		24.86913	-4.8309327E-1	.0	.576066
		5.3812753E-5	.0	4.265E-3	-----
		-5.5887470E-8	.0	-3.088	-----
C0	7	6.589326E-12	.0	.0	-----
		-4.4640952	.0	-90	-----
		1.1109914E-2	1.4068095E-3	.0	-----
		-----	-----	9.1547E-3	2.41831E-7
		-----	-----	.0	-2.49949E-4
		-----	-----	.0	5.95320E-2
B1	7	NaCl (79,8)	B1	K2S04 (79,7)	MgC12 (79,9)
		B0	B1	.0	5.93915E-7
		-656.81518	119.31966	7.476E-4	-9.31654E-4
		24.86913	-4.8309327E-1	.0	.576066
		5.3812753E-5	.0	4.265E-3	-----
		-5.5887470E-8	.0	-3.088	-----
C0	7	6.589326E-12	.0	.0	-----
		-4.4640952	.0	-90	-----
		1.1109914E-2	1.4068095E-3	.0	-----
		-----	-----	9.1547E-3	2.41831E-7
		-----	-----	.0	-2.49949E-4
		-----	-----	.0	5.95320E-2

Attachment C.5 - continued

NaCl (79.8)		NaOH (79.11)		NaCl (77)		Na2SO4 (77)		CaCl2 (77)		CaSO4 (77)	
B0	B1	B0	B1	B0	B1	B0	B1	B0	B1	B0	B1
8	-2.6573399E-7	1	2.7682478E+2	1	1.43783204d+01	1	8.16920027d+01	1	-9.41895832d+01	1	-9.0550901E-8
9	1.7460070E-10	2	-2.8131778E-3	2	5.60767406d-03	2	3.011049570d-02	2	-4.04750026d-02	2	.0
10	1.0462619E-14	3	-7.3755443E+3	3	-4.22185236d+02	3	-2.32193726d+03	3	2.34550368d+03	3	.0
11	-5.3070129E-6	4	3.7012540E-1	4	-2.51226677d00	4	-1.43780207d+01	4	1.70912300d+01	4	-1.5386008E-8
12	8.6340233E-10	5	-4.9359970E+1	5	.0	5	-6.66496111d-01	5	-9.22885841d-01	5	8.6926600E-11
13	-4.1785962E-13	6	1.0945106E-1	6	-2.61718135d-06	6	-1.03923656d-05	6	1.51488122d-05	6	.0
14	-1.579366	7	7.1788733E-6	7	4.43854508d00	7	.0	7	-1.39082000d00	7	.0
15	2.2022821E-3	8	-4.0218506E-5	8	-4.83060685d-01	8	1.00463018d+03	8	3.47870000d00	8	.0
16	-1.3105503E-7	9	-5.8847404E-9	9	-1.70502337d00	9	5.77453682d-01	9	-1.54170000d-02	9	3.0d00
17	-6.3813683E-11	10	1.1931122E-1	10	1.40677479d-03	10	-2.18434467d+04	10	.0	10	.0
18	9.7065780	11	2.4824963	11	1.19311989d+02	11	-1.89110656d+02	11	.0	11	.0
19	-2.6860396E-2	12	-4.8217410E-3	12	.0	12	-2.035505488-01	12	.0	12	.0
20	1.5344744E-5						-3.23949532d-04		.0	20	5.1904777E-4
21	-3.2153983E-9						1.46772243d+03		.0	21	.0
							.0		.0		
							-4.23433299d00		.0		

Attachment C.5 - continued

	NaCl (77)	Na2SO4 (77)	CaCl2 (77)	CaSO4 (77)
C0 1	-1.00588714d-01	-8.07816886d+01	-3.03578731d+01	.0
2	-1.80529413d-05	-3.54521126d-02	-1.36264728d-02	.0
3	8.61185543d00	2.02438830d+03	7.64582238d+02	.0
4	1.24880954d-02	1.46197730d+01	5.50458061d00	.0
5	.0	-9.16974740d-02	-3.27377782d-01	.0
6	3.41172108d-08	1.43946005d-05	5.69405869d-06	.0
7	6.83040995d-02	-2.42272049d00	-5.36231106d-01	.0
8	2.93922611d-01	.0	.0	.0
B2 1	-----	-----	-----	-----
2	-----	-----	-----	-----
3	-----	-----	-----	-----
4	-----	-----	-----	-----
5	-----	-----	-----	-----
6	-----	-----	-----	-----
7	-----	-----	-----	-----
8	-----	-----	-----	-----

Equations

- 1) $B(0) = q1 + q2(1/T - 1/Tr) + q3 \ln(T/Tr) + q4 \ln(T - Tr) + q5(T^{**2} - Tr^{**2})$
 $B(1) = q6 + q7(1/T - 1/Tr) + q9(T - Tr) + q10(T^{**2} - Tr^{**2})$
 $C0 = q11 + q12(1/T - 1/Tr) + q13 \ln(T/Tr) + q14(T - Tr) + q15(T^{**2} - Tr^{**2})$
- 2) $f(T) = p1 + p2(1/T - 1/Tr) + p3 \ln(T/Tr) + p4(T - Tr) + p5(T^{**2} - Tr^{**2}) + p6 \ln(T - 260)$
- 3) $B(0) = U34 + U6P + (U35 + U7P)/T + U17 \ln T + (U36 + U8P)T + (U37 + U9P)T^{**2} + U20/(T - 227) + (U38 + U10P)/(647 - T)$
 $B(1) = U22 + U23/T + U24 \ln T + U25T + U26T^{**2}$
 $C0 = U39 + U11P + (U40 + U12P)/T + U29 \ln T + (U41 + U13P)T + (U42 + U14P)T^{**2} + U32/(T - 227) + U33/(647 - T)$
- 4) $B^* = q1 + q2(1/T - 1/Tr) + q3 \ln(T/Tr) + q4(T - Tr) + q5(T^{**2} - Tr^{**2})$
- 5) $B(0) = q1 + q2(1/T - 1/Tr) + q3 \ln(T/Tr) + q4(T - Tr) + q5(T^{**2} - Tr^{**2})$
 $B(1) = q6 + q9(T - Tr) + q10(T^{**2} - Tr^{**2})$
 $C0 = q11 + q12(1/T - 1/Tr) + q13 \ln(T/Tr) + q14(T - Tr)$
- 6) $MgSO4$
 $f(T) = q1(T/2 + 298^{**2}/2T - 298) + q2(T^{**2}/6 + 298^{**3}/3T - 298^{**2}/2) + q3(T^{**3}/12 + 298^{**4}/4T - 298^{**3}/3)$
 $+ q4(T^{**4}/20 + 298^{**5}/5T - 298^{**4}/4) + (298 - 298^{**2}/T)Q5 + Q6$
- 7) $K2SO4$
 $f(T) = q1 + q2(T - Tr^{**2}/T) + q3(T^{**2} + 2Tr^{**3}/T - 3Tr^{**2}) + q4(T + Tr^{**2}/T - 2Tr) + q5[\ln(T/Tr) + Tr/T - 1]$
 $+ q6\{1/(T - 263) + (263T - Tr^{**2})/[T(Tr - 263)^2]\} + q7\{1/(680 - T) + (Tr^{**2} - 680T)/[T(680 - Tr)^{**2}]\}$
- 8) $NaCl$
 $f(T) = Q1/T + Q2 + Q3P + Q4P^{**2} + Q5P^{**3} + Q6 \ln(T) + (Q7 + Q8P + Q9P^{**2} + Q10P^{**3})T + (Q11 + Q12P + Q13P^{**2})T^{**2}$
 $+ (Q14 + Q15P + Q16P^{**2} + Q17P^{**3})/(T - 227) + (Q18 + Q19P + Q20P^{**2} + Q21P^{**3})/(630 - T)$
- 9) $MgCl2$
 $f(T) = Q1T^{**2} + Q2T + Q3$
- 10) $Na2SO4$
 $f(T) = Q1 + Q2(T - Tr^{**2}/T) + Q3(T^{**2} + 2Tr^{**3}/T - 3Tr^{**2}) + Q4(T + Tr^{**2}/T - 2Tr) + Q5[\ln(T/Tr) + Tr/T - 1]$
 $+ Q6\{1/(T - 263) + (263T - Tr^{**2})/[T(Tr - 263)^2]\} + q7\{1/(680 - T) + (Tr^{**2} - 680T)/[T(680 - Tr)^{**2}]\}$
- 11) $NaOH$
 $f(T) = Q1 + Q2P + (Q3 + Q4P)/T + Q5 \ln(T) + (Q6 + Q7P)T + (Q8 + Q9P)T^{**2} + Q10/(T - 227) + (Q11 + Q12P)/(647 - T)$

* numbers in parentheses following designation of a salt correspond to reference number (Attachment C.1) and equation number given below.

Attachment C.6. -- Sources of data in the file PITZER.DATA

Attachment C.6 - Sources of data in file PITZER.DATA

SPECIES	21	1	4	5	6	7	8	9	11	12	14	15	16	21	22	31	34	35	40	76	85	86
80																						
NA+	CL-								0.0765 (48)					-777.03 (120)								
K+	CL-								0.04835 (48)											0.008946 (120)		-3.3158E-6 (120)
MG+2	CL-								0.35235 (48)											5.794E-4 (121)		
CA+2	CL-								0.3159 (48)											-1.943E-4 (121)		
MGOH+	CL-								-0.1 (48)											-1.725E-4 (121)		
H+	CL-								0.1775 (48)											-3.081E-4 (121)		
LI+	CL-								0.1494 (89)											-1.685E-4 (121)		
SR+2	CL-								0.28575 (89)											0.717E-3 (121)		
FE+2	CL-								0.335925 (89)													
MN+2	CL-								0.327225 (89)													
BA+2	CL-								0.2628 (89)													
NA+	BR-								0.0973 (89)											0.6405E-3 (121)		
K+	BR-								0.0569 (89)											7.692E-4 (121)		
H+	BR-								0.1960 (89)											7.39E-4 (121)		
MG+2	BR-								0.43268 (89)											-2.049E-4 (121)		
CA+2	BR-								0.3816 (89)											-5.625E-5 (121)		
LI+	BR-								0.1748 (89)											-5.2275E-4 (121)		
SR+	BR-								0.331125 (89)											-1.819E-4 (121)		
BA+	BR-								0.31455 (89)											-0.32775E-3 (121)		
NA+	S04-2								0.01958 (48)											-0.33825E-3 (121)		
K+	S04-2								0.04995 (48)											2.367E-3 (121)		
MG+2	S04-2								0.221 (48)											1.44E-3 (121)		
CA+2	S04-2								0.2 (48)											-0.69E-3 (121)		
H+	S04-2								0.0298 (48)													
LI+	S04-2								0.136275 (89)											0.5055E-3 (121)		
SR+2	S04-2								0.220 (102)											-2.9E-3 (102)		
FE+2	S04-2								0.2568 (103)													
MN+2	S04-2								0.2065 (40)													
NA+	HS04-								0.0454 (48)													
K+	HS04-								-0.0003 (48)													
MG+2	HS04-								0.4746 (48)													
CA+2	HS04-								0.2145 (48)													
H+	HS04-								0.2065 (48)													
FE+2	HS04-								0.4273 (103)													
NA+	OH-								0.0864 (48)											7.00E-4 (121)		
K+	OH-								0.1298 (48)													
CA+2	OH-								-0.1747 (48)													
LI+	OH-								0.015 (89)													
BA+2	OH-								0.17175 (89)													
NA+	HC03-								0.0277 (48)											1.00E-3 (81)		
K+	HC03-								0.0296 (48)											0.996E-3 (81)		
MG+2	HC03-								0.329 (48)													
CA+2	HC03-								0.4 (48)													
SR+2	HC03-								0.12 (75)											1.79E-3 (81)		
NA+	C03-2								0.0399 (48)											1.788E-3 (81)		
K+	C03-2								0.1488 (48)													
SR+2	C03-2								-11.9 (75)													
81																						
NA+	CL-								0.2664 (49)											6.1608E-5 (120)		1.0715E-6 (120)
K+	CL-								0.2122 (48)											10.71E-4 (121)		
MG+2	CL-								1.6815 (48)											3.6525E-3 (86)		

Attachment C.6 - continued

CA+2	CL-	1.614 (48)	3.9E-3 (121)
MGOH+	CL-	1.658 (48)	
H+	CL-	0.2945 (48)	1.419E-4 (121)
LI+	CL-	0.3074 (89)	5.366E-4 (121)
SR+2	CL-	1.66725 (89)	2.8425E-3 (121)
FE+2	CL-	1.53225 (89)	
MN+2	CL-	1.55025 (89)	
BA+2	CL-	1.49625 (89)	
NA+	BR-	0.2791 (89)	3.2325E-3 (121)
K+	BR-	0.2212 (89)	10.79E-4 (121)
H+	BR-	0.3564 (89)	17.40E-4 (121)
MG+2	BR-	1.75275 (89)	4.467E-4 (121)
CA+2	BR-	1.613 (89)	3.8625E-3 (121)
LI+	BR-	0.2547 (89)	6.0375E-3 (121)
SR+2	BR-	1.7115 (89)	6.636E-4 (121)
BA+2	BR-	1.56975 (89)	6.5325E-3 (121)
NA+	S04-2	1.113 (48)	6.78E-3 (121)
K+	S04-2	0.77925 (48)	5.6325E-3 (121)
MG+2	S04-2	3.343 (48)	6.6975E-3 (121)
CA+2	S04-2	3.1973 (48)	1.53E-2 (121)
LI+	S04-2	1.2705 (89)	5.46E-2 (121)
SR+2	S04-2	2.88 (102)	1.41E-3 (121)
FE+2	S04-2	3.063 (103)	27.0 (102)
MN+2	S04-2	2.9511 (40)	
NA+	HS04-	0.398 (48)	
K+	HS04-	0.1735 (48)	
MG+2	HS04-	1.729 (48)	
CA+2	HS04-	2.53 (48)	
H+	HS04-	0.5556 (48)	
FE+2	HS04-	3.48 (103)	1.34E-4 (121)
NA+	OH-	0.253 (48)	
K+	OH-	0.32 (48)	
CA+2	OH-	-0.2303 (48)	
LI+	OH-	0.14 (89)	
BA+2	OH-	1.2 (89)	
NA+	HC03-	0.0411 (48)	1.10E-3 (81)
K+	HC03-	-0.013 (48)	1.104E-3 (81)
MG+2	HC03-	0.6072 (48)	
CA+2	HC03-	2.977 (48)	
NA+	C03-2	1.389 (48)	2.05E-3 (81)
K+	C03-2	1.43 (48)	2.051E-3 (81)
B2			
MG+2	S04-2	-37.23 (48)	-0.253 (121)
CA+2	S04-2	-54.24 (48)	-0.516 (121)
SR+2	S04-2	-41.8 (48)	-0.42 (102)
FE+2	S04-2	-42.0 (103)	
MN+2	S04-2	-40.0 (40)	
CA+2	OH-	-5.72 (48)	
C0			
NA+	CL-	0.00127 (48)	-4.655E-5 (120)
K+	CL-	-0.00084 (48)	-5.095E-5 (121)
MG+2	CL-	0.00519 (48)	-1.6493266 (121)
CA+2	CL-	-0.00034 (48)	
H+	CL-	0.0008 (48)	6.213E-5 (121)
LI+	CL-	0.00359 (89)	-4.520E-5 (121)

Attachment C.6 - continued

SR+2	CL-	-0.00130461 (89)	
FE+2	CL-	-0.00860725 (89)	
MN+2	CL-	-0.0204972 (89)	
BA+2	CL-	-0.0193782 (89)	
NA+	BR-	0.00116 (89)	-1.5379573 (121)
K+	BR-	-0.00180 (89)	-9.30E-5 (121)
H+	BR-	0.00827 (89)	-7.004E-5 (121)
MG+2	BR-	0.00312 (89)	-5.685E-5 (121)
CA+2	BR-	-0.00257 (89)	
LI+	BR-	0.0053 (89)	-2.813E-5 (121)
SR+2	BR-	0.00122506 (89)	
BA+2	BR-	-0.0159576 (89)	
NA+	S04-2	0.00497 (48)	-4.879037E-4 (121)
MG+2	S04-2	0.025 (48)	0.523E-3 (121)
H+	S04-2	0.0438 (48)	
LI+	S04-2	0.00399338 (89)	
SR+2	S04-2	0.019 (102)	
FE+2	S04-2	0.0209 (103)	
MN+2	S04-2	0.01636 (40)	-2.3334525 (121)
NA+	OH-	0.0044 (48)	3.0E-3 (102)
K+	OH-	0.0041 (48)	
K+	HC03-	-0.008 (48)	-18.94E-5 (121)
NA+	C03-2	0.0044 (48)	
K+	C03-2	-0.0015 (48)	
THETA			
K+	NA+	-0.012 (48)	
MG+2	NA+	0.07 (48)	
CA+2	NA+	0.07 (48)	
H+	NA+	0.036 (48)	
CA+2	K+	0.032 (48)	
H+	K+	0.005 (48)	
CA+2	MG+2	0.007 (48)	
H+	MG+2	0.1 (48)	
H+	CA+2	0.092 (48)	
S04-2	CL-	0.02 (48)	
HS04-	CL-	-0.006 (48)	
OH-	CL-	-0.05 (48)	
HC03-	CL-	0.03 (48)	
C03-2	CL-	-0.02 (48)	
OH-	BR-	-0.065 (88)	
OH-	S04-2	-0.013 (48)	
HC03-	S04-2	0.01 (48)	
C03-2	S04-2	0.02 (48)	
C03-2	OH-	0.1 (48)	
C03-2	HC03-	-0.04 (48)	
LAMDA			
NA+	H2C03	0.1 (48)	
K+	H2C03	0.051 (48)	
MG+2	H2C03	0.183 (48)	
CA+2	H2C03	0.183 (48)	
CL-	H2C03	-0.005 (48)	
S04-2	H2C03	0.097 (48)	
HS04-	H2C03	-0.003 (48)	
PSI			
NA+	K+	CL-	-0.0018 (48)

NA+	K+	BR-	-0.0022 (88)
NA+	K+	S04-2	-0.010 (48)
NA+	K+	HC03-	-0.003 (48)
NA+	K+	C03-2	0.003 (48)
NA+	CA+2	CL-	-0.007 (48)
NA+	CA+2	S04-2	-0.055 (48)
NA+	MG+2	CL-	-0.012 (48)
NA+	MG+2	S04-2	-0.015 (48)
NA+	H+	CL-	-0.004 (48)
NA+	H+	BR-	-0.012 (88)
NA+	H+	HS04-	-0.0129 (48)
K+	CA+2	CL-	-0.025 (48)
K+	MG+2	CL-	-0.022 (48)
K+	MG+2	S04-2	-0.048 (48)
K+	H+	CL-	-0.011 (48)
K+	H+	BR-	-0.021 (88)
K+	H+	S04-2	0.197 (48)
K+	H+	HS04-	-0.0265 (48)
CA+2	MG+2	CL-	-0.012 (48)
CA+2	MG+2	S04-2	0.024 (48)
CA+2	H+	CL-	-0.015 (48)
MG+2	MG0H+	CL-	0.028 (48)
MG+2	H+	CL-	-0.011 (48)
MG+2	H+	HS04-	-0.0178 (48)
CL-	BR-	K+	0.0000 (88)
CL-	S04-2	NA+	0.0014 (48)
CL-	S04-2	CA+2	-0.018 (48)
CL-	S04-2	MG+2	-0.004 (48)
CL-	HS04-	NA+	-0.006 (48)
CL-	HS04-	H+	0.013 (48)
CL-	OH-	NA+	-0.006 (48)
CL-	OH-	K+	-0.006 (48)
CL-	OH-	CA+2	-0.025 (48)
CL-	HC03-	NA+	-0.015 (48)
CL-	HC03-	MG+2	-0.096 (48)
CL-	C03-2	NA+	0.0085 (48)
CL-	C03-2	K+	0.004 (48)
S04-2	HS04-	NA+	-0.0094 (48)
S04-2	HS04-	K+	-0.0677 (48)
S04-2	HS04-	MG+2	-0.0425 (48)
S04-2	OH-	NA+	-0.009 (48)
S04-2	OH-	K+	-0.050 (48)
S04-2	HC03-	NA+	-0.005 (48)
S04-2	HC03-	MG+2	-0.161 (48)
S04-2	C03-2	NA+	-0.005 (48)
S04-2	C03-2	K+	-0.009 (48)
OH-	C03-2	NA+	-0.017 (48)
OH-	C03-2	K+	-0.01 (48)
OH-	BR-	NA+	-0.018 (88)
OH-	BR-	K+	-0.014 (88)
HC03-	C03-2	NA+	0.002 (48)
HC03-	C03-2	K+	0.012 (48)

Attachment D. -- Listing of the source code to program PHRQPITZ

	PROGRAM PHRQPITZ	00010
C	*****	00020
C		00030
C	PROGRAM PHRQPITZ	00040
C		00050
C	L.N. PLUMMER, D.L. PARKHURST, G.W. FLEMING, AND S.A. DUNKLE	00060
C		00070
C	VERSION CURRENT - NOVEMBER, 1988	00080
C		00090
C	*****	00100
C		00110
C		00120
	\$INSERT COMMON.BLOCKS	00130
	OPEN (UNIT=11,FILE='PHRQPITZ.DATA',STATUS='OLD')	00140
	MAXT=30	00150
	MAXS=250	00160
	MAXT1=31	00170
	MAXM=20	00180
	MAXEQ=50	00190
	DO 10 I=1,MAXS	00200
	NSP(I)=0	00210
10	CONTINUE	00220
	IFILE=11	00230
	CALL RDATA(IFILE)	00240
	CALL INITPZ	00250
20	CONTINUE	00260
C	SINGLE SIMULATION LOOP	00270
	CALL READ	00280
	CLOSE (UNIT=11)	00290
	ISTEP=0	00300
C	SOLVE INITIAL SOLUTIONS	00310
	DO 40 ISOL=1,2	00320
	IF (ISOLV(ISOL).EQ.0) GO TO 40	00330
	JSOL=ISOL	00340
	CALL SOLN(JSOL)	00350
	IESPEC=IOPT(2)	00360
	IF (IOPT(2).GE.2) IESPEC=0	00370
	CALL PICK1	00380
	CALL PTOT	00390
	CALL KTEMP	00400
	CALL MODEL	00410
	IF (IOPT(2).LT.2) GO TO 30	00420
	IESPEC=LNEG	00430
	IF (ELECT.LT.0D0) IESPEC=LPOS	00440
	CALL PICK1	00450
	CALL KTEMP	00460
	CALL MODEL	00470
	TOTAL(ISOL,IESPEC)=TOT(IESPEC)	00480
30	CONTINUE	00490
	IF (IASPEC.GT.0) TOTAL(ISOL,IASPEC)=TOT(IASPEC)	00500
	CALL THORIT(T)	00510
	TH(ISOL)=T	00520
	THOR=T	00530
	DIFFZ(ISOL)=ELECT	00540

	CALL PSPEC	00550
	CALL PLOOK	00560
40	CONTINUE	00570
	IF (IOPT(3).EQ.0) GO TO 20	00580
C	TAKE REACTION STEPS	00590
	DO 50 ISTEP=1,MAX0(1,NSTEPS)	00600
	CALL STEP	00610
	CALL PTOT	00620
	CALL PICK2	00630
	CALL KTEMP	00640
	CALL KMINO	00650
	CALL MODEL	00660
	CALL PPHASE	00670
	CALL PLOOK	00680
	CALL PTOT	00690
	CALL PSPEC	00700
50	CONTINUE	00710
	CALL SAVE	00720
	GO TO 20	00730
	END	00740
C	-----	00750
	BLOCK DATA	00760
C	-----	00770
C		00780
C	SEE SUBROUTINE PITZER	00790
C		00800
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	00810
	INTEGER TRANS(40),IN(20)	00820
	LOGICAL LELEM(4:30)	00830
	DOUBLE PRECISION LAM(40,20),BCX(4,20,21:40)	00840
	DOUBLE PRECISION PSI(40,40,40),BC(4,20,21:40,5),THETA(40,40)	00850
	COMMON / MX1 / BC	00860
	COMMON / MX3 / BCX,OTEMP	00870
	COMMON / MX6 / LAM,TRANS,PSI,IN	00880
	COMMON / MX7 / THETA	00890
	COMMON / MX8 / AK,BK,DK	00900
	COMMON / MX9 / ALPHA	00910
	COMMON / MX0 / B	00920
	COMMON / PI1C / SPECS,NEUTRL	00930
	COMMON / PI1 / M1,M2,M3	00940
	COMMON / PEL / LELEM	00950
	CHARACTER *8 SPECS(40),NEUTRL(20)	00960
	DIMENSION ALPHA(5),Z(40)	00970
	DIMENSION AK(0:20,2),BK(0:22),DK(0:22)	00980
C		00990
C	LINES 01020-01300 FROM HAAR, GALLAGHER, AND KELL (1984)	01000
C		01010
	COMMON /ACONST/ GASCON,TZ,AA,ZB,DZB,YB	01020
	COMMON /NCONST/ G(40),II(40),JJ(40),NC	01030
	COMMON /ELLCON/ G1,G2,GF,B1,B2,B1T,B2T,B1TT,B2TT	01040
	COMMON /BCONST/ P(10),Q(10)	01050
	COMMON /ADDCON/ ATZ(4),ADZ(4),AAT(4),AAD(4)	01060
	DATA ATZ/2*64.D1,641.6D0,27.D1/,ADZ/3*.319D0,1.55D0/,AAT/2*2.D4	01070
	1,4.D4,25.D0/,AAD/34.D0,4.D1,3.D1,1.05D3/	01080
	DATA GASCON/.461522D0/,TZ/647.073D0/,AA/1.D0/,NC/36/	01090
	DATA G1,G2,GF/11.D0,44.3333333333333D0,3.5D0/	01100

```

DATA P/.7478629D0,-.3540782D0,2*0.D0,.007159876D0,0.D0,-.003528426 01110
1D0,3*0.D0/ 01120
DATA Q/1.1278334D0,0.D0,-.5944001D0,-5.010996D0,0.D0,.63684256D0, 01130
1 4*0.D0/ 01140
DATA G/- .53062968529023D3,.22744901424408D4,.78779333020687D3 01150
1,-.69830527374994D2,.17863832875422D5,-.39514731563338D5 01160
2,.33803884280753D5,-.13855050202703D5,-.25637436613260D6 01170
3,.48212575981415D6,-.34183016969660D6,.12223156417448D6 01180
4,.11797433655832D7,-.21734810110373D7,.10829952168620D7 01190
5,-.25441998064049D6,-.31377774947767D7,.52911910757704D7 01200
6,-.13802577177877D7,-.25109914369001D6,.46561826115608D7 01210
7,-.72752773275387D7,.41774246148294D6,.14016358244614D7 01220
8,-.31555231392127D7,.47929666384584D7,.40912664781209D6 01230
9,-.13626369388386D7,.69625220862664D6,-.10834900096447D7 01240
A,-.22722827401688D6,.38365486000660D6,.68833257944332D4 01250
B,.21757245522644D5,-.26627944829770D4,-.70730418082074D5 01260
C,-.225D0,-1.68D0,.055D0,-93.0D0/ 01270
DATA II/4*0,4*1,4*2,4*3,4*4,4*5,4*6,4*8,2*2,0,4,3*2,4/ 01280
DATA JJ/2,3,5,7,2,3,5,7,2,3,5,7,2,3,5,7,2,3,5,7,2,3,5,7 01290
1,2,3,5,7,1,3*4,0,2,0,0/ 01300
DATA M1,M3,M2 / 2*0,20 /, LELEM / 27*.FALSE. / 01310
DATA BK,DK / 46*0.0D0 / 01320
DATA BC,LAM,THETA / 10400*0.0D0 / 01330
DATA BCX / 1600*0.0 /, OTEMP / 0.0 / 01340
DATA B / 1.2D0 / 01350

```

```

C 01360
C AK IS USED TO CALCULATE HIGHER ORDER ELECTROSTATIC TERMS IN 01370
C SUBROUTINE PITZER 01380
C 01390

```

```

DATA AK / 1.925154014814667,-.060076477753119,-.029779077456514 01400
1, -.007299499690937,0.000388260636404,0.000636874599598 01410
2, 0.000036583601823,-.000045036975204,-.000004537895710 01420
3, 0.000002937706971,0.000000396566462,-.000000202099617 01430
4, -.0000000025267769,0.000000013522610,0.000000001229405 01440
5, -.0000000000821969,-.0000000000050847,0.000000000046333 01450
6, 0.000000000001943,-.0000000000002563,-.000000000010991 01460
7, 0.628023320520852,0.462762985338493,0.150044637187895 01470
8, -.028796057604906,-.036552745910311,-.001668087945272 01480
9, 0.006519840398744,0.001130378079086,-.000887171310131 01490
A, -.000242107641309,0.000087294451594,0.000034682122751 01500
B, -.000004583768938,-.000003548684306,-.000000250453880 01510
C, 0.0000000216991779,0.000000080779570,0.000000004558555 01520
D, -.0000000006944757,-.000000002849257,0.000000000237816/ 01530

```

```

DATA ALPHA / 2.0D0, 1.4D0, 12.0D0, 2.0D0, 50.0D0 / 01540
END 01550

```

```

SUBROUTINE MODEL 01560

```

```

$INSERT COMMON.BLOCKS 01570

```

```

CALL SET 01580

```

```

ITER=1 01590

```

```

ITER1=0 01600

```

```

10 CONTINUE 01610

```

```

ITER1=ITER1+1 01620

```

```

IF (IOPT(8).EQ.1) WRITE (6,25) ITER,ITER1 01630

```

```

25 FORMAT('/ BEGINNING OF ITERATION ',I4,' SUB-ITERATION ',I4,') 01640

```

```

DO 20 N=1,LASTS 01650

```

```

LG(N)=0.0D0 01660

```

20	CONTINUE	01670
	CALL PITZER(M, LG, AW, MU, IOPT(10))	01680
	TOT(3)=AW	01690
	LA(3)=DLOG10(AW)	01700
	IF (IOPT(8).EQ.1) WRITE (6,15) MU,AW	01710
15	FORMAT (' AFTER SUBROUTINE PITZER ',	01720
1	' MU =',1PD15.7,' AW',0PF15.7)	01730
	CALL ZEROAR	01740
	CALL AQMOD	01750
	CALL CHECK(ISTAT)	01760
	IF (ISTAT.EQ.0) THEN	01770
	RETURN	01780
	ELSE IF (ISTAT.EQ.1) THEN	01790
	KMIN=NMIN	01800
	CALL SOLVE	01810
	CALL RESET	01820
	ITER=ITER+1	01830
	ITER1=0	01840
	GO TO 10	01850
	ELSE IF (ISTAT.EQ.2) THEN	01860
	CALL RESET	01870
	GO TO 10	01880
	END IF	01890
	STOP 'SUBROUTINE MODEL, UNKNOWN STATUS'	01900
	END	01910
	SUBROUTINE SET	01920
\$	INSERT COMMON.BLOCKS	01930
	LA(1)=DLOG10(TOT(1))	01940
	LA(2)=DLOG10(TOT(2))	01950
	LA(3)=0.0D0	01960
	PH=-LA(1)	01970
	IF(LA(1).GT.-1.0D0) LA(1)=-1.0D0	01980
	M(1)=1D1**(LA(1))	01990
	LG(1)=0.0	02000
	DO 20 I=4,MAXS	02010
	M(I)=0.0D0	02020
	LG(I)=0.0	02030
20	CONTINUE	02040
	MU=0D0	02050
	DO 10 I=4,LASTT	02060
	IF (IIN(I).LE.0) GO TO 10	02070
	IF (TOT(I).LE.0.0D0) TOT(I)=1D-10	02080
	IF (I.EQ.IASPEC) THEN	02090
C	SET CARBON SPECIES	02100
	M(15)=TOT(I)*1D-3	02110
	IF (PH.LT.6.3) M(35)=TOT(I)*0.8	02120
	IF (PH.LT.10.3.AND.PH.GE.6.3) THEN	02130
	M(34)=TOT(I)*0.8	02140
	MU=MU+0.5*M(34)	02150
	END IF	02160
	IF (PH.GE.10.3) M(15)=TOT(I)*0.8	02170
	ELSE	02180
	M(I)=TOT(I)	02190
	END IF	02200
	MU=MU+M(I)*ZSP(I)*ZSP(I)*0.5D0	02210
	LA(I)=DLOG10(M(I))	02220

10	CONTINUE	02230
	RETURN	02240
	END	02250
	SUBROUTINE GAMMA	02260
\$	INSERT COMMON.BLOCKS	02270
C		02280
C	THIS SUBROUTINE IS NOT USED	02290
C		02300
	SAVE DSUM	02310
C	CALCULATE IONIC STRENGTH	02320
	DOLDMU=MU	02330
	SUM=M(1)	02340
	MU=M(1)	02350
	DO 10 I=4, LASTS	02360
	IF (SFLAG(I).EQ.0) GO TO 10	02370
	SUM=SUM+M(I)	02380
	MU=MU+M(I)*ZSP(I)*ZSP(I)	02390
10	CONTINUE	02400
	IF (ITER.EQ.1) GO TO 20	02410
	IF (SUM-DSUM.LT.1D0) GO TO 20	02420
	SUM=DSUM+1D0	02430
20	CONTINUE	02440
	DSUM=SUM	02450
	MU=MU*0.5	02460
	MU=DMIN1(MU, 1D1)	02470
	IF (ITER.EQ.1) GO TO 30	02480
	IF (DABS(DOLDMU-MU).LE.0.75D0) GO TO 30	02490
	MU=DOLDMU+((MU-DOLDMU)/DABS(MU-DOLDMU))*0.5D0	02500
30	CONTINUE	02510
	MUHALF=DSQRT(MU)	02520
C	ACTIVITY OF WATER	02530
	AH20=1.0-SUM*0.017	02540
	IF (SUM.GT.40.0D0) AH20=0.32	02550
	LA(3)=DLOG10(AH20)	02560
	TOT(3)=AH20	02570
C	CALCULATE ACTIVITY COEFFICIENTS	02580
	AMU=-A*MUHALF	02590
	BMU=B*MUHALF	02600
	CMU=-A*(MUHALF/(1.0+MUHALF)-0.3*MU)	02610
	ZCHRG=0.1*MU	02620
	LG(1)=AMU/(1.0+DHA(1)*BMU)	02630
	IF (IOPT(6).EQ.1) LG(1)=CMU	02640
	LG(2)=0.0D0	02650
	LG(3)=0.0D0	02660
	DO 70 I=4, LASTS	02670
	IF (SFLAG(I).EQ.0) GO TO 70	02680
	IF (DABS(ZSP(I)).LE.1.0D-40) GO TO 40	02690
	IF (GFLAG(I).EQ.1) GO TO 50	02700
	IF (DHA(I).LE.0D0) GOTO 60	02710
	IF (IOPT(6).EQ.1) GO TO 60	02720
	LG(I)=AMU*ZSP(I)*ZSP(I)/(1.0+DHA(I)*BMU)	02730
	GO TO 70	02740
40	LG(I)=ZCHRG	02750
	GO TO 70	02760
50	LG(I)=AMU*ZSP(I)*ZSP(I)/(1.0+ADHSP(I,1)*BMU)+ADHSP(I,2)*MU	02770
	GO TO 70	02780

60	LG(I)=CMU*ZSP(I)*ZSP(I)	02790
70	CONTINUE	02800
C	LG FOR CARBONIC ACID	02810
	RETURN	02820
	END	02830
	SUBROUTINE AQMOD	02840
\$	INSERT COMMON.BLOCKS	02850
	DIMENSION OLDG(250),GAMMA(250),DELGAM(250)	02860
	COMMON /GBLOCK/ OLDG,OLDMU,OLDAW,GAMMA,DELGAM	02870
	DIMENSION DM2(30)	02880
C		02890
C	INITIALIZE VARIABLES FOR THIS ITERATION	02900
C		02910
	ELECT=0.0D0	02920
	THSOLN=0.0D0	02930
	DALKS=0D0	02940
	IF(ITER.LE.1) THEN	02950
	OLDMU=0.0D0	02960
	DO 7 I=1,LASTS	02970
	OLDG(I)=0.99D0	02980
7	CONTINUE	02990
	END IF	03000
	DMU=MU-OLDMU	03010
	IF(DABS(DMU).LE.1.0D-40) DMU=1.0D0	03020
C	SET GAMMA ARRAY	03030
	DO 8 I=1,LASTS	03040
	GAMMA(I)=1D1**LG(I)	03050
8	CONTINUE	03060
	SUM=M(1)	03070
C	CALCULATE DELTA ACTIVITY OF WATER PER MOLE SOLUTE	03080
	DO 9 I=1,LASTS	03090
	IF (SFLAG(I).EQ.0) GO TO 9	03100
	SUM=SUM+M(I)	03110
9	CONTINUE	03120
	DAW=-(1D0-TOT(3))/SUM	03130
	IF (IOPT(8).GT.0) WRITE(6,4) DAW	03140
4	FORMAT(' DELTA ACTIVITY WATER PER MOLE SOLUTE ',1PD12.4)	03150
	DAW=DAW/TOT(3)	03160
C	INITIALIZE ACTIVITY OF WATER AND IONIC STRENGTH ARRAYS	03170
	DO 2 I=1,50	03180
	AR(50,I)=0D0	03190
	AR(30,I)=0D0	03200
	AR(I,30)=0D0	03210
2	CONTINUE	03220
C		03230
C	SET ACTIVITY OF MASTER SPECIES USING MOLALITY AND GAMMA	03240
C		03250
	IF(IESPEC.NE.1) M(1)=1D1**(LA(1)-LG(1))	03260
C	HOLD HYDROXIDE CONSTANT FOR ENTIRE ITERATION IF PH > 12	03270
	IF(M(1).GT.1D-12.AND.M(31).LT.5D-3) THEN	03280
	LA(1)=DLOG10(M(1))+LG(1)	03290
	LM(31)=LKSP(31)+LA(3)-LG(31)-LA(1)	03300
	M(31)=1D1**LM(31)	03310
	ELSE	03320
	IF(IESPEC.EQ.1) THEN	03330
	LA(1)=LKSP(31)+LA(3)-LM(31)-LG(31)	03340

	LM(1)=LA(1)-LG(1)	03350
	M(1)=1D1**LM(1)	03360
	END IF	03370
	END IF	03380
	DO 5 I=4,30	03390
	IF(SFLAG(I).GT.0) LA(I)=DLOG10(M(I))+LG(I)	03400
	5 CONTINUE	03410
C		03420
C	RECALCULATE MOLALITIES OF MASTER SPECIES TO OBTAIN	03430
C	MASS BALANCE. UPDATE CATIONS THEN ANIONS.	03440
C		03450
	DO 80 IZ=1,2	03460
	DO 10 I=1,MAXT	03470
	DM2(I)=OD0	03480
	10 CONTINUE	03490
C	CALCULATE MOLALITIES WITH NEW GAMMAS, NEW ACTIVITIES.	03500
	DO 40 I=1,LASTS	03510
	IF (SFLAG(I).EQ.0) GO TO 40	03520
	K=NSP(I)	03530
	DM=LKSP(I)-LG(I)	03540
	DO 20 J=1,K	03550
	DM=DM+LA(LSP(I,J))*CSP(I,J)	03560
	20 CONTINUE	03570
	IF (I.NE.31) THEN	03580
	LM(I)=DM	03590
	M(I)=UNDER(DM)	03600
	END IF	03610
	IF (DABS(M(I)).LE.1.0D-40) GO TO 40	03620
C	SUM MOLALITIES FOR EACH ELEMENT	03630
	DO 30 I1=1,K	03640
	DM=M(I)*CSP(I,I1)	03650
	K2=LSP(I,I1)	03660
	IF (K2.LT.4) GO TO 30	03670
	AR(K2,K2)=AR(K2,K2)+DM	03680
	30 CONTINUE	03690
	40 CONTINUE	03700
C	CALCULATE NEW MOLALITIES OF MASTER SPECIES	03710
	DO 70 K=4,LASTT	03720
	IF (IIN(K).EQ.0) GO TO 70	03730
	IF (IESPEC.EQ.K) GO TO 60	03740
	IF (IASPEC.EQ.K) GO TO 60	03750
	IF (ZSP(K).LE.0.0.AND.IZ.EQ.1) GO TO 60	03760
	IF (ZSP(K).GT.0.0.AND.IZ.EQ.2) GO TO 60	03770
	IF (AR(K,K).LE.0D0) AR(K,K)=1D-35	03780
	LM(K)=LM(K)+(DLOG10(TOT(K))-DLOG10(AR(K,K)))	03790
	LA(K)=LM(K)+LG(K)	03800
	60 CONTINUE	03810
	AR(K,K)=0D0	03820
	70 CONTINUE	03830
	80 CONTINUE	03840
	ZSP(2)=0.0D0	03850
C		03860
C	SET UP ARRAYS FOR SUBROUTINE SOLVE	03870
C		03880
	DO 120 I=1,LASTS	03890
	IF (SFLAG(I).EQ.0) GO TO 120	03900

	IF (I.EQ.2.OR.I.EQ.3) GO TO 120	03910
	K=NSP(I)	03920
	DM=LKSP(I)-LG(I)	03930
C	CALCULATE MOLALITY	03940
	DO 90 J=1,K	03950
	DM=DM+LA(LSP(I,J))*CSP(I,J)	03960
90	CONTINUE	03970
	IF (I.NE.31) THEN	03980
	LM(I)=DM	03990
	M(I)=UNDER(DM)	04000
	END IF	04010
	IF (DABS(M(I)).LE.1.0D-40) GO TO 120	04020
C	CALCULATE EFFECT OF IONIC STRENGTH ON GAMMA	04030
	DGI=ODO	04040
	IF(DABS(DMU).GT.1D-11)	04050
1	DGI=(GAMMA(I)-OLDG(I))/DMU/GAMMA(I)*M(I)	04060
	DELGAM(I)=(GAMMA(I)-OLDG(I))/DMU	04070
C	ELECT, THOR, ALK, MU, AND H2O EQNS.	04080
	AR(1,30)=AR(1,30)-DGI*ZSP(I)	04090
	AR(2,30)=AR(2,30)-DGI*THSP(I)	04100
	AR(3,30)=AR(3,30)-DGI*ALKSP(I)	04110
	AR(30,30)=AR(30,30)-DGI*0.5D0*ZSP(I)*ZSP(I)	04120
	AR(50,30)=AR(50,30)-DGI*DAW	04130
C	MORE TERMS FOR MATRIX	04140
	DO 110 I1=1,K	04150
	DM=M(I)*CSP(I,I1)	04160
	K2=LSP(I,I1)	04170
	DG=ODO	04180
	IF(DABS(DMU).GT.1D-11)	04190
1	DG=(GAMMA(K2)-OLDG(K2))/DMU/GAMMA(K2)*DM	04200
C	EFFECT ON MOLALITY FROM CHANGE IN IONIC STRENGTH	04210
	IF(K2.GT.3) AR(K2,30)=AR(K2,30)-DGI*CSP(I,I1)	04220
C	MASS BALANCE	04230
	DO 100 J1=1,K	04240
	K1=LSP(I,J1)	04250
	IF (K1.LT.4) GO TO 100	04260
	AR(K1,K2)=AR(K1,K2)+DM*CSP(I,J1)	04270
	AR(K1,30)=AR(K1,30)+DG*CSP(I,J1)	04280
100	CONTINUE	04290
C	IONIC STRENGTH	04300
	AR(30,K2)=AR(30,K2)+DM*0.5D0*ZSP(I)*ZSP(I)	04310
	AR(30,30)=AR(30,30)+DG*0.5D0*ZSP(I)*ZSP(I)	04320
C	ELECTRICAL BALANCE	04330
	AR(1,K2)=AR(1,K2)+DM*ZSP(I)	04340
	AR(1,30)=AR(1,30)+DG*ZSP(I)	04350
C	THOR BALANCE	04360
	AR(2,K2)=AR(2,K2)+DM*THSP(I)	04370
	AR(2,30)=AR(2,30)+DG*THSP(I)	04380
C	ALKALINITY EQUATION	04390
	AR(3,K2)=AR(3,K2)+DM*ALKSP(I)	04400
	AR(3,30)=AR(3,30)+DG*ALKSP(I)	04410
C	H2O DERIVATIVE	04420
	AR(50,K2)=AR(50,K2)+DM*DAW	04430
	AR(50,30)=AR(50,30)+DG*DAW	04440
110	CONTINUE	04450
	THSOLN=THSOLN+M(I)*THSP(I)	04460

	ELECT=ELECT+M(I)*ZSP(I)	04470
	DALKS=DALKS+M(I)*ALKSP(I)	04480
120	CONTINUE	04490
	ZSP(2)=-1.0	04500
C		04510
C	H2O EQUATION	04520
C		04530
	AR(50,3)=AR(50,3)-1.0DO	04540
	AR(30,30)=AR(30,30)-1.0DO	04550
C	*****	04560
	ENTRY SI	04570
C	*****	04580
	IF (NMINs.NE.0) THEN	04590
	DO 140 I=1,NMINs	04600
	J=MAXT+I	04610
	K=NMIN(I)	04620
C	CALCULATE INVERSE SATURATION INDEX	04630
	DSI=LKMIN(I)	04640
	DO 130 I1=1,K	04650
	K1=LMIN(I,I1)	04660
	IF(DABS(DMU).GT.1D-11) THEN	04670
	AR(J,30)=AR(J,30)+(GAMMA(K1)-OLDG(K1))/DMU/GAMMA(K1)	04680
1	*CMIN(I,I1)	04690
	DG=(GAMMA(K1)-OLDG(K1))/DMU/GAMMA(K1)*CMIN(I,I1)	04700
	END IF	04710
	DSI=DSI-LA(LMIN(I,I1))*CMIN(I,I1)	04720
130	CONTINUE	04730
	CR(J)=DSI*DLOG(1D1)	04740
140	CONTINUE	04750
	END IF	04760
	IF(IOPT(8).GT.0) THEN	04770
	WRITE(6,107)	04780
107	FORMAT(/' I SPECIES DEL MU GAMMA ,	04790
1	' OLD GAMMA ',' MOLALITY ')	04800
	DO 105 I=1,LASTS	04810
	IF (SFLAG(I).EQ.0) GO TO 105	04820
	D=(GAMMA(I)-OLDG(I))/DMU/GAMMA(I)	04830
	DM=D*M(I)	04840
	WRITE(6,106) I,SNAME(I),DMU,GAMMA(I),OLDG(I),M(I)	04850
106	FORMAT(I4,1X,A8,1PD10.3,1P2D13.4,3D10.3)	04860
105	CONTINUE	04870
	END IF	04880
	RETURN	04890
C		04900
	END	04910
	SUBROUTINE CHECK(ISTAT)	04920
\$	INSERT COMMON.BLOCKS	04930
	DIMENSION OLDG(250),GAMMA(250),DELGAM(250)	04940
	COMMON /GBLOCK/ OLDG,OLDMU,OLDAW,GAMMA,DELGAM	04950
	DOUBLE PRECISION MCLOSE	04960
	SAVE OLDMU1	04970
	DATA ECLOSE,TCLOSE,THCLOS,MCLOSE /5D-9,1D-7,1D-7,1D-5/	04980
	DATA CHKMU /1D-3/	04990
	ICHECK=0	05000
	JCHECK=0	05010
C		05020

C	CALCULATE RESIDUAL ERROR	05030
C		05040
	CR(1)=DZOFF-ELECT	05050
	CR(2)=THOR-THSOLN	05060
	CR(3)=DALKT-DALKS	05070
C	CALCULATE MASS BALANCE	05080
	THTOT=DABS(THOR)	05090
	DO 10 I=4, LASTT	05100
	IF (IIN(I).LE.0) GO TO 10	05110
	CR(I)=TOT(I)-AR(I,I)	05120
	THTOT=DABS(THSP(I))*TOT(I)+THTOT	05130
	10 CONTINUE	05140
C		05150
C	CHECK IONIC STRENGTH	05160
C		05170
	IF(ITER1.GT.1) THEN	05180
	IF(DABS(MU-OLDMU1).GT.CHKMU) THEN	05190
	JCHECK=1	05200
	IF(IOPT(8).EQ.1) WRITE(6,18) DABS(MU-OLDMU1),CHKMU	05210
18	FORMAT(' FAILED IONIC STRENGTH CHECK. MU-ERROR ',1PD11.3,	05220
1	' TOLER ',1PD11.3)	05230
	END IF	05240
	END IF	05250
	OLDMU1=MU	05260
C		05270
C	CHECK ELECTRICAL BALANCE	05280
C		05290
	DCHECK=1D1**(DABS(PH-7DO))*ECLOSE	05300
	DCHECK=DMIN1(DCHECK,1D-5)	05310
	IF (IESPEC.GT.0.AND.DABS(CR(1)).GT.DCHECK) THEN	05320
	ICHECK=1	05330
	IF (IOPT(8).EQ.1) THEN	05340
	WRITE (6,12) IESPEC,CR(1),DCHECK	05350
12	FORMAT(1X,'FAILED ELECTRICAL BALANCE. '	05360
1	' E-SPECIES',I3,' E-ERROR',1PD11.3,' TOLER',D11.3)	05370
	END IF	05380
	END IF	05390
C		05400
C	CHECK THOR BALANCE	05410
C		05420
	IF (IIN(2).GT.0.AND.DABS(CR(2)).GT.THCLOS*THTOT) THEN	05430
	ICHECK=1	05440
	IF (IOPT(8).EQ.1) THEN	05450
	WRITE (6,13) CR(2),THCLOS*THTOT	05460
13	FORMAT(1X,'FAILED THOR BALANCE. '	05470
1	' THOR ERROR',1PD11.3,' THOR TOLERANCE',D11.3)	05480
	END IF	05490
	END IF	05500
C		05510
C	CHECK ALKALINITY	05520
C		05530
	IF (IASPEC.GT.0) THEN	05540
	IF(DABS(CR(3)).GT.TCLOSE*DALKT) THEN	05550
	ICHECK=1	05560
	IF (IOPT(8).EQ.1) THEN	05570
	WRITE (6,14) IASPEC,CR(3),TCLOSE*DALKT	05580

14	FORMAT(1X,'FAILED ALKALINITY BALANCE. '	05590
1	' A-SPECIES',I3,' A-ERROR',1PD11.3,' TOLER',D11.3)	05600
	END IF	05610
	END IF	05620
	END IF	05630
C		05640
C	CHECK MASS BALANCES	05650
C		05660
	DO 20 I=4, LASTT	05670
	IF (IIN(I).GT.0.AND.I.NE.IESPEC.AND.I.NE.IASPEC) THEN	05680
	IF (DABS(CR(I)).GT.TCLOSE*TOT(I)) THEN	05690
	ICHECK=1	05700
	IF(IOPT(8).EQ.1) WRITE (6,25) I	05710
25	FORMAT(' FAILED MASS BALANCE. SPECIES',I4)	05720
	IF(DABS(CR(I)).GT.1D-4*TOT(I)) THEN	05730
	JCHECK=1	05740
	IF(IOPT(8).EQ.1) WRITE(6,16)	05750
16	FORMAT(' MASS-BALANCE-ONLY FLAG SET.')	05760
	END IF	05770
	END IF	05780
	END IF	05790
	20 CONTINUE	05800
C		05810
C	CHECK MINERAL EQUILIBRIA	05820
C		05830
	IF (NMIN.SGT.0) THEN	05840
	DO 30 I=1,NMIN.S	05850
	K=MAXT+I	05860
	IF (DABS(CR(K)).GT.MCLOSE) THEN	05870
	ICHECK=1	05880
	IF (IOPT(8).GT.0) THEN	05890
	WRITE(6,35) MNAME(I),CR(K),MCLOSE	05900
35	FORMAT(' FAILED SATURATION TEST. ',A12,' SI-ERROR ',	05910
1	1PD11.3,' SI-TOLER ',1PD10.2)	05920
	END IF	05930
	END IF	05940
	30 CONTINUE	05950
	END IF	05960
C		05970
C	STOP ON MAX ITERATIONS	05980
C		05990
	IF (ITER.GT.100.OR.ITER1.GT.50) THEN	06000
	WRITE (6,70) ITER, ITER1	06010
70	FORMAT(1H1,80('*'))/' CALCULATIONS TERMINATED ',/,	06020
1	' ITER ',I5,' ITER1 ',I5)	06030
	F=0.0D0	06040
	CALL PBUG(F,F)	06050
	CALL PTOT	06060
	CALL PSPEC	06070
	CALL PPHASE	06080
	CALL PLOOK	06090
	ENDFILE (UNIT=6)	06100
	STOP 'SUBROUTINE CHECK, MAXIMUM ITERATIONS'	06110
	END IF	06120
C		06130
C	SET ISTAT TO BRANCH ON CONVERGENCE CRITERIA	06140

C	DO 50 I=1,NEQ+1	06150
	DELTA(I)=ODO	06160
50	CONTINUE	06170
	IF (JCHECK.GT.0.OR.ITER1.LE.1) THEN	06180
C	POOR MASS BALANCE, BACK TO AQMOD	06190
	ISTAT=2	06200
	RETURN	06210
	ELSE IF (ICHECK.GT.0.OR.ITER.LT.5) THEN	06220
C	GOOD MASS BALANCE, GO TO SOLVE	06230
	DO 107 I=1,LASTS	06240
	OLDG(I)=GAMMA(I)	06250
107	CONTINUE	06260
	OLDMU=MU	06270
	OLDAW=TOT(3)	06280
	ISTAT=1	06290
	RETURN	06300
	ELSE IF (ICHECK.EQ.0.AND.JCHECK.EQ.0.AND.ITER.GT.4) THEN	06310
C	CONVERGED	06320
	ISTAT=0	06330
	RETURN	06340
	ELSE	06350
	WRITE(6,80) ICHECK, JCHECK, ITER, ITER1	06360
80	FORMAT(' UNKNOWN STATUS IN SUBROUTINE CHECK.')	06370
1	' ICHECK = ',I4,/'	06380
1	' JCHECK = ',I4,/'	06390
1	' ITER = ',I4,/'	06400
1	' ITER1 = ',I4)	06410
	END IF	06420
	STOP 'SUBROUTINE CHECK, UNKNOWN STATUS'	06430
C		06440
	END	06450
	SUBROUTINE SOLVE	06460
\$INSERT	COMMON.BLOCKS	06470
	DIMENSION OLDG(250),GAMMA(250),DELGAM(250)	06480
	COMMON /GBLOCK/ OLDG,OLDMU,OLDAW,GAMMA,DELGAM	06490
	COMMON / FLAGMU / INMU	06500
C		06510
C	COPY ROWS OF AR INTO MATRIX AS WHICH IS INVERTED IN SLNQ	06520
C		06530
	IF (NEQ.EQ.0) RETURN	06540
C		06550
C	SWITCH OUT DELTA MU CALCULATIONS FOR MASTER SPECIES	06560
C		06570
	INMU=0	06580
	DMU=DABS(MU-OLDMU)	06590
C	NO MU EQN IF ELECT > 1 OR DEL MU > 1 OR SI > 0.5	06600
	IF(DABS(CR(1)).GT.1DO.AND.IIN(1).GT.0) INMU=1	06610
	IF(DMU.GT.0.1DO.AND.IIN(1).GT.0) INMU=1	06620
	IF(NMINS.GT.0) THEN	06630
	DO 6 I=1,NMINS	06640
	J=MAXT+I	06650
	IF(DABS(CR(J)).GT.5DO) INMU=1	06660
6	CONTINUE	06670
	END IF	06680
	IF (INMU.GT.0) THEN	06690
		06700

	AR(1,30)=0D0	06710
	AR(3,30)=0D0	06720
	DO 5 I=4,49	06730
	IF(I.NE.30) AR(I,30)=0D0	06740
5	CONTINUE	06750
	AR(30,30)=-1.0D0	06760
	END IF	06770
C		06780
C	PUT ELECTRICAL BALANCE EQUATION IN PLACE OF A MASS BALANCE	06790
C		06800
	LASTTX=LASTT	06810
	LASTT=30	06820
	IIN(30)=ILT+1	06830
	IF (IESPEC.LT.2) GO TO 30	06840
	TOT(IESPEC)=AR(IESPEC,IESPEC)	06850
	DO 10 I=1,LASTT	06860
	IF (IIN(I).EQ.0) GO TO 10	06870
	AR(IESPEC,I)=AR(1,I)	06880
10	CONTINUE	06890
	CR(IESPEC)=-ELECT	06900
30	CONTINUE	06910
C		06920
C	PUT ALKALINITY EQUATION IN PLACE OF MASS BALANCE	06930
C		06940
	IF (IASPEC.EQ.0) GO TO 50	06950
	TOT(IASPEC)=AR(IASPEC,IASPEC)	06960
	DO 40 I=1,LASTT	06970
	AR(IASPEC,I)=AR(3,I)	06980
40	CONTINUE	06990
	CR(IASPEC)=CR(3)	07000
50	CONTINUE	07010
C		07020
C	FIX AH2O EQUATION.	07030
C		07040
	IF(IIN(3).GT.0) THEN	07050
	DO 65 I=1,50	07060
	AR(3,I)=AR(50,I)	07070
65	CONTINUE	07080
	CR(3)=0D0	07090
	END IF	07100
C		07110
C	COPY EQUATIONS INTO ARRAY AS.	07120
C		07130
	DO 80 I=1,LASTT	07140
	IF (IIN(I).EQ.0) GO TO 80	07150
	K1=IIN(I)	07160
	DO 60 J=1,LASTT	07170
	K2=IIN(J)	07180
	IF (K2.EQ.0) GO TO 60	07190
	AS(K1,K2)=AR(I,J)	07200
60	CONTINUE	07210
	AS(K1,NEQ1+1)=CR(I)	07220
	IF(NMINS.EQ.0) GO TO 80	07230
	DO 70 J=1,NMINS	07240
	AS(K1,ILT+1+J)=AR(I,MAXT+J)	07250
70	CONTINUE	07260

80	CONTINUE	07270
C		07280
C	COPY MINERAL EQUATIONS	07290
C		07300
	IF (NMIN.EQ.0) GO TO 120	07310
	DO 110 I=1,NMIN	07320
	K=MAXT+I	07330
	L=ILT+I+1	07340
	DO 90 J=1,LASTT	07350
	IF (IIN(J).LE.0) GO TO 90	07360
	AS(L,IIN(J))=AR(K,J)	07370
90	CONTINUE	07380
	DO 100 J=1,NMIN	07390
	AS(L,ILT+J+1)=AR(K,MAXT+J)	07400
100	CONTINUE	07410
	AS(L,NEQ1+1)=CR(K)	07420
110	CONTINUE	07430
120	CONTINUE	07440
C		07450
C	INVERT MATRIX. (DX/X VALUES IN DELTA)	07460
C		07470
	CALL SLNQ	07480
	LASTT=LASTTX	07490
	IIN(30)=0	07500
	RETURN	07510
	END	07520
	SUBROUTINE SLNQ	07530
	IMPLICIT DOUBLE PRECISION(A-H,O-Z),INTEGER(I-N)	07540
	CHARACTER *12 SNAME,TNAME,MNAME	07550
	DOUBLE PRECISION LM,LA,LG,LKSP,LKMIN,MU	07560
	COMMON /REAL8/ LM(250),E(250),LA(250),LG(250),LKSP(250),	07570
1	TOT(50), DELTA(50),DELTOT(50),AR(50,50),	07580
2	AS(50,50),CR(50),CS(50),LKMIN(20),	07590
3	THOR,ELECT,THSOLN,PH,PE,G,H,MU,TOTAL(2,30),	07600
4	DALKT,DALKS,DIFFZ(2),DZOFF	07610
	INTEGER GFLAG,SFLAG	07620
	COMMON /INT2/ NSP(250),LSP(250,6),KFLAG(250),GFLAG(250),	07630
1	SFLAG(250),LASTT,LASTS,IIN(50),IOUT(50),IFE,ILE,	07640
2	IFTH,ILTH,IFT,ILT,IFM,ILM,NEQ,NEQ1,IESPEC,ISOLV(2),	07650
3	NMIN(20),LMIN(20,10),MFLAG(20),LMCON(20,5),NMCON(20),	07660
4	LMINO(20,10),NMINO(20),IOPT(10),NMIN,NSTEPS,NCOMPS,	07670
5	NELTS,NSPCS,ISTEP,LREAC(30),MAXT,MAXT1,MAXM,MAXEQ,	07680
6	MAXS,NRMIN,ITER,ISOL,IASPEC,IALK(2),IUNITS(2)	07690
C		07700
C	SOLUTION OF N LINEAR EQUATIONS IN N UNKNOWN BY GAUSSIAN	07710
C	ELIMINATION OR DETERMINANT EVALUTAION.	07720
C	A CONTAINS THE MATRIX OF THE COEFFICIENTS AND N INDICATES THE	07730
C	ORDER OF THE MATRIX. IF J EQUALS ZERO, D CONTAINS THE VALUE OF THE	07740
C	DETERMINANT. IF J DOES NOT EQUAL ZERO, X CONTAINS THE N VALUES OF	07750
C	THE UNKNOWNNS.	07760
C		07770
	J=1	07780
	N=NEQ+1	07790
	D=0.0D0	07800
	NP1=N+1	07810
	NM1=N-1	07820

IF (IOPT(9).NE.1) GO TO 20	07830
DO 10 II=1,N	07840
WRITE (6,240) (AS(II,JJ),JJ=1,NP1)	07850
10 CONTINUE	07860
20 CONTINUE	07870
IF (N.EQ.1) GO TO 210	07880
DO 120 I=1,NM1	07890
M=I	07900
K=I+1	07910
B=AS(I,I)	07920
DO 40 L=K,N	07930
IF (DABS(B)-DABS(AS(L,I))) 30,40,40	07940
30 M=L	07950
B=AS(L,I)	07960
40 CONTINUE	07970
IF (B) 50,160,50	07980
50 CC=AS(M,I)	07990
IF (I-M) 60,80,60	08000
60 D=-D	08010
IF (DABS(CC).LE.1.0D-40) GO TO 300	08020
DO 70 L=I,NP1	08030
B=AS(I,L)	08040
AS(I,L)=AS(M,L)/CC	08050
70 AS(M,L)=B	08060
GO TO 100	08070
80 CONTINUE	08080
IF (DABS(CC).LE.1.0D-40) GO TO 300	08090
DO 90 L=I,NP1	08100
90 AS(I,L)=AS(I,L)/CC	08110
100 D=D*CC	08120
DO 110 M=K,N	08130
B=AS(M,I)	08140
DO 110 L=I,NP1	08150
110 AS(M,L)=AS(M,L)-B*AS(I,L)	08160
120 CONTINUE	08170
D=D*AS(N,N)	08180
IF (J) 130,200,130	08190
130 CONTINUE	08200
IF(DABS(AS(N,N)).GE.1.0D-40) GO TO 135	08210
WRITE (6,320)N	08220
320 FORMAT(1X,'AS(N,N)=ODO. N=',I3)	08230
DELTA(N)=ODO	08240
GO TO 145	08250
135 DELTA(N)=AS(N,NP1)/AS(N,N)	08260
145 CONTINUE	08270
K=NM1	08280
140 M=K+1	08290
B=0.0D0	08300
DO 150 L=M,N	08310
150 B=B+AS(K,L)*DELTA(L)	08320
DELTA(K)=AS(K,NP1)-B	08330
K=K-1	08340
IF (K) 160,180,140	08350
160 IF (J) 170,180,170	08360
170 WRITE (6,260)	08370
ENDFILE (UNIT=6)	08380

	STOP 'SINGULAR MATRIX IN SUBROUTINE SLNQ'	08390
180	D=0.0D0	08400
	IF (IOPT(9).LT.1) GO TO 190	08410
	WRITE (6,220)	08420
	WRITE (6,230) (DELTA(I),I=1,N)	08430
	WRITE (6,250)	08440
190	CONTINUE	08450
200	RETURN	08460
210	CONTINUE	08470
	IF (DABS(AS(1,1)).LE.1.0D-40) GO TO 170	08480
	DELTA(1)=AS(1,2)/AS(1,1)	08490
	GO TO 180	08500
300	CONTINUE	08510
	WRITE (6,310)I,L,M,ITER	08520
310	FORMAT(1X,'CC=0',' I=',I3,' L=',I3,' M=',I3,' ITER=',I3)	08530
	ENDFILE (UNIT=6)	08540
	STOP	08550
C		08560
C		08570
C		08580
	220 FORMAT (1X,'DELTA VALUES')	08590
	230 FORMAT (3X,1P6D12.3)	08600
	240 FORMAT (/(3X,1P6D12.3))	08610
	250 FORMAT (1X)	08620
	260 FORMAT (1H1,'THE PHASE RULE HAS BEEN VIOLATED')	08630
	END	08640
	SUBROUTINE RESET	08650
	\$INSERT COMMON.BLOCKS	08660
	DIMENSION OLDG(250),GAMMA(250),DELGAM(250)	08670
	COMMON /GBLOCK/ OLDG,OLDMU,OLDAW,GAMMA,DELGAM	08680
	COMMON / FLAGMU / INMU	08690
C	INITIALIZE	08700
	DMIN=0.6D0	08710
	DMAX=3.0D0	08720
	F=1D0	08730
	DPHPE=1D0	08740
	DO 20 I=1,LASTT	08750
	DELTOT(I)=0.0D0	08760
20	CONTINUE	08770
C		08780
C		08790
C	CALCULATE DELTA TOTALS IN MASSES	08800
	DELMU=DELTA(ILT+1)	08810
	IF (NMIN.GT.0) THEN	08820
	DO 25 I=IFM,ILM	08830
	DELTA(I)=DELTA(I+1)	08840
25	CONTINUE	08850
	DO 40 I=IFT,ILT	08860
	DT=0.0D0	08870
	DO 30 J=IFM,ILM	08880
	K=MAXT+J-ILT	08890
	DT=DT-DELTA(J)*AR(1OUT(I),K)	08900
30	CONTINUE	08910
	DELTOT(I)=DT	08920
40	CONTINUE	08930
C	MAKE SURE DELTA MINERALS IS SMALL	08940

	DO 510 I=IFM,ILM	08950
	IF (DELTA(I).GT.0.5DO) THEN	08960
C	MAXIMUM INCREASE	08970
	DF=0.5DO/DELTA(I)	08980
	F=DMIN1(DF,F)	08990
	ELSE IF (DELTA(I).LT.-0.7DO) THEN	09000
C	MAXIMUM DECREASE	09010
	DF=0.7DO/DABS(DELTA(I))	09020
	F=DMIN1(DF,F)	09030
	END IF	09040
510	CONTINUE	09050
	DO 50 I=IFT,ILT	09060
C	MAXIMUM INCREASE IN ELEMENT CONCENTRATION	09070
	IF (DABS(DELTOT(I)).GT.5.0DO) THEN	09080
	DF=5.0DO/DABS(DELTOT(I))	09090
	F=DMIN1(F,DF)	09100
	END IF	09110
C	MAKE SURE NO NEGATIVE CONCENTRATIONS RESULT	09120
	IF (TOT(IOUT(I))+DELTOT(I).LE.0.0DO) THEN	09130
	DF=-0.9DO*TOT(IOUT(I))/DELTOT(I)	09140
	F=DMIN1(F,DF)	09150
	END IF	09160
50	CONTINUE	09170
	END IF	09180
C		09190
C	ENSURE DECREASE OR INCREASE IN ACTIVITY IS NOT TOO LARGE	09200
C		09210
	DO 55 I=IFT,ILT	09220
	IF (DELTA(I).LT.-0.9DO) THEN	09230
	DF=-0.9DO/DELTA(I)	09240
	F=DMIN1(F,DF)	09250
	END IF	09260
	IF (DELTA(I).GT.3.0DO) THEN	09270
	DF=3.0DO/DELTA(I)	09280
	F=DMIN1(F,DF)	09290
	END IF	09300
55	CONTINUE	09310
C		09320
C	ENSURE DELTA IONIC STRENGTH IS SMALL	09330
C		09340
	DMU=MU+DELMU	09350
	IF (INMU.GT.0) DELMU=0DO	09360
	IF(DELMU.GT.0.5DO) THEN	09370
C	MAXIMUM INCREASE ABSOLUTE	09380
	DF=0.5DO/DELMU	09390
	F=DMIN1(F,DF)	09400
	ELSE IF (DELMU.LE.-MU*.4DO) THEN	09410
C	MAXIMUM DECREASE BY FRACTION	09420
	DF=MU*0.4/DABS(DELMU)	09430
	F=DMIN1(F,DF)	09440
	END IF	09450
	IF (DELMU.LT.-0.4DO) THEN	09460
C	MAXIMUM DECREASE ABSOLUTE	09470
	DF=0.4DO/DABS(DELMU)	09480
	F=DMIN1(F,DF)	09490
	END IF	09500

C		09510
C	ENSURE PH CHANGE REASONABLE	09520
C		09530
	IF (IIN(1).GT.0) THEN	09540
	IF (DELTA(1).LT.-DMIN.OR.DELTA(1).GT.DMAX) THEN	09550
	DF=DABS(DMIN/DELTA(1))	09560
	IF (DELTA(1).GT.ODO) DF=DMAX/DELTA(1)	09570
	DPHPE=DF	09580
	END IF	09590
	END IF	09600
C		09610
C	ENSURE PE CHANGE REASONABLE.	09620
C		09630
	IF (IFTH.GT.0) THEN	09640
	IF (DELTA(IFTH).LT.-DMIN.OR.DELTA(IFTH).GT.DMAX) THEN	09650
	DF=DABS(DMIN/DELTA(IFTH))	09660
	IF (DELTA(IFTH).GT.ODO) DF=DMAX/DELTA(IFTH)	09670
	DPHPE=DMIN1(DF,DPHPE)	09680
	END IF	09690
	END IF	09700
C		09710
C	ENSURE ALKALINITY CHANGE IS REASONABLE.	09720
C		09730
	IF (IASPEC.GT.0.AND.IOPT(2).NE.1) THEN	09740
	I=IIN(IASPEC)	09750
	IF (DELTA(I).LT.-DMIN.OR.DELTA(I).GT.DMAX) THEN	09760
	DF=DABS(DMIN/DELTA(I))	09770
	IF (DELTA(I).GT.ODO)DF=DMAX/DELTA(I)	09780
	DPHPE=DMIN1(DF,DPHPE)	09790
	END IF	09800
	END IF	09810
	F=DMIN1(DPHPE,F)	09820
	DPHPE=F	09830
C		09840
C	SCALE DELTAS	09850
C		09860
	IF(DABS(1DO-F).GT.1D-8) THEN	09870
	DO 95 I=IFT,ILT	09880
	DELTA(I)=DELTA(I)*F	09890
95	CONTINUE	09900
	DELMU=DELMU*F	09910
	END IF	09920
	IF(DABS(1DO-F).GT.1D-8.AND.NMINS.GT.0) THEN	09930
	DO 90 I=IFT,ILT	09940
	DELTOT(I)=DELTOT(I)*F	09950
90	CONTINUE	09960
	DO 100 I=IFM,ILM	09970
	DELTA(I)=DELTA(I)*F	09980
100	CONTINUE	09990
	END IF	10000
	IF (IIN(1).GT.0) DELTA(1)=DELTA(1)*DPHPE	10010
	IF (IIN(2).GT.0) DELTA(IIN(2))=DELTA(IIN(2))*DPHPE	10020
C		10030
C	CALL DEBUG PRINT	10040
C		10050
	IF (IOPT(8).GT.0) WRITE(6,97) DELMU	10060

97	FORMAT(' CALCULATED DELTA MU, SCALED = ',1PD10.2)	10070
	IF (ITER.GT.190.OR.IOPT(8).EQ.1.OR.ITER1.GT.40) CALL PBUG(F,DPHPE)	10080
C		10090
C	RESET ELECTRICAL BALANCE SPECIES	10100
C		10110
	IF (IESPEC.GT.0) THEN	10120
	D=DELTA(IIN(IESPEC))	10130
	LM(IESPEC)=LM(IESPEC)+DLOG10(1D0+D)	10140
	M(IESPEC)=1D1**LM(IESPEC)	10150
	DGI=GAMMA(IESPEC)+DELGAM(IESPEC)*DELMU	10160
	IF (DGI.LE.0) THEN	10170
	DGI=LG(IESPEC)	10180
	ELSE	10190
	DGI=DLOG10(DGI)	10200
	END IF	10210
	LA(IESPEC)=LM(IESPEC)+DGI	10220
	TOT(1)=1D1**LA(1)	10230
	PH=-LA(1)	10240
	END IF	10250
C		10260
C	RESET ALKALINITY SPECIES	10270
C		10280
	IF (IASPEC.GT.0) THEN	10290
	D=DELTA(IIN(IASPEC))*DPHPE	10300
	IF (IOPT(2).NE.1) THEN	10310
	LM(IASPEC)=LM(IASPEC)+DLOG10(1D0+D)	10320
	M(IASPEC)=1.0D1**LM(IASPEC)	10330
	DGI=GAMMA(IASPEC)+DELGAM(IASPEC)*DELMU	10340
	IF (DGI.LE.0) THEN	10350
	DGI=LG(IASPEC)	10360
	ELSE	10370
	DGI=DLOG10(DGI)	10380
	END IF	10390
	LA(IASPEC)=LM(IASPEC)+DGI	10400
	END IF	10410
	END IF	10420
C		10430
C	RESET ACTIVITY OF ELECTRON	10440
C		10450
	IF (IFTH.GT.0) THEN	10460
	TOT(2)=TOT(2)*(1D0+DELTA(IFTH))	10470
	LA(2)=DLOG10(TOT(2))	10480
	PE=-LA(2)	10490
	END IF	10500
C		10510
C	ESTIMATE NEW ACTIVITY OF WATER	10520
C		10530
	IF(DELMU.GT.0D0.AND.IIN(3).GT.0.AND.INMU.LE.0) THEN	10540
	DAW=DELTA(IIN(3))*F	10550
	IF(DABS(DAW).GT.0D0) THEN	10560
	LA(3)=LA(3)+DLOG10(1D0+DAW)	10570
	END IF	10580
	END IF	10590
	M(3)=1D1**LA(3)	10600
C		10610
C	RESET MASTER SPECIES	10620

C		10630
	DO 165 I=IFT,ILT	10640
	K=IOUT(I)	10650
	IF(K.EQ.IASPEC) GO TO 165	10660
	IF(K.EQ.IESPEC) GO TO 165	10670
	LM(K)=LM(K)+DLOG10(1D0+DELTA(I))	10680
	M(K)=UNDER(LM(K))	10690
	DGI=GAMMA(K)+DELGAM(K)*DELMU	10700
	IF (DGI.LE.0) THEN	10710
	DGI=LG(K)	10720
	ELSE	10730
	DGI=DLOG10(DGI)	10740
	END IF	10750
	LA(K)=LM(K)+DGI	10760
165	CONTINUE	10770
	DO 122 I=4,LASTS	10780
	IF (SFLAG(I).EQ.0) GO TO 122	10790
	DGI=GAMMA(I)+DELGAM(I)*DELMU	10800
	IF (DGI.LE.0) THEN	10810
	DGI=LG(I)	10820
	ELSE	10830
	DGI=DLOG10(DGI)	10840
	END IF	10850
	DM=LKSP(I)-DGI	10860
	K=NSP(I)	10870
C	CALCULATE MOLALITY	10880
	DO 92 J=1,K	10890
	DM=DM+LA(LSP(I,J))*CSP(I,J)	10900
92	CONTINUE	10910
	LM(I)=DM	10920
	M(I)=UNDER(DM)	10930
122	CONTINUE	10940
	IF(IOPT(8).GT.0) THEN	10950
	WRITE(6,107)	10960
107	FORMAT('PRINTED IN RESET AFTER RECALCULATION OF M(I).',	10970
1	/, ' I GAMMA MOLALITY')	10980
	DO 105 I=1,LASTS	10990
	IF (SFLAG(I).EQ.0) GO TO 105	11000
	D=1D1**LG(I)	11010
	WRITE(6,106) I,D,M(I)	11020
106	FORMAT(I5,1P2D10.3)	11030
105	CONTINUE	11040
	END IF	11050
C	RESET TOTALS	11060
	IF (NMINS.EQ.0) GO TO 180	11070
	DO 160 I=IFT,ILT	11080
	K=IOUT(I)	11090
	TOT(K)=TOT(K)+DELTOT(I)	11100
	IF (TOT(K).LE.1D-30) TOT(K)=1D-30	11110
160	CONTINUE	11120
C	RESET TOTAL MINERALS ADDED OR REMOVED AND THE THOR.	11130
	DO 170 I=IFM,ILM	11140
	K=MAXT+I-ILT	11150
	TOT(K)=TOT(K)+DELTA(I)	11160
	IF (IOPT(3).EQ.6.AND.I.EQ.IFM) THEN	11170
	THOR=THOR+DELTA(I)*THREAC	11180

ELSE	11190
THOR=THOR+THMIN(I-ILT)*DELTA(I)	11200
END IF	11210
170 CONTINUE	11220
180 CONTINUE	11230
RETURN	11240
END	11250
FUNCTION UNDER(D)	11260
DOUBLE PRECISION D,UNDER	11270
UNDER=ODO	11280
IF (D.LT.-33.0D0) GO TO 10	11290
UNDER=1D2	11300
IF (D.GT.2D0) GO TO 10	11310
UNDER=1D1**D	11320
10 CONTINUE	11330
RETURN	11340
END	11350
SUBROUTINE PTOT	11360
\$INSERT COMMON.BLOCKS	11370
LOGICAL EVAP	11380
COMMON /EVAP/ EVAP	11390
COMMON / MX10 / VP,DWO	11400
CHARACTER *12 DALK,DN,DN1,DN2,DN3,DNAME	11410
CHARACTER *80 CARD	11420
DIMENSION DAO(250),DM(250),COR(1:34),DGO(40)	11430
LOGICAL LHSP,LNHSP	11440
COMMON / COS / COSMOT	11450
DATA IDATA /0/, DALK/'TOT ALK'/	11460
DATA DN1/'O2'/,DN2/'H2'/,DN3/'CHARGE'/	11470
WRITE (6,220)	11480
WRITE (6,230)	11490
KK=0	11500
DO 10 I=4,MAXT	11510
IF (DABS(TOT(I)).LE.1.0D-40) GO TO 10	11520
KK=1	11530
DLT=DLOG10(TOT(I))	11540
DNAME=TNAME(I)	11550
IF (IASPEC.EQ.I) DNAME=DALK	11560
WRITE (6,240) DNAME,TOT(I),DLT	11570
10 CONTINUE	11580
IF (KK.EQ.0) WRITE (6,250)	11590
WRITE (6,260)	11600
RETURN	11610
C *****	11620
ENTRY PSPEC	11630
C *****	11640
WRITE (6,270)	11650
WRITE (6,280) PH,TOT(3),COSMOT,MU,TC,VP,DWO,ELECT,DALKS,ITER	11660
IF (IASPEC.LE.0) GO TO 20	11670
WRITE (6,290) TOTAL(ISOL,IASPEC)	11680
20 CONTINUE	11690
IF (IESPEC.LE.1) GO TO 30	11700
DD=TOT(IESPEC)-DNEUT	11710
WRITE (6,300) TNAME(IESPEC),DD	11720
30 CONTINUE	11730
WRITE (6,310)	11740

IF(IOPT(10).EQ.0) THEN	11750
WRITE(6,312)	11760
ELSE	11770
WRITE(6,311)	11780
END IF	11790
WRITE (6,313)	11800
LM(1)=LA(1)-LG(1)	11810
LG(2)=0.0D0	11820
LM(2)=LA(2)	11830
LG(3)=0.0D0	11840
LM(3)=LA(3)	11850
DO 40 I=1,LASTS	11860
IF (SFLAG(I).EQ.0) GO TO 40	11870
DAO(I)=0D0	11880
DM(I)=0D0	11890
IF (LM(I).LT.-30D0.AND.I.GT.30) GO TO 40	11900
IF (LM(I).GE.-60D0) DM(I)=1D1**LM(I)	11910
DLA=LM(I)+LG(I)	11920
IF (DLA.GE.-60D0) DAO(I)=1D1**DLA	11930
DG=1D1**LG(I)	11940
IF (I.NE.2) WRITE (6,320) I,SNAME(I),ZSP(I),DM(I),LM(I),DAO(I)	11950
1,DLA,DG,LG(I)	11960
40 CONTINUE	11970
DO 41 I=1,34	11980
DGO(I)=0.0D0	11990
COR(I)=0.0D0	12000
41 CONTINUE	12010
DO 52 I=35,LASTS	12020
IF (SFLAG(I).EQ.0) GO TO 52	12030
LHSP=.FALSE.	12040
LNHSP=.FALSE.	12050
DO 45 J=1,NSP(I)	12060
IF (LSP(I,J).NE.1) GO TO 45	12070
LHSP=.TRUE.	12080
IF (CSP(I,J).GT.0.0D0) LNHSP=.TRUE.	12090
ITH=J	12100
45 CONTINUE	12110
DO 42 J=1,NSP(I)	12120
IF (LSP(I,J).EQ.3.AND.LHSP) GO TO 47	12130
IF (LSP(I,J).EQ.3.OR.LSP(I,J).EQ.2) GO TO 42	12140
IF (LSP(I,J).EQ.15.AND.LNHSP) GO TO 48	12150
IF (LSP(I,J).EQ.16.AND.LNHSP) GO TO 42	12160
IF (LSP(I,J).NE.1) GO TO 49	12170
IF (.NOT.LNHSP) GO TO 42	12180
DO 46 N=1,NSP(I)	12190
IF (LSP(I,N).EQ.15) GO TO 42	12200
46 CONTINUE	12210
49 COR(LSP(I,J))=COR(LSP(I,J))+CSP(I,J)*DM(I)	12220
GO TO 42	12230
47 DCSP=DABS(CSP(I,J)-DABS(CSP(I,ITH)))	12240
IF (DCSP.GT.0.1D0) GO TO 42	12250
COR(31)=COR(31)+CSP(I,J)*DM(I)	12260
GO TO 42	12270
48 IF (CSP(I,ITH).GT.1.1D0) GO TO 42	12280
COR(34)=COR(34)+CSP(I,J)*DM(I)	12290
42 CONTINUE	12300

52	CONTINUE	12310
C		12320
	IF(IOPT(10).EQ.0) THEN	12330
	WRITE(6,36)	12340
36	FORMAT(//39X,'UNSCALED UNSCALED')	12350
	ELSE	12360
	WRITE(6,35)	12370
35	FORMAT(//39X,'MACINNES MACINNES')	12380
	END IF	12390
	WRITE (6,31)	12400
31	FORMAT (14X,'SPECIES',4X,'TOTAL MOL',5X,'ACTIVITY',3X,	12410
	1'TOTAL GAMMA'/)	12420
	DO 43 I=1,35	12430
	IF (SFLAG(I).EQ.0) GO TO 43	12440
	IF (I.EQ.2.OR.I.EQ.3) GO TO 43	12450
	IF (I.EQ.35) GO TO 32	12460
	DM(I)=DM(I)+COR(I)	12470
32	DGO(I)=DAO(I)/DM(I)	12480
	WRITE (6,44) SNAME(I),DM(I),DAO(I),DGO(I)	12490
44	FORMAT (14X,A8,3(3X,1PD10.4))	12500
43	CONTINUE	12510
C		12520
	IF (NMGS.EQ.0) GO TO 700	12530
	STOT=0.0D0	12540
	DO 590 I=4,30	12550
	STOT=STOT+TOT(I)	12560
590	CONTINUE	12570
	IF (STOT.LT.1.0D-20) GO TO 700	12580
600	WRITE (6,610)	12590
610	FORMAT (//22X,'----- MEAN ACTIVITY COEFFICIENT -----',//	12600
	1,27X,'FORMULA',7X,'MEAN GAMMA',/)	12610
	DO 620 I=1,NMGS	12620
	EG=0.0D0	12630
	GG=1.0D0	12640
	DO 630 J=1,IMEANG(I)	12650
	GG=GG*DGO(LMEANG(I,J))**CMEANG(I,J)	12660
	EG=EG+CMEANG(I,J)	12670
630	CONTINUE	12680
	IF (GG.LT.1.0D-30) GO TO 620	12690
	GG=GG**(1.0D0/EG)	12700
	WRITE (6,640) NMEANG(I),GG	12710
640	FORMAT (27X,A8,6X,1PD10.4)	12720
620	CONTINUE	12730
C	*****	12740
	ENTRY PSUM	12750
C	*****	12760
700	IF (NSUMS.EQ.0) GO TO 80	12770
	KK=0	12780
	DO 70 I=1,NSUMS	12790
	LL=0	12800
	DSUM=0D0	12810
	K=NSUM(I)	12820
	DO 50 J=1,K	12830
	IF (SFLAG(LSUM(I,J)).LE.0) GO TO 50	12840
	LL=1	12850
	DSUM=DSUM+M(LSUM(I,J))	12860

50	CONTINUE	12870
	IF (LL.EQ.0) GO TO 70	12880
	IF (LL.EQ.0.OR.KK.NE.0) GO TO 60	12890
	WRITE (6,330)	12900
	KK=1	12910
60	CONTINUE	12920
	WRITE (6,340) SNAME(I),DSUM	12930
70	CONTINUE	12940
80	CONTINUE	12950
	RETURN	12960
C	*****	12970
	ENTRY PBUG(D,DPHPE)	12980
C	*****	12990
	WRITE (6,350) D,CR(1),DPHPE	13000
C	PRINT CHANGES IN PH AND PE	13010
	IF (IIN(1).LE.0) GO TO 90	13020
	DPH=-DLOG10(1D0+DELTA(1))	13030
	WRITE (6,360) PH,DPH	13040
90	CONTINUE	13050
	IF (IIN(2).LE.0) GO TO 100	13060
	DPE=-DLOG10(1D0+DELTA(IIN(2)))	13070
	WRITE (6,370) PE,DPE,CR(2)	13080
100	CONTINUE	13090
C	PRINT TOTALS AND ACTIVITIES	13100
	WRITE(6,105)	13110
105	FORMAT(' ERROR EQN TOT DELTOT FRAC TOT '	13120
1	'SPECIES MOLALITY DELTA MOL')	13130
	DO 110 K=4,LASTT	13140
	I=IIN(K)	13150
	IF (I.LE.0) GO TO 110	13160
	DRT=DELTOT(I)/TOT(K)	13170
	TEMPM=1D1**LM(K)	13180
	DRA=DELTA(I)*DA	13190
	WRITE (6,380) CR(K),TNAME(K),TOT(K),DELTOT(I),DRT,SNAME(K),TEMPM	13200
	1,DELTA(I)	13210
110	CONTINUE	13220
C	PRINT MINERAL TOTALS AND DELTAS	13230
	IF (NMINS.LE.0) GO TO 130	13240
	DO 120 I=1,NMINS	13250
	K=MAXT+I	13260
	J=ILT+I	13270
	WRITE (6,390) CR(K),MNAME(I),TOT(K),DELTA(J)	13280
120	CONTINUE	13290
130	CONTINUE	13300
	RETURN	13310
C	*****	13320
	ENTRY PPHASE	13330
C	*****	13340
C		13350
C	PRINT MASS TRANSFER FOR MINERAL PHASES	13360
C		13370
	IF (NMINS.EQ.0) RETURN	13380
	IF (EVAP) THEN	13390
	FACTOR=TOT(MAXT+1)+1D0	13400
	WRITE (6,401)	13410
401	FORMAT (27X,'----PHASE BOUNDARIES----'/	13420

1	/12X, ' MASS PRECIPITATED/DISSOLVED FROM INITIAL KILOGRAM WATER'	13430
2	//, 9X, 'PHASE', 5X, 'DELTA PHASE*', 6X, 'LOG IAP', 6X, 'LOG KT',	13440
3	6X, 'LOG IAP/KT'/)	13450
	ELSE	13460
	FACTOR=1DO	13470
	WRITE (6,400)	13480
	END IF	13490
	DO 150 I=1,NMINS	13500
	K=NMINO(I)	13510
	DIAP=0.0DO	13520
	DO 140 J=1,K	13530
	DIAP=DIAP+(LG(LMINO(I,J))+LM(LMINO(I,J)))*CMINO(I,J)	13540
140	CONTINUE	13550
	DSI=DIAP-LKMINO(I)	13560
	K=MAXT+I	13570
	D1=TOT(K)/FACTOR	13580
	IF (IOPT(3).EQ.6.AND.I.EQ.1) D1=0DO	13590
	WRITE (6,410) MNAME(I),D1,DIAP,LKMINO(I),DSI	13600
	IF (IOPT(3).EQ.6.AND.I.EQ.1) WRITE (6,420)	13610
150	CONTINUE	13620
	WRITE (6,430)	13630
C		13640
C	PRINT IF ADDING REACTION TO A PHASE BOUNDARY	13650
C		13660
	IF (IOPT(3).EQ.6) THEN	13670
	IF (EVAP) THEN	13680
	FACTOR=TOT(MAXT+1)+1DO	13690
	WRITE (6,445) FACTOR, MNAME(1)	13700
445	FORMAT (/,1X,'** ',1PD13.6,' IS THE EVAPORATION FACTOR NECESSARY'	13710
1	, ' TO REACH THE ',/4X,14X,A8,' PHASE BOUNDARY.')	13720
	WRITE (6,221)	13730
221	FORMAT(///17X,'MOLES OF ELEMENTS REMAINING AFTER REACTION'	13740
1	/,17X,'----- -- -----')	13750
	WRITE (6,231)	13760
231	FORMAT (/16X,'ELEMENT',10X,' MOLES ',9X,' LOG MOLES'/)	13770
	DO 11 I=4,MAXT	13780
	IF (DABS(TOT(I)).LE.1.0D-40) GO TO 11	13790
	DMT=TOT(I)/FACTOR	13800
	DLT=DLOG10(DMT)	13810
	DNAME=TNAME(I)	13820
	WRITE (6,240) DNAME,DMT,DLT	13830
11	CONTINUE	13840
	WRITE (6,242) 1DO/FACTOR	13850
242	FORMAT(/15X,1PD13.6, ' KILOGRAMS OF WATER REMAINING')	13860
	WRITE (6,260)	13870
	ELSE	13880
	WRITE (6,440) TOT(MAXT+1),MNAME(1)	13890
	END IF	13900
	END IF	13910
	IF (IOPT(3).NE.6.OR.EVAP) RETURN	13920
C	*****	13930
	ENTRY PREAC	13940
C	*****	13950
	WRITE (6,450)	13960
	DO 160 I=1,NCOMPS	13970
	L=LREAC(I)	13980

	IF (L.EQ.0) DN=DN3	13990
	IF (L.LT.31) DN=TNAME(L)	14000
	IF (L.GT.30) DN=DN1	14010
	IF (L.GT.30.AND.THMEAN(I).LT.0) DN=DN2	14020
	WRITE (6,460) CREAC(I),DN,THMEAN(I)	14030
160	CONTINUE	14040
	RETURN	14050
C	*****	14060
	ENTRY PLOOK	14070
C	*****	14080
	IF (NLOOKS.EQ.0) RETURN	14090
	KK=0	14100
	DO 190 I=1,NLOOKS	14110
	K=NLOOK(I)	14120
	DIAP=ODO	14130
	DO 170 J=1,K	14140
	LL=LLOOK(I,J)	14150
	IF (SFLAG(LL).LE.0) GO TO 190	14160
	DIAP=DIAP+(LG(LL)+LM(LL))*CLOOK(I,J)	14170
170	CONTINUE	14180
	IF (KK.NE.0) GO TO 180	14190
	KK=1	14200
	WRITE (6,470)	14210
180	CONTINUE	14220
	DSI=DIAP-LKLOOK(I)	14230
	WRITE (6,480) NAMELK(I),DIAP,LKLOOK(I),DSI	14240
190	CONTINUE	14250
	RETURN	14260
C	*****	14270
	ENTRY PDATA	14280
C	*****	14290
	IF (IOPT(1).NE.1) RETURN	14300
	IF (IDATA.GT.0) RETURN	14310
	IDATA=1	14320
	REWIND (UNIT=11)	14330
	WRITE (6,490)	14340
200	CONTINUE	14350
	READ (11,500,END=210) CARD	14360
	WRITE (6,510) CARD	14370
	GO TO 200	14380
210	CONTINUE	14390
	RETURN	14400
C		14410
220	FORMAT (///24X,'TOTAL MOLALITIES OF ELEMENTS'/,24X,'-----	14420
	1-- -- -----')	14430
230	FORMAT (/16X,'ELEMENT',10X,'MOLALITY',9X,'LOG MOLALITY'/)	14440
240	FORMAT (16X,A8,6X,1PD13.6,8X,OPF9.4)	14450
250	FORMAT (16X,'PURE WATER')	14460
260	FORMAT (//)	14470
270	FORMAT (//,24X,'----DESCRIPTION OF SOLUTION----')	14480
280	FORMAT (/ ,39X,'PH = ',F8.4/29X,'ACTIVITY H2O = ',F8.4/22X	14490
	1,'OSMOTIC COEFFICIENT = ',F8.4/,27X,'IONIC STRENGTH = ',F8.4/30X	14500
	2,'TEMPERATURE = ',F8.4/33X,'PRESSURE = ',F8.4,' ATM',/27X	14510
	3,'DENSITY OF H2O = ',F8.4,' G/CC',/23X,'ELECTRICAL BALANCE = '	14520
	4,1PD12.4/25X,'TOTAL ALKALINITY = ',D12.4/31X,'ITERATIONS = ',OPI3)	14530
290	FORMAT (29X,'TOTAL CARBON = ',1PD12.4)	14540

```

300 FORMAT (20X,'MOLES OF ',A8,' ADDED = ',D12.4) 14550
310 FORMAT (///,26X,23('-')/26X,'DISTRIBUTION OF SPECIES'/26X,23('-')) 14560
311 FORMAT (43X,'MACINNES SCALE      MACINNES SCALE ') 14570
312 FORMAT (43X,'      UNSCALED      UNSCALED') 14580
313 FORMAT (2X,'I',2X,'SPECIES',4X,'Z',3X,'MOLALITY',2X,'LOG MOLAL', 14590
1 3X,'ACTIVITY',2X,'LOG ACT',4X,'GAMMA',3X,'LOG GAM'/) 14600
320 FORMAT (1X,I3,1X,A8,1X,F4.1,1X,1PE10.3,2X,OPF7.3,2X,2(1PE10.3,1X, 14610
10PF7.3,1X)) 14620
330 FORMAT (///25X,'SUMS OF SPECIES'/) 14630
340 FORMAT (22X,A8,' = ',1PD13.6) 14640
350 FORMAT (/1X,'REDUCTION FACTOR: ',1PD12.5,3X,'ELECT: ',D12.5,3X, 14650
1'DPHPE: ',D12.5) 14660
360 FORMAT (1X,'PH = ',F8.4,5X,'DPH = ',F8.4) 14670
370 FORMAT (1X,'PE = ',F8.4,5X,'DPE = ',F8.4,5X,'DTHOR = ',1PD12.5) 14680
380 FORMAT (1X,1PD12.5,1X,A4,1X,D11.4,2(1X,D8.1),1X,A8,1X,D11.4,1X, 14690
1D8.1) 14700
390 FORMAT (1X,1PD12.5,1X,A4,1X,D11.4,2(1X,D8.1)) 14710
400 FORMAT (27X,'----PHASE BOUNDARIES----'//,9X,'PHASE',5X,'DELTA PHAS 14720
1E*',6X,'LOG IAP',6X,'LOG KT',6X,'LOG IAP/KT'/) 14730
410 FORMAT (8X,A8,2X,1PD13.6,3(4X,OPF9.4)) 14740
420 FORMAT (1H+,5X,'**') 14750
430 FORMAT (/,1X,'* NEGATIVE DELTA PHASE INDICATES PRECIPITATION',/,3X 14760
1,'AND POSITIVE DELTA PHASE INDICATES DISSOLUTION.')

```

	ISUB(I)=0	15110
20	CONTINUE	15120
C	READ TITLE CARD	15130
	READ (5,420,END=410) (TITLE(I),I=1,20)	15140
C	READ OPTION CARD	15150
	READ (5,430) (IOPT(I),I=1,10),NSTEPS,NCOMPS,VO	15160
C	SET MANDATORY OPTION	15170
	IOPT(5)=0	15180
	IOPT(6)=2	15190
	IF (IOPT(1).EQ.1) CALL PDATA	15200
	WRITE (6,440) (TITLE(I),I=1,20)	15210
	WRITE (6,450) (IOPT(I),I=1,10),NSTEPS,NCOMPS,VO	15220
	IREAD=0	15230
	IFILE=5	15240
	GO TO 30	15250
C	*****	15260
	ENTRY RDATA (KFILE)	15270
C	*****	15280
	JFILE=6	15290
	WRITE (JFILE,460)	15300
	IREAD=1	15310
	IFILE=KFILE	15320
	30 CONTINUE	15330
C	READ SUBROUTINE CARD AND BRANCH	15340
40	CONTINUE	15350
	READ (IFILE,470) SUB,ISOLN	15360
	DO 50 I=1,12	15370
	IF (SUB.NE.SUBS(I)) GO TO 50	15380
	GO TO (60,70,110,150,170,180,220,400,260,350,390,1000), I	15390
50	CONTINUE	15400
	WRITE (6,480) SUB	15410
	GO TO 40	15420
C		15430
C	READ TEMP DATA	15440
C		15450
60	CONTINUE	15460
	WRITE (6,490)	15470
	IF (IOPT(4).EQ.0) GO TO 40	15480
	NTEMP=IOPT(4)	15490
	IF (NTEMP.EQ.3) NTEMP=NSTEPS	15500
	READ (5,500) (TSTEP(I),I=1,NTEMP)	15510
	WRITE (6,510) (TSTEP(I),I=1,NTEMP)	15520
	ISUB(1)=1	15530
	GO TO 40	15540
C		15550
C	READ MINERAL DATA	15560
C		15570
70	CONTINUE	15580
	WRITE (6,520)	15590
	I=0	15600
80	CONTINUE	15610
	I=I+1	15620
C	READ NAME ETC. FOR MINERAL	15630
	READ (5,530) MNAME(I),NMINO(I),THMIN(I),LKTOM(I),DHMIN(I),MFLAG(I)	15640
	1,SIMIN(I)	15650
	WRITE (6,540) MNAME(I),NMINO(I),THMIN(I),LKTOM(I),DHMIN(I),MFLAG	15660

	1(I),SIMIN(I)	15670
	IF (NMINO(I).EQ.0) GO TO 100	15680
	NRMINS=I	15690
C	READ COEFFICIENTS FOR MINERAL	15700
	K=NMINO(I)	15710
	READ (5,550) (LMINO(I,J),CMINO(I,J),J=1,K)	15720
	WRITE (6,560) (LMINO(I,J),CMINO(I,J),J=1,K)	15730
C	READ ANALYTIC EXPRESSION COEFFICIENTS	15740
	IF (MFLAG(I).EQ.0) GO TO 90	15750
	READ (5,570) (AMIN(I,J),J=1,5)	15760
	WRITE (6,580) (AMIN(I,J),J=1,5)	15770
90	CONTINUE	15780
	GO TO 80	15790
100	CONTINUE	15800
	CALL MINCON	15810
	CALL CHKMIN(IERR)	15820
	ISUB(2)=1	15830
	GO TO 40	15840
C		15850
C	READ A SOLUTION	15860
C		15870
110	CONTINUE	15880
	WRITE (6,590) ISOLN	15890
C	READ SOLUTION TITLE	15900
	READ (5,600) (HEAD(ISOLN,J),J=1,20)	15910
	WRITE (6,610) (HEAD(ISOLN,J),J=1,20)	15920
C	READ NTOTS,PH,EH	15930
	READ (5,620) NTOTS,IALK(ISOLN),IUNITS(ISOLN),PH,PE,TEMP(ISOLN),	15940
	1SDENS(ISOLN)	15950
C	FIX PE AT 4.0	15960
	PE=4.0	15970
	IF (DABS(SDENS(ISOLN)).LE.1.0D-40) SDENS(ISOLN)=1.0	15980
	WRITE (6,630) NTOTS,IALK(ISOLN),IUNITS(ISOLN),PH,PE,TEMP(ISOLN),	15990
	1SDENS(ISOLN)	16000
C	READ TOTALS	16010
	DO 120 I=1,MAXT	16020
	TOTAL(ISOLN,I)=ODO	16030
120	CONTINUE	16040
	IF (NTOTS.EQ.0) GO TO 140	16050
	READ (5,640) (LT(J),DTOT(J),J=1,NTOTS)	16060
	WRITE (6,650) (LT(J),DTOT(J),J=1,NTOTS)	16070
C	ZERO TOTAL AND INSERT VALUES	16080
	NTOTM=0	16090
	DO 130 I=1,NTOTS	16100
C	THROW OUT ELEMENTS NOT IN PITZER'S MODEL	16110
	IF (LELEM(LT(I))) GO TO 125	16120
	NTOTM=NTOTM+1	16130
	GO TO 130	16140
125	TOTAL(ISOLN,LT(I))=DTOT(I)	16150
130	CONTINUE	16160
	NTOTS=NTOTS-NTOTM	16170
C		16180
	CALL UNITS(ISOLN)	16190
140	CONTINUE	16200
	TOTAL(ISOLN,1)=1D1**(-PH)	16210
	TOTAL(ISOLN,2)=1D1**(-PE)	16220

	ISOLV(ISOLN)=1	16230
	ISUB(3)=1	16240
	GO TO 40	16250
C		16260
C	READ REACTION DATA	16270
C		16280
150	CONTINUE	16290
	IF (NCOMPS.EQ.0) GO TO 40	16300
	WRITE(6,660)	16310
C	READ COEFFICIENTS OF REACTION	16320
	READ (5,670) (LREAC(I),CREAC(I),THMEAN(I),I=1,NCOMPS)	16330
	WRITE (6,680) (LREAC(I),CREAC(I),THMEAN(I),I=1,NCOMPS)	16340
	THREAC=ODO	16350
	DO 160 I=1,NCOMPS	16360
	THREAC=THREAC+THMEAN(I)*CREAC(I)	16370
160	CONTINUE	16380
	IF(IOPT(3).EQ.6.AND.NCOMPS.EQ.1.AND.LREAC(1).EQ.3) THEN	16390
	EVAP=.TRUE.	16400
	END IF	16410
	ISUB(4)=1	16420
	GO TO 40	16430
C		16440
C	READ STEPS DATA	16450
C		16460
170	CONTINUE	16470
	IF (NSTEPS.EQ.0.OR.IOPT(3).EQ.5) GO TO 40	16480
	WRITE(6,690)	16490
	K=NSTEPS	16500
	IF (IOPT(3).EQ.4) K=1	16510
	READ (5,700) (XSTEP(I),I=1,K)	16520
	WRITE (6,710) (XSTEP(I),I=1,K)	16530
	ISUB(5)=1	16540
	GO TO 40	16550
C		16560
C	READ ELEMENT CARDS	16570
C		16580
180	CONTINUE	16590
	WRITE(6,720)	16600
C	READ ELEMENTS UNTIL BLANK CARD	16610
190	CONTINUE	16620
	READ (IFILE,730) TNAM,NELT,TGFW	16630
	IF (IREAD.EQ.0) WRITE (6,740) TNAM,NELT,TGFW	16640
	IF (NELT.EQ.0) GO TO 210	16650
	IF (NELT.GT.3.AND.NELT.LE.29) GO TO 200	16660
	WRITE (6,750) NELT,TNAM,SUBS(6)	16670
	ENDFILE (UNIT=6)	16680
	STOP 'INDEX ERROR IN SUBROUTINE READ'	16690
200	CONTINUE	16700
	IF (IREAD.EQ.1) LELEM(NELT)=.TRUE.	16710
	TNAME(NELT)=TNAM	16720
	GFW(NELT)=TGFW	16730
	GO TO 190	16740
210	CONTINUE	16750
	ISUB(6)=1	16760
	GO TO 40	16770
C		16780

C	READ SPECIES CARDS	16790
C		16800
220	CONTINUE	16810
	WRITE (6,760)	16820
C	READ SPECIES CARDS UNTIL BLANK CARD	16830
230	CONTINUE	16840
	READ (IFILE,770) I	16850
	IF (IREAD.EQ.0) WRITE (6,780) I	16860
C	READ SPECIES NAME ETC.	16870
	IF (I.EQ.0) GO TO 250	16880
	IF (I.GT.0.AND.I.LE.MAXS) GO TO 240	16890
	READ (IFILE,790) TNAM	16900
	WRITE (6,750) I,TNAM,SUBS(7)	16910
	ENDFILE (UNIT=6)	16920
	STOP 'INDEX ERROR IN SUBROUTINE READ'	16930
240	CONTINUE	16940
	READ (IFILE,790) SNAME(I),NSP(I),KFLAG(I),GFLAG(I),ZSP(I),THSP(I),	16950
	1DHA(I),(ADHSP(I,J),J=1,2),ALKSP(I)	16960
	IF (IREAD.EQ.0) WRITE (6,800)SNAME(I),NSP(I),KFLAG(I),GFLAG(I)	16970
	1,ZSP(I),THSP(I),DHA(I),(ADHSP(I,J),J=1,2),ALKSP(I)	16980
	IF (NSP(I).EQ.0) GO TO 230	16990
C	READ LOG K DATA FOR SPECIES	17000
	READ (IFILE,810) LKTOSP(I),DHSP(I),(ASP(I,J),J=1,5)	17010
	IF (IREAD.EQ.0) WRITE (6,820) LKTOSP(I),DHSP(I),(ASP(I,J),J=1,5)	17020
C	READ COEFFICIENTS	17030
	K=NSP(I)	17040
	READ (IFILE,830) (LSP(I,J),CSP(I,J),J=1,K)	17050
	IF (IREAD.EQ.0) WRITE (6,840) (LSP(I,J),CSP(I,J),J=1,K)	17060
	GO TO 230	17070
250	CONTINUE	17080
	ISUB(7)=1	17090
	CALL CHKSPE(IERR)	17100
	GO TO 40	17110
C		17120
C	READ MINERALS TO LOOK AT	17130
C		17140
260	CONTINUE	17150
	WRITE (6,850)	17160
	I=NLKS	17170
270	CONTINUE	17180
	I=I+1	17190
	IF (I.LT.41) GO TO 280	17200
	WRITE (6,860)	17210
	I=40	17220
280	CONTINUE	17230
C	READ NAME ETC.	17240
	READ (IFILE,870) NAMELK(I),NLOOK(I),LKOLK(I),DHLOOK(I),LOOKFL(I)	17250
	IF (IREAD.EQ.0) WRITE (6,880) NAMELK(I),NLOOK(I),LKOLK(I),DHLOOK	17260
	1(I),LOOKFL(I)	17270
	IF (NAMELK(I).NE.DELETE) GO TO 290	17280
	I=0	17290
	NLOOKS=I	17300
	GO TO 270	17310
290	CONTINUE	17320
	IF (NLOOK(I).EQ.0) GO TO 340	17330
C	CHECK TO SEE IF MINERAL IS ALREADY IN THE LIST	17340

IF (I.LE.1) GO TO 310	17350
L=I-1	17360
DO 300 N=1,L	17370
IF (NAMELK(I).NE.NAMELK(N)) GO TO 300	17380
NLOOK(N)=NLOOK(I)	17390
LKOLK(N)=LKOLK(I)	17400
DHLOOK(N)=DHLOOK(I)	17410
LOOKFL(N)=LOOKFL(I)	17420
I=I-1	17430
GO TO 320	17440
300 CONTINUE	17450
310 CONTINUE	17460
N=I	17470
NLOOKS=I	17480
C N IS THE NUMBER FOR THE MINERAL BEING READ.	17490
320 CONTINUE	17500
C READ COEFFICIENTS	17510
K=NLOOK(N)	17520
READ (IFILE,550) (LLOOK(N,J),CLOOK(N,J),J=1,K)	17530
IF (IREAD.EQ.0) WRITE (6,560) (LLOOK(N,J),CLOOK(N,J),J=1,K)	17540
C READ ANALYTIC EXPRESSION	17550
IF (LOOKFL(N).EQ.0) GO TO 330	17560
READ (IFILE,570) (ALOOK(N,J),J=1,5)	17570
IF (IREAD.EQ.0) WRITE (6,580) (ALOOK(N,J),J=1,5)	17580
330 CONTINUE	17590
GO TO 270	17600
340 CONTINUE	17610
NLKS=NLOOKS	17620
CALL CHKLK	17630
ISUB(9)=1	17640
GO TO 40	17650
C	17660
C READ LISTS FOR SUMS OF SPECIES	17670
C	17680
350 CONTINUE	17690
WRITE (6,890)	17700
I=NSMS	17710
360 CONTINUE	17720
I=I+1	17730
READ (IFILE,900) Surname(I),K	17740
IF (IREAD.EQ.0) WRITE (6,910) Surname(I),K	17750
IF (Surname(I).NE.DELETE) GO TO 370	17760
I=0	17770
NSUMS=0	17780
GO TO 360	17790
370 CONTINUE	17800
IF (K.EQ.0) GO TO 380	17810
NSUM(I)=K	17820
NSUMS=I	17830
READ (IFILE,920) (LSUM(I,J),J=1,K)	17840
IF (IREAD.EQ.0) WRITE (6,930) (LSUM(I,J),J=1,K)	17850
GO TO 360	17860
380 CONTINUE	17870
NSMS=NSUMS	17880
ISUB(10)=1	17890
GO TO 40	17900

C		17910
C	READ SPECIES TO ADJUST FOR ELECTRICAL NEUTRALITY	17920
C		17930
	390 CONTINUE	17940
	WRITE (6,940)	17950
	READ (5,950) LPOS,LNEG	17960
	WRITE (6,960) LPOS,LNEG	17970
	ISUB(11)=1	17980
	GO TO 40	17990
C		18000
C	END DATA ENTRY FOR THIS SIMULATION	18010
C		18020
	400 CONTINUE	18030
	IF (IREAD.EQ.1) RETURN	18040
	IF (IERR.LE.0) GO TO 405	18050
	ENDFILE (UNIT=6)	18060
	STOP 'ERROR IN SPECIES OR MINERAL EQUATIONS'	18070
	405 IF (IOPT(4).GT.0.AND.ISUB(1).EQ.0) IERR=1	18080
	IF (IOPT(3).GT.0.AND.IOPT(3).LT.5.AND.ISUB(5).EQ.0) IERR=5	18090
	IF (IOPT(3).GE.3.AND.IOPT(3).LT.5.AND.ISUB(4).EQ.0) IERR=4	18100
	IF (IOPT(3).EQ.5.AND.ISUB(2).EQ.0) IERR=2	18110
	IF (IOPT(3).EQ.6.AND.ISUB(2).EQ.0) IERR=2	18120
	IF (IOPT(3).EQ.6.AND.ISUB(4).EQ.0) IERR=4	18130
	IF (IOPT(2).EQ.2.AND.ISUB(11).EQ.0) IERR=11	18140
	IF (IERR.EQ.0) RETURN	18150
	WRITE (6,970) SUBS(IERR)	18160
	IF (IOPT(7).EQ.0) GO TO 10	18170
	410 CONTINUE	18180
	ENDFILE (UNIT=6)	18190
	STOP 'NORMAL TERMINATION'	18200
C		18210
C	READ MEAN GAMMAS	18220
C		18230
	1000 WRITE (6,980)	18240
	1005 NMG=NMG+1	18250
	IF (NMG.LE.40) GO TO 1100	18260
	WRITE (6,990)	18270
	1010 READ (IFILE,1020) NM	18280
	1020 FORMAT (8X,I2)	18290
	IF (NM.EQ.0) GO TO 40	18300
	GO TO 1010	18310
	1100 READ (IFILE,995) NMEANG(NMG),IMEANG(NMG),(LMEANG(NMG,J)	18320
	1,CMEANG(NMG,J),J=1,3)	18330
	IF (IREAD.EQ.0) WRITE (6,996) NMEANG(NMG),IMEANG(NMG)	18340
	1,(LMEANG(NMG,J),CMEANG(NMG,J),J=1,IMEANG(NMG))	18350
	IF (NMEANG(NMG).EQ.DELETE) GO TO 1150	18360
	IF (IMEANG(NMG).NE.0) GO TO 1110	18370
	NMG=NMG-1	18380
	GO TO 1200	18390
	1110 DO 1120 J=1,NMG-1	18400
	IF (NMEANG(J).EQ.NMEANG(NMG)) GO TO 1130	18410
	1120 CONTINUE	18420
	GO TO 1005	18430
	1130 IMEANG(J)=IMEANG(NMG)	18440
	DO 1140 I=1,IMEANG(J)	18450
	LMEANG(J,I)=LMEANG(NMG,I)	18460

	CMEANG(J,I)=CMEANG(NMG,I)	18470
1140	CONTINUE	18480
	NMG=NMG-1	18490
	GO TO 1005	18500
1150	NMG=0	18510
	GO TO 1005	18520
1200	NMGS=NMG	18530
	GO TO 40	18540
C		18550
420	FORMAT (20A4)	18560
430	FORMAT (10I1,I2,I2,6X,F10.5)	18570
440	FORMAT (1H1,20A4)	18580
450	FORMAT (1X,10I1,I2,I2,6X,F10.5)	18590
460	FORMAT (10X,'DATA READ FROM DISK'/)	18600
470	FORMAT (A8,1X,I1)	18610
480	FORMAT (1X,'** ',A8,' ** INPUT UNKNOWN')	18620
490	FORMAT (1X,'TEMP')	18630
500	FORMAT (8F10.1)	18640
510	FORMAT (1X,8G10.3)	18650
520	FORMAT (1X,'MINERALS')	18660
530	FORMAT (A8,2X,I2,3X,3F10.2,5X,I1,9X,F10.3)	18670
540	FORMAT (1X,A8,2X,I2,3X,3G10.2,5X,I1,9X,F10.3)	18680
550	FORMAT ((5(I4,F11.3)))	18690
560	FORMAT ((1X,5(I4,G11.3)))	18700
570	FORMAT (5E12.5)	18710
580	FORMAT (1X,1P5E12.4)	18720
590	FORMAT (1X,'SOLUTION ',I1)	18730
600	FORMAT (20A4)	18740
610	FORMAT (1X,20A4)	18750
620	FORMAT (I2,I3,I2,3X,4F10.0)	18760
630	FORMAT (1X,I2,I3,I2,3X,4G10.3)	18770
640	FORMAT ((5(I4,D11.3)))	18780
650	FORMAT ((1X,5(I4,1PD11.3)))	18790
660	FORMAT (1X,'REACTION')	18800
670	FORMAT ((4(I4,2F8.3)))	18810
680	FORMAT ((1X,4(I4,2F8.3)))	18820
690	FORMAT (1X,'STEPS')	18830
700	FORMAT (8F10.3)	18840
710	FORMAT (1X,8G10.3)	18850
720	FORMAT (1X,'ELEMENTS')	18860
730	FORMAT (A9,1X,I2,3X,F10.0)	18870
740	FORMAT (1X,A9,2X,I2,3X,E16.5)	18880
750	FORMAT (1X,I4,2X,A8,5X,A8,' INDEX NUMBER IS OUT OF RANGE.'//1X,'CA	18890
	1LCULATION TERMINATED.')	18900
760	FORMAT (1X,'SPECIES')	18910
770	FORMAT (I3)	18920
780	FORMAT (1X,I3)	18930
790	FORMAT (A8,2X,I3,I1,I1,6F10.3)	18940
800	FORMAT (1X,A8,2X,I3,I1,I1,6F10.3)	18950
810	FORMAT (2F10.3,5E12.5)	18960
820	FORMAT (1X,2F10.3,5G12.5)	18970
830	FORMAT (6(I3,F7.3))	18980
840	FORMAT ((1X,6(I3,F7.3)))	18990
850	FORMAT (1X,'LOOK MIN')	19000
860	FORMAT (1X,'***** LOOK MIN ARRAY ALREADY HAS 40 MINERALS. *****')	19010
870	FORMAT (A8,2X,I2,13X,2F10.2,5X,I1)	19020

880	FORMAT (1X,A8,2X,I2,13X,2G10.3,5X,I1)	19030
890	FORMAT (1X,'SUMS')	19040
900	FORMAT (A8,2X,I2)	19050
910	FORMAT (1X,A8,2X,I2)	19060
920	FORMAT (20I4)	19070
930	FORMAT (1X,20I4)	19080
940	FORMAT (1X,'NEUTRAL')	19090
950	FORMAT (2I5)	19100
960	FORMAT (1X,2I5)	19110
970	FORMAT (/1X,14('*')/1X,'TERMINAL ERROR INPUT LACKED ',A8,' CARDS 1'/1X,14('*'))	19120 19130
980	FORMAT (1X,'MEAN GAM')	19140
990	FORMAT (1X,'***** MEAN GAM ARRARY ALREADY HAS 40 SALTS. *****')	19150
995	FORMAT (A8,I2,3(I4,F6.0))	19160
996	FORMAT (1X,A8,I2,3(I4,F6.2))	19170
	END	19180
	SUBROUTINE INOUT	19190
\$	INSERT COMMON.BLOCKS	19200
C		19210
C	DETERMINE TOTALS TO BE INCLUDED AS EQUATIONS IN MATRIX.	19220
C		19230
	DO 10 I=4,MAXT	19240
	IIN(I)=0	19250
	IF (TOT(I).LE.0.0D0) GO TO 10	19260
	IIN(I)=1	19270
10	CONTINUE	19280
	LAST=0	19290
	LASTT=0	19300
	DO 20 I=1,MAXT	19310
	IF (IIN(I).EQ.0) GO TO 20	19320
	LAST=LAST+1	19330
	IIN(I)=LAST	19340
	LASTT=I	19350
	IOUT(LAST)=I	19360
20	CONTINUE	19370
	ILT=LAST	19380
C	PICK SPECIES TO BE INCLUDED	19390
	IF (LASTT.LT.4) LASTT=4	19400
	SFLAG(1)=1	19410
	SFLAG(2)=1	19420
	SFLAG(3)=1	19430
	DO 30 I=4,MAXT	19440
	SFLAG(I)=0	19450
	IF (IIN(I).GT.0) SFLAG(I)=1	19460
30	CONTINUE	19470
	DO 50 I=MAXT1,MAXS	19480
	SFLAG(I)=0	19490
	K=NSP(I)	19500
	IF (K.EQ.0) GO TO 50	19510
	DO 40 J=1,K	19520
	IF (SFLAG(LSP(I,J)).LE.0) GO TO 50	19530
40	CONTINUE	19540
	SFLAG(I)=1	19550
	LASTS=I	19560
50	CONTINUE	19570
C	ZERO ARRAY AR	19580

DO 70 I=1,MAXEQ	19590
DO 60 J=1,MAXEQ	19600
AR(I,J)=ODO	19610
60 CONTINUE	19620
70 CONTINUE	19630
C PUT MINERAL COEFFICIENTS INTO FULL ARRAY	19640
IF (NMIN.LE.0) GO TO 120	19650
DO 90 I=1,NMIN	19660
K=NMIN(I)	19670
L=MAXT+I	19680
DO 80 J=1,K	19690
K1=LMIN(I,J)	19700
AR(K1,L)=-CMIN(I,J)	19710
AR(L,K1)=CMIN(I,J)	19720
80 CONTINUE	19730
AR(1,L)=ODO	19740
AR(2,L)=-THMIN(I)	19750
90 CONTINUE	19760
IF (IOPT(3).NE.6) GO TO 120	19770
C ADD REACTION TO EQUILBRATE WITH MINERAL # 1.	19780
L=MAXT+1	19790
DO 100 I=1,MAXT	19800
AR(I,L)=ODO	19810
100 CONTINUE	19820
DO 110 I=1,NCOMPS	19830
J=LREAC(I)	19840
AR(J,L)=AR(J,L)-CREAC(I)	19850
AR(2,L)=AR(2,L)-CREAC(I)*THMEAN(I)	19860
110 CONTINUE	19870
120 CONTINUE	19880
RETURN	19890
END	19900
SUBROUTINE KTEMP	19910
\$INSERT COMMON.BLOCKS	19920
DATA C/2.302585092/,R/1.98719D-3/	19930
C	19940
C COMPUTE TEMPERATURE DEPENDENCE OF A AND B FOR DEBYE-HUCKEL	19950
C	19960
S1=374.11-TC	19970
S2=S1**0.33333333	19980
S3=1.0+0.1342489*S2-3.946263E-03*S1	19990
S3=S3/(3.1975-0.3151548*S2-1.203374E-03*S1+7.48908E-13*S1**4.0)	20000
S3=DSQRT(S3)	20010
IF (TK.LT.373.15) GO TO 10	20020
C1=5321.0/TK+233.76-TK*(TK*(8.292E-07*TK-1.417E-03)+0.9297)	20030
GO TO 20	20040
10 C1=87.74-TC*(TC*(1.41E-06*TC-9.398E-04)+0.4008)	20050
20 CONTINUE	20060
C1=DSQRT(C1*TK)	20070
A=1824600.0*S3/C1**3.0	20080
B=50.29*S3/C1	20090
C	20100
C COMPUTE VANT HOFF CONSTANTS	20110
C	20120
C1=(298.15-TK)/(298.15*TK*C*R)	20130
DO 30 I=1,LASTS	20140

```

      IF (SFLAG(I).EQ.0) GO TO 30                                20150
      IF (KFLAG(I).EQ.1) LKSP(I)=ASP(I,1)+ASP(I,2)*TK+ASP(I,3)/TK+ASP(I, 20160
14)*DLOG10(TK)+ASP(I,5)/TK**2DO                                20170
      IF (KFLAG(I).EQ.0) LKSP(I)=LKTOSP(I)-DHSP(I)*C1          20180
30 CONTINUE                                                    20190
      IF (NLOOKS.EQ.0) GO TO 50                                20200
      DO 40 I=1,NLOOKS                                         20210
      IF (LOOKFL(I).EQ.0) LKLOOK(I)=LKOLK(I)-DHLOOK(I)*C1      20220
      IF (LOOKFL(I).EQ.1) LKLOOK(I)=ALOOK(I,1)+ALOOK(I,2)*TK+ALOOK(I,3)/ 20230
1TK+ALOOK(I,4)*DLOG10(TK)+ALOOK(I,5)/TK**2DO                  20240
40 CONTINUE                                                    20250
50 CONTINUE                                                    20260
      RETURN                                                    20270
C *****                                                    20280
      ENTRY KMINO                                              20290
C *****                                                    20300
      IF (NMINI.EQ.0) GO TO 80                                20310
      DO 70 I=1,NMINI                                          20320
      IF (MFLAG(I).EQ.0) LKMINO(I)=LKTOM(I)-DHMIN(I)*C1        20330
      IF (MFLAG(I).EQ.1) LKMINO(I)=AMIN(I,1)+AMIN(I,2)*TK+AMIN(I,3)/TK+A 20340
1MIN(I,4)*DLOG10(TK)+AMIN(I,5)/TK**2DO                        20350
C      CONVERT K TO MASTER SPECIES                            20360
      LKMIN(I)=LKMINO(I)+SIMIN(I)                             20370
      IF (NMCON(I).EQ.0) GO TO 70                             20380
      K=NMCON(I)                                               20390
      DO 60 J=1,K                                              20400
      LKMIN(I)=CMCON(I,J)*LKSP(LMCON(I,J))+LKMIN(I)           20410
60 CONTINUE                                                    20420
70 CONTINUE                                                    20430
80 CONTINUE                                                    20440
      RETURN                                                    20450
      END                                                       20460
      SUBROUTINE MINCON                                         20470
$INSERT COMMON.BLOCKS                                         20480
C      LOOP ON NUMBER OF MINERALS                             20490
      DO 100 I=1,NRMINS                                        20500
      K=NMINO(I)                                                20510
      ICOMP=0                                                  20520
      NONMAS=0                                                  20530
C      LOOP ON INDIVIDUAL MINERAL                             20540
      DO 90 J=1,K                                              20550
      IF (LMINO(I,J).GT.MAXT) GO TO 40                         20560
      IF (ICOMP.EQ.0) GO TO 20                                  20570
C      CHECK TO SEE IF COMPONENT IS ALREADY IN LIST          20580
      DO 10 I1=1,ICOMP                                         20590
      IF (LMIN(I,I1).EQ.LMINO(I,J)) GO TO 30                  20600
10 CONTINUE                                                    20610
20 CONTINUE                                                    20620
C      COMPONENT NOT IN LIST                                   20630
      ICOMP=ICOMP+1                                             20640
      LMIN(I,ICOMP)=LMINO(I,J)                                  20650
      CMIN(I,ICOMP)=CMINO(I,J)                                  20660
      GO TO 90                                                  20670
C      COMPONENT ALREADY IN LIST                               20680
30 CONTINUE                                                    20690
      CMIN(I,I1)=CMIN(I,I1)+CMINO(I,J)                        20700

```

	GO TO 90	20710
C	COMPONENT MUST BE REDUCED TO MASTER SPECIES	20720
40	CONTINUE	20730
	J1=LMINO(I,J)	20740
	K1=NSP(J1)	20750
	NONMAS=NONMAS+1	20760
	LMCON(I, NONMAS)=J1	20770
	CMCON(I, NONMAS)=-CMINO(I,J)	20780
C	LOOP THROUGH MASTER SPECIES OF NON-MASTER SPECIES COMPONENT	20790
	DO 80 I1=1,K1	20800
	IF (ICOMP.EQ.0) GO TO 60	20810
	DO 50 L1=1,ICOMP	20820
	IF (LSP(J1,I1).EQ.LMIN(I,L1)) GO TO 70	20830
50	CONTINUE	20840
C	MASTER SPECIES NOT IN LIST	20850
60	CONTINUE	20860
	ICOMP=ICOMP+1	20870
	LMIN(I,ICOMP)=LSP(J1,I1)	20880
	CMIN(I,ICOMP)=CMINO(I,J)*CSP(J1,I1)	20890
	GO TO 80	20900
C	MASTER SPECIES IN LIST	20910
70	CONTINUE	20920
	CMIN(I,L1)=CMIN(I,L1)+CMINO(I,J)*CSP(J1,I1)	20930
80	CONTINUE	20940
C	END LOOP MASTER SPECIES OF NON MASTER SPECIES COMPONENT	20950
90	CONTINUE	20960
	NMIN(I)=ICOMP	20970
	NMCON(I)=NONMAS	20980
C	END LOOP ON SINGLE MINERAL	20990
100	CONTINUE	21000
C	END ALL MINERALS	21010
	RETURN	21020
	END	21030
	SUBROUTINE PICK1	21040
\$	INSERT COMMON.BLOCKS	21050
C		21060
C	THIS SUBROUTINE CHOOSES THE MODEL FOR THE INITIAL SOLUTIONS	21070
C		21080
	DZOFF=ODO	21090
	IIN(1)=1	21100
	IIN(2)=0	21110
	IIN(3)=0	21120
	NMINS=0	21130
	IFE=1	21140
	ILE=1	21150
	IF (IESPEC.EQ.1) GO TO 10	21160
	IIN(1)=0	21170
	IFE=0	21180
	ILE=0	21190
10	CONTINUE	21200
C	CHECK ALKALINITY TOTAL	21210
	DALKT=ODO	21220
	IASPEC=IALK(ISOL)	21230
	IF (IASPEC.EQ.0) GO TO 30	21240
	IF (TOTAL(ISOL,IASPEC).GT.ODO) GO TO 20	21250
	WRITE (6,140)	21260

	ENDFILE (UNIT=6)	21270
	STOP 'ERROR IN ALKALINITY OPTIONS'	21280
20	CONTINUE	21290
	DALKT=TOTAL(ISOL,IASPEC)	21300
	IF (IESPEC.EQ.0.OR.IESPEC.NE.IASPEC) GO TO 30	21310
	WRITE (6,150)	21320
	ENDFILE (UNIT=6)	21330
	STOP 'ERROR IN ALKALINITY/NEUTRAL OPTIONS'	21340
30	CONTINUE	21350
	IF (IOPT(2).EQ.1.AND.IASPEC.GT.0) THEN	21360
	WRITE(6,146)	21370
146	FORMAT(' CAN NOT ADJUST PH AND ALKALINITY SIMULTANEOUSLY')	21380
	ENDFILE (UNIT=6)	21390
	STOP 'ERROR IN ALKALINITY/ELECTRICAL BALANCE OPTIONS'	21400
	END IF	21410
	DNEUT=0.0D0	21420
	IF (IESPEC.LT.4) GO TO 40	21430
C	MAKE SURE NEUTRAL SPECIES .GT. 0	21440
	DNEUT=TOT(IESPEC)	21450
	IF (TOT(IESPEC).LE.0D0) TOT(IESPEC)=1D-3	21460
40	CONTINUE	21470
	CALL INOUT	21480
C	SET COUNTERS ON THE ARRAY	21490
	IFTH=0	21500
	ILTH=0	21510
	IFT=ILE+1	21520
	ILT=ILT	21530
	IFM=0	21540
	ILM=0	21550
	NEQ=ILT	21560
	NEQ1=NEQ+1	21570
	RETURN	21580
C	*****	21590
	ENTRY PICK2	21600
C	*****	21610
C		21620
C	THIS SUBROUTINE CHOOSES THE MODEL FOR THE REACTION SOLUTIONS	21630
C		21640
	IIN(1)=1	21650
	IIN(2)=0	21660
	IF (IOPT(5).GT.0) IIN(2)=1	21670
	IIN(3)=1	21680
	IESPEC=1	21690
	IASPEC=0	21700
	NMINS=NRMINS	21710
C	ADD IN ANY MINERALS NOT INCLUDED IN TOTALS	21720
	IF (NMINS.EQ.0) GO TO 110	21730
	TOT(3)=1D0	21740
	ICK=0	21750
	DO 50 I=1,NMINS	21760
	TOT(MAXT+I)=0D0	21770
50	CONTINUE	21780
55	CONTINUE	21790
C	CHECK TO MAKE SURE NECESSARY TOTALS > 0.	21800
	JCK=0	21810
	ICK=ICK+1	21820

	DO 70 I=1,NMINS	21830
	IF (IOPT(3).EQ.6.AND.I.EQ.1) GO TO 70	21840
	K=NMIN(I)	21850
C	IF TOTALS > 0 NEXT MINERAL	21860
	DO 60 J=1,K	21870
	IF (LMIN(I,J).LT.4) GO TO 60	21880
	IF(TOT(LMIN(I,J)).GT.0D0) GO TO 60	21890
	GO TO 65	21900
60	CONTINUE	21910
	GO TO 70	21920
65	CONTINUE	21930
C	SOME TOTAL <= 0. IF ALL CMIN > 0 ADD MINERAL.	21940
	DO 75 J=1,K	21950
	IF (LMIN(I,J).LT.4) GO TO 75	21960
	IF (CMIN(I,J).GT.0) GO TO 75	21970
	GO TO 85	21980
75	CONTINUE	21990
	D=1D-6	22000
C	ADD D OF MINERAL # I.	22010
95	CONTINUE	22020
	TOT(MAXT+I)=D	22030
	THOR=THOR+THMIN(I)*D	22040
	DO 105 J=1,K	22050
	K1=LMIN(I,J)	22060
	IF(K1.LT.4)GO TO 105	22070
	TOT(K1)=TOT(K1)+D*CMIN(I,J)	22080
105	CONTINUE	22090
	GO TO 70	22100
C	SOME TOTAL <= 0. CHECK IF TOTALS <= 0 HAVE SAME SIGN FOR CMI	22110
85	CONTINUE	22120
	NCNEG=0	22130
	NCPOS=0	22140
	D=1D1	22150
	DO 115 J=1,K	22160
	K1=LMIN(I,J)	22170
	D1=TOT(K1)	22180
	IF (D1.LE.0D0) GO TO 125	22190
	IF (D1.LT.D) D=D1	22200
	GO TO 115	22210
125	CONTINUE	22220
	IF(CMIN(I,J).LE.0D0) NCNEG=1	22230
	IF(CMIN(I,J).GT.0D0) NCPOS=1	22240
115	CONTINUE	22250
	IF(NCNEG.GT.0.AND.NCPOS.GT.0) GO TO 135	22260
	D=D/5D1	22270
	IF(NCNEG.EQ.1) D=-D	22280
	GO TO 95	22290
135	CONTINUE	22300
	JCK=1	22310
70	CONTINUE	22320
	IF(JCK.EQ.0) GO TO 145	22330
	IF(ICK.LE.NMINS+1) GO TO 55	22340
	WRITE (6,155)	22350
155	FORMAT(1H1,90('*'))/'I HAD TROUBLE ADDING MINERALS SO THAT',	22360
1	' ALL THE NECESSARY ELEMENTS HAD POSITIVE CONCENTRATIONS.'/'	22370
2	'PLEASE ADJUST INITIAL CONCENTRATIONS TO SOME SMALL ',	22380

3	'POSITIVE QUANTITY.'/90('*'))	22390
	ENDFILE (UNIT=6)	22400
	STOP	22410
145	CONTINUE	22420
C	ADD ANY REACTION ELEMENTS IF IOPT(3)=6.	22430
	IF (IOPT(3).NE.6) GO TO 100	22440
	DO 90 I=1,NCOMPS	22450
	IF (TOT(LREAC(I)).GT.0D0) GO TO 90	22460
	D=DSIGN(1D-6,CREAC(I))	22470
	TOT(MAXT1)=D	22480
	DO 80 J=1,NCOMPS	22490
	THOR=THOR+CREAC(J)*THMEAN(J)*D	22500
	TOT(LREAC(J))=TOT(LREAC(J))+CREAC(J)*D	22510
80	CONTINUE	22520
	GO TO 100	22530
90	CONTINUE	22540
100	CONTINUE	22550
	TOT(1)=1D1**(-PH)	22560
	TOT(2)=1D1**(-PE)	22570
	TOT(3)=1D0	22580
110	CONTINUE	22590
	CALL INOUT	22600
C	SET COUNTERS FOR ARRAY	22610
	IFE=1	22620
	ILE=1	22630
	LAST=1	22640
	IFTH=0	22650
	ILTH=0	22660
	IF (IIN(2).LE.0) GO TO 120	22670
	IFTH=2	22680
	ILTH=2	22690
	LAST=2	22700
120	CONTINUE	22710
	LAST=LAST+1	22720
C	THIS IS FOR H2O EQUATION.	22730
	IFT=LAST+1	22740
	ILT=ILT	22750
	IFM=0	22760
	ILM=0	22770
	NEQ=ILT	22780
	IF (NMINSEQ.0) GO TO 130	22790
	IFM=ILT+1	22800
	ILM=ILT+NMINSEQ	22810
	NEQ=ILM	22820
130	CONTINUE	22830
	NEQ1=NEQ+1	22840
	RETURN	22850
C		22860
140	FORMAT (1X,'CARBON WAS SPECIFIED AS ALKALINITY BUT WAS LESS ', 'THA	22870
	1N OR EQUAL 0.0.'/1X,'CALCULATIONS TERMINATED')	22880
150	FORMAT (1X,'ALKALINITY SPECIES CAN NOT BE THE SAME AS THE ', 'NEUTR	22890
	1AL SPECIES'/1X,'CALCULATIONS TERMINATED')	22900
	END	22910
	SUBROUTINE STEP	22920
\$INSERT	COMMON.BLOCKS	22930
	LOGICAL EVAP	22940

	COMMON /EVAP/ EVAP	22950
	IF (IOPT(3).EQ.0) GO TO 120	22960
	I=IOPT(3)	22970
	WRITE (6,140) ISTEP	22980
	GO TO (10,30,50,60,100,100), I	22990
	GO TO 100	23000
C	MIX SOLUTION 1 WITH SOLUTION 2	23010
10	CONTINUE	23020
	X1=XSTEP(ISTEP)	23030
	X2=1.0-X1	23040
	DO 15 I=1,3	23050
	TOT(I)=(TOTAL(1,I)*TOTAL(2,I))**0.5	23060
15	CONTINUE	23070
	DO 20 I=4,MAXT	23080
	TOT(I)=X1*TOTAL(1,I)+X2*TOTAL(2,I)	23090
20	CONTINUE	23100
	THOR=TH(1)*X1+TH(2)*X2	23110
	TC=X1*TEMP(1)+X2*TEMP(2)	23120
	DZOFF=X1*DIFFZ(1)+X2*DIFFZ(2)	23130
	WRITE (6,150) X1,X2	23140
	GO TO 120	23150
C	TITRATE	23160
30	CONTINUE	23170
	X1=XSTEP(ISTEP)	23180
	VTOT=V0+X1	23190
	DO 40 I=1,MAXT	23200
	TOT(I)=(TOTAL(1,I)*V0+TOTAL(2,I)*X1)/VTOT	23210
40	CONTINUE	23220
	TC=(TEMP(1)*V0+TEMP(2)*X1)/VTOT	23230
	THOR=(TH(1)*V0+TH(2)*X1)/VTOT	23240
	DZOFF=(DIFFZ(1)*V0+DIFFZ(2)*X1)/VTOT	23250
	WRITE (6,160) X1	23260
	GO TO 120	23270
C	ADD REACTION IN INCREMENTS	23280
50	CONTINUE	23290
	X1=XSTEP(ISTEP)	23300
	GO TO 70	23310
C	ADD REACTION LINEARLY	23320
60	CONTINUE	23330
	X1=(ISTEP*XSTEP(1))/NSTEPS	23340
70	CONTINUE	23350
	DZOFF=DIFFZ(1)	23360
	DO 80 I=1,MAXT	23370
	TOT(I)=TOTAL(1,I)	23380
80	CONTINUE	23390
	DO 90 I=1,NCOMPS	23400
	K=LREAC(I)	23410
	IF (K.EQ.0) DZOFF=DIFFZ(1)+X1*CREAC(I)	23420
	IF (K.LT.4.OR.K.GT.MAXT) GO TO 90	23430
	TOT(K)=TOT(K)+X1*CREAC(I)	23440
90	CONTINUE	23450
	THOR=TH(1)+THREAC*X1	23460
	TC=TEMP(1)	23470
	IF (ISTEP.GT.1) TOT(1)=1D1**LA(1)	23480
	WRITE (6,170) X1	23490
	CALL PREAC	23500

	GO TO 120	23510
C	FOLLOW PHASE BOUNDARIES ONLY	23520
100	CONTINUE	23530
C		23540
C	SPECIAL CASE OF EVAPORATION	23550
C		23560
	IF (EVAP) THEN	23570
	NCOMPS=0	23580
	THREAC=ODO	23590
	DO 115 I=4,29	23600
	IF (TOTAL(1,I).GT.0.ODO) THEN	23610
	NCOMPS=NCOMPS+1	23620
	LREAC(NCOMPS)=I	23630
	CREAC(NCOMPS)=TOTAL(1,I)	23640
C	EVAPORATION DOES NOT CONSIDER REDOX	23650
	THMEAN(I)=ODO	23660
	END IF	23670
115	CONTINUE	23680
	END IF	23690
	DO 110 I=1,MAXT	23700
	TOT(I)=TOTAL(1,I)	23710
110	CONTINUE	23720
	THOR=TH(1)	23730
	TC=TEMP(1)	23740
	DZOFF=DIFFZ(1)	23750
	WRITE (6,180)	23760
120	CONTINUE	23770
C	TEMPERATURE	23780
	IF (IOPT(4).EQ.0) GO TO 130	23790
	IF (IOPT(4).EQ.1) TC=TSTEP(1)	23800
	IF (IOPT(4).EQ.2) TC=(ISTEP*(TSTEP(2)-TSTEP(1)))/NSTEPS+TSTEP(1)	23810
	IF (IOPT(4).EQ.3) TC=TSTEP(ISTEP)	23820
	WRITE (6,190) TC	23830
130	CONTINUE	23840
	TK=TC+273.15	23850
	RETURN	23860
C		23870
140	FORMAT (/1H1,'STEP NUMBER',I3,/1H0,14('-'))	23880
150	FORMAT (1H0,F10.3,' = FRACTION OF SOLUTION 1.',3X,F10.3,' = FR',	23890
	1'ACTION OF SOLUTION 2.'//)	23900
160	FORMAT (1H0,F10.3,' VOLUME UNITS OF SOLUTION 2 HAVE BEEN ',	23910
	1'ADDED.'//)	23920
170	FORMAT (1H0,1PD10.3,' MOLES OF REACTION HAVE BEEN ADDED.'//)	23930
180	FORMAT (//)	23940
190	FORMAT (23X,F10.2,' = NEW TEMPERATURE (C).')//)	23950
	END	23960
	SUBROUTINE THORIT(T)	23970
\$INSERT	COMMON.BLOCKS	23980
	T=0.ODO	23990
	DALKS=ODO	24000
	DO 10 I=1,LASTS	24010
	IF (SFLAG(I).EQ.0) GO TO 10	24020
	DALKS=DALKS+M(I)*ALKSP(I)	24030
	T=T+M(I)*THSP(I)	24040
10	CONTINUE	24050
	RETURN	24060

C	*****	24070
	ENTRY SOLN(J1)	24080
C	*****	24090
	WRITE (6,260) J1,(HEAD(J1,I),I=1,20)	24100
	DO 20 I=1,MAXT	24110
	TOT(I)=TOTAL(J1,I)	24120
20	CONTINUE	24130
	THOR=0.0D0	24140
	TC=TEMP(J1)	24150
	TK=TC+273.15	24160
	IF (DABS(TC-25.0).GT.30.0) WRITE (6,265)	24170
	PH=-DLOG10(TOT(1))	24180
	PE=-DLOG10(TOT(2))	24190
	RETURN	24200
C	*****	24210
	ENTRY SAVE	24220
C	*****	24230
	IF (IOPT(7).EQ.0) RETURN	24240
	K=IOPT(7)	24250
	DO 30 I=1,MAXT	24260
	TOTAL(K,I)=TOT(I)	24270
30	CONTINUE	24280
	TEMP(K)=TC	24290
	TH(K)=THOR	24300
	DIFFZ(K)=DZOFF	24310
	RETURN	24320
C	*****	24330
	ENTRY CHKSPE (IERR1)	24340
C	*****	24350
C	CHECK MASTER SPECIES IN AQUEOUS SPECIES	24360
	DO 80 I=1,MAXS	24370
	IF (NSP(I).LE.0) GO TO 80	24380
	K=NSP(I)	24390
	DO 70 J=1,K	24400
	L=LSP(I,J)	24410
	IF (L.GE.1.AND.L.LE.MAXT) GO TO 70	24420
	WRITE (6,270) I,SNAME(I)	24430
	IERR1=1	24440
70	CONTINUE	24450
80	CONTINUE	24460
C	CHECK SPECIES	24470
	DO 100 I=MAXT1,MAXS	24480
	K=NSP(I)	24490
	IF (K.EQ.0) GO TO 100	24500
	Z=0.0D0	24510
	DO 90 J=1,K	24520
	Z=Z+ZSP(LSP(I,J))*CSP(I,J)	24530
90	CONTINUE	24540
	IF (DABS(ZSP(I)-Z).LT.1D-5) GO TO 100	24550
	IERR1=1	24560
	WRITE (6,280) I,SNAME(I)	24570
100	CONTINUE	24580
	RETURN	24590
C	*****	24600
	ENTRY CHKMIN (IERR2)	24610
C	*****	24620

C	CHECK MINERALS FOR CHARGE AND VALENCE	24630
	IF (NRMINS.EQ.0) GO TO 140	24640
	THSP(2)=-1D0	24650
	DO 130 I=1,NRMINS	24660
	K=NMINO(I)	24670
	Z=0.0D0	24680
	THR=0.0D0	24690
	DO 110 J=1,K	24700
	Z=Z+ZSP(LMINO(I,J))*CMINO(I,J)	24710
	THR=THR+THSP(LMINO(I,J))*CMINO(I,J)	24720
110	CONTINUE	24730
	IF (DABS(Z).LT.1D-5) GO TO 120	24740
	IERR2=2	24750
	WRITE (6,290) I,MNAME(I)	24760
120	CONTINUE	24770
	IF (DABS(THR-THMIN(I)).LT.1D-5) GO TO 130	24780
	IERR2=3	24790
	WRITE (6,300) I,MNAME(I)	24800
130	CONTINUE	24810
	THSP(2)=0D0	24820
140	CONTINUE	24830
	RETURN	24840
C	*****	24850
	ENTRY CHKLK	24860
C	*****	24870
C	CHECK LOOK MINERALS	24880
	IF (NLOOKS.EQ.0) GO TO 170	24890
	DO 160 I=1,NLOOKS	24900
	K=NLOOK(I)	24910
	Z=0.0D0	24920
	DO 150 J=1,K	24930
	Z=Z+ZSP(LLOOK(I,J))*CLOOK(I,J)	24940
150	CONTINUE	24950
	IF (DABS(Z).LT.1D-5) GO TO 160	24960
	WRITE (6,310) I,NAMELK(I)	24970
160	CONTINUE	24980
170	CONTINUE	24990
	RETURN	25000
C	*****	25010
	ENTRY UNITS(ISOLN)	25020
C	*****	25030
	IASPEC=IALK(ISOLN)	25040
	IF (IUNITS(ISOLN).EQ.0) GO TO 250	25050
	IF (IUNITS(ISOLN).NE.1) GO TO 190	25060
C	MMOLES/L	25070
	DO 180 I=4,MAXT	25080
	TOTAL(ISOLN,I)=TOTAL(ISOLN,I)*GFW(I)	25090
180	CONTINUE	25100
	IUNITS(ISOLN)=2	25110
190	CONTINUE	25120
	IF (IUNITS(ISOLN).NE.2) GO TO 210	25130
C	MG/L	25140
	DO 200 I=4,MAXT	25150
	TOTAL(ISOLN,I)=TOTAL(ISOLN,I)/SDENS(ISOLN)	25160
200	CONTINUE	25170
	IUNITS(ISOLN)=3	25180

210	CONTINUE	25190
	IF (IUNITS(ISOLN).NE.3) GO TO 256	25200
C	PPM	25210
	TOTMG=0.0D0	25220
	DO 220 I=4,MAXT	25230
	TOTMG=TOTMG+TOTAL(ISOLN,I)	25240
220	CONTINUE	25250
	C=1.0-TOTMG*1E-06	25260
	IF (C.GT.0.0) GO TO 230	25270
	WRITE (6,320)	25280
	ENDFILE (UNIT=6)	25290
	STOP	25300
230	CONTINUE	25310
	C=1.0/(C*1E+03)	25320
	DO 240 I=4,MAXT	25330
	IF (TOTAL(ISOLN,I).LE.0.0D0) GO TO 240	25340
	TOTAL(ISOLN,I)=C*TOTAL(ISOLN,I)/GFW(I)	25350
240	CONTINUE	25360
	GO TO 250	25370
256	CONTINUE	25380
	IF(IUNITS(ISOLN).NE.4) GO TO 400	25390
C	MMOL/KG SOLN	25400
	DO 257 I=4,MAXT	25410
	IF(TOTAL(ISOLN,I).LE.0.0D0) GO TO 257	25420
	TOTAL(ISOLN,I) = TOTAL(ISOLN,I)*GFW(I)	25430
257	CONTINUE	25440
C	NOW IN PPM	25450
	IUNITS(ISOLN)=3	25460
	GO TO 210	25470
400	WRITE(6,410)	25480
	ENDFILE (UNIT=6)	25490
	STOP	25500
250	CONTINUE	25510
C	UNITS ARE NOW MOLALITY	25520
	RETURN	25530
C		25540
260	FORMAT (1H1,'SOLUTION NUMBER ',I1,/,1X,20A4/)	25550
265	FORMAT (/ ' WARNING -- TEMPERATURE IS MORE THAN 30 DEGREES FROM '	25560
	1,'THE REFERENCE',/,12X,'TEMPERATURE (25C); LARGE ERRORS IN '	25570
	2,'PITZER PARAMETERS',/,12X,'CAN BE EXPECTED IN THE CALCULATIONS '	25580
	3,'THAT FOLLOW.',/)	25590
270	FORMAT (1X,I4,2X,A8,5X,'SPECIES HAS A MASTER SPECIES NUMBER OUT OF	25600
	1 RANGE')	25610
280	FORMAT (1H , '***** ERROR IN SPECIES NUMBER',I4,' ',A8,'REACTION D	25620
	10ES NOT CHARGE BALANCE.')	25630
290	FORMAT (1H , '***** ERROR IN MINERAL NUMBER',I4,2X,A8,' REACTION DO	25640
	1ES NOT CHARGE BALANCE.')	25650
300	FORMAT (1H , '***** ERROR IN MINERAL NUMBER',I4,2X,A8,' REACTION',	25660
	1/,7X,'DOES NOT BALANCE THE SPECIFIED VALENCE.')	25670
310	FORMAT (1H , '***** WARNING: LOOK MINERAL NUMBER',I4,2X,A8,' REACT	25680
	1ION',/,7X,'DOES NOT CHARGE BALANCE.')	25690
320	FORMAT (1H1,80('*'))// 'TERMINAL ERROR'/'TOTAL SALT PPM GREATER ', 'T	25700
	1HAN 1 MILLION.'/'CHECK CONCENTRATION UNITS AND ELEMENT', ' GRAM FOR	25710
	2MULA WEIGHT INPUT',/,1X,80('*'))	25720
410	FORMAT(///,10X,'*** TERMINAL ERROR *** UNITS OF CONCENTRATION UNK	25730
	1NOWN',///)	25740

END	25750
SUBROUTINE ZEROAR	25760
\$INSERT COMMON.BLOCKS	25770
C	25780
DO 60 I=1,30	25790
DO 40 J=1,30	25800
AR(I,J)=0.0D0	25810
40 CONTINUE	25820
60 CONTINUE	25830
RETURN	25840
END	25850
C	25860
SUBROUTINE PITZER (MO, LG, AW, MU, ICON)	25870
IMPLICIT DOUBLE PRECISION (A-H, O-Z)	25880
CHARACTER *8 SPECS(40), NEUTRL(20)	25890
INTEGER TRANS(40), IN(20)	25900
DOUBLE PRECISION MU, MO(250), LG(250), LAM(40,20), MN(20), LGN(20), I	25910
1, LGAMMA(40), Z(40), PSI(40,40,40), BCX(4,20,21:40), M(40)	25920
LOGICAL IPRSNT(40), LNEUT	25930
COMMON / MX2 / I, J, K	25940
COMMON / MX3 / BCX, OTEMP	25950
COMMON / MX4 / A0	25960
COMMON / MX5 / Z	25970
COMMON / MX6 / LAM, TRANS, PSI, IN	25980
COMMON / MX0 / B	25990
COMMON / PI1C / SPECS, NEUTRL	26000
COMMON / PI1 / M1, M2, M3	26010
COMMON / COS / COSMOT	26020
DATA CONV / 0.4342944819D0 /	26030
C	26040
C THIS IS THE MAIN ROUTINE FOR CALCULATING GAMMAS USING HARVIE AND	26050
C WEARE (1980) MODEL. VALUES OF CONCENTRATION ARE PASSED INTO	26060
C SUBROUTINE PITZER VIA THE VARIABLE 'MO'. 'MO' CONTAINS MOLALITY	26070
C OF SPECIES, NUMBERED BY PHREEQE'S CONVENTION. ONCE INSIDE PITZER,	26080
C VALUES OF CONCENTRATION OF SPECIES USED IN PITZER'S MODEL ARE	26090
C TRANSFERED TO THE VARIABLES 'M' AND 'MN'. FOR A LIST OF SPECIES,	26100
C SEE SPECIES INPUT IN 'PITZER.DATA'.	26110
C	26120
C THE VARIABLE 'M' CONTAINS CONCENTRATIONS OF IONIZED SPECIES:	26130
C M(1:20) => CATIONS	26140
C M(21:40) => ANIONS	26150
C 'MN' CONTAINS CONCENTRATION OF NEUTRAL SPECIES	26160
C	26170
C VARIABLES 'SPECIES' AND 'NEUT' CONTAIN NAMES OF IONIZED AND	26180
C NEUTRAL SPECIES RESPECTIVELY, USING THE SAME NUMBERING SYSTEM AS	26190
C 'M' AND 'MN'.	26200
C	26210
C OTHER IMPORTANT VARIABLES:	26220
C M1 => MAX INDEX # FOR CATIONS (<21)	26230
C M2 => MAX INDEX # FOR ANIONS (20 < M2 < 41)	26240
C M3 => MAX INDEX # FOR NEUTRALS (<21)	26250
C IPRESENT, NEUTRAL => LOGICAL VARIABLE ARRAYS TO INDICATE	26260
C WHICH SPECIES HAVE NON-ZERO	26270
C CONCENTRATION.	26280
C I => IONIC STRENGTH	26290
C BCX(1,-,-) => BETA (0) SECOND SUBSCRIPT IS FOR CATION	26300

C	BCX(2,-,-) => BETA (1) INDEX;	26310
C	BCX(3,-,-) => BETA (2) THIRD SUBSCRIPT IS FOR ANION	26320
C	BCX(4,-,-) => C PHI INDEX.	26330
C	LAM => LAMBDA, IONIZED-NEUTRAL INTERACTION PARAMETER	26340
C	PSI => PSI, 1 CATION - 2 ANION INTERACTION OR 2 CATION -	26350
C	1 ANION INTERACTION PARAMETER.	26360
C	CONV => CONVERSION FACTOR FROM LOG BASE E TO LOG BASE 10	26370
C	LGAMMA => LOG OF GAMMAS	26380
C	PHIMAC => CONVERSION FACTOR FOR MACINNES CONVENTION	26390
C	COSMOTIC => OSMOTIC COEFFICIENT	26400
C	AW => ACTIVITY OF WATER	26410
C	GAMCLM => LOG GAMMA OF KCL SYSTEM AT SAME IONIC STRENGTH	26420
C	AND TEMPERATURE. USED BY 'PHIMAC'	26430
C	Z => CHANGE OF SPECIES	26440
C		26450
C	EQUATION NUMBERS REFER TO HARVIE AND WEARE (1980)	26460
C		26470
C	INITIALIZE	26480
C		26490
	XI=0.000	26500
	XX=0.000	26510
	OSUM=0.000	26520
	LNEUT=.FALSE.	26530
C		26540
C	TRANSFER DATA FROM MO TO M	26550
C		26560
	DO 20 N=1,M2	26570
	M(N)=0.000	26580
	IF (N.GT.M1.AND.N.LT.21) GO TO 20	26590
	M(N)=MO(TRANS(N))	26600
20	CONTINUE	26610
	DO 30 N=1,M3	26620
	MN(N)=MO(IN(N))	26630
	IF (MN(N).GT.1.0D-40) LNEUT=.TRUE.	26640
30	CONTINUE	26650
C		26660
	DO 40 N=1,M2	26670
	IF (M(N).GT.1.0D-40) GO TO 50	26680
	IPRSNT(N)=.FALSE.	26690
	GO TO 40	26700
50	IPRSNT(N)=.TRUE.	26710
40	CONTINUE	26720
C		26730
C	COMPUTE PITZER COEFFICIENTS' TEMPERATURE DEPENDENCE	26740
C		26750
	CALL PTEMP	26760
C		26770
	DO 10 N=1,M2	26780
	IF (.NOT.IPRSNT(N)) GO TO 10	26790
	XX=XX+M(N)*DABS(Z(N))	26800
	XI=XI+M(N)*Z(N)*Z(N)	26810
	OSUM=OSUM+M(N)	26820
10	CONTINUE	26830
	DO 15 N=1,M3	26840
	OSUM=OSUM+MN(N)	26850
15	CONTINUE	26860

	I=XI/2.0D0	26870
C		26880
C	EQUATION (8)	26890
C		26900
	BIGZ=XX	26910
	DI=DSQRT(I)	26920
C		26930
C	CALCULATE F & GAMCLM	26940
C		26950
	F=-A0*(DI/(1.0D0+B*DI)+2.0D0*DLOG(1.0D0+B*DI)/B)	26960
	XXX=2.0D0*DI	26970
	XXX=(1.0D0-(1.0D0+XXX-XXX*XXX*0.5D0)*DEXP(-XXX))/(XXX*XXX)	26980
	IK=ISPEC('K+')	26990
	IC=ISPEC('CL-')	27000
	GAMCLM=F+I*2.0D0*(BCX(1,IK,IC)+BCX(2,IK,IC)*XXX)+1.5D0	27010
	1*BCX(4,IK,IC)*I*I	27020
	DO 75 J=1,M1	27030
	IF (.NOT.IPRSNT(J)) GO TO 75	27040
	DO 70 K=21,M2	27050
	IF (.NOT.IPRSNT(K)) GO TO 70	27060
C		27070
C	EQUATION (3) PART 1	27080
C		27090
	F=F+M(J)*M(K)*BMXP()	27100
70	CONTINUE	27110
75	CONTINUE	27120
	DO 85 J=1,M1-1	27130
	IF (.NOT.IPRSNT(J)) GO TO 85	27140
	DO 80 K=J+1,M1	27150
	IF (.NOT.IPRSNT(K)) GO TO 80	27160
C		27170
C	EQUATION (3) PART 2	27180
C		27190
	F=F+M(J)*M(K)*ETHEAP()	27200
80	CONTINUE	27210
85	CONTINUE	27220
	DO 90 J=21,M2-1	27230
	IF (.NOT.IPRSNT(J)) GO TO 90	27240
	DO 110 K=J+1,M2	27250
	IF (.NOT.IPRSNT(K)) GO TO 110	27260
C		27270
C	EQUATION (3) PART 3	27280
C		27290
	F=F+M(J)*M(K)*ETHEAP()	27300
110	CONTINUE	27310
90	CONTINUE	27320
C		27330
	CSUM=0.0D0	27340
	DO 125 J=1,M1	27350
	IF (.NOT.IPRSNT(J)) GO TO 125	27360
	DO 120 K=21,M2	27370
	IF (.NOT.IPRSNT(K)) GO TO 120	27380
C		27390
C	EQUATION (2B) PART 4	27400
C		27410
	CSUM=CSUM+M(J)*M(K)*CMX()	27420

120	CONTINUE	27430
125	CONTINUE	27440
C		27450
C	CALCULATE LGAMMA FOR CATIONS	27460
C		27470
	DO 130 J=1,M1	27480
	LGAMMA(J)=Z(J)*Z(J)*F+DABS(Z(J))*CSUM	27490
	DO 140 K=21,M2	27500
	IF (.NOT.IPRSNT(K)) GO TO 140	27510
C		27520
C	EQUATION (2B) PART 3	27530
C		27540
	LGAMMA(J)=LGAMMA(J)+M(K)*(2.0D0*BMX()+BIGZ*CMX())	27550
140	CONTINUE	27560
	DO 150 K=1,M1	27570
	IF (.NOT.IPRSNT(K)) GO TO 150	27580
	LGAMMA(J)=LGAMMA(J)+2.0D0*M(K)*PHI()	27590
	DO 160 KK=21,M2	27600
	IF (.NOT.IPRSNT(KK)) GO TO 160	27610
C		27620
C	EQUATION (2B) PART 2	27630
C		27640
	LGAMMA(J)=LGAMMA(J)+M(KK)*M(K)*PSI(J,K,KK)	27650
160	CONTINUE	27660
150	CONTINUE	27670
	DO 170 K=21,M2-1	27680
	IF (.NOT.IPRSNT(K)) GO TO 170	27690
	DO 180 KK=K+1,M2	27700
	IF (.NOT.IPRSNT(KK)) GO TO 180	27710
C		27720
C	EQUATION (2B) PART 3	27730
C		27740
	LGAMMA(J)=LGAMMA(J)+M(K)*M(KK)*PSI(K,KK,J)	27750
180	CONTINUE	27760
170	CONTINUE	27770
	IF (.NOT.LNEUT) GO TO 130	27780
	DO 190 K=1,M3	27790
	LGAMMA(J)=LGAMMA(J)+2.0D0*MN(K)*LAM(J,K)	27800
190	CONTINUE	27810
130	CONTINUE	27820
C		27830
C	CALCULATE LGAMMA OF ANIONS	27840
C		27850
	DO 230 K=21,M2	27860
C		27870
C	EQUATION (2C) PART 1	27880
C		27890
	LGAMMA(K)=Z(K)*Z(K)*F+DABS(Z(K))*CSUM	27900
	DO 240 J=1,M1	27910
	IF (.NOT.IPRSNT(J)) GO TO 240	27920
C		27930
C	EQUATION (2C) PART 2	27940
C		27950
	LGAMMA(K)=LGAMMA(K)+M(J)*(2.0D0*BMX()+BIGZ*CMX())	27960
240	CONTINUE	27970
	DO 250 J=21,M2	27980

	IF (.NOT.IPRSNT(J)) GO TO 250	27990
	LGAMMA(K)=LGAMMA(K)+2.0D0*M(J)*PHI()	28000
	DO 260 KK=1,M1	28010
	IF (.NOT.IPRSNT(KK)) GO TO 260	28020
C		28030
C	EQUATION (2C) PART 3	28040
C		28050
	LGAMMA(K)=LGAMMA(K)+M(KK)*M(J)*PSI(K,J,KK)	28060
260	CONTINUE	28070
250	CONTINUE	28080
	DO 270 J=1,M1-1	28090
	IF (.NOT.IPRSNT(J)) GO TO 270	28100
	DO 280 KK=J+1,M1	28110
	IF (.NOT.IPRSNT(KK)) GO TO 280	28120
C		28130
C	EQUATION (2C) PART 4	28140
C		28150
	LGAMMA(K)=LGAMMA(K)+M(J)*M(KK)*PSI(J,KK,K)	28160
280	CONTINUE	28170
270	CONTINUE	28180
	IF (.NOT.LNEUT) GO TO 230	28190
	DO 290 J=1,M3	28200
	LGAMMA(K)=LGAMMA(K)+2.0D0*MN(J)*LAM(K,J)	28210
290	CONTINUE	28220
230	CONTINUE	28230
C		28240
C	CONVERT TO MACINNES CONVENTION	28250
C		28260
	IF (ICON.EQ.0) GO TO 300	28270
	PHIMAC=GAMCLM-LGAMMA(IC)	28280
	DO 220 K=1,M2	28290
	IF (.NOT.IPRSNT(K)) GO TO 220	28300
	LGAMMA(K)=LGAMMA(K)+Z(K)*PHIMAC	28310
220	CONTINUE	28320
C		28330
300	IF (.NOT.LNEUT) GO TO 860	28340
C		28350
C	CALCULATE THE GAMMA OF NEUTRAL IONS	28360
C		28370
	DO 800 K=1,M3	28380
	LGN(K)=0.0D0	28390
	DO 870 J=1,M2	28400
	IF (.NOT.IPRSNT(J)) GO TO 870	28410
	LGN(K)=LGN(K)+2.0D0*M(J)*LAM(J,K)	28420
870	CONTINUE	28430
800	CONTINUE	28440
C		28450
C	CALCULATE THE OSMOTIC COEFFICIENT	28460
C		28470
C	EQUATION (2A) PART 1	28480
C		28490
860	OSMOT=- (A0)*I**1.5D0/(1.0D0+B*DI)	28500
	DO 420 J=1,M1	28510
	IF (.NOT.IPRSNT(J)) GO TO 420	28520
	DO 430 K=21,M2	28530
	IF (.NOT.IPRSNT(K)) GO TO 430	28540

C		28550
C	EQUATION (2A) PART 2	28560
C		28570
	OSMOT=OSMOT+M(J)*M(K)*(BMXPHI()+BIGZ*CMX())	28580
430	CONTINUE	28590
420	CONTINUE	28600
	DO 440 J=1,M1-1	28610
	IF (.NOT.IPRSNT(J)) GO TO 440	28620
	DO 450 K=J+1,M1	28630
	IF (.NOT.IPRSNT(K)) GO TO 450	28640
	OSMOT=OSMOT+M(J)*M(K)*PHIPHI()	28650
	DO 460 KK=21,M2	28660
	IF (.NOT.IPRSNT(KK)) GO TO 460	28670
		28680
C		28690
C	EQUATION (2A) PART 3	28700
C		28710
	OSMOT=OSMOT+M(J)*M(K)*M(KK)*PSI(J,K,KK)	28720
460	CONTINUE	28730
450	CONTINUE	28740
440	CONTINUE	28750
	DO 470 J=21,M2-1	28760
	IF (.NOT.IPRSNT(J)) GO TO 470	28770
	DO 480 K=J+1,M2	28780
	IF (.NOT.IPRSNT(K)) GO TO 480	28790
	OSMOT=OSMOT+M(J)*M(K)*PHIPHI()	28800
	DO 490 KK=1,M1	28810
	IF (.NOT.IPRSNT(KK)) GO TO 490	28820
		28830
C		28840
C	EQUATION (2A) PART 4	28850
C		28860
	OSMOT=OSMOT+M(J)*M(K)*M(KK)*PSI(J,K,KK)	28870
490	CONTINUE	28880
480	CONTINUE	28890
470	CONTINUE	28900
	IF (.NOT.LNEUT) GO TO 850	28910
	DO 810 K=1,M3	28920
	DO 820 J=1,M2	28930
	IF (.NOT.IPRSNT(J)) GO TO 820	28940
		28950
C		28960
C	EQUATION (A.3A) PART 5 HARVIE, MOLLER, WEARE (1984)	28970
C		28980
	OSMOT=OSMOT+MN(K)*M(J)*LAM(J,K)	28990
820	CONTINUE	29000
810	CONTINUE	29010
850	COSMOT=1.0DO+2.0DO*OSMOT/OSUM	29020
		29030
C		29040
C	CALCULATE THE ACTIVITY OF WATER	29050
C		29060
	AW=DEXP(-OSUM*COSMOT/55.50837DO)	29070
		29080
C		29090
C	SET APPROPRIATE VALUES FOR RETURN	29100
C		
	MU=I	
	DO 900 N=1,M2	
	IF (.NOT.IPRSNT(N)) GO TO 900	
	LG(TRANS(N))=LGAMMA(N)*CONV	

900	CONTINUE	29110
	IF (.NOT.LNEUT) RETURN	29120
	DO 910 N=1,M3	29130
	LG(IN(N))=LGN(N)*CONV	29140
910	CONTINUE	29150
	RETURN	29160
	END	29170
C		29180
C	THESE FUNCTIONS CALCULATE THE BM'S	29190
C		29200
	DOUBLE PRECISION FUNCTION BMXPHI()	29210
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	29220
	COMMON / MX2 / I,J,K	29230
	COMMON / MX3 / BCX,OTEMP	29240
C		29250
C	ALPHA DEFINED IN BLOCK DATA	29260
C		29270
	COMMON / MX9 / ALPHA	29280
	DIMENSION ALPHA (5),BCX(4,20,21:40)	29290
	DOUBLE PRECISION I	29300
	LOGICAL TWOTWO,ONEANY	29310
C		29320
C	STATEMENT FUNCTIONS	29330
C		29340
C	EQUATION (5)	29350
C		29360
	G(Y)=2.0D0*(1.0D0-(1.0D0+Y)*DEXP(-Y))/Y**2.0D0	29370
C		29380
C	EQUATION (6)	29390
C		29400
	GP(Y)=-2.0D0*(1.0D0-(1.0D0+Y+Y**2.0D0/2.0D0)*DEXP(-Y))/	29410
	1 Y**2.0D0	29420
C		29430
	IF (ONEANY()) GO TO 40	29440
	IF (TWOTWO()) GO TO 50	29450
C		29460
C	EQUATION (7A)	29470
C		29480
	BMXPHI=BCX(1,J,K)+BCX(2,J,K)*DEXP(-ALPHA(4)*DSQRT(I))+BCX(3,J,K)*	29490
	1 DEXP(-ALPHA(5)*DSQRT(I))	29500
	RETURN	29510
C		29520
C	EQUATION (4A)	29530
C		29540
40	BMXPHI=BCX(1,J,K)+BCX(2,J,K)*DEXP(-ALPHA(1)*DSQRT(I))	29550
	RETURN	29560
C		29570
C	EQUATION (7A)	29580
C		29590
50	BMXPHI=BCX(1,J,K)+BCX(2,J,K)*DEXP(-ALPHA(2)*DSQRT(I))+BCX(3,J,K)*	29600
	1 DEXP(-ALPHA(3)*DSQRT(I))	29610
	RETURN	29620
C	*****	29630
	ENTRY BMX	29640
C	*****	29650
	IF (ONEANY()) GO TO 55	29660

	IF (TWO TWO()) GO TO 60	29670
C		29680
C	EQUATION (7B)	29690
C		29700
	BMX=BCX(1,J,K)+BCX(2,J,K)*G(ALPHA(4)*DSQRT(I))+BCX(3,J,K)*	29710
	1 G(ALPHA(5)*DSQRT(I))	29720
	RETURN	29730
C		29740
C	EQUATION (4B)	29750
C		29760
	55 BMX=BCX(1,J,K)+BCX(2,J,K)*G(ALPHA(1)*DSQRT(I))	29770
	RETURN	29780
C		29790
C	EQUATION (7B)	29800
C		29810
	60 BMX=BCX(1,J,K)+BCX(2,J,K)*G(ALPHA(2)*DSQRT(I))+BCX(3,J,K)*	29820
	1 G(ALPHA(3)*DSQRT(I))	29830
	RETURN	29840
C	*****	29850
	ENTRY BMXP	29860
C	*****	29870
	IF (ONEANY()) GO TO 65	29880
	IF (TWO TWO()) GO TO 70	29890
C		29900
C	EQUATION (7C)	29910
C		29920
	BMXP=(BCX(2,J,K)*GP(ALPHA(4)*DSQRT(I))+BCX(3,J,K)*GP(ALPHA(5)*	29930
	1 DSQRT(I)))/I	29940
	RETURN	29950
C		29960
C	EQUATION (4C)	29970
C		29980
	65 BMXP=BCX(2,J,K)*GP(ALPHA(1)*DSQRT(I))/I	29990
	RETURN	30000
C		30010
C	EQUATION (7C)	30020
C		30030
	70 BMXP=(BCX(2,J,K)*GP(ALPHA(2)*DSQRT(I))+BCX(3,J,K)*GP(ALPHA(3)*	30040
	1 DSQRT(I)))/I	30050
	RETURN	30060
	END	30070
C		30080
C	SUBROUTINE TO DETERMINE ONE-ANY ELECTROLYTES	30090
C		30100
C	ONEANY = .TRUE. IF EITHER ION IS UNIVALENT	30110
C		30120
	LOGICAL FUNCTION ONEANY()	30130
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	30140
	COMMON / MX2 / I,J,K	30150
	COMMON / MX5 / Z	30160
	DOUBLE PRECISION I	30170
	DIMENSION Z(40)	30180
	ONEANY=.FALSE.	30190
	IF (DABS(Z(J)).LT.2.0D0.AND.DABS(Z(J)).GT.0.0D0) ONEANY=.TRUE.	30200
	IF (DABS(Z(K)).LT.2.0D0.AND.DABS(Z(K)).GT.0.0D0) ONEANY=.TRUE.	30210
	RETURN	30220

	END	30230
C		30240
C	SUBROUTINE TO DETERMINE TWO-TWO ELECTROLYTES	30250
C		30260
C	TWOTWO = .TRUE. IF BOTH IONS ARE DOUBLY CHARGED	30270
C		30280
	LOGICAL FUNCTION TWOTWO()	30290
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	30300
	COMMON / MX2 / I,J,K	30310
	COMMON / MX5 / Z	30320
	DOUBLE PRECISION I	30330
	DIMENSION Z(40)	30340
	TWOTWO=.FALSE.	30350
	ZDIF=DABS(DABS(Z(J)*Z(K))-4.0D0)	30360
	IF (ZDIF.LT.1.0D0) TWOTWO=.TRUE.	30370
	RETURN	30380
	END	30390
C		30400
C	FUNCTION TO CALCULATE C(MX)	30410
C		30420
	DOUBLE PRECISION FUNCTION CMX()	30430
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	30440
	COMMON / MX2 / I,J,K	30450
	COMMON / MX3 / BCX,OTEMP	30460
	COMMON / MX5 / Z	30470
	DOUBLE PRECISION I,Z(40),BCX(4,20,21:40)	30480
C		30490
C	EQUATION (9)	30500
C		30510
	CMX=BCX(4,J,K)/(2.0D0*DSQRT(DABS(Z(J)*Z(K))))	30520
	RETURN	30530
	END	30540
C		30550
C	FUNCTIONS TO CALCULATE ETHETA AND ETHEAP	30560
C		30570
	DOUBLE PRECISION FUNCTION ETHETA()	30580
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	30590
	COMMON / MX2 / I,J,K	30600
	COMMON / MX4 / A0	30610
	COMMON / MX5 / Z	30620
	DIMENSION Z(40)	30630
	DOUBLE PRECISION I,JAY,JPRIME	30640
	IFLAG=1	30650
	ETHETA=0.0D0	30660
	GO TO 10	30670
C	*****	30680
	ENTRY ETHEAP	30690
C	*****	30700
	IFLAG=2	30710
	ETHEAP=0.0D0	30720
10	XCON=6.0D0*A0*DSQRT(I)	30730
	IF (DABS(Z(J)-Z(K)).LE.1.0D-40) RETURN	30740
	ZZ=Z(J)*Z(K)	30750
C		30760
C	NEXT 3 ARE EQUATION (A1)	30770
C		30780

	XJK=XCON*ZZ	30790
	XJJ=XCON*Z(J)*Z(J)	30800
	XKK=XCON*Z(K)*Z(K)	30810
C		30820
C	EQUATION (A2)	30830
C		30840
	ETHETA=ZZ*(JAY(XJK)-JAY(XJJ)/2.0D0-JAY(XKK)/2.0D0)/(4.0D0*I)	30850
	IF (IFLAG.EQ.1) RETURN	30860
C		30870
C	EQUATION (A3)	30880
C		30890
	ETHEAP=ZZ*(JPRIME(XJK)-JPRIME(XJJ)/2.0D0-JPRIME(XKK)/	30900
	1 2.0D0)/(8.0D0*I**2.0D0) - ETHETA/I	30910
	RETURN	30920
	END	30930
C		30940
C	FUNCTION TO CALCULATE JAY AND JPRIME	30950
C		30960
C	J0 AND J1, USED IN CALCULATION OF ETHETA AND ETHEAP	30970
C		30980
	DOUBLE PRECISION FUNCTION JAY(X)	30990
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	31000
	COMMON / MX8 / AK,BK,DK	31010
	DIMENSION AK(0:20,2),BK(0:22),DK(0:22)	31020
	DOUBLE PRECISION JPRIME	31030
	CALL BDK (X)	31040
	JAY=X/4.0D0-1.0D0+0.5D0*(BK(0)-BK(2))	31050
	RETURN	31060
C	*****	31070
	ENTRY JPRIME(Y)	31080
C	*****	31090
	CALL BDK (Y)	31100
	IF (Y.GT.1.0D0) GO TO 10	31110
	DZ=0.8D0*Y**(-0.8D0)	31120
	GO TO 20	31130
10	DZ=-4.0D0*Y**(-1.1D0)/9.0D0	31140
20	JPRIME=Y*(.25D0+DZ*(DK(0)-DK(2))/2.0D0)	31150
	RETURN	31160
	END	31170
C		31180
	SUBROUTINE BDK (X)	31190
C		31200
C	NUMERICAL APPROXIMATION TO THE INTEGRALS IN THE EXPRESSIONS FOR J0	31210
C	AND J1. CHEBYSHEV APPROXIMATION IS USED. THE CONSTANTS 'AK' ARE	31220
C	DEFINED IN BLOCK COMMON.	31230
C		31240
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	31250
	COMMON / MX8 / AK,BK,DK	31260
	DIMENSION AK(0:20,2),BK(0:22),DK(0:22)	31270
	II=1	31280
	IF (X.LE.1.0D0) GO TO 10	31290
	II=2	31300
	Z=40.0*X**(-1.0D-1)/9.0D0-22.0D0/9.0D0	31310
	GO TO 20	31320
10	Z=4.0D0*X**(0.2D0)-2.0D0	31330
20	DO 30 M=20,0,-1	31340

	BK(M)=Z*BK(M+1)-BK(M+2)+AK(M,II)	31350
	DK(M)=BK(M+1)+Z*DK(M+1)-DK(M+2)	31360
30	CONTINUE	31370
	RETURN	31380
	END	31390
C		31400
C	FUNCTION TO CALCULATE PHI (PHI' IS EQUAL TO ETHEAP)	31410
C		31420
	DOUBLE PRECISION FUNCTION PHI()	31430
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	31440
	COMMON / MX2 / I,J,K	31450
	COMMON / MX7 / THETA	31460
	DIMENSION THETA(40,40)	31470
	DOUBLE PRECISION I	31480
C		31490
C	EQUATION (10B)	31500
C		31510
	PHI=THETA(J,K)+ETHETA()	31520
	RETURN	31530
C	*****	31540
	ENTRY PHIPHI	31550
C	*****	31560
C		31570
C	EQUATION (10A)	31580
C		31590
	PHIPHI=THETA(J,K)+ETHETA()+I*ETHEAP()	31600
	RETURN	31610
	END	31620
C		31630
C	SUBROUTINE TO CALUCLATE TEMPERATURE DEPENDENCE OF PITZER PARAMETER	31640
C		31650
	SUBROUTINE PTEMP	31660
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	31670
\$	INSERT COMMON.BLOCKS	31680
	DOUBLE PRECISION BC(4,20,21:40,5),BCX(4,20,21:40)	31690
	CHARACTER *8 SPECS(40),NEUTRL(20)	31700
	COMMON / MX1 / BC	31710
	COMMON / MX3 / BCX,OTEMP	31720
	COMMON / MX4 / AO	31730
	COMMON / MX10 / VP,DWO	31740
	COMMON / PI1C / SPECS,NEUTRL	31750
	COMMON / PI1 / M1,M2,M3	31760
	DATA TR /298.15/	31770
C		31780
	IF (DABS(TK-OTEMP).LT.0.01) RETURN	31790
	OTEMP=TK	31800
	DWO=DW(TK)	31810
	IF (DABS(TK-TR).LT.0.01) GO TO 10	31820
	DCO=DC(TK)	31830
	AO=1.400684D6*(DWO/(DCO*TK)**3.0D0)**0.5D0	31840
	DO 30 L=1,4	31850
	DO 40 K=1,M1	31860
	DO 50 N=21,M2	31870
	BCX(L,K,N)=BC(L,K,N,1)+BC(L,K,N,2)*(1.0/TK-1.0/TR)+BC(L,K,N,3)	31880
	1*DLOG(TK/TR)+BC(L,K,N,4)*(TK-TR)+BC(L,K,N,5)*(TK*TK-TR*TR)	31890
50	CONTINUE	31900

40	CONTINUE	31910
30	CONTINUE	31920
	RETURN	31930
10	DO 20 L=1,4	31940
	DO 60 K=1,M1	31950
	DO 70 N=21,M2	31960
	BCX(L,K,N)=BC(L,K,N,1)	31970
70	CONTINUE	31980
60	CONTINUE	31990
20	CONTINUE	32000
	AO=0.392D0	32010
	RETURN	32020
	END	32030
C		32040
C	SUBROUTINE TO INITIALIZE DATA	32050
C		32060
	SUBROUTINE INITPZ	32070
C		32080
C	'PITZER.DATA' IS READ IN THIS SUBROUTINE.	32090
C		32100
C	SOME VARIABLES:	32110
C	IS => TOTAL NUMBER OF SPECIES USED IN PITZER'S MODEL	32120
C	NS => INDIVIDUAL SPECIES NUMBERS (IN PHREEQE'S NUMBERING	32130
C	SYSTEM) USED IN PITZER'S MODEL.	32140
C	TRANS,IN => ARRAYS TO TRANSLATE PHREEQE'S NUMBERING SYSTEM	32150
C	INTO NUMBERING SYSTEM USED IN SUBROUTINE PITZER.	32160
C		32170
	PARAMETER (IF=12)	32180
\$INSERT	COMMON.BLOCKS	32190
	INTEGER TRANS(40),IN(20),NS(40)	32200
	CHARACTER *8 SPEC1,SPEC2,SPEC3,NEUTRL(20),SPECS(40)	32210
	DOUBLE PRECISION PSI(40,40,40),BC(4,20,21:40,5),THETA(40,40),Z(40)	32220
	1,LAM(40,20),VALUE(5)	32230
	COMMON / MX1 / BC	32240
	COMMON / MX5 / Z	32250
	COMMON / MX6 / LAM,TRANS,PSI,IN	32260
	COMMON / MX7 / THETA	32270
	COMMON / PI1C / SPECS,NEUTRL	32280
	COMMON / PI1 / M1,M2,M3	32290
	OPEN (UNIT=IF,FILE='PITZER.DATA')	32300
	READ (IF,10) SPEC1,IS	32310
10	FORMAT (A8,I3)	32320
	READ (IF,20) (NS(J),J=1,IS)	32330
20	FORMAT (25(I3))	32340
	DO 90 J=1,IS	32350
	IF (NS(J).GT.0.AND.NS(J).LT.251) GO TO 40	32360
	WRITE (6,30)	32370
30	FORMAT (//,' PITZER DATA FILE ERROR.')	32380
	STOP	32390
40	IF (DABS(ZSP(NS(J))).LE.1.0D-40) GO TO 80	32400
	IF (ZSP(NS(J)).LT.-1.0D-5) GO TO 60	32410
C		32420
C	CATIONS (UP TO 20)	32430
C		32440
	M1=M1+1	32450
	IF (M1.LE.20) GO TO 58	32460

53	WRITE (6,55)	32470
55	FORMAT (' ***** ERROR: ONLY 20 IONS OF EACH TYPE ARE ALLOWED.')	32480
	STOP	32490
58	TRANS(M1)=NS(J)	32500
	SPECS(M1)=SNAME(NS(J))	32510
	Z(M1)=ZSP(NS(J))	32520
	GO TO 90	32530
C		32540
C	ANIONS (UP TO 20)	32550
C		32560
60	M2=M2+1	32570
	IF (M2.GT.40) GO TO 53	32580
	TRANS(M2)=NS(J)	32590
	SPECS(M2)=SNAME(NS(J))	32600
	Z(M2)=ZSP(NS(J))	32610
	GO TO 90	32620
C		32630
C	NEUTRAL SPECIES (UP TO 20)	32640
C		32650
80	M3=M3+1	32660
	IF (M3.GT.20) GO TO 53	32670
	IN(M3)=NS(J)	32680
	NEUTRL(M3)=SNAME(NS(J))	32690
90	CONTINUE	32700
C		32710
	READ (IF,95)	32720
95	FORMAT (1X)	32730
	K=1	32740
C		32750
C	READ IN BETA(0), BETA(1), BETA(2), AND C(PHI)	32760
C		32770
100	READ (IF,105) SPEC1,SPEC2,(VALUE(J),J=1,5)	32780
105	FORMAT (2(2X,A8),5(1X,F11.0))	32790
	IF (DABS(VALUE(1)).GT.1.0D-30) GO TO 107	32800
	K=K+1	32810
	IF (K.GT.4) GO TO 120	32820
	GO TO 100	32830
107	K1=ISPEC(SPEC1)	32840
	K2=ISPEC(SPEC2)	32850
	DO 110 J=1,5	32860
	BC(K,K1,K2,J)=VALUE(J)	32870
110	CONTINUE	32880
	GO TO 100	32890
120	READ (IF,105) SPEC1,SPEC2,VALUE(1)	32900
	IF (DABS(VALUE(1)).LE.1.0D-40) GO TO 130	32910
	I1=ISPEC(SPEC1)	32920
	I2=ISPEC(SPEC2)	32930
	THETA(I1,I2)=VALUE(1)	32940
	THETA(I2,I1)=VALUE(1)	32950
	GO TO 120	32960
C		32970
C	READ IN LAMBDA	32980
C		32990
130	READ (IF,105) SPEC1,SPEC2,VALUE(1)	33000
	IF (DABS(VALUE(1)).LE.1.0D-40) GO TO 140	33010
	LAM(ISPEC(SPEC1),ISPEC(SPEC2))=VALUE(1)	33020

	GO TO 130	33030
C	INITIALIZE PSI (TOO BIG TO DO IT IN BLOCK DATA)	33040
140	I1=1	33050
	I2=1	33060
	I3=1	33070
142	PSI(I1,I2,I3)=0.0D0	33080
	I1=I1+1	33090
	IF (I1.GT.M2) GO TO 143	33100
	IF (I1.GT.M1.AND.I1.LT.21) I1=21	33110
	GO TO 142	33120
143	I1=1	33130
	I2=I2+1	33140
	IF (I2.GT.M2) GO TO 144	33150
	IF (I2.GT.M1.AND.I2.LT.21) I2=21	33160
	GO TO 142	33170
144	I2=1	33180
	I3=I3+1	33190
	IF (I3.GT.M2) GO TO 145	33200
	IF (I3.GT.M1.AND.I3.LT.21) I3=21	33210
	GO TO 142	33220
C		33230
C	READ IN PSI	33240
C		33250
145	READ (IF,150,END=160) SPEC1,SPEC2,SPEC3,VALUE(1)	33260
150	FORMAT (3(2X,A8),2X,F10.0)	33270
	I1=ISPEC(SPEC1)	33280
	I2=ISPEC(SPEC2)	33290
	I3=ISPEC(SPEC3)	33300
	PSI(I1,I2,I3)=VALUE(1)	33310
	PSI(I1,I3,I2)=VALUE(1)	33320
	PSI(I2,I1,I3)=VALUE(1)	33330
	PSI(I2,I3,I1)=VALUE(1)	33340
	PSI(I3,I1,I2)=VALUE(1)	33350
	PSI(I3,I2,I1)=VALUE(1)	33360
	GO TO 145	33370
160	CLOSE (UNIT=IF)	33380
	RETURN	33390
	END	33400
C		33410
C	FUNCTION TO TRANSFORM SPECIES NAME INTO SPECIES NUMBER	33420
C		33430
	FUNCTION ISPEC(SPEC)	33440
	CHARACTER *8 SPEC,SPECS(40),NEUTRL(20)	33450
	COMMON / PI1C / SPECS,NEUTRL	33460
	COMMON / PI1 / M1,M2,M3	33470
	DO 10 I=1,M2	33480
	IF (I.GT.M1.AND.I.LT.21) GO TO 10	33490
	IF (SPEC.EQ.SPECS(I)) GO TO 20	33500
10	CONTINUE	33510
	DO 40 I=1,M3	33520
	IF (SPEC.EQ.NEUTRL(I)) GO TO 20	33530
40	CONTINUE	33540
	WRITE (6,30) SPEC	33550
30	FORMAT (' ***** ERROR IN DATA BASE ***** ',A8,' IS NOT A '	33560
	1,'VALID ION.')	33570
	STOP	33580

20	ISPEC=I	33590
	RETURN	33600
	END	33610
C		33620
C	SUBROUTINE TO CALCULATE THE DENSITY OF WATER AS A FUNCTION OF	33630
C	TEMPERATURE. T IS IN KELVIN, P IS IN PASCALS, DW IS IN G/CM^3	33640
C		33650
C	FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	33660
C		33670
	DOUBLE PRECISION FUNCTION DW (T)	33680
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	33690
	COMMON / MX10 / VP,DWO	33700
	COMMON /ACONST/ GASCON,TZ,AA,Z,DZ,Y	33710
	DATA FP / 9.869232667D0 /	33720
	CALL BB (T)	33730
	P=1.0D0/FP	33740
	IF (T.GT.373.149D0) P=PS(T)	33750
	DGSS=P/T/.4D0	33760
	IF (T.GE.TZ) GO TO 10	33770
	DGSS=1.0D0/(VLEST(T))	33780
10	CALL DFIND (D,P,DGSS,T)	33790
	DW=D	33800
	VP=P*FP	33810
	RETURN	33820
	END	33830
C		33840
C	THIS SUBROUTINE CALCULATES THE B'S NEEDED FOR FUNCTION DW.	33850
C	THE B'S CALCULATED HERE ARE IN CM3/G.	33860
C		33870
C	FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	33880
C		33890
	SUBROUTINE BB(T)	33900
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	33910
	COMMON /ELLCON/ G1,G2,GF,B1,B2,B1T,B2T,B1TT,B2TT	33920
	COMMON /ACONST/ GASCON,TZ,AA,Z,DZ,Y	33930
	COMMON /BCONST/ P(10),Q(10)	33940
	DIMENSION V(10)	33950
	V(1)=1.D0	33960
	DO 2 I=2,10	33970
2	V(I)=V(I-1)*TZ/T	33980
	B1=P(1)+P(2)*DLOG(1.D0/V(2))	33990
	B2=Q(1)	34000
	B1T=P(2)*V(2)/TZ	34010
	B2T=0.D0	34020
	B1TT=0.D0	34030
	B2TT=0.D0	34040
	DO 4 I=3,10	34050
	B1=B1+P(I)*V(I-1)	34060
	B2=B2+Q(I)*V(I-1)	34070
	B1T=B1T-(I-2)*P(I)*V(I-1)/T	34080
	B2T=B2T-(I-2)*Q(I)*V(I-1)/T	34090
	B1TT=B1TT+P(I)*(I-2)**2*V(I-1)/T/T	34100
4	B2TT=B2TT+Q(I)*(I-2)**2*V(I-1)/T/T	34110
	B1TT=B1TT-B1T/T	34120
	B2TT=B2TT-B2T/T	34130
	RETURN	34140

	END	34150
C		34160
C	THIS FUNCTION CALCULATES THE Z (=PBASE/(DRT)) NEEDED FOR FUNCTION	34170
C		34180
C	FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	34190
C		34200
	FUNCTION BASE(D,T)	34210
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	34220
	COMMON /ELLCON/ G1,G2,GF,B1,B2,B1T,B2T,B1TT,B2TT	34230
C		34240
C	G1,G2 AND GF ARE THE ALPHA, BETA AND GAMMA FOR DENSITY OF WATER	34250
C	CALCULATIONS. B1 AND B2 ARE THE 'EXCLUDED VOLUME' AND '2ND VIRIAL	34260
C	SUPPLIED BY THE SUBROUTINE BB(T), WHICH ALSO SUPPLIES THE 1ST AND	34270
C	2ND DERIVATIVES WITH RESPECT TO T (B1T,B2T,B1TT,B2TT).	34280
C		34290
	COMMON /ACONST/ GASCON,TZ,A,Z,DZ,Y	34300
	Y=.25D0*B1*D	34310
	X=1.D0-Y	34320
	Z0=(1.D0+G1*Y+G2*Y*Y)/X**3	34330
	Z=Z0+4.D0*Y*(B2/B1-GF)	34340
	DZ0=(G1+2.D0*G2*Y)/X**3 + 3.D0*(1.D0+G1*Y+G2*Y*Y)/X**4	34350
	DZ=DZ0+4.D0*(B2/B1-GF)	34360
	BASE=Z	34370
	RETURN	34380
	END	34390
C		34400
C	THIS ROUTINE CALCULATES, FOR A GIVEN T(K) AND D(G/CM3), THE RESIDU	34410
C	CONTRIBUTIONS TO: PRESSURE (Q), DP/DRHO (Q5)	34420
C	THIS SUBROUTINE IS USED IN DENSITY OF WATER CALCULATION.	34430
C		34440
C	FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	34450
C		34460
	SUBROUTINE QQ(T,D)	34470
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	34480
	COMMON /QQQQ/ Q,Q5	34490
	DIMENSION QR(11),QT(10),QZR(9),QZT(9)	34500
	EQUIVALENCE (QR(3),QZR(1)),(QT(2),QZT(1))	34510
	COMMON /NCONST/ G(40),II(40),JJ(40),N	34520
	COMMON /ACONST/ GASCON,TZ,AA,Z,DZ,Y	34530
	COMMON /ADDCON/ ATZ(4),ADZ(4),AAT(4),AAD(4)	34540
	RT=GASCON*T	34550
	QR(1)=0.D0	34560
	Q5=0.D0	34570
	Q=0.D0	34580
	E=DEXP(-AA*D)	34590
	Q10=D*D*E	34600
	Q20=1.D0-E	34610
	QR(2)=Q10	34620
	V=TZ/T	34630
	QT(1)=T/TZ	34640
	DO 4 I=2,10	34650
	QR(I+1)=QR(I)*Q20	34660
4	QT(I)=QT(I-1)*V	34670
	DO 10 I=1,N	34680
	K=II(I)+1	34690
	L=JJ(I)	34700

	ZZ=K	34710
	QP=G(I)*AA*QR(K+1)*QZT(L)	34720
	Q=Q+QP	34730
10	Q5 = Q5 + AA*(2.DO/D-AA*(1.DO-E*(K-1)/Q20))*QP	34740
	QP=0.DO	34750
C		34760
	DO 20 J=37,40	34770
	IF(DABS(G(J)).LT.1.0D-20) GO TO 20	34780
	K=II(J)	34790
	KM=JJ(J)	34800
	DDZ = ADZ(J-36)	34810
	DEL = D/DDZ - 1.DO	34820
	IF(DABS(DEL).LT.1.D-10) DEL=1.D-10	34830
	EX1 = -AAD(J-36)*DEL**K	34840
	IF(EX1.GT.-88.028D0) GO TO 5	34850
	DEX=0.DO	34860
	GO TO 6	34870
5	CONTINUE	34880
	DEX=DEXP(EX1)*DEL**KM	34890
6	CONTINUE	34900
	ATT = AAT(J-36)	34910
	TX = ATZ(J-36)	34920
	TAU = T/TX-1.DO	34930
	EX2 = -ATT*TAU*TAU	34940
	IF(EX2.GT.-88.028D0) GO TO 7	34950
	TEX=0.DO	34960
	GO TO 8	34970
7	CONTINUE	34980
	TEX = DEXP(EX2)	34990
8	CONTINUE	35000
	Q10 = DEX*TEX	35010
	QM = KM/DEL - K*AAD(J-36)*DEL**(K-1)	35020
	FCT=QM*D**2*Q10/DDZ	35030
	Q5T = FCT*(2.DO/D+QM/DDZ)-(D/DDZ)**2*Q10*(KM/DEL/DEL+	35040
1	K*(K-1)*AAD(J-36)*DEL**(K-2))	35050
	Q5 = Q5 + Q5T*G(J)	35060
	QP = QP + G(J)*FCT	35070
20	CONTINUE	35080
	Q=Q+QP	35090
	RETURN	35100
	END	35110
C		35120
C	ROUTINE TO FIND DENSITY CORRESPONDING TO INPUT PRESSURE P(MPA), AN	35130
C	TEMPERATURE T(K), USING INITIAL GUESS DENSITY D(G/CM3). THE OUTPUT	35140
C	DENSITY IS IN G/CM3.	35150
C		35160
C	FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	35170
C		35180
	SUBROUTINE DFIND(DOUT,P,D,T)	35190
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	35200
	COMMON /QQQQ/ Q0,Q5	35210
	COMMON /ACONST/ GASCON,TZ,AA,Z,DZ,Y	35220
	DD=D	35230
	RT=GASCON*T	35240
	IF(DD.LE.0.DO) DD=1.D-8	35250
	IF(DD.GT.1.9D0) DD=1.9D0	35260

L=0	35270
9 L=L+1	35280
IF(DD.LE.0.D0) DD=1.D-8	35290
IF(DD.GT.1.9D0) DD=1.9D0	35300
CALL QQ(T,DD)	35310
PP = RT*DD*BASE(DD,T)+Q0	35320
DPD=RT*(Z+Y*DZ)+Q5	35330
C	35340
C THE FOLLOWING 3 LINES CHECK FOR NEGATIVE DP/DRHO, AND IF SO ASSUME	35350
C GUESS TO BE IN 2-PHASE REGION, AND CORRECT GUESS ACCORDINGLY.	35360
C	35370
IF(DPD.GT.0.D0) GO TO 13	35380
IF(D.GE..2967D0) DD=DD*1.02D0	35390
IF(D.LT..2967D0) DD=DD*.98D0	35400
IF(L.LE.10) GO TO 9	35410
13 DPDX=DPD*1.1D0	35420
IF(DPDX.LT..1D0) DPDX=.1D0	35430
DP=DABS(1.D0-PP/P)	35440
IF(DP.LT.1.D-8) GO TO 20	35450
IF(D.GT..3D0 .AND. DP.LT.1.D-7) GO TO 20	35460
IF(D.GT..7D0 .AND. DP.LT.1.D-6) GO TO 20	35470
X=(P-PP)/DPDX	35480
IF(DABS(X).GT..1D0) X=X*.1D0/DABS(X)	35490
DD=DD+X	35500
IF(DD.LE.0.D0) DD=1.D-8	35510
IF(L.LE.30) GO TO 9	35520
STOP 1	35530
20 CONTINUE	35540
DOUT=DD	35550
RETURN	35560
END	35570
C	35580
C THIS FUNCTION CALCULATES AN APPROXIMATION TO THE VAPOR PRESSURE, P	35590
C AS A FUNCTION OF THE INPUT TEMPERATURE. THE VAPOR PRESSURE	35600
C CALCULATED AGREES WITH THE VAPOR PRESSURE PREDICTED BY THE SURFACE	35610
C TO WITHIN .02% TO WITHIN A DEGREE OR SO OF THE CRITICAL TEMPERATUR	35620
C AND CAN SERVE AS AN INITIAL GUESS FOR FURTHER REFINEMENT BY	35630
C IMPOSING THE CONDITION THAT GL=GV.	35640
C	35650
C FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	35660
C	35670
C FUNCTION PS(T)	35680
IMPLICIT DOUBLE PRECISION (A-H,O-Z)	35690
DIMENSION A(8)	35700
DATA A/-7.8889166D0,2.5514255D0,-6.716169D0	35710
1,33.239495D0,-105.38479D0,174.35319D0,-148.39348D0	35720
2,48.631602D0/	35730
IF(T.GT.314.D0) GO TO 2	35740
PL=6.3573118D0-8858.843D0/T+607.56335D0*T**(-.6)	35750
PS=.1*DEXP(PL)	35760
RETURN	35770
2 V=T/647.25D0	35780
W=DABS(1.D0-V)	35790
B=0.D0	35800
DO 4 I=1,8	35810
Z=I	35820

4	B=B+A(I)*W**((Z+1.D0)/2.D0)	35830
	Q=B/V	35840
	PS=22.093D0*DEXP(Q)	35850
	RETURN	35860
	END	35870
C		35880
	FUNCTION VLEST (T)	35890
C		35900
C	FROM L. HAAR, J. S. GALLAGHER, AND G. S. KELL, (1984)	35910
C		35920
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	35930
	DATA A,B,C,D,E,F,G/-1.59259D1,6.57886D-2,-1.12666D-4,7.33191D-8,	35940
1	1.60229D3,2.88572D0,650.0D0/	35950
	VLEST=A+B*T+C*T*T+D*T**3+E/T+F/(G-T)	35960
	RETURN	35970
	END	35980
C		35990
C	THIS FUNCTION CALCULATES THE RELATIVE DIELECTRIC CONSTANT AS A	36000
C	FUNCTION OF TEMPERATURE, ASSUMING ONE ATMOSPHERE PRESSURE	36010
C	ACCORDING TO D. J. BRADLEY AND K. S. PITZER, (1979)	36020
C		36030
	DOUBLE PRECISION FUNCTION DC (T)	36040
	IMPLICIT DOUBLE PRECISION (A-H,O-Z)	36050
	COMMON / MX10 / VP,DWO	36060
	DIMENSION U(9)	36070
	DATA U / 3.4279D2, -5.0866D-3, 9.4690E-7, -2.0525D0, 3.1159D3	36080
1	, -1.8289D2, -8.0325D3, 4.2142D6, 2.1417D0 /	36090
	D1000=U(1)*DEXP(U(2)*T+U(3)*T*T)	36100
	C=U(4)+U(5)/(U(6)+T)	36110
	B=U(7)+U(8)/T+U(9)*T	36120
	DC=D1000+C*DLOG((B+VP*1.01325)/(B+1000.0D0))	36130
	RETURN	36140
	END	36150

C		00010
C	THE FILE COMMON.BLOCKS TO BE INSERTED IN PHRQPITZ AT \$INSERT	00020
C	STATEMENTS.	00030
C		00040
	IMPLICIT DOUBLE PRECISION(A-H,O-Z),INTEGER(I-N)	00050
	CHARACTER *4 TITLE,HEAD	00060
	CHARACTER *12 SNAME,TNAME,MNAME,NAMELK,NMEANG,SUNAME	00070
	COMMON /CHARS/ TITLE(20),HEAD(2,20),SNAME(250),TNAME(30),	00080
1	MNAME(20),NAMELK(40),NMEANG(40),SUNAME(10)	00090
	DOUBLE PRECISION LM,M,LA,LG,LKSP,LKMIN,MU	00100
	COMMON /REAL8/ LM(250),M(250),LA(250),LG(250),LKSP(250),	00110
1	TOT(50),DELTA(50),DELTOT(50),AR(50,50),	00120
2	AS(50,50),CR(50),CS(50),LKMIN(20),	00130
3	THOR,ELECT,THSOLN,PH,PE,A,B,MU,TOTAL(2,30),	00140
4	DALKT,DALKS,DIFFZ(2),DZOFF	00150
	DOUBLE PRECISION LKTOSP,LKTOM,LKMINO	00160
	COMMON /REAL4/ CSP(250,6),ZSP(250),THSP(250),LKTOSP(250),	00170
1	DHSP(250),ASP(250,5),ADHSP(250,2),TH(2),	00180
2	TEMP(2),CMIN(20,10),THMIN(20),LKTOM(20),	00190
3	DHMIN(20),AMIN(20,5),CMCON(20,5),CMINO(20,10),	00200
4	LKMINO(20),VO,TITRPH(50),TITRML(50),TK,TC,XSTEP(50),	00210
5	TSTEP(50),CREAC(30),THREAC,	00220
6	THMEAN(30),DHA(250),ALKSP(250),SDENS(2),GFW(30),	00230
7	IMEANG(40),LMEANG(40,3),CMEANG(40,3),NMGS	00240
	INTEGER GFLAG,SFLAG	00250
	COMMON /INT2/ NSP(250),LSP(250,6),KFLAG(250),GFLAG(250),	00260
1	SFLAG(250),LASTT,LASTS,IIN(50),IOUT(50),IFE,ILE,	00270
2	IFTH,ILTH,IFT,ILT,IFM,ILM,NEQ,NEQ1,IESPEC,ISOLV(2),	00280
3	NMIN(20),LMIN(20,10),MFLAG(20),LMCON(20,5),NMCON(20),	00290
4	LMINO(20,10),NMINO(20),IOPT(10),NMINs,NSTEPS,NCOMPS,	00300
5	NELTS,NSPECS,ISTEP,LREAC(30),MAXT,MAXT1,MAXM,MAXEQ,	00310
6	MAXS,NRMINS,ITER,ISOL,IASPEC,IALK(2),IUNITS(2)	00320
	DOUBLE PRECISION LKLOOK,LKOLK	00330
	COMMON /LOOK/ LKOLK(40),LKLOOK(40),DHLOOK(40),	00340
1	ALOOK(40,5),CLOOK(40,10),LLOOK(40,10),NLOOK(40),	00350
2	LOOKFL(40),NLOOKS	00360
	COMMON /OFFSET/ SIMIN(20)	00370
	COMMON /CHKIT/ JCHECK,ICHECK	00380
	COMMON /H2O/ ICKH2O,ITER1	00390
	COMMON /NEUT/ DNEUT,NSUM(10),NSUMS,LSUM(10,50),LPOS,LNEG	00400
	COMMON / PMIN / KMIN	00410

**Attachment E. -- Listing of the source code to the interactive input
program PITZINPT**

PROGRAM PITZINPT	00010
C	00020
C	00030
C	00040
C	00050
C	00060
C	00070
C	00080
C	00090
C	00100
C	00110
C	00120
C	00130
C	00140
C	00150
C	00160
C	00170
C	00180
C	00190
C	00200
C	00210
C	00220
C	00230
C	00240
C	00250
C	00260
C	00270
C	00280
C	00290
COMMON / PHRQ1 / ADHSP(250,2),ALKSP(250),AMIN(2,100,5),ASP(250,5)	00300
1,CMIN(2,100,10),CREAC(30),CSP(250,6),DHA(250),DHMIN(2,100)	00310
2,DHSP(250),DTOT(2,30),GFLAG(250),HEAD(2),IALK(2),IOPT(10)	00320
3,ISDEL(250),IUNITS(2),KFLAG(250),LKTOM(2,100),LKTOSP(250)	00330
4,LMIN(2,100,10),LNEG,LPOS,LREAC(30),LSP(250,6),LSUM(10,50)	00340
COMMON / PHRQ2 / LT(2,30),MFLAG(2,100),MNAME(2,100),NCOMPS	00350
1,NELT(30),NI(250),NMINO(2,100),NSOLUTION(2),NSP(250),NSTEPS	00360
2,NSUM(10),PE(2),PH(2),SDENS(2),SIMIN(2,100),SNAME(250),SUNAME(10)	00370
3,TEMP(2),TGF(30),THMEAN(30),THMIN(2,100),THSP(250),TITLE	00380
4,TNAME(30),VO,XSTEP(50),XTEMP(50),ZSP(250)	00390
DIMENSION JSUB(11),IMINERAL(2),IMORE(11),NTOTS(2),IREQUIRED(11)	00400
2,ICHECK(2,4:30),MUSED(100),IORDER(11),IMAX(10)	00410
COMMON /IUNIT/ J	00420
COMMON /REFF/ JOPTION,IOPEN,JLINE	00430
COMMON /REFF1/ I	00440
COMMON /PN/ NFLAG	00450
COMMON /EE/ EXIT,LIST,LINE,JFLAG	00460
COMMON /PT/ MNAME0,NMINO0(100),THMINO,LKTOMO,DHMINO,MFLAGO(100)	00470
1,SIMINO,LMINO(100,10),CMINO,AMINO,IMINO,SUB,ICOL1,ICOL2(0:30)	00480
2,ICOL3,ICOL4,SPEC,OPV(30)	00490
PARAMETER (IOUT=10)	00500
REAL ICOL3(0:30)	00510
LOGICAL YN,EXIT,LIST,STOP,OK	00520
INTEGER GFLAG,OUT	00530
CHARACTER *80 TITLE,LINE,HEAD,JLINE(3)	00540

	CHARACTER *72 TFILE	00550
	CHARACTER *8 SUB(11),KEYWORD(20),TNAME,SNAME,MNAME,SUNAME,DELETE	00560
	1,ICOL1(0:30),ICOL4(0:30),CSP,MNAMEO(100),SPEC(30),CREAC,THMEAN	00570
	2,EQUIVWT,SUBO(11)	00580
	CHARACTER *12 OPTION,VO,PH,PE,TEMP,SDENS,TGFW,ZSP,THSP,DHA	00590
	1,ADHSP,ALKSP,LKTOSP,THMIN,LKTOM,DHMIN,SIMIN	00600
	2,XTEMP,XSTEP,DTOT,CMIN,DHSP,AMIN,ASP,BLANK,CMINO(100,10)	00610
	3,THMINO(100),LKTOMO(100),DHMINO(100),SIMINO(100),AMINO(100,5)	00620
	DATA SUB/'SOLUTION','ELEMENTS','SPECIES ','MINERALS','LOOK MIN',	00630
	1'TEMP ','STEPS ','REACTION','NEUTRAL ','SUMS ','END '/'	00640
	2SUBO/'solution','elements','species ','minerals','look min',	00650
	3'temp ','steps ','reaction','neutral ','sums ','end '/'	00660
	DATA OPTION/'OPTION CARD'/BLANK/' 0.0 '/'	00670
	DATA IMAX/ 1,2,6,3,0,2,2,1,1,1 /IORDER/2,3,1,4,5,6,7,8,9,10,11/	00680
	1IMORE/1,1,1,1,1,0,0,0,0,1,0/DELETE/'DELETE '/'	00690
C		00700
C	READ NECESSARY DATA AND INITIALIZE VARIABLES.	00710
C		00720
	CALL READFILE	00730
	WRITE (J,1)	00740
	1 FORMAT ('ENTER OUTPUT FILE NAME')	00750
	READ (J,20) TFILE	00760
	20 FORMAT (A72)	00770
	OPEN (UNIT=IOUT,FILE=TFILE,STATUS='UNKNOWN')	00780
C		00790
	ITIME=0	00800
	2 CALL OPEN (ITIME)	00810
	3 JOPEN=IOPEN	00820
C		00830
	DO 101 K=1,11	00840
	IREQUIRED(K)=0	00850
101	JSUB(K)=0	00860
	DO 103 K=1,100	00870
103	MUSED(K)=0	00880
	DO 104 K=1,250	00890
104	ISDEL(K)=0	00900
	DO 105 K=4,30	00910
	DO 105 K2=1,2	00920
105	ICHECK(K2,K)=0	00930
	DO 106 K=1,50	00940
	XSTEP(K)=BLANK	00950
106	XTEMP(K)=BLANK	00960
	DO 102 K=1,30	00970
	LREAC(K)=0	00980
	CREAC(K)=BLANK	00990
102	THMEAN(K)=BLANK	01000
	ISUB=0	01010
	IELEMENT=0	01020
	IMINERAL(1)=0	01030
	IDDELETE=0	01040
	JDELETE=0	01050
	IMINERAL(2)=0	01060
	ISPECIE=0	01070
	ISUM=0	01080
	NSTART=1	01090
	OUT=IOUT	01100

I=0	01110
ISOL=0	01120
IERROR=0	01130
JOPTION=0	01140
C	01150
IF (IOPEN.EQ.0)GO TO 10	01160
C	01170
EXPLANATION OF SOME FLAGS:	01180
C	01190
ICHECK -->MAKE SURE THAT FOR EVERY MASTER SPECIES	01200
ENTERED, THERE IS A CORRESPONDING ELEMENT.	01210
IERROR -->FLAG TO SEE IF IT IS OK TO END THE	01220
SIMULATION.	01230
JSUB -->KEEPS TRACK OF WHICH DATA BLOCK IS USED,	01240
AND WHICH ISN'T.	01250
MUSED -->KEEPS TRACK OF WHICH MINERAL DATA HAS BEEN	01260
CALLED.	01270
ISDEL -->SPECIES TO BE DELETED.	01280
IELEMENT -->NUMBER OF ELEMENTS.	01290
IMINERAL(1)-->NUMBER OF MINERALS.	01300
IMINERAL(2)-->NUMBER OF LOOK MINS.	01310
IMINO -->NUMBER OF PRE-CONSTRUCTED MINERAL DATA.	01320
ISPECIE -->NUMBER OF SPECIES.	01330
ISUM -->NUMBER OF SUMS.	01340
IDDELETE -->SET TO 1 IF ALL MINERALS ARE TO BE DELETED.	01350
JDELETE -->SET TO 1 IF ALL OLD SUMS ARE TO BE DELETED.	01360
I -->WHICH DATA BLOCK IS THE USER USING.	01370
NFLAG -->SET TO 1 IF THE VARIABLE IS SEEN BY THE USER	01380
THE SECOND TIME.	01390
JFLAG -->USER'S RESPONSE FLAG	01400
IREQUIRED -->WHICH DATA BLOCKS ARE REQUIRED.	01410
IMASTER -->IS THE SPECIES OR ELEMENT ENTERED 'MASTER'?	01420
IOPEN -->IF IT IS > 0, A REFERENCE FILE IS USED.	01430
JOPTION -->IF IOPEN > 0, WHAT DOES THE USER WANT TO DO	01440
WITH EACH CARD.	01450
ITIME -->NUMBER OF TIMES SUBROUTINE OPEN IS CALLED.	01460
ILINE -->NUMBER OF LINES TO BE READ FROM THE REFERENCE	01470
FILE.	01480
IMAX -->MAXIMUM VALUES PERMITTED FOR EACH OPTION.	01490
IORDER -->THE ORDER THAT THE DATA BLOCKS ARE TO BE	01500
PRINTED.	01510
IMORE -->IS IT OK TO CALL A CERTAIN DATA BLOCK MORE	01520
THAN ONCE.	01530
C	01540
C	01550
*** THE TITLE ***	01560
C	01570
WRITE (J,94)	01580
94 FORMAT ('TITLE CARD:')	01590
CALL REF (1,1,4)	01600
TITLE=JLINE(1)	01610
IF (IOPEN.EQ.0.OR.JOPTION.NE.1)GO TO 10	01620
GO TO 90	01630
10 NFLAG=0	01640
CALL QUESTA (*1000,0,'THE TITLE',9,TITLE,'A80 ',0,0)	01650
WRITE (J,9110)	01660

	WRITE (J,1731) TITLE	01670
	IF (.NOT.OK()) GO TO 10	01680
C		01690
C	*** OPTIONS ***	01700
C		01710
	90 IF (IOPEN.EQ.0)GO TO 91	01720
	WRITE (J,95) OPTION	01730
	95 FORMAT (/ ,A11,':')	01740
	CALL REF (1,1,1)	01750
	READ (JLINE(1),9991)(IOPT(N),N=1,10),NSTEPS,NCOMPS,VO	01760
	IOPT03=IOPT(3)	01770
	IOPT04=IOPT(4)	01780
	NSTEPS0=NSTEPS	01790
	NCOMPS0=NCOMPS	01800
	IF (JOPTION.EQ.3.OR.IOPEN.EQ.0)GO TO 91	01810
	IF (JOPTION.EQ.1)GO TO 92	01820
	GO TO 91	01830
	91 IOPT(5)=0	01840
	IOPT(6)=2	01850
	93 DO 200 N=NSTART,10	01860
	IF (N.EQ.5.OR.N.EQ.6) GO TO 200	01870
	CALL QUESTB (*290,N,'IOPT',4,IOPT(N),'I2 ',0,IMAX(N),0)	01880
	200 CONTINUE	01890
	IF ((IOPT(3).LE.4.AND.IOPT(3).GE.1).OR.IOPT(4).GE.2) GO TO 250	01900
	NSTEPS=0	01910
	GO TO 300	01920
C		01930
C	*** NSTEPS ***	01940
C		01950
	250 CALL QUESTB (*2200,0,'NSTEPS',6,NSTEPS,'I2 ',1,50,0)	01960
	300 IF (IOPT(3).EQ.3.OR.IOPT(3).EQ.4.OR.IOPT(3).EQ.6)GO TO 301	01970
	NCOMPS=0	01980
	GO TO 400	01990
C		02000
C	*** NCOMPS ***	02010
C		02020
	301 CALL QUESTB (*2300,0,'NCOMPS',6,NCOMPS,'I2 ',1,100,0)	02030
	400 IF (IOPT(3).EQ.2) GO TO 430	02040
	VO=' 0.0'	02050
	GO TO 500	02060
C		02070
C	*** VO ***	02080
C		02090
	430 CALL QUESTA (*2400,0,'VO',2,VO,'A10 ',1,0)	02100
	500 WRITE (J,9110)	02110
	WRITE (J,9991) (IOPT(I),I=1,10),NSTEPS,NCOMPS,VO	02120
	IF (OK()) GO TO 92	02130
	NSTART=1	02140
	GO TO 91	02150
C		02160
C	SET FLAGS FOR THE DATA BLOCKS THAT ARE REQUIRED.	02170
C		02180
	92 IF (IOPT(3).EQ.5.OR.IOPT(3).EQ.6)IREQUIRED(4)=1	02190
	IF (IOPT(2).EQ.2)IREQUIRED(9)=1	02200
	IF (IOPT(3).GE.1.AND.IOPT(3).LE.4)IREQUIRED(7)=1	02210
	IF (IOPT(3).EQ.3.OR.IOPT(3).EQ.4.OR.IOPT(3).EQ.6)IREQUIRED(8)=1	02220

IF (IOPT(4).NE.0)IREQUIRED(6)=1	02230
C	02240
C	02250
C	02260
C	02270
C	02280
C	02290
C	02300
C	02310
C	02320
C	02330
C	02340
C	02350
C	02360
C	02370
C	02380
531 ISUB=1	02390
621 WRITE (J,601)	02400
601 FORMAT (/,19('*'),/, 'KEYWORD DATA BLOCKS',/,19('*'),//)	02410
IF (IOPEN.EQ.0)GO TO 604	02420
612 READ (13,9001,END=608,ERR=650) KEYWORD(ISUB),NSOL	02430
DO 609 I=1,11	02440
609 IF (KEYWORD(ISUB).EQ.SUB(I))GO TO 615	02450
650 GO TO 612	02460
C	02470
C	02480
C	02490
C	02500
615 IF (I.EQ.6.AND.IOPT(4).EQ.0)GO TO 612	02510
IF (I.EQ.7.AND.(IOPT(3).EQ.0.OR.IOPT(3).EQ.5.OR.IOPT(3).EQ.6))	02520
1GO TO 612	02530
IF (I.EQ.8.AND.NCOMPS.EQ.0)GO TO 612	02540
IF (I.EQ.9.AND.IOPT(2).NE.2)GO TO 612	02550
WRITE (J,613)	02560
613 FORMAT (/, 'KEYWORD:')	02570
IF (I.EQ.1)GO TO 660	02580
WRITE (J,9050)KEYWORD(ISUB)	02590
GO TO 670	02600
660 WRITE (J,9001)KEYWORD(ISUB),NSOL	02610
670 CALL REF (0,1,2)	02620
IF (I.NE.11)GO TO 120	02630
IOPEN=0	02640
120 IF (JOPTION.EQ.1)GO TO 620	02650
JOPTION=0	02660
IF (I.EQ.11) GO TO 604	02670
GO TO 621	02680
608 IOPEN=0	02690
GO TO 531	02700
604 WRITE (J,603)	02710
603 FORMAT (/, 'ENTER KEYWORD.')	02720
READ (J,20) LINE	02730
READ (LINE,*,ERR=622) IHELP	02740
IF (IHELP.GT.11.OR.IHELP.LT.1)GO TO 4200	02750
C	02760
C	02770
C	02780
HE ASKED FOR HELP..	

GO TO (4148,4150,4146,4152,4154,4156,4158,4160,4162,4164,4166),	02790
11HELP	02800
622 READ (LINE,9050,ERR=4200) KEYWORD(ISUB)	02810
READ (LINE,810,ERR=820) NSOL	02820
810 FORMAT (9X,I1)	02830
GO TO 830	02840
820 NSOL=0	02850
830 DO 610 I=1,11	02860
IF (KEYWORD(ISUB).EQ.SUB0(I)) GO TO 620	02870
IF (KEYWORD(ISUB).EQ.SUB(I)) GO TO 620	02880
610 CONTINUE	02890
GO TO 4200	02900
C	02910
C AH, A LEGITIMATE KEYWORD...	02920
C	02930
620 IF(JSUB(I).EQ.1.AND.IMORE(I).EQ.0)GO TO 630	02940
JSUB(I)=1	02950
GO TO 640	02960
630 WRITE (J,635)	02970
635 FORMAT (/, 'YOU HAVE ALREADY USED IT ONCE' /)	02980
GO TO 621	02990
640 WRITE (J,639) KEYWORD(ISUB)	03000
639 FORMAT (//,8(' '),/,A8,/,8(' '),//)	03010
II=I-3	03020
GO TO (1700,1725,1750,3260,3250,3270,3280,3900,6000,7000,950), I	03030
STOP	03040
C	03050
C -----	03060
C SOLUTION DATA BLOCK	03070
C -----	03080
C	03090
1700 CONTINUE	03100
ISOL=ISOL+1	03110
IF (ISOL.LE.2) GO TO 1510	03120
WRITE (J,1300)	03130
1300 FORMAT ('ONLY 2 SOLUTIONS ARE ALLOWED. ')	03140
ISOL=2	03150
GO TO 531	03160
1510 IF (IOPEN.EQ.0.AND.NSOL.EQ.0)GO TO 1739	03170
IF (NSOL.EQ.1.OR.NSOL.EQ.2) GO TO 1500	03180
JOPTION=0	03190
GO TO 1492	03200
1500 IF (ISOL.NE.2.OR.NSOL.NE.NSOLUTION(1)) GO TO 1501	03210
WRITE (J,1502) NSOL	03220
1502 FORMAT ('SOLUTION #',I1,' HAS ALREADY BEEN ENTERED.')	03230
ISOL=1	03240
GO TO 531	03250
1501 NSOLUTION(ISOL)=NSOL	03260
IF (IOPEN.EQ.0) GO TO 1400	03270
WRITE (J,1736) SUB(I)	03280
1736 FORMAT (/,A8,' CARD:')	03290
CALL REF (2,1,0)	03300
READ (JLINE(1),1731) HEAD(ISOL)	03310
1731 FORMAT (A80)	03320
READ (JLINE(2),9003) NTOTS(ISOL),IALK(ISOL),IUNITS(ISOL),PH	03330
1 (ISOL),PE(ISOL),TEMP(ISOL),SDENS(ISOL)	03340

IF (NTOTS(ISOL).EQ.0)GO TO 1734	03350
CALL REF(1,1,0)	03360
READ (JLINE(1),9004) (LT(ISOL,M),DTOT(ISOL,M),M=1,5)	03370
IF (NTOTS(ISOL).LE.5)GO TO 1734	03380
LLINE=INT((NTOTS(ISOL)-1)/5)	03390
DO 1738 KLINE=1,LLINE	03400
CALL REF (1,1,0)	03410
1738 READ (JLINE(1),9004) (LT(ISOL,M),DTOT(ISOL,M),M=KLINE*5+1,	03420
1 (KLINE+1)*5)	03430
1734 CALL REF (0,1,1)	03440
IF (JOPTION.EQ.1.AND.IALK(ISOL).GT.0)GO TO 1752	03450
IF (JOPTION.EQ.1)GO TO 531	03460
C	03470
C *** SOLUTION NUMBER ***	03480
C	03490
1739 CALL QUESTB (*1492,0,'SOLUTION NUMBER',15,NSOLUTION(ISOL),'I1	03500
1,1,2,1)	03510
IF (.NOT.EXIT)GO TO 1730	03520
JSUB(1)=0	03530
1753 GO TO 531	03540
1730 IF (ISOL.NE.2.OR.NSOLUTION(2).NE.NSOLUTION(1)) GO TO 1732	03550
WRITE (J,1502) NSOLUTION(2)	03560
GO TO 1739	03570
1732 WRITE (J,1001) NSOLUTION(ISOL)	03580
1001 FORMAT (/,'SOLUTION',1X,I1)	03590
IF (.NOT.OK()) GO TO 1739	03600
1400 IREQUIRED(1)=0	03610
C	03620
C *** HEAD ***	03630
C	03640
IF (JOPTION.EQ.2) NFLAG=1	03650
1488 CALL QUESTA (*4300,0,'HEAD',4,HEAD(ISOL),'A80 ',0,0)	03660
IF (HEAD(ISOL).EQ.'*') HEAD(ISOL)='Pure water'	03670
WRITE (J,9110)	03680
WRITE (J,1731)HEAD(ISOL)	03690
IF (.NOT.OK()) GO TO 1488	03700
IF (HEAD(ISOL).NE.'Pure water') GO TO 1713	03710
NTOTS(ISOL)=0	03720
IALK(ISOL)=0	03730
IUNITS(ISOL)=0	03740
PH(ISOL)=' 7.0'	03750
PE(ISOL)=' 4.0'	03760
TEMP(ISOL)=' 25.0'	03770
GO TO 1719	03780
C	03790
C *** NTOTS ***	03800
C	03810
1713 CALL QUESTB (*4320,0,'NTOTS',5,NTOTS(ISOL),'I2 ',0,100,0)	03820
1714 IF (NTOTS(ISOL).NE.0)GO TO 1727	03830
IALK(ISOL)=0	03840
IUNITS(ISOL)=0	03850
GO TO 1718	03860
C	03870
C *** IALK ***	03880
C	03890
1727 CALL QUESTB (*4340,0,'IALK',4,IALK(ISOL),'I3 ',0,0,0)	03900

JUMP=0	03910
IF (IALK(ISOL)) 4340,1715,1708	03920
1708 IF (IALK(ISOL).GT.30.OR.IALK(ISOL).LT.4) GO TO 4340	03930
C	03940
C *** IUNITS ***	03950
C	03960
1715 CALL QUESTB (*4360,0,'IUNITS',6,IUNITS(ISOL),'I2' ',0,4,0)	03970
C	03980
C *** PH ***	03990
C	04000
1718 CALL QUESTA (*4380,0,'PH',2,PH(ISOL),'A10 ' ',1,0)	04010
C	04020
C *** PE ***	04030
C	04040
1723 PE(ISOL)=' 4.0	04050
C	04060
C *** TEMP ***	04070
C	04080
1724 CALL QUESTA (*4420,0,'TEMP',4,TEMP(ISOL),'A10 ' ',1,0)	04090
IF (IUNITS(ISOL).EQ.1.OR.IUNITS(ISOL).EQ.2) GO TO 1482	04100
1719 SDENS(ISOL)=' 1.0'	04110
GO TO 1490	04120
C	04130
C *** SDENS ***	04140
C	04150
1482 CALL QUESTA (*4440,0,'SDENS',5,SDENS(ISOL),'A10 ' ',1,0)	04160
1490 WRITE (J,9110)	04170
WRITE (J,9003) NTOTS(ISOL),IALK(ISOL),IUNITS(ISOL),PH(ISOL)	04180
1,PE(ISOL),TEMP(ISOL),SDENS(ISOL)	04190
IF (.NOT.OK()) GO TO 1713	04200
IF (IALK(ISOL).EQ.0)GO TO 1726	04210
1752 WRITE (J,1006)	04220
1006 FORMAT (/, 'HAS THE APPROPRIATE CARBON ELEMENT-CARD BEEN',	04230
1' CONSTRUCTED?')	04240
JUMP=1	04250
IF (.NOT.YN()) GO TO 1728	04260
GO TO (1726,1753,1726,1726),JOPTION+1	04270
C	04280
C ADD AN EXTRA ELEMENT CARD FOR HIM...	04290
C	04300
1728 WRITE (J,1709)	04310
1709 FORMAT (/, 'WHAT IS THE GRAM EQUIVALENT WEIGHT (GRAMS/EQUIVALENT'	04320
1,') OF THE',/, 'CHEMICAL SPECIES IN WHICH THE ALKALINITY IS',	04330
2' REPORTED?')	04340
WRITE (J,1710)	04350
1710 FORMAT (//, 'FOR EXAMPLE:',/,19X, 'CACO3 50.0446 G/EQ',/,19X	04360
1, 'HCO3- 61.0171 G/EQ',/,19X, 'CO3-- 30.0046 G/EQ.',//)	04370
READ (J,1731) LINE	04380
READ (LINE,*,ERR=1728) EQ	04390
EQUIVWT=LINE(1:8)	04400
JSUB(2)=2	04410
WRITE (J,1712)	04420
1712 FORMAT (//, 'ELEMENTS DATA BLOCK WILL BE CONSTRUCTED AUTOMATICALLY'	04430
1,//)	04440
IF (1ELEMENT.LT.27)GO TO 1004	04450
WRITE (J,1002)	04460

1002	FORMAT (1X,52('*')),/,1X,'WARNING: NUMBER OF ELEMENTS EXCEEDS ',	04470
	1'MAXIMUM POSSIBLE.',/,1X,'CARBON CARD IS DELETED.',/,1X,52('*'))/)	04480
	GO TO (1726,1753),JOPTION+1	04490
1004	IELEMENT=IELEMENT+1	04500
	TNAME(IELEMENT)='C'	04510
	NELT(IELEMENT)=IALK(ISOL)	04520
	TGFW(IELEMENT)=EQUIVWT	04530
	GO TO 1600	04540
1601	IF (JOPTION.EQ.1)GO TO 1753	04550
1726	IF (IUNITS(ISOL).GE.2.AND.IUNITS(ISOL).LE.3.AND.IELEMENT.EQ.0)	04560
	1 IREQUIRED(2)=1	04570
	IF (NTOTS(ISOL).EQ.0)GO TO 531	04580
1320	M=1	04590
1310	N2=1	04600
C		04610
C	*** LT ***	04620
C		04630
1326	CALL QUESTB (*4100,M,'LT',2,LT(ISOL,M),'I4 ',1,10000,2)	04640
	IF (.NOT.LIST)GO TO 1057	04650
1361	CALL LISTM	04660
	GO TO (1326,1356),N2	04670
1057	N2=2	04680
C		04690
C	*** DTOT ***	04700
C		04710
1356	CALL QUESTA (*4100,M,'DTOT',4,DTOT(ISOL,M),'A11 ',1,2)	04720
	IF (LIST)GO TO 1361	04730
	M=M+1	04740
	IF (M.LE.NTOTS(ISOL)) GO TO 1310	04750
	WRITE (J,9110)	04760
	WRITE (J,9004) (LT(ISOL,M),DTOT(ISOL,M),M=1,NTOTS(ISOL))	04770
	IF (.NOT.OK())GO TO 1320	04780
	GO TO 531	04790
1725	CONTINUE	04800
C		04810
C	-----	04820
C	ELEMENTS DATA BLOCK	04830
C	-----	04840
C		04850
	IELEMENT=IELEMENT+1	04860
	IF (IOPEN.EQ.0)GO TO 1729	04870
1536	CALL REF (1,0,0)	04880
	IF (JOPTION.EQ.0)GO TO 1530	04890
	WRITE (J,1736) SUB(I)	04900
	WRITE (J,1731)JLINE(1)	04910
	CALL REF (0,0,3)	04920
	GO TO (1530,1531,1536), JOPTION+1	04930
1531	READ (JLINE(1),9102) TNAME(IELEMENT),NELT(IELEMENT),TGFW(IELEMENT)	04940
	IREQUIRED(2)=0	04950
	IF (JOPTION.EQ.1) GO TO 1600	04960
C		04970
C	*** TNAME ***	04980
C		04990
1729	CALL QUESTA (*4102,0,'TNAME',5,TNAME(IELEMENT),'A8 ',0,3)	05000
	IF (.NOT.LIST)GO TO 1549	05010
	CALL LISTM	05020

GO TO 1729	05030
1549 IF (.NOT.EXIT)GO TO 1550	05040
1530 IELEMENT=IELEMENT-1	05050
IF (IELEMENT.EQ.0)JSUB(2)=0	05060
GO TO 531	05070
1550 IREQUIRED(2)=0	05080
C	05090
C *** NELT ***	05100
C	05110
CALL QUESTB (*4104,0,'NELT',4,NELT(IELEMENT),'I2 ',4,30,2)	05120
IF (.NOT.LIST) GO TO 1558	05130
CALL LISTM	05140
GO TO 1550	05150
1558 DO 1547 M=1,IELEMENT-1	05160
1547 IF (NELT(IELEMENT).EQ.NELT(M))GO TO 1548	05170
GO TO 1570	05180
1548 WRITE (J,1551) NELT(M)	05190
1551 FORMAT (/, 'WARNING: ELEMENT #',I2,' IS ALREADY ENTERED.'/)	05200
C	05210
C *** TGFW ***	05220
C	05230
1570 CALL QUESTA (*4106,0,'TGFW',4,TGFW(IELEMENT),'A10 ',1,2)	05240
IF (.NOT.LIST) GO TO 1580	05250
CALL LISTM	05260
GO TO 1570	05270
1580 WRITE (J,9110)	05280
WRITE (J,9102) TNAME(IELEMENT),NELT(IELEMENT),TGFW(IELEMENT)	05290
IF (.NOT.OK()) GO TO 1729	05300
C	05310
C KEEP TRACK OF ELEMENTS AND SPECIES	05320
C	05330
1600 IF(ICHECK(2,NELT(IELEMENT)).EQ.1)GO TO 1557	05340
ICHECK(1,NELT(IELEMENT))=1	05350
GO TO 1575	05360
1557 ICHECK(2,NELT(IELEMENT))=0	05370
1575 IF (I.EQ.1)GO TO 1601	05380
IF (IELEMENT.EQ.27) GO TO 531	05390
IF (IOPEN.NE.0)GO TO 1725	05400
WRITE (J,1573)	05410
1573 FORMAT (/, 'MORE ELEMENTS?')	05420
IF (YN()) GO TO 1725	05430
GO TO 531	05440
1750 CONTINUE	05450
C	05460
C -----	05470
C SPECIES DATA BLOCK	05480
C -----	05490
C	05500
ISPECIE=ISPECIE+1	05510
IF (IOPEN.EQ.0)GO TO 1590	05520
1523 CALL REF (1,0,0)	05530
IF (JOPTION.EQ.0)GO TO 1529	05540
WRITE (J,1736)SUB(1)	05550
WRITE (J,1731)JLINE(1)	05560
READ (JLINE(1),9202) NI(ISPECIE)	05570
CALL REF (3,1,3)	05580

GO TO (1521,1521,1523),JOPTION	05590
CALL REF (0,0,2)	05600
IF (JOPTION.EQ.2) GO TO 1523	05610
ISDEL(ISPECIE)=1	05620
GO TO 1750	05630
1521 READ (JLINE(1),9203) SNAME(ISPECIE),NSP(ISPECIE),KFLAG	05640
1(ISPECIE),GFLAG(ISPECIE),ZSP(ISPECIE),THSP(ISPECIE),DHA(ISPECIE)	05650
2,ADHSP(ISPECIE,1),ADHSP(ISPECIE,2),ALKSP(ISPECIE)	05660
READ (JLINE(2),9204) LKTOSP(ISPECIE),DHSP(ISPECIE),	05670
1(ASP(ISPECIE,MMM),MMM=1,5)	05680
READ (JLINE(3),9205) (LSP(ISPECIE,MMM),CSP(ISPECIE,MMM),	05690
1MMM=1,NSP(ISPECIE))	05700
IF (JOPTION.EQ.1) GO TO 1670	05710
C	05720
C *** I ***	05730
C	05740
1590 CALL QUESTB (*4110,0,'I',1,NI(ISPECIE),'I3 ',4,250,1)	05750
IF (.NOT.EXIT) GO TO 1654	05760
1529 ISPECIE=ISPECIE-1	05770
IF (ISPECIE.EQ.0)JSUB(3)=0	05780
IF (IOPEN.NE.0)GO TO 531	05790
GO TO 1933	05800
1654 WRITE (J,9110)	05810
WRITE (J,9202) NI (ISPECIE)	05820
IF (.NOT.OK()) GO TO 1590	05830
IF (ISDEL(ISPECIE).NE.1) GO TO 1660	05840
IF (JOPTION.EQ.2)GO TO 1750	05850
GO TO 1934	05860
C	05870
C *** SNAME ***	05880
C	05890
1660 CALL QUESTA (*4112,0,'SNAME',5,SNAME(ISPECIE),'A8 ',0,0)	05900
C	05910
C IS IT A MASTER SPECIE....	05920
C	05930
1670 IMASTER=0	05940
IF (NI(ISPECIE).LE.30.AND.NI(ISPECIE).GT.3.AND.ISDEL(ISPECIE)	05950
1.NE.1)IMASTER=1	05960
IF (JOPTION.EQ.1.AND.IMASTER.EQ.0)GO TO 1750	05970
IF (IMASTER.EQ.0.OR.(NFLAG.EQ.1.AND.JOPTION.NE.2))GO TO 1676	05980
C	05990
C KEEP TRACK OF ELEMENTS AND SPECIES...	06000
C	06010
IF (ICHECK(1,NI(ISPECIE)).EQ.1)GO TO 1679	06020
ICHECK(2,NI(ISPECIE))=1	06030
GO TO 1595	06040
1679 ICHECK(1,NI(ISPECIE))=0	06050
1595 IF (JOPTION.EQ.1)GO TO 1750	06060
C	06070
C IF IT IS A MASTER SPECIES, A LOT OF VARIABLES CAN BE	06080
C DETERMINED WITHOUT ASKING THE USER...	06090
C	06100
NSP(ISPECIE)=1	06110
KFLAG(ISPECIE)=0	06120
CSP(ISPECIE,1)=' 1.0'	06130
LSP(ISPECIE,1)=NI(ISPECIE)	06140

LKTOSP(ISPECIE)=' 0.0'	06150
DHSP(ISPECIE)=' 0.0'	06160
DO 1678 NII=1,5	06170
1678 ASP(ISPECIE,NII)=' 0.0'	06180
GO TO 1690	06190
C	06200
C *** NSP ***	06210
C	06220
1676 CALL QUESTB (*4114,0,'NSP',3,NSP(ISPECIE),'I3 ',0,6,0)	06230
C	06240
C *** KFLAG ***	06250
C	06260
1680 CALL QUESTB (*4116,0,'KFLAG',5,KFLAG(ISPECIE),'I2 ',0,1,0)	06270
C	06280
C *** GFLAG ***	06290
C	06300
1690 CALL QUESTB (*4118,0,'GFLAG',5,GFLAG(ISPECIE),'I2 ',0,1,0)	06310
C	06320
C *** ZSP ***	06330
C	06340
1780 CALL QUESTA (*4120,0,'ZSP',3,ZSP(ISPECIE),'A10 ',1,0)	06350
C	06360
C *** THSP ***	06370
C	06380
1790 CALL QUESTA (*4122,0,'THSP',4,THSP(ISPECIE),'A10 ',1,0)	06390
C	06400
C *** DHA ***	06410
C	06420
1880 CALL QUESTA (*4124,0,'DHA',3,DHA(ISPECIE),'A10 ',1,0)	06430
IF (GFLAG(ISPECIE).EQ.1)GO TO 1891	06440
ADHSP(ISPECIE,1)=' 0.0'	06450
ADHSP(ISPECIE,2)=' 0.0'	06460
GO TO 1740	06470
C	06480
C *** ADHSP ***	06490
C	06500
1891 M=1	06510
1892 CALL QUESTA (*1897,M,'ADHSP',5,ADHSP(ISPECIE,M),'A10 ',1,0)	06520
M=M+1	06530
IF (M.EQ.3)GO TO 1740	06540
GO TO 1892	06550
1897 GO TO (4126,4128), M	06560
C	06570
C *** ALKSP ***	06580
C	06590
1740 CALL QUESTA (*4130,0,'ALKSP',5,ALKSP(ISPECIE),'A10 ',1,0)	06600
WRITE (J,9110)	06610
WRITE (J,9203)SNAME(ISPECIE),NSP(ISPECIE),KFLAG(ISPECIE),	06620
1GFLAG(ISPECIE),ZSP(ISPECIE),THSP(ISPECIE),DHA(ISPECIE),	06630
2ADHSP(ISPECIE,1),ADHSP(ISPECIE,2),ALKSP(ISPECIE)	06640
IF (.NOT.OK()) GO TO 1660	06650
1960 IF(IMASTER.EQ.0)GO TO 1967	06660
GO TO 1930	06670
C	06680
C *** LKTOSP ***	06690
C	06700

1967	CALL QUESTA (*4132,0,'LKTOSP',6,LKTOSP(ISPECIE),'A10 ',1,0)	06710
C		06720
C	*** DHSP ***	06730
C		06740
1970	CALL QUESTA (*4136,0,'DHSP',4,DHSP(ISPECIE),'A10 ',1,0)	06750
	IF (KFLAG(ISPECIE).EQ.1)GO TO 1986	06760
C		06770
C	*** ASP ***	06780
C		06790
	DO 1988 K=1,5	06800
1988	ASP(ISPECIE,K)=' 0.0'	06810
	GO TO 1990	06820
1986	K=1	06830
1987	CALL QUESTA (*4138,K,'ASP',3,ASP(ISPECIE,K),'A12 ',1,0)	06840
	K=K+1	06850
	IF (K.LT.6)GO TO 1987	06860
1990	WRITE (J,9110)	06870
	WRITE (J,9204) LKTOSP(ISPECIE),DHSP(ISPECIE),(ASP(ISPECIE,I8),	06880
	1I8=1,5)	06890
	IF (.NOT.OK()) GO TO 1960	06900
1910	K=1	06910
	IF (NSP(ISPECIE).EQ.0)GO TO 1930	06920
1912	K2=1	06930
C		06940
C	*** LSP, CSP ***	06950
C		06960
1915	IF (K2.EQ.1) CALL QUESTB (*4142,K,'LSP',3,LSP(ISPECIE,K),'I3 ',1	06970
	1,30,2)	06980
	IF (K2.EQ.2) CALL QUESTA (*4144,K,'CSP',3,CSP(ISPECIE,K),'A7 ',1	06990
	1,2)	07000
	IF (.NOT.LIST)GO TO 1922	07010
	CALL LISTM	07020
	GO TO 1915	07030
1922	IF (K2.EQ.2) GO TO 1923	07040
	K2=2	07050
	GO TO 1915	07060
1923	K=K+1	07070
	IF (K.GT.NSP(ISPECIE)) GO TO 1940	07080
	GO TO 1912	07090
1940	WRITE (J,9110)	07100
	WRITE (J,9205)(LSP(ISPECIE,MM),CSP(ISPECIE,MM),MM=1,NSP(ISPECIE))	07110
	IF (.NOT.OK())GO TO 1910	07120
1930	IF (IOPEN.NE.0)GO TO 1750	07130
	WRITE (J,1931)	07140
1931	FORMAT (/, 'MORE SPECIES?')	07150
	IF (YN())GO TO 1750	07160
1933	WRITE (J,1932)	07170
1932	FORMAT (/, 'ANY SPECIES TO DELETE?')	07180
1936	IF (.NOT.YN())GO TO 531	07190
	ISDEL(ISPECIE+1)=1	07200
	JSUB(3)=1	07210
	GO TO 1750	07220
1934	WRITE (J,1935)	07230
1935	FORMAT (/, 'MORE SPECIES TO DELETE?')	07240
	GO TO 1936	07250
C		07260

C	-----	07270
C	MINERALS DATA BLOCK	07280
C	-----	07290
C		07300
3260	CONTINUE	07310
	IF (IMINO.EQ.0.OR.IOPEN.NE.0)GO TO 3100	07320
C		07330
C	MINERAL DATA AVAILABLE...	07340
C		07350
	WRITE (J,3002)	07360
3002	FORMAT (/, 'PRE-CONSTRUCTED MINERAL DATA ARE AVAILABLE.'/,	07370
	1 'DO YOU WISH TO HAVE ANY OF THEM?')	07380
	IF (.NOT.YN())GO TO 3100	07390
C		07400
C	PRINT THE LIST OF MINERALS.....	07410
C		07420
3004	WRITE (J,3003)(K,MNAMEO(K),K=1,IMINO)	07430
3003	FORMAT (/20(1X,5(13,1X,A8,2X,:)/))	07440
3021	WRITE (J,3005)	07450
3005	FORMAT (/, 'ENTER THE INDEX NUMBER OF MINERAL. (TYPE <STOP> TO'	07460
	1, ' EXIT)')	07470
	READ (J,1731) LINE	07480
	IF (LINE(1:4).EQ. 'STOP'.OR.LINE(1:4).EQ. 'stop') GO TO 3010	07490
	READ (LINE,*,ERR=3004) INDEX	07500
	IF (INDEX.GT.IMINO.OR.INDEX.LT.1)GO TO 3004	07510
	IF (MUSED(INDEX).EQ.1)GO TO 3012	07520
	MUSED(INDEX)=1	07530
	GO TO 3013	07540
3012	WRITE (J,3014) INDEX	07550
3014	FORMAT (//, 'MINERAL # ',13, ' HAS ALREADY BEEN ENTERED.'//)	07560
	GO TO 3021	07570
3013	IMINERAL(1)=IMINERAL(1)+1	07580
	IM=IMINERAL(1)	07590
	IREQUIRED(4)=0	07600
	MNAME(1,IM)=MNAMEO(INDEX)	07610
	NMINO(1,IM)=NMINOO(INDEX)	07620
	THMIN(1,IM)=THMINO(INDEX)	07630
	LKTOM(1,IM)=LKTOMO(INDEX)	07640
	DHMIN(1,IM)=DHMINO(INDEX)	07650
	MFLAG(1,IM)=MFLAGO(INDEX)	07660
	SIMIN(1,IM)=SIMINO(INDEX)	07670
	DO 3006 K2=1,NMINO(1,IM)	07680
	LMIN(1,IM,K2)=LMINO(INDEX,K2)	07690
3006	CMIN(1,IM,K2)=CMINO(INDEX,K2)	07700
3015	LFLAG=1	07710
	GO TO 3032	07720
3020	DO 3008 K2=1,5	07730
3008	AMIN(1,IM,K2)=AMINO(INDEX,K2)	07740
	GO TO 3210	07750
3024	IF (MFLAG(11,IM).EQ.1) GO TO 3020	07760
	DO 3001 K2=1,5	07770
3001	AMIN(1,IM,K2)=BLANK	07780
3000	IF (OK()) GO TO 3016	07790
	WRITE (J,3017)	07800
3017	FORMAT ('DO YOU WISH TO CHANGE SIMIN?')	07810
	IF (.NOT.YN())GO TO 3019	07820

LFLAG=3	07830
GO TO 3167	07840
3009 WRITE (J,3022)	07850
3022 FORMAT ('DO YOU WISH TO CHANGE ANYTHING ELSE?')	07860
IF (.NOT.YN())GO TO 3015	07870
3019 LFLAG=2	07880
GO TO 3102	07890
3016 LFLAG=0	07900
GO TO 3021	07910
3010 WRITE (J,3011)	07920
3011 FORMAT (/, 'MORE MINERALS TO BE TYPED IN FROM THE TERMINAL?')	07930
IF (.NOT.YN()) GO TO 3061	07940
3100 IF (IOPEN.EQ.0)GO TO 3221	07950
IMINERAL(II)=IMINERAL(II)+1	07960
IM=IMINERAL(II)	07970
3062 CALL REF (1,0,0)	07980
IF (JOPTION.NE.0)GO TO 3056	07990
3059 IMINERAL(II)=IMINERAL(II)-1	08000
3061 IF (IMINERAL(II).EQ.0)JSUB(I)=0	08010
GO TO 531	08020
3056 WRITE (J,1736) SUB(I)	08030
WRITE (J,1731)JLINE(1)	08040
READ (JLINE(1),9302)MNAME(II,IM),NMINO(II,IM),THMIN(II,IM)	08050
1,LKTOM(II,IM),DHMIN(II,IM),MFLAG(II,IM),SIMIN(II,IM)	08060
IF (MNAME(II,IM).NE.DELETE)GO TO 3051	08070
CALL REF (0,0,2)	08080
GO TO (3063,3062),JOPTION	08090
3063 IDELETE=1	08100
GO TO 3062	08110
3051 NLINE=2	08120
IF (NMINO(II,IM).GT.5)NLINE=3	08130
IF (MFLAG(II,IM).NE.1)NLINE=NLINE-1	08140
CALL REF (NLINE,1,3)	08150
READ (JLINE(1),9303)(LMIN(II,IM,MMM),CMIN(II,IM,MMM),MMM=1,5)	08160
IF (NMINO(II,IM).LE.5)GO TO 3054	08170
READ (JLINE(2),9303)(LMIN(II,IM,MMM),CMIN(II,IM,MMM),MMM=6	08180
1,NMINO(II,IM))	08190
3054 IF (MFLAG(II,IM).NE.1)GO TO 3057	08200
READ (JLINE(NLINE),9304) (AMIN(II,IM,MMM),MMM=1,5)	08210
3057 GO TO (3102,3062), JOPTION-1	08220
IREQUIRED(I)=0	08230
GO TO 3100	08240
C	08250
C *** MNAME ***	08260
C	08270
3221 IMINERAL(II)=IMINERAL(II)+1	08280
IM=IMINERAL(II)	08290
LFLAG=0	08300
3102 CALL QUESTA (*8000,0,'MNAME',5,MNAME(II,IM),'A8 ',0,1)	08310
IF (EXIT) GO TO 3059	08320
3106 IREQUIRED(I)=0	08330
C	08340
C *** NMINO ***	08350
C	08360
3110 CALL QUESTB (*8002,0,'NMINO',5,NMINO(II,IM),'I2 ',1,10,0)	08370
C	08380

C	*** THMIN ***	08390
C		08400
	3120 CALL QUESTA (*8004,0,'THMIN',5,THMIN(II,IM),'A10 ',1,0)	08410
C		08420
C	*** LKTOM ***	08430
C		08440
	3130 CALL QUESTA (*8006,0,'LKTOM',5,LKTOM(II,IM),'A10 ',1,0)	08450
C		08460
C	*** DHMIN ***	08470
C		08480
	3140 CALL QUESTA (*8008,0,'DHMIN',5,DHMIN(II,IM),'A10 ',1,0)	08490
C		08500
C	*** MFLAG ***	08510
C		08520
	3150 CALL QUESTB (*8010,0,'MFLAG',5,MFLAG(II,IM),'I1 ',0,1,0)	08530
	IF (LFLAG.EQ.2) GO TO 3032	08540
C		08550
C	*** SIMIN ***	08560
C		08570
	3167 CALL QUESTA (*8012,0,'SIMIN',5,SIMIN(II,IM),'A10 ',1,0)	08580
	IF (LFLAG.EQ.3)GO TO 3009	08590
	3032 WRITE (J,9110)	08600
	WRITE (J,9302) MNAME(II,IM),NMINO(II,IM),THMIN(II,IM),LKTOM(II,IM)	08610
	1,DHMIN(II,IM),MFLAG(II,IM),SIMIN(II,IM)	08620
	IF (LFLAG.EQ.1)GO TO 3187	08630
	IF (LFLAG.EQ.2) LFLAG=4	08640
	IF (.NOT.OK()) GO TO 3102	08650
	IF (LFLAG.EQ.4) NFLAG=1	08660
C		08670
C	*** LMIN, CMIN ***	08680
C		08690
	3190 K=1	08700
	3192 K2=1	08710
	3195 IF (K2.EQ.1) CALL QUESTB (*8014,K,'LMIN',4,LMIN(II,IM,K),'I4 ',0	08720
	1,0,2)	08730
	IF (K2.EQ.2) CALL QUESTA (*8016,K,'CMIN',4,CMIN(II,IM,K),'A11 ',1	08740
	1,2)	08750
	IF (.NOT.LIST)GO TO 3191	08760
	CALL LISTM	08770
	GO TO 3195	08780
	3191 IF (K2.EQ.2) GO TO 3186	08790
	K2=2	08800
	GO TO 3195	08810
	3186 K=K+1	08820
	IF (K.LE.NMINO(II,IM)) GO TO 3192	08830
	WRITE (J,9110)	08840
	3187 WRITE (J,9303) (LMIN(II,IM,K),CMIN(II,IM,K),K=1,NMINO(II,IM))	08850
	IF (LFLAG.EQ.1) GO TO 3024	08860
	IF (.NOT.OK()) GO TO 3190	08870
	IF(MFLAG(II,IM).EQ.1)GO TO 3201	08880
	IF (LFLAG.EQ.4) GO TO 3016	08890
	IF (IOPEN.NE.0)GO TO 3100	08900
	GO TO 3010	08910
C		08920
C	*** AMIN ***	08930
C		08940

3201 IF (LFLAG.EQ.4) NFLAG=1	08950
3200 K=1	08960
3207 CALL QUESTA (*8018,K,'AMIN',4,AMIN(II,IM,K),'A12 ',1,0)	08970
K=K+1	08980
IF (K.LT.6)GO TO 3207	08990
WRITE (J,9110)	09000
3210 WRITE (J,9304)(AMIN(II,IM,K),K=1,5)	09010
IF (LFLAG.EQ.1) GO TO 3000	09020
IF (.NOT.OK()) GO TO 3200	09030
IF (LFLAG.EQ.4) GO TO 3016	09040
IF (IOPEN.NE.0)GO TO 3100	09050
GO TO 3010	09060
3250 CONTINUE	09070
C	09080
C	09090
C	09100
C	09110
C	09120
IF(IMINERAL(2).NE.0.OR.IOPEN.NE.0)GO TO 3100	09130
WRITE (J,3251)	09140
3251 FORMAT (/, 'DO YOU WANT TO DELETE ALL OLD MINERALS?')	09150
IF (YN())IDELETE=1	09160
IF (IDELETE.EQ.1) JSUB(I)=1	09170
C	09180
C	09190
C	09200
C	09210
GO TO 3100	09220
3270 CONTINUE	09230
C	09240
C	09250
C	09260
C	09270
C	09280
IF(IOPT(4).EQ.0)GO TO 3320	09290
NTEMP=IOPT(4)	09300
IF (IOPT(4).EQ.3)NTEMP=NSTEPS	09310
IF (IOPEN.EQ.0)GO TO 3308	09320
NSTEPS2=IOPT04	09330
IF (IOPT04.EQ.3) NSTEPS2=NSTEPS0	09340
READ (13,9501)(XTEMP(K3),K3=1,NSTEPS2)	09350
WRITE (J,1736) SUB(I)	09360
WRITE (J,9501)(XTEMP(K3),K3=1,NTEMP)	09370
CALL REF (0,0,1)	09380
IF (JOPTION.NE.1) GO TO 3308	09390
IREQUIRED(6)=0	09400
GO TO 531	09410
C	09420
C	09430
C	09440
3308 K=1	09450
3307 CALL QUESTA (*8020,K,'XTEMP',5,XTEMP(K),'A10 ',1,1)	09460
IF (.NOT.EXIT)GO TO 3312	09470
JSUB(6)=0	09480
GO TO 531	09490
3312 K=K+1	09500

IF (K.LE.NTEMP)GO TO 3307	09510
IREQUIRED(6)=0	09520
WRITE (J,9110)	09530
WRITE (J,9501)(XTEMP(K),K=1,NTEMP)	09540
IF (OK())GO TO 531	09550
GO TO 3308	09560
3320 WRITE (J,3321)	09570
3321 FORMAT (/,'ERROR: TEMP DATA BLOCK NOT REQUIRED WHEN IOPT(4) ='	09580
1,' 0',//)	09590
JSUB(6)=0	09600
GO TO 531	09610
3280 CONTINUE	09620
C	09630
C	09640
C	09650
C	09660
C	09670
IF(IOPT(3).EQ.0.OR.IOPT(3).EQ.5.OR.IOPT(3).EQ.6)GO TO 3630	09680
NSTEP=NSTEPS	09690
IF (IOPT(3).EQ.4)NSTEP=1	09700
IF (IOPEN.EQ.0)GO TO 3608	09710
NSTEPS1=NSTEPS0	09720
IF (IOPT03.EQ.4) NSTEPS1=1	09730
READ (13,9501)(XSTEP(K3),K3=1,NSTEPS1)	09740
WRITE (J,1736) SUB(1)	09750
WRITE (J,9501)(XSTEP(K3),K3=1,NSTEP)	09760
CALL REF (0,0,1)	09770
GO TO (3801,3801,3608), JOPTION	09780
3801 IREQUIRED(7)=0	09790
IF (JOPTION.NE.2) GO TO 531	09800
C	09810
C	09820
C	09830
3608 K=1	09840
3607 CALL QUESTA (*3620,K,'XSTEP',5,XSTEP(K),'A10 ',1,1)	09850
IF (.NOT.EXIT)GO TO 3606	09860
JSUB(7)=0	09870
GO TO 531	09880
3606 K=K+1	09890
IF (K.LE.NSTEP)GO TO 3607	09900
IREQUIRED(7)=0	09910
WRITE (J,9110)	09920
WRITE (J,9501)(XSTEP(K),K=1,NSTEP)	09930
IF (OK())GO TO 531	09940
GO TO 3608	09950
3620 GO TO (8022,8024,8026,8028),IOPT(3)	09960
3630 WRITE (J,3631) IOPT(3)	09970
3631 FORMAT (/,'ERROR: STEPS DATA BLOCK NOT ALLOWED WHEN IOPT(3) ='	09980
1,1X,I1,//)	09990
JSUB(7)=0	10000
GO TO 531	10010
3900 CONTINUE	10020
C	10030
C	10040
C	10050
C	10060

C		10070
	IF (NCOMPS.NE.0)GO TO 3902	10080
	WRITE (J,3901)	10090
3901	FORMAT (//,'ERROR: REACTION DATA BLOCK NOT ALLOWED WHEN',	10100
	1' NCOMPS = 0',//)	10110
	JSUB(8)=0	10120
	GO TO 531	10130
3902	IF (IOPEN.EQ.0)GO TO 5010	10140
	READ (13,9701)(LREAC(K3),CREAC(K3),THMEAN(K3),K3=1,NCOMPS0)	10150
	WRITE (J,1736) SUB(I)	10160
	WRITE (J,9701)(LREAC(K3),CREAC(K3),THMEAN(K3),K3=1,NCOMPS)	10170
	CALL REF (0,0,1)	10180
	IF (JOPTION.EQ.2) GO TO 5010	10190
	IREQUIRED(8)=0	10200
	IF (JOPTION.EQ.1) GO TO 531	10210
C		10220
C	*** THMEAN, LREAC, CREAC ***	10230
C		10240
5010	K=1	10250
5012	K2=1	10260
5015	IF (K2.EQ.1) CALL QUESTB (*8032,K,'LREAC',5,LREAC(K),'I4 ',0,0,3)	10270
	IF (K2.EQ.2) CALL QUESTA (*8034,K,'CREAC',5,CREAC(K),'A8 ',1,3)	10280
	IF (K2.EQ.3) CALL QUESTA (*8036,K,'THMEAN',6,THMEAN(K),'A8 ',1,3)	10290
	IF (.NOT.EXIT)GO TO 5023	10300
5300	JSUB(8)=0	10310
	GO TO 531	10320
5023	IF (.NOT.LIST) GO TO 5025	10330
	CALL LISTM	10340
	GO TO 5015	10350
5025	IF (K2.EQ.3) GO TO 5026	10360
	K2=K2+1	10370
	IREQUIRED(8)=0	10380
	GO TO 5015	10390
5026	K=K+1	10400
	IF (K.GT.NCOMPS) GO TO 5027	10410
	GO TO 5012	10420
5027	WRITE (J,9110)	10430
	WRITE (J,9701) (LREAC(K),CREAC(K),THMEAN(K),K=1,NCOMPS)	10440
	IF (OK()) GO TO 531	10450
	GO TO 5010	10460
6000	CONTINUE	10470
C		10480
C	-----	10490
C	NEUTRAL DATA BLOCK	10500
C	-----	10510
C		10520
	IF(IOPT(2).NE.2)GO TO 6040	10530
	IF (IOPEN.EQ.0)GO TO 6010	10540
	CALL REF (1,1,1)	10550
	GO TO (6100,6110,6110,6010),JOPTION+1	10560
6100	JSUB(9)=0	10570
	GO TO 531	10580
6110	READ(JLINE(1),9801)LPOS,LNEG	10590
	IREQUIRED(9)=0	10600
	IF (JOPTION.EQ.1)GO TO 531	10610
C		10620

C	*** LPOS ***	10630
C		10640
6010	CALL QUESTB (*8038,0,'LPOS',4,LPOS,'15 ',0,0,3)	10650
	IF (.NOT.EXIT)GO TO 6016	10660
	JSUB(9)=0	10670
	GO TO 531	10680
6016	IF (.NOT.LIST)GO TO 6017	10690
	CALL LISTM	10700
	GO TO 6010	10710
6017	IREQUIRED(9)=0	10720
C		10730
C	*** LNEG ***	10740
C		10750
6020	CALL QUESTB (*8040,0,'LNEG',4,LNEG,'15 ',0,0,2)	10760
	IF (.NOT.LIST)GO TO 6030	10770
	CALL LISTM	10780
	GO TO 6020	10790
6030	WRITE (J,9110)	10800
	WRITE (J,9801)LPOS,LNEG	10810
	IF (OK()) GO TO 531	10820
	GO TO 6000	10830
6040	WRITE (J,6041)	10840
6041	FORMAT (/,'ERROR: NEUTRAL DATA BLOCK IS REQUIRED ONLY IF IOPT'	10850
	1,'(2) = 2',//)	10860
	JSUB(9)=0	10870
	GO TO 531	10880
7000	CONTINUE	10890
C		10900
C	-----	10910
C	SUMS DATA BLOCK	10920
C	-----	10930
C		10940
	IF(ISUM.NE.0.OR.IOPEN.NE.0)GO TO 7001	10950
C		10960
C	DELETE ALL OLD SUMS....	10970
C		10980
	WRITE (J,7003)	10990
7003	FORMAT (/,'DO YOU WANT TO DELETE ALL OLD SUMS?')	11000
	IF (YN())JDELETE=1	11010
7001	ISUM=ISUM+1	11020
	IF (IOPEN.EQ.0)GO TO 7010	11030
7116	CALL REF (1,0,0)	11040
	IF (JOPTION.EQ.0)GO TO 7111	11050
	WRITE (J,1736) SUB(I)	11060
	WRITE (J,1731)JLINE(1)	11070
	READ (JLINE(1),9901) SUNAME(ISUM),NSUM(ISUM)	11080
	IF (SUNAME(ISUM).NE.DELETE)GO TO 7120	11090
	CALL REF (0,0,2)	11100
	GO TO (7111,7117,7116),JOPTION+1	11110
7117	JDELETE=1	11120
	GO TO 7116	11130
7120	READ (13,9902)(LSUM(ISUM,M3),M3=1,NSUM(ISUM))	11140
	WRITE (J,9902) (LSUM(ISUM,M3),M3=1,NSUM(ISUM))	11150
	CALL REF (0,0,3)	11160
	GO TO (7111,7001,7010,7116),JOPTION+1	11170
C		11180

C	*** SUNAME ***	11190
C		11200
	7010 CALL QUESTA (*8042,0,'SUNAME',6,SUNAME(ISUM),'A8 ',0,1)	11210
	IF (.NOT.EXIT)GO TO 7020	11220
	7111 ISUM=ISUM-1	11230
	IF (ISUM.EQ.0)JSUB(10)=0	11240
	GO TO 531	11250
C		11260
C	*** NSUM ***	11270
C		11280
	7020 CALL QUESTB (*8044,0,'NSUM',4,NSUM(ISUM),'I2 ',0,50,0)	11290
	WRITE (J,9110)	11300
	WRITE (J,9901) SUNAME(ISUM),NSUM(ISUM)	11310
	IF (.NOT.OK())GO TO 7010	11320
C		11330
C	*** LSUM ***	11340
C		11350
	7030 K=1	11360
	7037 CALL QUESTB (*8046,K,'LSUM',4,LSUM(ISUM,K),'I4 ',0,10000,0)	11370
	K=K+1	11380
	IF (K.LE.NSUM(ISUM))GO TO 7037	11390
	WRITE (J,9110)	11400
	WRITE (J,9902) (LSUM(ISUM,K),K=1,NSUM(ISUM))	11410
	IF (.NOT.OK())GO TO 7030	11420
	IF (JOPTION.EQ.2)GO TO 7001	11430
	WRITE (J,7051)	11440
	7051 FORMAT (/, 'MORE SUMS?')	11450
	IF (YN())GO TO 7000	11460
	GO TO 531	11470
C		11480
C	-----	11490
C	END DATA BLOCK	11500
C	-----	11510
C		11520
	950 CONTINUE	11530
	JOPTION=0	11540
	KKK=1	11550
	954 IF (IREQUIRED(KKK).EQ.1)GO TO 952	11560
	953 KKK=KKK+1	11570
	IF (KKK.EQ.11)GO TO 958	11580
	GO TO 954	11590
C		11600
C	OOPS!! HE LEFT OUT A DATA BLOCK.	11610
C		11620
	952 WRITE (J,955)SUB(KKK)	11630
	955 FORMAT (/, 'ERROR: ',A8,' DATA BLOCK IS REQUIRED'/)	11640
	IERROR=1	11650
	GO TO 953	11660
	958 IF (IERROR.EQ.0)GO TO 957	11670
	JSUB(11)=0	11680
	IERROR=0	11690
	GO TO 531	11700
C		11710
C	EXCLUDE ELEMENTS OR SPECIES THAT ARE ALREADY IN PHRQPITZ'S	11720
C	THERMODYNAMIC DATA.	11730
C		11740

957	DO 951 JL=4,30	11750
	IF (ICOL2(JL).EQ.0)GO TO 951	11760
	ICHECK(1,ICOL2(JL))=0	11770
	ICHECK(2,ICOL2(JL))=0	11780
951	CONTINUE	11790
	JJ=1	11800
943	JJJ=4	11810
942	IF (ICHECK(JJ,JJJ).EQ.1)GO TO (944,945),JJ	11820
949	JJJ=JJJ+1	11830
	IF (JJJ.EQ.31)GO TO 946	11840
	GO TO 942	11850
946	JJ=JJ+1	11860
	IF (JJ.EQ.3)GO TO 947	11870
	GO TO 943	11880
944	WRITE (J,948)JJJ	11890
948	FORMAT ('ERROR: SPECIES #',I2,' WAS NOT ENTERED.')	11900
938	IERROR=1	11910
	GO TO 949	11920
945	WRITE (J,937)JJJ	11930
937	FORMAT ('ERROR: ELEMENT #',I2,' WAS NOT ENTERED.')	11940
	GO TO 938	11950
947	IF (IERROR.EQ.0)GO TO 959	11960
C		11970
C	HE ISN'T ALLOWED TO END.	11980
C		11990
	IERROR=0	12000
	JSUB(11)=0	12010
	GO TO 531	12020
C		12030
C	EVERYTHING NEEDED IS PRESENT.	12040
C		12050
	959 GO TO 9010	12060
5001	WRITE (J,5000)	12070
5000	FORMAT ('MORE SIMULATIONS?')	12080
	IF (.NOT.YN())GO TO 50	12090
	IF (JOPEN.EQ.0)GO TO 5003	12100
5006	READ (13,1731,END=5003)LINE	12110
	WRITE (J,5005)	12120
5005	FORMAT(/,'DO YOU WISH TO USE THE NEXT SIMULATION IN YOUR',	12130
	1' REFERENCE?')	12140
	IF (.NOT.YN())GO TO 5008	12150
	BACKSPACE 13	12160
	IOPEN=1	12170
	GO TO 3	12180
5008	READ (13,5009,END=5003)LINE	12190
5009	FORMAT (A3)	12200
	IF (LINE(1:3).EQ.'END')GO TO 5006	12210
	GO TO 5008	12220
5003	WRITE (J,5002)	12230
5002	FORMAT (/,'DO YOU WISH TO DEFINE THE PREVIOUS OUTPUT AS '	12240
	1,/ ,1X,'YOUR NEW REFERENCE?')	12250
	IF (.NOT.YN())GO TO 2	12260
C		12270
C	WRITE THE OUTPUT ONTO THE REFERENCE FILE.■	12280
C		12290
	OUT=13	12300

CLOSE (UNIT=OUT)	12310
OPEN (UNIT=OUT,STATUS='SCRATCH')	12320
REWIND OUT	12330
GO TO 9010	12340
999 ENDFILE (UNIT=OUT)	12350
REWIND OUT	12360
IOPEN=1	12370
C	12380
C LET'S TRY AGAIN...	12390
C	12400
GO TO 3	12410
50 ENDFILE (UNIT=10)	12420
STOP	12430
C	12440
9010 CONTINUE	12450
C	12460
C WRITE TO FILE.....	12470
C	12480
ITER=1	12490
C	12500
C PRINT TITLE CARD.	12510
C	12520
WRITE (OUT,1731)TITLE	12530
C	12540
C PRINT OPTION CARD.	12550
C	12560
WRITE (OUT,9991) (IOPT(I),I=1,10),NSTEPS,NCOMPS,VO	12570
9991 FORMAT (10I1,2I2,6X,A10)	12580
DO 9999 L=1,11	12590
IF (JSUB(IORDER(L)).LT.1) GO TO 9999	12600
IF (IORDER(L).EQ.1)GO TO 9000	12610
WRITE (OUT,9050)SUB(IORDER(L))	12620
9050 FORMAT (A8)	12630
GO TO (9100,9200,9300,9300,9500,9600,9700,9800,9900,9999),IORDER(L	12640
1)-1	12650
C	12660
C PRINT SOLUTION DATA BLOCK.	12670
C	12680
9000 WRITE (OUT,9001) SUB(1),NSOLUTION(ITER)	12690
9001 FORMAT (A8,1X,I1)	12700
WRITE (OUT,1731) HEAD(ITER)	12710
WRITE (OUT,9003) NTOTS(ITER),IALK(ITER),IUNITS(ITER),PH(ITER)	12720
1, PE(ITER),TEMP(ITER),SDENS(ITER)	12730
9003 FORMAT(I2,I3,I2,3X,4A10)	12740
IF (NTOTS(ITER).EQ.0)GO TO 9005	12750
WRITE (OUT,9004)(LT(ITER,M),DTOT(ITER,M),M=1,NTOTS(ITER))	12760
9004 FORMAT (6(5(I4,A11),:,:/))	12770
9005 IF (ISOL.EQ.ITER) GO TO 9999	12780
ITER=2	12790
GO TO 9000	12800
C	12810
C PRINT ELEMENTS DATA BLOCK.	12820
C	12830
9100 DO 9103 MM=1,IELEMENT	12840
9103 WRITE (OUT,9102) TNAME(MM),NELT(MM),TGFWM(MM)	12850
9102 FORMAT(A8,2X,I2,3X,A10)	12860

WRITE (OUT,9110)	12870
9110 FORMAT (1X)	12880
GO TO 9999	12890
C	12900
C PRINT SPECIES DATA BLOCK.	12910
C	12920
9200 DO 9206 MM=1,ISPECIE	12930
WRITE (OUT,9202) NI(MM)	12940
9202 FORMAT (I3)	12950
IF(ISDEL(MM).EQ.0)GO TO 9207	12960
WRITE (OUT,9110)	12970
GO TO 9206	12980
9207 WRITE (OUT,9203) SNAME(MM),NSP(MM),KFLAG(MM),GFLAG(MM),ZSP(MM)	12990
1,THSP(MM),DHA(MM),ADHSP(MM,1),ADHSP(MM,2),ALKSP(MM)	13000
9203 FORMAT (A8,2X,I3,2I1,6A10)	13010
WRITE (OUT,9204) LKTOSP(MM),DHSP(MM),(ASP(MM,M),M=1,5)	13020
9204 FORMAT (2A10,5A12)	13030
WRITE (OUT,9205)(LSP(MM,M),CSP(MM,M),M=1,NSP(MM))	13040
9205 FORMAT (6(I3,A7))	13050
9206 CONTINUE	13060
WRITE (OUT,9110)	13070
GO TO 9999	13080
C	13090
C PRINT MINERALS OR LOOK MIN DATA BLOCK.	13100
C	13110
9300 NM=IORDER(L)-3	13120
IF (NM.NE.2.OR.IDELETE.NE.1)GO TO 9301	13130
WRITE (OUT,9050) DELETE	13140
9301 IM=IMINERAL(NM)	13150
DO 9305 MM=1,IM	13160
WRITE (OUT,9302) MNAME(NM,MM),NMINO(NM,MM),THMIN(NM,MM),LKTOM(NM	13170
1,MM),DHMIN(NM,MM),MFLAG(NM,MM),SIMIN(NM,MM)	13180
9302 FORMAT (A8,2X,I2,3X,3A10,5X,I1,9X,A10)	13190
WRITE (OUT,9303) (LMIN(NM,MM,M),CMIN(NM,MM,M),M=1,NMINO(NM,MM))	13200
9303 FORMAT (2(5(I4,A11,:),/))	13210
IF(MFLAG(NM,MM).NE.1)GO TO 9305	13220
WRITE (OUT,9304) (AMIN(NM,MM,M),M=1,5)	13230
9305 CONTINUE	13240
WRITE (OUT,9110)	13250
9304 FORMAT (5A12)	13260
GO TO 9999	13270
C	13280
C PRINT TEMP DATA BLOCK.	13290
C	13300
9500 WRITE (OUT,9501) (XTEMP(IF),IF=1,NTEMP)	13310
9501 FORMAT (7(8(A10,:),/))	13320
GO TO 9999	13330
C	13340
C PRINT STEP DATA BLOCK.	13350
C	13360
9600 WRITE (OUT,9501) (XSTEP(IF),IF=1,NSTEP)	13370
GO TO 9999	13380
C	13390
C PRINT REACTION DATA BLOCK.	13400
C	13410
9700 WRITE (OUT,9701)(LREAC(M),CREAC(M),THMEAN(M),M=1,NCOMPS)	13420

9701	FORMAT (8(4(I4,2A8,:),/))	13430
	GO TO 9999	13440
C		13450
C	PRINT NEUTRAL DATA BLOCK.	13460
C		13470
9800	WRITE (OUT,9801)LPOS,LNEG	13480
9801	FORMAT (2I5)	13490
	GO TO 9999	13500
C		13510
C	PRINT SUMS DATA BLOCK.	13520
C		13530
9900	IF (JDELETE.NE.1)GO TO 9904	13540
	WRITE (OUT,9302) DELETE	13550
9904	DO 9903 MM=1,ISUM	13560
	WRITE (OUT,9901)SUNAME(MM),NSUM(MM)	13570
9901	FORMAT (A8,2X,12)	13580
9903	WRITE (OUT,9902) (LSUM(MM,M),M=1,NSUM(MM))	13590
9902	FORMAT(20I4)	13600
	WRITE (OUT,9110)	13610
9999	CONTINUE	13620
C		13630
	IF (OUT.EQ.10) GO TO 5001	13640
	GO TO 999	13650
C		13660
C	HELP MESSAGES....	13670
C		13680
290	NSTART=N	13690
	GO TO (2000,2020,2040,2060,2080,2080,2120,2140,2160,2180),N	13700
1000	WRITE (J,1010)	13710
1010	FORMAT (/1X,'TITLE CARD',2X,'TITLE',/,1X,'FORMAT (A80)',	13720
	1/,1X,'EIGHTY CHARACTERS OF TITLES OR COMMENTS.'//)	13730
	GO TO 10	13740
2000	WRITE (J,2010) OPTION	13750
2010	FORMAT (/1X,A11,/,13X,'IOPT(1) = 0, NO PRINT OF',	13760
	1' THERMODYNAMIC DATA OR COEFFICIENTS',/,26X,'OF AQUEOUS',	13770
	2' SPECIES.',/,21X,'= 1, PRINT THE AQUEOUS MODEL DATA',	13780
	3' (WHICH ARE STORED',/26X,'ON DISK) ONCE DURING THE ENTIRE',	13790
	4' COMPUTER RUN.',/,1X,'FORMAT (I1)',//)	13800
	GO TO 93	13810
2020	WRITE (J,2030) OPTION	13820
2030	FORMAT(/1X,A11,/,13X,'IOPT(2) = 0, INITIAL SOLUTIONS',	13830
	1' ARE NOT TO BE CHARGE BALANCED.',/,26X,'REACTION SOLUTIONS',	13840
	2' MAINTAIN THE INITIAL CHARGE',/,26X,'IMBALANCE.',/,21X,'= 1,',	13850
	3' PH IS ADJUSTED IN INITIAL SOLUTION(S) TO',/,26X,'OBTAIN CHARGE',	13860
	4' BALANCE.',/,21X,'= 2, THE TOTAL CONCENTRATION OF ONE OF THE',	13870
	5' ELEMENTS',/,26X,'(EXCEPT H OR O) IS ADJUSTED TO OBTAIN',	13880
	6' ELECTRICAL',/,26X,'BALANCE. NEUTRAL INPUT IS REQUIRED.',/,1X,'F	13890
	7FORMAT (I1)',//)	13900
	GO TO 93	13910
2040	WRITE (J,2050) OPTION	13920
2050	FORMAT (/1X,A11,/,13X,'IOPT(3) = 0, NO REACTIONS ARE',	13930
	1' MODELED. ONLY THE INITIAL',/,26X,'SOLUTIONS ARE SOLVED.'//,	13940
	221X,'= 1, SOLUTION 1 IS MIXED (A HYPOTHETICAL CONSTANT',/,26X,	13950
	3'VOLUME PROCESS) WITH SOLUTION 2 IN SPECIFIED',/,26X,'REACTION',	13960
	4' STEPS. STEPS INPUT AND A VALUE FOR',/,26X,'NSTEPS ARE REQUIRED'	13970
	5,',. MINERALS INPUT MAY BE',/,26X,'INCLUDED.',/,21X,'= 2,',	13980

```

6' SOLUTION 1 IS TITRATED WITH SOLUTION 2 IN SPECI-',/,26X,      13990
7' FIED REACTION STEPS. STEPS INPUT, A VALUE FOR',/,26X,'NSTEPS,', 14000
8' AND A VALUE FOR V0 ARE REQUIRED.',/,26X,'MINERALS INPUT MAY', 14010
9' BE INCLUDED.',/,21X,'= 3, A', 14020
!' STOICHIOMETRIC REACTION IS ADDED IN SPECIFIED',/,26X,'REACTION', 14030
@' STEPS. REACTION INPUT, STEPS INPUT,',/,26X,'A VALUE FOR', 14040
#' NSTEPS, AND A VALUE FOR NCOMPS ARE',/,26X,'REQUIRED. MINERALS', 14050
$' INPUT MAY BE INCLUDED.'///' HIT <RETURN> FOR MORE INFORMATION.') 14060
  READ (J,20) LINE      14070
  WRITE (J,2055)        14080
2055 FORMAT(21X,'= 4, A NET STOICHIOMETRIC', 14090
  ^' REACTION IS ADDED IN NSTEPS',/,26X,'EQUAL INCREMENTS. REACTION' 14100
&,' INPUT, STEPS INPUT,',/,26X,'A VALUE FOR NSTEPS, AND A VALUE FOR 14110
* NCOMPS ARE',/,26X,'REQUIRED. MINERALS INPUT MAY BE INCLUDED.', 14120
(' ONLY',/,26X,'ONE VALUE FOR THE TOTAL REACTION IS READ IN STEPS. 14130
)',/,21X,'= 5, SOLUTION NUMBER 1 IS EQUILIBRATED WITH MINERAL', 14140
-/,26X,'PHASES ONLY. NO OTHER REACTION IS PERFORMED.',/,26X, 14150
='MINERALS INPUT IS REQUIRED.',/,21X,'= 6, A REACTION IS ADDED', 14160
+' TO SOLUTION 1 UNTIL EQUI-',/,26X,'LIBRIUM IS ATTAINED WITH THE', 14170
|' FIRST PHASE IN',/,26X,'MINERAL INPUT (EQUILIBRIUM WITH OTHER', 14180
:' MINERALS',/,26X,'PHASES IS MAINTAINED THROUGHOUT THE REACTION)', 14190
/.'',/,26X,'REACTION INPUT, A VALUE FOR NCOMPS, AND MINERALS',/, 14200
]26X,'INPUT ARE REQUIRED. NO STEPS INPUT IS REQUIRED.',/,26X, 14210
['NOTE: THERE SHOULD BE A COMMON ELEMENT IN THE',/,26X,'REACTION', 14220
~' AND THE FIRST PHASE IN MINERALS INPUT.',/,1X,'FORMAT (I1)',/) 14230
  GO TO 93      14240
2060 WRITE (J,2070) OPTION 14250
2070 FORMAT (/1X,A11,/,13X,'IOPT(4) = 0, THE TEMPERATURE OF', 14260
1' THE REACTION SOLUTION IS',/,26X,'A) THE SAME AS THE INITIAL', 14270
2' SOLUTION IF ADDING',/,26X,'A REACTION, OR B) CALCULATED LINEARLY 14280
3 FROM THE',/,26X,'END MEMBERS IF MIXING OR TITRATING. NO TEMP', 14290
4/,26X,'INPUT IS REQUIRED.',/,21X,'= 1, THE TEMPERATURE IS', 14300
5' CONSTANT DURING THE REACTION',/,26X,'STEPS AND DIFFERS FROM THAT 14310
6 OF THE INITIAL SOL-',/,26X,'UTION(S). ONE VALUE IS READ IN THE', 14320
7' TEMP INPUT.'///,21X,'= 2, THE TEMPERATURE IS VARIED FROM T(0) TO' 14330
8,' T(F) IN',/,26X,'NSTEPS EQUAL INCREMENTS, DURING THE REACTION',/ 14340
9,26X,'STEPS.',/,21X,'= 3, THE TEMPERATURE', 14350
$' OF EACH REACTION STEP IS SPEC-',/,26X,'IFIED IN TEMP INPUT, IN' 14360
%, ' ORDER. NSTEPS VALUES',/,26X,'ARE READ.',/,1X,'FORMAT (I1)',/) 14370
2080 GO TO 93      14380
2120 WRITE (J,2130) OPTION 14390
2130 FORMAT (/1X,A11,/,13X,'IOPT(7) = 0, DO NOT SAVE THE', 14400
1' AQUEOUS PHASE COMPOSITION AT',/,26X,'THE END OF A REACTION FOR', 14410
2' ADDITIONAL SIMU-',/,26X,'LATIONS.',/,21X,'= 1, SAVE THE FINAL', 14420
3' REACTION SOLUTION IN SOLUTION',/,26X,'NUMBER 1.',/,21X,'= 2,', 14430
4' SAVE THE FINAL REACTION SOLUTION IN SOLUTION',/,26X,'NUMBER 2.' 14440
5,/,1X,'FORMAT (I1)',/) 14450
  GO TO 93      14460
2140 WRITE (J,2150) OPTION 14470
2150 FORMAT (/1X,A11,/,13X,'IOPT(8) = 0, THE DEBUGGING PRINT' 14480
1,' ROUTINE IS NOT CALLED.',/,21X,'= 1, A LONG PRINTOUT IS OUTPUT' 14490
2,' AT EACH ITERATION IN',/,26X,'EACH PROBLEM. THIS PRINT IS TO BE 14500
3 USED ONLY',/,26X,'IF THERE ARE CONVERGENCE PROBLEM WITH THE', 14510
4/,26X,'PROGRAM. (SEE SUBROUTINE PBUG)',/,1X,'FORMAT (I1)',/) 14520
  GO TO 93      14530
2160 WRITE (J,2170) OPTION 14540

```

```

2170 FORMAT (/1X,A11,/,13X,'IOPT(9) = 0, NO PRINTOUT OF EACH',      14550
1' ARRAY INVERTED.',/,21X,'= 1, A LONG PRINTOUT OCCURS OF THE',      14560
2' ENTIRE ARRAY TO',/,26X,'BE INVERTED AT EACH ITERATION. THIS',    14570
3' PRINT IS',/,26X,'USED ONLY IF THERE ARE CONVERGENCE PROBLEMS.',    14580
4/,26X,'(SEE SUBROUTINE SLNQ)',/,1X,'FORMAT (I1)',/)                14590
GO TO 93                                                                14600
2180 WRITE (J,2190) OPTION                                             14610
2190 FORMAT (/1X,A11,/,13X,'IOPT(10) = 0, NO CONVENTION'             14620
1,' FOR ACTIVITY',/,27X,'COEFFICIENT IS USED.',/,22X,'= 1,'         14630
2,' MACINNES CONVENTION IS USED.',/,1X,'FORMAT (I1)',/)              14640
GO TO 93                                                                14650
2200 WRITE (J,2210) OPTION                                             14660
2210 FORMAT (/1X,A11,/,13X,'NSTEPS',7X,'THE NUMBER OF'               14670
1,' STEPS. A VALUE IS',/,26X,'REQUIRED IF IOPT(3) = 1, 2, 3,',      14680
2'OR 4, OR IF',/,26X,'IOPT(4) = 2 OR 3.',/,1X,'FORMAT (I2)',/)      14690
GO TO 250                                                                14700
2300 WRITE (J,2310) OPTION                                             14710
2310 FORMAT (/1X,A11,/,13X,'NCOMPS',7X,'THE NUMBER OF CONSTIT',     14720
1'UENTS IN A NET STOICHIOMETRIC REACTION. A CONSTITUENT M           14730
2AY BE ANY',/,26X,'ELEMENT WITH AN INDEX NUMBER BETWEEN 4 AND',     14740
3/,26X,'30 INCLUSIVE. NO AQUEOUS SPECIES WITH INDEX',/,26X,'NUMBER   14750
4S GREATER THAN 30 MAY BE INCLUDED AS',/,26X,'REACTION CONSTITUENTS 14760
5 EXCEPT H2 AND O2. ANY',/,26X,'CONSTITUENT WITH AN INDEX NUMBER G 14770
6REATER THAN',/,26X,'30 IS ASSUMED TO BE EITHER H2 OR O2 AND HAS',/ 14780
7,26X,'THE EFFECT OF RAISING OR LOWERING THE REDOX',/,26X,'STATE OF 14790
8 THE SOLUTION DEPENDING ON THE AS-',/,26X,'SIGNED VALENCE (THMEAN) 14800
9. A VALUE FOR NCOMPS',/,26X,'IS REQUIRED IF IOPT(3) = 3, 4, OR 6.   14810
!',/,1X,'FORMAT (I2)',/)                                              14820
GO TO 301                                                                14830
2400 WRITE (J,2410) OPTION                                             14840
2410 FORMAT (/1X,A11,/,13X,'VO',11X,'THE INITIAL VOLUME OF SO',      14850
1'LUTION NUMBER 1 WHEN',/,26X,'MODELING A TITRATION. THIS UNIT OF V 14860
20 MUST',/,26X,'BE THE SAME AS THAT OF XSTEP (SEE STEPS INPUT)',/,2 14870
36X,'IF IOPT(3) = 2. OTHERWISE, VO IS NOT REQUIRED',/,1X,'FORMAT (   14880
4F10.5)',/)                                                            14890
GO TO 430                                                                14900
4146 WRITE (J,4147) SUB(1)                                             14910
4147 FORMAT (/1X,A8,10X,'THIS INPUT IS USED TO DEFINE A STARTING'    14920
1,' SOLU-',/,19X,'TION.',/)                                           14930
GO TO 621                                                                14940
4148 WRITE (J,4149) SUB(2)                                             14950
4149 FORMAT (/1X,A8,10X,'THIS INPUT DEFINES THE NAMES AND INDICES',   14960
1' OF ALL',/,19X,'ELEMENTS IN THE AQUEOUS MODEL DATA BASE.',/)      14970
GO TO 621                                                                14980
4150 WRITE (J,4151) SUB(3)                                             14990
4151 FORMAT (/1X,A8,10X,'THIS INPUT DEFINES THE NAMES, INDEX NUMB'    15000
1'ERS AND',/,19X,'COMPOSITION OF ALL AQUEOUS SPECIES IN THE',       15010
2' AQUEOUS',/,19X,'MODEL DATA BASE.',/)                              15020
GO TO 621                                                                15030
4152 WRITE (J,4153) SUB(4)                                             15040
4153 FORMAT (/1X,A8,10X,'THIS INPUT DEFINES THE PHASES WHICH WILL'    15050
1,' BE MAIN-',/,19X,'TAINED AT EQUILIBRIUM WITH EACH OF THE'       15060
2,' REACTION',/,19X,'SOLUTIONS',/)                                    15070
GO TO 621                                                                15080
4154 WRITE (J,4155) SUB(5)                                             15090
4155 FORMAT (/1X,A8,10X,'THE PURPOSE OF THIS INPUT IS SIMPLY TO      15100

```

1, 'PROVIDE',/, 19X, 'INFORMATION ON THE SATURATION STATE OF THE'	15110
2, ' AQUEOUS',/, 19X, 'PHASE WITH RESPECT TO DESIRED MINERALS. '	15120
3, 'THE',/, 19X, 'MINERALS IN THIS BLOCK OF INPUT DO NOT AFFECT'	15130
4, ' THE',/, 19X, 'CALCULATIONS OF THE INITIAL SOLUTION OR ANY '	15140
5, 'OF THE',/, 19X, 'REACTION SOLUTIONS. THIS INPUT IS NEVER'	15150
6, ' MANDATORY.',//)	15160
GO TO 621	15170
4156 WRITE (J,4157) SUB(6)	15180
4157 FORMAT (/1X,A8,10X, 'THIS INPUT VARIES THE TEMPERATURE DURING'	15190
1, ' THE',/, 19X, 'REACTION STEPS.',//)	15200
GO TO 621	15210
4158 WRITE (J,4159) SUB(7)	15220
4159 FORMAT (/1X,A8,10X, 'THIS INPUT DEFINES THE STEPS OF THE '	15230
1, 'REACTION PROCESS.',//)	15240
GO TO 621	15250
4160 WRITE (J,4161) SUB(8)	15260
4161 FORMAT (/1X,A8,10X, 'THIS INPUT DESCRIBES THE STOICHIOMETRY'	15270
1, ' AND VALENCE',/, 19X, 'OF THE ELEMENTS TO BE ADDED AS A'	15280
2, ' REACTION.',//)	15290
GO TO 621	15300
4162 WRITE (J,4163) SUB(9)	15310
4163 FORMAT (/1X,A8,10X, 'THIS INPUT DEFINES THE ELEMENTS TO BE '	15320
1, 'USED TO',/, 19X, 'ADJUST THE INITIAL SOLUTION(S) TO ELECTRI'	15330
2, 'CAL',/, 19X, 'NEUTRALITY.',//)	15340
GO TO 621	15350
4164 WRITE (J,4165) SUB(10)	15360
4165 FORMAT (/1X,A8,10X, 'THIS INPUT SUMS MOLALITIES OF AQUEOUS'	15370
1, ' SPECIES WHICH',/, 19X, 'ARE THEN PRINTED IN THE OUTPUT OF THE'	15380
2, ' RUN. THESE',/, 19X, 'SUMS DO NOT AFFECT THE CALCULATIONS IN'	15390
3, ' ANY WAY AND',/, 19X, 'ARE NEVER MANDATORY.',//)	15400
GO TO 621	15410
4166 WRITE (J,4167) SUB(11)	15420
4167 FORMAT (/1X,A8,10X, 'THIS CARD TERMINATES INPUT OPERATIONS'	15430
1, ' FOR A',/, 19X, 'SINGLE SIMULATION. ANY COMPUTER RUN HAS '	15440
2, 'AT LEAST ONE',/, 19X, 'END CARD.',//)	15450
GO TO 621	15460
4200 WRITE (J,4210)	15470
4210 FORMAT (/1X, 'POSSIBLE KEYWORDS:',//1X, '(1)ELEMENTS, (2)SPECIES'	15480
1, ', (3)SOLUTION, (4)MINERALS, (5)LOOK MIN',/, 1X, '(6)TEMP, '	15490
2, '(7)STEPS, (8)REACTION, (9)NEUTRAL, (10)SUMS, (11)END',//, 1X	15500
3, 'FORMAT (A8)',//, ' FOR MORE INFORMATION ON ANY OF THE DATA'	15510
4, ' BLOCK, ENTER THE',/, 1X, 'NUMBER ASSOCIATED WITH IT.',/)	15520
GO TO 621	15530
1492 WRITE (J,1493)SUB(1)	15540
1493 FORMAT (/1X,A8,/, 13X, 'N',12X, 'A NUMBER OF EITHER 1 OR 2 INDICAT'	15550
1, 'ING THE',/, 26X, 'SOLUTION NUMBER OF THE FOLLOWING DATA.',/, 1X	15560
2, 'FORMAT (I1)'//)	15570
GO TO 1739	15580
4300 WRITE (J,4310) SUB(1)	15590
4310 FORMAT (/1X,A8,/, 13X, 'HEAD',/, 13X, 'FORMAT (A80)',/, 13X,	15600
1 'TITLE OR COMMENTS ABOUT THE SOLUTION.',//)	15610
GO TO 1488	15620
4320 WRITE (J,4330) SUB(1)	15630
4330 FORMAT (/1X,A8,/, 13X, 'NTOTS',8X, 'THE NUMBER OF TOTAL CONCENTRATION	15640
1S TO BE',/, 26X, 'READ FROM CARD INPUT. FOR EXAMPLE, IF THE',/, 26X,	15650
2 'STARTING SOLUTION IS MGCL2-NAHCO3 SOLUTION',/, 26X, 'NTOTS = 4 (FOR	15660

```

3 MG, CL, NA, AND C)',/,1X,'FORMAT (I2)',//) 15670
GO TO 1713 15680
4340 WRITE (J,4350) SUB(1) 15690
4350 FORMAT (/1X,A8,/,13X,'IALK',9X,'FLAG WHICH INDICATES WHETHER TOTAL 15700
1 CARBON OR',/,26X,'TOTAL ALKALINITY IS TO BE INPUT.',/,21X,'= 0, 15710
2INDICATES THE TOTAL CONCENTRATION OF CARBON',/,26X,'(NOT ALKALINIT 15720
3Y) IS INPUT IN THE UNITS',/,26X,'SPECIFIED BY IUNITS.',/,21X,'= N 15730
4, 4<=N<=30, WHERE N IS THE INDEX NUMBER FOR',/,26X,'THE ELEMENT CA 15740
5RBON, (IN OUR DATA BASE N=15)',/,26X,'INDICATES TOTAL ALKALINITY I 15750
6S BEING ENTERED.',/,26X,'ELEMENTS INPUT MAY BE REQUIRED. THE UNIT 15760
7S OF',/,26X,'ALKALINITY ARE SPECIFIED BY IUNITS',/,26X,'AND IF ', 15770
9'IUNITS > 0, THE GRAM FORMULA WEIGHT (GFW)',/,26X,'OF THE ELEMENT' 15780
!,' CARBON IS CRITICALLY IMPORTANT.',/,26X,'THE GFW IN THE CASE OF' 15790
@,' ALKALINITY MUST BE THE',/,26X,'GRAM EQUIVALENT WEIGHT (GRAMS/EQ 15800
#UIVALENT) OF',/,26X,'THE CHEMICAL SPECIES IN WHICH THE ALKALINTY', 15810
$/,26X,'IS REPORTED. THE FOLLOWING IS A LIST OF',/,/, ' HIT <RETURN 15820
/> FOR MORE EXPLANATION') 15830
READ (J,1731) LINE 15840
WRITE (J,4355) 15850
4355 FORMAT(//26X,'SPECIES ', 15860
%COMMONLY USED FOR REPORTING ALKA-',/,26X,'LINITY AND THEIR CORRES 15870
^PONDING EQUIVALENT',/,26X,'WEIGHTS:',/,31X,'CaCO3',4X,'50.0446 G/ 15880
&EQ',/,31X,'HCO3-',4X,'61.0171 G/EQ',/,31X,'CO3--',4X,'30.0046 G/ 15890
*EQ',/,26X,'IN OUR DATA BASE 44.010 IS THE GFW OF CARBON',/,26X,'W 15900
(HICH IS SUITABLE FOR ENTERING CARBON AS',/,26X,'TOTAL CO2. THIS G 15910
)FW MUST BE CHANGED VIA',/,26X,'ELEMENTS INPUT IF ALKALINITY IS TO' 15920
-,' BE ENTERED',/,26X,'AS MG/L OR PPM (IUNITS = 2 OR 3). IF IUNITS 15930
=',/,26X,'= 0 ALKALINITY MUST BE INPUT AS EQ/KG H2O AND',/,26X,'IN' 15940
+,' THIS CASE THE GFW NEED NOT BE CHANGED SINCE',/,26X,'NO CONVERSI 15950
:ON IS NECESSARY.',/,1X,'FORMAT (I3)',//) 15960
GO TO (1714,1752),JUMP+1 15970
4360 WRITE (J,4370) SUB(1) 15980
4370 FORMAT (/1X,A8,/,13X,'IUNITS',7X,'FLAG DESCRIBING UNITS OF INPUT C 15990
1ONCENTRATIONS.',/,26X,'THE PROGRAM MAKES ALL OF ITS CALCULATIONS' 16000
2,/,26X,'IN TERMS OF MOLALITY AND ANY OTHER ALLOWED',/,26X,'CONCEN' 16010
3,'TRATION UNITS (MMOLES/L, MG/L,',/,26X,'PPM OR MMOL/KG) MUST BE' 16020
4,' CONVERTED TO MOLALITY',/,26X,'BEFORE THE CALCULATION MAY BEGIN.' 16030
5,' TO MAKE',/,26X,'THE CONVERSIONS IT IS NECESSARY TO KNOW THE' 16040
6,/,26X,'GRAM FORMULA WEIGHT (GFW), IN G/MOLE, OF THE',/,26X 16050
7,'CHEMICAL FORMULA IN WHICH ELEMENTAL ANALYSES',/,26X,'ARE' 16060
8,' REPORTED. THE GFW IS AN INPUT PARAMETER',/,26X,'UNDER' 16070
9,' ELEMENTS INPUT AND MUST BE IN AGREEMENT',/,26X,'WITH THE' 16080
!,' ANALYTICAL UNITS FOR EACH SOLUTION',/,26X,'DATA SET. (IF THE' 16090
@,' UNITS ARE MOLALITY, NO',/,26X,'CONVERSION IS NECESSARY AND THE' 16100
#,' GFWs ARE NOT',/,26X,'USED.) NOTE: ALL ELEMENTS MUST HAVE',/ 16110
$,26X,'THE SAME UNITS. IT IS NOT POSSIBLE TO ENTER MG/L',/,26X 16120
/, 'OF ONE ELEMENT AND MOLALITY OF ANOTHER.',/,/, ' HIT <RETURN> FOR' 16130
*, ' MORE EXPLANATION') 16140
READ (J,1731) LINE 16150
WRITE (J,4375) 16160
4375 FORMAT (/21X,'= 0, CONCENTRATION ', 16170
%OF ELEMENTS ENTERED AS MOLALITY',/,26X,'OF EACH ELEMENT, OR FO', ' 16180
^R ALKALINITY, EQUIV-',/,26X,'ALENTS/KG H2O.',/,21X,'= 1, CONCENTR 16190
&ATION OF ELEMENTS ENTERED AS MMOLES/L',/,26X,'OF EACH ELEMENT, OR' 16200
*, ' FOR ALKALINITY, MEQ/L.',/,21X,'= 2, CONCENTRATION OF ELEMENTS 16210
(ENTERED AS MG/L',/,26X,'OF THE SPECIES WHICH HAS A GRAM FORMULA', 16220

```



```

),26X,'WEIGHT GIVEN IN ELEMENTS INPUT. (ELEMENTS',/,26X,'INPUT MA 16230
-Y BE REQUIRED.)',/,21X,'= 3, CONCENTRATION OF ELEMENTS ENTERED AS 16240
= PPM',/,26X,'OF THE SPECIES WHICH HAS A GRAM FORMULA',/,26X,'WEIGH 16250
+T GIVEN IN ELEMENTS INPUT. (ELEMENTS',/,26X,'INPUT MAY BE REQUIRE 16260
:D.)',/,21X,'= 4, CONCENTRATION OF ELEMENTS ENTERED AS MMOL/KG',/, 16270
!26X,'SOLUTION.',/,1X,'FORMAT (I1)') 16280
GO TO 1715 16290
4380 WRITE (J,4390) SUB(1) 16300
4390 FORMAT (/1X,A8,/,13X,'PH',11X,'THE PH OF THE SOLUTION (THE APPROXI 16310
1MATE PH',/,26X,'IF IOPT(2) = 1)',/,1X,'FORMAT (F10.3)',/) 16320
GO TO 1718 16330
4420 WRITE (J,4430) SUB(1) 16340
4430 FORMAT (/1X,A8,/,13X,'TEMP',9X,'THE TEMPERATURE OF THE SOLUTION IN 16350
1 DEGREES',/,26X,'CELCIUS.',/,1X,'FORMAT (F10.3)',/) 16360
GO TO 1724 16370
4440 WRITE (J,4450) SUB(1) 16380
4450 FORMAT (/1X,A8,/,13X,'SDENS',8X,'THE DENSITY OF THE SOLUTION.',/,1 16390
1X,'FORMAT (F10.3)',/) 16400
GO TO 1482 16410
4100 WRITE (J,4101) SUB(1) 16420
4101 FORMAT (/1X,A8,/,13X,'LT,DTOT',6X,'TOTAL CONCENTRATIONS OF ELEMENT 16430
1S.',/,26X,'LT',8X,'INDEX NUMBER OF THE ELEMENT.',/,26X,'DTOT', 16440
26X,'TOTAL CONCENTRATION OF THE ELEMENT IN',/,36X,'MOLALITY, MMOLES 16450
3/L, MG/L, OR',/,36X,'PPM ACCORDING TO IUNITS.',/,1X,'FORMAT (I4' 16460
4,',1X,F11.4)') 16470
WRITE (J,4000) 16480
GO TO (1326,1356),N2 16490
4102 WRITE (J,4103) SUB(2) 16500
4103 FORMAT (/1X,A8,/,13X,'TNAME',8X,'ALPHANUMERIC NAME OF ELEMENT.', 16510
1/,1X,'FORMAT (A8)') 16520
WRITE (J,4000) 16530
4000 FORMAT (' NOTE: FOR A LISTING OF MASTER SPECIES AND ELEMENTS,' 16540
1,/,8X,'ENTER <LIST>.',/) 16550
GO TO 1729 16560
4104 WRITE (J,4105) SUB(2) 16570
4105 FORMAT (/1X,A8,/,13X,'NELT',9X,'INDEX NUMBER ASSIGNED TO THE ELEME 16580
1NT. NUMBER',/,26X,'MUST BE BETWEEN 4 AND 30, INCLUSIVE.',/,1X, 16590
2'FORMAT (I2)') 16600
WRITE (J,4000) 16610
GO TO 1550 16620
4106 WRITE (J,4107) SUB(2) 16630
4107 FORMAT (/1X,A8,/,13X,'TGFW',9X,'GRAM FORMULA WEIGHT OF THE SPECIES 16640
1 USED TO',/,26X,'REPORT THE ANALYTICAL DATA. IF SOLUTION DATA' 16650
2,/,26X,'IS TO INCLUDE ALKALINITY, TGFW FOR THE ELEMENT',/,26X,'CAR 16660
3BON MUST BE THE EQUIVALENT WEIGHT OF THE',/,26X,'REPORTED ALKALINI 16670
4TY SPEC S.',/,1X,'FORMAT (F10.4)') 16680
WRITE (J,4000) 16690
GO TO 1570 16700
4110 WRITE (J,4111) SUB(3) 16710
4111 FORMAT (/1X,A8,/,13X,'I',12X,'THE INDEX NUMBER ASSIGNED TO THE AQU 16720
1EOUS',/,26X,'SPECIES. NUMBER 4 THROUGH 30 ARE RESERVED',/,26X, 16730
2'FOR MASTER SPECIES. 250 IS THE MAXIMUM INDEX',/,26X,'NUMBER FOR 16740
3AN AQUEOUS SPECIES.',/,1X,'FORMAT (I3)',/) 16750
GO TO 1590 16760
4112 WRITE (J,4113) SUB(3) 16770
4113 FORMAT (/1X,A8,/,13X,'SNAME',8X,'ALPHANUMERIC SPECIES NAME.',/,1X, 16780

```

```

1'FORMAT (A8)',//) 16790
GO TO 1660 16800
4114 WRITE (J,4115) SUB(3) 16810
4115 FORMAT (/1X,A8,/,13X,'NSP',10X,'THE TOTAL NUMBER OF MASTER SPECIES 16820
1 IN THE',/,26X,'ASSOCIATION REACTION THAT FORMS THIS SPECIES;',/,2 16830
26X,'DO NOT COUNT THE SPECIES ITSELF UNLESS THE',/,26X,'SPECIES IS 16840
3A MASTER SPECIES.',/,1X,'FORMAT (I3)',//) 16850
GO TO 1676 16860
4116 WRITE (J,4117) SUB(3) 16870
4117 FORMAT (/1X,A8,/,13X,'KFLAG',3X,'= 0, THE VAN'T HOFF EXPRESSION IS 16880
1 USED TO CALCU-',/,26X,'LATE TEMPERATURE DEPENDENCE OF THE ASSOCIA 16890
TION',/,26X,'CONSTANT FOR THIS SPECIES.',/,21X,'= 1, AN ANALYTICA 16900
3L EXPRESSION IS USED TO CALCU-',/,26X,'LATE TEMPERATURE DEPENDENCE 16910
4 OF THE ASSOCIATION',/,26X,'CONSTANT.',/,1X,'FORMAT (I1)',//) 16920
GO TO 1680 16930
4118 WRITE (J,4119) SUB(3) 16940
4119 FORMAT (/1X,A8,/,13X,'GFLAG',3X,'= 0, THE EXTENDED DEBYE-HUCKEL OR 16950
1 DAVIS EXPRES-',/,26X,'SION (ACCORDING TO IOPT(6)) IS USED TO CALC 16960
2ULATE',/,26X,'THE ACTIVITY COEFFICIENT FOR THIS SPECIES.',/,21X, 16970
3'= 1, THE WATEQ DEBYE-HUCKEL EXPRESSION IS USED',/,26X,'TO CALCULA 16980
4TE THE ACTIVITY COEFFICIENT OF THIS',/,26X,'SPECIES REGARDLESS OF 16990
5THE VALUE OF IOPT(6)',/,1X,'FORMAT (I1)',//) 17000
GO TO 1690 17010
4120 WRITE (J,4121) SUB(3) 17020
4121 FORMAT (/1X,A8,/,13X,'ZSP',10X,'THE CHARGE ON THIS AQUEOUS SPECIES 17030
1.',/,1X,'FORMAT (F10.3)',//) 17040
GO TO 1780 17050
4122 WRITE (J,4123) SUB(3) 17060
4123 FORMAT (/1X,A8,/,13X,'THSP',9X,'THE SUM OF THE FORMAL VALENCE OF T 17070
1HE REDOX',/,26X,'SPECIES IN THIS SPECIES. (E.G. FESO4 HAS',/,26X, 17080
2'A THSP = 2 + 6 = 8.)',/,1X,'FORMAT (F10.3)',//) 17090
GO TO 1790 17100
4124 WRITE (J,4125) SUB(3) 17110
4125 FORMAT (/1X,A8,/,13X,'DHA',10X,'THE EXTENDED DEBYE-HUCKEL A 0 TERM 17120
1.',/,1X,'FORMAT (F10.3)'//) 17130
GO TO 1880 17140
4126 WRITE (J,4127) SUB(3) 17150
4127 FORMAT (/1X,A8,/,13X,'ADHSP(1)',5X,'THE A(I) TERM FOR THE WATEQ DE 17160
1BYE-HUCKEL EXPRESSION.',/,1X,'FORMAT (F10.3)',//) 17170
GO TO 1892 17180
4128 WRITE (J,4129) SUB(3) 17190
4129 FORMAT (/1X,A8,/,13X,'ADHSP(2)',5X,'THE B(I) TERM FOR THE WATEQ DE 17200
1BYE-HUCKEL EXPRESSION.',/,1X,'FORMAT (F10.3)',//) 17210
GO TO 1892 17220
4130 WRITE (J,4131) SUB(3) 17230
4131 FORMAT (/1X,A8,/,13X,'ALKSP',8X,'THE ALKALINITY ASSIGNED TO THIS A 17240
1QUEOUS SPECIES.',/,1X,'FORMAT (F10.3)',//) 17250
GO TO 1740 17260
4132 WRITE (J,4133) SUB(3) 17270
4133 FORMAT (/1X,A8,/,13X,'LKTOSP',7X,'LOG (K) AT 25 DEGREES CELCIUS', 17280
1', WHERE',/) 17290
WRITE (J,4134) 17300
4134 FORMAT (24X,'LOG K = A1 + A2*T + A3/T + A4*LOG(T) + A5/T^2 ') 17310
WRITE (J,4135) 17320
4135 FORMAT (/1X,'FORMAT (F10.3)'//) 17330
GO TO 1960 17340

```

4136	WRITE (J,4137) SUB(3)	17350
4137	FORMAT (/1X,A8,/,13X,'DHSP',9X,'STANDARD ENTHALPY OF THE ASSOCIAT'	17360
	1,'ION REACTION AT',/,26X,'25 DEGREES CELCIUS (H(R), IN KCAL/'	17370
	2,'MOLE).',/,1X,'FORMAT (F10.3)',//)	17380
	GO TO 1970	17390
4138	WRITE (J,4139) SUB(3),K,K	17400
4139	FORMAT (/1X,A8,/,13X,'ASP(',I1,')',7X,'A',I1,' OF THE FOLLOWING',	17410
	1' EQUATION:',/)	17420
	WRITE (J,4134)	17430
	WRITE (J,4140)	17440
4140	FORMAT (/1X,'FORMAT (F12.5)')	17450
	GO TO 1987	17460
4142	WRITE (J,4143) SUB(3)	17470
4143	FORMAT (/1X,A8,/,13X,'LSP',10X,'INDEX NUMBER OF MASTER SPECIES.'	17480
	1,/,1X,'FORMAT (I3)')	17490
	WRITE (J,4000)	17500
	GO TO 1915	17510
4144	WRITE (J,4145) SUB(3)	17520
4145	FORMAT (/1X,A8,/,13X,'CSP',10X,'STOICHIOMETRIC COEFFICIENT OF',	17530
	1' MASTER SPECIES',/,26X,'IN THIS AQUEOUS SPECIES.',/,1X,'FORMAT'	17540
	2,' (F7.3)')	17550
	WRITE (J,4000)	17560
	GO TO 1915	17570
8000	WRITE (J,8001) SUB(I)	17580
8001	FORMAT (/1X,A8,/,13X,'MNAME',8X,'ALPHANUMERIC NAME OF MINERAL',/,	17590
	11X,'FORMAT (A8)',//)	17600
	GO TO 3102	17610
8002	WRITE (J,8003) SUB(I)	17620
8003	FORMAT (/1X,A8,/,13X,'NMINO',8X,'NUMBER OF DIFFERENT SPECIES IN TH	17630
	1E MINERAL',/,26X,'DISSOCIATION REACTION (INCLUDING H+, E-, AND',/	17640
	2,26X,'H2O). NMINO MUST BE LESS THAN OR EQUAL TO 10.',/,1X,	17650
	3'FORMAT (I2)',//)	17660
	GO TO 3110	17670
8004	WRITE (J,8005) SUB(I)	17680
8005	FORMAT (/1X,A8,/,13X,'THMIN',8X,'THE SUM OF THE VALENCES OF THE RE	17690
	1DOX SPECIES',/,26X,'IN THE MINERAL DISSOCIATION REACTION.',/,1X,	17700
	2'FORMAT (F10.3)',//)	17710
	GO TO 3120	17720
8006	WRITE (J,8007)SUB(I)	17730
8007	FORMAT (/1X,A8,/,13X,'LKTOM',8X,'LOG OF THE EQUILIBRIUM CONSTANT'	17740
	1,' AT 25 DEGREES',/,26X,'CELCIUS FOR THE REACTION',/,1X,'FORM AT'	17750
	2,' (F10.3)',//)	17760
	GO TO 3130	17770
8008	WRITE (J,8009) SUB(I)	17780
8009	FORMAT (/1X,A8,/,13X,'DHMIN',8X,'DELTA H(R) (KCAL/MOLE) FOR THE '	17790
	1,'VANT HOFF',/,26X,'EXPRESSION.',/,1X,'FORMAT (F10.3)'//)	17800
	GO TO 3140	17810
8010	WRITE (J,8011) SUB(I)	17820
8011	FORMAT (/1X,A8,/,13X,'MFLAG',3X,'= 0, THE VANT HOFF EXPRESSION IS'	17830
	1,' USED TO CALCU-',/,26X,'LATE THE TEMPERATURE DEPENDENCE OF THE',	17840
	2' EQUIL-',/,26X,'IBRIUM CONSTANT',/,21X,'= 1, THE ANALYTICAL EXP'	17850
	3,'RESSION IS USED TO CALCU-',/,26X,'LATE THE TEMPERATURE DEPENDENCE	17860
	4 OF THE EQUIL-',/,26X,'IBRIUM CONSTANT.',/,1X,'FORMAT (I1)'//)	17870
	GO TO 3150	17880
8012	WRITE (J,8013) SUB(I)	17890
8013	FORMAT (/1X,A8,/,13X,'SIMIN',8X,'SATURATION INDEX (LOG(ION ACTIVI'	17900

```

1, 'TY PRODUCT/K(SP)))',/,26X, 'DESIRED IN THE FINAL SOLUTION. SIMIN 17910
2 = 0.0',/,26X, 'WOULD PRODUCE EQUILIBRIUM WITH THE MINERAL',/,26X, 17920
3 'WHILE 1.0 WOULD PRODUCE A SOLUTION 10 TIMES',/,26X, 'SUPERSATURATE 17930
4D (SI = 1.0). THIS VARIABLE IS',/,26X, 'USEFUL IN SPECIFYING THE ' 17940
5, 'PARTIAL PRESSURE OF',/,26X, 'A GAS. THE HENRY(S) LAW CONSTANT ', 17950
6 'FOR THE GAS',/,26X, 'WOULD BE ENTERED USING THE VANT HOFF (LKTOM) 17960
7',/,26X, 'OR ANALYTICAL EXPRESSION (AMIN) AND THE LOG OF',/,25X, ' THE 17970
8 PARTIAL PRESSURE WOULD BE ENTERED FOR SIMIN.',/,1X, 'FORMAT (F10.3 17980
9)',/) 17990
GO TO 3167 18000
8014 WRITE (J,8015) SUB(I) 18010
8015 FORMAT (/1X,A8,/,13X,'LMIN',9X,'INDEX NUMBER OF SPECIES (NOT NECE' 18020
1, 'SSARILY MASTER',/,26X, 'SPECIES) IN THE DISSOCIATION REACTION ', 18030
2 'FOR',/,26X, 'THIS MINERAL',/,1X, 'FORMAT (I4)') 18040
WRITE (J,3800) 18050
GO TO 3195 18060
8016 WRITE (J,8017) SUB(I) 18070
8017 FORMAT (/1X,A8,/,13X,'CMIN',9X,'STOICHIOMETRIC COEFFICIENT OF ', 18080
1 'SPECIES IN',/,26X, 'DISSOCIATION REACTION.',/,1X, 'FORMAT (F11.3)') 18090
WRITE (J,3800) 18100
GO TO 3195 18110
8018 WRITE (J,8019) SUB(I),K,K 18120
8019 FORMAT (/1X,A8,/,9X,'AMIN(',I1,')',6X,'A',I1,1X,'OF THE FOLLOWING 18130
1 MINERAL DISSOCIATION',/,22X, 'REACTION:',/,22X, 'LOG K = A1 + ' 18140
2, ' A2*T + A3/T + A4*LOG(T) + A5/T^2 ',/,22X, 'WHERE T IS' 18150
3, ' IN DEGREES KELVIN',/,1X, 'FORMAT (F12.5)',/) 18160
GO TO 3207 18170
8020 WRITE (J,8021) SUB(6) 18180
8021 FORMAT (/1X,A8,/,13X,'XTEMP',8X,'TEMPERATURE IN DEGREES CELSIUS.' 18190
1,/,1X, 'FORMAT (F10.3)',/) 18200
GO TO 3307 18210
8022 WRITE (J,8023) SUB(7) 18220
8023 FORMAT (/1X,A8,/,13X,'XSTEP',8X,'THE FRACTION OF SOLUTION 1 TO BE' 18230
1,/,26X, 'MIXED WITH SOLUTION 2.') 18240
8030 WRITE (J,8029) 18250
8029 FORMAT (1X, 'FORMAT (F10.9)',/) 18260
GO TO 3607 18270
8024 WRITE (J,8025) SUB(7) 18280
8025 FORMAT (/1X,A8,/,13X,'XSTEP',8X,'THE VOLUME OF SOLUTION 2 TO', 18290
1 ' BE',/,26X, 'TITRATED INTO SOLUTION 1. XSTEP MUST',/,26X, 'HAVE' 18300
2, ' THE SAME UNITS AS VO.') 18310
GO TO 8030 18320
8026 WRITE (J,8027) SUB(7) 18330
8027 FORMAT (/1X,A8,/,13X,'XSTEP',8X,'THE MOLES OF REACTION TO BE ', 18340
1 'ADDED',/,26X, 'TO SOLUTION 1') 18350
GO TO 8030 18360
8028 WRITE (J,8031) SUB(7) 18370
8031 FORMAT (/1X,A8,/,13X,'XSTEP',8X,'THE TOTAL NUMBER OF MOLES OF' 18380
1,/,26X, 'REACTION TO BE ADDED IN NSTEPS',/,26X, 'STEPS. NSTEPS' 18390
2, ' REACTION SOLUTIONS',/,26X, 'WILL BE CALCULATED. THE I(TH) SOLU-' 18400
3,/,26X, 'TION WILL HAVE I*XSTEP/NSTEPS MOLES',/,26X, 'OF REACTION' 18410
4, ' ADDED TO SOLUTION 1.') 18420
GO TO 8030 18430
8032 WRITE (J,8033) SUB(8) 18440
8033 FORMAT (/1X,A8,/,13X,'LREAC',8X,'INDEX NUMBER OF ELEMENT FOR THE' 18450
1, ' REACTION.',/,26X, 'LREAC MUST BE BETWEEN 4 AND 30 INCLUSIVE.',/, 18460

```

226X,'IF LREAC IS GREATER THAN 30 THE PROGRAM',/,26X,'CONSIDERS',	18470
3' THIS CONSTITUENT TO BE H2 OR ',/,26X,'O2 AND ONLY USES CREAC',	18480
4' AND THMEAN',/,26X,'TO CHANGE THE OXIDATION STATE OF THE',/,26X	18490
5,'REACTION SOLUTION.',/,1X,'FORMAT (I4)')	18500
WRITE (J,3800)	18510
GO TO 5015	18520
8034 WRITE (J,8035) SUB(8)	18530
8035 FORMAT (/1X,A8,/,13X,'CREAC',8X,'STOICHIOMETRIC COEFFICIENT OF'	18540
1,' THE ELEMENT',/,26X,'IN THE REACTION',/,1X,'FORMAT (F8.3)')	18550
WRITE (J,3800)	18560
GO TO 5015	18570
8036 WRITE (J,8037) SUB(8)	18580
8037 FORMAT (/1X,A8,/,13X,'THMEAN',7X,'THE VALENCE OF THE ELEMENT IN'	18590
1,' THE REACTION.',/,26X,'AN ELEMENT MAY BE INCLUDED MORE THAN',	18600
2' ONCE IN',/,26X,'A REACTION TO ACCOMODATE DIFFERENT VALENCE',/,	18610
326X,'STATES OF THE ELEMENT.',/,1X,'FORMAT (F8.3)')	18620
WRITE (J,3800)	18630
GO TO 5015	18640
8038 WRITE (J,8039) SUB(9)	18650
8039 FORMAT (/1X,A8,/,13X,'LPOS',9X,'INDEX NUMBER OF AN ELEMENT WITH'	18660
1,' A CATION',/,26X,'MASTER SPECIES.',/,1X,'FORMAT (I5)')	18670
WRITE (J,3800)	18680
GO TO 6010	18690
8040 WRITE (J,8041) SUB(9)	18700
8041 FORMAT (/1X,A8,/,13X,'LNEG',9X,'INDEX NUMBER OF AN ELEMENT WITH'	18710
1,' AN ANION',/,26X,'MASTER SPECIES.',/,1X,'FORMAT (I5)')	18720
WRITE (J,3800)	18730
3800 FORMAT (1X,'NOTE: FOR A LISTING OF MASTER SPECIES AND ELEMENTS,'	18740
1,/,8X,'ENTER <LIST>.',/)	18750
GO TO 6020	18760
8042 WRITE (J,8043) SUB(10)	18770
8043 FORMAT (/1X,A8,/,13X,'SUNAME',7X,'ALPHANUMERIC NAME TO BE PRINT'	18780
1'ED TO IDENTIFY',/,26X,'THE SUM.',/,1X,'FORMAT (A8)')//)	18790
GO TO 7010	18800
8044 WRITE (J,8045) SUB(10)	18810
8045 FORMAT (/1X,A8,/,13X,'NSUM',9X,'THE NUMBER OF INDEX NUMBER TO BE'	18820
1,' READ.',/,26X,'NSUM<=50.',/,1X,'FORMAT (I2)')//)	18830
GO TO 7020	18840
8046 WRITE (J,8047) SUB (10)	18850
8047 FORMAT (/1X,A8,/,13X,'LSUM',9X,'INDEX NUMBERS OF SPECIES IN '	18860
1,'SUM.',/,1X,'FORMAT (I4)')//)	18870
GO TO 7037	18880
END	18890
SUBROUTINE READFILE	18900
C	18910
C THIS SUBROUTINE READS FROM LOGICAL FILE #11,	18920
C PRE-CONSTRUCTED MINERAL CARDS; AND FROM LOGICAL FILE #12,	18930
C THE THERMODYNAMIC DATA OF PHREEQE, THE NECESSARY INFORMATIONS	18940
C TO CONSTRUCT A MASTER SPECIES AND ELEMENTS TABLE.	18950
C	18960
PARAMETER (IU1=11,IU2=12)	18970
REAL ICOL3(0:30),NOUSE6	18980
CHARACTER *12 THMIN(100),LKTOM(100),DHMIN(100),SIMIN(100)	18990
1,CMIN(100,10),AMIN(100,5)	19000
CHARACTER *8 NONE,MNAME(100),SUB(11),ICOL1(0:30),ICOL4(0:30)	19010
1,SPEC(30),KEYWORD,NOUSE2	19020

	COMMON /IUNIT/ J	19030
	COMMON /PT/ MNAME,NMINO(100),THMIN,LKTOM,DHMIN,MFLAG(100)	19040
	1,SIMIN,LMIN(100,10),CMIN,AMIN,IMINERAL,SUB,ICOL1,ICOL2(0:30)	19050
	2,ICOL3,ICOL4,SPEC,OPV(30)	19060
	DATA NONE/'*****'/	19070
C		19080
	LREAD=0	19090
	J=1	19100
	DO 40 M=1,30	19110
40	ICOL2(M)=0	19120
	NUM=1	19130
C		19140
C	READ MINERAL CARDS.	19150
C	NOTE: MAXIMUM NUMBER OF MINERALS IS 100.	19160
C		19170
	OPEN (UNIT=IU1,FILE='RATEMOD>PITZ>INPUT>MINERALS.2.DATA',STATUS=	19180
	1'OLD')	19190
	DO 26 M=1,100	19200
	READ (IU1,27,END=7)MNAME(M),NMINO(M),THMIN(M),LKTOM(M),DHMIN(M)	19210
	1,MFLAG(M),SIMIN(M)	19220
27	FORMAT (A8,2X,I2,3X,3A10,5X,I1,9X,A10)	19230
	NO2=NMINO(M)	19240
	IF (NO2.LE.5) GO TO 29	19250
	NO2=5	19260
29	READ (IU1,28) (LMIN(M,MO),CMIN(M,MO),MO=NUM,NO2)	19270
28	FORMAT (5(I4,A11))	19280
	IF (NMINO(M).LE.5.OR.NUM.EQ.6) GO TO 30	19290
	NUM=6	19300
	NO2=NMINO(M)	19310
	GO TO 29	19320
30	IF (MFLAG(M).EQ.0) GO TO 26	19330
10	READ (IU1,31) (AMIN(M,MO),MO=1,5)	19340
31	FORMAT (5A12)	19350
26	CONTINUE	19360
	7 IMINERAL=M-1	19370
C		19380
C	READ THERMODYNAMIC DATA OF PHREEQE.	19390
C	NOTE: THE DATA BLOCKS OF PHREEQE DATA CAN BE ARRANGED	19400
C	IN ANY ORDER.	19410
C		19420
	CLOSE (UNIT=IU1)	19430
	OPEN (UNIT=IU2,FILE='RATEMOD>PITZ>LATEST>PHRQPITZ.DATA'	19440
	1,STATUS='OLD')	19450
48	READ (IU2,11,END=50) KEYWORD	19460
11	FORMAT (A8)	19470
C		19480
C	READ IS REPEATED UNTIL A KEYWORD IS FOUND.	19490
C		19500
	IF (KEYWORD.EQ.SUB(2))GO TO 42	19510
	IF (KEYWORD.EQ.SUB(3))GO TO 44	19520
	GO TO 48	19530
46	LREAD=LREAD+1	19540
	IF (LREAD.NE.2) GO TO 48	19550
	CLOSE (UNIT=IU2)	19560
	RETURN	19570
C		19580

C	READ SPECIES DATA BLOCK.	19590
C		19600
44	DO 80 JL=1,250	19610
	READ (IU2,51) NOUSE1	19620
	IF (NOUSE1.EQ.0)GO TO 46	19630
	READ (IU2,52)NOUSE2,NOUSE3,NOUSE6	19640
51	FORMAT (I3)	19650
52	FORMAT (A8,2X,I3,12X,F10.3,//)	19660
	IF (NOUSE3.GT.6)READ (22,53)	19670
53	FORMAT (1X)	19680
	IF (NOUSE1.GT.30)GO TO 80	19690
	SPEC(NOUSE1)=NOUSE2	19700
	OPV(NOUSE1)=NOUSE6	19710
80	CONTINUE	19720
	GO TO 46	19730
C		19740
C	READ ELEMENTS DATA BLOCK.	19750
C		19760
42	DO 8 JL=1,50	19770
	READ (IU2,9)ICOL1(0),ICOL2(0),ICOL3(0),ICOL4(0)	19780
9	FORMAT (A8,2X,I2,3X,F10.0,5X,A8)	19790
	IF (ICOL2(0).GT.30)GO TO 8	19800
	IF (ICOL2(0).NE.0)GO TO 100	19810
	JCOL=JL-1	19820
	GO TO 46	19830
100	ICOL1(ICOL2(0))=ICOL1(0)	19840
	ICOL2(ICOL2(0))=ICOL2(0)	19850
	ICOL3(ICOL2(0))=ICOL3(0)	19860
	ICOL4(ICOL2(0))=ICOL4(0)	19870
8	CONTINUE	19880
50	STOP	19890
C	*****	19900
	ENTRY LISTM	19910
C	*****	19920
C		19930
C	THIS SUBROUTINE FORMATS AND PRINTS A TABLE OF MASTER	19940
C	SPECIES AND ELEMENTS.	19950
C		19960
	DO 200 MO=1,3	19970
	ICOL1(MO)=NONE	19980
	ICOL2(MO)=MO	19990
	ICOL3(MO)=0.0	20000
	ICOL4(MO)=NONE	20010
	OPV(MO)=0.0	20020
200	CONTINUE	20030
	ICOL3(3)=18.0152	20040
	SPEC(1)='H+ '	20050
	SPEC(2)='E- '	20060
	SPEC(3)='H2O '	20070
	WRITE (J,201)	20080
201	FORMAT (/34X,'INPUT FORMULA',/,34X,'CORRESPONDING',5X,'MASTER',/ 1,1X,'ELEMENTS',5X,'#',9X,'GFW',10X,'TO GFW',9X,'SPECIES',6X,'OPV' 2,/,1X,8('-'),4X,3('-'),4X,10('-'),4X,14('-'),4X,8('-'),4X,5('-') 3,/))	20090 20100 20110 20120
	DO 203 MO=1,30	20130
	IF (ICOL2(MO).EQ.0)GO TO 203	20140

WRITE(J,202)ICOL1(MO),ICOL2(MO),ICOL3(MO),ICOL4(MO),SPEC(MO)	20150
1,OPV(MO)	20160
203 CONTINUE	20170
202 FORMAT (1X,A8,4X,I2,5X,F10.4,8X,A8,6X,A8,4X,SP,F4.1,SS)	20180
RETURN	20190
END	20200
C	20210
SUBROUTINE REF (ILINE,JOPT,JOPEN)	20220
C	20230
C THIS SUBROUTINE READS CERTAIN NUMBER OF LINES--ACCORDING	20240
C TO THE VALUE OF ILINE--FROM THE REFERENCE FILE, PRINTS	20250
C IT(THEM) ON THE SCREEN, AND, DEPENDING ON THE VALUE OF	20260
C IOPEN, ASKS A QUESTION CONCERNING THE FATE OF THE LINE(S).	20270
C	20280
C JOPT: 0 = NO PRINT OF LINES	20290
C 1 = PRINT THE LINES	20300
C JOPEN: 0 = NO QUESTION ASKED	20310
C > 0 = QUESTION ASKED	20320
C	20330
COMMON /IUNIT/ J	20340
COMMON /REFF/ JOPTION,IOPEN,JLINE	20350
COMMON /REFF1/ I	20360
CHARACTER *80 JLINE(3),LINE	20370
C	20380
DO 310 II=1,ILINE	20390
READ (13,10,END=99) JLINE(II)	20400
10 FORMAT (A80)	20410
IF (JLINE(II)(1:10).NE.' ')	20420
IF (II.NE.1) GO TO 311	20430
JOPTION=0	20440
RETURN	20450
311 IF (JOPT.EQ.0) GO TO 310	20460
WRITE (J,300) JLINE(II)	20470
300 FORMAT (A80)	20480
310 CONTINUE	20490
IF (JOPEN.EQ.0)RETURN	20500
402 GO TO (303,400,500,600), JOPEN	20510
303 WRITE (J,301)	20520
301 FORMAT ('ENTER OPTION. (1=KEEP, 2=MODIFY, 3=REPLACE)')	20530
305 READ (J,300) LINE	20540
READ (LINE,*,ERR=402) JOPTION	20550
IF (JOPTION.LE.0.OR.JOPTION.GT.3)GO TO 402	20560
IF ((JOPTION.EQ.3).AND.(JOPEN.EQ.2.OR.JOPEN.EQ.4))GO TO 402	20570
RETURN	20580
99 IOPEN=0	20590
RETURN	20600
400 WRITE (J,401)	20610
401 FORMAT ('ENTER OPTION. (1=KEEP, 2=ELIMINATE)')	20620
GO TO 305	20630
500 WRITE (J,501)	20640
501 FORMAT ('ENTER OPTION. (1=KEEP, 2=MODIFY, 3=ELIMINATE)')	20650
GO TO 305	20660
600 WRITE (J,601)	20670
601 FORMAT ('ENTER OPTION. (1=KEEP, 2=REPLACE)')	20680
GO TO 305	20690
END	20700

C		20710
	SUBROUTINE SPECL(I)	20720
C		20730
C	THIS SUBROUTINE CHECKS FOR SPECIAL CHARACTERS 'EXIT' AND 'LIST'	20740
C		20750
	COMMON /EE/ EXIT,LIST,LINE,JFLAG	20760
	LOGICAL EXIT,LIST	20770
	CHARACTER *80 LINE	20780
	CHARACTER *4 QUIT(6)	20790
	DATA QUIT /'EXIT','exit','STOP','stop','QUIT','quit'/	20800
C		20810
	EXIT=.FALSE.	20820
	LIST=.FALSE.	20830
	IF (I.EQ.0) RETURN	20840
	IF ((LINE(1:4).EQ.'LIST'.OR.LINE(1:4).EQ.'list').AND.(I.EQ.2.OR.	20850
	11.EQ.3)) LIST=.TRUE.	20860
	DO 10 N=1,6	20870
	IF (LINE(1:4).EQ.QUIT(N).AND.(I.EQ.1.OR.I.EQ.3)) EXIT=.TRUE.	20880
10	CONTINUE	20890
	RETURN	20900
	END	20910
C		20920
	LOGICAL FUNCTION OK()	20930
C		20940
C	THIS SUBROUTINE ASKS WHETHER IT IS O.K.	20950
C		20960
	COMMON /IUNIT/ J	20970
	COMMON /PN/ NFLAG	20980
	LOGICAL ANSWER,YN	20990
C		21000
	WRITE (J,10)	21010
10	FORMAT ('O.K.?')	21020
	NFLAG=0	21030
	OK=.TRUE.	21040
	ANSWER=YN()	21050
	IF (ANSWER) RETURN	21060
	NFLAG=1	21070
	OK=.FALSE.	21080
	RETURN	21090
	END	21100
C		21110
C	THIS SUBROUTINE READS VALUES FROM THE TERMINAL AND CHECKS TO MAKE	21120
C	SURE THAT IT IS PROPER	21130
C		21140
	SUBROUTINE QUESTA (*,J1,CHAR1,L1,STRING,F1,K1,M1)	21150
	COMMON /REFF/ JOPTION,IOPEN,JLINE	21160
	COMMON /PN/ NFLAG	21170
	COMMON /EE/ EXIT,LIST,LINE,JFLAG	21180
	COMMON /IUNIT/ N	21190
	CHARACTER *(*) STRING	21200
	CHARACTER *80 LINE,JLINE(3)	21210
	CHARACTER *72 QUESTION	21220
	CHARACTER *40 FMT	21230
	CHARACTER *16 CHAR,CHAR1,CHAR2	21240
	CHARACTER *4 F,F1,F2,FF	21250
	CHARACTER *2 FO(16)	21260

	LOGICAL EXIT,LIST,NONB	21270
	DATA FO/'1 ','2 ','3 ','4 ','5 ','6 ','7 ','8 ','9 ','10','11'	21280
	1,'12','13','14','15','16' /	21290
	DATA QUESTION /'KEEP THE OLD VALUE? (HIT <RETURN> IF YOU DO; IF NO	21300
	1T, REENTER THE DATA)'/	21310
C		21320
C	K=0 ==> STRING	21330
C	K=1 ==> REAL	21340
C	K=2 ==> INTEGER	21350
C		21360
	FF=F1	21370
	IF (F1.EQ.'A80') FF='/A80'	21380
	F1=FF	21390
	J=J1	21400
	CHAR=CHAR1	21410
	F=F1	21420
	K=K1	21430
	L=L1	21440
	M=M1	21450
	GO TO 100	21460
C	*****	21470
	ENTRY QUESTB (*,J2,CHAR2,L2,IVAL,F2,IMIN,IMAX,M2)	21480
C	*****	21490
	K=2	21500
	J=J2	21510
	CHAR=CHAR2	21520
	F=F2	21530
	L=L2	21540
	M=M2	21550
C		21560
100	IF (J.GT.0.AND.J.LT.10) I 1	21570
	IF (J.GE.10.AND.J.LT.100) I=2	21580
	IF (J.GE.100) I=3	21590
	IF (NFLAG.GT.0.OR.JOPTION.EQ.2) GO TO 30	21600
	IF (J.GT.0) GO TO 10	21610
	FMT='(/,'INPUT ','A'//FO(L)///)'	21620
	WRITE (N,FMT) CHAR	21630
	GO TO 300	21640
10	FMT='(/,'INPUT ','A'//FO(L)///','('','I'//FO(I)///','')')'	21650
	WRITE (N,FMT) CHAR,J	21660
	GO TO 300	21670
30	IF (J.GT.0) GO TO 40	21680
	FMT='(/,'OLD ','A'//FO(L)///','='','//F//','/A72)'	21690
	IF (K.EQ.2) WRITE (N,FMT) CHAR,IVAL,QUESTION	21700
	IF (K.NE.2) WRITE (N,FMT) CHAR,STRING,QUESTION	21710
	GO TO 300	21720
40	FMT='(/,'OLD ','A'//FO(L)///','('','I'//FO(I)///','')='','//F//','	21730
	1/,A72)'	21740
	IF (K.EQ.2) WRITE (N,FMT) CHAR,J,IVAL,QUESTION	21750
	IF (K.NE.2) WRITE (N,FMT) CHAR,J,STRING,QUESTION	21760
C		21770
C	READ RESPONSE	21780
C		21790
300	JFLAG=0	21800
	READ (N,105) LINE	21810
	105 FORMAT(A80)	21820

C		21830
C	JFLAG=0 ---> PROPER ENTRY	21840
C	JFLAG=1 ---> EXPLANATION IS ASKED	21850
C	JFLAG=2 ---> A NULL ANSWER	21860
C	JFLAG=3 ---> ANSWER OF WRONG TYPE	21870
C	JFLAG=4 ---> SPECIAL CHARACTERS	21880
C		21890
	IF (LINE(1:1).NE.'') GO TO 110	21900
	JFLAG=2	21910
	IF (K.EQ.2) IVALO=IVAL	21920
110	IF (LINE(1:1).EQ.'?') JFLAG=1	21930
	CALL SPECL (M)	21940
	IF (EXIT.OR.LIST) JFLAG=4	21950
	IF (JFLAG.NE.0) GO TO 200	21960
	GO TO (140,150,120), K+1	21970
140	STRING=LINE	21980
	RETURN	21990
120	READ (LINE,*,ERR=130) IVALO	22000
	GO TO 170	22010
150	READ (LINE,*,ERR=130) G	22020
C		22030
C	MAKE SURE THAT DECIMAL POINT IS PRESENT	22040
C		22050
	NONB=.FALSE.	22060
	DO 155 I2=1,80	22070
	IF (LINE(I2:I2).NE.' ') NONB=.TRUE.	22080
	IF (NONB.AND.LINE(I2:I2).EQ.' ') GO TO 130	22090
	IF (LINE(I2:I2).EQ.'.') GO TO 156	22100
155	CONTINUE	22110
	GO TO 130	22120
156	STRING=LINE	22130
	GO TO 200	22140
C		22150
130	JFLAG=3	22160
200	IF (JFLAG.EQ.2.AND.(NFLAG.NE.0.OR.JOPTION.EQ.2)) GO TO 170	22170
	IF (JFLAG.EQ.4) RETURN	22180
	IF (JFLAG.NE.0) RETURN 1	22190
170	IF (K.NE.2) RETURN	22200
	IF (IMAX.GT.IMIN) GO TO 180	22210
	GO TO 190	22220
180	IF (IVALO.GT.IMAX.OR.IVALO.LT.IMIN) RETURN 1	22230
190	IVAL=IVALO	22240
	RETURN	22250
	END	22260
C		22270
	LOGICAL FUNCTION YN()	22280
C		22290
C	THIS FUNCTION CHECKS FOR YES OR NO ANSWER	22300
C		22310
	CHARACTER *1 YESNO	22320
	COMMON /IUNIT/ J	22330
C		22340
5	READ (J,10) YESNO	22350
10	FORMAT (A1)	22360
	YN=.FALSE.	22370
	IF (YESNO.EQ.'Y'.OR.YESNO.EQ.'y') GO TO 20	22380

IF (YESNO.EQ.'N'.OR.YESNO.EQ.'n') RETURN	22390
WRITE (J,30)	22400
30 FORMAT ('PLEASE ANSWER ''Y'' OR ''N''.')	22410
GO TO 5	22420
20 YN=.TRUE.	22430
RETURN	22440
END	22450
C	22460
SUBROUTINE OPEN(ETIME)	22470
C	22480
C THIS SUBROUTINE GIVES THE USER A CHOICE OF CHOOSING A	22490
C REFERENCE FILE OR NOT TO USE IT AT ALL. IF THE	22500
C REFERENCE FILE IS DESIRED, IT OPENS THE FILE AND	22510
C POSITION THE POINTER SO THAT THE PROGRAM CAN USE IT;	22520
C IF REFERENCE FILE IS NOT WANTED, IT SETS IOPEN TO ZERO	22530
C AND RETURN TO 'MAIN.	22540
C	22550
COMMON /REFF/ JOPTION,IOPEN,JLINE	22560
COMMON /IUNIT/ J	22570
CHARACTER *36 TREAD	22580
CHARACTER *80 LINE,JLINE(3)	22590
SAVE TREAD,KOPEN	22600
LOGICAL YN	22610
C	22620
ETIME=ETIME+1	22630
IF (ETIME.EQ.1.OR.KOPEN.EQ.0)GO TO 100	22640
120 WRITE (J,110)	22650
110 FORMAT (/, 'DO YOU WISH TO USE THE SAME REFERENCE FILE?')	22660
IF (.NOT.YN()) GO TO 100	22670
CLOSE (UNIT=13)	22680
GO TO 25	22690
100 WRITE (J,10)	22700
10 FORMAT (/, 'ENTER REFERENCE FILE NAME: (HIT <CR> TO OMIT)')	22710
READ (J,20) TREAD	22720
20 FORMAT (A36)	22730
IF (TREAD(1:1).NE.'')GO TO 25	22740
IOPEN=0	22750
GO TO 30	22760
25 OPEN (UNIT=13,FILE=TREAD,STATUS='OLD')	22770
IOPEN=1	22780
30 KOPEN=IOPEN	22790
RETURN	22800
END	22810

Attachment F. -- Listing of the file MINERALS.2.DATA read by the
interactive input program PITZINPT

ANHYDRIT	2		6.00	-4.362		1			
4	1.000	16		1.000					
422.950	0.0			-18431.	-147.708				
ARAGONIT	2		4.00	-8.220		1			
15 1.0		4	1.0						
-171.8607	-.077993			2903.293	71.595				
ARCANITE	2		6.00	-1.776		1			
7	2.000	16		1.000					
2.823	0.0			-1371.2					
BISCHOFI	3		0.00	4.455		1			
5	1.000	14		2.000	3		6.000		
3.524	0.0			277.6					
BLOEDITE	4		12.00	-2.347		0			
6	2.000	5		1.000	16		2.000	3	4.000
BRUCITE	2		0.00	-10.884.85		0			
5	1.000	31		2.000					
BURKEITE	3		16.00	-.772		0			
6	6.000	15		1.000	16		2.000		
CALCITE	2		4.00	-8.406		1			
15 1.0		4	1.0						
-171.8329	-.077993			2839.319	71.595				
CARNALLI	4		0.00	4.330		0			
7	1.000	5		1.000	14		3.000	3	6.000
DOLOMITE	3		8.00	-17.083	-9.436	0			
4 1.0		5	1.0		15 2.0				
EPSOMITE	3		6.00	-1.881		1			
5	1.000	16		1.000	3		7.000		
1.718	0.0			-1073.					
GAYLUSSI	4		8.00	-9.421		0			
4	1.000	6		2.000	15		2.000	3	5.000
GLASERIT	3		24.00	-3.803		0			
6	1.000	7		3.000	16		2.000		
GLAUBERI	3		12.00	-5.245		0			
6	2.000	4		1.000	16		2.000		
GYP SUM	3		6.00	-4.581		1			
4 1.0		16	1.0		3 2.0				
90.318	0.0			-4213.	-32.641				
HALITE	2		0.00	1.570		1			
6	1.000	14		1.000					
-713.4616	-.1201241			37302.21	262.4583		-2106915.		
HEXAHYDR	3		6.00	-1.635		1			
5	1.000	16		1.000	3		6.000		
-62.666	0.0			1828.	22.187				
KAINITE	5		6.00	-0.193		0			
7	1.000	5		1.000	14		1.000	16	1.000 3 3.000
KALICINI	3		4.00	-10.058		0			
7	1.000	1		1.000	15		1.000		
KIESERIT	3		6.00	-0.123		0			
5	1.000	16		1.000	3		1.000		
LABILE S	4		18.00	-5.672		0			
6	4.000	4		1.000	16		3.000	3	2.000
LEONHARD	3		6.00	-0.887		0			
5	1.000	16		1.000	3		4.000		
LEONITE	4		12.00	-3.979		0			

7	2.000	5	1.000	16	2.000	3	4.000		
MAGNESIT	2		4.00	-7.834	-6.169	0			
5	1.000	15	1.000						
MIRABILI	3		6.00	-1.214		1			
6	2.000	16	1.000	3	10.000				
-3862.234	-1.19856		93713.54	1577.756					
MISENITE	3		42.00	-10.806		0			
7	8.000	1	6.000	16	7.000				
NAHCOLIT	3		4.00	-10.742		0			
6	1.000	1	1.000	15	1.000				
NATRON	3		4.00	-0.825		0			
6	2.000	15	1.000	3	10.000				
NESQUEHO	3		4.00	-5.167		0			
5	1.000	15	1.000	3	3.000				
PCO2	1		4.0	-1.468	-4.776	1			
35 1.0									
108.3865	0.01985076		-6919.53	-40.45154	669365.0				
PENTAHYD	3		6.00	-1.285		0			
5	1.000	16	1.000	3	5.000				
PIRSSONI	4		8.00	-9.234		0			
6	2.000	4	1.000	15	2.000	3	2.000		
POLYHALI	5		24.00	-13.744		0			
7	2.000	5	1.000	4	2.000	16	4.000	3	2.000
PORTLAND	2		0.00	-5.190		0			
4	1.000	31	2.000						
SCHOENIT	4		12.00	-4.328		0			
7	2.000	5	1.000	16	2.000	3	6.000		
SYLVITE	2		0.00	.900		1			
7	1.000	14	1.000						
3.984	0.0		-919.55						
SYNGENIT	4		12.00	-7.448		0			
7	2.000	4	1.000	16	2.000	3	1.000		
TRONA	4		8.00	-11.384		0			
6	3.000	1	1.000	15	2.000	3	2.000		

**Attachment G. -- Example of use of the interactive input program
PITZINPT to construct the input file for
test problem 3**

ENTER OUTPUT FILE NAME
TEST.3

ENTER REFERENCE FILE NAME: (HIT <CR> TO OMIT)

INPUT THE TITLE
Test Problem 3: Find the anhydrite-gypsum phase boundary with NaCl addition.
Test Problem 3: Find the anhydrite-gypsum phase boundary with NaCl addition.
O.K.?
Y

INPUT IOPT(1)
0

INPUT IOPT(2)
0

INPUT IOPT(3)
6

INPUT IOPT(4)
0

INPUT IOPT(7)
0

INPUT IOPT(8)
0

INPUT IOPT(9)
0

INPUT IOPT(10)
0

INPUT NCOMPS
2

0060020000 0 2 0.0
O.K.?
Y

KEYWORD DATA BLOCKS

ENTER KEYWORD.
SOLUTION 1

SOLUTION

INPUT HEAD

*

Pure water

O.K.?

Y

0 0 0 7.0 4.0 25.0 1.0

O.K.?

Y

KEYWORD DATA BLOCKS

ENTER KEYWORD.

MINERALS

MINERALS

PRE-CONSTRUCTED MINERAL DATA ARE AVAILABLE.

DO YOU WISH TO HAVE ANY OF THEM?

Y

1 ANHYDRIT	2 ARAGONIT	3 ARCANITE	4 BISCHOFI	5 BLOEDITE
6 BRUCITE	7 BURKEITE	8 CALCITE	9 CARNALLI	10 DOLOMITE
11 EPSOMITE	12 GAYLUSSI	13 GLASERIT	14 GLAUBERI	15 GYPSUM
16 HALITE	17 HEXAHYDR	18 KAINITE	19 KALICINI	20 KIESERIT
21 LABILE S	22 LEONHARD	23 LEONITE	24 MAGNESIT	25 MIRABILI
26 MISENITE	27 NAHCOLIT	28 NATRON	29 NESQUEHO	30 PCO2
31 PENTAHYD	32 PIRSSONI	33 POLYHALI	34 PORTLAND	35 SCHOENIT
36 SYLVITE	37 SYNGENIT	38 TRONA		

ENTER THE INDEX NUMBER OF MINERAL. (TYPE <STOP> TO EXIT)

1

ANHYDRIT 2 6.00 -4.362 1

4 1.000 16 1.000
422.950 0.0 -18431. -147.708

O.K.?

Y

ENTER THE INDEX NUMBER OF MINERAL. (TYPE <STOP> TO EXIT)

15

GYPSUM 3 6.00 -4.581 1
 4 1.0 16 1.0 3 2.0
90.318 0.0 -4213. -32.641
O.K.?
Y

ENTER THE INDEX NUMBER OF MINERAL. (TYPE <STOP> TO EXIT)
STOP

MORE MINERALS TO BE TYPED IN FROM THE TERMINAL?
N

KEYWORD DATA BLOCKS

ENTER KEYWORD.
REACTION

REACTION

INPUT LREAC(1)
6

INPUT CREAC(1)
1.

INPUT THMEAN(1)
0.

INPUT LREAC(2)
14

INPUT CREAC(2)
1.

INPUT THMEAN(2)
0.

 61. 0. 141. 0.
O.K.?
Y

KEYWORD DATA BLOCKS
